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# Reactions of [60]fullerene with alkynes promoted by OH<sup>−</sup>†

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In the current work, the reactions of [60]fullerene with alkynes promoted by OH<sup>−</sup> (base) are addressed. The treatment of C<sub>60</sub> with alkynes in the presence of TBAOH produces alkynylation products (R-C<sub>60</sub>-H) with high selectivity in *o*-DCB at 100 °C. Plausible reaction mechanisms were proposed. This work provides a convenient and environmental friendly method for the functionalization of fullerenes.

Fullerenes functionalized with alkynyl groups are particularly valuable in materials science and organic chemistry due to their excellent electrochemical properties, among others. The combination of fullerene and acetylene units is not only a potential precursor for developing novel fullerene-based dimers or polymers,<sup>1,2</sup> but also can promote the conjunction pathway with a supramolecular framework providing new properties. For example, it has been reported that fullerene covalently linked to acetylenic polythiophene produces unique photoelectronic properties.<sup>3,4</sup> On the other hand, alkynes are considered as one of the most outstanding building blocks for unsaturated molecular scaffolds in organic synthesis, owing to their transformation potential into various functional groups.<sup>5–8</sup>

Up to now, the easiest method to achieve acetylene derivatives of C<sub>60</sub> is based on reactions of fullerene with alkynyllithium. Komatsu *et al.* reported the synthesis, characterization and properties of the first acetylene derivative of C<sub>60</sub> containing a trimethylsilyl ethynyl group or a phenylethynyl group.<sup>9,10</sup> Diedrich *et al.* prepared a series of alkynylfullerenes using the same method.<sup>1,11</sup> However, the use of organolithium reagents led to complex experimental conditions, due to them being extremely sensitive to air and moisture.<sup>12</sup> Hence, new methods for the preparation of alkynylfullerene compounds are necessary, reducing the steps involved in the synthesis, and

responding to new regulations associated with sustainability. In this sense, these new procedures would not only be efficient, selective, high-yielding, but also easy to operate and environmentally friendly.

In this work, our group deeply worked on the new functionalized chemistry of fullerenes with alkynes and found that the nucleophilic carbon functionalization of C<sub>60</sub> with alkynes could be mediated by OH<sup>−</sup>. In this sense, TBAOH (tetra-*n*-butylammoniumhydroxide) was used as a source of OH<sup>−</sup>, because it is cheap, environment friendly, and allows getting a high efficiency in the transformations. This new process allows getting highly functionalized alkynylfullerene derivatives through a one-step reaction. Furthermore, the protocol is simple, and works with a wide range of alkynes. Interestingly, different types of alkynes undergo diverse reaction pathways.

A series of phenylacetylene compounds were chosen as model substrates to optimize the reaction. The experimental procedure for the new nucleophilic reaction is very simple: a mixture composed by C<sub>60</sub> (36 mg) and a series of alkynes (20 equiv.) was stirred in *o*-DCB (*o*-dichlorobenzene) with a solution 1.0 mol L<sup>−1</sup> of TBAOH/CH<sub>3</sub>OH (150 μL, 3 equiv.), as the OH<sup>−</sup> source, under argon gas for 1 h at 100 °C. The colour of the reaction gradually changed from purple to dark green after the addition of TBAOH/CH<sub>3</sub>OH. It was observed that the use of an excess TBAOH or longer reaction time did not improve the yield because of the formation of polar side products, like multiple addition compounds.<sup>13</sup> Subsequently, a trifluoroacetic acid treatment, produced the different alkynylfullerenes **2a–2d** with 1,2-addition and the yield is 35.2%, 29%, 21.1%, 31.5%, respectively (listed in Table 1) after separation by Buckyprep column. The trifluoroacetic acid is working as the donor of protons. The 1,4-substituted phenylacetylene with both electron-donating and electron-withdrawing groups on the phenyl could be used in the reaction as it can be observed in Table 1.

The structures of 1,2-addition C<sub>60</sub> compounds **2a–2d** were fully characterized and elucidated by their HRMS, <sup>1</sup>H NMR, <sup>13</sup>C

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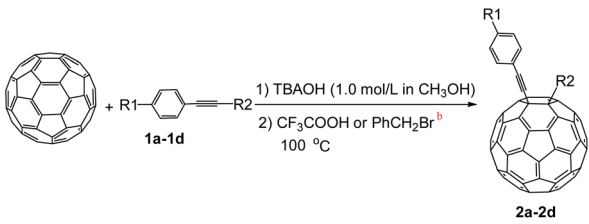
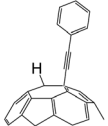
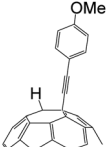
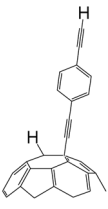
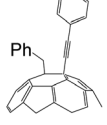
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Table 1 Alkynylation of C<sub>60</sub><sup>a</sup>

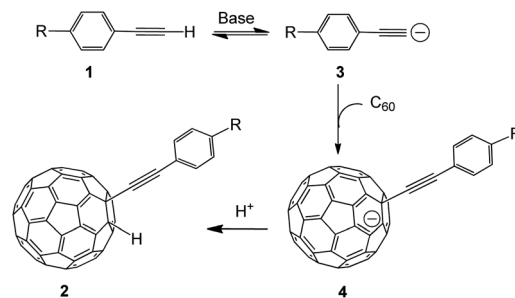
					
Entry	1	R1	R2	2a-2d <sup>b</sup>	Yield (%) <sup>c</sup>
1	1a	H	H		35.2
2	1b	OMe	H		29.0
3	1c	CCH	H		21.1
4	1d	H	CO <sub>2</sub> Et		31.5

<sup>a</sup> All reactions were performed in *o*-DCB using an oil bath at 100 °C. Molar ratio of compound 1 : C<sub>60</sub> = 20 : 1. <sup>b</sup> CF<sub>3</sub>COOH was used in the reaction of C<sub>60</sub> with 1a–1c. While PhCH<sub>2</sub>Br was used in the case of C<sub>60</sub> with 1d. <sup>c</sup> Isolated yield.

NMR, and UV-vis spectra. The products 2a–2c exhibited proper molecular weights in their mass spectra. The <sup>13</sup>C NMR spectra for 2a–2c clearly exhibited no more than 29 peaks in the range of 134–152 ppm, associated with the sp<sup>2</sup>-carbons of the fullerene cage, which is consistent with the *Cs* symmetry of their molecule structure, and the peak placed at 54–55 ppm for the sp<sup>3</sup>-carbon of fullerene and alkyne. The UV-vis spectra for all products showed a peak at 434 nm, which is the characteristic peak of the 1,2-adducts of C<sub>60</sub>.

The previous analysis allows proposing a reaction mechanism for the formation of 2a–2d as is shown in Scheme 1. First, the reaction of terminal alkynes 1 with TBAOH/CH<sub>3</sub>OH generates an intermediate compound 3, which suffers a nucleophilic addition to produce 4 (RC<sub>60</sub><sup>−</sup>). The protonation of the intermediate species 4 leads to the formation of the product 2.

The reactivity of other types of alkynes in the same conditions was also studied. The derivatives 2a (in general conditions), and 2d (when CF<sub>3</sub>COOH changed to PhCH<sub>2</sub>Br) were



Scheme 1 Proposed mechanism for the formation of 2a–2c.

curiously obtained when the terminal alkynes were transformed into ethylphenylpropiolate. The structures composed by two different C<sub>60</sub>-with alkynes were studied and determined through their HRMS, <sup>1</sup>H NMR, <sup>13</sup>C NMR, and UV-vis spectra. The results indicated that the cleavage of ethylphenylpropiolate occurred at the C–C bond, specifically, between phenylpropynyl and ethoxycarbonyl. This fact is confirmed by the crystal structure of 2d (Fig. 1).

A plausible mechanism for the formation of 2a and 2d compounds is deduced in the Scheme 2. The substrate 5 undergoes hydrolysis (under base conditions) to produce a carboxylic anion, which decomposes to CO<sub>2</sub>. Because the conjugated phenylethynyl group can disperse the negative charge, it makes it possible to form the intermediate 6. The α,β-unsaturated carboxylic acid 6 undergo decarboxylation<sup>14</sup> and generates a carbanion 3 in base condition to attack C<sub>60</sub>. The protonation or nucleophilic addition of benzyl bromide produces 2a and 2d, respectively.

The reactivity of the ester group was deeply studied using ethyl propiolate as substrate, which has a weak acid proton and an electron-withdrawing ester group in its structure. It is interesting to observe that the reactions of C<sub>60</sub> with 7 using TBAOH/CH<sub>3</sub>OH (1.0 mol L<sup>−1</sup>) and TBAOH/isopropanol (1.0 mol L<sup>−1</sup>) led to different products as the unprecedented

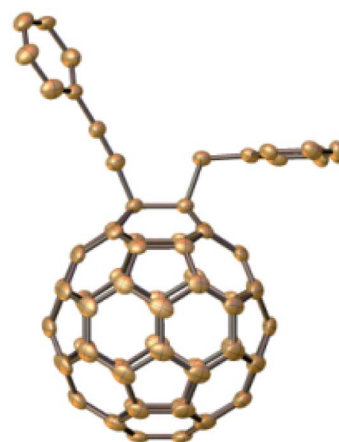
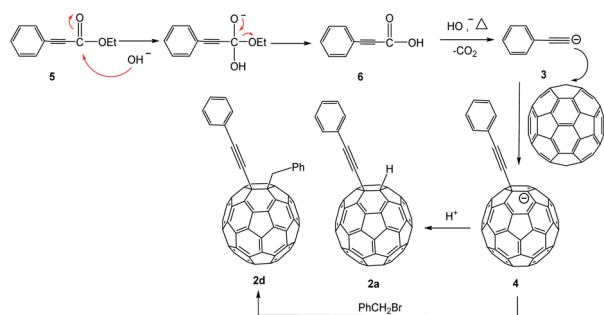


Fig. 1 Single crystal structure of 2d with 30% ellipsoid probability (hydrogens are omitted). The CCDC number of the crystal structure is: 2082000.





Scheme 2 Proposed mechanism for the formation of 2d.

product **7a** and **7b** with an isolated yield of 35.2% and 11.8% respectively. In addition, three  $\text{CH}_3\text{O}^-$  groups were incorporated in the adduct **7a**, apparently, methanol participates in the reaction and two bonds of alkyne are broken. Nevertheless, the alkynylfullerene derivative **7c** with an isolated yield of 27.5% was obtained at room temperature (Fig. 2). The reaction between  $\text{C}_{60}$  with **7** and TBAOH/ $\text{CH}_3\text{OH}$  ( $1.0 \text{ mol L}^{-1}$ ), suggests that methanol participates in the reaction. Obviously, the reactions follow different routes in the presence of methanol and isopropanol.

The structure of derivatives **7a–7c** were estimated through HRMS,  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, HMBC, HSQC and UV-vis spectra. HRMS (+ESI) of **7a** and HRMS (-ESI) of **7c** show the protonated molecular ions ( $[\text{M} + \text{H}]^+$ ) and ( $[\text{M} - \text{H}]^-$ ), which are placed at 869.6473 and 817.0260, respectively (Fig. S23 and S29 in ESI†). These are in good agreement with the assigned structures (**7a**:  $\text{C}_{66}\text{H}_{13}\text{O}_4^+$ , calcd. 869.0814; **7c**:  $\text{C}_{65}\text{H}_5\text{O}_2^-$ , calcd. 817.0290). The 1,2-addition pattern of **7a** and **7c** was further confirmed through UV-vis spectra which shown a characteristic sharp peak placed at 435 nm. Fig. 3 displays the  $^1\text{H}$  NMR spectrum of **7a**, where the methoxy protons are characterized by the three singlets situated at 3.48, 3.51, and 3.65 ppm, respectively.<sup>13,15</sup> The  $\text{C}_{60}\text{-H}$  proton was observed at 6.97 ppm. The ratio defined by peak area of the methoxy protons and the  $\text{C}_{60}\text{-H}$  proton is 3 : 1, consist with the addends added to  $\text{C}_{60}$  is methoxy groups, and indicated that the methanol from TBAOH solution took part in the reaction. Two methoxy protons (3.51 and 3.65 ppm) show  $^3J_{\text{CH}}$  correlations with the carbon of the methenyl group respectively (Fig. S26 in ESI†), which indicates that these two methoxyl groups were linked to the same carbon of methenyl. HSQC NMR (Fig. S27 in ESI†) confirmed the structure of **7a**. It should be noted that

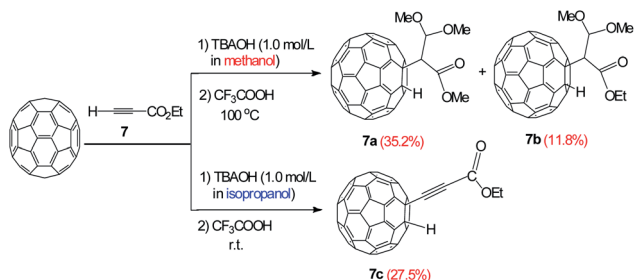


Fig. 2 Reaction of  $\text{C}_{60}$  and **7** in the presence of  $\text{OH}^-$ .

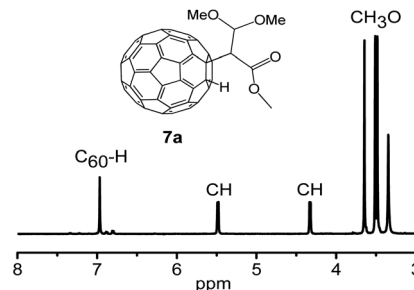
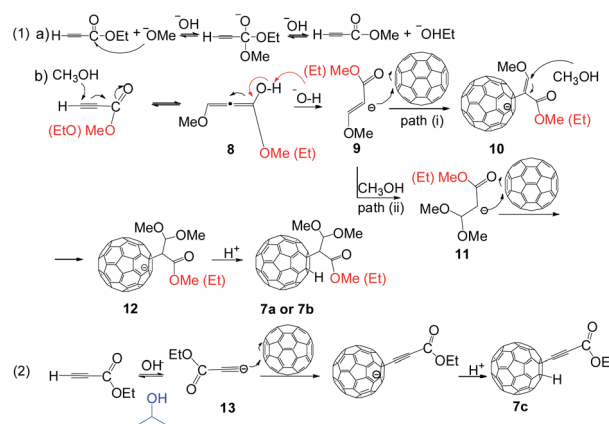


Fig. 3 Expanded  $^1\text{H}$  NMR spectrum of **7a** recorded using  $\text{CS}_2/(\text{DMSO}-d_6)$  as the external lock) measured in a 600 MHz instrument.

there is another peak with a retention time at 6.8 min associated with a HPLC trace of the reaction of  $\text{C}_{60}$  with TBAOH/ $\text{CH}_3\text{OH}$  ( $1.0 \text{ mol L}^{-1}$ ) (Fig. S5 in ESI†). Data extracted from HRMS,  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, HMBC, HSQC and UV-vis spectra, lead us to inferred that the peak placed before **7a** belongs to product **7b** (Fig. S29–S33 in ESI†).

A reaction mechanism is proposed and shown in the Scheme 3. The key is that the alcoholic solution (in the TBAOH) could affect the structure of the final products. Ethyl propiolate **7** is an  $\alpha,\beta$ -unsaturated carbonyl compound in the reaction. Methanol attacks to  $\beta$ -carbon (basic conditions), and produces the intermediate **8**. To verify the alcohol exchange reaction, a control experiment between **7** and TBAOH/ $\text{CH}_3\text{OH}$  ( $1.0 \text{ mol L}^{-1}$ , 150  $\mu\text{L}$ , 3 equiv.) in *o*-DCB was carried out for 1 h at  $100^\circ\text{C}$ . As expected,  $^1\text{H}$  NMR spectra of the reaction shows peaks at 4.26 ppm (q, 2H) and 1.32 ppm (t, 3H), which are associated with methylene and methyl protons of compound **7**, respectively. In addition, the peak around 3.80 ppm is associated with methyl proton of methyl propiolate (Fig. S42 in ESI†), and it reveals that the crude contains both ethyl propiolate **7** and methyl propiolate, confirming the alcohol exchange reaction. Therefore, **7a** and **7b** products are both detected. However, similar reactions do not take place in the case of phenylacetylene and ethyl-phenylpropiolate are used as substrates, probably due to the steric hindrance of the phenyl ring in the molecule. The produced intermediate **8** generates **9** via Michael addition.<sup>15,16</sup> A



Scheme 3 Proposed mechanism for the formation of **7a–7c**.

part of **9** produces **11** through alcohol exchange reaction. The Michael addition, and subsequent protonation, would afford the generation of **7a** and **7b**. However, the Michael addition reaction was restricted when methanol was changed to isopropanol. This fact was ascribed to two reasons: (1) the steric effect of isopropanol; (2) the lower reactivity of isopropanol.<sup>17–19</sup> Thus, a deprotonation reaction of **7** happens when isopropanol is present. The generated intermediate **13** undergoes through nucleophilic addition and protonation to produce **7c**.

## Conclusions

In summary, we have found that the addition of alkynes to C<sub>60</sub> can now be achieved without the use of heavy metals and alkynyllithium, under simple reaction conditions which were mediated by OH<sup>−</sup>. In the reaction of C<sub>60</sub> with ethyl propiolate, the use of alcoholic solution in the TBAOH was found to be crucial. The reaction performed under methanol afforded a Michael-addition product, but Michael-addition was restrained while the reaction under isopropanol. Through the nucleophilic addition, we have introduced alkynes, a versatile functional group to C<sub>60</sub> under mild conditions, and the monoadducts may also prove to be useful for other applications.

## Conflicts of interest

There are no conflicts of interest.

## Acknowledgements

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