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Unveiling the charge transfer behavior within ZSM-5 and carbon nitride composites for enhanced photocatalytic degradation of methylene blue

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ZSM-5/graphitic carbon nitride (g-C₃N₄) composites were successfully prepared using a simple solvothermal method. By varying the amount of ZSM-5 and g-C₃N₄ in the composites, the charge carrier (electrons and holes) transfer within the materials, which contributes to the enhanced photocatalytic performance, was unraveled. The X-ray diffraction (XRD), Fourier-transform infrared (FTIR), and scanning electron microscopy (SEM) analysis revealed that more ZSM-5 component leads to a stronger interaction with g-C₃N₄. The photocatalytic performance test toward methylene blue (MB) degradation shows that more ZSM-5 in the composites is beneficial in enhancing photocatalytic activity. Meanwhile, the impedance electron spectroscopy (EIS) and photoluminescence (PL) analysis revealed that ZSM-5 facilitates the charge carrier transfer of photogenerated electrons and holes from g-C₃N₄ to the catalyst surface due to its lower charge transfer resistance. During the charge carrier migration, the interface between g-C₃N₄ and ZSM-5 particles may induce higher resistance for the charge carrier transfer, however after passing through the interface from g-C₃N₄ to ZSM-5 particles, the charge carrier can be efficiently transferred to the surface, hence suppressing the charge carrier recombination.

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Introduction

Dye waste is considered as an environmental pollutant and has negative impacts on the environment and human health.^{1,2} Therefore, the removal of dye molecules from the water body becomes crucial. Several water-treatment processes have been employed for dye removals, such as filtration, adsorption, and chemical deposition.^{3–5} However, those processes cannot totally remove the dye molecules since they only encapsulate the dye molecules into certain materials, for example, in the adsorbents or the membranes. Hence, further treatments are still needed. In this regard, photocatalysis reaction has attracted considerable attention^{6,7} since the photocatalytic process is able to degrade the dye molecules into CO₂, water, or other less harmful ions or molecules as the final products.⁸ Thus, further treatments are not needed, which is beneficial in terms of practicality and the economy.

Several semiconductors have been employed as photocatalyst materials for the photodegradation of organic pollutants. For example, WO₃ was reported to be an efficient photocatalyst due to its tunable bandgap, unique band positions, and excellent electronic properties that promoted its stability and photoactivity.⁹ Generating heterojunction in WO₃, such as in Bi₂WO₆/Fe₂O₃/WO₃ S-scheme ternary system, will further enhance the performance toward bisphenol-A degradation due to the efficient charge carriers separation and its ability to maintain redox reaction.¹⁰ Another material such as Cu_{2–x}Se microsheet was reported to be an efficient photocatalyst for rhodamine B degradation, showing a complete discoloration within 30 minutes due to the highly accessible active sites and wider light absorption ranges as a result of the narrower bandgap.¹¹ In terms of rhodamine B degradation, recently MoS₂/tourmaline composite was also reported to have enhanced performance due to the increased dispersion of MoS₂ nanosheet on the tourmaline support, which exposed more active sites. Moreover, the strong electric field in the tourmaline surface suppressed the recombination of charge carriers.¹²

The aforementioned photocatalysts are based on metal catalysts. Nowadays, nonmetal photocatalysts were also obtained considerable attention due to their synthesis practicality and abundant source. Among nonmetal photocatalysts, graphitic carbon nitride (g-C₃N₄) has been considered as one of the promising materials for photocatalytic process¹³ because it has a narrow bandgap (~2.7 eV), good chemical and thermal stability,

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nontoxicity, and relatively easy large-scale production with readily raw materials.^{14–16} Recently, $g\text{-C}_3\text{N}_4$ is also considered as an emerging material for biomedical applications, such as for bio-sensing, diagnosis, therapy, and antibacterial materials due to its good biocompatibility and unique fluorescence properties.¹⁷ However, pristine $g\text{-C}_3\text{N}_4$ still suffers from the rapid charge carriers (electrons and hole pairs) recombination. Hence, some modifications can be applied to suppress this limitation, such as by preparing porous $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4$ -based composites. Porous $g\text{-C}_3\text{N}_4$ was reported to enhance photocatalytic activity due to the increased surface area and narrowed bandgap, which shortened the charge carriers transfer and improves the light absorption.^{18–20} Moreover, the preparation of $g\text{-C}_3\text{N}_4$ -based composites can efficiently suppress the charge carriers recombination.^{21,22} Among supported materials, ZSM-5 is considered as an appropriate support for $g\text{-C}_3\text{N}_4$ due to its chemical, mechanical, and thermal stability, superior adsorption ability, large surface area, and unique pore structure. Furthermore, ZSM-5 has also been reported as suitable photocatalyst support due to its ability to delocalize the bandgap energy of the excited electrons and to decrease the recombination rate.^{23–25}

Some researchers have utilized ZSM-5 as support material for $g\text{-C}_3\text{N}_4$. For example, Prakash *et al.* reported the use of ZSM-5/ $g\text{-C}_3\text{N}_4$ as a photocatalyst for rhodamine-B and crystal violet degradation.²⁵ Preparing ZSM-5/ $g\text{-C}_3\text{N}_4$ consists of ZSM-5 and $g\text{-C}_3\text{N}_4$ with a mass ratio of 4 : 1; it was observed that the ZSM-5 was decorated on the surface of $g\text{-C}_3\text{N}_4$ particles with strong interaction. Also, the charge recombination can be suppressed as indicated by the lower intensity in the photoluminescence (PL) spectra of ZSM-5/ $g\text{-C}_3\text{N}_4$ compared to bare $g\text{-C}_3\text{N}_4$. The enhanced photocatalytic performance was attributed to the efficient separation of charge carriers and the formation of heterojunction between ZSM-5 and $g\text{-C}_3\text{N}_4$.²⁵ Similar results were also reported by Pham *et al.* in ZSM-5 decorated O-doped $g\text{-C}_3\text{N}_4$.²⁶ Despite the enhanced performance of ZSM-5 and $g\text{-C}_3\text{N}_4$ composites, the mechanism of charge carriers transfer within the composite's components and how the charge carrier recombination can be suppressed are still not well understood.

In this present study, we investigate the charge carrier transfer behavior within ZSM-5/ $g\text{-C}_3\text{N}_4$ composites, which determines their photocatalytic performances. In this regard, three ZSM-5/ $g\text{-C}_3\text{N}_4$ composites with significant differences of ZSM-5 and $g\text{-C}_3\text{N}_4$ composition were prepared to obtain noticeable differences of each composite. The photocatalytic test was performed toward methylene blue (MB) degradation as a reaction model. The study revealed that the photogenerated electrons and holes generated by $g\text{-C}_3\text{N}_4$ migrate to the catalyst surface facilitated by ZSM-5 particles due to their higher conductivity. The facilitated charge carriers transfer by ZSM-5 particles suppressed the charge carriers recombination, hence enhancing the catalytic performance of the ZSM-5/ $g\text{-C}_3\text{N}_4$ composites compared to bare $g\text{-C}_3\text{N}_4$.

Experimental

Materials

The materials used in this work are melamine ($\text{C}_3\text{N}_3(\text{NH}_2)_3$, Sigma-Aldrich), citric acid monohydrate ($\text{C}_6\text{H}_8\text{O}_7 \cdot \text{H}_2\text{O}$, Merck,

ACS grade), methylene blue ($\text{C}_{16}\text{H}_{18}\text{ClN}_3\text{S}$, Merck), sodium hydroxide (NaOH, Merck), acetic acid (CH_3COOH , Merck), ethylene glycol ($(\text{CH}_2\text{OH})_2$, Merck), ZSM-5 (Pingxiang Naiké Chemical), dimethyl sulfoxide (DMSO, $(\text{CH}_3)_2\text{SO}$, Sigma-Aldrich), methanol (CH_3OH , SAP chemicals), isopropyl alcohol (IPA, $(\text{CH}_3)_2\text{CHOH}$, SAP chemicals) and L-ascorbic acid ($\text{C}_6\text{H}_8\text{O}_6$, SAP chemicals) and deionized water.

Procedure

Synthesis of $g\text{-C}_3\text{N}_4$. $g\text{-C}_3\text{N}_4$ used in this work was porous $g\text{-C}_3\text{N}_4$. The synthesis of porous $g\text{-C}_3\text{N}_4$ was performed as reported by Kong *et al.*, using citric acid as a pore-forming agent.¹⁹ In a typical procedure, 2.52 g (0.02 mol) melamine was poured into 40 mL deionized water, then was heated at 80 °C for 30 min under vigorous stirring. Meanwhile, 0.2 mmol citric acid was dissolved in 10 mL deionized water and was dropped into a melamine solution. The mixture was then continuously stirred for 2 h at 80 °C. The resulted suspension was then transferred into a 250 mL Teflon autoclave and was heated in an oven at 180 °C for 12 h. After heating, the powders were filtered and washed with deionized water. The precipitate was then dried at 80 °C for 2 h. The dried powder was calcined at 550 °C for 2 h.

Preparation of ZSM-5/ $g\text{-C}_3\text{N}_4$ composites. Preparation of ZSM-5/ $g\text{-C}_3\text{N}_4$ composites was carried out by solvothermal method using the same procedure as reported by Prakash *et al.*²⁵ Before preparation, the commercial ZSM-5 was firstly modified using the desilication and dealumination processes. The desilication process was performed using NaOH. In brief, a certain amount of ZSM-5 was mixed with 0.25 M NaOH with the ratio between the mass of ZSM-5 to the volume of NaOH was 1 : 30. The mixture was then put into reflux equipment and was heated at 65 °C for 30 min. After reflux, the mixture was directly cooled down in an ice bath. The ZSM-5 powders were filtered and washed with water until reaching neutral pH. The powder was then dried in an oven at 105 °C for 3 h. The dried samples were then treated for the dealumination process with 0.01 M acetic acid with the same procedure as the desilication process. However, for the dealumination process, the ratio between the mass of ZSM-5 and the volume of acetic acid was set at 1 : 20. After the dealumination process, the resulting ZSM-5 powders were calcined at 550 °C for 6 h.² ZSM-5/ $g\text{-C}_3\text{N}_4$ composites were prepared as follows. 0.5 g of $g\text{-C}_3\text{N}_4$ was mixed with various amounts of ZSM-5 (0.0 g, 0.25 g, 0.5 g, and 1.0 g) in 25 mL of ethylene glycol. The mixture was then ultrasonicated for 2 h. The well-dispersed mixture was then transferred into Teflon autoclave and was heated at 120 °C for 8 h. After cooling, the resulting powders were filtered and were dried at 105 °C for 3 h. The resulting samples were then assigned as bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, GZ-3, respectively.

Material characterizations. Crystalline phase of the prepared samples was characterized by X-ray diffraction (XRD). The XRD analysis was performed using XRD PANalytical under Cu K α irradiation ($\lambda = 1.5406$), with accelerating voltage and current of 40 kV and 30 mA. The analysis was carried out at a diffraction angle of 2θ (20–80°) with a scan-interval of 0.017°. The Fourier-transformed infra-red (FTIR) analysis was performed using



Shimadzu IRAFFINITY-s using conventional KBr pellets. The surface morphology was observed using a scanning electron microscope (SEM) ZEISS EVO 10 with an accelerating voltage of 20 kV with a secondary electron (SE1) detector. The elemental composition and distribution were analyzed using energy dispersive X-ray (EDX APEX). The surface area was measured using Brunauer–Emmett–Teller (BET) N₂ sorption Quantachrome NovaWin. The optical properties were characterized by UV diffuse reflectance spectroscopy (DRS) using Shimadzu UV2600-I, in which BaSO₄ was used as the reference. The photoluminescence (PL) spectroscopy was performed using spectrofluorophotometer Shimadzu RF-5301 with an excitation wavelength of 350 nm. The photoelectrochemical measurements and electrochemical impedance spectroscopy (EIS) were performed using the CHI 600E instrument. The samples were dropped cast on the surface of fluorinated-tin oxide (FTO) prior to measurement. The measurements were conducted in 0.5 M Na₂SO₄ under full-spectrum illumination. The EIS measurement was performed at -0.2 V vs. Ag/AgCl at a frequency of 0.1 to 10⁵ Hz.

Photocatalytic test

Photocatalytic activity of samples was evaluated using MB degradation under UV-LED illumination (365 nm, 12 × 2 watt). Typically, 50 mg of the photocatalyst was added to 50 mL of a 20 ppm MB solution. Prior to the MB degradation test, the solution was allowed to stand for 1 hour under stirring in the dark to reach MB adsorption/desorption equilibrium. After 1 hour in the dark, the UV lamp was turned on. 2 mL of aliquot solution was taken periodically every 60 min up to 180 min. The concentration was determined using a UV-Vis spectrophotometer at a wavelength of 655 nm. The degradation rate constant was calculated based on the pseudo-first-order Langmuir–Hinshelwood equation as shown in eqn (1).

$$\ln \frac{C_t}{C_0} = kt \quad (1)$$

where C_t is the concentration at time t min (ppm), C_0 is the initial concentration (ppm), t is the reaction time (minutes), and k is the reaction rate constant.¹⁵

The detection of reactive species was performed by trapping experiment using different scavengers, namely DMSO, methanol, IPA, and L-ascorbic acid as scavengers for electrons (e^-), holes (h^+), hydroxyl radicals ($\cdot\text{OH}$), superoxide radicals ($\cdot\text{O}_2^-$), respectively. The experiments were similar to the photodegradation experiments with addition of the corresponding scavenger.¹⁵

Results and discussion

Material characterizations

The crystalline phase and the purity of the samples were characterized using XRD.^{27,28} As shown in Fig. 1, bare g-C₃N₄ exhibits pronounce peaks at 2θ of 13.2° and 27.5°. The peak 13.2° refers to the (100) crystal lattice of g-C₃N₄, which can be attributed to the in-plane repetitive unit of tri-*s*-triazine. The strongest peak

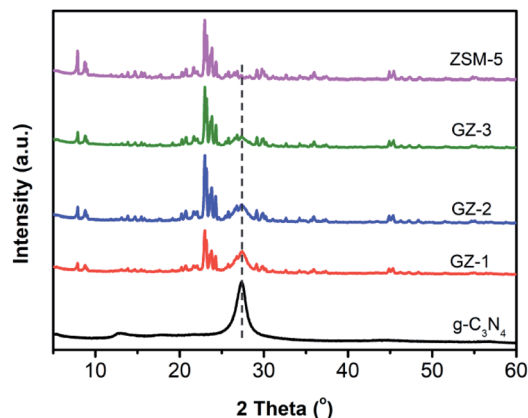


Fig. 1 XRD patterns of bare g-C₃N₄, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts.

at 27.5° corresponds to the (002) crystal plane, attributed to the interplanar stacking peak of the conjugated aromatic rings.²⁹ Based on these XRD patterns, the g-C₃N₄ has been successfully formed with high purity and crystallinity. The XRD patterns of the ZSM-5 display the pronounce peaks at 2θ of 23.3°, 26.4°, 27.5°, and 30.1°, which correspond to the (533), (642), (733), and (555) reflection planes of ZSM-5, respectively. The XRD patterns of the GZ composites show mixed-phases between g-C₃N₄ and ZSM-5. As expected, the intensity of the g-C₃N₄ characteristic peak at 27.3° decreases from GZ-1 to GZ-3 along with the less portion of g-C₃N₄ in the composites.²⁵

The FTIR analysis was then performed to further investigate the chemical composition of the GZ composite. As shown in FTIR spectra (Fig. 2), the GZ composites exhibit a mixed characteristic feature of g-C₃N₄ and ZSM-5 FTIR spectra. A broad peak around 3200 cm⁻¹ can be assigned to the N–H bond stretching from amino groups ($-\text{NH}_2$ or $=\text{NH}$ groups).³⁰ The peaks at 1234, 1319, 1404, 1543, and 1627 cm⁻¹ can be attributed to the vibration modes of the C–N stretching and C=N stretching in the aromatic ring.³¹ The peak at 810 cm⁻¹ corresponds to the out-of-plane skeletal bending modes of the tri-*s*-triazine ring.^{16,32} In the FTIR spectra of ZSM-5, the peak at 450 cm⁻¹ refers to the buckling vibration of T–O–T, where T can be Si or Al atoms. The peak at 546 cm⁻¹ can be observed, attributing to the formation of the pentasil ring by the tetrahedral SiO₄ and AlO₄ units.³³ The symmetric stretching of the Si–O–Si group can be indicated by the peak at 789 cm⁻¹. The peak at 1049 cm⁻¹ and 1218 cm⁻¹ can be attributed to the symmetric stretching of the Si–O–T group from the intra-tetrahedral mode of the zeolite framework.³⁴ Moreover, the slight shifting in some peaks, especially at wavenumber of 1234 cm⁻¹, 1404 cm⁻¹, and 1543 cm⁻¹, can also be observed in the GZ-3 samples, which indicates the strong chemical interaction between ZSM-5 and g-C₃N₄ particles in GZ-3.^{25,26} This result is in good agreement with the XRD results.

The surface morphology of the composites was then observed using SEM.³⁵ As shown in Fig. 3, the bare g-C₃N₄ displays a wrinkled-interplanar structure with sheet-like morphology. The rough surface of the sheet particles



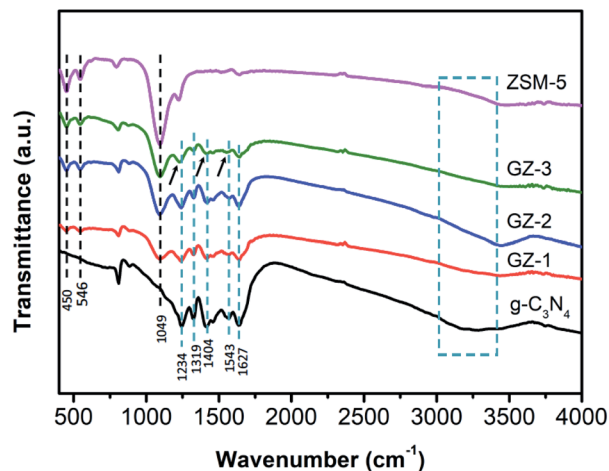


Fig. 2 FTIR spectra of bare $g\text{-C}_3\text{N}_4$, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts.

indicates the formation of pores due to the utilization of acetic acid as a pore-forming agent during $g\text{-C}_3\text{N}_4$ synthesis.¹⁹ Meanwhile, the ZSM-5 shows a cubic-like or prism-like structure with some rough surface due to the desilication and dealumination process. The SEM images of the composite materials exhibit mixed particles between the $g\text{-C}_3\text{N}_4$ and ZSM-5 particles. In particular, the GZ-1 sample shows the dominance of $g\text{-C}_3\text{N}_4$ particles decorated with ZSM-5 particles on the surface. Unfortunately, it seems that the contact area between the ZSM-5 and of $g\text{-C}_3\text{N}_4$ is still limited. The GZ-2 shows more ZSM-5 particles together with $g\text{-C}_3\text{N}_4$ particles. As expected, the GZ-3 samples were dominated by the ZSM-5 particles. In this case, the $g\text{-C}_3\text{N}_4$ particles were surrounded by many ZSM-5 particles, creating more contact area and strong interaction between $g\text{-C}_3\text{N}_4$ and ZSM-5 particles. The more contact area between $g\text{-C}_3\text{N}_4$ and ZSM-5 particles are favorable in terms of the

photogenerated charge carrier transport from $g\text{-C}_3\text{N}_4$ to the surface of the catalyst to further contact with the reactant.²⁵

The elemental composition of GZ composites was also confirmed using EDX analysis (Fig. 4 and Table 1).³⁶ As presented in Fig. 4a and Table 1, all GZ samples composed of major elements such as carbon (C) and nitrogen (N) as representative elements for $g\text{-C}_3\text{N}_4$ as well as silicon (Si), aluminium (Al), and oxygen (O) as representative elements for ZSM-5. The amount of C and N elements decrease from GZ-1 to GZ-3 as the proportion of $g\text{-C}_3\text{N}_4$ decreases (Table 1), which is in a good agreement with the SEM analysis. The elemental mapping of GZ-3 (Fig. 4b–g) show overlapping elemental distribution between elements from $g\text{-C}_3\text{N}_4$ and ZSM-5, which suggest a uniform distribution between the two components.

The surface area measurement was carried out using the N_2 adsorption/desorption method on bare $g\text{-C}_3\text{N}_4$ and ZSM-5. As shown in Fig. 5, bare $g\text{-C}_3\text{N}_4$ and ZSM-5 show typical type IV isotherms with hysteresis loops according to the Brunauer–Deming–Teller classification. The specific surface area for $g\text{-C}_3\text{N}_4$ and ZSM-5 are $13.97 \text{ m}^2 \text{ g}^{-1}$ and $281.97 \text{ m}^2 \text{ g}^{-1}$, respectively. As typical type IV isotherms, both samples possess mesopores structure, with pore radii being distributed around $\sim 15.7 \text{ nm}$ and $\sim 3.6 \text{ nm}$ for $g\text{-C}_3\text{N}_4$ and ZSM-5, respectively. Based on the surface area measurement, the ZSM-5 may increase the specific surface area of the GZ composites, where the more ZSM-5 content in the GZ composites, the higher the surface area.^{25,26} In this regard, the higher specific surface area and the larger mesopores in $g\text{-C}_3\text{N}_4$ makes the GZ composites promising candidates for MB photodegradation with higher efficiency due to the facilitated mass transport and enhanced multilight scattering.^{37,38}

Optical properties

The optical properties of the GZ composites were firstly characterized by UV-vis diffuse reflectance spectroscopy (DRS). As

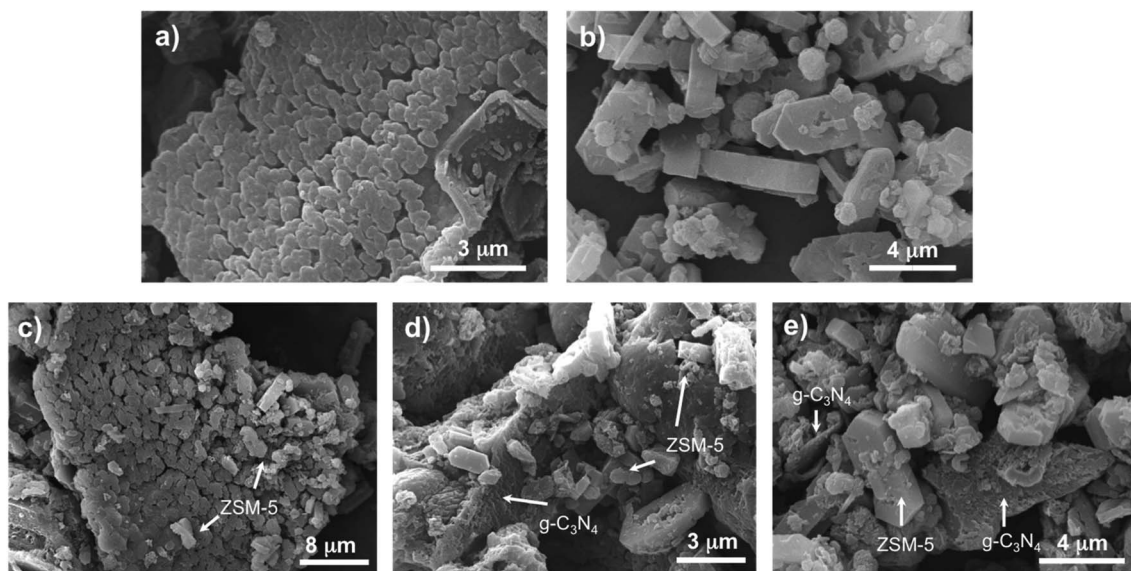


Fig. 3 SEM images of (a) bare $g\text{-C}_3\text{N}_4$, (b) ZSM-5, (c) GZ-1, (d) GZ-2, and (e) GZ-3 catalysts.



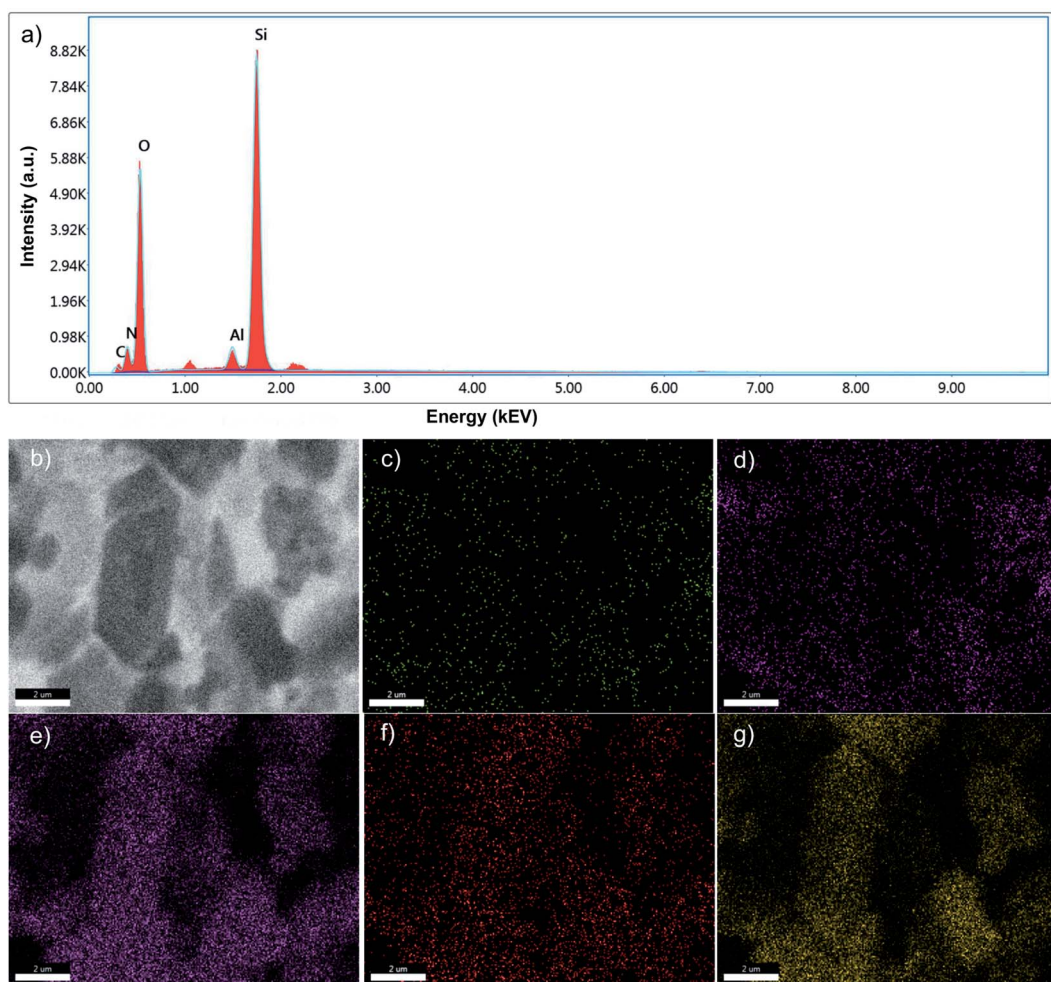


Fig. 4 (a) EDX spectrum and (b) SEM micrograph with elemental mapping of (c) carbon (C), (d) nitrogen (N), (e) silicon (Si), (f) aluminium (Al), and (g) oxygen (O) of the GZ-3 composite. The scale bars represent 2 μm .

Table 1 Elemental composition of GZ composites

Element	GZ-1		GZ-2		GZ-3	
	Weight (%)	Atomic (%)	Weight (%)	Atomic (%)	Weight (%)	Atomic (%)
Carbon (C)	15.28	20.22	8.02	11.91	3.92	6.10
Nitrogen (N)	28.33	32.15	13.66	17.40	9.91	13.21
Silicon (Si)	18.29	10.35	32.58	20.69	36.95	24.56
Aluminium (Al)	1.39	0.82	2.20	1.45	2.73	1.89
Oxygen (O)	36.71	36.46	43.54	48.54	46.48	54.24

shown in Fig. 6a, the DRS spectra of bare $g\text{-C}_3\text{N}_4$ and GZ composites exhibit two absorption peaks at 270 nm and 370 nm. The peak at 270 nm corresponds to the $\pi\text{-}\pi^*$ transition of C=N groups in the conjugated aromatic triazine ring, while the peak at 370 nm can be attributed to the $n\text{-}\pi^*$ transition of C-N terminal in $g\text{-C}_3\text{N}_4$.^{39,40} All samples show strong absorption peaks both in UV and visible regions (400–450 nm). It is worth noting that the ZSM-5 also shows a slight absorption in the UV region. Moreover, the absorption edge of GZ composites exhibits a blue shift compared to the bare $g\text{-C}_3\text{N}_4$. Furthermore,

the bandgap energy of $g\text{-C}_3\text{N}_4$ was then obtained from a plot of $(\alpha h\nu)^{1/2}$ versus $h\nu$ as presented in the Tauc plots in Fig. 6b. A careful examination shows that the tailing edge of the $g\text{-C}_3\text{N}_4$ is 2.73 eV. The conduction band potential (E_{CB}) and valence band potential (E_{VB}) energies were calculated using the Mulliken electronegativity formula, as presented in eqn (2) and (3)

$$E_{\text{VB}} = X - E^0 + 0.5E_{\text{g}} \quad (2)$$

$$E_{\text{CB}} = E_{\text{VB}} - E_{\text{g}} \quad (3)$$



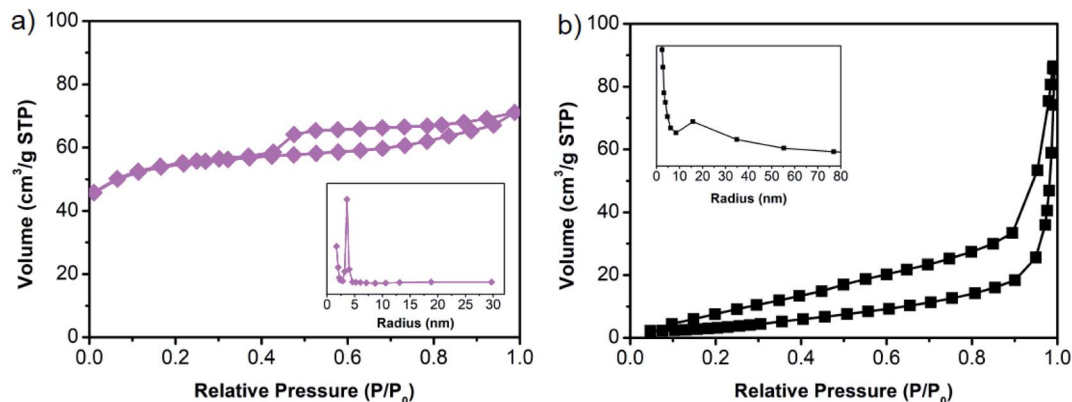


Fig. 5 N_2 adsorption/desorption isotherms of (a) $g-C_3N_4$ and (b) ZSM-5. Inset is the corresponding pore size distribution.

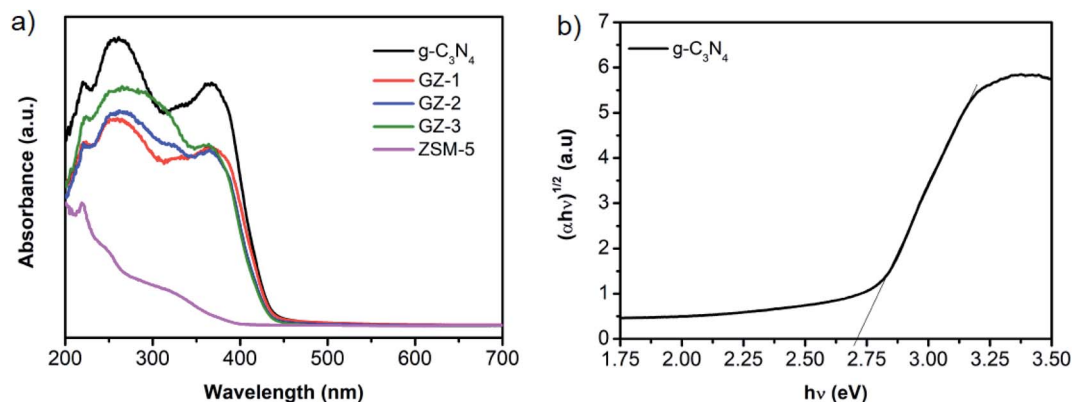


Fig. 6 (a) DRS spectra of bare $g-C_3N_4$, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts. (b) Tauc plot of bare $g-C_3N_4$.

where X is the electronegativity of material ($g-C_3N_4$, 4.73 eV) and E^0 is the free electron energy (4.5 eV vs. NHE). Based on eqn (2) and (3), the E_{CB} and E_{VB} values of GZ-3 are -1.14 V and 1.60 V, which are similar to previous reports.^{15,25}

Photocatalytic activity test

The photocatalytic activity of the photocatalyst materials was examined towards MB degradation. The main objective of this experiment is to find the MB photodegradation trend over various portions of ZSM-5 in bare $g-C_3N_4$ in GZ samples, hence it can indicate the determining component for photocatalytic activity and charge carrier behavior in the catalysts. At the first stage, we need to verify the component of the GZ, whether the $g-C_3N_4$ or ZSM-5, which has a dominant role in determining the performance. To achieve this objective, photodegradation experiments using the same amount of catalyst, *i.e.*, 50 mg, were performed (detail in the experimental section).

Prior to the photocatalytic degradation, the adsorption experiment using 50 mg samples was performed to reveal the adsorption capacity of the photocatalyst toward MB. In this regard, the adsorption capacity of the photocatalyst may differ due to the different proportions of $g-C_3N_4$ and ZSM-5. As presented in Fig. 7, the adsorption capacity of all samples increases rapidly within 30 min of adsorption time, followed by the

relatively stable adsorption capacity up to 90 min. These results suggest that after 30 min, the adsorption process has reached equilibrium. The rapid adsorption process within 30 min was attributed to the presence of mesopores in the samples, which allows the diffusion of MB molecules into the pores.⁴¹ Moreover,

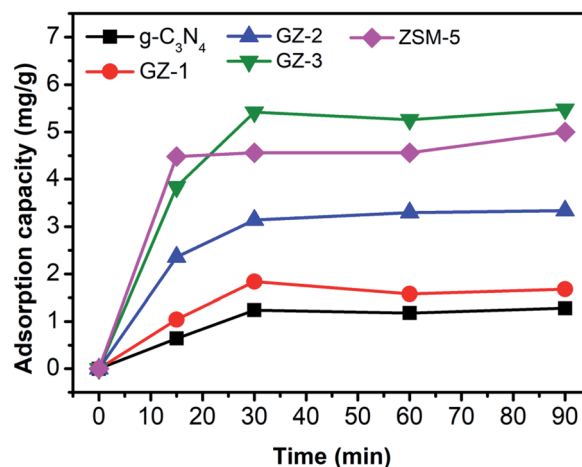


Fig. 7 Adsorption capacity of bare $g-C_3N_4$, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts.



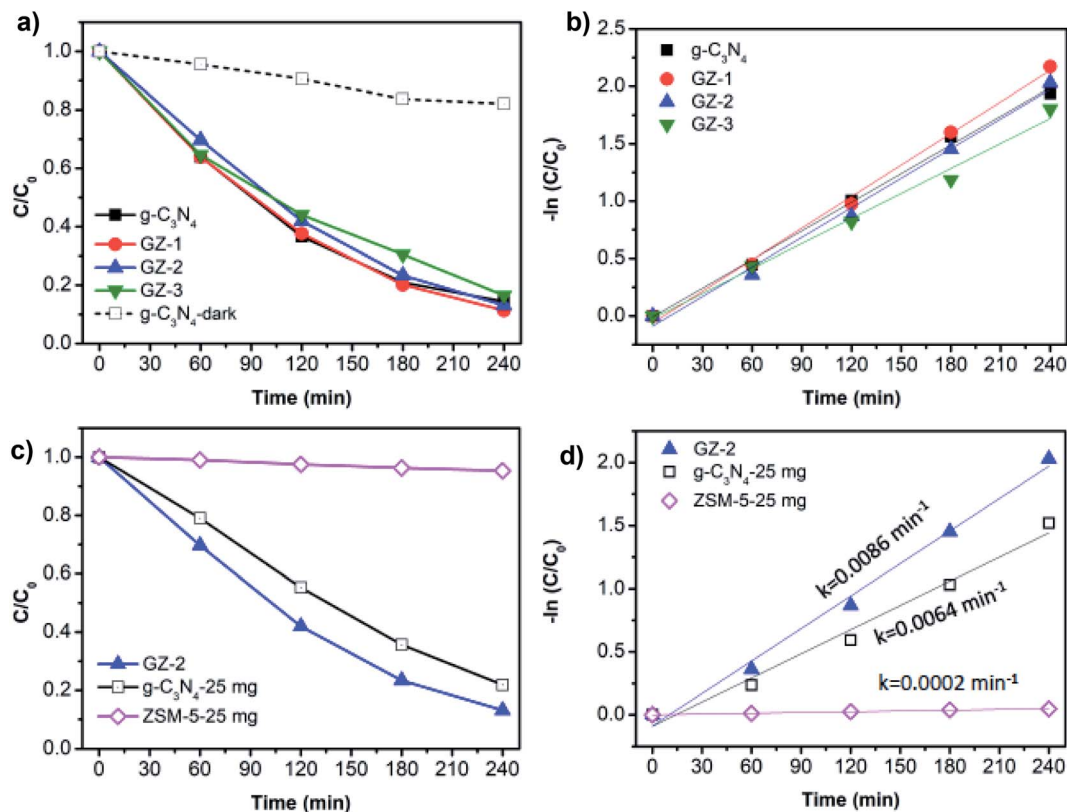


Fig. 8 (a) Photodegradation of MB over 50 mg $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3 catalysts and (b) corresponding Langmuir–Hinshelwood plots. (c) Photodegradation of MB over 50 mg of GZ-2 as a representative sample with the control experiments using the bare $g\text{-C}_3\text{N}_4$ and ZSM-5, and (d) corresponding Langmuir–Hinshelwood plots for the control experiments.

the ZSM-5 seems to be the component, which determines the adsorption ability due to the porous nature. However, when high portion of ZSM-5 is combined with $g\text{-C}_3\text{N}_4$, as in GZ-3, the adsorption capacity was further enhanced, due to the combined effect between the porosity from ZSM-5 and the larger mesopores in $g\text{-C}_3\text{N}_4$. The lesser amount of ZSM-5 lead to the decrease on the adsorption capacity. Since the adsorption reached the equilibrium after 30 min, performing the stirring process for 1 h in dark before irradiating the reactor with UV light is sufficient to exclude the contribution of the adsorption process during the photocatalytic reaction. Hence, after 1 h of stirring in dark, the decrease of MB concentration under UV irradiation is contributed by the photocatalytic process.

Without light irradiation, the photocatalyst is not able to degrade the MB as presented by the degradation curve of MB using bare $g\text{-C}_3\text{N}_4$ in the dark, as a primary photoactive material (Fig. 8a). Under UV light illumination, all materials are able to degrade the MB, in which the photodegradation trends of bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3 samples, are very comparable. Moreover, the degradation rates are also comparable, as indicated by the rate constant values (Fig. 8b and Table 2). Even though the degradation rate of all samples shows a similar trend and values, it does not mean that there is no enhancement in the photocatalytic activity upon the compositing of ZSM-5 and $g\text{-C}_3\text{N}_4$. In fact, it shows an indirect evidence of the enhancement and indicates the determining component which

responsible for catalytic activity. Considering the composition of the catalysts, in 50 mg of GZ composites, the portion of $g\text{-C}_3\text{N}_4$ and ZSM-5 in each sample is different. As tabulated in Table 2, in 50 mg of tested catalysts, the amount of $g\text{-C}_3\text{N}_4$ are 50 mg, 33.3 mg, 25.0 mg, and 16.7 mg (or equal to 100%, 66.6%, 50.0%, and 33.3%) for each bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3, respectively. The remaining components are ZSM-5. Based on this component composition, we know that even with the lower amount of $g\text{-C}_3\text{N}_4$ in GZ-1, GZ-2, and GZ-3, the exhibited photocatalytic performances are still comparable with the pure $g\text{-C}_3\text{N}_4$. These results indirectly indicate the enhancement of photocatalytic activity since the photocatalytic activity is generally decreases along with the significantly lower photocatalyst dose.⁴² Please note that the degradation experiments were started after the adsorption process reached equilibrium

Table 2 The rate constants of various $g\text{-C}_3\text{N}_4$ -based catalysts with their respective $g\text{-C}_3\text{N}_4$ and ZSM-5 composition

Sample	Rate constant (k , min^{-1})	Component portion (%)	
		$g\text{-C}_3\text{N}_4$	ZSM-5
$g\text{-C}_3\text{N}_4$	0.0083	100.0	0.0
GZ-1	0.0092	66.7	33.3
GZ-2	0.0086	50.0	50.0
GZ-3	0.0073	33.3	66.7



(after 1 hour of adsorption in the dark), hence the contribution of the adsorption process can be excluded.

To further verify the contribution of ZSM-5 and the bare $g\text{-C}_3\text{N}_4$ in the GZ composites, we performed the MB photodegradation process using GZ-2 as representative GZ samples. 50 mg of GZ-2 consists of 25 mg of $g\text{-C}_3\text{N}_4$ and 25 mg of ZSM-5. Thus, performing the MB degradation experiment using 25 mg of $g\text{-C}_3\text{N}_4$ and 25 mg of ZSM-5 can give information on each component contribution and how much the contribution is. As presented in Fig. 8c and d, the 25 mg bare $g\text{-C}_3\text{N}_4$ exhibits a slower MB degradation rate compared to the GZ-2. Moreover, the MB degradation over 25 mg of ZSM-5 is negligible. These results verify that the determining component for the photocatalytic activity is the $g\text{-C}_3\text{N}_4$. It is reasonable considering the nature of $g\text{-C}_3\text{N}_4$, which can generate electrons and holes upon light irradiation. In the photocatalytic process, the photo-generated electrons and holes are two species that initiate the photodegradation process *via* direct or indirect degradation by forming radicals from water ($\cdot\text{OH}$ and $\cdot\text{O}_2^-$). In this regard, the ZSM-5 can serve as the support for $g\text{-C}_3\text{N}_4$, where the supporting mechanism will be discussed in the following sections.

Unveiling that the $g\text{-C}_3\text{N}_4$ is the determining component in GZ composites, we then performed the MB photodegradation using the same amount of $g\text{-C}_3\text{N}_4$ component for each sample. In this regard, the mass of total photocatalyst is 50 mg, 75 mg,

100 mg, and 150 mg for bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3, since all samples will have 50 mg of $g\text{-C}_3\text{N}_4$. In this case, the influence of the ZSM-5 amount in the GZ-composites can be investigated. The portion of ZSM-5 increases in the following order: GZ-3 > GZ-2 > GZ-1. The MB photodegradation plots of all samples are depicted in Fig. 9a. GZ-2 and GZ-3 show enhancement of the MB degradation rate compared to GZ-1 and bare $g\text{-C}_3\text{N}_4$. Surprisingly, the GZ-1 shows a comparable MB degradation with bare $g\text{-C}_3\text{N}_4$. Based on these results, the addition of more ZSM-5 into $g\text{-C}_3\text{N}_4$ is favorable for enhancing photocatalytic performance. The low amount of ZSM-5 as the support seems to have a small effect on the catalytic performance. Among the three GZ composites, the GZ-3 shows the highest MB degradation rate. The same trends are also observed and become more obvious using the higher initial concentration of MB (*i.e.* 40 ppm) as presented in Fig. 9c and d. To verify these results, the control experiments were conducted (Fig. 10a and b). The control experiment performed over 150 mg GZ-3 in the dark shows a negligible MB concentration decrease. Since the GZ-3 contains 100 mg of ZSM-5, the control experiment using 100 mg of ZSM-5 under light illumination was also performed and the result shows significantly slower MB concentration changes compared to the GZ-3. All of these results confirm that the MB degradation is mainly attributed to the activity of GZ-3, due to the synergistic effect between ZSM-5 and $g\text{-C}_3\text{N}_4$.

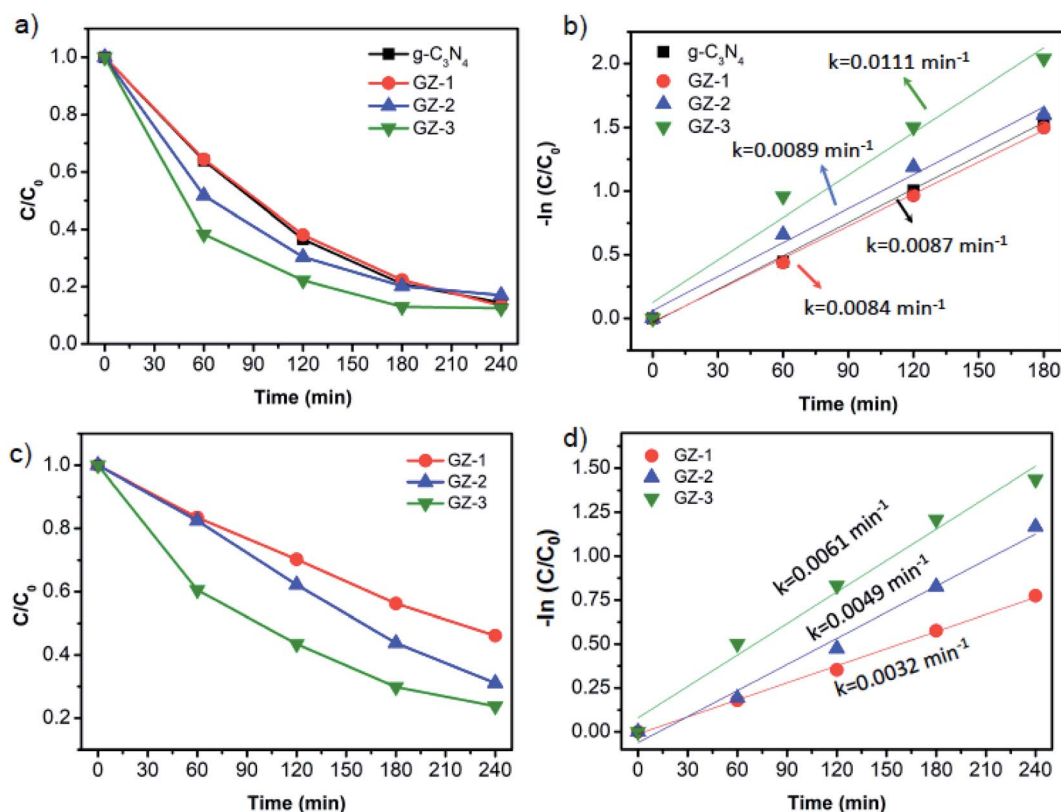


Fig. 9 (a) Photodegradation of MB over bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3 catalysts with a fixed amount of $g\text{-C}_3\text{N}_4$ as photoactive material in 20 ppm of MB, and (b) corresponding Langmuir-Hinshelwood plots. (c) Photodegradation of MB over bare $g\text{-C}_3\text{N}_4$, GZ-1, GZ-2, and GZ-3 catalysts with a fixed amount of $g\text{-C}_3\text{N}_4$ as photoactive material with a higher initial concentration of MB (40 ppm), and (d) corresponding Langmuir-Hinshelwood plots.



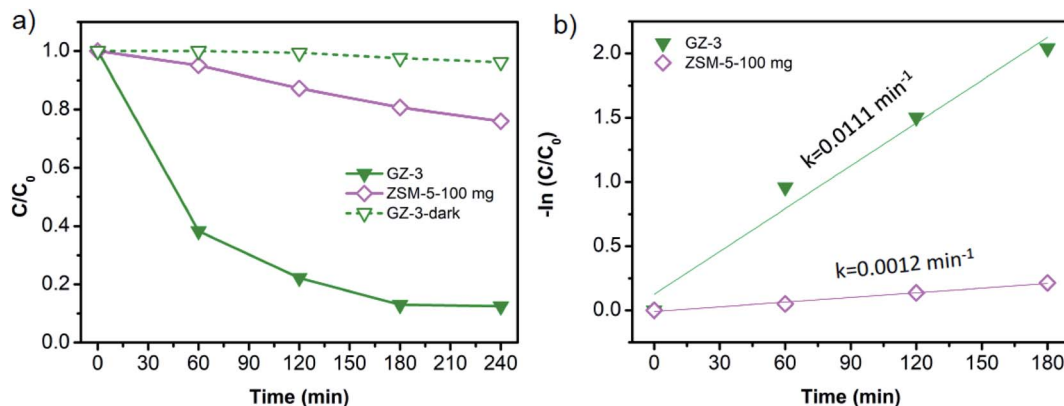


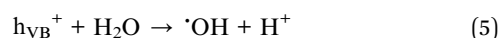
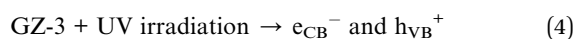
Fig. 10 (a) Photodegradation of MB and (b) corresponding Langmuir–Hinshelwood plots over 150 mg of GZ-3 with respective control experiments. The control experiments were performed in the dark and using 100 mg of ZSM-5 under UV illumination. The MB initial concentration is 20 ppm.

Apart from the synergistic effect between the $g\text{-C}_3\text{N}_4$ and ZSM-5 in GZ-3, the enhanced photocatalytic performance can also be attributed to the higher specific surface area of GZ-3 due to its higher ZSM-5 content. The higher specific surface area and the abundant mesopores in GZ-3 both originated from ZSM-5 and $g\text{-C}_3\text{N}_4$ enhance the multilight scattering. In this case, the large mesopores, especially in $g\text{-C}_3\text{N}_4$, induce a multiple-reflection effect on incident light, which facilitates the light transfer onto inner surface of GZ-3. As a result, the light absorption is improved.^{18,37,38} In addition to the enhanced multilight scattering, the large mesopores also facilitates fast mass transport.^{43–47}

Mechanism of photocatalytic activity

The trapping experiments were performed to identify the reactive species, which are responsible for the MB photodegradation process. For this purpose, different scavengers such as DMSO, methanol, IPA, and L-ascorbic acid were added into GZ-3 as scavengers for electrons (e^-), holes (h^+), hydroxyl radicals ($\cdot\text{OH}$), $\cdot\text{O}_2^-$, respectively. As shown in Fig. 11a, comparable degradation rates as in reaction without scavenger can still be observed by the addition of methanol and L-ascorbic acid, suggesting that the h^+ and $\cdot\text{O}_2^-$ do not play a significant role in MB photodegradation. In contrast, the MB photodegradation rate significantly decreases with the addition of IPA into the reaction system and moderately decreases with the addition of DMSO. These results indicate that during the photocatalytic reaction $\cdot\text{OH}$ was produced and became the reactive species together with e^- on the MB photodegradation process. This result is in good agreement with the previous report.²⁵

Based on the aforementioned results, the reaction mechanism can be predicted. When the GZ-3 composite was irradiated by UV light, the electrons (e^-) are excited from valence band (VB) to conduction band (CB) with immediately generated h^+ in the VB (eqn (4)). The h^+ reacts with water to produce $\cdot\text{OH}$ (eqn (5)).²⁵ The $\cdot\text{OH}$ and e^- species finally involves in MB degradation as presented in eqn (6) and (7).



Charge carriers behavior

The photocatalytic activity depends on the charge carriers (electrons and holes) transfer and recombination. Therefore, to investigate the charge carriers behavior, the photoelectrochemical measurements and steady-state PL analysis were performed. Fig. 11b exhibits the amperometric $I-t$ curves of the samples coated on the surface of FTO glass in several on-off cycles. The exhibited photocurrent reflects the ability of the material to generate charge carriers under light illumination.^{48,49} All $g\text{-C}_3\text{N}_4$ -based samples exhibit obvious photoresponse under full-spectrum illumination. Interestingly, ZSM-5 also shows slight photoresponse, which agrees with the UV-visible DRS spectra showing the absorbance in the UV region and rationalizes the slight MB degradation as presented in Fig. 10. While the bare $g\text{-C}_3\text{N}_4$, GZ-1, and GZ-2 show a comparable trend, the GZ-3 displays the highest photocurrent density, suggesting that the photoresponse ability of GZ-3 is improved and more charge carriers are generated and transferred.^{50,51}

To further investigate the charge transfer and conductivity of the composites, the EIS analysis was performed. Fig. 11c shows the Nyquist plots of the composites samples with their fitted circuit. In general, the smaller semicircle in the high-frequency region indicates the smaller charge transfer resistance (R_{CT}). In this regard, the R_s represents the series resistance, mainly from the solution impedance, while the R_{CT} represents the charge transfer resistance within the material.²⁸ As shown in Table 3, the R_s values are relatively comparable among the measured samples. Therefore, the resistance is dominantly affected by the R_{CT} . It can be seen from Fig. 11c that all of the GZ composites generally have lower semicircles compared to the bare $g\text{-C}_3\text{N}_4$, suggesting the lower charge transfer resistance.^{52–54} It is worth



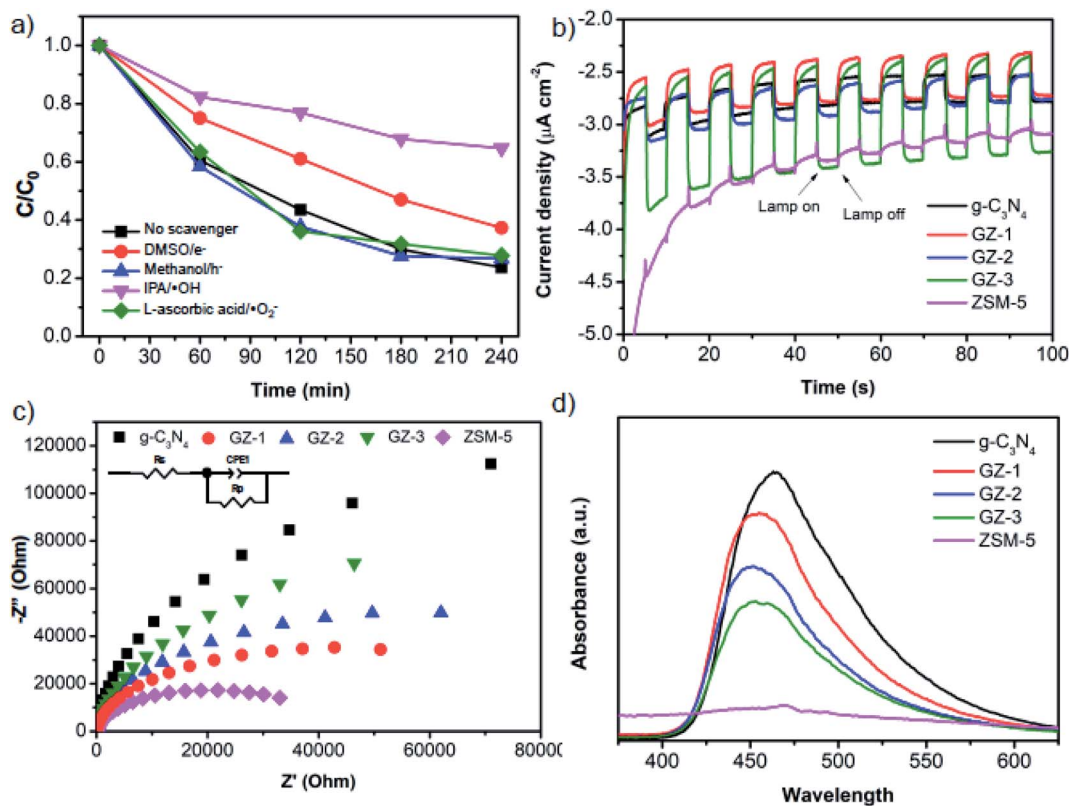


Fig. 11 (a) The reactive species trapping experiments for MB photodegradation over GZ-3 catalyst. Experiments were performed at MB initial concentration of 40 ppm. (b) Photocurrent spectra, (c) EIS Nyquist plots, and (d) PL spectra photocurrent spectra of bare g-C₃N₄, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts.

noting that the ZSM-5 exhibits the smallest charge transfer resistance, which indicates the highest conductivity among the samples. Based on this observation, the presence of ZSM-5 in the GZ-composites can potentially facilitate the charge carrier transfer from g-C₃N₄. Possibly, the more conductive ZSM-5 facilitates the charge transfer from the bulk phase of g-C₃N₄ to the composite's surface.

The PL spectra were then performed to investigate the separation and recombination of the photoinduced excitons. In general, a lower PL peak intensity reflects lower electron-hole pair recombination.⁵⁵ As shown in Fig. 11d, the PL emission peaks of bare g-C₃N₄ was observed at 473 nm and exhibits the highest emission intensity, suggesting the rapid recombination of the electron-hole pairs. Upon combination with ZSM-5, the intensity of the PL emission peaks decreases and exhibits a blue

shift, which agrees with the UV-vis DRS spectra.⁵⁶ The decrease of the intensity is the following order: g-C₃N₄ > GZ-1 > GZ-2 > GZ-3. The more ZSM-5 portion in the composites, the lower the intensity. The results suggest that the presence of ZSM-5 in the g-C₃N₄-based composites reduces the recombination of electron-hole pairs, which agrees with the previous report.²⁵ Considering a more inhibition of electron-hole pair recombination along with a more portion of ZSM-5 in the composites, it seems that more ZSM-5 in the GZ composites facilitates the electrons and hole transfer to the surface. Specifically, when the electrons and holes are generated in g-C₃N₄, the electrons and holes can be directly migrated to the ZSM-5, hence promoting the contact between the charge carriers (electrons and holes) with the reactant molecules. Note that ZSM-5 also has a porous structure, which provides more surface area. Because of this, the recombination can be suppressed by the presence of more ZSM-5 in the composites.

The results from PL analysis seem to have a different trend from the EIS results. However, it can be rationalized based on the charge transfer migration within the g-C₃N₄ and ZSM-5 particles. Based on EIS analysis, the more ZSM-5 portion in the composites, the higher charge transfer resistance. In this regard, the resistance in the composites may be mainly contributed by the interface resistance between the g-C₃N₄ and ZSM-5 particles. That is why the GZ-3 shows the highest charge transfer resistance since it has more contact area or interface

Table 3 R_s and R_{CT} values of over bare g-C₃N₄, ZSM-5, GZ-1, GZ-2, and GZ-3 catalysts

Samples	R_s (ohm)	R_{CT} (ohm)
g-C ₃ N ₄	43.26	274 250
GZ-1	46.82	73 906
GZ-2	49.05	116 170
GZ-3	43.27	169 560
ZSM-5	52.63	40 985



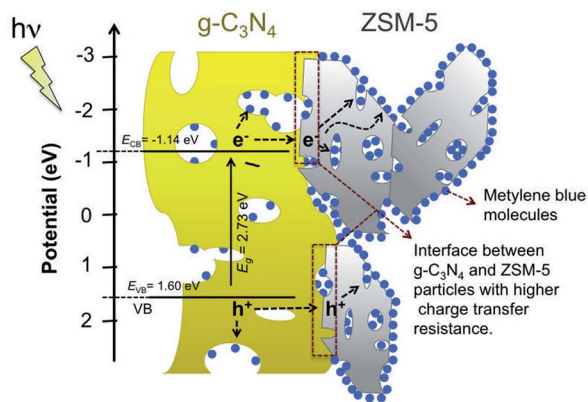


Fig. 12 Schematic illustration of charge carriers transfer in ZSM-5/g-C₃N₄ composite.

between the ZSM-5 and g-C₃N₄ particles. When the charge carriers migrate from the bulk g-C₃N₄, they can be highly resisted in the interface between g-C₃N₄ and ZSM-5 particles. However, once the electrons or holes reach the ZSM-5 particles, the charges carrier can be migrated efficiently to the surface and makes contact with reactants (Fig. 12). Note that the ZSM-5 shows the lowest charge transfer resistance (Fig. 11c). In other words, in the GZ composites, the g-C₃N₄ serves as the main photoactive material, which generates the electrons and holes, while the ZSM-5 facilitates the transfer of photogenerated charge carriers to the surface and optimizes the contact with the reactant by providing a larger surface area. Moreover, because of the facilitated charge transfer migration in the ZSM-5 particles, either volume recombination or surface recombination of electron and hole pairs can be significantly suppressed. Therefore, although GZ-3 has the highest charge transfer resistance than GZ-1 and GZ-2, it shows the lowest charge carriers recombination due to the significantly suppressed charge carrier pairs recombination.

Conclusions

In summary, the g-C₃N₄ was successfully composited with ZSM-5 to form ZSM-5/g-C₃N₄ composites with enhanced photocatalytic activity towards MB degradation. All samples exhibit well dispersion of g-C₃N₄ and ZSM-5 components with strong interaction as characterized by SEM, XRD, and FTIR analysis. The more portion of ZSM-5, as in GZ-3, leads to a stronger contact between the g-C₃N₄ and ZSM-5 particles. The UV-visible DRS analysis shows that bare g-C₃N₄ and GZ composites samples are responsive in the UV-visible region. The photocatalytic activity toward MB degradation confirms the photoactivity of bare g-C₃N₄ and GZ composites. More portion of ZSM-5 in the composites leads to better photocatalytic activity. The highest MB photodegradation activity was performed by GZ-3. The enhanced performance was attributed to the synergistic effect between g-C₃N₄ and ZSM-5, enhanced light scattering and the facilitated mass transfer. Unveiling the charge transfer behavior, the EIS and PL analysis reveal the enhanced charge

carrier transfer from g-C₃N₄ to the surface of the catalyst facilitated by ZSM-5, which suppresses the charge carriers recombination. Additionally, the higher resistance may raise from the interface between g-C₃N₄ and ZSM-5 particles, suggesting that the strong and highly intact interaction between the g-C₃N₄ and ZSM-5 particles is an important factor to be considered in order to further facilitate the charge transfer in ZSM-5/g-C₃N₄ composites.

Conflicts of interest

There are no conflicts to declare.

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