PAPER



View Article Online View Journal | View Issue



RSC Advances

Cite this: RSC Adv., 2022, 12, 5080

Photocatalytic turnover of CO₂ under visible light by [Re(CO)₃(1-(1,10) phenanthroline-5-(4-nitronaphthalimide))Cl] in tandem with the sacrificial donor BIH†

Alyssa Spear, Da Robson L. Schuarca, Jesse Q. Bond, Db Timothy M. Korter, Da Jon Zubieta*a and Robert P. Doyle *D *a

Optimized photocatalytic conversion of CO₂ requires new potent catalysts that can absorb visible light. The photocatalytic reduction of CO₂ using rhenium(i) has been demonstrated but suffers from low turnover. Herein, we describe a $[Re(CO)_3(1-(1,10))$ phenanthroline-5-(4-nitro-naphthalimide))Cl] photocatalyst, which when combined with the sacrificial donor 1,3-dimethyl-2-phenyl-2,3-dihydro-1H-benzo[d]imidazole, results in selective production of formic acid and a high turnover number of 533 and turnover frequency of 356 h^{-1} . Single-crystal X-ray diffraction and DFT studies are also discussed.

Received 11th November 2021 Accepted 4th February 2022

DOI: 10.1039/d1ra08261b

rsc.li/rsc-advances

Introduction 1.

Increasing levels of atmospheric and anthropogenic CO₂ require the development of new technologies and strategies in the pursuit of sustainable energy and climate preservation. 1-3 Homogenous mononuclear rhenium(1) complexes have been explored as reductive catalysts for the conversion of CO₂ to fuels and feedstocks, primarily CO.4,5 While this subset of catalysts has been tuned for high turnover numbers (TN, [CO₂ reduction product]/[catalyst]), notably as high as 350,6 rarely is selective CO₂ to formic acid conversion reported. Formate and formic acid are attractive products due to ease of transportation and storage, and high energy densities, and as favourable candidates for hydrogen storage.7-9 Any reports of CO2 to formate reduction by mononuclear rhenium(1) complexes are typically in trace amounts and hence there is an unmet need to tune such complexes towards formate/formic acid selectivity if they are to provide a viable option for CO2 conversion to more diverse products, especially under visible light.

(1,10)-phenanthroline-5-(4-nitro-naphthalimide)Cl] (1). 1 consists of a rhenium(1) tricarbonyl core with

a naphthalimide based photosensitizing ligand, (1,10)phenanthroline-5-(4-nitro-naphthalimide) (5-PAN), for the visible light photoreduction of CO₂ to formate (Fig. 1).10 A sacrificial electron donor, triethylamine (TEA), was required for photocatalysis.7 The catalytic activity of 1 increased with a concomitant increase in TEA, and no CO2 conversion was obtained in the absence of light (450-460 nm). Catalytic saturation was reached at 0.87 mmol of HCOO-, corresponding to a TN ([HCOO-]/[1]) of 48.10

In the present research, we aim to discern a mechanism for the system and improve the photocatalytic ability of 1. We synthesized bromide analogue, $[Re(CO)_3(1-(1,10)-$

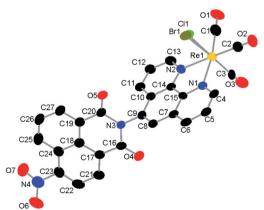


Fig. 1 Single-crystal X-ray diffraction structure of a mixture of 1 and 2. H-atoms and toluene solvent molecule are omitted for clarity. The halido ligand site is modelled as a disorder of Br: Cl, 0.80: 0.20. Structure deposited as CCDC # 2086581.

Previously, we have reported on the synthesis, characterization, electrochemistry, and catalytic activity of a [Re(CO)₃(1-

^aDepartment of Chemistry, Syracuse University, Syracuse, NY 13244, USA. E-mail: jazubiet@syr.edu; rpdoyle@syr.edu

^bDepartment of Biomedical and Chemical Engineering, Syracuse University, Syracuse, NY 13244, USA

[†] Electronic supplementary information (ESI) available: Details of the X-ray determinations, bond lengths and angles. CCDC 2086581. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d1ra08261b

phenanthroline-5-(4-nitro-naphthalimide)Br] (2), to evaluate the effect of the halide. Moreover, we obtained crystals of combined 1 and 2 which allowed X-ray diffraction studies and subsequent DFT calculations reported herein. To further drive CO_2 reduction, we replaced TEA with a 2-electron proton donor in the form of 1,3-dimethyl-2-phenyl-2,3-dihydro-1*H*-benzo[*d*] imidazole (BIH), which has been shown to significantly quench the excited state of rhenium(i) photocatalysts thereby yielding better turnover. 11,12 Consequently, we have combined BIH with 1 and 2 in the presence of blue light (450–460 nm) to produce formate in high selectivity over CO.

2. Experimental

2.1. Materials and methods

CDCl₃, CD₃OD, deuterated DMSO, Re(CO)₅(X) (X = Cl, Br), 2-phenylbenzimidazole, iodomethane, sodium borohydride, sulfuric acid, and diethyl ether were purchased through Sigma Aldrich (Milwaukee, WI, USA). Sodium hydroxide was purchased through Fisher Scientific (Agawam, MA, USA). Methanol (MeOH), ethanol, and dichloromethane (DCM) were purchased through Sigma Aldrich (Milwaukee, WI, USA) and dried over 4 Å molecular sieves (Sigma-Aldrich) in-house to remove water. Barium carbonate (13 C, 98%+) was purchased from Cambridge Isotope Laboratories, Inc (Tewksbury, MA, USA). 1 M H₂SO₄ was prepared in-house. A 12 W LED blue light (450–460 nm) was purchased from Abi, USA (Vernon, CA, USA).

¹H/¹³C NMR were conducted on a Bruker Avance III HD 400 MHz instrument and calibrated to the residual CDCl₃ signal at 7.24 ppm, D1 = 5s for 13 C NMR. FT-IR was carried out on a Nicolet Infrared Spectrophotometer utilizing potassium (KBr) disks. Crystals bromide of $[Re(CO)_3(1-(1,10))]$ phenanthroline-5-(4-nitro-naphthalimide))X] (X = Cl or Br) suitable for single crystal X-ray diffraction studies were only obtained by combining 1 and 2 (1:4 by mass) in MeOH/DCM (1:9 v/v) and gently layering toluene on top, then sealing the glass vial with a cap and parafilm. Gas chromatography (GC) was conducted on an HP-6890 instrument with a ShinCarbon packed column (Restek, PA, USA), a Flame Ionization Detector (FID), and a methanizer. The FID was calibrated with the standards 1% CO in helium, 1% CO₂ in helium, and 0.5% CH₄ in helium (Airgas). All injections were done with a calibrated 1 mL syringe (Agilent). Ion chromatography was conducted by Intertek Pharmaceuticals (Whitehouse, NJ, USA).

2.2. Synthetic procedures

1 was prepared as previously reported. ¹⁰ 2 and BIH were prepared as described below. $Re(CO)_5Br$ was refluxed in water for 24 h to produce $[Re(CO)_3(H_2O)_3Br]$. ¹³ The ligand 5-PAN was prepared as previously reported. ¹⁰ $[Re(CO)_3(H_2O)_3Br]$ (0.26 mmol, 105 mg) and 5-PAN (0.26 mmol, 110 mg) were combined in 30 mL MeOH/DCM (1:9 v/v), put under argon, and stirred for 1.5 hours. The solution volume was reduced to half *in vacuo* and diethyl ether was added to precipitate the product. The mixture was centrifuged, and the ether decanted.

The yellow solid noted was collected and dried under vacuum overnight giving 68% yield based on rhenium.

The synthesis of BIH was modified from literature procedures. 14,15 2-Phenylbenzimidazole (3 g) was mixed with methyl iodide (8 g) in 15 mL of anhydrous MeOH containing NaOH (0.64 g) in a pressure tube. The tube was purged with argon then sealed and heated at 110 °C with stirring for 3 days. After cooling to room temperature, crystals formed. The crude intermediate was filtered and washed with MeOH and dried *in vacuo*. The intermediate was dissolved in 80 mL of anhydrous MeOH and NaBH₄ (1.5 g) was slowly added and stirred for 1 hour under argon. The solvent was then evaporated, and the product was recrystallized in hot ethanol/water 5:1 to give white crystals of BIH in 75% yield.

2.3. Density functional theory

The electronic structure of the metal complex was simulated with Gaussian 16 Rev C.01 using the B3LYP density functional and the def2-TZVP basis set (all-electron basis functions for atoms except Re, which utilized an effective-core potential). The starting geometry of the complex was taken from the experimental crystallographic structure and the structure was then fully geometry optimized. Solvent effects were included in all calculations using the Polarizable Continuum Model (nonequilibrium for the excited states) with DCM as the specified solvent. Vertical electronic transitions were calculated with time-dependent DFT based on the Tamm-Dancoff approximation (TDA) and included 100 states (singlets and triplets) to explore the possible transitions.

2.4. Typical photocatalysis experiments

The following is a typical experiment using 1. In each sample tested, 0–110.3 equivalences of BIH (0–2 mmol), were premixed in MeOH/DCM (1 : 9 v/v) for 10 min, after which, 13.5 mg (0.018 mmol) of 1 was added. All samples were kept at a 5 mL total reaction volume, irradiated with blue light positioned 1 inch from the reaction vessel and at a 21° angle of incident light. The solution was bubbled with $^{13}\mathrm{CO}_2$ evolved *in situ* with 0.500 g Ba $^{13}\mathrm{CO}_3$ and 1 M $\mathrm{H}_2\mathrm{SO}_4$ drip for a reaction time of 1.5 hours. For each trial, the drip rate was set such that the Schlenk flask containing Ba $^{13}\mathrm{CO}_3$ was full at the end of 1.5 hours and the $\mathrm{H}_2\mathrm{SO}_4$ was well in excess. All catalytic experiments were conducted at least in triplicate under blue light irradiation.

At the end of each trial, liquid products were quantified *via* 13 C NMR. 400 μ L of the reaction solution and 250 μ L of CDCl $_3$ were added to an NMR tube and wrapped in aluminium foil until it was ready to be loaded into the NMR. The carbon peaks for BIH and its oxidative product, BI $^+$, at \sim 33 ppm were integrated and set to 1. The H 13 COO $^-$ peak at \sim 160 ppm was integrated and quantified with a calibration curve (S12 $^+$).

For the detection of gaseous products, 1 and BIH were dissolved in MeOH/DCM (1:9 v/v) in a vial with a stir bar and septum cap. The vial was purged with CO_2 for 20 min, then irradiated with blue light for 30 min. A headspace sample was taken with a 1 mL airtight syringe and analysed by GC.

RSC Advances Paper

3. Results and discussion

3.1. **Crystal structure**

We sought to obtain a solid-state structure of the catalytic species to facilitate DFT studies. Single-crystal X-ray diffraction analysis revealed that the catalyst crystalizes in the triclinic space group $P\bar{1}$ with two molecules in the unit cell, one catalyst and one toluene solvate molecule (S8†). Diffraction quality crystals were only obtained from a mix of 1 and 2; thus, the six-coordinate, octahedral rhenium center includes a disordered halido ligand comprised of Br: Cl, 0.80: 0.20. The CO ligands are facially coordinated, and the ligand, 5-PAN, is coordinated to the rhenium center in a bidentate fashion at N1 and N2 (Fig. 1). Detailed description of the structure is included in ESI†.

3.2. Density functional theory

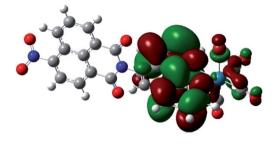
The full geometry optimization yielded a final structure that was in good agreement with the X-ray diffraction results described herein for 1.

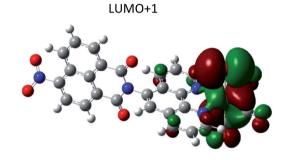
The TDA simulation yielded a predicted a UV/VIS spectrum that matched very well with the experimental data7 (S9†), indicating the high quality of the calculated orbital energies and participating states. Of particular interest to the photocatalytic behaviour are the four transitions with non-zero oscillator strength in the 400 to 500 nm region. These singlet transitions are detailed in Table 1 and the strongest transition, $HOMO-1 \rightarrow LUMO+1$, is visualised in Fig. 2. Other transitions are visualised in the ESI (S10)†.

3.3. Homogenous photocatalysis

Previously reported¹⁰ cyclic voltammetry of 1 indicated sufficient potentials of -0.95 V and -1.18 V for the two-electron reduction of CO₂ to formate (-0.53 V).4 Thus, 1 was evaluated for catalytic reduction of CO2 in the presence and absence of blue light (450-460 nm) utilizing ¹³C NMR tracking of H¹³COO⁻ (typical ¹³C NMR included in S11†).

It should be noted that at higher concentrations of BIH (≥1 mmol), the limit of solubility in MeOH/DCM (1:9 v/v) in the presence of catalyst was reached, requiring that the BIH be dissolved in the solvent system prior to addition of the catalyst. Without this order of dissolution, the catalytic production of $H^{13}COO^{-}$ at concentrations of BIH ≥ 1 mmol, was highly





HOMO-1

Fig. 2 Orbital density plots for $HOMO-1 \rightarrow LUMO+1$. The trend in these excitations is a migration of electron density from the rhenium atom to the ligand.

variable (S12†) as undissolved BIH would be misrepresented in the calibration curve (S13†).

The production of H¹³COO⁻ correlates with the addition of excess BIH and blue light in an exponential trend with R^2 0.9932 (Fig. 3). A plateau at 9.48 mmol indicated the maximum $\mathrm{H}^{13}\mathrm{COO}^{-}$ produced and catalytic saturation. The maximum TN achieved was 533, and turnover frequency (TOF = TN/reaction time) 356 h^{-1} at \sim 110 equivalences of BIH, where the concentration of 1 was 3.62 mM in 5 mL of MeOH/DCM (1:9 v/v). 2 was also tested for photocatalytic activity and showed no functional difference from that of 1, which is in contradiction to some literature reports which have seen halide degradation kinetics differences.16,17 No formate was detected in the absence of catalyst, BIH, or ¹³CO₂ and a negligible amount was detected in the absence of irradiation. The calculated TN of 533 not only surpasses our previously reported 1/TEA system tenfold but introduces a standard for selective turnover of CO2 to formate via a mononuclear rhenium(1) catalyst.

Table 1 Singlet transitions of 1

Excitation wavelength (nm)	Oscillator strength (f)	Main participating orbitals and percentage contributions
470.63	0.0016	HOMO → LUMO+1 (99%)
446.59	0.0680	$HOMO-1 \rightarrow LUMO+1 (90\%)$
		$HOMO \rightarrow LUMO+2 (8\%)$
426.62	0.0345	$HOMO-1 \rightarrow LUMO+1 (8\%)$
416.46	0.0172	HOMO → LUMO+2 (90%) HOMO-1 → LUMO+2 (98%)

Paper

15-(Town) 10-0.0 0.5 1.0 1.5 2.0 BIH mmol

Fig. 3 A plot of $\rm H^{13}COO^-$ (mmol) vs. BIH (mmol) upon blue light (450–460 nm) irradiation of 1 (blue circles, $R^2=0.9932$) and 2 (red squares, $R^2=0.9864$). $\rm H^{13}COO^-$ increased with an exponential curve fit and reached a plateau at y=9.45 and 8.96 mmol $\rm H^{13}COO^-$, respectively.

The production of $\mathrm{HCO_3}^-$ has also been reported appearing on $^{13}\mathrm{C}$ NMR at $\sim\!158$ ppm. 18 Ion chromatography was therefore conducted and confirmed the identity of formic acid, and lack of carbonate (S14†). Gaseous products of photocatalysis were also analysed by GC. Trace quantities of CO were detected at a maximum production of 820 nmol CO corresponding to a TN of 0.04 and a TOF of 0.09 h $^{-1}$ (S15†).

3.4. Solvent effects

The solvent system MeOH/DCM (1:9 v/v) was chosen because it best solubilized all of the reaction components, however MeOH has been reported to enhance the catalytic activity of rhenium diimine complexes by increasing the current in electrolysis experiments. To reveal the effect of MeOH in photocatalysis, experiments were done in the presence of deuterated methanol (along with DMF at typical 1:9 v/v ratio used before). When DCM was used as the lone solvent, no formate was produced as evidenced by the loss of the $\sim\!160$ ppm peak in the 13 C NMR and the subsequent emersion of a weak peak at $\sim\!156$ ppm with an integration of 0.6. When MeOH was replaced with CD₃OD (CD₃OD/DCM 1:9 v/v), the $\sim\!160$ ppm peak was reduced to an integration of 0.9 (S16†).

In DMF alone, a common solvent used in rhenium photocatalysis, the formate peak was also not seen in the 13 C NMR and a peak at \sim 156 ppm was seen with an integration of 7.6. It should be noted that the aldehyde carbon of DMF has a peak at 161 ppm, therefore a formate peak could be occluded. Nonetheless, the absence of MeOH resulted in the loss of the signature formate peak at 160 ppm and subsequent appearance of a peak slightly upfield at 156 ppm with greatly diminished integration. When experiments were conducted in MeOH/DCM with no CO₂, no carbon signal was observed above 150 ppm, thus the source of the carbon is indeed CO₂ with no degradation of MeOH itself. MeOH is likely necessary for catalysis as a major proton source.

3.5. Photocatalytic mechanism

A variety of factors including solvent, sacrificial donor, and halide ligand can affect the catalytic activity of rhenium diimine complexes. TLC analysis of the reaction mixture before and after irradiation showed no obvious degradation of the catalyst while BI+ did separate from BIH (S17†), which was also indicated by NMR. With the addition of BIH, compared to TEA, TN was dramatically increased possibly due to the increase in reductive power and efficient formation of the one-electron reduced catalytic species.²⁰ When TEMPO was added in equimolar concentrations to BIH, the integration of formate increased, suggesting TEMPO acting as a hydride donor. These observations allude to a hydride mechanism.

A possible mechanism is described (Fig. 4). After reductive quenching of the photoexcited catalyst by BIH, the halide dissociates, and the active catalyst is produced. Literature descriptions of formate generation then propose the formation of a metal hydride intermediate. $^{4,21-23}$ Insertion of CO_2 into the M–H bond then produces the intermediate. This process is in competition with H_2 generation, which cannot be detected as H_2 is used a co-feed gas for the methanizer during GC analysis.

In control experiments in the absence of light, formate was still detected, reaching a maximum amount of 3.20 mmol ($R^2 = 0.9037$) corresponding to a TN of 176 at 110 equivalents of BIH to 1. To explain the cause of this activity, we noted a literature report of BIH forming a radical cation in the presence of atmospheric O_2 without irradiation.²⁴ Under blue light irradiation, BIH alone, in MeOH/DCM (1:9 v/v), was able to reduce CO_2 to formate with a low turnover of 5.7 (TN = [formate]/ [BIH]). Given the sizable difference in TN from 5.7 to 533, we conclude that any BIH-only mechanism is overridden in the presence of 1 by the hydride mechanism.

The hydride mechanism shown in Fig. 4 only accounts for one BIH per catalytic cycle, whereas the TN is 533 with 110 equivalences of BIH. This suggests multiple mechanisms or components are contributing to the formation of formate herein. Notably, the role of methanol is underappreciated in the proposed mechanism and may account for the enhanced TN.

Fig. 4 Proposed hydride mechanism of CO_2 reduction to formate *via* 1 or 2. X = Cl or Br, Sol = DCM or MeOH.

4. Conclusion

We have reported an extraordinarily improved photocatalytic TN and TOF of a rhenium(1) naphthalimide complex in the presence of the sacrificial electron and proton donor, BIH. The crystal structure of [Re(CO)₃(1-(1,10)phenanthroline-5-(4-nitronaphthalimide)) X] (X = Cl or Br) was also reported, indicating facial coordination of three CO ligands, coordination of a halide in the solid state, and bidentate coordination of the naphthalimide-based photosensitizing ligand. The generated H¹³COO⁻ increased with excess BIH until a plateau was reached at 9.45 mmol of H13COOH. The corresponding TN at 110 equivalents of BIH to 1 is 533, a tenfold improvement to catalysis using 1 with the sacrificial donor TEA10 and marked selectivity over CO production, which was observed in nanomole quantities. MeOH/DCM was not only used for its ability to solvate the catalyst, but to optimize the production of H¹³COO⁻; DCM or DMF lead to diminished H¹³COO⁻ signal. While the addition of MeOH improved catalytic activity, the source of the carbon was CO₂. To date, this TN is the only reported value for a rhenium(1) mononuclear homogeneous photocatalyst for the selective photoreduction of CO₂ to formate.

1 is an exciting complex to apply towards CO_2 reduction as the field is developing new methods of synthesizing valuable chemicals while maintaining sustainable and mild conditions. Notably, at similar reduction potentials to those reported for $\mathbf{1}$, ¹⁰ electrosynthesis of methylamine from CO_2 and NO_3^- was recently reported²⁵ and will be explored with $\mathbf{1}/BIH$. Further mechanistic studies are needed to elaborate on the complex mechanism at play, especially the role of MeOH. The complex itself is an interesting avenue for future investigation in terms of development through derivatization of the bidentate naphthalimide ligand and optimization of the solvent system to better harness the quenching capabilities of BIH.

Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 D. Archer, J. Geophys. Res., 2005, 110, C09S05.
- 2 E. T. Sundquist, Quat. Sci. Rev., 1991, 10, 283-296.
- 3 G. Shaffer, S. M. Olsen and J. O. P. Pedersen, *Nat. Geosci.*, 2009, 2, 105–109.
- 4 A. J. Morris, G. J. Meyer and E. Fujita, *Acc. Chem. Res.*, 2009, 42, 1983–1994.

- 5 H. Takeda and O. Ishitani, *Coord. Chem. Rev.*, 2010, 254, 346–354.
- 6 L. Q. Qiu, K. H. Chen, Z. W. Yang and L. N. He, *Green Chem.*, 2020, 22, 8614–8622.
- 7 D. Mellmann, P. Sponholz, H. Junge and M. Beller, *Chem. Soc. Rev.*, 2016, 45, 3954–3988.
- 8 J. Eppinger and K. W. Huang, *ACS Energy Lett.*, 2017, 2, 188–195.
- 9 L. Wang, S. Nitopi, A. B. Wong, J. L. Snider, A. C. Nielander, C. G. Morales-Guio, M. Orazov, D. C. Higgins, C. Hahn and T. F. Jaramillo, *Nat. Catal.*, 2019, 2(8), 702–708.
- 10 D. R. Case, A. Spear, A. F. Henwood, M. Nanao, S. Dampf, T. M. Korter, T. Gunnlaugsson, J. Zubieta and R. P. Doyle, *Dalton Trans.*, 2021, 50, 3479–3486.
- 11 Y. Tamaki, K. Koike, T. Morimoto and O. Ishitani, *J. Catal.*, 2013, **304**, 22–28.
- 12 P. Lang, M. Pfrunder, G. Quach, B. Braun-Cula, E. G. Moore and M. Schwalbe, *Chem.-Eur. J.*, 2019, **25**, 4509–4519.
- 13 N. Lazarova, S. James, J. Babich and J. Zubieta, *Inorg. Chem. Commun.*, 2004, 7, 1023–1026.
- 14 I. S. H. Lee, E. H. Jeoung and M. M. Kreevoy, *J. Am. Chem. Soc.*, 1997, **119**, 2722–2728.
- 15 A. Rosas-Hernández, C. Steinlechner, H. Junge and M. Beller, *Green Chem.*, 2017, 19, 2356–2360.
- 16 J. Hawecker, J. Lehn and R. Ziessel, *Helv. Chim. Acta*, 1986, 69, 1990.
- 17 J. Hawecker, J.-M. Lehn and R. Ziessel, *Chem. Commun.*, 1983, 536–538.
- 18 Y. Tamaki, K. Koike, T. Morimoto, Y. Yamazaki and O. Ishitani, *Inorg. Chem.*, 2013, **52**, 11902–11909.
- 19 L. Rotundo, E. Azzi, A. Deagostino, C. Garino, L. Nencini, E. Priola, P. Quagliotto, R. Rocca, R. Gobetto and C. Nervi, Front. Chem., 2019, 7, 417.
- 20 L. Rotundo, D. C. Grills, R. Gobetto, E. Priola, C. Nervi, D. E. Polyansky and E. Fujita, *ChemPhotoChem*, 2021, 5, 526–537.
- 21 N. D. Loewen, T. V. Neelakantan and L. A. Berben, *Acc. Chem. Res.*, 2017, 50, 2362–2370.
- 22 D. B. Cluff, A. Arnold, J. C. Fettinger and L. A. Berben, *Organometallics*, 2018, **38**, 1230–1235.
- 23 D. W. Cunningham and J. Y. Yang, Chem. Commun., 2020, 56, 12965–12968.
- 24 E. Hasegawa, S. Nakamura, K. Oomori, T. Tanaka, H. Iwamoto and K. Wakamatsu, *J. Org. Chem.*, 2021, **86**, 47.
- 25 Y. Wu, Z. Jiang, Z. Lin, Y. Liang and H. Wang, *Nat. Sustain.*, 2021, 4, 725–730.