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# Preparation of carboxylic arylphosphines by hydrolysis of the trifluoromethyl group†

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Triarylphosphines substituted with carboxylic and trifluoromethyl groups have been prepared by the hydrolysis of trifluoromethyl groups using fuming sulfuric acid and boric acid. The reaction has been studied in a set of homoleptic and heteroleptic trifluoromethylated triarylphosphines and offers a new synthetic procedure for the preparation of carboxylic phosphines with a relatively simple methodology. The degree of carboxylation is modulated by the reaction conditions and is sensitive to the substitution pattern of the starting trifluoromethylated phosphines. A pH-dependent procedure based on the amphiphilic character of these phosphines was developed for their separation and purification. The electronic properties of the synthesized carboxylic-trifluoromethylated phosphines have been analyzed by  $^{31}\text{P}$  NMR of the corresponding selenide derivatives. Finally, the structures of two palladium complexes, containing the *para* and *meta* carboxylic-trifluoromethylated phosphines are also described, showing different dimeric structures.

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## Introduction

Phosphines containing carboxylic groups have been widely reported in the literature. The amphiphilic behaviour that the carboxylic groups confer to the phosphines has been found of interest in several applications such as their use as water-soluble reducing agents of disulfide bonds<sup>1</sup> or as ligands in transition metal-based catalysis, either by coordination through the phosphorus atom<sup>2</sup> or acting as chelating ligands.<sup>3</sup> Moreover, the carboxylic derivatives of triphenylphosphine have been used as building blocks in the preparation of porous coordination materials.<sup>4</sup> In spite of the versatility of carboxylic phosphines, certain considerations should be born in mind in regard to their preparation. The most extended procedures for the general synthesis of phosphines, such as the reaction of chlorophosphines or phosphorus trichloride with Grignard reagents or organolithium derivatives, as well as the nucleophilic reaction between alkali phosphides and haloderivatives, are not compatible with the presence of acidic protons. In the last case, the incompatibility between the phosphide anions and the carboxylic acid can be overcome using strong basic conditions, although sometimes this is accompanied by solubility issues. In this regard, several other alternative strategies have been developed for the synthesis of carboxylic phosphines. Ravindar *et al.* reported the reaction of carbonation of organolithium derivatives to yield monocarboxylic triarylphosphines.<sup>5</sup> The

hydrolysis of the nitrile group has also been found to yield mono-, di- and tri-carboxylic triarylphosphines, and has also been applied to some alkyl phosphines.<sup>6</sup> Finally, Pd(II)-catalysed P–C coupling has been also used for the preparation of carboxylic phosphines.<sup>7</sup>

Carbonyl compounds can be prepared from a trifluoromethyl group under strong acidic conditions. In 1943, Simons and Ramler<sup>8</sup> and Gilman and Blume<sup>9</sup> reported the hydrolysis of the trifluoromethyl group, using sulfuric acid, as a methodology to establish the structure of different trifluoromethylbenzene derivatives. Later, Le Fave *et al.* reported the formation of benzoic acids from the hydrolysis of trifluoromethylbenzene and derivatives, using concentrated sulfuric acid, followed by the treatment of the mixture with water.<sup>10</sup> They also proved that benzoate esters could be obtained when treating the reaction mixtures with alcohols instead of water.<sup>11</sup> Furthermore, the same authors described the preparation of phthalic acids by hydrolysis of the corresponding bis(trifluoromethyl)benzenes in an equimolar mixture of sulfuric acid and chlorosulfuric acid.<sup>12</sup> In spite of these early publications, the reaction of hydrolysis of the trifluoromethyl group has not attracted much interest and only a few more examples have been reported after the pioneering studies of Le Fave. Wen and co-workers applied the reaction of hydrolysis, using fuming sulfuric acid, in the preparation of cholesteryl *p*-perfluoroalkyl benzoates.<sup>13</sup> Platonov and co-workers showed the formation of perfluorobenzoic acids from the reaction of the corresponding perfluorotoluene derivative with  $\text{CF}_3\text{CO}_2\text{H}/\text{SbF}_5$  and the subsequent aqueous treatment.<sup>14</sup> Finally, Kethe *et al.* studied the hydrolysis of one trifluoromethyl group in 1,3,5-tris(trifluoromethyl)benzene with either triflic acid or fluorosulfuric acid.<sup>15</sup>

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Despite the above mentioned applications of the hydrolysis of the trifluoromethyl group, to the best of our knowledge, no precedents are known about the application of this reaction in the synthesis of phosphines. We envisioned the use of fuming sulfuric acid in combination with boric acid to promote the hydrolysis of the trifluoromethyl group in trifluoromethylated triarylphosphines. Amongst the available superacid systems, fuming sulfuric acid is strong ( $-H_0$  up to 14.9 in 65 wt%  $\text{SO}_3$  oleum),<sup>16</sup> non-expensive and, most importantly, with a relatively lower oxidation power compared to other stronger super acids such as magic acid (*i.e.*  $\text{FSO}_3\text{H}\cdot\text{SbF}_5$ ). Moreover, it has been previously studied in the sulfonation of phosphines.<sup>17</sup> Herrmann and co-workers showed that the oxidizing power of fuming sulfuric acid could be reduced in combination with boric acid thus preventing the oxidation of the phosphines during the process of sulfonation.<sup>17a</sup> In this sense, the studies of Gillespie and co-workers showed that the addition of boric acid to sulfuric acid generates a strong superacid of the sulfuric acid system,  $\text{H}_3\text{SO}_4[\text{B}(\text{OSO}_3\text{H})_4]$ .<sup>18</sup> More recently, Peral *et al.* reported the use of oleum and boric acid in the sulfonation of the strongly deactivated trifluoromethylated phosphines with minor yields of phosphine oxides.<sup>17b,c</sup>

Based on the previous examples, the reaction of hydrolysis of the trifluoromethyl group offers a new synthetic procedure to obtain carboxylic arylphosphines through a relatively simple procedure. In this methodology, the trifluoromethyl group will be used as a masked carboxylic acid, overcoming the difficulties that arise from the incompatibility of carboxylic acids with the reagents commonly used in the preparation of phosphines. Most interestingly, this methodology provides the preparation of arylphosphines containing both carboxylic and trifluoromethyl groups. On the one hand, the carboxylic group confers chemical versatility to phosphines, covering a wide variety of applications, including those derived from the amphiphilic character of the carboxylic group. On the other hand, the trifluoromethyl group confers greater  $\pi$ -acidity to the phosphines, which turns into a higher stability towards their oxidation.<sup>17c,19</sup> Phosphines with  $\pi$ -acidic character have proven to increase the activity of certain transition metal-catalysed reactions such as some C–C couplings<sup>20</sup> and carbonylations.<sup>17c,21</sup>

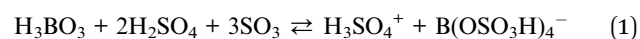
A few examples of carboxylic and trifluoromethylated arylphosphines have been previously described in the literature. Reek's group reported the synthesis of carboxylic and trifluoromethylated phosphines through a Pd catalysed P–C coupling.<sup>22</sup> More recently, Starkov *et al.* described the synthesis of an *ortho*-carboxylic trifluoromethylated phosphine by a several-steps methodology in which the P–C bond is constructed through a Cu-mediated coupling with the carboxylic group protected as methyl ester.<sup>23</sup>

Here we report the preparation of carboxylic phosphines by the reaction of hydrolysis of the trifluoromethyl group in trifluoromethylated triarylphosphines, using fuming sulfuric acid and boric acid. The carboxylic-trifluoromethylated phosphines have been successfully separated and purified, and their electronic properties have been studied by  $^{31}\text{P}\{^1\text{H}\}$  NMR using the selenium derivatives. Palladium complexes of the *para* and *meta* carboxylic and bis(trifluoromethylated) triarylphosphine were

also structurally characterized showing different supramolecular patterns, illustrating the capability of this family of ligands to build diverse intermolecular networks.

## Results and discussion

The synthesis of carboxylic triarylphosphines was carried out from the hydrolysis of the trifluoromethyl group of a series of triarylphosphines with different substitution pattern, Scheme 1. The reaction was performed using a highly acidic media prepared with a mixture of boric acid and fuming sulfuric acid (oleum). The fine tuning of the reaction conditions allows the control of the yield and selectivity of the reaction. The protection towards the oxidation of the phosphines under these highly acidic conditions takes place through the protonation of the phosphorus atom.<sup>17a</sup> For this reason, complete solubility of the phosphines is necessary before the addition of fuming sulfuric acid. It is important to notice that water is formed when mixing  $\text{H}_3\text{BO}_3$  and sulfuric acid, previous to the addition of the phosphines, eqn (1),<sup>17a,18</sup> thus changing the polarity of the media and decreasing the acidity of the mixture. Therefore, sulfur trioxide from oleum is required to react with water shifting the equilibrium of eqn (1) towards the formation of the superacid.



As it will be explained below, the superacid plays an important role, not only protecting the phosphines from oxidation, but also in the hydrolysis of the trifluoromethyl groups.

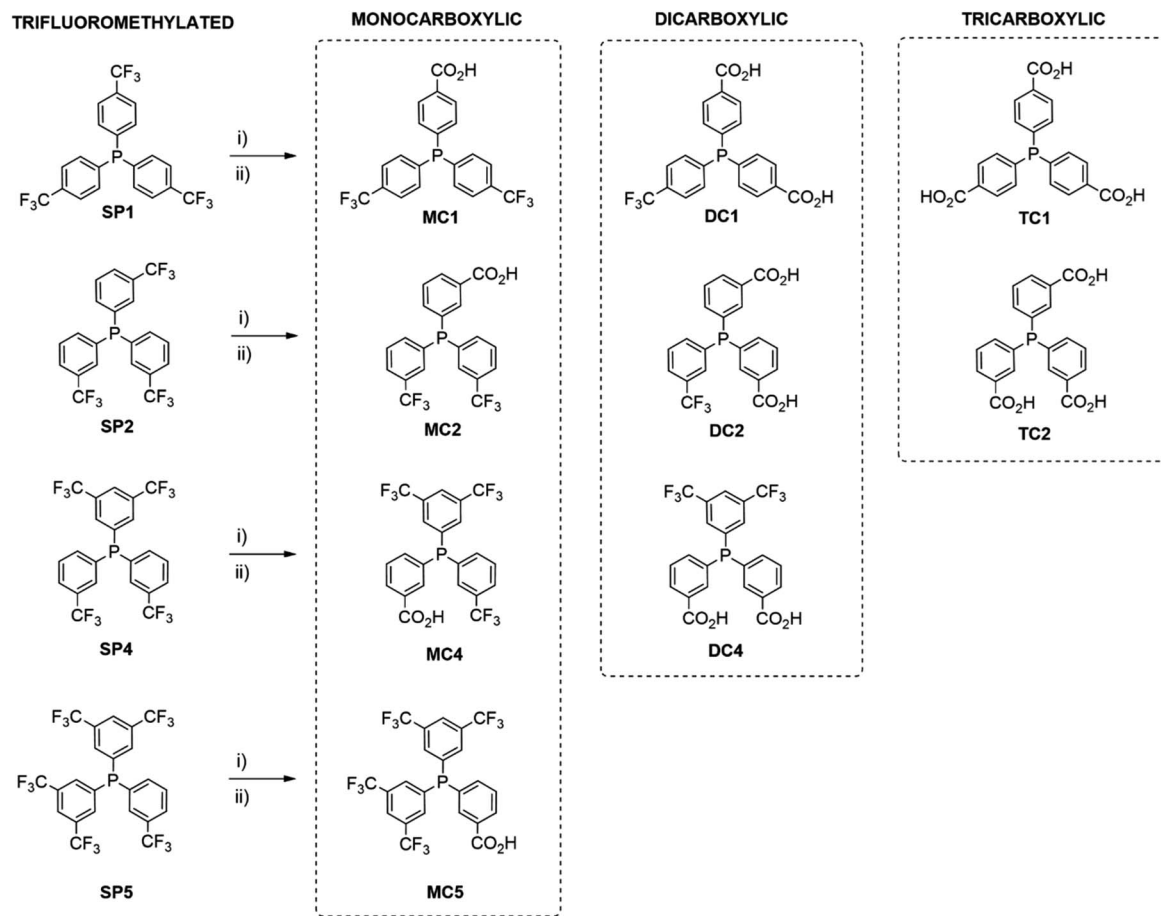
### Synthesis of carboxylic phosphines from homoleptic trifluoromethylated phosphines

The hydrolysis of the trifluoromethyl groups of three homoleptic phosphines (**SP1**, **SP2**, **SP3**) were firstly studied.

In a preliminary set of experiments, the *para*-substituted tris(4-trifluoromethylphenyl)phosphine (**SP1**) and the chosen amount of boric acid were dissolved in sulfuric acid before the addition of 65%  $\text{SO}_3$  fuming sulfuric acid. The evolution of the reaction was monitored by aliquoting the sulfuric mixture. After quenching the samples in cold water, the phosphines were extracted with diethyl ether and analysed by  $^{31}\text{P}\{^1\text{H}\}$  NMR to quantify the distribution phosphorus compounds present in the mixture. Control analysis of the aqueous layers showed that all phosphorus compounds were completely extracted into the organic phase. Three carboxylic products were observed as result of the hydrolysis of one, two or all the trifluoromethyl groups.

The reaction was explored and optimized at different concentrations of phosphine,  $\text{H}_3\text{BO}_3$  and  $\text{SO}_3$  (ESI Fig. S1†). The reaction rates of the hydrolysis of the trifluoromethyl group were observed to strongly depend on the concentration of  $\text{H}_3\text{BO}_3$  and  $\text{SO}_3$ . High concentrations of  $\text{SO}_3$  (above 12 M) gave faster reaction rates although led to the complete oxidation of the phosphine products at short reaction times. As expected, increasing  $\text{H}_3\text{BO}_3$  concentration was found to decrease the rate of phosphine oxidation. However, and contrary to what was





Scheme 1 Synthesis of carboxylic triarylphosphines from trifluoromethylated phosphines. (i)  $\text{H}_3\text{BO}_3/\text{H}_2\text{SO}_4/\text{SO}_3$ . (ii)  $\text{H}_2\text{O}$ .

observed for the reactions of sulfonation of triarylphosphines,<sup>17b,c</sup> the use of  $\text{H}_3\text{BO}_3$  showed a positive effect on the rate of the trifluoromethyl hydrolysis reaction. The most suitable conditions for the preparation of the monocarboxylic phosphine **MC1** and the dicarboxylic phosphine **DC1** were found at concentrations of  $\text{SO}_3$  and  $\text{H}_3\text{BO}_3$  of 10.5 M and 0.8 M, respectively, with initial phosphine concentration of 0.15 M. Fig. 1 shows the graphical representation of the consumption of the trifluoromethylated starting phosphine **SP1** together with the evolution of the monocarboxylic phosphine **MC1**, the dicarboxylic phosphine **DC1** and the tricarboxylic phosphine **TC1** under the optimized conditions. In addition to the molar fraction of the phosphines, the sum of phosphine oxides present in the mixture is also represented.

In the optimized conditions, after 1 h, the reaction mixture showed 25% of the monocarboxylic compound **MC1**, together with 11% of phosphine oxides. After 3 h of reaction, the amount of **MC1** increased to 45% and 10% of the dicarboxylic phosphine **DC1** was detected, with no significant variation of the amount of phosphine oxides. When the reaction was analyzed after 21 h, the mixture showed complete consumption of the starting phosphine **SP1**, it being the dicarboxylic phosphine **DC1** the major component of the mixture (53%) together with 22% of the monocarboxylic phosphine **MC1** and 23% of oxides.

Moreover, in this experiment, only traces of the tricarboxylic phosphine **TC1** were detected (*ca.* 2%).

Fine tuning of the reaction conditions and reaction time allows the control of the composition of the reaction mixture. For the preparation of the monocarboxylic phosphine **MC1**, reaction times can be shortened to 4 hours at the optimized conditions described above, to get 50% of this carboxylic phosphine together with the unreacted phosphine **SP1**, which

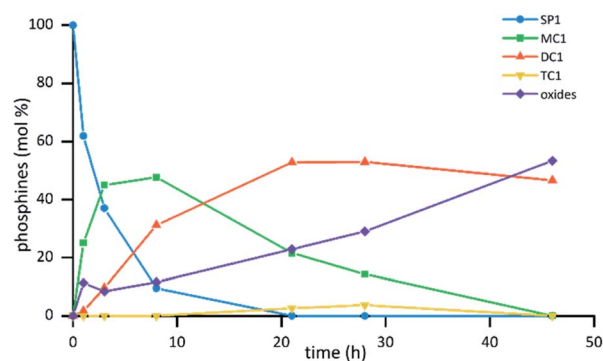


Fig. 1 Evolution of the reaction of hydrolysis in phosphine **SP1** in sulfuric acid with time. Reaction conditions [**SP1**] = 0.15 M, [ $\text{H}_3\text{BO}_3$ ] = 0.8 M, [ $\text{SO}_3$ ] = 10.5 M.



after proper separation, could be reused in a new reaction. If the desired product is the dicarboxylic phosphine **DC1**, the reaction mixture is quenched after 14 h to obtain a 50% mixture of **MC1** and **DC1**, Table 1. Moreover, increasing the concentration of  $\text{H}_3\text{BO}_3$  to 2.4 M, it is possible to obtain 46% of the tricarboxylic phosphine **TC1**, together with 36% of the dicarboxylic phosphine **DC1** and below 25% of phosphine oxides after 20 hours. Although such conditions need fairly long reaction times, they open a new synthetic approach to the already known procedures for preparing this tricarboxylic phosphine.<sup>6</sup>

The synthesis of *meta*-carboxylic phosphines was also essayed from the homoleptic trifluoromethylated phosphine **SP2** using the same optimized conditions than for the *para*-substituted phosphine **SP1**. The evolution of the reaction analysed by  $^{31}\text{P}$  NMR also showed the hydrolysis of one, two and three trifluoromethyl groups producing the monocarboxylic phosphine **MC2**, the dicarboxylic phosphine **DC2** and the tricarboxylic phosphine **TC2**, respectively, Scheme 1. Interestingly, the rate of the reaction of hydrolysis of the trifluoromethyl groups in *meta* is much faster than that of the trifluoromethyl groups in *para* position (ESI Fig. S2†). For the synthesis of the monocarboxylic phosphine **MC2**, the reaction is quenched after 0.5 h to obtain a mixture of 48% of **SP2** and 41% of **MC2** with only traces of the dicarboxylic phosphine. For the preparation of the dicarboxylic phosphine **DC2**, the reaction was left for 20 h to obtain a mixture of 6% of **MC2**, 60% of **DC2** and 20% of the tricarboxylic phosphine **TC2**, Table 1. As previously shown for **SP1**, with the *meta*-isomer **SP2**, it is also possible to tune the composition of the mixture before the separation process adjusting the reaction conditions.

Based on the faster reactivity of the trifluoromethyl groups in *meta* position with respect to the *para* substituted phosphines, the reaction of hydrolysis of the trifluoromethyl groups of the 3,5-substituted phosphine **SP3** was tested, Chart 1. Unexpectedly, the hydrolysis in this doubly *meta*-substituted phosphine **SP3** was much slower than for the hydrolysis of the single-substituted phosphine **SP2**. Using the optimized conditions for the preparation of the *para*- and *meta*-carboxylic phosphines, the progress of the reaction for **SP3**, analysed by NMR, showed considerable slower rates for the hydrolysis of the trifluoromethyl groups than in the case of **SP1** or **SP2**, yielding phosphine oxides as main products. Mass spectrometry analysis of the reaction mixture after 20 h confirmed the presence of the monocarboxylic phosphine, the monocarboxylic phosphine

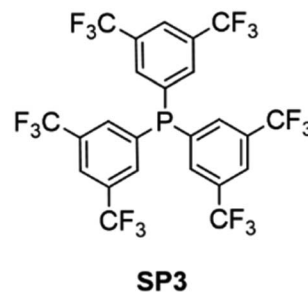


Chart 1 Trifluoromethylated phosphine **SP3**.

oxide and a dicarboxylic phosphine oxide. However, the pattern of carboxylation in the dicarboxylic compound (*i.e.* in the same or different ring), could not be determined.

### Synthesis of carboxylic phosphines from heteroleptic trifluoromethylated phosphine

Given the feasibility of the synthesis of carboxylic phosphines from the hydrolysis of the *para* (**SP1**) and *meta*-substituted (**SP2**) trifluoromethylated phosphines and the poor reactivity of the trifluoromethyl groups in the 3,5-substituted phosphines (**SP3**), it was appealing to study the reaction of hydrolysis in heteroleptic phosphines containing both patterns of substitution, (**SP4**, **SP5** in Scheme 1 and **SP6** and **SP7** in Chart 2).

As a general trend, low solubility of the heteroleptic phosphines in the  $\text{H}_3\text{BO}_3/\text{H}_2\text{SO}_4$  mixture, previous to the addition of oleum, was observed yielding to high percentages of oxides at short reaction times. The solubility of the phosphines decreased when increasing the number of trifluoromethyl groups and this was more evident with the *para*-substituted heteroleptic phosphines (**SP6** and **SP7**). Moreover, the hydrolysis of the trifluoromethyl group with the heteroleptic phosphines showed slower reaction rates for the *para*-isomers (**SP6** and **SP7**) than for the *meta*-substituted phosphines (**SP4** and **SP5**), a trend already observed for the homoleptic phosphines **SP1** and **SP2**. Furthermore, hydrolysis of the trifluoromethyls at 3,5-positions was not detected. It is important to note that the reaction rates were found to be lower when increasing the number of trifluoromethyls in the phosphines (ESI Fig. S3†).

Optimized conditions were found for the synthesis of the *meta*-carboxylic phosphines **MC4**, **DC4** and **MC5**, as shown in

Table 1 Optimized conditions for the synthesis of carboxylic and trifluoromethylated phosphines

Carboxylic phosphine	Starting phosphine	Reaction conditions				Hydrolysis yield (%)				
		[SP] (M)	[SO <sub>3</sub> ] (M)	[H <sub>3</sub> BO <sub>3</sub> ] (M)	Time (h)	SP	MC	DC	TC	Oxides
<b>MC1</b>	<b>SP1</b>	0.15	10.5	0.8	4	46	46	0	0	8
<b>DC1</b>	<b>SP1</b>	0.15	10.5	0.8	14	4	40	48	0	8
<b>MC2</b>	<b>SP2</b>	0.15	10.5	0.8	0.5	48	41	6	0	5
<b>DC2</b>	<b>SP2</b>	0.15	10.5	0.8	20	0	6	60	20	14
<b>MC4–DC4</b>	<b>SP4</b>	0.07	9	0.8	18	9	48	34	—	9
<b>MC5</b>	<b>SP5</b>	0.08	9	0.4	50	0	63	—	—	37





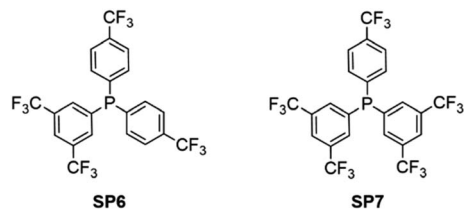


Chart 2 Heteroleptic trifluoromethylated phosphines SP6 and SP7.

Table 1. Long reaction times were needed to obtain the carboxylic phosphines in moderate yields, due to the diluted phosphine concentrations required to minimize the initial oxidation of SP4 and SP5.

For the *para*-heteroleptic phosphines SP6 and SP7 the reaction was found to be very slow. The large amount of phosphine oxides and the slow reactivity of the *para*-trifluoromethyls jeopardize the preparation and isolation of the corresponding carboxylic phosphines. After several attempts to optimize the reactions, no suitable conditions were found for the preparation and isolation of the corresponding carboxylic phosphines. Therefore, other reported procedures would be more appropriate for the formation of these carboxylic phosphines.<sup>5,7,22,23</sup>

### Separation and purification of the carboxylic phosphines

In all the synthetic procedures described above, a mixture of different phosphines is obtained after the reaction of the trifluoromethylated starting phosphines with H<sub>3</sub>BO<sub>3</sub> and oleum. A general methodology for the separation and purification of the

carboxylic-trifluoromethylated phosphines has been developed based on the pH-dependent solubility of carboxylic acids. The separation process has been optimized for a mixture containing all the possible phosphine compounds derived from the reaction of hydrolysis of trifluoromethylated phosphines, that is, starting phosphine (SP), monocarboxylic (MC), dicarboxylic (DC), tricarboxylic (TC) and phosphine oxides, to cover all possible scenarios. A scheme of the separation is presented in Fig. 2.

After the hydrolysis reaction, the mixture was poured on ice and the phosphorus compounds were extracted from the aqueous sulfuric media using diethyl ether. Water was added to the previous Et<sub>2</sub>O solution, yielding a new biphasic organic-aqueous system. The aqueous phase was brought to pH 6.0 using an aqueous solution of NaOH to completely extract the tricarboxylated phosphine (TC) into the aqueous phase, together with its oxide and the oxide of DC, when present. After phase separation, water was added again to the ethereal phase and further basification to reach a pH of about 7.5 allowed the extraction of the dicarboxylic phosphine (DC) into the aqueous solution. Within this range of pH, also the oxide of the monocarboxylic compound (MC) is extracted. The acidification of each of the aqueous phases containing the sodium salts of the dicarboxylate and tricarboxylate phosphines to pH 2, allowed the re-extraction of the phosphines in the acidic form in diethyl ether.

The extraction of the monocarboxylic phosphine (MC) from the ethereal phase by further increasing the pH of the aqueous phase is not possible because of the formation of stable emulsions. Therefore, an alternative procedure was developed starting by vacuum evaporation of the diethyl ether. To the solid

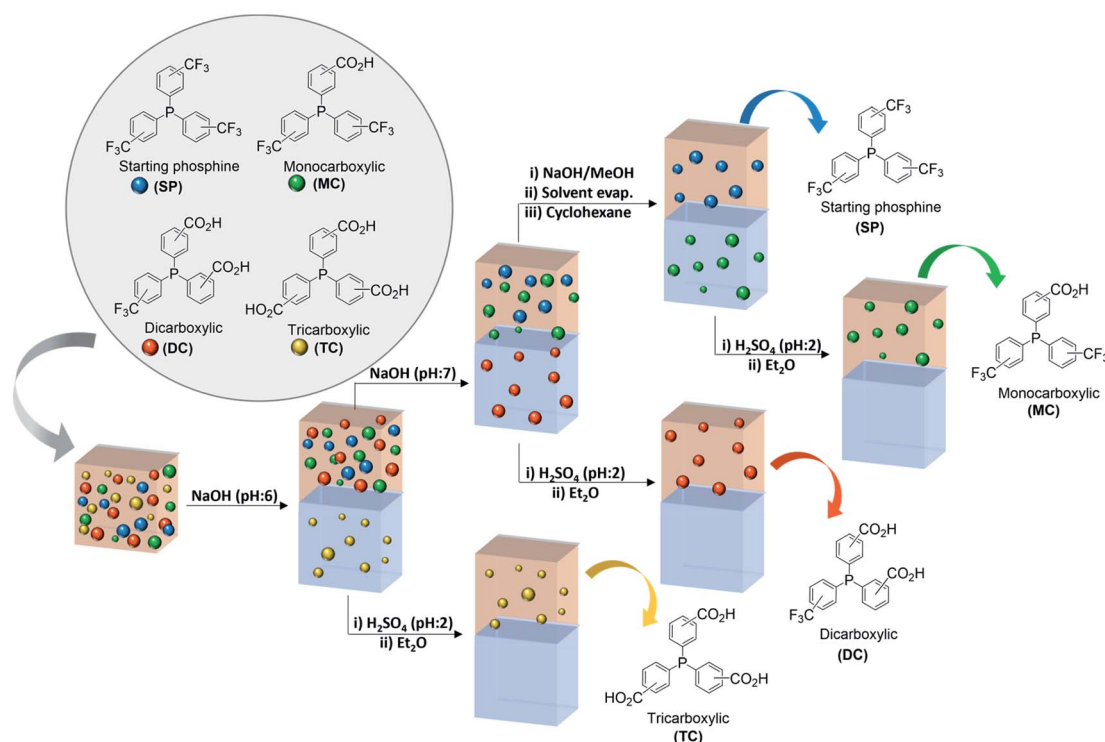


Fig. 2 Separation of the phosphines products of the reaction of hydrolysis of the trifluoromethyl group.

residue containing SP and MC, a solution of 10% NaOH in methanol was added dropwise. After complete deprotonation of the carboxylic group of MC, monitored by TLC ( $\text{CH}_2\text{Cl}_2$ ), methanol was vacuum evaporated and cyclohexane was added to the solid residue to extract the unreacted starting phosphine (SP). Finally, phosphine MC was recovered again as carboxylic acid by acidification to pH 2 followed by extraction in diethyl ether.

With this procedure, it was possible to prepare a whole set of mono- and dicarboxylic trifluoromethylated triarylphosphines, Scheme 1.<sup>24,25</sup> These phosphines are a proof concept of the application of the hydrolysis of the trifluoromethyl group in trifluoromethylated arylphosphines using sulfuric fuming acid and  $\text{H}_3\text{BO}_3$ , and open a new synthetic route in the synthesis of these sort of compounds.

### Reaction mechanism

The nature of the highly acidic medium used for the hydrolysis of the trifluoromethyl groups hampers the study of the reaction mechanism. Nevertheless, with  $^{31}\text{P}\{^1\text{H}\}$  and  $^{19}\text{F}\{^1\text{H}\}$  NMR experiments it is possible to analyse *in situ* the progression of the reaction. Fig. 3 shows the evolution of the hydrolysis of the trifluoromethyl group in the *p*-tris(trifluoromethylphenyl)phosphine (SP1) monitored by  $^{31}\text{P}\{^1\text{H}\}$  and  $^{19}\text{F}\{^1\text{H}\}$  NMR.

Importantly, a singlet in the  $^{19}\text{F}\{^1\text{H}\}$  spectra at 42.6 ppm, corresponding to the  $-\text{OSO}_2\text{F}$  moiety<sup>26</sup> was detected. The intensity of this signal increased as new phosphine signals were observed in the  $^{31}\text{P}\{^1\text{H}\}$  NMR spectra. Besides the signal of the trifluoromethyl groups (around  $-63$  ppm), no other fluorinated species were found in these experiments.

Based on the NMR experiments and the previously reported reactions of trifluoromethylated aryl compounds in acidic media,<sup>15,27</sup> a mechanism for the hydrolysis of the trifluoromethyl group in trifluoromethylated triarylphosphines using fuming sulfuric acid and  $\text{H}_3\text{BO}_3$  can be postulated, Scheme 2.

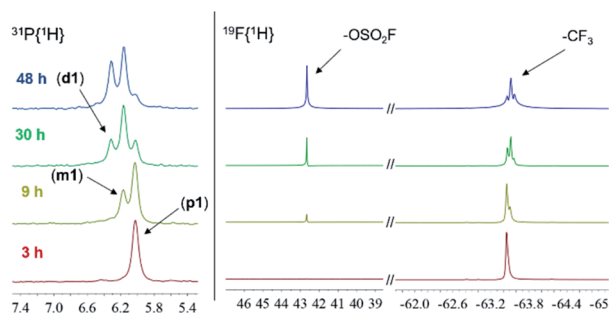
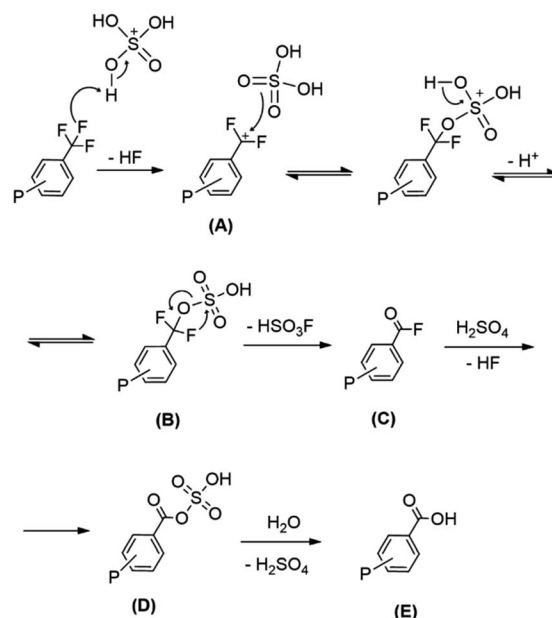


Fig. 3 Evolution of the reaction of hydrolysis in phosphine SP1. Analysis of the crude reaction by  $^{31}\text{P}\{^1\text{H}\}$  (101.27 MHz) (left) and  $^{19}\text{F}\{^1\text{H}\}$  NMR (235.4 MHz) (right). Signals (intensity and position) relative to  $\text{NH}_4\text{PF}_6$  used as internal standard ( $-144.45$  ppm and  $-72.65$  ppm, respectively).  $\delta$  in ppm. In  $^{31}\text{P}\{^1\text{H}\}$  NMR, only phosphine region is shown. p1, m1 and d1 refer to the precursors of phosphines SP1, MC1, and DC1 in sulfuric medium. Reaction conditions: [phosphine] =  $0.16$  M,  $[\text{H}_3\text{BO}_3]$  =  $0.8$  M,  $[\text{SO}_3]$  =  $11.5$  M.



Scheme 2 Proposed reaction mechanism of the hydrolysis of the trifluoromethyl group in trifluoromethylated phosphines with sulfuric acid.

In a first step, strong interaction between the acidic protons of the superacid and fluorine promotes the C–F bond cleavage releasing a HF molecule and generating the difluorobenzylic carbocation (A), in analogy with previously reported experimental and theoretical studies.<sup>27</sup> Next, the nucleophilic attack of a molecule of  $\text{H}_2\text{SO}_4$  on the carbocation A yields the formation of the hydrogensulfuric ester B. The reaction may be favoured by the highly electrophilic fluorinated carbocation. Following, the concerted elimination of  $\text{HSO}_3\text{F}$  (*ca.* 40–42 ppm in  $^{19}\text{F}\{^1\text{H}\}$ ) would give the corresponding acyl fluoride (C) similarly to previously reported studies.<sup>15,27b</sup> Finally, the acyl fluoride (C) would react with sulfuric acid to yield the mixed sulfonic-carboxylic anhydride (D).<sup>28</sup> Although the formation of mixed anhydrides by addition of  $\text{SO}_3$  to acyl fluorides has been reported in the literature,<sup>26</sup> this option was discarded considering that such reaction requires the use of pure  $\text{SO}_3$ .<sup>29</sup> Moreover, the signal for the corresponding fluorinated mixed anhydride in  $^{19}\text{F}\{^1\text{H}\}$  NMR (above 45 ppm) was not detected. After quenching the reaction in water, D is hydrolysed to form the carboxylic acid (E). It is worth mentioning that neither the signal for the fluorinated carbocation (A) (*ca.* 23 ppm)<sup>30</sup> nor the corresponding signal for the acyl fluoride (C) (*ca.* 17.7 ppm)<sup>31</sup> or any other intermediates were detected in the  $^{19}\text{F}\{^1\text{H}\}$  NMR experiments. Therefore, the elimination of HF to form the carbocation (A) is likely to be the rate determining reaction step, in agreement with the mechanism proposed by Kethe *et al.* in the hydrolysis of the trifluoromethyl group in trifluoromethylbenzenes using triflic acid or fluorosulfuric acid.<sup>15</sup>

In the proposed mechanism,  $\text{H}_3\text{BO}_3$  and  $\text{SO}_3$  would be involved in the generation of the superacid mixture. The  $^{11}\text{B}$  NMR spectra of the reaction mixtures containing phosphine, sulfuric acid, oleum and boric acid only showed the tetrahedral



species  $\text{B}(\text{OSO}_3\text{H})_4^-$  (ESI Fig. S4†). Gillespie and co-workers showed that binary mixtures of sulfuric acid with either  $\text{H}_3\text{BO}_3$  or  $\text{SO}_3$  give rise to an increase in the acidity of the media.<sup>16,18</sup> In the case of the present work, the increase in the acidity produced by  $\text{SO}_3$  and  $\text{H}_3\text{BO}_3$  would contribute to the C–F bond cleavage in the first step of the reaction.

Based on the proposed mechanism, the different reaction rates between phosphines with different substitution pattern can be explained through the resonant structures. Since phosphines are protonated under the reaction conditions, the formation of the difluorobenzyl carbocation (A) implies the generation of a second positive charge on the molecule. While in *para*-substituted phosphines the positive charge of the phosphonium can be delocalised at the adjacent atom to the carbocation, in *meta*-substituted phosphines this charge is never localised next to the other (ESI Fig. S5†). Therefore, the formation of the difluorobenzyl carbocation would be more favourable when the trifluoromethyl group is in *meta*-position than when located in *para*. Moreover, the extremely low reaction rates observed with those trifluoromethyls in 3,5-positions might be attributed to inductive effects.

### Study of the $\sigma$ -basicity of the trifluoromethylated-carboxylic phosphines: synthesis of the phosphine selenides

Phosphines have been widely studied as ligands in transition-metal catalysed reactions. The stereoelectronic properties of these ligands are determined by their substituents and play an essential role on the catalytic performance. Selenide derivatives have been used as a simple tool to study the  $\sigma$ -donation properties of phosphines through the  $^1J_{\text{PSe}}$  coupling of the  $^{77}\text{Se}$  isotopomer, measured by  $^{31}\text{P}$  NMR.<sup>19,32,33</sup> The selenide derivatives of the carboxylic-trifluoromethylated phosphines have been prepared following a previously reported methodology in trifluoromethylated triarylphosphines.<sup>19</sup>

As previously observed in trifluoromethylated phosphines, the  $^1J_{\text{PSe}}$  increases with the number of trifluoromethyl groups, thus decreasing the basicity of the ligands (ESI Table S1†). The

hydrolysis of each trifluoromethyl group producing a new carboxylic group causes a moderate decrease, between 2 and 5 Hz, in the  $^1J_{\text{PSe}}$  of the phosphine selenides, consistent with the decrease in the electron withdrawing power of the substituents. In previous works,<sup>19,33</sup> it has been proven that the introduction of trifluoromethyl groups in triarylphosphines was cumulative and resulted in the increase of 12 Hz in the  $^1J_{\text{PSe}}$ , regardless of the position of the group. Same observation was found in the case of the trifluoromethylated-carboxylic phosphines. Similarly, the introduction of carboxylic moieties is also cumulative and adds 9 Hz to the  $^1J_{\text{PSe}}$ . Consequently, eqn (2) predicts the values of the  $^1J_{\text{PSe}}$  in trifluoromethylated and/or carboxylic triarylphosphines with an error  $\leq 3$  Hz

$$^1J_{\text{PSe}} = 731 + 12n + 9m \quad (2)$$

where  $n$  and  $m$  are the number of trifluoromethyl and carboxylic groups in the phosphine, respectively, regardless of their position in the ring.

### *Trans*- $\text{PdCl}_2\text{L}_2$ complexes of phosphines MC1 and MC2

The two monocarboxylic phosphines were used to synthesise the corresponding *trans*- $\text{PdCl}_2\text{L}_2$  ( $\text{L} = \text{MC1}$  and  $\text{MC2}$ ) complexes crystals structures of both complexes were solved showing the formation of different dimeric structures by hydrogen bonding.

Both palladium complexes showed a slightly distorted square-planar geometry. The *para*-isomer  $\text{PdCl}_2(\text{MC1})_2$  was obtained as a dimeric structure where complexes are bridged by a carboxylic acid group of one of the phosphines of each of the monomers. The other carboxylic phosphine interacts through H-bonding with a molecule of acetone, Fig. 4. The *meta*-isomer  $\text{PdCl}_2(\text{MC2})_2$  also forms a dimeric structure. However, in this case, both carboxylic phosphines participate in the H-bond network, resulting in a cage type structure, Fig. 5. Inside this cage, badly disordered solvent molecules were observed, in the shape of residual peaks, fenced by the aromatic rings with the trifluoromethyl substituents.

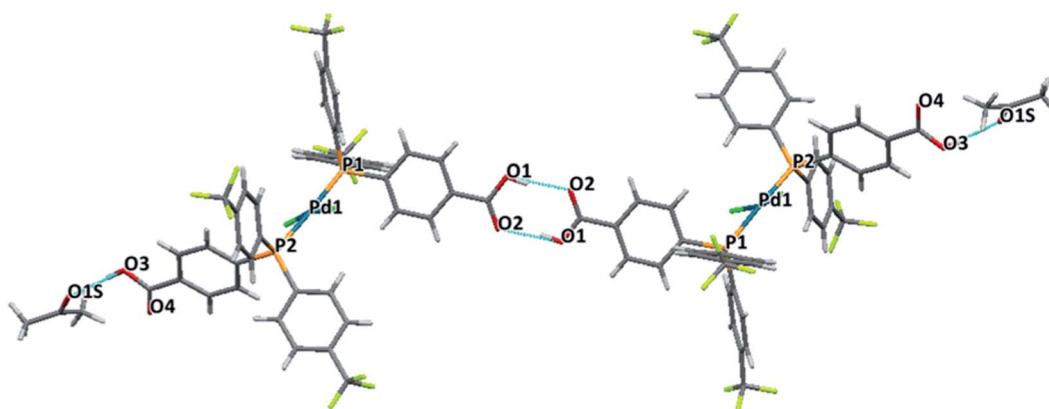


Fig. 4 Stick representation of the dimeric structure of the complex of *trans*- $[\text{PdCl}_2(\text{MC1})_2] \cdot \text{acetone}$ . Only one conformation of the disordered  $\text{CF}_3$  is shown. Hydrogen bond distances (Å): O1–O2, 2.612(3); O3–O15, 2.708(5).



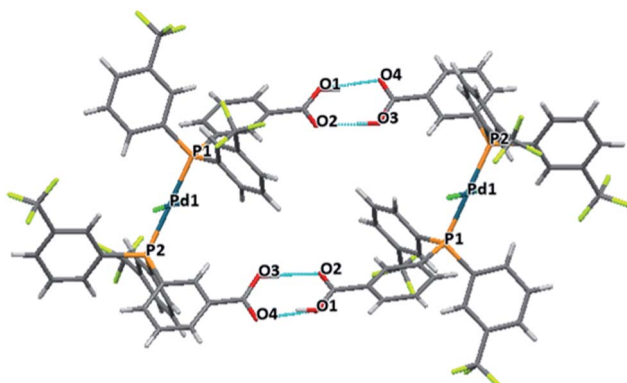


Fig. 5 Stick representation of the dimeric structure of the complex of *trans*-[PdCl<sub>2</sub>(MC2)<sub>2</sub>]. Only one conformation of the disordered CF<sub>3</sub> is shown. Hydrogen bond distances (Å): O1–O4, 2.673(9); O2–O3, 2.608(8).

## Conclusions

The reaction of hydrolysis of the trifluoromethyl group using fuming sulfuric acid and H<sub>3</sub>BO<sub>3</sub> allows the preparation of carboxylic and trifluoromethylated triarylphosphines. The reactivity of the phosphines towards the hydrolysis of the trifluoromethyl group depends on the relative position of this group with respect to the phosphorus atom, with *meta*-position showing the fastest rate. The number of trifluoromethyl groups present in the phosphines also influences the rate of the reaction. Lower reactivity was found when increasing the number of trifluoromethyl groups. The reaction mechanism of the hydrolysis of the trifluoromethylated phosphines in the conditions described in the present work has been proposed based on *in situ* NMR experiments. The differences observed in the rate of the reaction between the *para*- and the *meta*-substituted phosphines are consistent with the formation of a difluorobenzyl carbocation in the rate determining step of the reaction.

The separation and purification of the phosphines was carried out by a pH-dependent procedure, based on the amphiphilic character of the carboxylic moiety. The reaction has been successfully optimized to prepare seven carboxylic and trifluoromethylated phosphines in fair to good yields.

The  $\sigma$ -donation properties of the carboxylic-trifluoromethylated phosphines was studied by means of the <sup>31</sup>P{<sup>1</sup>H} NMR spectra of the selenide derivatives. Both carboxylic and trifluoromethyl groups decrease the  $\sigma$ -basicity character of the phosphines although, as expected, trifluoromethyls have a larger influence. The correlation between the experimental values of <sup>1</sup>J<sub>PSe</sub> and the number of carboxylic and trifluoromethyl groups was empirically studied allowing the prediction of the  $\sigma$ -donation properties of *para*- and *meta*-substituted carboxylic-trifluoromethylated triarylphosphines.

The *trans*-PdCl<sub>2</sub>(MC)<sub>2</sub> complexes of carboxylic phosphines MC1 and MC2 were synthesised. Their crystal structures show the diversity of framework that complexes of these ligands can produce.

## Experimental

### General procedure and reagents

Unless otherwise mentioned, all the synthetic manipulations were performed under nitrogen atmosphere using standard Schlenk techniques.<sup>34</sup> Solvents and liquid reagents were deoxygenated by bubbling nitrogen for 10–15 minutes. General chemical reagents and solvents were purchased from commercial sources and, otherwise specified were used without further purification. Anhydrous solvents and reagents were used when required and were dried by using different standard procedures.<sup>35</sup> The NMR spectra were recorded in a Bruker Avance 250 and a Bruker Avance III 400 at the *Servei de Ressonància Magnètica Nuclear* at the UAB. Chemical shifts in <sup>1</sup>H and <sup>13</sup>C NMR are relative to the signal of the solvent.<sup>36</sup> In <sup>31</sup>P NMR chemical shifts are relative to 85% H<sub>3</sub>PO<sub>4</sub> (0.0 ppm), used as external standard. For <sup>19</sup>F NMR and <sup>11</sup>B NMR experiments the signals are relative to fluorene (–113.15 ppm with respect to CFCl<sub>3</sub>) and BF<sub>3</sub>·Et<sub>2</sub>O (0.0 ppm), respectively. Unless otherwise stated, all the NMR experiments were performed at 298 K. High resolution mass-spectrometry analyses and elemental analyses (C and H) were performed by the *Servei d'Anàlisi Química* at UAB in a Bruker microTOF spectrometer coupled to an Apollo II electrospray source (ESI) and a Thermo Fisher Scientific Flash EA 2000 CHNS, respectively.

Synthesis of the trifluoromethylated triarylphosphines SP1–SP7 was performed according to the previously reported methodologies in the literature.<sup>19,37</sup>

### Synthesis of carboxylic phosphines<sup>38</sup>

4-(Bis(4-(trifluoromethyl)phenyl)phosphanyl)benzoic acid (MC1) and 4,4'-((4-(trifluoromethyl)phenyl)phosphanediy) dibenzoic acid (DC1). In a typical reaction, 3.5 g of tris(4-trifluoromethylphenyl)phosphine (SP1) (7.5 mmol) were dissolved in 15 ml of a solution of boric acid (2.3 g, 37.2 mmol) in 98% sulfuric acid at 0 °C. A strong red or orange solution was obtained once the phosphine was completely dissolved. At this point, the solution was kept at 0 °C and 32.5 ml of 65% SO<sub>3</sub> fuming sulfuric acid (0.53 mol SO<sub>3</sub>) were slowly added during 30 minutes with vigorous stirring. The dark brown mixture was then further stirred at room temperature for 14 h to give a mixture of *ca.* 42% of monocarboxylic phosphine MC1, 50% of dicarboxylic DC1 and 8% of unreacted phosphine SP1. It is advisable to monitor the evolution of the reaction by <sup>31</sup>P{<sup>1</sup>H} NMR since the rate of the reaction is fairly sensitive to the concentration of SO<sub>3</sub> in oleum. The reaction mixture was then gently poured on 150 g of deoxygenated ice, which was prepared by freezing 150 ml of deoxygenated water. The resulting brown slurry was extracted 3 times with 100 ml of Et<sub>2</sub>O. The organic phases were combined and washed 2 times with 50 ml of brine. The separation of the phosphine compounds was then carried out by means of extractions in water at different pH. Firstly, the volume of organic solvent was reduced to *ca.* 100 ml and 100 ml of water were added. The aqueous phase was brought to pH 6.0 by addition of 5% aqueous NaOH, while stirring the biphasic mixture. Once the phases were separated, the pH of the aqueous





phase was measured and corrected if necessary. At this pH, both the oxide of phosphine **DC1** and the tricarboxylic phosphine **TC1**, if present, are extracted as sodium salts in the aqueous phase which was therefore discarded. Next fresh water (100 ml) was added to the organic phase and the pH of the aqueous solution was brought to 7.6, following the procedure described above. At this pH, phosphine **DC1** is extracted into the aqueous phase as sodium salt. After decantation, the organic phase was extracted again with 50 ml of water at pH 7.6. The organic layer was kept under nitrogen in a round-bottom flask for later treatment. The aqueous phases were combined and, after acidification to *ca.* pH 2 with 10% aqueous  $\text{H}_2\text{SO}_4$ , phosphine **DC1** was extracted in the acidic form with 3 portions of 30 ml of  $\text{Et}_2\text{O}$ . The organic layers were dried over  $\text{MgSO}_4$  and the solvent was vacuum evaporated to give 1.1 g (35% yield) of the dicarboxylic phosphine (**DC1**) in its hemihydrate form, as a white solid in purity higher than 95%. The organic phase separated at pH 7.6, which contained the monocarboxylic phosphine **MC1** and the starting phosphine **SP1**, was vacuum evaporated and treated with a solution of NaOH (5%, in methanol) until complete neutralisation of the carboxylic group of **MC1**. This reaction can be monitored by TLC in silica. Then, the solvent was vacuum evaporated to dryness and the resulting solid residue was extracted with cyclohexane to remove phosphine **SP1**. Best results were obtained when the extraction was performed in a Soxhlet extractor overnight. After evaporation of the cyclohexane, 0.2 g of **SP1** were obtained (6% recovered), which can be used again in a new reaction. The solid residue that remained after the extraction with cyclohexane, corresponding to the sodium salt of **MC1**, was suspended in a biphasic mixture of 30 ml of water and 30 ml of  $\text{Et}_2\text{O}$ . After acidification to *ca.* pH 2 with 10% aqueous  $\text{H}_2\text{SO}_4$ , **MC1** was extracted to the organic phase in its acid form. After decantation, the aqueous phase was further extracted 2 times with 30 ml of  $\text{Et}_2\text{O}$ . The organic layers were dried over  $\text{MgSO}_4$  and finally the solvent was vacuum evaporated to give 1.40 g (42% yield) of the carboxylic phosphine (**MC1**), as white solid, in purity above 95%.

**Data for MC1.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (101.27 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm):  $-5.27$  (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 8.10 (dd, 2H,  $\text{H}_{\text{C}_3}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 8.5$  Hz,  $^4J_{\text{HP}} = 1.5$  Hz); 7.63 (d, 4H,  $\text{H}_{\text{C}_3}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = 8.0$  Hz); 7.41 (m, 6H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}-\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 171.7 (s,  $\text{CO}_2\text{H}$ ); 142.77 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 14.4$  Hz); 140.4 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CP}} = 14.2$  Hz); 134.20 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 20.2$  Hz); 133.7 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 19.6$  Hz); 131.7 (q,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 32.6$  Hz); 130.5 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.9$  Hz); 130.2 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 125.8 (dq,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 7.4$  Hz,  $^3J_{\text{CF}} = 3.6$  Hz); 125.5 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 272.5$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (376.50 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm):  $-62.27$  (s). HR-MS ( $\text{ESI}^- m/z$ ) [ $\text{M}-\text{H}^-$ ]: calculated for  $[\text{C}_{21}\text{H}_{12}\text{F}_6\text{O}_2\text{P}]^-$  441.0485; found 441.0486. Elemental analysis: calculated for  $\text{C}_{21}\text{H}_{13}\text{F}_6\text{O}_2\text{P}$ : C, 57.03; H, 2.96; found: C, 57.40; H, 3.06.

**Data for DC1.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (101.27 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm):  $-4.66$  (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 8.0 (bd, 4H,  $\text{H}_{\text{C}_3}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 8.3$  Hz); 7.7 (d, 2H,  $\text{H}_{\text{C}_3}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = 8.1$  Hz); 7.5 (*pseudo-t*, 2H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.7$  Hz); 7.4 (*bt*, 2H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.7$  Hz).  $^{13}\text{C}\{^1\text{H}\}$

NMR (100.61 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 169.2 (s,  $\text{CO}_2\text{H}$ ); 142.8 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 13.7$  Hz); 142.7 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CP}} = 14.2$  Hz); 135.3 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 20.0$  Hz); 134.7 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 20.1$  Hz); 132.9 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{COOH}\}$ ); 132.2 (q,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 32.7$  Hz); 131.0 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 7.0$  Hz); 126.5 (dq,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 7.0$  Hz,  $^3J_{\text{CF}} = 3.7$  Hz); 125.4 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 271.2$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (235.39 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm):  $-62.0$  (s). HR-MS ( $\text{ESI}^- m/z$ ) [ $\text{M}-\text{H}^-$ ]: calculated for  $[\text{C}_{21}\text{H}_{13}\text{F}_3\text{O}_4\text{P}]^-$  417.0509; found 417.0503. Elemental analysis: calculated for  $\text{C}_{21}\text{H}_{14}\text{F}_3\text{O}_4\text{P} \cdot 1/2\text{H}_2\text{O}$ : C, 59.03; H, 3.54; found: C, 59.19; H, 3.36.

**3-(Bis(3-(trifluoromethyl)phenyl)phosphanyl)benzoic acid (MC2) and 3,3'-(3-(trifluoromethyl)phenyl)phosphanediyl)dibenzoic acid (DC2).** The same procedure and reaction conditions as described for the *para*-substituted phosphines were followed in the hydrolysis of the *meta*-substituted phosphine **SP2**. In this case, when the dicarboxylic phosphine **DC2** was the aimed product of the synthesis, the trifluoromethylated phosphine **SP2** was allowed to react with oleum for 20 h, yielding a mixture composed of 6% of **MC2**, 60% of **DC2** and 20% of the tricarboxylic phosphine **TC2**. If, instead, the monocarboxylic phosphine **MC2** was aimed, the reaction with oleum was stopped after 0.5 h, yielding 48% of unreacted phosphine **SP2**, 41% of phosphine **MC2** and 6% of **DC2**. Reaction times between 0.5 and 20 h gave intermediate results. In an example of reaction 5.5 g of **SP2** (11.8 mmol) and 3.6 g of boric acid (58.2 mmol) were dissolved in 22.8 ml of 98% sulfuric acid. Following 51.0 ml of 65%  $\text{SO}_3$  fuming sulfuric acid (0.8 mol of  $\text{SO}_3$ ) were added and allowed to react for 3 h. After quenching with water, NMR analysis of the mixture showed *ca.* 8% of **SP2**, 35% of **MC2**, 46% of **DC2** and 5% of **TC2**. Finally, after the process of separation described above for the *para*-isomers, 1.8 g of **MC2** (35% yield) and 2.1 g of **DC2** (43% yield) were isolated as white solids.

**Data for MC2.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (101.27 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm):  $-4.93$  (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 8.14 (m, 1H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 8.10 (d, 1H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HP}} = 8.7$  Hz); 7.65 (d, 2H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = 7.7$  Hz); 7.60 (d, 2H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HP}} = 7.9$  Hz); 7.51 (m, 4H,  $\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CF}_3\}-\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}-\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 7.44 (*pseudo-t*, 2H,  $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.1$  Hz).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 171.4 (s,  $\text{CO}_2\text{H}$ ); 138.8 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 17.2$  Hz); 137.4 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CP}} = 14.0$  Hz); 136.8 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 16.1$  Hz); 136.6 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 13.3$  Hz); 135.5 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 23.9$  Hz); 131.5 (qd,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 32.5$  Hz,  $^3J_{\text{CP}} = 8.1$  Hz); 131.4 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 130.4 (dq,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 25.0$  Hz,  $^3J_{\text{CF}} = 3.8$  Hz); 130.15 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 7.7$  Hz); 129.5 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 5.6$  Hz); 129.4 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.1$  Hz); 126.4 (q,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CF}} = 3.6$  Hz); 124.0 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 272.2$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (235.39 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm):  $-62.40$  (s). HR-MS ( $\text{ESI}^- m/z$ ) [ $\text{M}-\text{H}^-$ ]: calculated for  $[\text{C}_{21}\text{H}_{12}\text{F}_6\text{O}_2\text{P}]^-$  441.0485; found 441.0489. Elemental analysis: calculated for  $\text{C}_{21}\text{H}_{13}\text{F}_6\text{O}_2\text{P}$ : C, 57.03; H, 2.96; found: C, 57.15; H, 3.02.

**Data for DC2.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (161.98 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm):  $-4.61$  (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 8.07 (m, 2H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 7.98 (bd, 2H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HP}} = 8.0$  Hz); 7.70 (bd, 1H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = 7.9$  Hz); 7.59 (bt, 1H,

$\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HP}} = 7.7$  Hz); 7.52 (m, 6H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CF}_3\}$ – $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CF}_3\}$ – $\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ – $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 169.08 (s,  $\text{CO}_2\text{H}$ ); 139.7 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 14.7$  Hz); 139.0 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 19.3$  Hz); 138.2 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CP}} = 17.9$  Hz); 137.9 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 12.9$  Hz); 135.8 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 21.9$  Hz); 132.7 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 7.0$  Hz); 132.2 (qd,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 32.2$  Hz,  $^3J_{\text{CP}} = 7.2$  Hz); 131.8 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 130.9 (dq,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 22.4$  Hz,  $^3J_{\text{CF}} = 3.9$  Hz); 130.7 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 5.9$  Hz); 130.2 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.7$  Hz); 127.06 (q,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 3.7$  Hz); 125.33 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 272.2$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (376.50 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): –62.07(s). HR-MS ( $\text{ESI}^- m/z$ )  $[\text{M}-\text{H}]^-$ : calculated for  $[\text{C}_{21}\text{H}_{13}\text{F}_3\text{O}_4\text{P}]^-$  417.0509; found 417.0512. Elemental analysis: calculated for  $\text{C}_{21}\text{H}_{14}\text{F}_3\text{O}_4\text{P}$ : C, 60.30; H, 3.37; found: C, 59.58; H, 3.34.

**Data for TC2.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (101.27 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): –5.07 (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 8.06 (m, 3H,  $\text{H}_{\text{C}_4}$ ); 7.98 (bd, 3H,  $\text{H}_{\text{C}_2}$ ,  $^3J_{\text{HP}} = 8.1$  Hz); 7.52 (m, 6H,  $\text{H}_{\text{C}_5} + \text{H}_{\text{C}_6}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CD}_3\text{OD}$ ),  $\delta$  (ppm): 169.2 (s,  $\text{CO}_2\text{H}$ ); 140.0 (d,  $\text{C}_6$ ,  $^2J_{\text{CP}} = 19.0$  Hz); 138.4 (d,  $\text{C}_1$ ,  $^1J_{\text{CP}} = 12.9$  Hz); 135.8 (d,  $\text{C}_2$ ,  $^2J_{\text{CP}} = 21.9$  Hz); 132.6 (d,  $\text{C}_3$ ,  $^3J_{\text{CP}} = 7.0$  Hz); 131.6 (s,  $\text{C}_4$ ); 130.2 (d,  $\text{C}_5$ ,  $^3J_{\text{CP}} = 6.5$  Hz). HR-MS ( $\text{ESI}^- m/z$ )  $[\text{M}-\text{H}]^-$ : calculated for  $[\text{C}_{21}\text{H}_{13}\text{F}_3\text{O}_4\text{P}]^-$  393.0533; found 393.0535.

**3-((3,5-Bis(trifluoromethyl)phenyl)phenyl)(3-(trifluoromethyl)phenyl)phosphanylbenzoic acid (MC4) and 3,3'-((3,5-bis(trifluoromethyl)phenyl)-phosphanediyldibenzoic acid (DC4).** The same procedure described for the hydrolysis of the trifluoromethyl groups in the homoleptic phosphines **SP1** and **SP2** was followed for the heteroleptic phosphine **SP4**. In a typical experiment, a solution of 2.5 g of phosphine **SP4** (4.7 mmol) and 3.25 g of boric acid (52.6 mmol) in 30 ml of 98% sulfuric acid was cooled down to 0 °C in an ice bath. Following, 37 ml of 65%  $\text{SO}_3$  fuming sulfuric acid (0.6 mol of  $\text{SO}_3$ ) were added dropwise and the mixture was allowed to react for 18.5 h at room temperature. The mixture was then quenched on 180 g of deoxygenated ice and extracted three times with 100 ml of  $\text{Et}_2\text{O}$ . Biphasic extractions at pH 6.7 allows the separation of the oxide of phosphine **DC4** in the aqueous phase while the dicarboxylic phosphine **DC4** is extracted pH 8.5. The monocarboxylic phosphine **MC4** is purified as reported for phosphines **MC1** and **MC2**. After full separation and purification, 0.67 g of **DC4** (28% yield) and 0.70 g of **MC4** (30% yield) were obtained as white solids.

**Data for MC4.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (161.98 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): –4.4 (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 8.17 (d, 1H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 7.4$  Hz); 8.12 (d, 1H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HP}} = 8.1$  Hz); 7.88 (s, 1H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ); 7.70 (m, 3H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ – $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ); 7.63 (d, 1H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HP}} = 8.3$  Hz); 7.55 (m, 2H,  $\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ – $\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ); 7.50 (*pseudo-t*, 1H,  $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.5$  Hz); 7.44 (*pseudo-t*, 1H,  $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.3$  Hz).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): 170.95 (s,  $\text{CO}_2\text{H}$ ); 140.22 (d,  $\text{C}_1\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^1J_{\text{CP}} = 18.2$  Hz); 138.72 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 17.0$  Hz); 136.72 (bd,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CP}} = 15.9$  Hz); 136.25 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CP}} = 13.8$  Hz); 135.68 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 25.3$  Hz); 135.32 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 12.9$  Hz); 133.20 (bdq,

$\text{C}_2\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CP}} = 19.8$  Hz,  $^3J_{\text{CF}} = 3.3$  Hz); 132.29 (qd,  $\text{C}_3\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CF}} = 33.3$  Hz,  $^3J_{\text{CP}} = 6.2$  Hz); 131.95 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 131.71 (qd,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 32.5$  Hz,  $^3J_{\text{CP}} = 8.4$  Hz); 130.65 (dq,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^2J_{\text{CF}} = 26.6$  Hz); 130.45 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 8.9$  Hz); 129.84 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CP}} = 5.8$  Hz); 129.70 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.1$  Hz); 126.97 (q,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^3J_{\text{CF}} = 3.7$  Hz); 123.80 (q,  $\text{CF}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ,  $^1J_{\text{CF}} = 272.8$  Hz); 123.36 (sept,  $\text{C}_4\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^3J_{\text{CF}} = 3.6$  Hz); 123.15 (q,  $\text{CF}_3\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^1J_{\text{CF}} = 272.9$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (376.50 MHz,  $\text{CDCl}_3$ ),  $\delta$  (ppm): –62.88 (s, 3F,  $\text{CF}_3\{\text{C}_6\text{H}_4\text{CF}_3\}$ ), –62.99 (s, 6F,  $\text{CF}_3\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ). HR-MS ( $\text{ESI}^- m/z$ )  $[\text{M}-\text{H}]^-$ : calculated for  $[\text{C}_{22}\text{H}_{11}\text{F}_9\text{O}_2\text{P}]^-$  509.0358; found 509.0361. Elemental analysis: calculated for  $\text{C}_{22}\text{H}_{12}\text{F}_9\text{O}_2\text{P}$ : C, 51.78; H, 2.37; found: C, 51.45; H, 2.49.

**Data for DC4.**  $^{31}\text{P}\{^1\text{H}\}$  NMR (161.98 MHz, acetone- $d_6$ ),  $\delta$  (ppm): –3.50 (s).  $^1\text{H}$  NMR (400.13 MHz, acetone- $d_6$ ),  $\delta$  (ppm): 8.14 (d, 2H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 6.8$  Hz); 8.08 (m, 3H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ – $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 7.91 (d, 2H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^3J_{\text{HP}} = 6.0$  Hz); 7.66 (m, 4H,  $\text{H}_{\text{C}_5} + \text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz, acetone- $d_6$ ),  $\delta$  (ppm): 166.90 (s,  $\text{CO}_2\text{H}$ ); 142.34 (d,  $\text{C}_1\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^1J_{\text{CP}} = 19.2$  Hz); 138.91 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 18.5$  Hz); 136.43 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 12.8$  Hz); 135.68 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 23.8$  Hz); 134.14 (bd,  $\text{C}_2\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CP}} = 20.5$  Hz); 132.41 (qd,  $\text{C}_3\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CF}} = 33.4$  Hz,  $^3J_{\text{CP}} = 5.8$  Hz); 132.33 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 7.6$  Hz); 131.97 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 130.46 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.6$  Hz); 124.18 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 272.6$  Hz); 123.86 (sept,  $\text{C}_4\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^3J_{\text{CF}} = 3.9$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (376.50 MHz, acetone- $d_6$ ),  $\delta$  (ppm): –62.38 (s). HR-MS ( $\text{ESI}^- m/z$ )  $[\text{M}-\text{H}]^-$ : calculated for  $[\text{C}_{22}\text{H}_{12}\text{F}_6\text{O}_4\text{P}]^-$  485.0383; found 485.0385. Elemental analysis: calculated for  $\text{C}_{22}\text{H}_{13}\text{F}_6\text{O}_4\text{P}$ : C, 54.34; H, 2.69; found: C, 54.25; H, 2.82.

**3-(Bis(3,5-bis(trifluoromethyl)phenyl)phosphanyl)benzoic acid (MC5).** Following the same procedure as reported for the monocarboxylic phosphines **MC1** and **MC2**, 2.6 g of the trifluoromethylated phosphine **SP5** (4.3 mmol) and 1.30 g of boric acid (21 mmol) were dissolved in 24 ml of sulfuric acid resulting in the formation of an orange solution. The mixture was cooled in an ice bath and 28 ml of 65%  $\text{SO}_3$  fuming sulfuric acid (0.46 mol of  $\text{SO}_3$ ) were added dropwise. After the complete addition of oleum, the dark mixture was stirred at room temperature for 62 h before being poured on 150 g of deoxygenated ice. The resulting slurry was extracted 4 times with 50 ml of  $\text{Et}_2\text{O}$ . The organic solvent was evaporated giving a white solid. At this point,  $^{31}\text{P}\{^1\text{H}\}$  NMR analysis showed a composition of ca. 65% of **MC5** and 35% of phosphine oxides. Small amount of the starting phosphine **SP5** were also detected by TLC in silica ( $\text{CH}_2\text{Cl}_2$ ). Therefore, the solid was treated with NaOH (5% in methanol) until neutralisation of the carboxylic groups. Following, the solvent was vacuum evaporated and the **SP5** was extracted from the solid with cyclohexane (Soxhlet extractor, overnight). The solid residue containing the carboxylic phosphine was mixed with 20 ml of water and 20 ml of diethyl ether and the pH was brought to 5–6. The two phases were then decanted and the aqueous phase was extracted with 2 portions of 20 ml of  $\text{Et}_2\text{O}$ . The organic layers were dried over  $\text{MgSO}_4$  and the solvent was vacuum evaporated to afford 0.88 g



of the carboxylic phosphine **MC5** (35% yield).  $^{31}\text{P}\{^1\text{H}\}$  NMR (161.98 MHz,  $\text{CDCl}_3$ , at 323 K),  $\delta$  (ppm):  $-3.49$  (s).  $^1\text{H}$  NMR (400.13 MHz,  $\text{CDCl}_3$ , at 323 K),  $\delta$  (ppm): 8.22 (d, 1H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 7.3$  Hz); 8.14 (d, 1H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HP}} = 8.5$  Hz); 7.93 (s, 2H,  $\text{H}_{\text{C}_4}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ); 7.73 (d, 4H,  $\text{H}_{\text{C}_2}\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^3J_{\text{HP}} = 6.6$  Hz); 7.59 (bt, 1H,  $\text{H}_{\text{C}_5}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = 7.3$  Hz); 7.52 (bt, 1H,  $\text{H}_{\text{C}_6}\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{HH}} = ^3J_{\text{HP}} = 7.3$  Hz).  $^{13}\text{C}\{^1\text{H}\}$  NMR (100.61 MHz,  $\text{CDCl}_3$ , at 323 K),  $\delta$  (ppm): 170.31 (s,  $\text{CO}_2\text{H}$ ); 139.17 (d,  $\text{C}_1\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^1J_{\text{CP}} = 18.3$  Hz); 138.66 (d,  $\text{C}_6\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 17.1$  Hz); 135.83 (d,  $\text{C}_2\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^2J_{\text{CP}} = 26.0$  Hz); 134.40 (d,  $\text{C}_1\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^1J_{\text{CP}} = 13.1$  Hz); 133.31 (bd,  $\text{C}_2\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CP}} = 20.5$  Hz); 132.86 (qd,  $\text{C}_3\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^2J_{\text{CF}} = 33.6$  Hz,  $^3J_{\text{CP}} = 6.3$  Hz); 132.50 (s,  $\text{C}_4\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ); 130.85 (d,  $\text{C}_3\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 8.7$  Hz); 130.05 (d,  $\text{C}_5\{\text{C}_6\text{H}_4\text{CO}_2\text{H}\}$ ,  $^3J_{\text{CP}} = 6.3$  Hz); 123.91 (sept,  $\text{C}_4\{\text{C}_6\text{H}_3(\text{CF}_3)_2\}$ ,  $^3J_{\text{CF}} = 3.7$  Hz); 123.10 (q,  $\text{CF}_3$ ,  $^1J_{\text{CF}} = 273.8$  Hz).  $^{19}\text{F}\{^1\text{H}\}$  NMR (376.50 MHz,  $\text{CDCl}_3$ , at 323 K),  $\delta$  (ppm):  $-63.36$  (s). HR-MS ( $\text{ESI}^- m/z$ )  $[\text{M}-\text{H}]^-$ : calculated for  $[\text{C}_{23}\text{H}_{10}\text{F}_{12}\text{O}_2\text{P}]^-$  577.0232; found 577.0251. Elemental analysis: calculated for  $\text{C}_{23}\text{H}_{11}\text{F}_{12}\text{O}_2\text{P}$ : C, 47.77; H, 1.96; found: C, 47.40; H, 1.97.

### X-ray crystallography

Crystallographic data for *trans*- $[\text{PdCl}_2(\text{MC1})_2]$ -acetone and *trans*- $[\text{PdCl}_2(\text{MC}_2)_2]$  are provided free of charge from the Cambridge Crystallographic Data Centre with deposition numbers 2097178 and 2106238, respectively.

Crystallographic details about the two structures can be found in the ESI.†

## Conflicts of interest

There are no conflicts to declare.

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