



 Cite this: *RSC Adv.*, 2022, 12, 17499

 Received 26th April 2022
 Accepted 6th June 2022

DOI: 10.1039/d2ra02656b

rsc.li/rsc-advances

Visible-light-induced one-pot synthesis of sulfonic esters *via* multicomponent reaction of arylazo sulfones and alcohols†

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Sulfonic ester is a chemical structure common to many organic molecules, including biologically active compounds. Herein, a visible-light-induced synthetic method to prepare aryl sulfonic ester from arylazo sulfones, DABSO (DABCO·(SO₂)₂), and alcohols in the presence of CuI as a coupling catalyst and HCl as an additive to yield sulfonic esters *via* multicomponent reaction. This synthetic method afforded a wide range of sulfonic esters with high yields under mild conditions.

Introduction

Organic compounds containing sulfonyl groups such as sulfones, sulfonamide, sulfonic esters, and sulfonyl hydrazine have been important molecules for many organic chemistry and material sciences.¹ The sulfonic ester motif in particular is featured as a key structure in many organic molecules.² The sulfonic ester moieties, including mesylate, camyslate, tosylate, and besylate, are frequently observed in biologically active compounds for agrochemicals and pharmaceuticals.³ For example, compounds containing a sulfonic ester moiety have been reported as effective anticancer agents, antimicrotubule agents, MAO-A inhibitor, and phosphor-STAT3 inhibitor (Fig. 1).⁴

Several synthetic protocols have been developed over the years for the preparation of sulfonic esters due to their importance. One common synthetic strategy for sulfonyl esters relied on the esterification of sulfonyl chlorides in the presence of a base.⁵ However, this procedure has drawbacks such as hazardous application of sulfonyl chlorides, limited tolerance of functional groups, poor stability, and challenging storage and handling protocols.⁶

Several other methods using different types of starting materials, including sulfonic acid,⁷ thiophenols,⁸ and dimethyl and diethyl sulfates,⁹ have also been reported in various studies for the synthesis of sulfonyl esters. In addition, a copper-

catalyzed formation of pentafluoroaryl-containing sulfonyl esters using boronic acids was developed.¹⁰ However, the methods have some limitations including harsh reaction conditions and use of expensive catalysts or reagents. Recently, reaction of aryl diazonium salts for the synthesis of pentafluoroaryl-bearing sulfonyl esters was carried out.¹¹ The reaction using aryl diazonium salts required harsh reaction conditions (Scheme 1).

Besides, in the synthesis of compounds containing sulfonic esters, several preparation procedures utilize sulfur dioxide (SO₂).¹² SO₂ is a crucial intermediate in organic synthesis with advantages including its prevalence in nature, affordable cost, and ease of transformation to the sulfonyl structural moiety.¹³ Therefore, SO₂ has been extensively used as a sulfonyl precursor in organic chemistry for several decades.¹⁴ However, less desirable attributes of SO₂ such as its toxicity and gaseous nature prevent its wide usage for organic reactions. Thus, SO₂

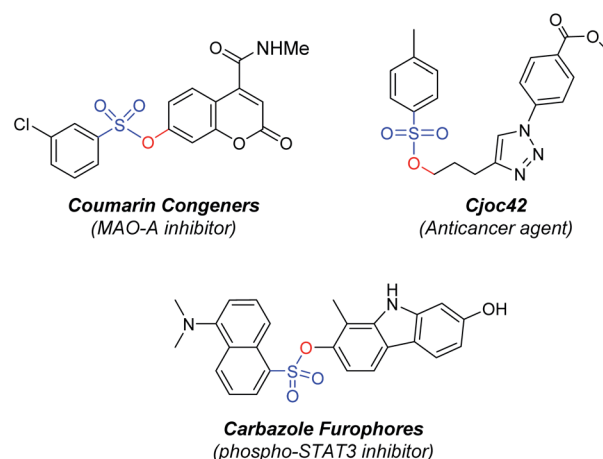


Fig. 1 Bioactive compounds containing sulfonic esters.

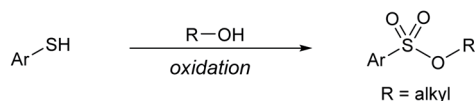
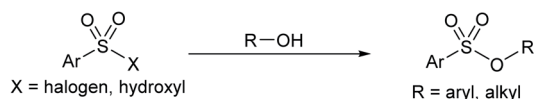
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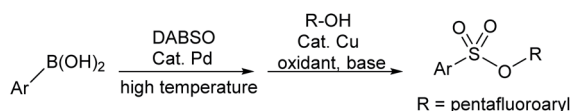
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 † Electronic supplementary information (ESI) available. See <https://doi.org/10.1039/d2ra02656b>


a) Traditional strategies from substrates bearing sulfur moiety



b) Synthetic strategy from boronic acids



This work:



Scheme 1 Synthetic procedures to give sulfonic esters.

surrogates have been developed as alternative options.¹⁵ Particularly, DABCO·(SO₂)₂ (DABSO) has recently been employed as one of the more useful SO₂ surrogates, allowing valuable advancement in the preparation of sulfonyl derivatives.¹⁶

The photochemical reaction strategy has recently been developed, which employs light power as a sustainable energy source because light enables activation of organic molecules. Visible-light-induced reactions could especially provide pathways to create free radical species, which are able to selectively generate C–C as well as C–X bonds.¹⁷ However, numerous visible-light-induced reactions have employed expensive and toxic photoredox catalysts such as ruthenium, iridium complexes, or organic dyes. Thus, it would be tremendously valuable to develop a photochemical reaction that does not require photoredox catalysts.

Arylazo sulfones have emerged as useful molecules for several organic synthetic processes. Arylazo sulfones were reported to be easily and efficiently produced, and readily divide to give aryl radical and sulfonyl methyl radical under light irradiation *via* homolytic cleavage of the N–S bond.¹⁸ So, they are employed for synthesis of valuable chemical structures such as aromatic amides,¹⁹ allylarenes,²⁰ arylboronates,²¹ triaryl-ethylenes,²² aryl stannanes,²³ aryl selenides and aryl tellurides,²⁴ aryl phosphonates,²⁵ as well as for non-catalytic arylation of heterocycles and inactivated arenes.²⁶

With the beneficial application of sulfonate ester moiety, discovering a novel and efficient synthetic procedure for sulfonyl esters is highly valuable. Herein, we report a novel photocatalyst-free visible-light-induced synthesis of sulfonate esters from arylazo sulfones under mild reaction conditions.

Result and discussion

In this study, we assumed that arylazo sulfones could yield aryl radical intermediates under irradiation by visible light.

Subsequent treatment with sulfur dioxide surrogate could then provide aryl sulfonyl radicals, which would be converted to sulfonic esters *via* reaction with alcohols.

To verify the feasibility of our concept, novel reaction conditions consisting of sulfonylating agents and alcohols were investigated for the synthesis of sulfonic esters. The initial reaction system using (*E*)-1-(4-methoxyphenyl)-2-(methylsulfonyl) diazene **1a** as starting material, DABSO as sulfur dioxide surrogate, MeOH **2a**, catalyst, and HCl as acid additive was designed, and the reaction was carried out under the radiation of blue LED light at room temperature for 4 hours.

First, various catalysts able to facilitate sulfonic ester generation were examined. When ZnCl₂ or SnCl₂ were used for the reaction, the target product was not observed (Table 1, entries 1 and 2). ZrCl₄, FeCl₃, and MnCl₂ provided the desired sulfonic ester product with low yields (21–25%) (Table 1, entries 3–5). TiCl₄ could facilitate the reaction for sulfonyl ester with an improved yield (62%) (Table 1, entry 6), while utilization of CuCl₂ provided the target product at good yield (74%) (Table 1,

Table 1 Screening of reaction conditions for synthesis of sulfonic esters^a

Entry	Catalyst	SO ₂ source agent	Additive	Yield ^b (%)
1	ZnCl ₂	DABSO	HCl	Trace
2	SnCl ₂	DABSO	HCl	Trace
3	ZrCl ₄	DABSO	HCl	21
4	FeCl ₃	DABSO	HCl	23
5	MnCl ₂	DABSO	HCl	25
6	TiCl ₄	DABSO	HCl	62
7	CuCl ₂	DABSO	HCl	74
8	CuBr	DABSO	HCl	Trace
9	CuBr ₂	DABSO	HCl	71
10	Cu(OAc) ₂	DABSO	HCl	75
11	Cu(acac) ₂	DABSO	HCl	72
12	CuI	DABSO	HCl	85
13	CuI	Na ₂ S ₂ O ₅	HCl	Trace
14	CuI	K ₂ S ₂ O ₅	HCl	Trace
15	CuI	DABSO	CF ₃ COOH	70
16	CuI	DABSO	NaHCO ₃	Trace
17	CuI	DABSO	K ₂ CO ₃	Trace
18	CuI	DABSO	DIEA	35
19	CuI	DABSO	DBU	37
20	CuI	DABSO	AcOH	45
21	CuI	DABSO	CH ₂ BrCOOH	47
22	CuI	DABSO	CF ₃ SO ₂ H	79
23 ^c	CuI	DABSO	HCl	82
24 ^d	CuI	DABSO	HCl	37
25 ^e	CuI	DABSO	HCl	75

^a Reaction conditions: compound **1** (1.0 mmol), SO₂ source agent (1.0 mmol), catalyst (0.1 mmol), MeOH (20 mmol), additive (1.0 mmol), CH₂Cl₂ (2 mL), 4 h, irradiation by 5 W blue LEDs for 4 h. ^b Isolated yield after purification by flash column chromatography. ^c Irradiation by 5 W white LEDs. ^d Irradiation by 5 W green LEDs. ^e Irradiation by 20 W white CFL.



entry 7). The results suggested that copper salts were effective for generating sulfonyl ester. Thus, investigations of several other copper salts were carried out. Reactions using CuBr_2 , CuCl_2 , $\text{Cu}(\text{OAc})_2$, and $\text{Cu}(\text{acac})_2$ provided the target product with good yields (71–75%) (Table 1, entries 9–12) while employment of CuBr did not yield the desired sulfonyl ester compound. When CuI was used as a catalyst for the reaction, the target sulfonyl ester was obtained with 85% yield. So, CuI was chosen as catalyst for generating sulfonyl ester in this reaction.

Next, several sulfur dioxide surrogates as SO_2 source agents were examined to optimize the synthesis of sulfonyl ester (Table 1, entries 12–14). Employment of $\text{Na}_2\text{S}_2\text{O}_5$ or $\text{K}_2\text{S}_2\text{O}_5$ did not provide the target sulfonyl esters (Table 1, entries 13 and 14). However, DABSO, a $\text{DABCO} \cdot (\text{SO}_2)_2$ complex, was successfully used for the reaction, and the target sulfonyl ester was synthesized with high yield (Table 1, entry 12).

Additives were an important factor affecting the reaction yield. Thus, different additives were tested. In reactions using inorganic bases such as NaHCO_3 and K_2CO_3 , no desired sulfonyl ester product was prepared. Utilization of organic bases such as DIEA and DBU afforded the corresponding products in low yield (35–37%) (Table 1, entries 18 and 19). In contrast, when acids such as AcOH and CH_2BrCOOH were used as additives, sulfonyl ester was synthesized with improved conversion yields (45–47%) (Table 1, entries 20–21). Additionally, in reactions employing $\text{CF}_3\text{SO}_2\text{H}$ and HCl as additives, the target products were prepared with high yields (79 and 85%) (Table 1, entries 22 and 12).

The influence of light source on reaction efficiency was examined using several types of light sources. The replacement of blue LEDs by white LEDs, green LEDs, or white CFLs did not give an enhanced yield (Table 1, entries 23–25). The reaction of arylazo sulfone **1a**, DABSO, and methanol **2a** under 5 W white LEDs affords the desired product **3a** in 82% yield which was lower than that of the reaction under 5 W blue LEDs with 85%. Therefore, using blue LEDs is suitable for this reaction. Thus, blue LEDs were indicated as a suitable light source for this reaction.

Solvent effect on the reaction was evaluated to find an optimal reaction protocol (Table 2). Employment of DMF, 1,4-dioxane did not yield the target sulfonyl ester product. In the reaction in THF, toluene, and MeCN, the desired products were prepared in less than 40% yield. When DCE was used as reaction solvent, the sulfonyl ester product was generated in increased yield. Especially, CH_2Cl_2 solvent showed a better reaction performance with high yield, suggesting that CH_2Cl_2 was a suitable solvent for sulfonyl ester synthesis.

Next, various amounts of reagents were examined to evaluate their efficiency for synthesis of sulfonyl esters (Table S1†). A series of different amounts of CuI were tested for the reaction to discover optimal amount of CuI (Table S1†). Addition of more than 0.1 equiv. of CuI did not provide any significant enhancement of synthetic yield (85% for 0.2, 0.5, 1.0 equiv. of CuI). The finding suggested that 0.1 equiv. of CuI was the proper amount to generate sulfonyl ester in this method.

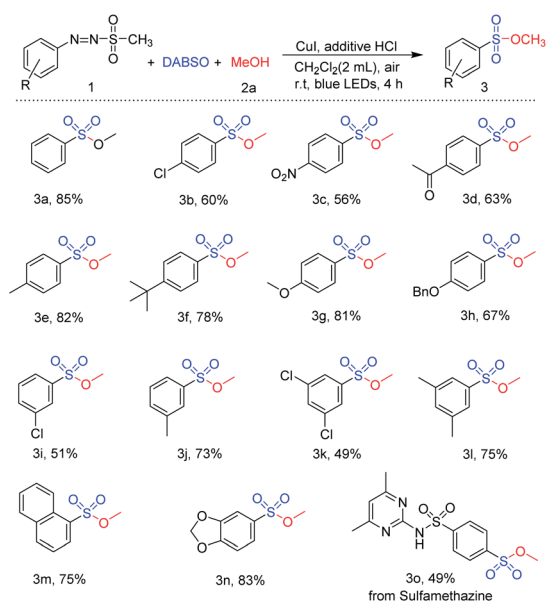
Table 2 Screening of solvents for the preparation of sulfonyl esters^a

Entry	Solvent	Yield ^b (%)
1	1,4-Dioxane	Trace
2	DMF	Trace
3	THF	22
4	Toluene	35
5	MeCN	39
6	DCE	73
7	CH_2Cl_2	85

^a Reaction conditions: compound **1** (1.0 mmol), MeOH (20 mmol), DABSO (1.0 mmol), CuI (0.1 mmol), HCl (1.0 mmol), irradiation by 5 W blue LEDs for 4 h. ^b Isolated yield after purification of flash column chromatography.

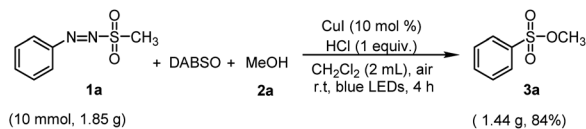
Investigation of the optimal amount of DABSO was also carried out for the reaction (Table S2†). Reactions using less than 0.1 equiv. of DABSO did not provide a satisfactory synthetic yield (43% for 0.8 equiv. of DABSO), while reactions in the presence of 1.0 equiv. of DABSO or higher successfully afforded the corresponding product in high yield. However, there was no significant increase in synthetic yield of the reaction when using higher than 1.0 equiv. of DABSO.

With the optimized reaction conditions in hand, the scope for visible-light-induced synthesis of sulfonyl ester was

Table 3 Substrate scope of arylazo sulfone^{a,b}

^a Reaction conditions: compound **1** (1.0 mmol), MeOH (20 mmol), DABSO (1.0 mmol), CuI (0.1 mmol), HCl (1.0 mmol), CH_2Cl_2 (2 mL), irradiation by 5 W blue LEDs for 4 h. ^b Isolated yield after purification of flash column chromatography.

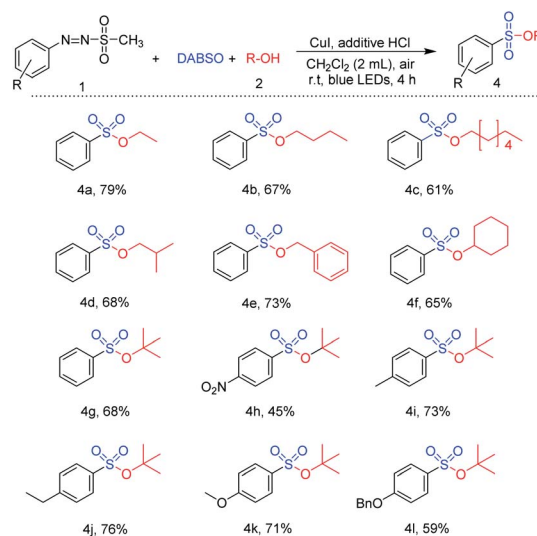


Scheme 2 Gram-scale reaction of arylazo sulfone (**1a**).

investigated. First, reactions of a variety of arylazo sulfones were carried out with methanol and DABSO in the presence of CuI and HCl (Table 3). Arylazo sulfones bearing different substituents were well tolerated for this synthetic method. Synthesis of sulfonic esters from arylazo sulfones with electron withdrawing substituents (chloro- and nitro-) at the *para* position was readily achieved (**3b** and **3c**). Similarly, reactions using arylazo sulfones with electron-donating groups (methyl, *tert*-butyl, and methoxy) at *para* position also afforded the target sulfonic esters in 78–82% yields (**3e–3g**). In addition, sulfonic esters with various functional groups including acetyl and benzyl protected hydroxyl groups were easily synthesized *via* visible-light-induced reaction of arylazo sulfone substrates (Scheme 2, compounds **3d** and **3h**). Arylazo sulfone substrates with substituents at *meta* position also reacted well to afford the desired sulfonic esters **3i–3j** in the yield range from 51 to 73%. Reactions of arylazo sulfones bearing di-substituted groups were investigated under the same reaction conditions, and the desired products such as 3,5-dichloro-substituted sulfonic ester **3k** and 3,5-dimethyl-substituted sulfonic ester **3l** were successfully synthesized in good yield. Moreover, substrates containing naphthalene and heterocyclic substituted rings were employed for novel visible-light-induced reaction, and they worked well with this reaction: methyl naphthalene-1-sulfonate and methyl benzo[*d*][1,3]dioxole-4-sulfonate were readily prepared in 75% and 83% yield, respectively (**3m** and **3n**). To investigate the applicability of this method in pharmaceuticals, substrate obtained from pharmaceutical compound was employed for this procedure. Reaction of the arylazo sulfone prepared from sulfamethazine, an antibacterial medicine, smoothly yielded the target sulfonic ester **3o**.

Next, the reaction was explored using various alcohols instead of methanol (Table 4). Synthesis of aryl sulfonyl ester was readily achieved *via* treatment of primary alcohol such as ethanol, *n*-butanol, *n*-heptanol, and iso-butyl alcohol with 1-(methylsulfonyl)-2-aryldiazene (61–79%) (**4a–4d**). Benzyl alcohol was also tested for the reaction, providing the desired product in 73% yield. A reaction with cyclohexanol, a secondary alcohol, was additionally evaluated, and cyclohexyl benzenesulfonate was readily synthesized in a good yield (**4f**). Moreover, *tert*-butanol, a bulky tertiary alcohol, was employed to react with several arylazo sulfones. Reactions of substrates bearing electron-withdrawing substituents (nitro-) and electron-donating substituents (methyl-, ethyl-, methoxyl-, and benzyloxy) with *tert*-butanol successfully gave the corresponding sulfonic esters in 45–76% yield under the same conditions of reaction.

It is note that novel synthetic method was better than the previous procedure reported by Han and co-workers (*J. Org.*

Table 4 Substrate scope of alcohols^{a,b}

^a Reaction conditions: compound **1** (1.0 mmol), alcohol (20 mmol), DABSO (1.0 mmol), CuI (0.1 mmol), HCl (1.0 mmol), CH₂Cl₂ (2 mL), irradiation by 5 W blue LEDs for 4 h. ^b Isolated yield after purification of flash column chromatography.

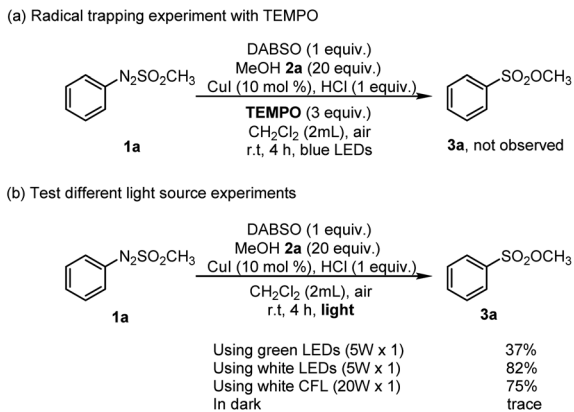
Chem., 2018, **83**, 4674–4680).³¹ The reaction time (4 h) of this study was shorter than 12 h of the previous method. In addition, these reaction yields of this study were better than those of the previous method (for example, compound **3a** in this study was generated with 85% yield *vs.* 50% yield in the previous method. Similarly, 82% *vs.* 64% for compound **3e**; 60% *vs.* 41% for compound **3b**; 73% *vs.* 29% for compound **4i**). Thus, the new visible-light-induced reaction of arylazo sulfones provided a more efficient synthetic protocol to prepare sulfonic esters rather than that of the previous method.

Next, a large-scale synthesis of sulfonic esters from arylazo sulfones was performed to see the utility of the procedure (Scheme 2). In reaction of arylazo sulfone **1a** (10.0 mmol, 1.85 g) with MeOH **4a**, the target sulfonic ester was successfully synthesized in 84% yield *via* the same reaction condition, suggesting that this reaction method can be effective and scalable.

To get an idea for the mechanism of this reaction, control reactions were performed (Scheme 3). First, 3.0 equiv. of TEMPO (2,2,6,6-tetramethyl-1-piperidinyloxy), a radical scavenger, was employed for a reaction system consisting of arylazo sulfone **1a**, DABSO, and methanol **2a**, and the conversion of arylazo sulfones **1a** into the corresponding sulfonic ester **3a** was fully inhibited by TEMPO (Scheme 3a). The fact that no sulfonic ester product **3a** was formed during reaction in the presence of TEMPO suggested that a radical pathway should be involved in this reaction. Next, the reaction was carried out in the dark, and it was found that no desired product **3a** was observed (Scheme 3b).

The UV-vis spectra of some arylazo sulfone compounds in reaction solvent CH₂Cl₂, as well as in the presence of DABSO and CuI, have been quantified (Fig. S2–S5†). The results show that arylazo sulfone compounds exhibit absorption maxima





Scheme 3 Control experiments.

located at 410–460 nm wavelength. This wavelength range belongs to the blue region in the visible light. Therefore, using blue LEDs is suitable for this reaction.

Moreover, we investigated the role of visible light for this transformation by performing light on/off experiments (Fig. S1†). The reaction was almost suppressed when carried out in the dark, while the desired product **3a** was smoothly prepared with increasing yield when starting substrate **2a** was irradiated by blue LEDs. The light on/off experiment of the model reaction indicated that continuous irradiation of visible light is important for the progress of this reaction.

With the information from the control experiments and previously published literature,¹⁹ we proposed a plausible mechanism for this reaction (Scheme 4). Initially, under irradiation of blue LED light, arylazo sulfone compounds **1a** were homolytically divided at N–S bond to generate aryl radical **4**.^{27–30} The radical **4** captured an SO₂ group of DABSO **7** to form arylsulfonyl radical **8** and the DABCO cation **9**. The catalyst Cu¹⁺ **10** was oxidized by DABCO cation **9** or other oxidizing agents (O₂)

to give Cu²⁺ **11**, which reacted with MeOH **2a** to form complex Cu²⁺–OMe **12**. The complex **12** then reacted with the arylsulfonyl radical **8** to give intermediate **13**. A reductive elimination process of **13** generated the desired compound **3a**, and recovered Cu¹⁺ **10**.

Conclusions

In summary, a new methodology for visible-light-induced synthesis of sulfonic esters from arylazo sulfones was described. In the present study, multi-component reaction using arylazo sulfones, DABSO, alcohols, and CuI was successfully achieved by the irradiation of blue LED light. A wide range of arylazo sulfone substrates with different substituents and several alcohols were well tolerated in this synthetic method. This novel visible-light-induced reaction produced various sulfonic esters under mild reaction conditions. The results suggested that this efficient synthetic approach using visible light can serve as a useful tool to afford a variety of the highly valuable arylsulfonic esters.

Experimental

General procedure of the synthesis of sulfonic esters (**3a–3o**, **4a–4l**)

In a typical synthetic procedure, arylazo sulfone compound **1a** (185 mg, 1.0 mmol), DABSO (240 mg, 1.0 mmol), and CuI (19 mg, 0.1 mmol) were added to a mixed solution of dichloromethane (2 mL) and methanol (640 mg, 20 mmol). Hydrochloric acid 37% (100 mg, 1.0 mmol) was dropped into the reaction mixture. The mixture was stirred at room temperature under irradiation by blue LEDs. After 4 hours, the mixture was extracted with 50 mL of CH₂Cl₂ and washed with 50 mL of brine solution. The organic layer was dried by sodium sulfate and concentrated under reduced pressure. The residue was purified using flash column chromatography on silica gel with hexane–EtOAc as the eluent to get the intended product **3a** (146 mg, 85%).

Author contributions

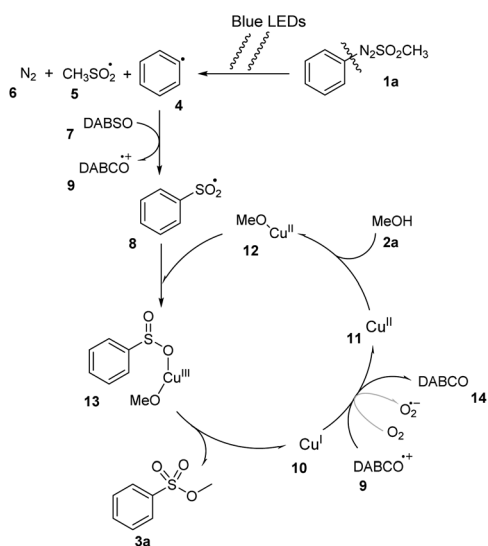
H.-K. Kim conceived and designed this work. T. G. Luu and T. T. Bui performed the experiments and collected data. All authors analysed the data, discussed the results, and wrote the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research was supported by the National Research Foundation of Korea (NRF) grant funded by the Korea government (MSIT) (NRF-2021R1A2C1011204).



Scheme 4 Proposed reaction mechanism.



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