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Influence of ITO electrode on the electrochromic performance outcomes of viologen-functionalized polyhedral oligomeric silsesquioxanes†

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Electrochromic devices (ECDs) exhibit reversible optical changes under applied electrical stimuli. Transparent conducting electrodes (TCOs), generally constructed with indium tin oxide (ITO), are a vital component determining transparency and switching behaviors. ITO specifications for TCO materials have not drawn much attention despite the critical role of these materials. Herein we investigate the influence of ITO electrodes in achieving high-performance ECDs containing viologen-functionalized polyhedral oligomeric silsesquioxane (POSS-viologen). Indeed, ITO electrodes exert significant effects on the electrochromic characteristics. A high ITO thickness shows superior color-switching with high optical density and coloration efficiency levels. Enhanced electrical conductivity facilitates diffusion behaviors, an outcome beneficial for electrochromic switching. The surface-charge capacity ratio values are measured and found to be close to one, indicating that no residual current remains, and the prepared devices provide good reversibility during the coloring and bleaching process. Furthermore, with an increase in the ITO thickness, the current required for the coloring and bleaching processes decreases, and the power consumption needed for the operation of the device becomes low. The superiority of POSS-viologen should also be noted, especially when compared to normal viologens, in terms of the electrochromic properties and long-term operational stability. These results demonstrate the critical role of electrical conductivity in ITO electrodes, providing a valuable guideline for TCO specifications for ECD fabrication using viologen derivatives.

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1. Introduction

Given the increased demand for information-providing devices, electrochromic devices (ECDs) have been developed as a promising candidate to fulfil such requests due to their low-voltage operation, simple fabrication process, and user-controllable properties.^{1–3} ECDs can be intelligent windows to protect privacy while blocking harmful ultra-violet radiation.⁴ Applications have recently been extended to antiglare mirrors, camouflage, and e-paper displays in the industrial market.⁵ When incorporated into architectural windows, ECDs deliver user-oriented information efficiently while eliminating the need for curtains and blinds to reduce indoor lighting and air conditioning loads. ECDs are electrochemical cells in which the optical properties of redox materials change reversibly under suitable electrical stimuli.⁶ The corresponding absorption bands are identified as color changes from transparent to

colored states.^{7,8} A wide variety of electrochromic materials have been developed, some of which are conjugated polymers,⁹ metal oxides,¹⁰ and viologens.¹¹ Viologens are popular cathodic materials that undergo a reduction when bias is applied. Various derivatives have been studied in-depth due to their desirable electron-accepting capabilities, stable redox states, and bias tunability of their nitrogen substituents.¹² However, viologens have critical drawbacks that prevent their practical applications, such as low coloration efficiency, slow switching speeds, and poor long-term stability.¹³

Transparent conducting electrodes (TCOs) are another vital component determining transparency and switching behaviors when constructing ECDs.^{14,15} Indium tin oxide (ITO) has been widely adopted as a TCO material due to its low visible contrast and high electrical conductivity, finding use in many applications in optoelectronics, sensors, and solar cells.^{16,17} However, the effects of the oxide network and surface roughness on the resulting optical contrast have not been studied often.^{18,19} ITO degrades under repeated anodic or cathodic polarization, slowly deteriorating redox performance capabilities.^{20,21} Numerous approaches have been presented to improve the electrochromic performances of viologen-based ECDs. Most reports emphasize the composition of the active layer and how it induces the charge transfer and coloration/bleaching process.^{22,23} Despite

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the crucial impacts of TCOs when in direct contact with the active layer, not much attention has been directed toward the specifications of ITO electrodes in viologen-based ECDs.

Our previous reports demonstrated that electron-donating/withdrawing substituents, counter anions, asymmetric structures, and nanofiber structures improved electrochromic transitions.^{24–27} Furthermore, viologen-functionalized polyhedral oligomeric silsesquioxane (POSS-viologen) demonstrated outstanding electrochromic performance capabilities.^{28,29} The superb electrochromic performances were attributed to the structural rigidity of the POSS fillers and the synthetic versatility of the viologen compounds.

Motivated by previous works, in the current study, we investigate the influence of ITO electrodes in achieving high-performance ECDs containing a POSS-viologen. ITO-coated glass specimens of three different thicknesses were adopted for the electrodes, and single-layered ECDs were prepared using an ion gel composition. The effects of the ITO electrodes on the electrochromic characteristics were examined by estimating the surface charge capacity, operating voltage, diffusion coefficients, power consumption, and coloration/bleaching times. In doing so, the critical role of electrical conductivity in the electrodes and the superiority of the POSS-viologens were verified, especially when compared to normal viologens, in terms of the electrochromic transitions and long-term operational stability.

2. Experimental

2.1. Materials

Poly(vinylidene fluoride-*co*-hexafluoropropylene) (PVDF-HFP), 1-butyl-3-methylimidazolium tetrafluoroborate ($[\text{BMIM}][\text{BF}_4]$), and 1-bromohexane were purchased from Sigma Aldrich and were used as received. (3-Chloropropyl)trimethoxysilane, 4,4'-bipyridine, ferrocene (Fc), and ammonium hexafluorophosphate (NH_4PF_6) were purchased from Alfa Aesar. Dimethylformamide, acetonitrile, methanol, acetone, and diethyl ether were obtained commercially and used without further purification. Indium tin oxide (ITO) coated glass specimens with thicknesses of 68.5 nm, 140.3 nm, and 252.1 nm were purchased from Wooyang GMS (Korea). Before use, the ITO-coated glasses were cleaned successively with deionized water, acetone, and isopropyl alcohol under sonication.

2.2. Synthesis

Octa-hexyl viologen-substituted polyhedral oligomeric silsesquioxane, $[\text{OHV-POSS}][\text{Br}]_8[\text{Cl}]_8$, was synthesized according to a procedure in the literature.²⁴ The anion-exchanged species, $[\text{OHV-POSS}][\text{PF}_6]_{16}$, was obtained after a reaction with NH_4PF_6 in methanol and was employed as the electrochromic species for the ion gel preparation process. Di-hexyl viologen, $[\text{DHV}][\text{PF}_6]_2$, was similarly synthesized for comparison purposes.²⁸

2.3. Property characterizations

The surface roughness (R_a) and morphology were evaluated using an atomic force microscope (NX10, Park systems) and a scanning electron microscope (FE-SEM, Zeiss SUPRA25),

respectively. The electrical conductivity values of the ITO-coated glass slides were evaluated using a four-point probe system (M4P302 and Source Meter, MS-TECH, Keithley). The thickness of the ITO film was confirmed using a KLA instrument (Tencor, D-500). The electrochromic properties were measured in single-layered solid-state ECDs. The ion gel solution was prepared by mixing $[\text{OHV-POSS}][\text{PF}_6]_{16}$ (300 mg) as an electroactive species, methyl-ferrocene (10 mg) as an anodic species, PVDF-HFP (400 mg) as a polymer matrix, and $[\text{BMIM}][\text{BF}_4]$ (1600 mg) as a room-temperature ionic liquid, in acetone. The solution was stirred at 50 °C for 24 h to form a homogeneous ion gel. The ion gel was coated onto the ITO-coated glass slides by a drop-casting method with a working area of $4 \times 5 \text{ cm}^2$. Then, Surlyn® film with a thickness of 60 nm was applied. The deposited ion gel was kept at 70 °C for 1 h to remove the acetone, and the formed film was sandwiched using another ITO-coated glass slide, followed by sealing with epoxy resin. By the same procedures, a single-layered ECD with $[\text{DHV}][\text{PF}_6]_2$ was prepared.

Optical and spectroelectrochemical measurements were taken using a UV-1800 UV-vis spectrophotometer (Shimadzu). The electrochemical and spectroelectrochemical properties were measured in a two-electrode system using an SP-150 potentiostat (Biologic). The kinetic stability was estimated by applying constant potential stepping with a potentiostat and monitoring the corresponding transmittance changes with a UV-vis spectrophotometer. The transmittance difference (ΔT), the change in the optical density (ΔOD), and the coloration efficiency (CE) were calculated using the following relationships:³⁰ $\Delta T = T_{\max} - T_{\min}$, $\Delta \text{OD} = \log \frac{T_b}{T_c}$, and $\text{CE} = \frac{\Delta \text{OD}}{Q_d}$, respectively, where T_b and T_c correspondingly represent the bleached and colored transmittance values, and Q_d denotes the injected charge per unit area.

For ionic conductivity measurement, cells were fabricated by placing the prepared ion gel solution on an ITO-coated glass with a long groove width of 100 μm and then sandwiching it with a cover glass using a Surlyn film spacer. The Nyquist impedance graphs were obtained by plotting $\text{Re}(Z)/\text{Ohm}$ vs. $\text{Im}(Z)/\text{Ohm}$, from which the ionic conductivity was calculated according to the Nyquist equation.³¹

3. Results and discussion

The surface morphology of ITO-coated glass slides of different ITO thicknesses was examined by scanning electron microscopy (SEM) measurements (Fig. 1a–c). The SEM images showed a smooth surface consisting of a network of high-porosity interconnected nanoparticles, which may improve ionic mobility and charge transfer.^{32,33} The surface roughness of the ITO glass specimens was determined by an atomic force microscopy (AFM) analysis (Fig. 1d–f). The AFM images revealed an island-like surface structure, which became smaller with an increase in the ITO thickness. The glass substrates with ITO thicknesses of 68.5, 140.3, and 252.1 nm showed surface roughness (R_a) values of 10.1, 6.40, and 6.18 nm, respectively. It should be noted that the sheet resistance decreased monotonically with an increase in the ITO thickness (Table 1). With



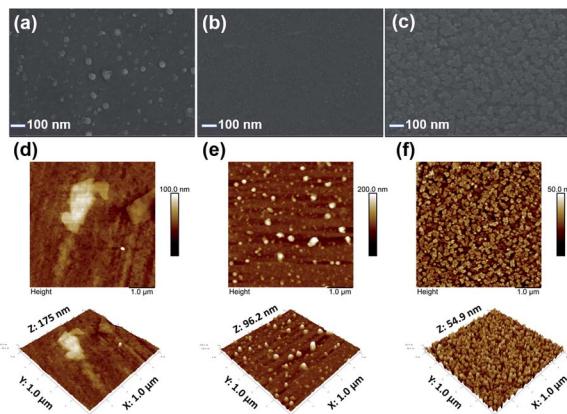


Fig. 1 SEM images of the ITO films having thicknesses of (a) 68.5 nm (ITO-I), (b) 140.3 nm (ITO-II), and (c) 252.1 nm (ITO-III). The corresponding AFM images are shown in (d)–(f).

the increase of ITO thickness, the surface roughness of glass decreased, which helped to provide more electrical conductivity and lower the sheet resistance.

Before evaluating the effects of the ITO electrodes, the ionic conductivity values of ion gel compositions made up of [OHV-POSS] $[\text{PF}_6]_{16}$ were estimated and compared with those of [DHV] $[\text{PF}_6]_{12}$ (Table S1, ESI[†]). The electrochemical impedance spectroscopy (EIS) results found ionic conductivity values of $6.24 \times 10^{-3} \text{ S cm}^{-1}$ and $3.63 \times 10^{-3} \text{ S cm}^{-1}$ for OHV-POSS and DHV, respectively (Fig. S1, ESI[†]). Higher ionic conductivity of OHV-POSS was observed because POSS acts as a nano-building block, facilitating the ion-transport ability of ionic liquids and leading to a significant increase in the ionic conductivity.^{34,35}

The effect of the ITO thickness on the electrochemical was examined by obtaining cyclic voltammetry (CV) curves with various scan rates (Fig. 2). The ECD containing the OHV-POSS-based ion gel (OHV-POSS ECD), fabricated with ITO-I, required relatively high voltage operation at -3.0 V (Fig. 2a). On the other hand, the OHV-POSS ECDs with ITO-II and ITO-III showed lower operation voltages of -2.3 V and -1.9 V , respectively (Fig. 2b and c). The observed difference is attributed to the surface resistance and electrical conductivity of the ITO electrodes. A thin ITO thickness in ITO-I led to the high sheet resistance of 26.8Ω and a high R_a value of 10.1 nm . In comparison, the thicker ITO layers in ITO-II and ITO-III featured low sheet resistance values of $11.9 \Omega \square^{-1}$ and $7.27 \Omega \square^{-1}$, respectively, and smooth surfaces with low corresponding R_a values of 6.40 nm and 6.18 nm .

With an increase in the scan rate, both the anodic (I_{pa}) and cathodic (I_{pc}) peak currents increased, showing linear

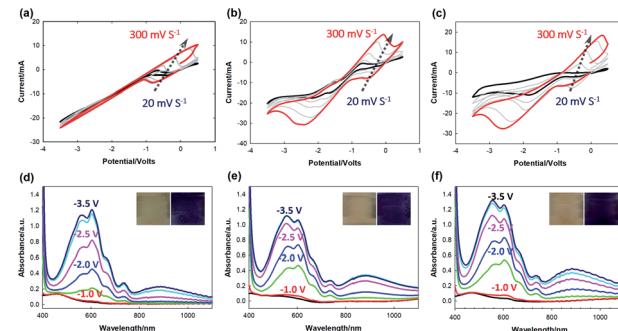


Fig. 2 Cyclic voltammetry curves with different scan rates of 20, 50, 100, 200, and 300 mV s^{-1} for [OHV-POSS] $[\text{PF}_6]_{16}$ -based ECDs with ITO thicknesses of (a) 68.5 nm, (b) 140.3 nm, and (c) 252.1 nm. The corresponding linear plots of I_{pa} and I_{pc} with the applied voltage are shown in the insets. Spectroelectrochemical curves obtained with different applied potentials for [OHV-POSS] $[\text{PF}_6]_{16}$ containing ion gel-based ECDs prepared using electrodes with ITO thicknesses of (d) 68.5 nm, (e) 140.3 nm, and (f) 252.1 nm.

dependency on the square root of the scan rate. According to the Randle–Sevcik equation, this behavior corresponded to a typical one-electron transfer process, which is diffusion-controlled.³⁶ From the above relationship, the diffusion coefficient values (D) depending on the ITO thickness were evaluated according to the following equation:³⁷

$$D = \frac{i_p^2 RT}{(0.4463nCAF)^2 nFV}$$

Here, n is the number of electrons, assumed here to be 1; i_p is the peak current (A); D is the diffusion coefficient of the anion ($\text{cm}^2 \text{ S}^{-1}$); A is the electrode area (cm^2); C is the concentration of active ion (mol cm^{-3}); V is the potential scan rate (V s^{-1}), F is the Faraday constant ($96\ 485 \text{ C mol}^{-1}$); R is the gas constant; T is the absolute temperature.

For the sake of convenience, the peak current values at a scan rate of 200 mV were extracted from the CV curves (Fig. S2, ESI[†]). ITO electrodes heavily impacted the diffusion behaviors of the OHV-POSS ECDs. With an increase in the ITO thickness, D on the cathodic side increased. The OHV-POSS ECDs prepared with ITO-III exhibited the highest D values of $9.30 \times 10^{-7} \text{ cm}^2 \text{ S}^{-1}$ during the reduction process (Table 2). The current flow process features an electrode reaction and charge transport through the polymeric ion gel film in the single-layered device. Clearly, efficient charge diffusion processes are necessary for smooth charge transport. In this regard, high D values are preferred for electrochromic switching, as the diffusion coefficient mainly

Table 1 Summarized electrical and surface properties of ITO glass specimens

ITO glass	ITO coating thickness (nm)	Sheet resistance ($\Omega \square^{-1}$)	Surface roughness (R_a)
ITO-I	68.5	26.8	10.1
ITO-II	140.3	11.9	6.40
ITO-III	252.1	7.27	6.18

Table 2 Diffusion coefficients of OHV-POSS-based ECDs prepared with various ITO thicknesses

ITO glass	Cathodic current I_{pc} (A)	D from I_{pc} ($\text{cm}^2 \text{ S}^{-1}$)
ITO-I	-2.04×10^{-2}	5.39×10^{-7}
ITO-II	-2.38×10^{-2}	7.31×10^{-7}
ITO-III	-2.68×10^{-2}	9.30×10^{-7}



governs the mass transfer process for electro-neutralizing ions during the electrochemical exchange reaction.

During the electrochromic switching redox process, charge transport is driven by an electron self-exchange reaction between adjacent redox sites. It is also dominated by charge diffusion based on the redox gradient produced under the electrolyte conditions. The steady-state redox gradient through the sandwiched electrochromic layer is maximized when the anode potential is more positive, and the cathode potential is more negative than $E_{1/2}$ of the sandwiched layer.

The redox gradient provides the diffusion-limited current density J_{lim} obtained for the sandwiched device according to the following equation:³⁹

$$J_{\text{lim}} = \frac{\omega^0 n F D C_T}{d}$$

In this equation, C_T is the concentration of the pendant redox sites, d is the film thickness, and ω^0 is the correction factor for electrostatic coupling between electron and counterion motions. The diffusion coefficient D reflects the mass transfer process for the electro-neutralizing ions during the

electrochemical exchange reaction in the layer. The observed J_{lim} for OHV-POSS ECDs prepared with ITO 252.1 nm ($-26.89 \text{ mA cm}^{-2}$) coincided with the theoretical value for these devices ($-26.83 \text{ mA cm}^{-2}$) (Fig. S3, ESI†).

The effects of the ITO thickness on kinetic behaviors were studied by applying alternating wave potentials between bleached and colored states (Fig. 3). During long-term stability measurements, the OHV-POSS ECDs with ITO-I exhibited a ΔT of 56.7% with an ΔOD of 0.575, producing a CE outcome of $149.4 \text{ cm}^2 \text{ C}^{-1}$. The switching times required for the colored and bleached states were 12 s and 9 s, respectively (Fig. 3a and d). As the ITO thickness was increased to 140.3 nm and 252.1 nm, the ΔT increased to 62.3% and 75.5%, respectively, with corresponding CE values of $172.08 \text{ cm}^2 \text{ C}^{-1}$ and $324.1 \text{ cm}^2 \text{ C}^{-1}$ (Fig. 3b and c). The switching times required for the colored and bleached states were almost unchanged (Fig. 3e and f). Regardless of the ITO thickness, the OHV-POSS ECDs demonstrated excellent cyclic stability for $>12\,000$ s with no transmittance losses. Thicker ITO electrodes clearly exhibited better electrochromic performances, both in terms of optical density and coloration efficiency. These outcomes are attributed to the low sheet resistance.

The electrochemical and kinetic behaviors of a normal violagen, $[\text{DHV}][\text{PF}_6]_2$, containing ECDs (DHV ECDs) with different ITO thicknesses were also examined for comparison. DHV ECDs began coloration to blue at -1.1 V with λ_{max} at 605 nm (Fig. S4, ESI†). These samples showed slow coloring and bleaching processes with poor cyclic stability and low CE values (Fig. S5, ESI†). The obtained electrochromic properties are summarized and compared in Table 3.

The effect of the ITO thickness on the surface charge capacity ratio (r) was estimated from the intercalation surface charge (Q_{in}) and extraction surface charge (Q_{out}) according to the following equation²³ (Fig. 4a):

$$r = \frac{Q_{\text{in}}}{Q_{\text{out}}}$$

Here, r is defined upon the insertion of the surface charge divided by the extraction of surface charge. For the OHV-POSS ECDs fabricated with ITO thicknesses of 68.5 nm, 140.3 nm, and 252.1 nm, the corresponding intercalated charge densities (Q_{in}) are -3.51 , -1.91 , and -1.30 mQ cm^{-2} and the extracted

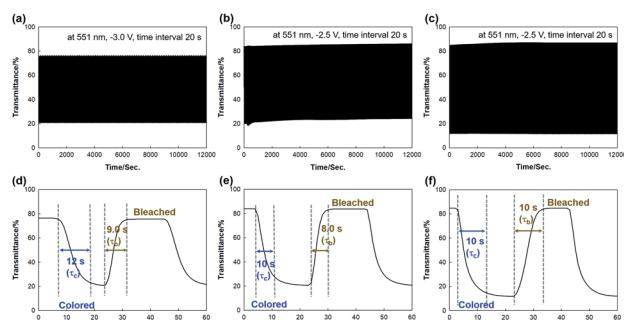


Fig. 3 Kinetic stability for $[\text{OHV-POSS}][\text{PF}_6]_{16}$ -based ECDs with ITO thicknesses of (a) 68.5 nm, (b) 140.3 nm, and (c) 252.1 nm during constant potential stepping between (coloration) and short-circuit (bleaching) conditions. The corresponding switching times required for coloration and bleaching of the ECDs are shown in (d)–(f). Resistance for faster electron movement. Moreover, it is also apparent that the OHV-POSS-based ion gel possessed high ionic conductivity, leading to a quicker charge transfer process.

Table 3 Summarized electrochromic properties of OHV-POSS-based ECDs in comparison with DHV ECDs

Electrochromic material	ITO glass	Transmittance difference ^a (%)	Optical density (OD)	Switching time to colored state ^b (s)	Switching time to bleached state ^c (s)	Coloration efficiency
[OHV-POSS][PF ₆] ₁₆	ITO-I	56.7	0.575	12	9.0	149.4
	ITO-II	62.3	0.561	10	8.0	172.1
	ITO-III	75.5	0.868	10	10	324.1
[DHV][PF ₆] ₂	ITO-I	4.41	0.751	20	—	14.47
	ITO-II	19.4	0.364	20	—	26.30
	ITO-III	17.05	1.42	20	—	126.9

^a The transmittance difference was calculated by the difference in maximum to minimum transmittance change (%). ^b Time required for percent transmittance changes from minimum to maximum. ^c Time required for percent transmittance changes from maximum to a minimum; applied voltage ranges from 0 to -3 V for all ECDs.



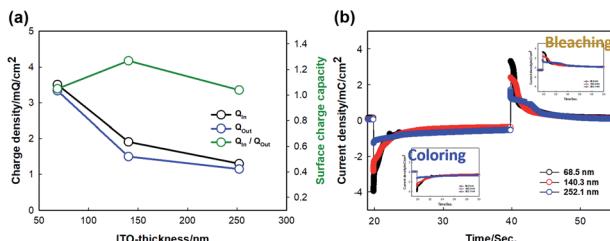


Fig. 4 (a) ITO thickness-dependent surface capacity for [OHV-POSS][PF₆]₁₆-based ECDs as a function of Q_{in} and Q_{out} . (b) Current density response of [OHV-POSS][PF₆]₁₆-based ECDs with ITO thicknesses of 68.5 nm, 140.3 nm, and 252.1 nm. Corresponding current density responses for coloration and bleaching are shown in the inset.

charge densities (Q_{out}) are 3.34, 1.50, and 1.15 mQ cm^{-2} . The calculated charge capacity ratio (r) values were estimated to be 1.05, 1.27, and 1.04, respectively. It should be noted that the r values are close to one, indicating that all devices exhibited good electrochemical reversibility during the coloring and bleaching process.

Fig. 4b shows the chronoamperometric results when applying bias ($-2.5 \text{ V}/+0.5 \text{ V}$). During intercalation with a coloring time from 20 to 30 s, the amounts of charge of the current density were -3.51 , -1.91 , and -1.30 mC cm^{-2} . The corresponding charge amounts with deintercalation with bleaching times from 40 to 50 s were 3.34, 1.50, and 1.15 mC cm^{-2} . These results indicate that with an increase in the ITO thickness, the current required for the coloring and bleaching processes is decreased.

Table 4 Calculated power consumption values for OHV-POSS-based ECDs prepared with various ITO thicknesses

ITO-thickness (nm)	Cathodic J_{mean} (mA cm^{-2})	Anodic J_{mean} (mA cm^{-2})	Applied voltage (V)	Power consumption ($\mu\text{W cm}^{-2}$)	
				Cathodic	Anodic
68.5	0.175	0.167	$-3.0/+0.5$	525.0	83.5
140.3	0.0955	0.075	$-2.5/+0.5$	238.7	37.5
252.1	-0.0651	0.0575	$-2.5/+0.5$	162.5	28.7

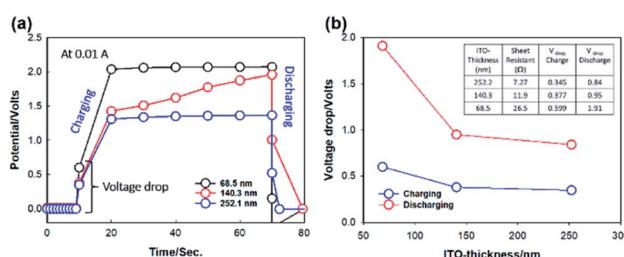


Fig. 5 (a) Galvanometric charge/discharge curves measured after applying a current of 0.01 A for [OHV-POSS][PF₆]₁₆-based ECDs with ITO thicknesses of 68.5 nm, 140.3 nm, and 252.1 nm. The corresponding voltage drops during the charging and discharging processes are shown in (b).

The power consumption (P) for OHV-POSS ECD operation was calculated from the relationship of $P = VJ_{\text{mean}}$, where V is the applied voltage and J_{mean} is the average current density when voltage is applied to the device. J_{mean} is obtained from the

relationship $J_{\text{mean}} = \frac{\int_{t_0}^{t_1} J dt}{(t_1 - t_0)}$, where J is the current density during the ECD time of t_0 to t_1 .^{11,38} The J_{mean} values of OHV-POSS-ECD with ITO-III at -2.5 V and $+0.5 \text{ V}$ were estimated to be -0.065 and $0.0575 \text{ mA cm}^{-2}$, respectively. The corresponding power consumption outcomes were $162.5 \mu\text{W cm}^{-2}$ and $28.7 \mu\text{W cm}^{-2}$. On the other hand, the devices with ITO-I and ITO-II exhibited higher power consumption values; the calculated values are summarized in Table 4.

An instantaneous voltage drop may occur after switching on the current during ECD operation.⁴⁰ This behavior is attributed to the internal cell resistance related to the electrode, electrolyte, and solid-solid and solid-liquid interfaces. In electrolytic media, a voltage drop often occurs due to current leakage without any chemical processes.⁴¹ Thus, the spontaneous voltage drop represents an energy loss, which should be avoided to ensure acceptable performance. Furthermore, the voltage drop at the beginning of the discharge process may cause the ECD to supply voltage exceeding that required by the external load, which is detrimental for the long-term operation. As a measure to evaluate the ITO effect, the voltage drop was estimated by charge/discharge measurements at an applied current of 0.01 A with applied reduction potential bias (Fig. 5a). It was observed that with an increase in the ITO thickness, the voltage drop in the OHV-POSS ECDs is decreased (Fig. 5b). The ECD with ITO-III nm showed the lowest voltage drop due to the more rapid charge insertion and extraction process. With relatively thick ITO, the sheet resistance is reduced, exhibiting high electrical conductivity, which helps to keep the current density uniform during the electrochemical reaction process.

4. Conclusion

Here, we studied the effects of the ITO thickness on the electrochromic performance capabilities of electrochromic devices (ECDs) containing viologen-functionalized polyhedral oligomeric silsesquioxanes (OHV-POSS). For this purpose, three ITO glass specimens with thicknesses of 68.5, 140.3, and 252.1 nm were chosen, and their surface properties were studied using morphological characterization techniques. The ITO electrode plays a vital role in determining the ECD performance outcome. The electrode with relatively thick ITO showed low sheet resistance and low surface roughness, which facilitated the superior electrochemical performance of OHV-POSS ECDs. Compared to normal viologen-containing ECDs, the OHV-POSS ECDs exhibited low voltage operation, high transmittance, and good long-term cyclic stability. In addition, the OHV-POSS ECDs featured a low voltage drop, low power consumption, and a high diffusion coefficient. These results suggest that it is necessary to use electrode with relatively thick ITO as it exhibits low sheet resistance of less than $10 \Omega \square^{-1}$ to achieve superior electrochromic performance outcomes in viologen-based ECDs. The



proposed ITO specifications may be useful for the fabrication of high-performance and durable electrochromic devices containing [OHV-POSS][PF₆]₁₆ for smart window and display applications.

Conflicts of interest

There are no conflicts to declare.

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