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DBU mediated one-pot synthesis of triazolo triazines *via* Dimroth type rearrangement†

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Herein we report an efficient one-pot synthesis of [1,2,4]triazolo[1,5-*a*][1,3,5]triazines from commercially available substituted aryl/heteroaryl aldehydes and substituted 2-hydrazinyl-1,3,5-triazines *via* *N*-bromosuccinimide (NBS) mediated oxidative C–N bond formation. Isomerisation of [1,2,4]triazolo[4,3-*a*][1,3,5]triazines to [1,2,4]triazolo[1,5-*a*][1,3,5]triazines is driven by 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) affording both isomers with good to excellent yields (70–96%).

Purines are nitrogen-containing heterocycles and are structural motifs in the nucleobases adenine and guanine of DNA as well as RNA. Purine nucleotides (ATP, GTP, cAMP, cGMP, NAD, FAD) also act as co-factors, substrates, or mediators in the functioning of numerous proteins.¹ Therefore, bioisosteres of purines are widely explored and exploited by pharmaceutical chemists in developing new drug entities. Heterocycles containing the 1,3,5-triazine ring act as bioisosteres of purine, which exist as two isomers, namely [1,2,4]triazolo[4,3-*a*][1,3,5]triazine and [1,2,4]triazolo[1,5-*a*][1,3,5]triazine (Fig. 1), which have been extensively studied as adenosine receptor antagonists,^{2,3} as well as for other pharmacological activities^{4,5} (Fig. 2). Literature reports suggest that [1,2,4]triazolo[1,5-*a*][1,3,5]triazine has been exploited extensively in drug discovery as compared to its corresponding isomer. Further, from the literature it is evident that symmetrical disubstituted triazines,^{6–9} especially morpholine^{10–12} substituted, have displayed broad pharmacological activities.

Various synthetic methods have been reported for the C–N bond formation by employing different starting materials^{13–15}

via oxidative cyclisation,¹⁶ high temperature condition^{17–19} and/or metal-catalysed reactions.²⁰ However, the current reported protocols were environmentally unfriendly as they suffered from drawbacks such as, multistep, and tedious procedures, use of carcinogenic solvents, high-temperature, expensive and toxic metal-catalysts, and other hazardous reagents.

Furthermore, there is very less reported research on these heterocycles, and that can be because of the unavailability of the efficient and cheaper methods. Thus, there is a need to develop new versatile synthetic method for the synthesis of disubstituted triazolo triazine heterocycles of pharmacological interest.

In 1970, for the first time Kobe²¹ *et al.* (scheme a) reported the synthesis of [1,2,4]triazolo[4,3-*a*][1,3,5]triazine utilizing lead tetraacetate in benzene under reflux conditions (Fig. 3). However, no further Isomerization was carried out and resulted low to moderate yield. Deshpande²² *et al.* reported (scheme b) the reaction of 2-hydrazinyl-1,3,5-triazine with various substituted benzoic acids. The product formed was treated with

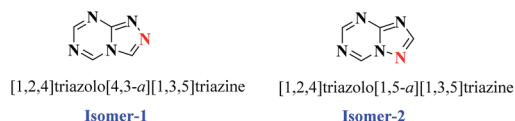


Fig. 1 Structure of isomers of triazolo-triazine.

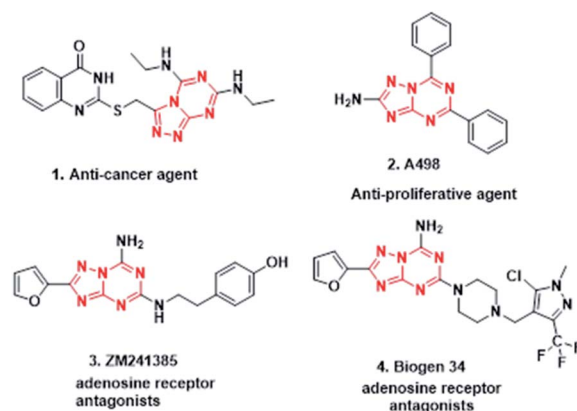


Fig. 2 Biologically active molecules of [1,2,4]triazolo[4,3-*a*][1,3,5]triazine (1) and [1,2,4]triazolo[1,5-*a*][1,3,5]triazine (2, 3, 4).

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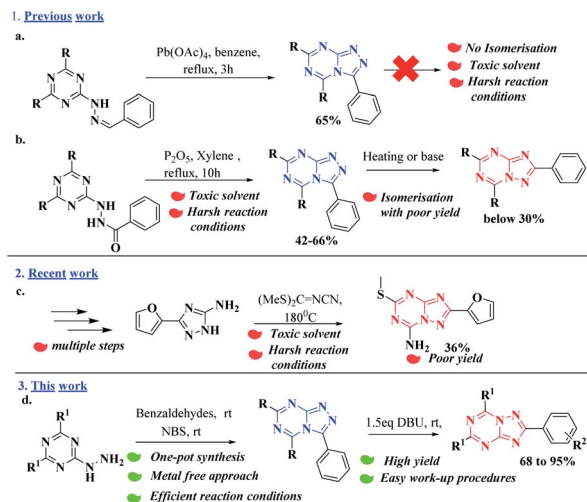


Fig. 3 Different approaches for the synthesis of triazolo triazines.

P_2O_5 , refluxed in xylene for **10h** to yield [1,2,4]triazolo[4,3-*a*] [1,3,5]triazine. Further, Isomerization of the resulting product was carried out in 2% methanolic-NaOH solution resulting in poor yields. Recently, Stefano²³ *et al.* (scheme c) reported, a multistep protocol by reacting the intermediate with bis-(methyl-sulfanyl)methylenecyanamide at 180 °C under N_2 for **3h** resulting in low yield due to the formation of several side products. In addition no isomerization studies were carried out, and only the [1,2,4]triazolo[1,5-*a*][1,3,5]triazine analogs were reported.

Herein, we report an economical one-pot synthesis of [1,2,4] triazolo[1,5-*a*][1,3,5]triazine analogs *via* Dimroth type rearrangement of [1,2,4]triazolo[4,3-*a*][1,3,5]triazine derivatives. This one pot novel methodology was carried out by reacting readily available, inexpensive starting materials such as substituted aryl/heteroaryl benzaldehydes and substituted 2-hydrazinyl-1,3,5-triazine in methanol (a mild solvent)²⁴ at room temperature giving excellent yields of the desired product. For cyclization reaction, an eco-friendly reagent NBS²⁵ was used and the resulting product was treated with DBU to yield its corresponding desired isomer. To the best of our knowledge this is the first report for the greener synthesis of symmetric disubstituted triazolotriazine heterocycles *via* Dimroth type rearrangement. We believe that this simple, yet novel methodology could be further exploited by the researchers in pharmaceutical industries and academics settings in drug discovery.

Formation of the Schiff base was initiated reacting 4,4'-(6-hydrazinyl-1,3,5-triazine-2,4-diyl)dimorpholine **1a** and unsubstituted benzaldehyde as shown in Table 1. The reaction (entry 1) was carried out using *N*-bromosuccinimide (NBS) as an oxidant followed by the addition of 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) in ethanol at room temperature for **16h**, resulting in desired product **2a** with 68% yield. The product appeared as a clean blue spot on the TLC, when visualized under UV (254 nm). Interestingly, the formation of isomer-1 was very straightforward; hence, optimization was achieved for its conversion into isomer-2. Solvent optimization (entry 1–6) was

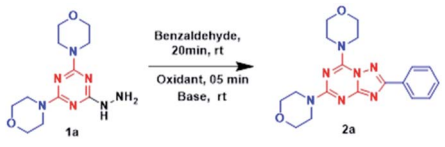
performed using different solvents, starting with DMF or DCM (entry 2, 3) which did not yield expected results. Yield improvement was recorded by the use of MeOH as a solvent, although *i*-PrOH (entry 4, 5) resulted in a trace amount of the products. The same reaction was performed in water (entry 6), but it did not give valued results. Next, base optimization (entry 7–11) was carried out using MeOH as the solvent. The first reaction (entry 7) was performed without base, which showed no further isomerization. Inorganic bases such as 2% NaOH and K_2CO_3 (entry 8,9) were also used but displayed no significant results. Triethylamine and DABCO (1,4-diazabicyclo[2.2.2]octane) were unable to give the desirable product **2a**.

Meanwhile, equivalents of DBU were adjusted (entry 12, 13) to attain the highest yield %. Reaction with 1.5 eq. showed drastic improvement in yields in 2 h whereas 2.0 eq. resulted in good yields with less reaction time. Considering the yield factor (entry 12), the oxidant optimization was achieved (entry 14–17). Using *N*-chlorosuccinimide (NCS) and *N*-iodosuccinimide (NIS) (entry 14, 15) offered 56% and 61% yield respectively. When the reaction was performed using phenyliodine(III) diacetate (PIDA) and KI/I_2 (entry 16, 17) it provided **2a** in 63% and 53% yields, respectively. Finally, methanol–water and ethanol–water systems in 3 : 1 were used, which gave approximately 70% yield, concluding that entry 12 gives the best result, demanding alcoholic solvents especially methanol as a key factor for Dimroth type rearrangement. Additionally, it was supported by the observation that reaction goes well in methanol, a little worse in ethanol (Table 1), and practically does not go in isopropanol, as the bulky *i*-PrO[−] cannot attack the sterically hindered C5 atom in isomer-1 from the side of the phenyl ring. It was also observed that isomerisation did not proceed in presence of acid (HCl or AcOH) at room temperature in methanol solvent. Therefore, MeO[−] as well as DBU were responsible for the rearrangement.

The detailed plausible reaction mechanism also supports the product formation and is described in Fig. 4. The reaction involves a Schiff base formation (II) by **1a** and benzaldehydes, the addition of NBS results in oxidative cyclization reaction to produce isomer 1. The addition of DBU in isomer-1 initiates a famous Dimroth type rearrangement, protonation of III results in ring-opening with the formation of unstable intermediate V. It undergoes tautomerism by 1,3 proton shift, bond rotation and proton abstraction by methoxide ion facilitating the intramolecular cyclization to afford the isomer 2. This reaction mechanism is supported by the formation of the Schiff base, isomer-1, and its conversion to isomer-2, which were easily monitored by TLC, isolated, and characterized (ESI[†]). In addition to that, a single crystal of compound **2e** was obtained, which further supports the Dimroth type of rearrangement (ESI[†]).

Having in hand the optimized conditions, the substrate scope was further explored by using different aldehydes (Fig. 5). The reaction was carried out using **1a** with benzaldehydes having electron-donating groups (**2a–f**) followed by NBS and DBU additions. The reaction was allowed to stir at room temperature till reaction completion, as monitored by TLC, which would approximately take 2 h. The reaction could



Table 1 Optimisation of the reaction condition^a


Entry no.	Solvent	Oxidant	Base	Base equiv.	Time (hour)	Yield%
1	EtOH	NBS	DBU	1.0	16	68
2	DCM	NBS	DBU	1.0	16	0
3	DMF	NBS	DBU	1.0	16	0
4	MeOH	NBS	DBU	1.0	16	72
5	i-PrOH	NBS	DBU	1.0	16	Trace
6	H ₂ O	NBS	DBU	1.0	16	0
7	MeOH	NBS	—	—	16	0
8	MeOH	NBS	2%NaOH	1.0	16	Trace
9	MeOH	NBS	K ₂ CO ₃	1.0	16	0
10	MeOH	NBS	TEA	1.0	16	0
11	MeOH	NBS	DABCO	1.0	16	0
12	MeOH	NBS	DBU	1.5	2	85
13	MeOH	NBS	DBU	2.0	1.5	80
14	MeOH	NCS	DBU	1.5	2	56
15	MeOH	NIS	DBU	1.5	2	61
16	MeOH	IBD	DBU	1.5	2	63
17	MeOH	KI/I ₂	DBU	1.5	2	53
18	MeOH/H ₂ O	NBS	DBU	1.5	2	70

^a Conditions: **1a** (1 mmol), benzaldehyde (1 mmol), solvent, rt, 20 min, then oxidant (1 mmol), 5 min, rt, then DBU, stir at rt till completion of the reaction.

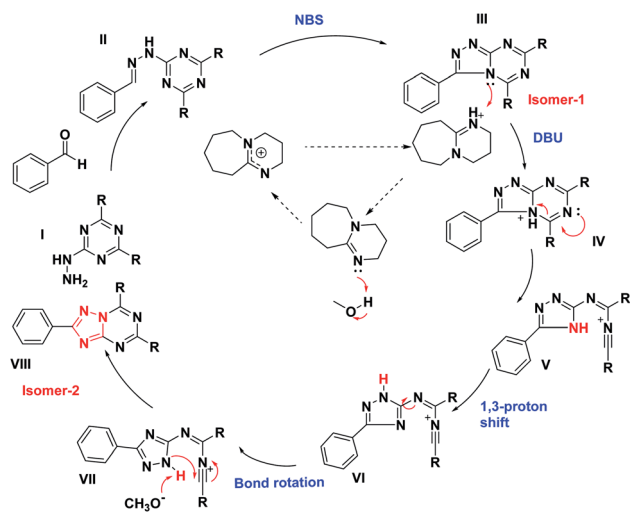


Fig. 4 Plausible mechanism of the reaction pathway.

smoothly give final products in excellent reaction yields (79 to 96%). Electron withdrawing groups such as chlorobenzaldehydes (**2g-i**) despite the position of substitution gave excellent yields (93–95%). With the 4-bromo and different fluoro-benzaldehydes, the reaction resulted in **2j** (89%) and **2k-m** (84–90%) with very good yields. Also, reaction afforded 70% and 77% yields with heterocyclic aldehydes such as 3-

pyridinecarboxyaldehyde (**2n**) and thiophene-2-carbaldehyde (**2o**), respectively.

The gram scale reaction was performed to further validate this synthetic procedure. The reaction was carried out using **1a** (1 g) with benzaldehyde (0.38 g) and stirred for 30 min. The NBS (1.26 g, 1 eq.) was added slowly and stirred till new spot appeared on TLC followed by the slow addition of DBU (1.5 eq.) resulted in isomerisation to give **2a** in good yields (1.1 g, 84%) (ESI⁺).

Encouraged from the potency of the reaction to generate in good to excellent yields [1,2,4]triazolo[1,5-*a*][1,3,5]triazine analogues, we shifted our investigation to isolate [1,2,4]triazolo[4,3-*a*][1,3,5]triazine derivatives (isomer-1) as described in Fig. 6. Different aldehydes were reacted with hydrazinyl-1,3,5-triazine analogs (**1a**, **3a**, and **4a**) and followed by the addition of NBS, stirred at room temperature till the completion of reaction as monitored by TLC. Compound **1a** on reaction with 4-bromobenzaldehyde and NBS resulted in **1b** with 90% yield. Similarly, **3a** with 4-bromobenzaldehydes and thiophene-2-carbaldehyde afforded **3b** and **3c** with 84% and 83% yields, respectively. The same reaction was carried out with **4a** and 4-bromobenzaldehyde bearing an electron-withdrawing group gave **3d** with 87% yield. Further, **4a** with aldehydes bearing electron-donating groups resulted **4c** and **4d**, with 88% and 84% yields, respectively (Fig. 7).

With these promising results, we explored the scope of aldehydes and 2-hydrazinyl-1,3,5-triazine analogs for the synthesis of [1,2,4]triazolo[1,5-*a*][1,3,5]triazine derivatives.



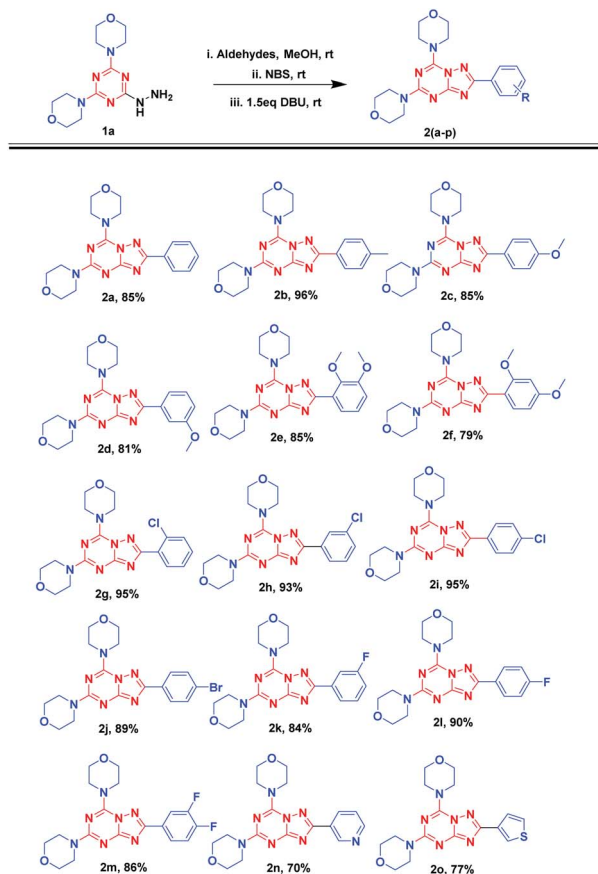


Fig. 5 Substrate scope for different aldehydes.

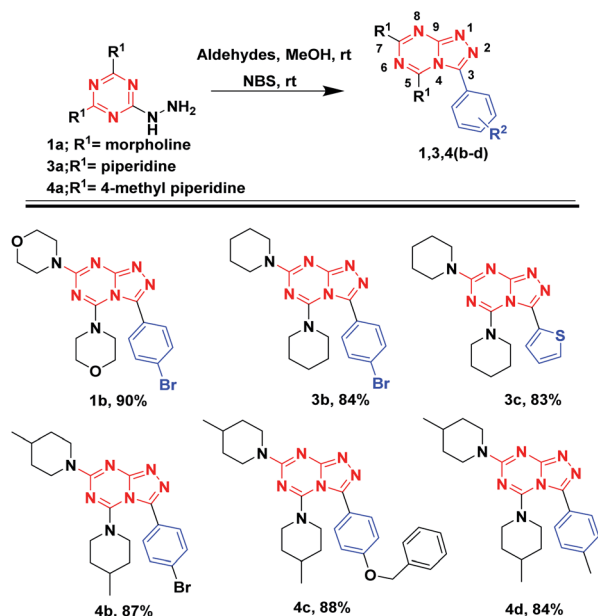


Fig. 6 Substrate scope for [1,2,4]triazolo[4,3-a][1,3,5]triazine analogues.

Compound **3a**, when reacted with different aldehydes and employing the optimized protocols gave **5a-e** with excellent yields (86–92%) in 2–4 h. Similarly compounds **6a**, **6b** and **6c**,

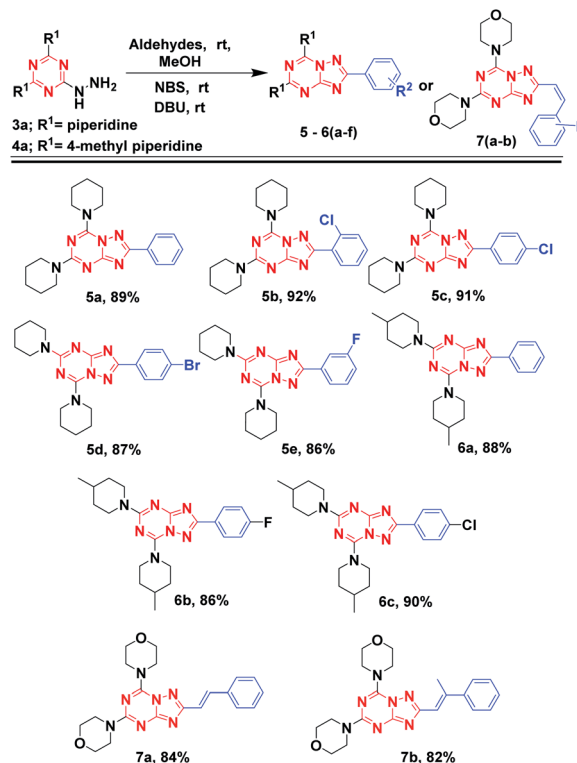


Fig. 7 Substrate scope for different disubstituted triazinyl-hydrazine and aldehydes.

synthesized from **4a** also afforded excellent yields 88%, 86% and 90% respectively. Further, substituted cinnamaldehydes were also explored and reacted with **1a** followed by the addition of NBS and DBU resulted in **7a** and **7b** in appreciable yields of 84% and 82%, respectively.

Finally, we investigated both the isomers (**3b** and **5d**) spectroscopically, to elucidate the change in proton chemical shifts during rearrangement. Interestingly, it was observed that the proton chemical shift for **5d** at the 7th position (Fig. 6) was not affected. However, 5th position of **5d** suffers a downfield shift due to change in its electronic environment as compared to its corresponding isomer **3b**. In general, the proton chemical shifts for the substitution at 5th position and for the aromatic region was found to be more for isomer-2 as compared to isomer-1.

Conclusions

We have developed a simple yet novel, one-pot, facile, rapid, efficient, and greener approach for the synthesis of pharmacologically important triazolo triazines. Economical and commercially available benzaldehydes were reacted with symmetrical disubstituted triazinyl-hydrazines to obtain a Schiff base followed by *in situ* reaction initiated by NBS to form isomer-1 [1,2,4]triazolo[4,3-a][1,3,5]triazine analogues, further addition of DBU yielded isomer-2 [1,2,4]triazolo[1,5-a][1,3,5]triazine derivatives via Dimroth type of rearrangement. Operational simplicity, short reaction time and high product yield make this method especially striking for library generation in drug discovery process.



Thus, we believe that these heterocycles could lead to hit in medicinal chemistry, if explored and exploited on different biological targets. Further applications of this methodology in developing drug like candidates is underway.

Author contributions

All authors contributed equally in the manuscript.

Conflicts of interest

The authors declare no conflict of interest.

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