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Ni₃FeN functionalized carbon nanofibers boosting polysulfide conversion for Li–S chemistry†

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Limiting the shuttle effect of polysulfides is an important means to realizing high energy density lithium–sulfur batteries (Li–S). In this study, an efficient electrocatalyst (CNFs@Ni₃FeN) is synthesized by anchoring Ni₃FeN in the carbon nanofibers (CNFs). The CNFs@Ni₃FeN shows electrocatalytic activity and enhances the conversion of polysulfides. After assembling a battery, a high initial capacity (1452 mA h g⁻¹) and favorable long-time cycling stability (100 cycles) with a capacity retention rate of 83% are obtained by the electrocatalysis of Ni₃FeN. Compared with unmodified CNFs, the cycling stability of CNFs@Ni₃FeN can be greatly improved. The catalytic mechanism is further deduced by X-ray photoelectron spectroscopy (XPS). Our work will inspire the rational design of CNFs@support hybrids for various electrocatalysis applications.

Introduction

In the past decades, the demand for high energy density energy storage devices has promoted the development of many energy storage systems, such as lithium-ion battery technology, sodium-ion battery technology and so on. The most representative energy storage system is the lithium-ion battery.^{1–3} Lithium-ion batteries are also widely used in our society because of their relatively high energy density.⁴ However, lithium-ion batteries have some shortcomings, such as low abundance of lithium resources, uneven distribution and low energy density, so their further development is limited.⁵ Previous research shows that lithium–sulfur batteries (Li–S battery) have an energy density as high as 2600 W h kg⁻¹ and a theoretical specific capacity of 1675 mA h g⁻¹. So, Li–S battery is promising to become the next generation energy storage system after lithium-ion battery.⁶

As the positive active material, sulfur (S₈) participates in the charge compensation in the process of charge and discharge in Li–S battery. The discharge process of Li–S battery can be divided into three stages.^{6,7} In the first stage, sulfur (S₈) is reduced to form S₄²⁻ by step-by-step sequence reduction of soluble polysulfide ions, and the reaction speed is fast (voltage platform ~2.3 V vs. Li/Li⁺). It is worth noting that various intermediate polysulfides (Li₂S_x, x = 8–4) will be formed in this process. These polysulfides are highly dissolved in organic electrolytes, resulting in irreversible capacity loss. For example, high-order polysulfides are converted into low-order

polysulfides (Li₂S or Li₂S₂) and deposited the surface of lithium anode. This process will reduce the utility of active material sulfur and the coulombic efficiency in charge–discharge process. In the second stage, the soluble lithium polysulfide Li₂S₄ is reduced to insoluble Li₂S₂, and the voltage platform is 2 V vs. Li/Li⁺. This process is controlled by the solid phase nucleation reaction, so the reaction rate is slow. In the third stage, solid phase Li₂S₂ is reduced to solid phase Li₂S, which is controlled by solid phase diffusion of Li ions, so the reaction rate is slow.⁸ The capacity contribution of these three stages is 25%, 25% and 50%, respectively.⁹ Compared with the discharge process, the charge process is relatively simple. In the cyclic voltammetry (CV) of sulfur or polysulfide electrodes, typically there is only one anodic peak.¹⁰ Although the lithium–sulfur battery has high specific capacity and rich mineral resources, its development still faces many challenges, such as low conductivity of positive active material sulfur (S₈),¹¹ shuttle effect of polysulfides, slow reaction kinetics and large volume change in charge–discharge process.

At present, there are mainly the following strategies to overcome the problem of polysulfide shuttle effect: (1) physically limiting the active material sulfur to porous conductive materials;^{12–14} (2) chemical adsorption the polysulfides with defective carbon,^{15–17} transition metal oxides,^{18–20} MXene,^{21,22} transition metal nitride,²³ and so on; (3) microstructure design with graphene^{24–27} or carbon materials^{28–30} to prepare free-standing electrodes or three-dimensional electrodes;^{30–32} and (4) introducing a conductive intermediate layer between separator and cathode to alleviate the shuttle effect of the polysulfides.^{33–38} Among them, transition metal nitride has attracted considerable attention in Li–S batteries because the catalysts efficiently accelerate the conversion of the lithium polysulfides and actually inhibit polysulfide shuttling. Especially for the

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Ni₃FeN material with highly exposed active sites and high electrical conductivity has been developed to enhance the electrocatalytic activity in oxygen evolution reaction (OER).³⁹ Considering that sulfur and oxygen have similar electrochemical redox mechanisms, Ni₃FeN is also used as a sulfur promoter to promote the conversion of polysulfides.

In this work, we use CNFs@Ni₃FeN composites to alleviate the shuttle effect of polysulfides in the charge–discharge process, and enhance the coulomb efficiency and cycle stability of Li–S battery in the charge–discharge process. The Ni₃FeN in the material can implement the adsorption of soluble polysulfides produced in the process of charge and discharge and alleviate the shuttle effect in the process of cycle. At the same time, it can also catalyze the conversion of polysulfides to S₂²⁻ and the oxidation of Li₂S during charging, so as to reduce the existence time of soluble polysulfides and alleviate the shuttle effect of polysulfides.^{40–42} The carbon nanofibers in the materials provide a three-dimensional supporting conductive skeleton for the cathode materials. Therefore, through the synergistic effect of Ni₃FeN and CNFs, high-performance Li–S battery can be achieved.

Materials

All the chemical materials and characterized apparatus are described in the ESI.†

Results and discussions

The crystal structural characteristics of CNFs, Ni₃FeN and CNFs@Ni₃FeN were detected by X-ray diffraction (XRD) analysis, illustrated in Fig. 1a. From the diagram, we can observe that the peak position of the Ni₃FeN material coincides well with that of the Ni₃FeN standard card (PDF: 00-050-1434), indicating that the ternary phase Ni₃FeN has been synthesized successfully.³⁹ The characteristic peak of Ni₃FeN is strong and sharp, implying that the synthesized Ni₃FeN has high crystallinity. Comparing the XRD spectra of CNFs, Ni₃FeN and CNFs@Ni₃FeN, we find that CNFs@Ni₃FeN not only has the characteristic peak of Ni₃FeN crystal, but also has the same peak as CNFs, certifying that Ni₃FeN and CNFs were successfully compounded by electrospinning. Fig. 1b shows the Raman spectra of CNFs, Ni₃FeN and CNFs@Ni₃FeN. It can be observed that there are two typical G-band peaks in CNFs and CNFs@Ni₃FeN materials, which are located at 1354.17 cm⁻¹ and 1587.34 cm⁻¹ respectively. The peak-to-intensity ratio of D-band and G-band is 1.05, proving that the structure of CNFs is mainly short-range ordered graphite structure.⁴³ The low graphitization degree of the carbon in CNFs@Ni₃FeN can promote the migration of Li⁺ during the cycle.⁴² In addition, we can find that there is a Raman peak of Ni₃FeN (546.988 cm⁻¹) in CNFs@Ni₃FeN materials, which further proves that CNFs@Ni₃FeN is successfully synthesized. As depicted in Fig. 1c–e, the scanning electron microscopy (SEM) and transmission electron microscopy (TEM) images show that Ni₃FeN nanoparticles are well dispersed on the carbon nanofiber. From the SEM images of CNFs and CNFs@Ni₃FeN, it can be observed that the average diameter of CNFs and CNFs@Ni₃FeN is about

300 nm. According to TEM figure, Ni₃FeN nanoparticles are crystallized on the inner of carbon nanofiber. The *d*-spacings of 0.215 nm correspond to the (111) planes of Ni₃FeN.³⁹ This further verifies that Ni₃FeN is well compounded with carbon nanofibers.

For the purpose of exploring the load of sulfur in the host materials, the thermogravimetric analysis (TGA) of CNFs@Ni₃FeN@S was collected and shown in Fig. 1f. It can be seen that the mass loss of CNFs@Ni₃FeN@S sample is 77.03 wt%, but there is no obvious mass change of CNFs@Ni₃FeN sample. The result reveals that the load of sulfur is about 77.03 wt%. In order to further analyze the total surface area and micropore volume of the CNFs and CNFs@Ni₃FeN, the nitrogen adsorption–desorption isotherms were carried out. The type of curve in Fig. S1a† is type IV, which manifests that there are abundant mesoporous structures. From the analysis, the total surface area and micropore volume of the CNFs are 228.521 m² g⁻¹, 0.1017 cm³ g⁻¹, respectively and the CNFs@Ni₃FeN are 107.265 m² g⁻¹, 0.04857 cm³ g⁻¹, respectively. As exhibited in the Fig. S1b,† the pore size distribution of CNFs@Ni₃FeN is between 2 and 5 nm, while that of CNFs is about 15 nm. It turns out that both materials are porous structure. The porous structure is conducive to electrolyte infiltration into the material and the transport of lithium ion in the electrode material.⁴⁴ Moreover, the relatively small specific surface area and pore size in CNFs@Ni₃FeN can inhibit the bulky volumetric corrosion of the CNFs@Ni₃FeN@S cathode occurring in the conversion reaction that might enhance the energy storage performances of CNFs@Ni₃FeN.⁴⁵

The valence states of elements in CNFs@Ni₃FeN and CNFs were determined by X-ray photoelectron spectroscopy (XPS), shown in Fig. 2a–d. From Fig. 2a, there are obvious peaks of Fe, Ni, N and S in Ni₃FeN and CNFs@Ni₃FeN/Li₂S₄. The C and O elements with strong peaks in the picture may come from the raw materials in the process of Ni₃FeN preparation. Fig. 2b and c are the XPS spectrum of Ni 2p and Fe 2p, respectively. The result reveals that the element of Ni is in the state of Ni²⁺ and Ni⁰, and the element of Fe is in the state of Fe³⁺ and Fe⁰. The existence of Ni²⁺ and Fe³⁺ may be related to the existence of some inevitable surface oxide layer in Ni₃FeN.^{46–48} Compared with CNFs@Ni₃FeN, the Ni 2p_{3/2} spectrum of CNFs@Ni₃FeN/Li₂S₄ shifts to the direction of low binding energy (0.3 eV). These changes indicate that there is a charge transfer between Ni and S in Li₂S₄, and the electrons on S atoms are transferred to Ni atoms.⁴⁹ In the Fe 2p_{3/2} spectrum of Fig. 2c, we can see that all the peaks in the Fe 2p_{3/2} spectrum (712.6 eV) of CNFs@Ni₃FeN after polysulfide adsorption shift 0.4 eV to the direction of low binding energy, revealing that electrons are transferred from S atoms in Li₂S₄ to Fe atoms.⁴⁹ Fig. 2d is the N 1s spectrum. After the reaction with polysulfides, the N 1s spectrum of CNFs@Ni₃FeN/Li₂S₄ shifts significantly to higher binding energy than that of CNFs@Ni₃FeN (0.6 eV). This phenomenon may result from the transfer of electrons from N atom to metal atoms. The XPS spectra of Ni 2p_{3/2}, Fe 2p_{3/2} and N 1s show that Ni₃FeN has a strong adsorption effect on polysulfides.

In addition, static visible light adsorption experiments were implemented to visualize the anchoring ability between Ni₃FeN and polysulfides. CNFs and CNFs@Ni₃FeN with the same specific surface area were immersed into the Li₂S₄



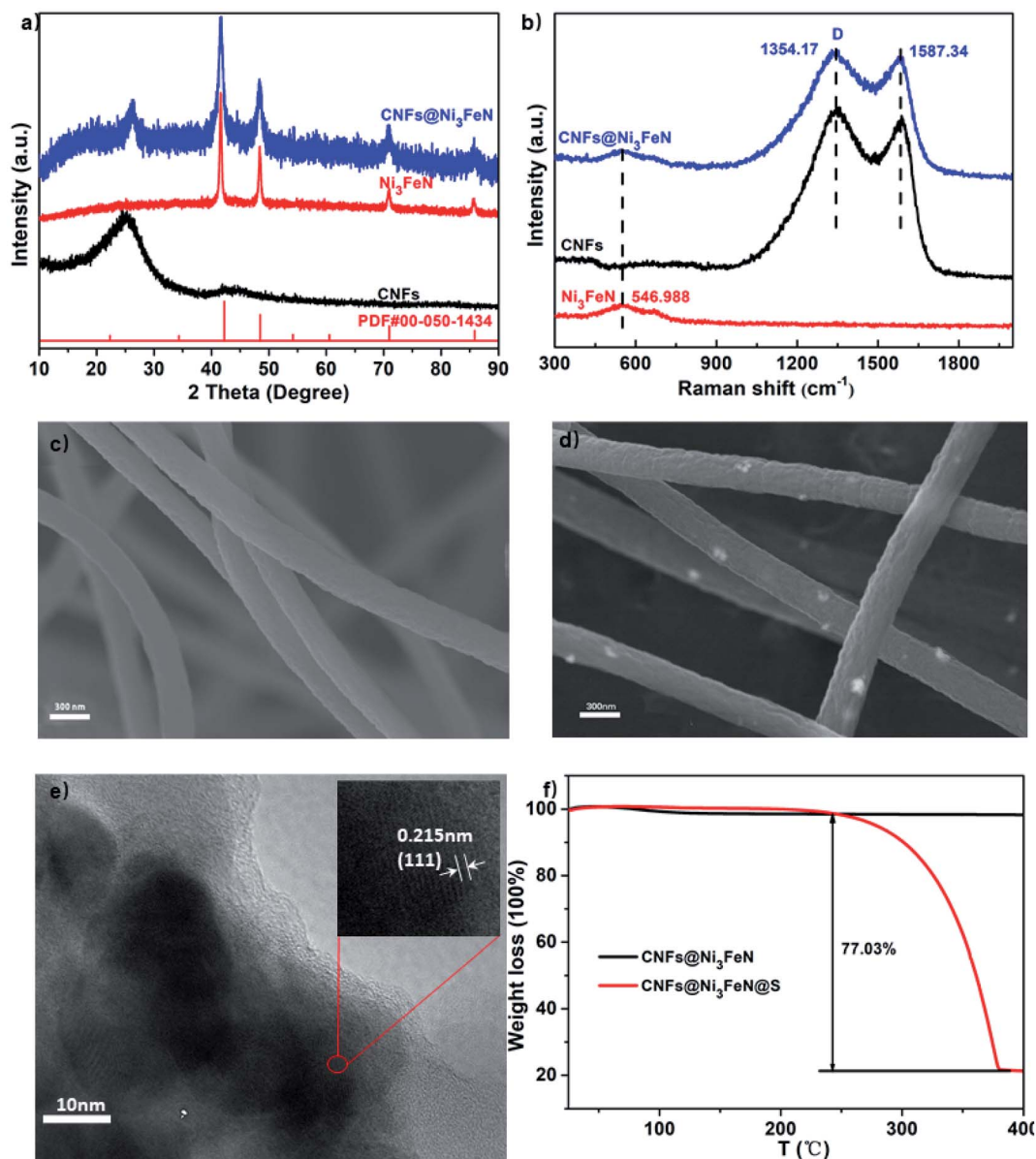


Fig. 1 (a) XRD patterns of Ni_3FeN , CNFs and $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$. (b) Raman spectra of Ni_3FeN , CNFs and $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$. SEM images of (c) CNFs and (d) $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$. (e) TEM images of $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$. (f) TGA curves of $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$ and $\text{CNFs}@_{\text{Ni}_3\text{FeN}}@_{\text{S}}$.

electrolyte of 1.5 mL. After that, they were placed in a glove box for 3 h. We can observe in Fig. 2e that the color of Li_2S_4 solution immersed with CNFs is still yellow and does not fade, while the color of Li_2S_4 solution immersed with $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$ has completely faded. The results show that Ni_3FeN has a strong adsorption effect on polysulfides. For further studying the interaction between $\text{CNFs}@_{\text{Ni}_3\text{FeN}}@_{\text{S}}$ and Li_2S_4 , XPS studies on Li_2S_4 and $\text{CNFs}@_{\text{Ni}_3\text{FeN}}@_{\text{S}}-\text{Li}_2\text{S}_4$ were carried out. As shown in Fig. 2f, the peaks at 161.8 eV and 163.6 eV in the S 2p spectrum correspond to terminal (S_T^{-1}) and bridged (S_B^0) sulfur atoms. However, when $\text{CNFs}@_{\text{Ni}_3\text{FeN}}@_{\text{S}}$ was added to the Li_2S_4 solution in Fig. 2g, S_T^{-1} (162.8 eV) and S_B^0 (163.9 eV) in the S 2p spectrum shifted to varying degrees, and they moved 1.0 and 0.3 eV, respectively. At the same time, there are

new peaks at 162.5/161.9 eV and 161.1 eV, which belongs to Ni–S bond and Fe–S bond, respectively. The appearance of these two bonds indicate that $\text{CNFs}@_{\text{Ni}_3\text{FeN}}@_{\text{S}}$ has a strong chemisorption to Li_2S_4 .^{50,51} In addition, we can see that there is a sulfate peak at 165–168 eV and a thiosulfate peak at 168 eV, which may be related to the disproportionation of intermediate Li_2S_x to produce higher and lower oxidized substances.^{52–55} Moreover, we can find that the peaks of these two places are weakened to some extent after the addition of $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$, which disclose that $\text{CNFs}@_{\text{Ni}_3\text{FeN}}$ can better catalyze the transformation of Li_2S_4 to low-valent reduced state. The above analyses imply that the electrocatalyst was successfully prepared and shown good adsorption and catalytic capacity for polysulfide.



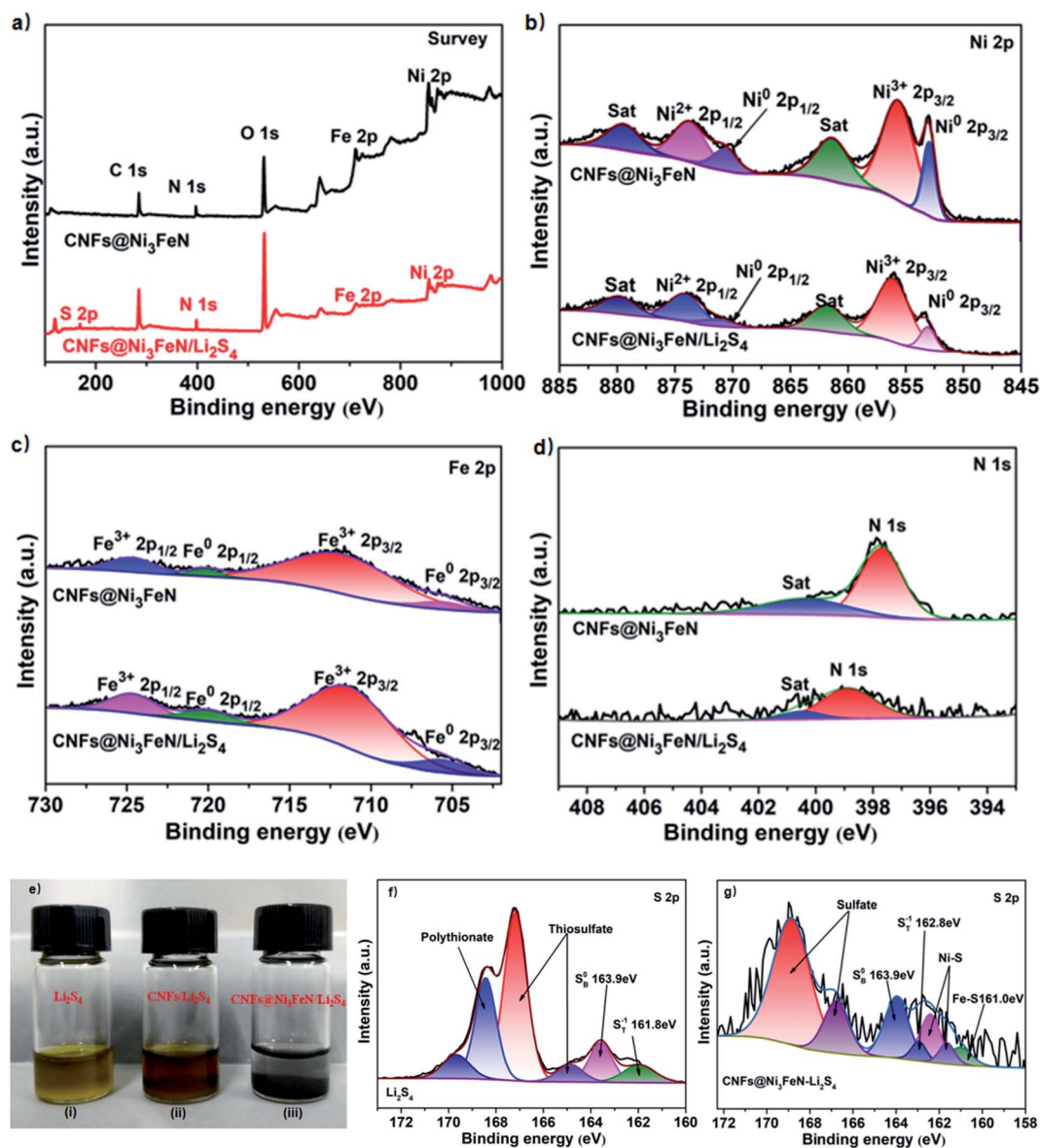


Fig. 2 (a–d) XPS spectra of the CNFs and CNFs@Ni₃FeN. (a) Full spectrum, (b) Ni 2p, (c) Fe 2p, (d) N 1s. (e) Photographs of a Li₂S₄ solution (i) Li₂S₄ solution added with CNFs (ii) and Li₂S₄ solution added with CNFs@Ni₃FeN (iii). (f) XPS S 2p spectra of Li₂S₄ and (g) XPS S 2p spectra of CNFs@Ni₃FeN/Li₂S₄.

The electrochemical storage properties of CNFs@S and CNFs@Ni₃FeN@S electrodes were studied by using coin cell and lithium sheet as counter electrode. The CV curves of CNFs@S and CNFs@Ni₃FeN@S are shown in Fig. 3a. The typical cathode peaks appear at 2.0 V and 2.3 V, which is attributed to the formation and conversion of long-chain polysulfides to Li₂S. The continuous anodizing peaks at 2.3–2.5 V are the transition from Li₂S/Li₂S₂ to polysulfide. CNFs@Ni₃FeN@S shows the sharpest and highest anodic peak, which implies that the conversion process from Li₂S to polysulfides is more thorough. At the same time, the CNFs@Ni₃FeN also shows better catalytic behavior to promote the transformation of polysulfides. Fig. 3b presents the CV curves of the five circles of CNFs@Ni₃FeN@S. It can be seen that the coincidence of

cathode peak and anode peak is very good. The result demonstrates that the Li–S battery assembled with CNFs@Ni₃FeN@S as positive electrode has excellent charge–discharge cycle stability and the reversible reaction is more favorable from S₈ to Li₂S/Li₂S₂.⁵⁶

To further quantitatively analyze the catalytic kinetic rate of polysulfide conversion process, the Tafel slopes of the first reduction and oxidation processes were obtained in Fig. 3c and d, respectively. For the first reduction peak, the Tafel slopes of CNFs@S and CNFs@Ni₃FeN@S are 41.0 and 57.6 mV dec⁻¹, respectively. For the first oxidation peak, the Tafel slopes of CNFs@S and CNFs@Ni₃FeN@S are 76.0 and 62.9 mV dec⁻¹, respectively. The Tafel slope of lithium–sulfur battery assembled with *in situ* synthesized CNFs@Ni₃FeN@S as positive



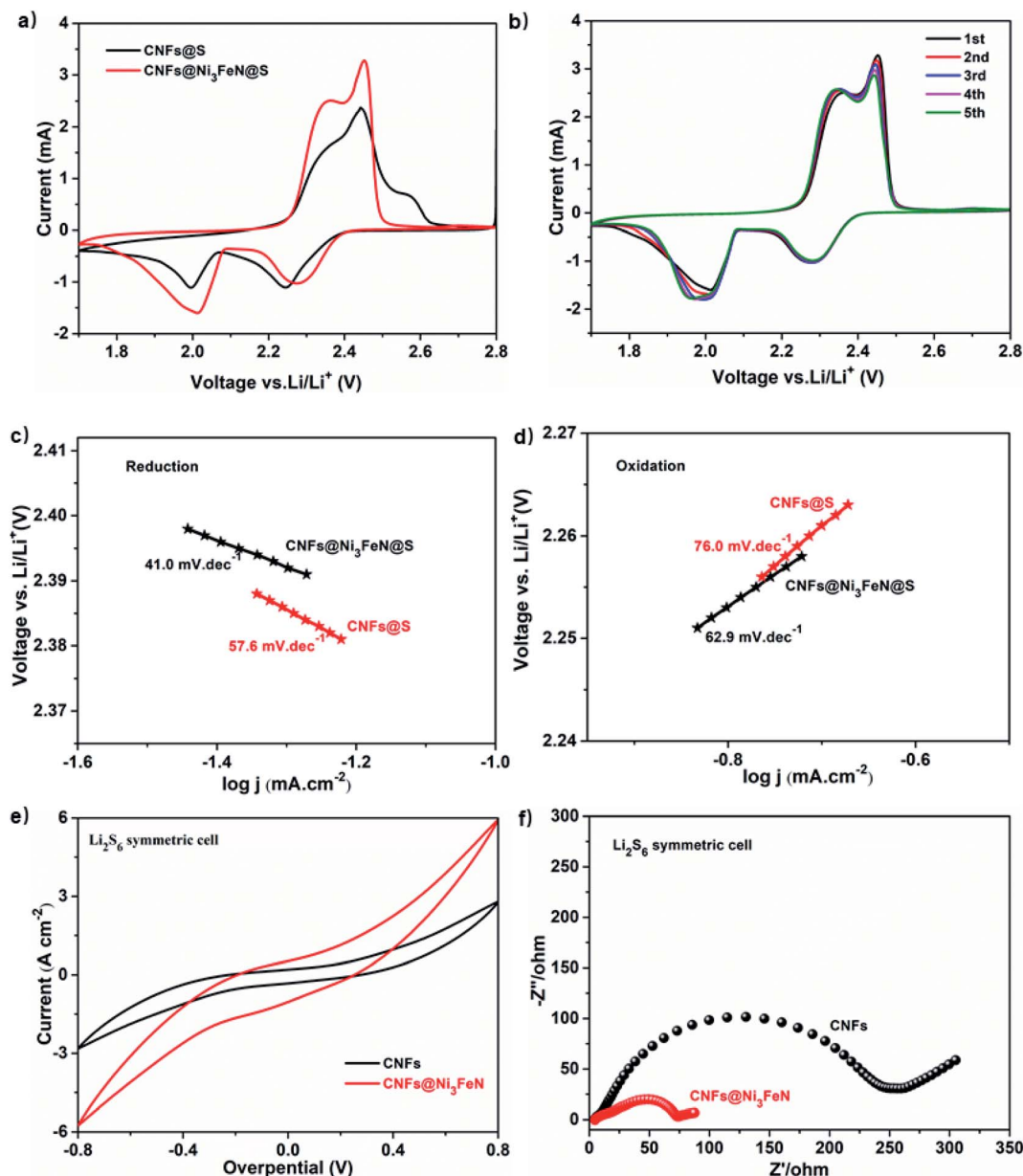


Fig. 3 CV curves at the scan rate of 0.1 mV s^{-1} in a potential window from 1.7 to 2.8 V of (a) the CNFs@Ni₃FeN and CNFs@Ni₃FeN@S electrodes and (b) the first five cycles of CNFs@Ni₃FeN@S. Tafel plots of the (c) first reduction peak and (d) the oxidation peak. (e) CV curves of symmetric cells and (f) EIS Nyquist plots obtained from CNFs@Ni₃FeN and CNFs@Ni₃FeN@S electrodes.

electrode is relatively small, which clearly confirms that the reaction between liquid polysulfide and solid sulfur is easier, and the electrocatalyst Ni₃FeN promotes the conversion of polysulfide.⁵⁷ Symmetrical cells were assembled for CV and Electrochemical impedance spectroscopy (EIS) tests to further explore the catalytic performance of CNFs and CNFs@Ni₃FeN. As showcase in Fig. 3e, when the voltage bias is 0.8 V, the redox current response of CNFs@Ni₃FeN is nearly twice that of CNFs, implying that Ni₃FeN can enhance the conversion efficiency and electrochemical reversibility of polysulfides to Li₂S.⁵⁸ The EIS measurements of CNFs and CNFs@Ni₃FeN are analyzed in Fig. 3f. It can be seen from the Nyquist diagram that the EIS of CNFs and CNFs@Ni₃FeN are composed of high frequency

semicircle, intermediate frequency semicircle and an inclined low frequency line. The high frequency semicircle is related to the charge transfer resistance (R_{ct}), the intermediate frequency semicircle is related to the establishment of the solid electrolyte interface (SEI) film resistance (R_s) produced during the irreversible aggregation of insoluble low-order polysulfide, and the inclined line is related to the Warburg impedance (W_e) of Li⁺ diffusion in the electrode.⁵⁹ Obviously, compared with the CNFs@S electrode, the CNFs@Ni₃FeN@S displays smaller the solid electrolyte interface (SEI) film resistance (R_s). The excellent interface impedance results suggest that due to the existence of Ni₃FeN, the amount of irreversible insoluble low-order polysulfide at the electrode interface is reduced, thereby resulting in



enhanced the transformation of insoluble low-order polysulfide to higher-order polysulfide and S_8 . Above results consistently demonstrate the efficient adsorption and catalytic polysulfide conversion capability of $CNFs@Ni_3FeN$.

The galvanostatic charge–discharge curves of $CNFs@S$ and $CNFs@Ni_3FeN@S$ composites at 0.1C ($1C = 1675 \text{ mA h g}^{-1}$) are witnessed in Fig. 4a. At $\sim 2.3 \text{ V}$ and $\sim 2.1 \text{ V}$, the two typical high discharge platforms and low discharge platforms refer to the adsorption level and conversion level of polysulfides, respectively.⁴² This is consistent with the analysis of the CV curve.

$CNFs@Ni_3FeN@S$ composite nanostructure materials can provide high initial discharge capacity of $1406.5 \text{ mA h g}^{-1}$, which is much higher than that of $CNFs@S$ ($1156.9 \text{ mA h g}^{-1}$). The charge–discharge platform of $CNFs@Ni_3FeN@S$ is longer, indicating that it has the ability to improve the utilization of active materials and promote the conversion between solid sulfide and long-chain polysulfides. In addition, to further confirm that $CNFs@Ni_3FeN@S$ can better adsorb and catalyze polysulfides, we used $CNFs@Ni_3FeN@S$ and $CNFs@S$ as positive electrodes to test the charge–discharge cycle performance of

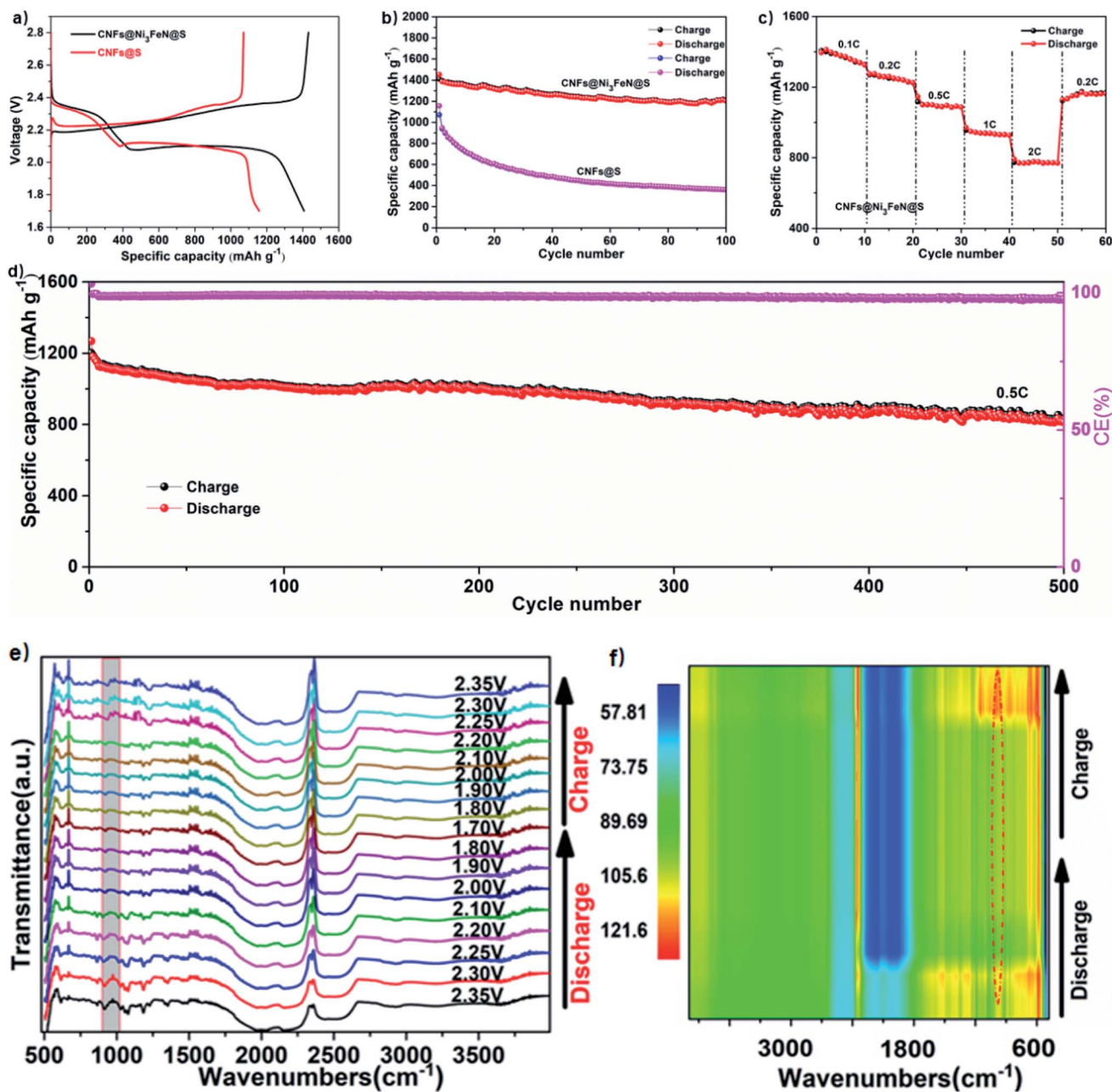


Fig. 4 (a) Reversible discharge/charge voltage profiles at 0.1C of $CNFs@S$ and $CNFs@Ni_3FeN@S$ electrodes. (b) Cycle performances at 0.1C of $CNFs@S$ and $CNFs@Ni_3FeN@S$ electrodes. (c) Rate performances of $CNFs@Ni_3FeN@S$ electrodes at 0.1C, 0.2C, 0.5C, 1C and 2C rates. (d) Long-term cycling performances of $CNFs@Ni_3FeN@S$ electrodes at 0.5C. (e) *In situ* FT-IR spectra of $CNFs@Ni_3FeN@S$ cathode at different discharge and charge states. (f) Contour and response surface analysis corresponding to the discharge and charge processes.



the battery at 0.1C. As can be seen from Fig. 4b, the initial discharge specific capacities of CNFs@Ni₃FeN@S and CNFs@S are 1452.2 mA h g⁻¹ and 1156.9 mA h g⁻¹, respectively. After 100 cycles, the discharge specific capacities of CNFs@Ni₃FeN@S and CNFs@S are 1203.7 mA h g⁻¹ and 358.2 mA h g⁻¹, respectively. The results imply that the Ni₃FeN can effectively adsorb polysulfides and catalyze the conversion of polysulfides.

The rate performance of CNFs@Ni₃FeN@S cathode at various current rates from 0.1C to 2C was evaluated in the voltage range from 1.7 to 2.8 V. As depicted in Fig. 4c, the discharge capacities of CNFs@Ni₃FeN@S cathode exhibit 1372.4, 1261.9, 1097.1, 937.1, and 781.2 mA h g⁻¹ at 0.1, 0.2, 0.5, 1, and 2C, respectively. Impressively, once the current rate switched back to 0.2C, the CNFs@Ni₃FeN@S cathode can harvest an excellent capacity of 1129.5 mA h g⁻¹, which is attributed to the robust and stable structure.⁴⁰ Long-cycle stability is very important for commercial applications of lithium batteries. For evaluate the long-term cycle behavior of CNFs@Ni₃FeN@S, batteries using CNFs@Ni₃FeN@S cathodes are cycled at 0.5C for 500 times in Fig. 4d. After 500 cycles, the discharge capacity descends from 1267 mA h g⁻¹ to 812 mA h g⁻¹, and the coulombic efficiency remains above 99%. The result displays that the CNFs@Ni₃FeN@S cathode has excellent long cycle performance, which can be attributed to the catalysis and adsorption of Ni₃FeN.

Furthermore, in order to fully understand the mechanism of action, we assembled the Li-S half cells for *in situ* Fourier transform infrared (FT-IR) measurement. The cathode material was CNFs@Ni₃FeN@S, the anode was lithium plate, the positive current collector was carbon paper, and the separator was Celgard polypropylene (PP) membrane. The electrolyte was 1 M LiTFSI dissolved in DME/DOL ($V_{\text{DME}} : V_{\text{DOL}} = 1 : 1$) solution, in addition, 2% LiNO₃ was used as additive. The result is illustrated in Fig. 4e and f. It is worth noticing that the intensities of the Li-S peak located at 933 cm⁻¹ decrease and nearly disappear at the end of discharging, followed by reappearance upon charging, signifying the excellent reversibility of electrochemical reactions and effective restriction of polysulfides within the CNFs@Ni₃FeN@S||Li@Li cell.

Conclusions

In summary, we proved that Ni₃FeN can be used as a highly efficient cathode catalyst for Li-S cells. In this paper, CNFs@Ni₃FeN materials with special micro-nano composite structure were synthesized *in situ* by electrospinning. Electrochemical tests and other physical characterizations show that, compared with CNFs, CNFs@Ni₃FeN not only has strong chemisorption effect on polysulfide intermediates, but also can significantly accelerate the transformation of polysulfides and alleviate the shuttle effect of Li-S cells. Our work explores a new way for the design and development of cathode materials for high-performance Li-S batteries, as well as providing a more reasonable method for the commercialization of Li-S cells.

Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

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