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# Synthesis and mechanism of biological action of morpholinyl-bearing arylsquaramides as small-molecule lysosomal pH modulators†

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Lysosomal pH is an important modulator for many cellular processes. An agent that is capable of regulating lysosomal pH may find a wide range of potential applications in the field of biomedicine. In this study, we describe the synthesis of a family of morpholinyl-bearing arylsquaramides as small-molecule lysosomal pH modulators. These compounds are able to efficiently facilitate the transmembrane transport of chloride anions as mobile carriers across vesicular and cellular phospholipid membranes. They are capable of specifically alkalinizing liposomes, disrupting the homeostasis of lysosomal pH and inactivating lysosomal Cathepsin B enzyme. Anion transport is considered as the probable mechanism of action for the high efficiency of these compounds to modulate lysosomal pH. The present findings present a novel means to efficiently regulate lysosomal pH, which is in contrast to the methods shown by conventional lysosomal pH modulators that generally function by either acting as a weak base/acid, or releasing a basic/acidic component in lysosomal environments to change lysosomal pH.

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## 1. Introduction

It is known that lysosomes are membrane-bound subcellular organelles and maintain cellular homeostasis by generating a highly acidic environment.<sup>1</sup> Lysosomal enzymes are responsible for breaking down carbohydrates, nucleic acids, lipids and proteins,<sup>2–6</sup> and usually exert the maximal activity in an acidic environment (pH 4.5–5.0).<sup>7–14</sup> Leakage of protons due to a defective proton pump or lysosomal permeabilization would lead to lysosomal alkalization and consequently a decrease or deactivation in the activity of lysosomal hydrolytic enzymes.<sup>15,16</sup> Thus, modulating lysosomal pH may serve as a practical strategy for regulating cellular processes as well as for developing therapeutic agents for lysosome-involved diseases.

It is reported that some agents are able to specifically modulate lysosomal pH.<sup>1</sup> These so-called lysosomal pH modulators generally function by acting as a weak base/acid, or releasing a basic/acidic component in lysosomal environments to elevate/lower lysosomal pH.<sup>1</sup> For example, chloroquine and hydroxychloroquine inhibit lysosomal acidification through

elevating lysosomal pH,<sup>17–19</sup> whereas poly(lactic-co-glycolic acid) (PLGA) degrades in an acidic environment to release lactic and glycolic component carboxylic acids to lower lysosomal pH (Fig. 1).<sup>20–22</sup> In contrast to these conventional lysosomal pH modulators, in previous studies we introduced the concept of “lysosome-targeting anionophores”, whereby arylsquaramides bearing fluorescent coumarinyl and lysosome-targeting morpholinyl subunits (e.g., compound A, Fig. 2),<sup>23,24</sup> are used to target lysosomes and disrupt the homeostasis of anions within liposomes. Significantly, we have found that such squaramides are able to elevate lysosomal pH,<sup>23,24</sup> and that enhancing their anion transport activity is favorable to the efficiency to alkalinize lysosomes. This represents a novel approach to efficiently alter lysosomal pH.<sup>24</sup>

In the work reported herein, we sought to gain insight into the correlation between anion transport activity and lysosomal pH-regulating efficiency, and to better characterize the

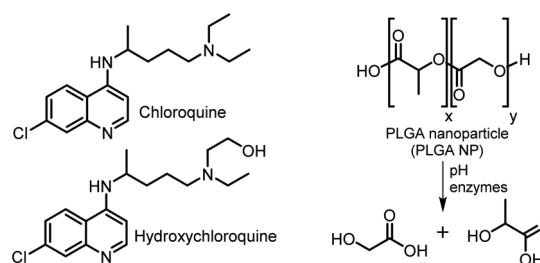


Fig. 1 Agents capable of modulating lysosomal pH.

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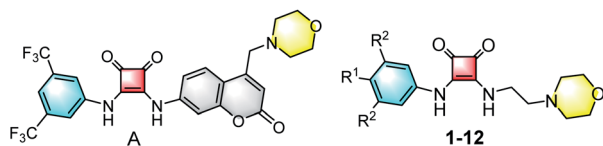


Fig. 2 Structures of compounds A and 1–12. For the specific substituents  $R^1$  and  $R^2$ , see Scheme 1.

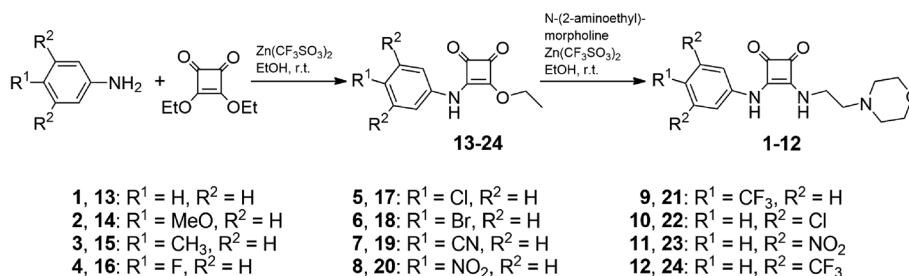
lysosomal pH-modulating properties. Specifically, we designed and synthesized a family of morpholinyl-bearing arylsquar-amides with no fluorophores (*i.e.*, compounds 1–12, Fig. 2), and systematically investigated their capability and probable

mechanism of biological action to modulate lysosomal pH by facilitating the transmembrane transport of anions across vesicular and cellular phospholipid membranes.

## 2. Results and discussion

### 2.1 Synthesis

Compounds 1–12 were synthesized as shown in Scheme 1. Thus, reaction of substituted anilines with excess diethyl squarate under the catalysis of  $Zn(CF_3SO_3)_2$  gave compounds 13–24.<sup>25–28</sup> Reaction of compounds 13–24 with excess *N*-(2-aminoethyl)morpholine under the catalysis of  $Zn(CF_3SO_3)_2$  afforded compounds 1–12. The structures of compounds 1–12 were



Scheme 1 Synthesis of compounds 1–12.

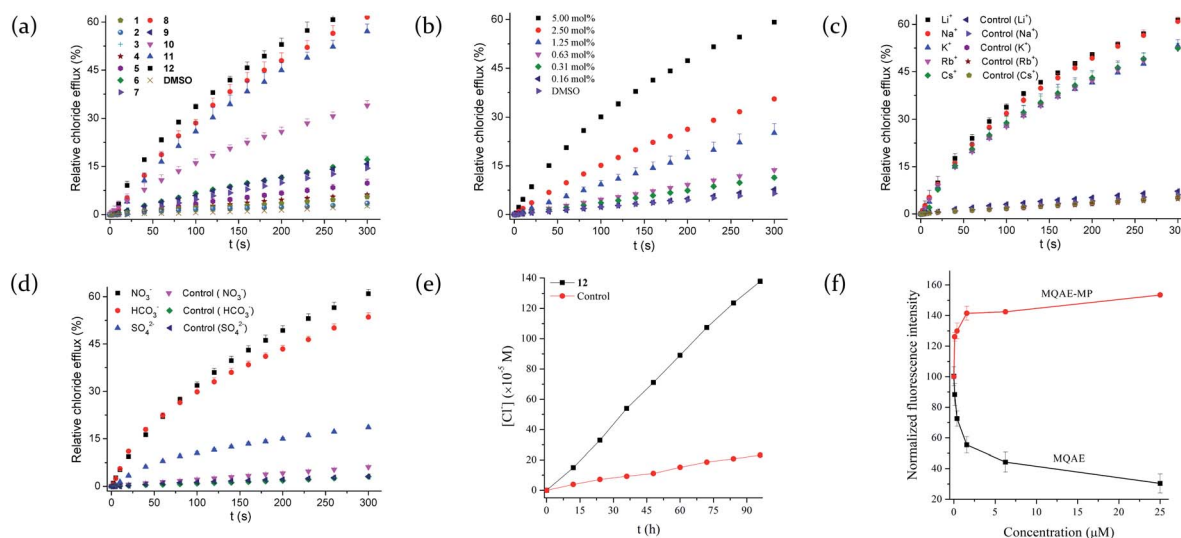


Fig. 3 (a) Relative efflux of chloride anions out of EYPC liposomes enhanced by compounds 1–12 (5 mol%). (b) Relative efflux of chloride anions out of EYPC liposomes, induced by compound 12 of varying concentrations. Assay conditions for (a) and (b): an intravesicular 500 mM NaCl solution in 25 mM HEPES buffer (pH 7.0) and extravesicular 500 mM  $NaNO_3$  in 25 mM HEPES buffer (pH 7.0). (c) Relative efflux of chloride anions out of EYPC liposomes enhanced by compound 12 (5 mol%), under the conditions of an intravesicular 500 mM MCl solution in 25 mM HEPES buffer (pH 7.0) and extravesicular 500 mM  $NaNO_3$  in 25 mM HEPES buffer (pH 7.0) ( $M = Li, Na, K, Rb$  or  $Cs$ ). (d) Relative efflux of chloride anions out of EYPC liposomes enhanced by compound 12 (5 mol%), under the conditions of an intravesicular 500 mM NaCl in 25 mM HEPES buffer (pH 7.0) and extravesicular 500 mM  $NaNO_3$ ,  $NaHCO_3$  or 250 mM  $Na_2SO_4$  in 25 mM HEPES buffer (pH 7.0). (e) Transport of chloride anions across a bulky nitrobenzene membrane, promoted by compound 12 (1 mM) and detected by a chloride ion selective electrode, under the conditions of 500 mM  $NaNO_3$  (25 mM HEPES, pH 7.0) for the U-tube chloride receiving aqueous phase, 500 mM NaCl (25 mM HEPES, pH 7.0) for the U-tube chloride donating aqueous phase and 2 mM TBAPF<sub>6</sub> in the nitrobenzene organic phase. (f) Normalized fluorescence intensity (FI) of MQAE and MQAE-MP in HeLa cells. For MQAE assays, the cells were incubated with MQAE (5 mM) for 4.0 h followed by the treatment with compound 12 of varying concentrations for 2 h, whereas for MQAE-MP assay, the cells were incubated with compound 12 of varying concentrations for 4 h, followed by the treatment with MQAE-MP (5 mM) for 0.5 h. The FI was recorded by a plate reader at  $\lambda_{em}/\lambda_{ex} = 460/350$  nm and normalized with respect to the FI of untreated cells. Each data point represents the mean FI of three independent experiments.



confirmed by means of NMR ( $^1\text{H}$  and  $^{13}\text{C}$ ) and ESI MS (LR and HR) (see experimental section and Fig. S1–S52†).

## 2.2 Anion transport

**Vesicular anionophoric activity.** Before we investigated the lysosomal pH-modulating ability of compounds 1–12, we firstly studied their anion transport activity by measuring the efflux of chloride anions out of liposomes (100 nm, extrusion) composed of egg-yolk phosphatidylcholine (EYPC), using chloride ion selective electrode (ISE) techniques.<sup>29,30</sup> The results are shown in Fig. 3 and Table 1, and indicate that compounds 1–12 are capable of mediating the transport of chloride anions across phospholipid liposomal membranes, but with obvious differences in the activity. For example, at the concentration of 5 mol% (relative to the concentrations of EYPC), compound 12 bearing bistrifluoromethyl groups exhibits relative efflux of higher than 60% chloride anions out of the EYPC vesicles, whereas compound 1 shows negligible chloride efflux. This result suggests that the chloride efflux ability may be regulated by the substituents. Generally, the compounds bearing strongly electron-withdrawing substituents, such as compounds 8, 11 and 12 exhibit relatively high activity. This is in good accord with literature report that electron-withdrawing substituents on the aryl subunits of an arylsquaramide are favorable to the anion transport activity.<sup>31</sup>

To quantitatively characterize the efficiency of the compounds showing relatively high chloride efflux activity, we carried out the concentration-dependent experiments and analyzed the relationship between the relative chloride efflux at 260 s and the concentrations of each compound, by using a Hill equation, relative chloride efflux =  $k_{\text{max}} \times [\text{compound}]^n / \{[\text{compound}]^n + (\text{EC}_{50})^n\}$ . This analysis gives the  $\text{EC}_{50}$  value of each compound that is defined as the effective transporter loading when 50% of the maximum chloride

efflux is reached. As shown in Table 1, compounds 8, 11 and 12 exhibit relatively high activity. It should be noted that these compounds are 5–6-fold less active than compound A, which may be ascribed to their lower lipophilicity relative to compound A.<sup>32</sup>

To clarify whether the chloride efflux activity is authentic, we chose compound 12 to measure its effect on the size distribution and integrity of the EYPC vesicles (Fig. S54 and S55†). As a result, the diameters of the vesicles were  $117.2 \pm 30.8$  nm and  $113.3 \pm 37.9$  nm in the absence and presence of 5 mol% of compound 12, respectively (Table S1†). This result suggests that compound 12 has no effect on the size distribution of the vesicles. Study by means of calcein leakage assay reveals that no calcein was leaked from the vesicles after the addition of compound 12.<sup>30,32</sup> These results suggest that compound 12 has no effect on the integrity of vesicles and the observed chloride efflux is authentic.

**Mechanism of action of chloride efflux.** To gain insight into the probable mechanism of action of compounds 1–12, we chose compound 12 to re-carry out the chloride efflux experiments in the presence of different alkali metal ions ( $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Rb}^+$  or  $\text{Cs}^+$ ) or anions ( $\text{NO}_3^-$ ,  $\text{HCO}_3^-$  or  $\text{SO}_4^{2-}$ ).<sup>33,34</sup> The results are shown in Fig. 3c and d, and clearly indicate that the chloride efflux activity of compound 12 is essentially independent of group I alkali metal ions, but varies with the anions. This result excludes the determining role of those metal ions in the permeation process, and strongly suggests that the efflux of chloride anions proceeds *via* an anion exchange process.<sup>33,34</sup> Then we carried out a U-tube experiment to address whether compound 12 functions as a channel or a carrier.<sup>35</sup> In this experiment, the chloride-donating phase was separated from the chloride-receiving phase by a bulky organic phase. As shown in Fig. 3e, the presence of compound 12 led to a time-dependent increase in the concentration of chloride anions in the receiving phase, suggesting that chloride anions were transported from the donating phase to the receiving phase.

Table 1 Lipophilicity and chloride transport efficiency of compounds 1–12

Compound	$R^1$	$R^2$	Lipophilicity ( $c \log P^d$ )	Chloride transport <sup>b</sup>	
				$\Delta R_{\text{Cl}^- \text{efflux}} (\%)^c$	$\text{EC}_{50} (\text{mol}\%, \text{ISE})^d$
1	H	H	0.79	$2.4 \pm 0.79$	$\text{NA}^e$
2	$\text{OCH}_3$	H	0.64	$0.7 \pm 0.16$	$\text{NA}^e$
3	$\text{CH}_3$	H	1.31	$0.5 \pm 0.01$	$\text{NA}^e$
4	F	H	0.94	$3.3 \pm 0.01$	$\text{NA}^e$
5	Cl	H	1.40	$6.2 \pm 0.59$	$\text{NA}^e$
6	Br	H	1.56	$12.6 \pm 0.19$	$>10^f$
7	CN	H	0.65	$10.4 \pm 1.40$	$>10^f$
8	$\text{NO}_2$	H	0.73	$54.3 \pm 2.00$	$3.87 \pm 0.19$
9	$\text{CF}_3$	H	1.67	$11.9 \pm 1.04$	$>10^f$
10	H	Cl	2.00	$28.5 \pm 1.75$	$6.81 \pm 0.29$
11	H	$\text{NO}_2$	0.67	$50.1 \pm 1.56$	$4.19 \pm 0.18$
12	H	$\text{CF}_3$	2.55	$58.6 \pm 3.30$	$4.46 \pm 0.29$

<sup>a</sup> Calculated using MarvinSketch (Version 6.1.0, weighted Model, ChemAxon, MA). <sup>b</sup> Measured on liposomal models by means of chloride ion selective electrode (ISE) techniques, under the conditions of an intravesicular 500 mM NaCl solution (25 mM HEPES buffer) and extravesicular 500 mM  $\text{NaNO}_3$  solution (25 mM HEPES buffer). <sup>c</sup>  $\Delta R_{\text{Cl}^- \text{efflux}} = R_{\text{Cl}^- \text{efflux}} [5.0 \text{ mol}\% \text{ compound}]_{260 \text{ s}} - R_{\text{Cl}^- \text{efflux}} [\text{DMSO}]_{260 \text{ s}}$ . <sup>d</sup> Compound A with a  $c \log P$  of 3.71 has an  $\text{EC}_{50}$  of  $0.73 \pm 0.28$  mol% under the same assay conditions. <sup>e</sup> Hill analysis not performed due to very low transport activity. <sup>f</sup> Estimated due to low to moderate transport activity.



Taken together, the above-mentioned results suggest that compound **12** facilitates the transport of chloride anions across phospholipid membranes as a mobile carrier *via* the process of an anion exchange.

**In vitro chloride influx activity.** As compound **12** is able to facilitate the transport of chloride anions across vesicular membranes, we are concerned about whether it exhibits anionophoric activity in living cells. To address this, we measured the activity of compound **12** to facilitate the influx of chloride anions into HeLa cervical cancer cells by means of an MQAE [*N*-(ethoxycarbonylmethyl)-6-methoxyquinolinium bromide] assay.<sup>36–38</sup> It is known that MQAE is a cell-permeable dye and its fluorescence is specifically quenched by chloride anions. As shown in Fig. 3f, incubation of MQAE with HeLa cells and subsequent treatment of the cells with compound **12**, led to a concentration-dependent decline in the fluorescence of MQAE. This provides strong evidences that compound **12** is involved in facilitating the influx of chloride anions into the intracellular matrix.

Since compound **12** has lysosome-targeting morpholinyl groups, we used MQAE-MP assays to explore its ability to facilitate the transport of chloride anions across lysosomal membranes.<sup>39</sup> MQAE-MP is an MQAE derivative that is able to specifically accumulate in lysosomes and has been used to detect the change in the concentration of chloride anions in lysosomes.<sup>40</sup> As shown in Fig. 3f, treatment of HeLa cells with compound **12** and subsequent incubation with MQAE-MP, led to a concentration-dependent increase in the fluorescence of MQAE-MP, indicative of a decrease in the concentration of chloride anions in lysosomes. The result lends strong support that compound **12** is able to mediate the efflux of chloride anions out of lysosomes.

### 2.3 Lysosomal pH-modulating effect

**Alkalinization of acidic organelles.** To demonstrate the ability of compounds **1–12** to modulate intracellular pH, we conducted staining experiments of HeLa cells with acridine orange (AO).<sup>16,41,42</sup> As a cell-permeable dye, AO exhibits a characteristic

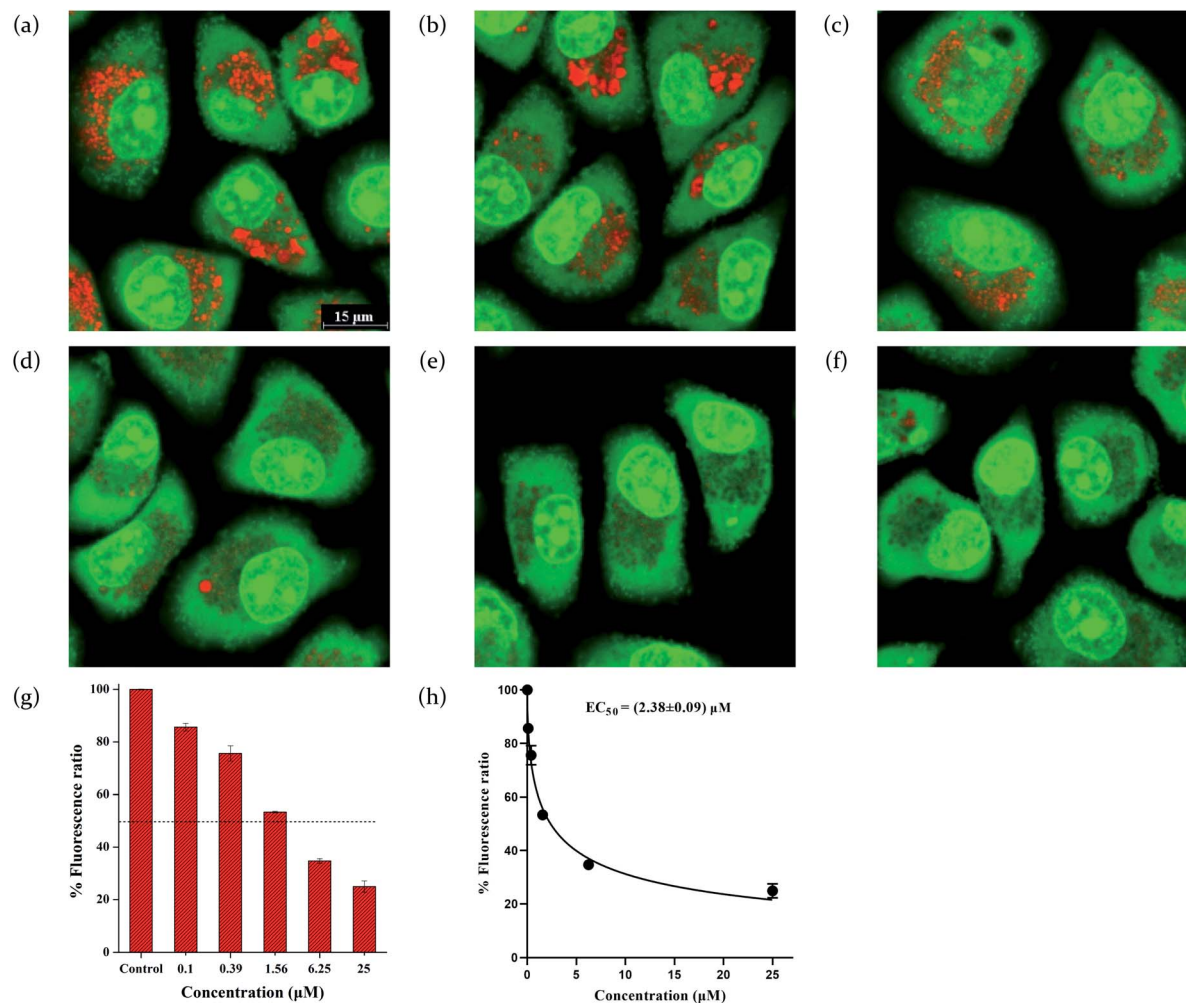


Fig. 4 AO staining of (a) untreated HeLa cells and (b–f) HeLa cells treated with compound **12** at the concentrations of (b) 0.10 μM; (c) 0.39 μM; (d) 1.56 μM; (e) 6.25 μM and (f) 25 μM for 4 h, respectively. Green fluorescence:  $\lambda_{\text{ex}}$  BP 470/40 nm,  $\lambda_{\text{em}}$  BP 525/50 nm; red fluorescence:  $\lambda_{\text{ex}}$  BP 546/12 nm,  $\lambda_{\text{em}}$  BP 575–640 nm. (g) Graph of the AO fluorescence against the concentrations of compound **12**. (h) Curve fitting of the AO fluorescence against the concentrations of compound **12**.

orange fluorescence emission when it is protonated and accumulated in acidic compartments such as lysosomes, whereas it emits green fluorescence when the acidic compartments are basified. It is clear from Fig. 4 and S56–S67† that, when HeLa cells were stained with AO, granular orange fluorescence was observed in the cytoplasm, suggesting that the orange fluorescence is due to acidic organelles. Treatment of the HeLa cells with compounds 1–12 led to a decline in the orange emission. This result suggests that compounds 1–12 are able to efficiently basify the acidic organelles.

To quantitatively characterize the efficiency of these compounds to modulate cellular pH, we carried out concentration-dependent experiments and determined the EC<sub>50</sub> values that are defined as the concentrations required to shift the orange/green ratios of AO by 50%.<sup>43</sup> The EC<sub>50</sub> values were listed in Table 2, and show that compounds 1–12 alkalize acidic organelles in the order of compound 12 > 11 > 8 > 10 > 9 > 7 > 6 > 5 > 1 ≈ 2 ≈ 3 ≈ 4, which is consistent with their transport activity on vesicles.

**Lysosomal alkalization.** The ability of compounds 1–12 to alkalize acidic organelles inspired us to explore whether they are able to alkalize lysosomes. Thus, we carried out the co-staining experiments of compounds 1–12 with commercially available LysoSensor Green DND-189, a pH-sensitive fluorescent probe that specifically targets lysosomes.<sup>44,45</sup> Chloroquine phosphate (CQP), a well-known lysosomal pH modulator, was used as a positive control. As showed in Fig. 5 and S68–S80,† treatment of HeLa cells with each compound led to a concentration-dependent reduction in the green fluorescence, indicative of an increase in the lysosomal pH. This observation indicates that compounds 1–12 are able to alkalize lysosomes.

To quantitatively assess the efficiency of these compounds, we analyzed the green fluorescence intensity against the concentrations of each compound to give the EC<sub>50</sub> values. This EC<sub>50</sub> value is defined as the concentrations required to result in 50% disappearance of the green fluorescence of LysoSensor

Green DND-189 (Table 2). The EC<sub>50</sub> values suggest that compounds 1–12 alkalize lysosomes in the order of compound 12 > 11 > 8 > 10 > 9 > 7 > 6 > 5 > 1 ≈ 2 ≈ 3 ≈ 4. This order is consistent with the ones obtained in the liposome and AO assays. Notably, compounds 11 and 12 exhibit comparable ability to alkalize lysosomes with CQP.

As CQP is inactive in facilitating the efflux of chloride anions (data not shown), we propose that compounds 1–12 modulate lysosomal pH in a different mechanism with CQP. The ability of compounds 1–12 to basify acidic organelles and lysosomes in HeLa cells parallels their anionophoric activity across liposomal membranes, which strongly suggests that it is the anion exchange promoted by compounds 1–12 that could be responsible for the increase of lysosomal pH. Specifically, compounds 1–12 facilitate the efflux of chloride anions out of lysosomes and meanwhile the influx of hydroxide or bicarbonate into the interior of lysosomes, leading to an increase in the lysosomal pH.<sup>46</sup>

**Quantitative determination of lysosomal pH.** To quantitatively determine the effect of compounds 1–12 on lysosomal pH, we firstly carried out staining experiments with HeLa cells, by using BCECF-AM [2',7'-bis-(2-carboxyethyl)-5-(and-6)-carboxyfluorescein, acetoxymethyl ester] assays.<sup>47</sup> BCECF-AM is a cell-permeable and intracellular pH-sensitive fluorescent probe that does not fluoresce by itself. When entering the cells, BCECF-AM may be sheared into BCECF and AM by intracellular esterases, which can be retained in the cells to emit fluorescence. BCECF-AM can be measured ratiometrically because the ratio of the two emitted fluorescence corresponding to its excitation has a clear linear relationship with pH. As shown in Fig. S81† and Table 2, after the treatment with compounds 1–12, the intracellular pH (pH<sub>i</sub>) in HeLa cells changes with time and increases by up to 1.0 unit after 3 h. Among them, compounds 11 and 12 showed the strongest effect, which is comparable to CQP.

Table 2 Lysosomal alkalization efficiency of compounds 1–12 and chloroquine phosphate<sup>a</sup>

Compound	Cellular alkalization		Lysosomal alkalization		
	EC <sub>50</sub> <sup>b</sup> (μM, AO)	ΔpH <sup>c</sup>	EC <sub>50</sub> <sup>d</sup> (μM, LysoSensor)	ΔpH <sup>e</sup>	Lysosomal enzymes <sup>f</sup> (%)
1	>50	0.03 ± 0.15	>50	−0.02 ± 0.28	92.7 ± 2.07
2	>50	0.08 ± 0.11	>50	0.01 ± 0.21	95.0 ± 1.46
3	>50	0.09 ± 0.20	>50	−0.14 ± 0.14	101.9 ± 1.77
4	>50	0.00 ± 0.04	>50	0.58 ± 0.05	87.7 ± 1.08
5	36.9 ± 2.25	0.04 ± 0.20	39.8 ± 2.48	0.66 ± 0.19	75.4 ± 1.02
6	17.3 ± 0.55	0.07 ± 0.16	17.3 ± 0.86	1.03 ± 0.05	72.3 ± 0.75
7	11.3 ± 0.60	0.15 ± 0.12	14.0 ± 0.45	1.17 ± 0.04	69.9 ± 1.21
8	5.56 ± 0.22	0.79 ± 0.15	5.19 ± 0.36	1.82 ± 0.12	46.2 ± 0.94
9	9.56 ± 0.60	0.38 ± 0.06	10.9 ± 0.54	1.28 ± 0.15	63.2 ± 1.91
10	8.25 ± 0.82	0.72 ± 0.18	8.47 ± 0.50	1.51 ± 0.10	57.2 ± 2.37
11	3.31 ± 0.30	1.04 ± 0.15	3.54 ± 0.18	1.92 ± 0.07	42.9 ± 2.10
12	2.38 ± 0.09	1.14 ± 0.03	3.09 ± 0.05	1.90 ± 0.01	43.2 ± 1.78
CQP	1.44 ± 0.18	1.36 ± 0.02	1.34 ± 0.23	2.17 ± 0.13	37.3 ± 2.45

<sup>a</sup> All the measurements were made on HeLa cells. <sup>b</sup> Measured by using AO assays. <sup>c</sup> Measured by using BCECF-AM assays, after treatment for 3 h. <sup>d</sup> Measured by using LysoSensor Green DND-189 assays. <sup>e</sup> Measured by using the assays based on dextran, fluorescein and tetramethylrhodamine, after treatment for 12 h. <sup>f</sup> Measured by using Magic Red Cathepsin assays and reported relative to the background.



Next, we used a ratiometric lysosomal pH probe, fluorescein-tetramethylrhodamine-labeled dextran (FRD),<sup>38,42,48</sup> to quantitatively measure the effect of compounds 1–12 on the lysosomal pH (Fig. 6, S82† and Table 2). FRD has a relative molecular weight of 70 000 and may accumulate in lysosomes by endocytosis. It emits two types of fluorescence, green from fluorescein and red from tetramethylrhodamine, respectively. The ratios of green/red fluorescence are well correlated with the pH changes in lysosomes, whereby fluorescein acts as a pH-dependent fluorophore and tetramethylrhodamine as a pH-independent fluorophore.<sup>49</sup> As a consequence, treatment of HeLa cells with compounds 1–12 increased the lysosomal pH by up to 1.92 units. These results clearly show that compounds 1–12 are able to cause lysosomal alkalinization, and notably compounds 11 and 12 are comparable to CQP.

**Decreased activity in lysosomal enzymes.** It is known that lysosomes contain a variety of degrading enzymes.<sup>2–5,50</sup> These enzymes, such as Cathepsin B enzyme exert the highest biological activity in an acidic environment (pH 4.5–5.0), and

lysosomal alkalinization leads to a decrease or deactivation in the activity of lysosomal hydrolytic enzymes.<sup>14,38</sup> Thus, we tested the effect of compounds 1–12 on the activity of Cathepsin B enzyme using a Magic Red Cathepsin assay. In this assay, when the Magic Red<sup>®</sup> MR-(RR)<sub>2</sub> Cathepsin B substrate, a cell-permeable, non-fluorescent probe, is hydrolyzed by lysosomal Cathepsin B enzymes, the two R–R peptide sequences are cleaved from the Magic Red molecule and as a consequence red fluorescence emits.<sup>42</sup>

As shown in Fig. 7 and Table 2, treatment of HeLa cells with compounds 1–12 led to a decay in the red fluorescence within the cells, which suggests that the activity of Cathepsin B enzyme decreases and compounds 1–12 are able to change the lysosomal pH. Quantitative analysis of the red fluorescence intensity relative to the control indicates that at the tested concentration, the red fluorescence decreases in the order of compounds 12 > 11 > 8 > 10 > 9 > 7 > 6 > 5 > 1 ≈ 2 ≈ 3 ≈ 4, which is in good agreement with the ability of compounds 1–12 to transport anions and alkalinize lysosomes.

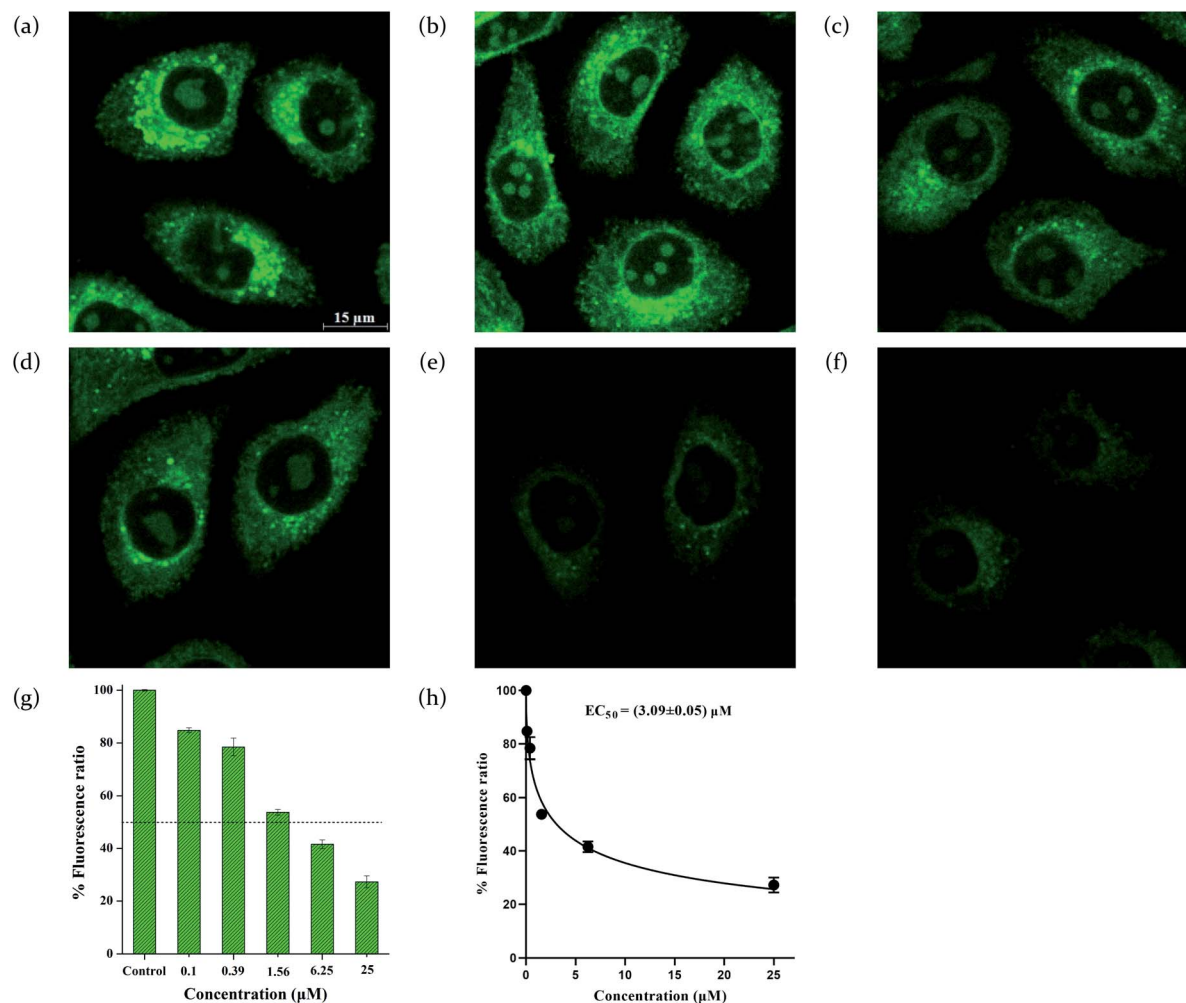
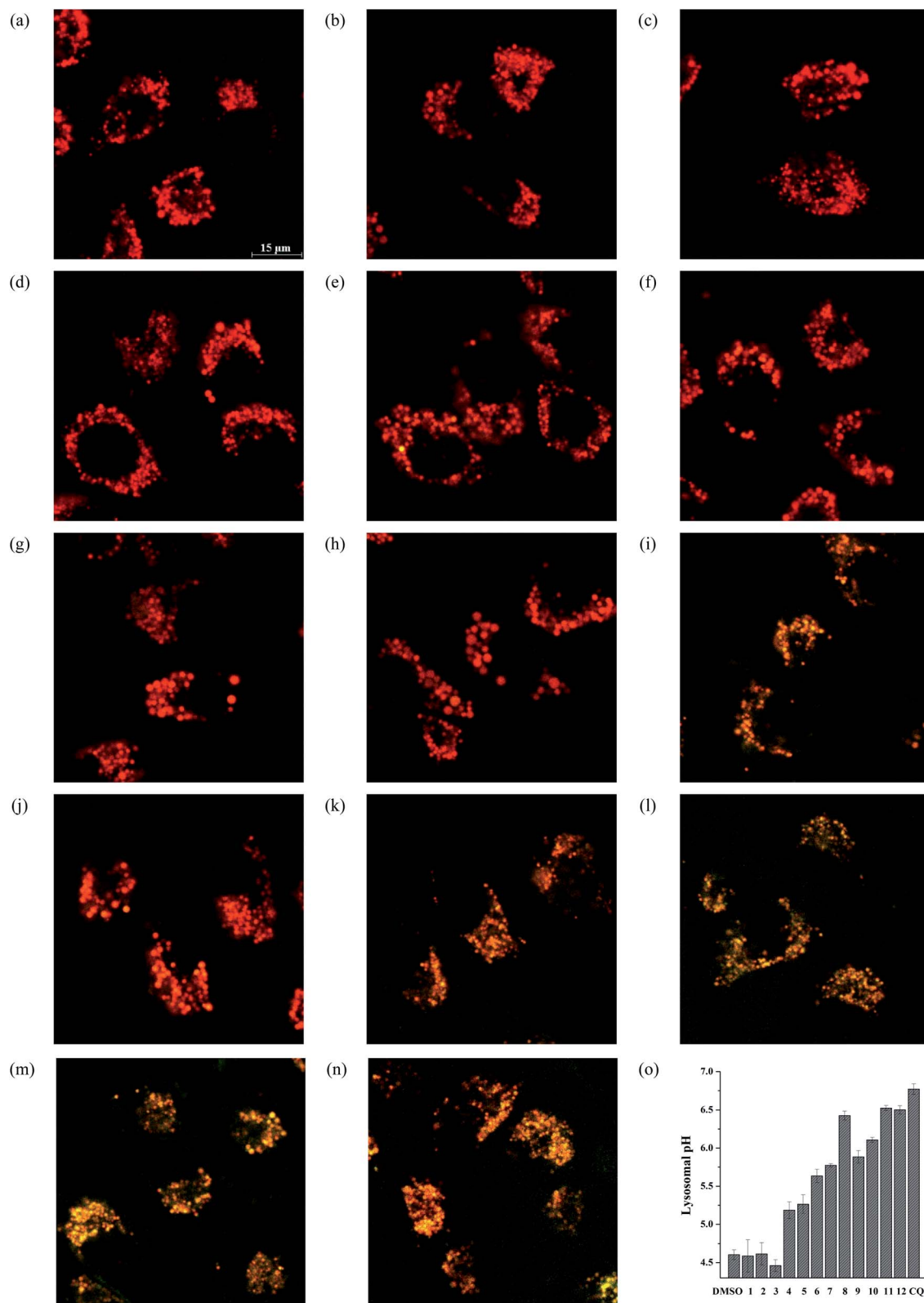


Fig. 5 LysoSensor Green DND-189 (1.0 μM) staining of (a) untreated HeLa cells and (b–f) HeLa cells treated with compound 12 at the concentrations of (b) 0.1 μM; (c) 0.39 μM; (d) 1.56 μM; (e) 6.25 μM and (f) 25 μM for 4 h, respectively. Green fluorescence:  $\lambda_{\text{ex}}$  BP 470/40 nm,  $\lambda_{\text{em}}$  BP 525/50 nm. (g) Graph of the LysoSensor Green DND-189 fluorescence against the concentrations of compound 12. (h) Curve fitting of the LysoSensor Green DND-189 fluorescence against the concentrations of compound 12.



**Fig. 6** (a–n) HeLa cells incubated with fluorescein-tetramethylrhodamine-labeled dextran ( $0.2 \text{ mg mL}^{-1}$ ) for 12 h, and then (a) untreated and (b–n) pretreated with  $6.25 \text{ }\mu\text{M}$  of compounds **1** (b), **2** (c), **3** (d), **4** (e), **5** (f), **6** (g), **7** (h), **8** (i), **9** (j), **10** (k), **11** (l), **12** (m) and CQP (n) for 12 h, respectively. Green fluorescence:  $\lambda_{\text{ex}}$  BP 470/40 nm,  $\lambda_{\text{em}}$  BP 525/50 nm; red fluorescence:  $\lambda_{\text{ex}}$  BP 546/12 nm,  $\lambda_{\text{em}}$  BP 590–700 nm (o) Impact of  $6.25 \text{ }\mu\text{M}$  of compounds **1–12** and CQP on lysosomal pH. The lysosomal pH was determined by using a pH titration curve.



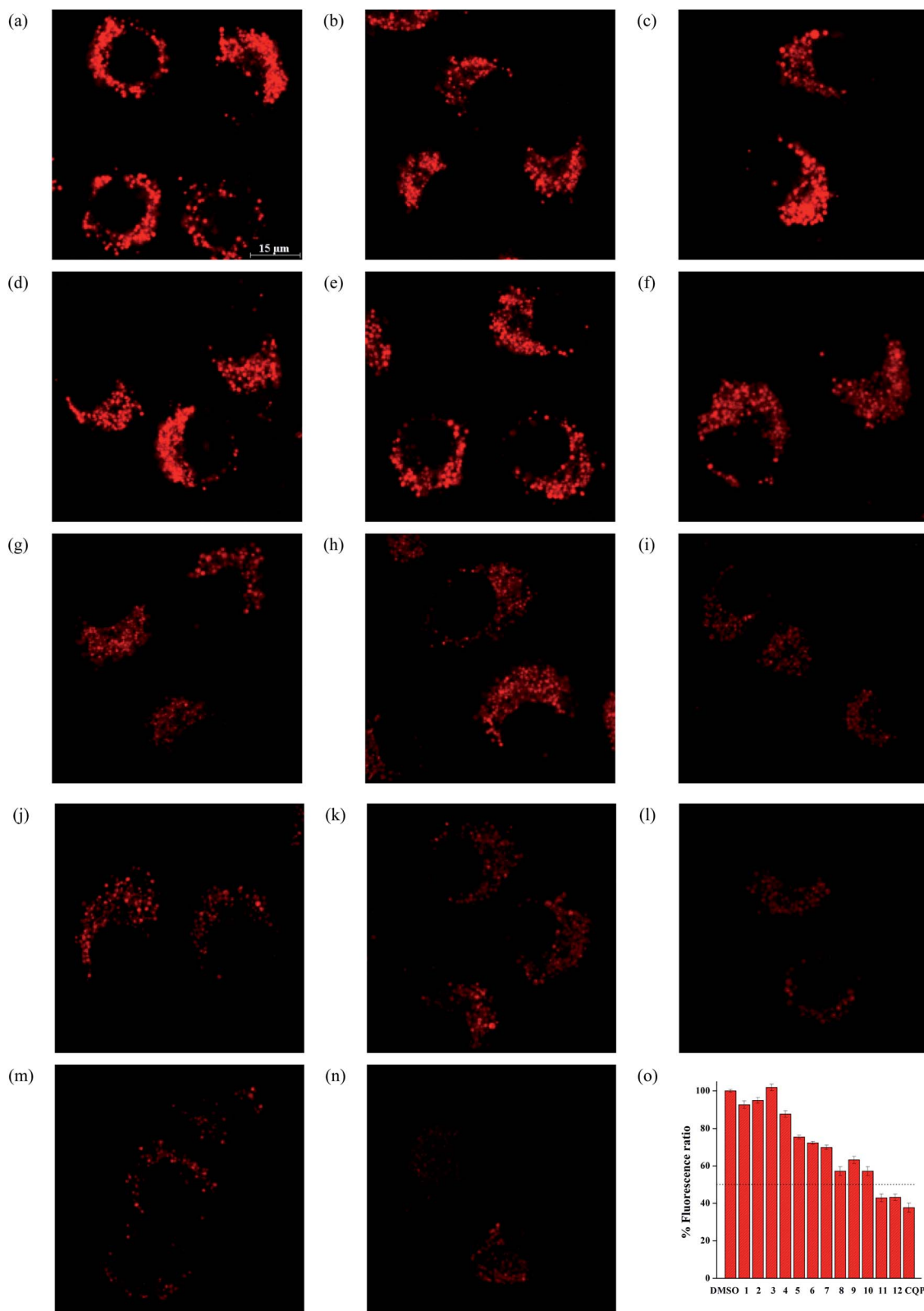
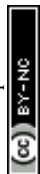


Fig. 7 (a–n) HeLa cells (a) untreated and (b–n) pretreated with 6.25  $\mu\text{M}$  of compounds 1 (b), 2 (c), 3 (d), 4 (e), 5 (f), 6 (g), 7 (h), 8 (i), 9 (j), 10 (k), 11 (l), 12 (m) and CQP (n) for 4 h, respectively, and then incubated with MR-(RR)<sub>2</sub> for 30 min. Red fluorescence:  $\lambda_{\text{ex}}$  BP 546/12 nm,  $\lambda_{\text{em}}$  BP 575–640 nm. (o) Relative FI of MR-(RR)<sub>2</sub> in the presence of 6.25  $\mu\text{M}$  of compounds 1–12 and CQP.



### 3. Conclusion

In conclusion, we have successfully synthesized a family of morpholinyl-bearing arylsquaramides as small-molecule lysosomal pH modulators. Model and *in vitro* studies show that these squaramides are capable of efficiently mediating the transmembrane transport of chloride anions across liposomal and cellular phospholipid membranes. In contrast to many reported lysosomal pH modulators that generally function by acting as a weak base/acid, or releasing a basic/acidic component in the lysosomal environments, these compounds are able to modulate lysosomal pH *via* a probable mechanism of transmembrane anion transport. The present findings that these compounds are able to alkalize lysosomes without producing any cytotoxicity or disturbance to the autophagy of cells,<sup>51</sup> strongly suggest that they may serve as ideal tool molecules to study the biological function of lysosomes.

### 4. Experimental

#### 4.1 Generals

The instruments used in this study included a Bruker Avance AV 500 NMR spectrometer for measuring <sup>1</sup>H and <sup>13</sup>C NMR spectra; Agilent 1260 Infinity II Prime-Ultivo and Thermo Scientific™ Orbitrap Fusion™ Tribrid™ mass spectrometers for recording the LR and HR ESI-MS spectra, respectively; an Avanti's Mini-Extruder (Avanti Polar Lipids, Inc., Alabaster, Alabama, USA) assembled with nuclepore track-etched polycarbonate membranes (100 nm, Whatman, Florham Park, New Jersey, USA) for vesicle preparation; Mettler-Toledo PerfectIon™ chloride ion selective electrode assembled with a Mettler-Toledo Seven Compact S220 ionometer for measuring chloride efflux; a BioTek Synergy Neo2 microplate reader for conducting MQAE, MQAE-MP and BCECF-AM assays; and a Leica TCS SP8 microscope for carrying out AO and LysoSensor Green DND-189 staining experiments, fluorescein-tetramethylrhodamine-labeled dextran and Magic Red Cathepsin assays.

The specific reagents used in this study were EYPC from Sigma Chemical Co. (St Louis, USA); calcein and acridine orange from J&K Chemical Co (Beijing, China); MQAE from Macklin (Shanghai, China); LysoSensor Green DND-189 from Yeasen (Shanghai, China); Magic Red Cathepsin assay kits from ImmunoChemistry Technologies (Bloomington, USA); BCECF-AM from Dojindo (Shanghai, China); fluorescein-tetramethyl rhodamine-labeled dextran from Invitrogen (Shanghai, China); and CQP from Aladdin (Shanghai, China). Compounds 13–21 and 23 and 24, MQAE-MP were prepared according to reported procedures.<sup>25–28,40</sup> All the other chemicals and reagents were obtained from commercial resources and used as received.

#### 4.2 Synthesis of compound 22

To a mixture of diethyl squarate (151 μL, 1.0 mmol) and Zn(CF<sub>3</sub>SO<sub>3</sub>)<sub>2</sub> (91 mg, 0.25 mmol) in EtOH (1 mL) was added drop wise a solution of 3,5-dichloroaniline (133 mg, 0.82 mmol) in EtOH (3 mL). The resulting mixture was stirred at room temperature and the reaction was monitored by TLC (PE/EA, 2/

1, v/v). 8 h later, the starting material disappeared on TLC and the reaction mixture was concentrated under reduced pressure. The obtained residue was purified by chromatography on a silica gel column, eluted with PE/EA (1/1, v/v) to afford compound 22 (159 mg, 70%) having <sup>1</sup>H NMR (500 MHz, DMSO-*d*<sub>6</sub>) δ 10.96 (s, 1H), 7.46 (s, 2H), 7.31 (t, *J* = 1.8 Hz, 1H), 4.78 (q, *J* = 7.1 Hz, 2H), 1.42 (t, *J* = 7.1 Hz, 3H); <sup>13</sup>C NMR (125 MHz, DMSO-*d*<sub>6</sub>) δ 187.4, 184.3, 179.2, 169.3, 140.7, 134.5, 122.8, 117.7, 70.1, 15.6; negative LR-ESI-MS: *m/z* 284.0 ([M–H]<sup>–</sup>) and HR-ESI-MS for C<sub>12</sub>H<sub>10</sub>Cl<sub>2</sub>NO<sub>3</sub> ([M + H]<sup>+</sup>) Calcd: 286.00322; found: 286.00308.

#### 4.3 Synthesis of compounds 1–12

**Compound 1.** To a mixture of compound 13 (22 mg, 0.1 mmol) and Zn(CF<sub>3</sub>SO<sub>3</sub>)<sub>2</sub> (7.3 mg, 0.02 mmol) in EtOH (3 mL) was added *N*-(2-aminoethyl)morpholine (14 mg, 0.11 mmol). The resulting mixture was stirred at room temperature and the reaction was monitored by TLC (DCM/MeOH, 20/1, v/v). 6 h later, the starting material disappeared on TLC. The reaction mixture was filtered under reduced pressure and the obtained solid was washed twice with methanol to give compound 1 (25 mg, 83%) as a white solid having <sup>1</sup>H NMR (500 MHz, DMSO-*d*<sub>6</sub>) δ 9.81 (s, 1H), 7.66 (s, 1H), 7.44 (d, *J* = 7.8 Hz, 2H), 7.34 (t, *J* = 7.8 Hz, 2H), 7.02 (t, *J* = 7.8 Hz, 1H), 3.73 (br, 2H), 3.58 (br, 4H), 2.52 (br, 2H), 2.42 (br, 4H); <sup>13</sup>C NMR (125 MHz, DMSO-*d*<sub>6</sub>) δ 184.4, 180.5, 169.5, 163.8, 139.3, 129.6, 122.9, 118.3, 66.4, 58.8, 53.5, 40.8; LR-ESI-MS: *m/z* 302.7 ([M + H]<sup>+</sup>) and HR-ESI-MS for C<sub>16</sub>H<sub>20</sub>N<sub>3</sub>O<sub>3</sub> ([M + H]<sup>+</sup>). Calcd: 302.14991; found: 302.14993.

**Compound 2.** Similar procedure as described for compound 1; from compound 14 (60 mg, 0.24 mmol) and *N*-(2-aminoethyl)morpholine (39 μL, 0.29 mmol). Yield: compound 2 (69 mg, 87%). <sup>1</sup>H NMR (500 MHz, DMSO-*d*<sub>6</sub>) δ 9.67 (s, 1H), 7.48 (s, 1H), 7.41–7.34 (m, 2H), 6.96–6.91 (m, 2H), 3.74 (s, 2H), 3.73 (s, 3H), 3.58 (t, *J* = 4.3 Hz, 4H), 2.51 (br, 2H), 2.42 (br, 4H); <sup>13</sup>C NMR (125 MHz, DMSO-*d*<sub>6</sub>) δ 183.6, 180.5, 168.9, 165.1, 163.6, 155.3, 132.3, 120.2, 114.7, 66.2, 58.7, 55.4, 53.4, 40.6; LR-ESI-MS: *m/z* 332.2 ([M + H]<sup>+</sup>) and HR-ESI-MS for C<sub>17</sub>H<sub>22</sub>N<sub>3</sub>O<sub>4</sub> ([M + H]<sup>+</sup>). Calcd: 332.16048; found: 332.16006.

**Compound 3.** Similar procedure as described for compound 1; from compound 15 (60 mg, 0.26 mmol) and 4-(2-aminoethyl)morphine (42 μL, 0.31 mmol). Yield: compound 3 (57 mg, 69%). <sup>1</sup>H NMR (500 MHz, DMSO-*d*<sub>6</sub>) δ 9.72 (s, 1H), 7.56 (s, 1H), 7.32 (d, *J* = 8.3 Hz, 2H), 7.14 (d, *J* = 8.3 Hz, 2H), 3.72 (br, 2H), 3.58 (t, *J* = 4.2 Hz, 4H), 2.51 (br, 2H), 2.42 (br, 4H), 2.25 (s, 3H); <sup>13</sup>C NMR (125 MHz, DMSO-*d*<sub>6</sub>) δ 184.1, 180.5, 168.2, 163.7, 136.7, 131.9, 129.9, 118.4, 103.9, 66.3, 58.8, 53.5, 40.7, 20.5; LR-ESI-MS: *m/z* 316.2 ([M + H]<sup>+</sup>) and HR-ESI-MS for C<sub>17</sub>H<sub>21</sub>N<sub>3</sub>O<sub>3</sub> ([M + H]<sup>+</sup>). Calcd: 316.16556; found: 316.16507.

**Compound 4.** Similar procedure as described for compound 1; from compound 16 (25 mg, 0.11 mmol) and 4-(2-aminoethyl)morphine (17 μL, 0.13 mmol). Yield: compound 4 (26 mg, 73%). <sup>1</sup>H NMR (500 MHz, DMSO-*d*<sub>6</sub>) δ 9.79 (s, 1H), 7.57 (s, 1H), 7.45–7.43 (m, 2H), 7.20–7.17 (m, 2H), 3.72 (br, 2H), 3.58 (t, *J* = 4.2 Hz, 4H), 2.51 (br, 2H), 2.42 (br, 4H); <sup>13</sup>C NMR (125 MHz, DMSO-*d*<sub>6</sub>) δ 184.2, 180.5, 169.3, 163.5, 158.3 (d, *J* = 237.8 Hz), 135.7, 120.1, 116.2 (d, *J* = 22.5 Hz), 66.3, 58.8, 53.5, 40.7; LR-ESI-MS: *m/z*



320.2 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{19}FN_3O_3$  ( $[M + H]^+$ )  
Calcd: 320.14049; found: 320.14014.

**Compound 5.** Similar procedure as described for compound 1; from compound 17 (77 mg, 0.31 mmol) and 4-(2-aminoethyl)morphine (50  $\mu$ L, 0.38 mmol). Yield: compound 5 (98 mg, 94%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  9.87 (s, 1H), 7.62 (s, 1H), 7.45 (d,  $J = 8.6$  Hz, 2H), 7.40–7.38 (m, 2H), 3.73 (br, 2H), 3.58 (t,  $J = 4.0$  Hz, 4H), 2.52 (br, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  184.5, 180.5, 169.5, 163.2, 138.3, 129.4, 126.6, 119.9, 66.3, 58.7, 53.5, 40.8; LR-ESI-MS:  $m/z$  336.2 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{19}ClN_3O_3$  ( $[M + H]^+$ ). Calcd: 336.11094; found: 336.11044.

**Compound 6.** Similar procedure as described for compound 1; from compound 18 (80 mg, 0.27 mmol) and 4-(2-aminoethyl)morphine (44  $\mu$ L, 0.33 mmol). Yield: compound 6 (86 mg, 84%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  9.73 (s, 1H), 7.62 (s, 1H), 7.51 (d,  $J = 8.9$  Hz, 2H), 7.39 (d,  $J = 8.9$  Hz, 2H), 3.72 (br, 2H), 3.58 (t,  $J = 4.0$  Hz, 4H), 2.51 (br, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  184.5, 180.5, 169.5, 163.2, 138.7, 132.3, 120.3, 114.6, 66.3, 58.7, 53.5, 40.8; LR-ESI-MS:  $m/z$  380.0 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{19}BrN_3O_3$  ( $[M + H]^+$ ). Calcd: 380.06043; found: 380.05960.

**Compound 7.** Similar procedure as described for compound 1; from compound 19 (80 mg, 0.33 mmol) and 4-(2-aminoethyl)morphine (53  $\mu$ L, 0.40 mmol). Yield: compound 7 (82 mg, 76%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  10.13 (s, 1H), 7.78 (d,  $J = 8.8$  Hz, 2H), 7.58 (d,  $J = 8.8$  Hz, 2H), 3.73 (br, 2H), 3.58 (br, 4H), 2.52 (br, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.3, 180.3, 170.1, 162.5, 143.5, 140.0, 119.5, 118.3, 103.9, 66.3, 58.7, 53.5, 40.9; negative LR-ESI-MS:  $m/z$  325.8 ( $[M - H]^-$ ) and HR-ESI-MS for  $C_{17}H_{19}N_4O_3$  ( $[M + H]^+$ ). Calcd: 327.14516; found: 327.14468.

**Compound 8.** Similar procedure as described for compound 1; from compound 20 (69 mg, 0.26 mmol) and 4-(2-aminoethyl)morphine (27  $\mu$ L, 0.20 mmol). Yield: compound 8 (62 mg, 90%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  10.29 (s, 1H), 8.23 (d,  $J = 9.1$  Hz, 2H), 7.81 (s, 1H), 7.61 (d,  $J = 9.1$  Hz, 2H), 3.74 (br, 2H), 3.58 (br, 4H), 2.52 (t,  $J = 6.0$  Hz, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.6, 180.3, 170.4, 162.3, 145.6, 141.6, 125.9, 117.9, 66.4, 58.6, 53.5, 40.9; LR-ESI-MS:  $m/z$  347.7 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{19}N_4O_5$  ( $[M + H]^+$ ). Calcd: 347.13499; found: 347.13391.

**Compound 9.** Similar procedure as described for compound 1; from compound 21 (77 mg, 0.25 mmol) and 4-(2-aminoethyl)morphine (37  $\mu$ L, 0.28 mmol). Yield: compound 9 (86 mg, 87%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  10.05 (s, 1H), 7.72 (s, 1H), 7.69 (d,  $J = 8.7$  Hz, 2H), 7.61 (d,  $J = 8.7$  Hz, 2H), 3.74 (br, 2H), 3.58 (br, 4H), 2.51 (t,  $J = 6.1$  Hz, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.1, 180.4, 169.9, 163.0, 142.9, 126.9, 126.9, 124.7 (q,  $J = 269.4$  Hz), 122.5 (q,  $J = 31.7$  Hz), 118.2, 66.4, 58.7, 53.5, 40.9; LR-ESI-MS:  $m/z$  370.7 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{17}H_{19}F_3N_3O_3$  ( $[M + H]^+$ ). Calcd: 370.13730; found: 370.13623.

**Compound 10.** Similar procedure as described for compound 1; from compound 22 (80 mg, 0.28 mmol) and 4-(2-aminoethyl)morphine (45  $\mu$ L, 0.34 mmol). Yield: compound 10 (85 mg, 82%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  10.03 (s, 1H), 7.69 (s, 1H), 7.48 (s, 2H), 7.20 (s, 1H), 3.73 (br, 2H), 3.57 (br, 4H), 2.52 (br, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.0,

180.4, 169.9, 162.6, 141.8, 134.9, 121.5, 116.6, 66.3, 58.7, 53.5, 40.8; LR-ESI-MS:  $m/z$  370.6 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{18}Cl_2N_3O_3$  ( $[M + H]^+$ ). Calcd: 370.07197; found: 370.07153.

**Compound 11.** Similar procedure as described for compound 1; from compound 23 (77 mg, 0.25 mmol) and 4-(2-aminoethyl)morphine (37  $\mu$ L, 0.28 mmol). Yield: compound 11 (67 mg, 68%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ /TFA- $d$ , 19/1, v/v)  $\delta$  11.17 (s, 1H), 8.71 (s, 2H), 8.57 (s, 1H), 8.41 (t,  $J = 1.8$  Hz, 1H), 4.00 (d,  $J = 5.9$  Hz, 4H), 3.68 (t,  $J = 11.6$  Hz, 2H), 3.52 (d,  $J = 11.6$  Hz, 2H), 3.44–3.42 (m, 2H), 3.17 (d,  $J = 10.0$  Hz, 2H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.2, 181.0, 170.2, 163.1, 149.0, 141.9, 117.9, 116.8, 114.5, 111.1, 63.6, 63.4, 56.4, 51.8, 51.6, 38.2; LR-ESI-MS:  $m/z$  392.2 ( $[M + H]^+$ ) and HR-ESI-MS for  $C_{16}H_{18}N_5O_7$  ( $[M + H]^+$ ). Calcd: 392.12007; found: 392.11890.

**Compound 12.** Similar procedure as described for compound 1; from compound 24 (50 mg, 0.14 mmol) and 4-(2-aminoethyl)morphine (15  $\mu$ L, 0.11 mmol). Yield: compound 12 (44 mg, 91%).  $^1H$  NMR (500 MHz, DMSO- $d_6$ )  $\delta$  10.32 (s, 1H), 8.05 (s, 2H), 7.72 (s, 1H), 7.66 (s, 1H), 3.75 (br, 2H), 3.58 (br, 4H), 2.51 (t,  $J = 6.1$  Hz, 2H), 2.42 (br, 4H);  $^{13}C$  NMR (125 MHz, DMSO- $d_6$ )  $\delta$  185.1, 180.6, 170.0, 162.5, 141.5, 131.6 (q,  $J = 33.3$  Hz), 123.4 (q,  $J = 271.3$  Hz), 118.3, 114.9, 66.4, 58.6, 53.5, 40.9; negative LR-ESI-MS:  $m/z$  436.8 ( $[M - H]^-$ ) and HR-ESI-MS for  $C_{18}H_{18}F_6N_3O_3$  ( $[M + H]^+$ ). Calcd: 438.12468; found: 438.12369.

#### 4.4 Measurement of anion transport, lysosomal targeting and alkalization

Literature procedures were adopted to carry out the chloride efflux;<sup>29,30,34,52</sup> calcein leakage;<sup>30,52</sup> U-tube experiments;<sup>30,35</sup> MQAE and MQAE-MP assays;<sup>40,49,52</sup> AO assays;<sup>41,42,52</sup> LysoSensor Green DND-189 assays;<sup>23,44,45</sup> BCECF-AM assays;<sup>47</sup> dextran, fluorescein and tetramethylrhodamine assays<sup>38,42</sup> and Magic Red Cathepsin assays.<sup>23,42</sup> Detailed experimental procedures may be found in the ESI.†

## Conflicts of interest

There are no conflicts to declare.

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