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Molecular evidence for sulfurization of molybdenum dithiocarbamates (MoDTC) by zinc dithiophosphates: a key process in their synergetic interactions and the enhanced preservation of MoDTC in formulated lubricants?†

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Molybdenum dithiocarbamates (MoDTC) are widely used in automotive industries as lubricant additives to reduce friction and to enhance fuel economy. Sulfur-containing additives such as zinc dithiophosphates (ZnDTP) are proposed to play a key role in the improvement of friction reducing properties of MoDTC in formulated lubricants by facilitating the formation of MoS₂ tribofilm at the rubbing contacts. This study focuses on the interactions between MoDTC and ZnDTP under conditions comparable with those prevailing in operating engines. The capacity of ZnDTP to sulfurize MoDTC in solution in a hydrocarbon base oil could be demonstrated. Sulfurized Mo complexes bearing one or two additional sulfur atoms (1S-MoDTC and 2S-MoDTC, respectively) which have replaced the genuine oxygen atom(s) from the MoDTC core were detected and quantified using a specifically developed HPLC-MS analytical method. A possible sulfurization mechanism relying on the higher affinity of phosphorus from ZnDTP for oxygen could be proposed. In parallel, the evolution and molecular transformation of the prepared 2S-MoDTC in hydrocarbon base oil under thermal and thermo-oxidative conditions were followed using HPLC-MS and compared with the evolution of their friction coefficients. 2S-MoDTC complexes were shown to exhibit a better retention of friction reducing capability under oxidative conditions than the “classical” MoDTC, although they did not seem to significantly reduce the friction coefficients of lubricants as compared to the “classical” MoDTC. Therefore, sulfurization of MoDTC by ZnDTP might contribute to delaying the progressive consumption of MoDTC and the loss of their friction-reducing efficiency in lubricants under thermo-oxidative conditions.

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1. Introduction

Molybdenum-based compounds have a wide range of applications in chemistry, and are used for instance as catalysts to speed-up chemical reactions^{1–3} or occur as lubricant additives in automotive applications.^{4,5} Among the different molybdenum-based additives, oil-soluble molybdenum dithiocarbamates (MoDTC) have been used as highly effective friction modifiers since the early 1970s^{6–8} in engine lubricants for energy saving. Indeed, MoDTC are able to generate molybdenum disulfide (MoS₂) at the tribo-contact, thereby providing ultralow friction coefficients which improve fuel efficiency.^{9–15} Different analytical techniques such as X-ray spectroscopy,^{16,17} electron

diffraction,¹⁸ Raman spectroscopy^{19,20} and high-resolution transmission electron microscopy⁸ have been employed to detect and to study MoS₂ tribofilms. Although numerous investigations have been conducted to explain the mechanisms of MoS₂ formation from Mo-based additives,^{9,13,21–25} the molecular transformations involved in the processes have not yet been completely clarified. The main limitation of employing MoDTC as friction modifiers is related to their deterioration over time upon oil oxidation, leading to the loss of their friction-reducing effectiveness during engine operation.^{11,26–30} Therefore, understanding the mechanisms of MoDTC decomposition and their interactions with other additives from formulated lubricants at the molecular level is crucial in order to (i) predict their friction reduction efficiency (ii) provide guidance for the optimization of MoDTC tribological performance in lubricant formulation (iii) extend the functional lifespan of MoDTC during engine oil ageing and (iv) provide information for future development of new eco-friendly additives.

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Modern lubricant formulations also require anti-wear action and thus additional multifunctional additives are used to achieve optimal engine efficiency. The presence of these additives in formulated lubricants could exhibit synergistic or antagonistic interactions with MoDTC, resulting in either enhancement or decrease in friction reduction efficiency of MoDTC. Zinc dithiophosphates (ZnDTP) show antiwear, anticorrosive and antioxidant properties,³¹ thus representing one of the most widely used multifunctional additives in engine lubricants. Tribo-chemical interactions between ZnDTP and MoDTC have been shown to exhibit positive effects on both friction and wear reduction in boundary lubrication regime.^{32–35}

In this connection, it has been proposed that the presence of both molybdenum and sulfur-containing species such as ZnDTP in lubricants is required for the generation of a MoS₂ tribofilm.³⁶ Similarly, Sogawa *et al.* (1999)³⁷ showed that *ca.* 40% of the sulfur atoms present in the MoS₂ tribofilm derive from ZnDTP when used in combination with MoDTC, suggesting that the presence of a sulfur source together with zinc atoms are important to promote the generation of MoS₂ at the tribo-contacts. Umehara *et al.* (2013)³⁸ also asserted that ZnDTP could provide sulfur atoms to MoDTC and affirmed the importance of sulfur suppliers to form MoS₂ layers at the sliding surfaces. In addition, some sulfur-free organic molybdenum-based friction modifiers were also proved to be effective when used in conjunction with sulfur-containing additives such as ZnDTP.^{39,40}

Therefore, given the numerous studies emphasizing on the importance of sulfur suppliers such as ZnDTP to improve the tribological performance of MoDTC in lubricants, this study focuses on the investigation at the molecular level of the interactions between MoDTC and ZnDTP using laboratory oil ageing experiments. We report here the transformation of MoDTC into mono- and di-sulfurized analogues (hereinafter noted as 1S- and 2S-MoDTC) induced by ZnDTP in lubricants under thermal non-oxidizing conditions (Fig. 1a). The formation of the sulfurized Mo complexes was investigated using a specifically developed HPLC-MS analytical method allowing their unambiguous detection. Besides, 2S-MoDTC reference compounds were synthesized using Lawesson's reagent (Fig. 1b) for the investigation of their tribological performances,

evolution and chemical transformation in hydrocarbon base oils under thermal and thermo-oxidative conditions. The possible implications of the sulfurization processes of MoDTC on the tribological performances of lubricants and on the persistence of MoDTC in lubricants under thermo-oxidative conditions have been investigated using the synthesized 2S-MoDTC.

2. Experimental section

2.1 Materials (base oil and additives)

The mineral base oil used in this work was provided by Total-Energies Marketing Services. The chemical structures of the Mo and Zn complexes discussed in this study are shown in Fig. 2. MoDTC **1a–1c** were purchased from Adeka as a mixture in base oil (MoDTC **1a–1c**: 10 wt% Mo; trade name: Sakuralube 525), whereas MoDTC **1d** (trade name Sakuralube 600) occurs as a pure solid material. ZnDTP **2a–2c** (in solution in a hydrocarbon base oil) were purchased from Infineum. ZnDTP **2a** (7.3 wt% P, trade name: OLOA 269R) corresponds to a primary dithiophosphate complex with C₈ (2-ethylhexyl) alkyl chains. Secondary ZnDTP (trade name: Infineum C9421) additives were bought as a mixture of compounds **2b** (8.0 wt% P) with C₃ and C₆ alkyl chains and ZnDTP **2c** (7.0 wt% P, trade name: Infineum C9426) is constituted of a mixture of ZnDTP with a complex mixture of isomeric C₈ (isooctyl) alkyl chains (see ESI Fig. 7S†). Lawesson's reagent **3** was purchased from Sigma-Aldrich. All the purchased compounds were used without any prior purification.

2.2 Laboratory ageing experiments

2.2.1 Heating experiments involving concentrated mixtures. Mixtures of commercial additives containing MoDTC **1a–1c** and ZnDTP (**2a**, mixture of **2b** or **2c**, respectively) (1 : 1, wt./wt.) were vacuum-sealed in NMR glass tubes without dilution in hydrocarbon base oil or solvents. The glass tubes were heated at 150 °C for 3 h or 6 h.

2.2.2 Heating experiments involving diluted solutions. Ageing experiments were carried out using hydrocarbon base oils containing the additives MoDTC **1a–1c** (1 wt%) or the

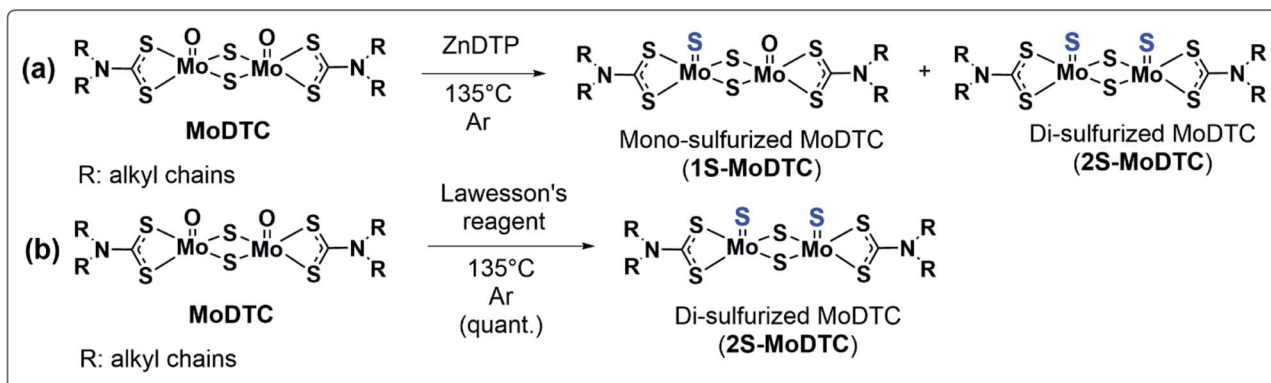


Fig. 1 Sulfurization of MoDTC by (a) ZnDTP and (b) Lawesson's reagent.



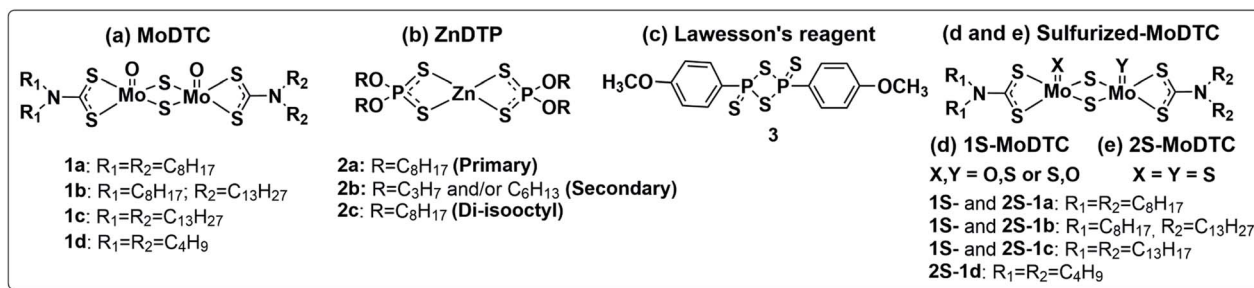


Fig. 2 Chemical structures of lubricant additives cited in this work. (a) Molybdenum dithiocarbamates (MoDTC); (b) zinc dithiophosphates (ZnDTP); (c) Lawesson's reagent; (d) mono-sulfurized molybdenum dithiocarbamates (1S-MoDTC); (e) di-sulfurized molybdenum dithiocarbamates (2S-MoDTC).

synthesized sulfurized MoDTC **2S-1a-1c** (1 wt%, Fig. 2) alone or in combination with ZnDTP (**2a**, **2b** or **2c**; 1 wt%) at concentrations compatible with those from commercial lubricants. Typically, 50 g of formulated oils or hydrocarbon base oils with additives were heated at 135 °C in a multi-necked round-bottomed flask under a constant flow (100 mL min⁻¹) of either argon, air or NO₂ (2000 ppm in air) bubbling into the reaction mixture. In the case of the experiments with NO₂, the oil mixture was first heated under a bubbling flow of argon and the gas flow was then switched to NO₂ in air as soon as the temperature of 135 °C was reached. Oil samples (1 mL aliquots) were collected through a septum at 2–3 h intervals using a syringe for analysis by HPLC-MS and for tribological tests. The first sample was collected when the temperature of the oil-additive mixture reached 135 °C (defined as $T = 0$ h).

2.3 Ball-on-flat tribotests

Tribological tests were carried out by using an Anton Paar MCR302 rheometer with a T-PTD-200 tribology assembly (ball on three flats). Test temperature was 110 °C, normal load 5 N and sliding speed was kept constant throughout 15 min test at 0.1 m s⁻¹. Upper ball was made of AISI 52100 hardened steel (roughness) and flats was made of 100Cr6 hardened steel (60–66 HRC). Roughness of the balls was 0.032 μm (R_a) and 0.06 μm (R_a) for the flats. Test conditions are representative of a “boundary lubrication regime” where important asperity to asperity contacts occur favouring the reaction of additives and tribofilm formation. Only final friction coefficients are reported in this paper.

2.4 High pressure liquid chromatography-mass spectrometry (HPLC-MS) analyses

MoDTC and their sulfurized analogues were analyzed by HPLC-MS using an Agilent HP 1100 series instrument equipped with an auto-injector, and connected to a Bruker Esquire 3000^{Plus} ion trap mass spectrometer as described in Kiw *et al.* (2020).⁴¹ Sample preparation and quantification of MoDTC derivatives were carried out as described in Kiw *et al.* (2020)⁴¹ using 2,3-bis(*n*-octadecyloxy)-propan-1-ol as internal standard which was added prior HPLC-MS analysis. ZnDTP **2a-2c** could not be detected using the developed HPLC-MS analytical method. In

the case of the experiments involving concentrated solutions, samples were prepared following a similar procedure, but MoDTC derivatives were not quantified (no internal standard added).

2.5 ¹H and ¹³C NMR (analysis of 2S-1d)

NMR spectra were recorded on a Bruker Avance I 500 MHz spectrometer (500 MHz for ¹H; 125 MHz for ¹³C) for NMR measurements. The ¹H- and ¹³C-chemical shifts are reported in ppm relative to tetramethylsilane with the residual protons and carbon atoms of the solvent used as internal standards (CDCl₃; δ ¹H 7.26 ppm; δ ¹³C 77.2 ppm). 5 mg of the synthesized **2S-1d** was diluted in *ca.* 0.8 mL of CDCl₃ for NMR analysis.

2.6 GC-MS (analysis of ligands from ZnDTP 2c)

GC-MS analysis was carried out using a Thermo Trace gas chromatograph (Thermo Scientific) coupled to a Thermo Scientific TSQ Quantum mass spectrometer equipped with a programmed temperature vaporizing (PTV) injector. The temperature of the source was set at 220 °C. The mass spectrometer was operating in the electron ionization (EI) mode at 70 eV and scanning m/z 50 to 700. Gas chromatographic separation was performed on a HP5-MS column (30 m × 0.25 mm; 0.1 μm film thickness) using He as carrier gas at a constant flow rate of 1.2 mL min⁻¹. The oven temperature was programmed as followed: 40 °C (1 min), 40–200 °C (10 °C min⁻¹), 200–300 °C (4 °C min⁻¹), isothermal at 300 °C for 30 min.

2.7 Synthesis of 2S-MoDTC

MoDTC **1d** (5 g, 7.1 mmol) and Lawesson's reagent **3** (2.9 g, 7.1 mmol) were stirred under reflux in toluene for 1.5 h. The reaction mixture was then cooled down to room temperature and washed with 200 mL of water. The organic phase was dried over magnesium sulfate, filtrated and concentrated *in vacuo* to afford **2S-1d** as a dark red solid. ¹H- and ¹³C- NMR data as well as HPLC-MS analysis showed that the starting product **1d** was completely converted into its corresponding disulfurized MoDTC complex **2S-1d**, indicating that the sulfurization reaction was quantitative. It is worth noting that **2S-1d** is barely soluble in a hydrocarbon medium, and it was thus not further used for laboratory experiments. It was synthesized essentially



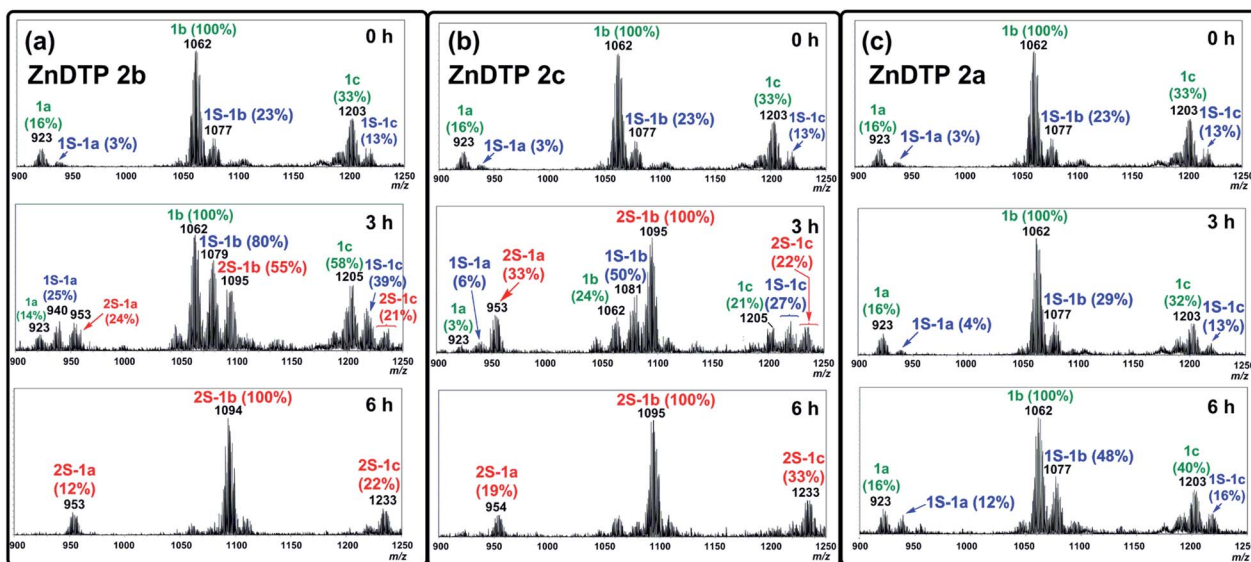


Fig. 3 Mass spectra (m/z 900–1250, APPI, positive mode) showing the evolution of the reaction mixture obtained by heating MoDTC 1a–1c with (a) secondary ZnDTP 2b; (b) di-isooctyl ZnDTP 2c; (c) primary ZnDTP 2a in a concentrated medium at 150 °C after 0 h, 3 h and 6 h. % corresponds to the normalized intensity.

to obtain a pure reference compound of 2S-MoDTC for ^1H - and ^{13}C -NMR characterization and allowed the yield of the sulfuration reaction to be evaluated.

Following this, the same procedure as above was carried out using a mixture containing MoDTC 1a–1c (8.2 g, 4.7 mmol) and Lawesson's reagent 3 (2.0 g, 4.7 mmol), leading to the complete conversion of 1a–1c into the corresponding mixture of 2S-1a–1c as a dark red liquid, as determined by HPLC-MS (APPI, positive mode). 2S-1a–1c existed in the form of a mixture of three 2S-MoDTC complexes, therefore their formation was hard to be confirmed by ^1H -NMR. However, the reaction was believed to give an almost quantitative yield of 2S-1a–1c, based on the ^1H - and ^{13}C -NMR results obtained from the pure disulfurized MoDTC complexes 2S-1d, which was synthesized using the same procedure (see above).

This sulfuration reaction involving either MoDTC 1a–1c or MoDTC 1d was repeated twice and led to the same results as those reported above (*i.e.*, complete sulfuration of the starting MoDTC products). Direct evaluation of the sulfuration yield using NMR would be difficult in the case of the reaction involving the mixture of MoDTC 1a–1c since the chemical shifts of the protons from the methylenes adjacent to N of the various MoDTC involved (1a–1c) and of their related sulfurized analogues 2S-1a–1c must have similar though not identical chemical shifts, making the overall complex signal hard to interpret.

2S-1d: ^1H -NMR (500 MHz; CDCl_3): 4.00–3.93 (m, 4H), 3.92–3.85 (m, 4H), 1.88–1.7 (m, 8H), 1.49–1.36 (m, 8H), 1.00 (t, $J = 7.5$ Hz, 12H). ^{13}C -NMR (125 MHz; CDCl_3): 207.4, 51.6, 29.7, 20.2, 13.9.

2S-1d: HPLC-MS (APPI, positive mode), $[\text{M} + \text{H}]^+$: m/z (relative intensity) 772 (11%), 723 (27), 724 (42), 725 (44), 726 (78), 727 (65), 728 (98), 729 (91), 730 (100), 731 (67), 732 (72), 733 (36), 734 (42), 735 (12), 736 (17) (see also ESI Fig. S1† for the mass spectrum of 2S-1d).

2S-1a: HPLC-MS (APPI, positive mode), $[\text{M} + \text{H}]^+$: m/z (relative intensity) 948 (13%), 949 (31), 950 (43), 951 (61), 952 (54), 953 (83), 954 (87), 955 (100), 956 (73), 957 (62), 958 (55), 959 (36), 960 (15).

2S-1b: HPLC-MS (APPI, positive mode), $[\text{M} + \text{H}]^+$: m/z (relative intensity) 1087 (17%), 1088 (24), 1089 (52), 1090 (48), 1091 (72), 1092 (67), 1093 (92), 1094 (100), 1095 (96), 1096 (82), 1097 (91), 1098 (57), 1099 (52), 1100 (25), 1101 (24).

2S-1c: HPLC-MS (APPI, positive mode), $[\text{M} + \text{H}]^+$: m/z (relative intensity) 1226 (10%), 1227 (17), 1228 (31), 1229 (34), 1230 (36), 1231 (74), 1232 (56), 1233 (83), 1234 (100), 1235 (86), 1236 (93), 1237 (75), 1238 (52), 1239 (45), 1240 (26), 1241 (19), 1242 (10) (see also ESI Fig. S2† for the mass spectra of 2S-1a–1c).

2.8 GC-MS ligand analysis of ZnDTP 2c

5 mg of ZnDTP 2c were dissolved in a 1 : 1 v/v mixture of methanol and dichloromethane (1 mL). Methyl iodide (500 μL) was then added and the mixture was left to react at room temperature for 1 h. After removal of methyl iodide and of the solvents under a stream of argon, the crude mixture was separated by column chromatography on silica gel (H: 4 cm, O : 1.5 cm). A first elution using cyclohexane (20 mL; 2 dead volumes) yielded a fraction containing the hydrocarbon base oil present with the additive. A second fraction was obtained by elution with dichloromethane (30 mL; 3 dead volumes) and contained the dithiophosphate ligands with isomeric C_8 (isooctyl) alkyl chains which were analyzed as methylated derivatives using GC-MS.

3. Results

3.1 Sulfuration of MoDTC by ZnDTP in a concentrated medium

The interactions between MoDTC and ZnDTP have first been investigated under concentrated conditions by heating



a mixture of commercial MoDTC **1a–1c** and primary ZnDTP **2a** (secondary ZnDTP **2b** or **2c**, respectively) in vacuum-sealed glass tubes at 150 °C for 3 h and 6 h. HPLC-MS analysis of the resulting crude mixtures showed the formation of two new series of compounds. Based on the interpretation of their mass spectra, it could be established that the newly-formed compounds corresponded to Mo derivatives, given the typical complex isotopic pattern observed for the pseudo-molecular ions (Fig. 3).⁴¹ Since these pseudo-molecular ions were shown to be shifted upwards by 16 mass units (first series) or 32 mass units (second series) relative to the MoDTC substrates, it could be deduced that the new compound series corresponded to MoDTC derivatives bearing one or two additional sulfur atoms (1S-MoDTC and 2S-MoDTC, respectively; Fig. 3) which have replaced the genuine oxygen atom(s) from the MoDTC core. It is worth mentioning that the formation of MoDTC derivatives showing an upward molecular shift of 16 and 32 mass units has been previously observed by electrospray mass spectrometry (direct inlet) by De Feo (2015)²⁶ in experiments involving MoDTC and ZnDTP in solution in a hydrocarbon base oil, these authors suggesting the possible sulfur–oxygen exchange on the genuine MoDTC, without providing further explanation on the sulfurization mechanisms of MoDTC by ZnDTP and on the possible implications of sulfurized MoDTC on tribological performances.

HPLC-MS analysis of MoDTC species showed that at $T = 0$ h, in the presence of ZnDTP **2b** or **2c**, MoDTC **1a–1c** could be mainly detected. The small amounts of 1S-MoDTC detected at $T = 0$ h were likely already present in the commercial MoDTC additive. After 3 h at 150 °C, the relative proportions of 1S-MoDTC (**1S-1a–1c**) increased and 2S-MoDTC species (**2S-1a–1c**) progressively appeared, the kinetics of formation of the sulfurized species differing in function of the nature of the ZnDTP substrates used (*i.e.*, mixture of ZnDTP **2b** or **2c**, respectively) (Fig. 3a and b). Comparison of the relative abundance of 2S-MoDTC after 3 h of reaction (Fig. 3a and b) showed that sulfurization of MoDTC occurred faster in the presence of di-isooctyl ZnDTP **2c** as compared to the secondary ZnDTP **2b**. After 6 h at 150 °C, the complete disappearance of the MoDTC substrates **1a–1c** and of the related 1S-MoDTC was observed as a result of their extensive transformation into 2S-MoDTC species (Fig. 3a and b). By contrast, the reaction with primary ZnDTP **2a** resulted in a very low degree of sulfurization of MoDTC under the same experimental conditions, leading to moderate formation of 1S-MoDTC even after 6 h at 150 °C (Fig. 3c). These results thus clearly demonstrate the capacity of di-isooctyl ZnDTP **2c** and secondary ZnDTP **2b** to efficiently sulfurize MoDTC in a concentrated medium without hydrocarbon base oil, whereas primary ZnDTP **2a** is apparently much less efficient.

3.2 Sulfurization of MoDTC by ZnDTP in diluted medium

Heating experiments involving MoDTC **1a–1c** and ZnDTP (**2a**, mixture of **2b**, or **2c**) have also been performed in a diluted medium at concentrations compatible with those used in commercial formulated oils (*i.e.*, 1 wt%) using a hydrocarbon

base oil at 135 °C under argon atmosphere for a period of 15 h. The evolution of the sulfurization reaction was monitored by HPLC-MS analysis of the samples collected at every 3 h intervals.

Fig. 4a (respectively Fig. 4d) shows the evolution of the relative concentrations of MoDTC substrates **1a–1c** upon heating with ZnDTP **2b** (respectively **2c**). It could be determined that the concentrations of the related sulfurized MoDTC species progressively increased over time (Fig. 4b and e), whereas those of the MoDTC substrates decreased gradually, notably due to oxygen–sulfur substitution on MoDTC derivatives. Detection of 1S-MoDTC and of traces of 2S-MoDTC at the starting point of the reaction ($T = 0$ h) could be partly explained, as mentioned above, by the presence of minor amounts of 1S-MoDTC (but not 2S-MoDTC) in the commercially purchased MoDTC substrates. Furthermore, since the initial sample ($T = 0$ h) was collected only when the reaction mixture has reached the temperature of 135 °C, it is likely that sulfurization of MoDTC by ZnDTP already started during the preliminary heating step, thus explaining the formation of 2S-MoDTC at this early stage. Nonetheless, the gradual increase of the different sulfurized MoDTC species over time clearly showed the capacity of the secondary ZnDTP **2b** and of the di-isooctyl ZnDTP **2c** to efficiently sulfurize MoDTC in solution at concentrations similar to those used in formulated lubricants. Fig. 4c and f show the global evolution of the proportion of sulfurized MoDTC (sum of 1S- and 2S-MoDTC) relative to the sum of the total MoDTC pool (*i.e.*, MoDTC and sulfurized MoDTC) over time. After 15 h of reaction, the percentage of sulfurized MoDTC increased gradually from 9% to 60% in the presence of secondary ZnDTP **2b** and from 24% to 63% in the presence of di-isooctyl ZnDTP **2c**. The higher proportion of sulfurized MoDTC already observed at $T = 0$ h in the presence of di-isooctyl ZnDTP **2c** is likely due to the higher sulfurization efficiency of this reagent during the initial heating step as compared to the secondary ZnDTP **2b**. Similar experiments involving primary 2-ethylhexyl ZnDTP **2a** showed a much lower degree of sulfurization of MoDTC (<20% of sulfurized MoDTC; data not shown). These results are coherent with those from the experiments under concentrated conditions (see Section 3.1).

3.3 Synthesis of 2S-MoDTC

Since the formation of sulfurized MoDTC can be expected to favor the generation of MoS₂ tribofilms, this process can be thought to contribute, at least to some extent, to the synergistic improvement of the tribological performances of MoDTC when used in combination with ZnDTP. In this respect, synthesis of 2S-MoDTC from MoDTC **1a–1c** has been carried out in order to evaluate their tribological properties and compare them to those of the “classical” MoDTC. In addition, since sulfurization of MoDTC may potentially result in the increase of the functional lifetime of MoDTC during engine operation, the evolution of 2S-MoDTC under thermal and thermo-oxidative conditions was investigated.

Lawesson's reagent (**3**) is a reagent able to induce the thionation of ketones and esters (*e.g.* Ozturk *et al.*, 2007⁴²). It has



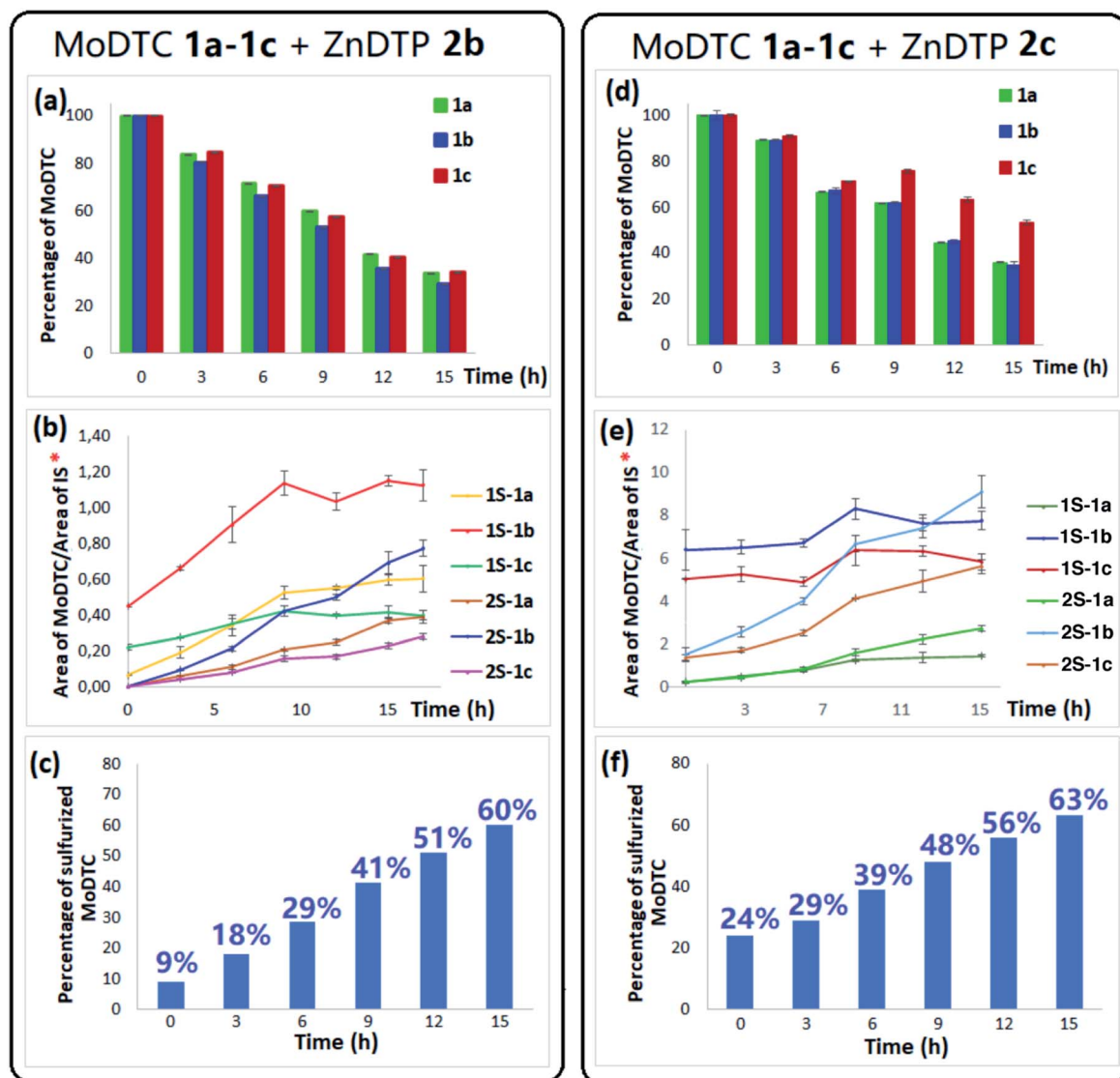


Fig. 4 Evolution of the relative abundances of (a and d) MoDTC substrates 1a–1c; (b and e) sulfurized MoDTC 1S-1a–1c and 2S-1a–1c formed during the heating experiment involving MoDTC 1a–1c and ZnDTP 2b and 2c, respectively, in lubricant base oil at 135 °C under argon atmosphere; (c and f) evolution of the proportion of sulfurized MoDTC (sum of 1S-1a–1c and 2S-1b–1c) relative to the sum of the (regular plus sulfurized) MoDTC. IS: Internal standard. *Y-axis: arbitrary units. Error bars (Fig. 4a–e) correspond to triplicate HPLC-MS analyses for each sample.

also been used to sulfurize the Mo=O bond of Mo complexes in a context different to that related to the present study.⁴³ We have, therefore, tested the capacity of this reagent to sulfurize MoDTC 1a–1c under the same conditions as those used for the sulfurization of MoDTC by ZnDTP under diluted conditions in a hydrocarbon base oil (see Section 2.2.2). Reaction of a mixture of MoDTC 1a–1c with Lawesson's reagent (3) at 135 °C under argon atmosphere led to the complete conversion of the starting MoDTC 1a–1c into their related disulfurized counterparts during the heating phase from room temperature to 135 °C as determined using HPLC-MS analysis of the crude mixture. The kinetic of the sulfurization reaction of MoDTC by Lawesson's reagent is thus obviously very fast under these conditions and this reagent can therefore be considered as an excellent

sulfurization agent of MoDTC. Efficient sulfurization could also be observed when MoDTC (mixture of 1a–1c or pure 1d) were heated with Lawesson's reagent under reflux in toluene (*i.e.*, at 110 °C). The latter procedure allowed synthesis in a quantitative yield of a mixture of 2S-1a–1c and of 2S-1d as determined by HPLC-MS analysis and NMR analysis (2S-1d, see Section 2.7).

3.4 The fate of 2S-MoDTC under thermal and thermo-oxidative conditions and evaluation of their tribological performances

3.4.1 The fate of 2S-MoDTC in a diluted medium. The stability of sulfurized MoDTC substrates 2S-1a–1c in hydrocarbon base oil with bubbling argon, air or NO₂ (2000 ppm in air) at 135 °C was investigated using laboratory ageing



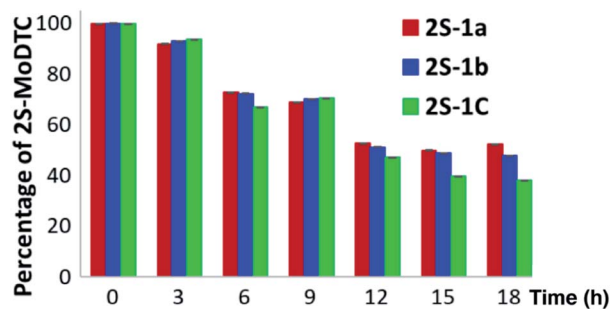


Fig. 5 Evolution of the relative abundance of MoDTC 2S-1a–1c in a hydrocarbon base oil at 135 °C under argon bubbling over a period of 18 h. IS: internal standard. Error bars correspond to the triplicate HPLC-MS analyses for each sample.

experiments (see ESI Fig. S3† for the visual aspect of the oil samples collected during the experiments involving 2S-MoDTC substrates under argon and air bubbling).

3.4.1.1 Argon atmosphere. For the experiment under argon bubbling, HPLC-MS analyses of the oil samples showed a gradual decrease of the relative concentrations of sulfurized MoDTC substrates 2S-1a–1c with time, with *ca.* 50% of the initial MoDTC remaining after 18 h (Fig. 5). This decrease indicates that 2S-MoDTC are unstable at elevated temperature

under an inert atmosphere. No thermal degradation products of 2S-MoDTC could, however, be detected under our HPLC-MS analytical conditions. Since it was shown previously that MoDTC 1a–1c in hydrocarbon base oil remained stable under the same conditions,⁴¹ it appears that sulfurized MoDTC 2S-1a–1c are less stable than their MoDTC 1a–1c counterparts and are susceptible to decompose by a yet unknown process under inert atmosphere conditions (no degradation products of Mo derivatives could be identified using our developed HPLC-MS analytical method). Despite being thermally unstable at high temperatures, 2S-MoDTC appeared, however, to be very stable at ambient temperature according to the NMR and HPLC-MS analysis of the synthesized 2S-MoDTC left at room temperature (*i.e.*, between 20 °C and 35 °C) under air atmosphere for a period of seven months.

3.4.1.2 Air atmosphere. Fig. 6a–c show the evolution of sulfurized MoDTC 2S-1a–1c and of MoDTC 1a–1c during laboratory oil ageing experiments under air bubbling at 135 °C. The concentrations of both series of MoDTC decreased over time, the degradation rate of 2S-MoDTC (Fig. 6a–c, blue colour) being however much faster than that of the genuine MoDTC under these conditions (Fig. 6a–c, red colour), leading to total disappearance of the 2S-MoDTC substrates after 8 h whereas *ca.* 50% of the MoDTC species 1a–1c were still present after 15 h of ageing.

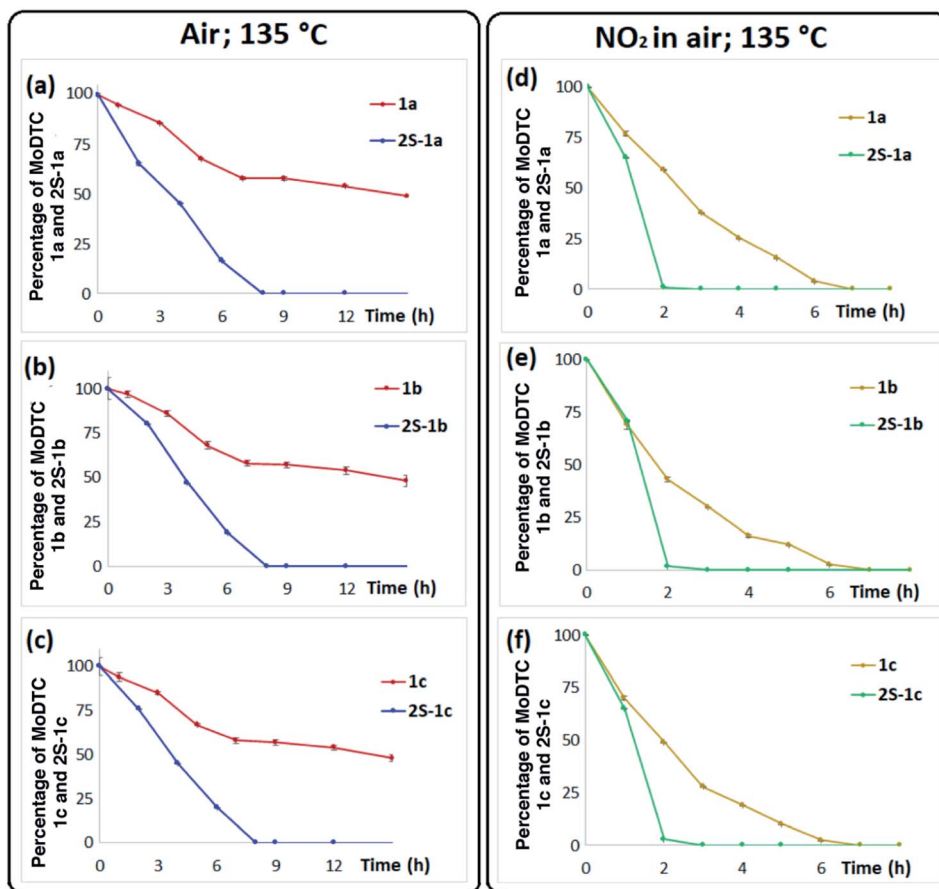


Fig. 6 Evolution of the relative abundance of sulfurized MoDTC 2S-1a–1c and of MoDTC 1a–1c upon heating at 135 °C in a hydrocarbon base oil under (a–c) air; (d–f) NO₂ (2000 ppm in air) bubbling. Error bars correspond to triplicate HPLC-MS analyses for each sample.



3.4.1.3 NO₂ atmosphere. An even higher degradation rate was observed when 2S-MoDTC and regular MoDTC were subjected to ageing experiments under nitro-oxidation conditions at 135 °C as compared to air conditions. Substrates **2S-1a–1c** were almost completely consumed after only 2 h of experiment (Fig. 6d–f, green colour) whereas MoDTC **1a–1c** persisted for a period of ca. 7 h (Fig. 6d–f, yellow colour). As already pointed out in the case of the experiment under air atmosphere, these results clearly confirm that 2S-MoDTC are much less stable than their non-sulfurized counterparts and tend to decompose at a significantly faster rate under thermo-oxidative conditions.

It could however be determined by HPLC-MS analysis that the progressive degradation of 2S-MoDTC under thermo-oxidative conditions (*i.e.*, under air and NO₂ in air atmosphere) was accompanied by the formation of new MoDTC species corresponding to mono-sulfurized MoDTC **1S-1a–1c** (Fig. 2: structure **d** and ESI Fig. S4† for the mass spectra of **1S-1a–1c**) and of the “classical” MoDTC **1a–1c**. These new species most likely originated from sulfur–oxygen substitution on the Mo₂S₄ core of the 2S-MoDTC. Thus, even though 2S-MoDTC tend to decompose very quickly under thermo-oxidative conditions, their detected transformation products (*i.e.*, 1S-MoDTC and “classical” MoDTC) persist in the oil for a longer period of time (Fig. 7a–d) and could potentially continue to contribute to friction reducing capabilities of lubricants.

On the other hand, the degradation pathway of MoDTC **1a–1c** under air and NO₂ (2000 ppm in air) at 135 °C that ultimately leads to the formation of molybdenum oxides could involve the intermediate formation of oxygenated MoDTC species, which

correspond to MoDTC derivatives with one or two sulfur atom(s) originally present in **1a–1c** being replaced by oxygen atom(s) (see ESI Fig. S5 and S6† for the mass spectra and quantitative analysis of oxygenated MoDTC species).

3.4.2 Tribological performances of 2S-MoDTC. The capacity of MoDTC and 2S-MoDTC to reduce friction under nitro-oxidation conditions was also studied. For this purpose, the 1 mL samples collected during the course of oil ageing experiments involving MoDTC **1a–1c** or synthesized sulfurized MoDTC **2S-1a–1c** in hydrocarbon base oil under a bubbling flow of NO₂ (2000 ppm in air) at 135 °C (see Section 2.2.2 and 3.4.1) were used to measure the friction coefficient using a ball-on-flat tribometer. The evolution with time of the concentrations of 2S-MoDTC, 1S-MoDTC and MoDTC in both experiments (see Section 3.4.1, Fig. 6 and 7) was compared to that of the measured friction coefficients (Fig. 8).

The base oil containing 2S-MoDTC displayed low values of friction coefficients (≈ 0.07) during the first 5 h of nitro-oxidation, before showing a sharp decrease of the friction reducing capabilities (Fig. 8). Even though 2S-MoDTC were shown to be unstable and were completely degraded within 2 h under nitro-oxidation conditions (Fig. 6d–f), their oxidative transformation products (*i.e.* 1S-MoDTC and “classical” MoDTC) remained in the oil for longer periods than 2S-MoDTC (Fig. 7c and d) and were most likely still active as friction reducers accounting for the preservation of the friction reducing capabilities up to 5 h. Surprisingly, although MoDTC in the experiment involving “classical” MoDTC as well as MoDTC resulting from oxidative O–S exchanges in the experiments involving 2S-MoDTC have been extensively consumed, in both cases, after 7 h under nitro-oxidation

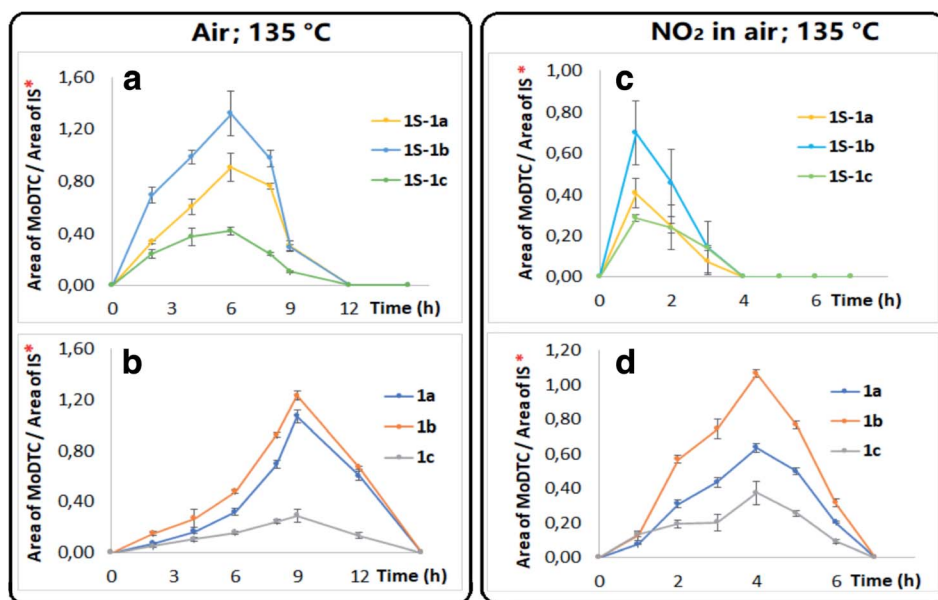


Fig. 7 Evolution of the concentration of MoDTC complexes **1S-1a–1c** and MoDTC complexes **1a–1c** as oxidative degradation products formed during the heating experiment involving 1 wt% of sulfurized MoDTC substrates **2S-1a–1c** under (a and b) air bubbling at 135 °C and (c and d) NO₂ bubbling (2000 ppm in air) at 135 °C. IS: internal standard. *Y-axis: arbitrary units. Error bars correspond to triplicate HPLC-MS analyses for each sample.



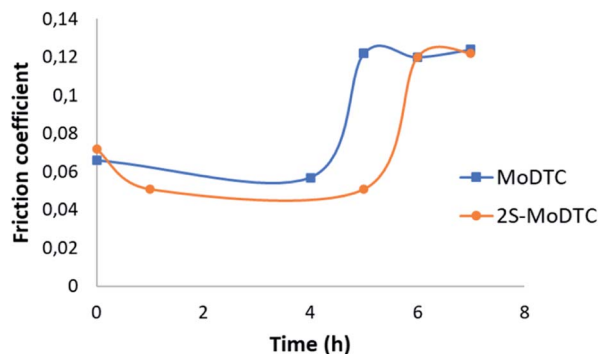


Fig. 8 Evolution with time of the friction coefficient of MoDTC 1a–1c and di-sulfurized MoDTC 2S-1a–1c in solution in a hydrocarbon base oil subjected to oil ageing experiments under NO₂ (2000 ppm in air).

conditions (*cf.* Fig. 6d–f, 7c and d, respectively), friction reducing capacities were preserved for a longer period of time in the case of the oil containing 2S-MoDTC (5 h vs. 4 h, Fig. 8). This result suggests that the use of 2S-MoDTC, although displaying a comparable friction coefficient than the “classical” MoDTC at $T = 0$ h, apparently results in a better retention of friction reducing capability with time under thermo-oxidative conditions.

4. Discussion

4.1 Mechanism of the sulfurization of MoDTC by ZnDTP and by Lawesson's reagent

The mechanism of sulfurization of MoDTC by ZnDTP additives is likely related to the affinity between phosphorus and oxygen atoms, which favours sulfur–oxygen exchange reactions on phosphorus. In this respect, sulfur-bearing phosphorus derivatives such as Lawesson's reagent **3** (*e.g.*, Ozturk *et al.*, 2007⁴²) and P₂S₅ (*e.g.* Campaigne and Foye, 1952;⁴⁴ Sakurai *et al.*, 1971;⁴⁵ Isomaya and Sakurai, 1974;⁴⁶ Malik *et al.*, 2008⁴⁷) are used in organic synthesis to induce sulfurization of oxygenated functions. In the case of the sulfurization of MoDTC by ZnDTP, a mechanism similar to that proposed by Partyka *et al.* (2003)⁴³ and Moula *et al.* (2013)⁴⁸ for the sulfurization of Mo=O bonds with Lawesson's reagent can be proposed (Fig. 9a).

Although the formation of mixed Mo(DTC)(DTP) complexes is not favored in the case of mixtures comprising MoDTC and ZnDTP, they seemed, however, to be formed in small proportions as demonstrated by Kiw *et al.* (2020)⁴¹ and could act as intermediates during the sulfurization of MoDTC following the mechanism presented in Fig. 9b. Step II of this mechanism presents some analogies with the rearrangement reaction proposed by several authors to play a role in the tribochemical formation of MoS₂ from MoDTC or in the thermo-oxidative degradation of MoDTC.^{27,49,50} It involves an exchange between a sulfur atom from the ligand and an oxygen atom on the MoDTC core. Another possible mode of formation of 2S-MoDTC could be the sulfurization of MoDTC by H₂S resulting from the thermal decomposition of ZnDTP complexes described by Dickert and Rowe (1967)⁵¹ and Spikes (2004).³¹

4.2 Possible influence of the ZnDTP structures on their capacity to sulfurize MoDTC

According to the present study, the laboratory heating experiments involving MoDTC 1a–1c and ZnDTP (2b or 2c respectively), either in concentrated or diluted solutions in hydrocarbon base oils clearly showed that some ZnDTP species are more able to induce sulfurization of MoDTC, leading to the formation of 1S- and 2S-MoDTC. It appeared that the secondary ZnDTP 2b with C₃ and C₆ alkyl chains seem to sulfurize MoDTC more efficiently than the primary ZnDTP 2a bearing 2-ethylhexyl (C₈) chains. This suggests that the chain length and structure of the side-chains and thus, the steric hindrance induced by these alkyl chains, might influence the kinetic of the sulfurization reaction. Nevertheless, the fact that ZnDTP 2c bearing isoocetyl chains is even a more efficient sulfurization agent than ZnDTP 2a – both bearing C₈ alkyl chains, but with different structures – rather seems to indicate that the length of the side chain is not the key structural feature of ZnDTP that controls their capacity to sulfurize MoDTC. Structural investigation of the nature of the ligands from ZnDTP 2c – no information being available for this commercial additive – showed that it bears DTP ligands with a great variety of isomeric C₈ alkyl chains (*cf.* ESI Fig. S7† for the gas chromatogram of the dithiophosphate ligands of ZnDTP 2c analyzed as methylated derivatives). It therefore appears that the alkyl chains from ZnDTP 2c have most likely different modes of linkage to the dithiophosphate core (*i.e.*, via C–O bonds at primary, secondary or even tertiary positions). Thus the capacity of ZnDTP 2c to sulfurize rapidly MoDTC is possibly related, as would be the case for ZnDTP 2b, to the occurrence of alkyl chains linked at secondary (and possibly also tertiary?) positions. This led us to propose that the modes of linkage of the alkyl chains to the dithiophosphate core play a key role in the efficiency of ZnDTP to sulfurize MoDTC. However, it is speculative at this stage to restrict the differences of reactivity of the various ZnDTP to a specific structural feature related to the nature of their ligands due to a lack of precise information on the structures of the isomeric C₈ alkyl chains occurring on ZnDTP 2c and the absence of pure commercially available ZnDTP reference compounds bearing single secondary or tertiary alkyl chains (instead of isomeric mixtures).

The kinetic of the formation of 2S-MoDTC from MoDTC 1a–1c showed, at least in the case of the reaction with ZnDTP 2c (Fig. 4), that MoDTC 1a and 1b bearing C₈ alkyl chains underwent sulfurization more rapidly than MoDTC 1c which bears exclusively C₁₃ alkyl chains (Fig. 4d). In parallel, the concentrations of sulfurized MoDTC 2S-1a and 2S-1b seemed to increase faster than that of the homologue 2S-1c (Fig. 4e). This could mean that the steric hindrance induced by the MoDTC alkyl chains also plays a role in their capacity to be sulfurized by ZnDTP. The apparently more rapid consumption of MoDTC with C₈ alkyl chains relative to the MoDTC bearing exclusively C₁₃ chains also implies that ligand exchange reactions occurring between the various MoDTC (sulfurized or not) are less rapid than the sulfurization reaction under the applied thermal conditions since, otherwise, no preferential consumption/formation of specific homologues would occur.



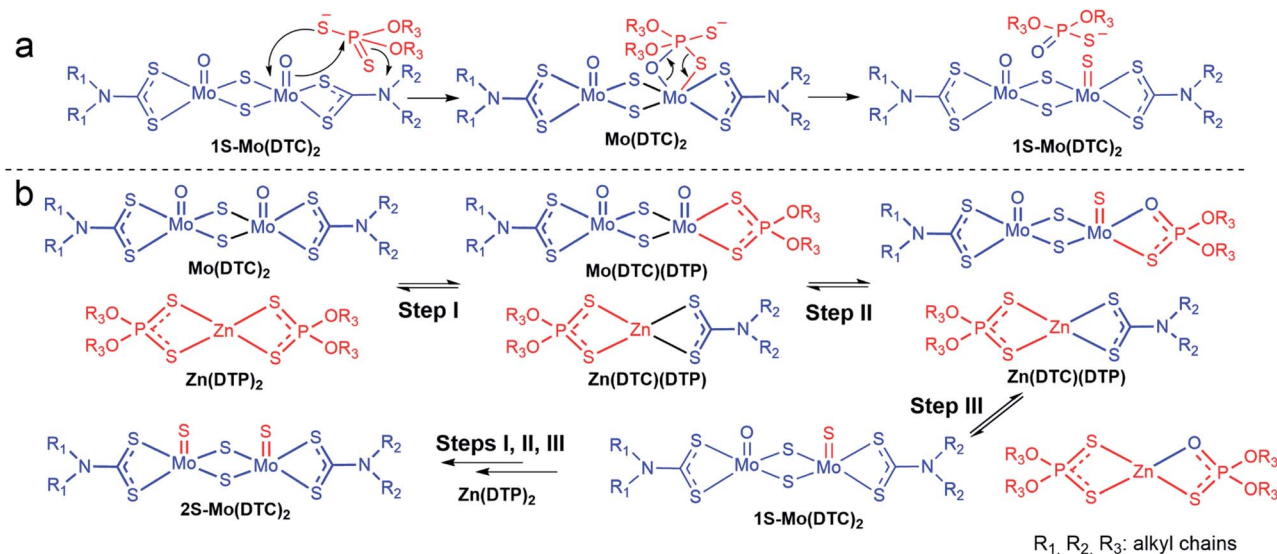


Fig. 9 Proposed mechanisms of sulfurization of MoDTC by ZnDTP through (a) direct exchange of O and S atoms between Mo=O and P=S bonds; or (b) O and S exchange involving Mo-bound DTP ligands. I, III: ligand exchange reactions, II: rearrangement involving an exchange between a sulfur atom from a DTP ligand and an oxygen atom on the MoDTC core.

4.3 Effect of 2S-MoDTC on friction reduction efficiency

According to the tribological studies carried out with the synthesized 2S-MoDTC (see Section 3.4.2; Fig. 8), 2S-MoDTC additives do not appear to reduce more significantly the friction coefficient than the “classical” MoDTC do. Therefore, sulfurization of MoDTC by ZnDTP in solution (sulfurization in bulk) does most likely not contribute to the synergistic improvement of friction coefficients sometimes observed when MoDTC are used in combination with ZnDTP.³⁴ Nonetheless, compared to the “classical” MoDTC species, 2S-MoDTC were shown to display extended periods of friction reducing capability in lubricants under thermo-oxidative conditions as compared to “classical” MoDTC. A possible explanation could be related to the presence of higher amounts of available sulfur for the formation of the MoS₂ tribofilm due to the initial presence of 2S-MoDTC. Alternatively, it can be considered that the newly-formed active Mo species deriving from 2S-MoDTC and favoring the formation of the tribofilm were out of our analytical window. Whatever the reason is, our study indicates that sulfurization of MoDTC by ZnDTP could be considered as a process contributing to delay the decrease of the tribological performances of lubricants containing MoDTC.

5. Conclusions

The present study involving thermal ageing experiments of MoDTC in the presence of ZnDTP species, combined with HPLC-MS monitoring of the evolution of the resulting MoDTC species, revealed the potential of ZnDTP additives to sulfurize MoDTC in lubricants. This process leads to substitution of Mo-bound oxygen atom(s) by sulfur atom(s) at temperatures and concentrations compatible with those undergone by lubricants under engine functioning conditions. A difference in the reactivity between ZnDTP additives towards sulfurization was

observed, primary ZnDTP species leading to a lower degree of sulfurization as compared to secondary ZnDTP. A mechanism of sulfurization of MoDTC by ZnDTP was proposed, possibly involving mixed Mo(DTC)(DTP) intermediates. In addition, the capacity of Lawesson’s reagent to (almost) quantitatively sulfurize MoDTC complexes was evidenced and was thus used to synthesize 2S-MoDTC reference compounds that were used for the tribological tests.

Contrary to what could be expected based on structural considerations (*i.e.*, enhanced capacity to form a MoS₂ tribofilm), tribological measurements indicate that the use of 2S-MoDTC does not significantly improve the coefficients of friction when compared to their non-sulfurized MoDTC counterparts. Nevertheless, it appeared that under nitro-oxidation conditions, the use of synthetic 2S-MoDTC as friction modifiers in a hydrocarbon base oil during laboratory ageing experiments led to the retention of the friction reducing capacities for a longer period of time than the “classical” MoDTC, although the mechanism(s) involved remain(s) to be elucidated. Sulfurization of MoDTC might thus contribute to delay the decrease with time of the overall tribological properties of lubricants containing MoDTC when exposed to thermo-oxidative conditions. To further confirm this hypothesis, future work will be dedicated to study the implications of the sulfurization processes of MoDTC by ZnDTP on the tribological performances of fully formulated lubricants and on the persistence MoDTC derivatives under thermo-oxidative conditions. The possibility of using 2S-MoDTC instead of MoDTC (or a combination of both) as additives in formulated engine lubricants will also be investigated.

Conflicts of interest

There are no conflicts to declare.



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