





Cite this: *Nanoscale*, 2022, **14**, 12668

¹H NMR global diatropicity in copper hydride complexes†

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Understanding the magnetic response of electrons in nanoclusters is essential to interpret their NMR spectra thereby providing guidelines for their synthesis towards various target applications. Here, we consider two copper hydride clusters that have applications in hydrogen storage and release under standard temperature and pressure. Through Born–Oppenheimer molecular dynamics simulations, we study dynamics effects and their contributions to the NMR peaks. Finally, we examine the electrons' magnetic response to an applied external magnetic field using the gauge-including magnetically induced currents theory. Local diatropic currents are generated in both clusters but an interesting global diatropic current also appears. This diatropic current has contributions from three μ_3 -H hydrides and six Cu atoms that form a chain together with three S atoms from the closest ligands resulting in a higher shielding of these hydrides' ¹H NMR response. This explains the unusual upfield chemical shift compared to the common downfield shift in similarly coordinated hydrides both observed in previous experimental reports.

Received 3rd May 2022,
Accepted 3rd August 2022

DOI: 10.1039/d2nr02415b

rsc.li/nanoscale

1. Introduction

Since the discovery of Stryker's reagent, $[\text{HCu-PPh}_3]_6$,¹ now widely employed to regioselectively hydrogenate $\alpha\beta$ -unsaturated carbonyl compounds, several copper hydrides complexes have been synthesized.^{2–28} The broad variety of applications for these complexes include^{23,24,29–31} hydrogen storage,^{5,32–36} catalysis,^{28,32,37–45} energy storage,^{38,32} and energy conversion,^{32,46} to give but a few examples. Within this family of compounds, several coordination modes have been observed for Cu–H hydrides such as adopting a bridge position, face-capping or interstitial positions to help to stabilize the cluster.^{24,47} Determining the hydrides' positions has been a fundamental part of the structural characterization necessary to tailor the potential targeted applications of these complexes. The standard X-ray diffraction technique is used to

determine the overall structure involving the heavy atoms but it fails to locate hydrides and hydrogen atoms in general due to the low electron density around the nuclei. More sophisticated techniques such as single crystal neutron diffraction are therefore often used to correctly assign the coordination modes for these challenging copper hydrides.^{8,47} Recently, due to the lack of neutron diffraction evidence, deep learning approaches have been employed to determine the location of hydrides.^{48,49} As a result, experimental methods that can be performed in solution to help identify the chemical environment around the hydrides, such as nuclear magnetic resonance (NMR), become fundamental in determining the hydrides' positions.

Indeed, NMR spectroscopy has been widely used to non-destructively characterize the reaction state, dynamic phenomena and the chemical environment around the constituent atoms of molecules, macromolecules, nanoparticles, proteins, etc.⁵⁰ Furthermore, calculations of magnetic properties and their response towards perturbations from an applied external magnetic field have become routine.⁵¹ An applied external magnetic field, **B**, generates a response from the electrons in the system studied which is observed as a current composed of both diamagnetic and paramagnetic contributions in a plane perpendicular to **B**. This induced current flows in two directions relative to the external magnetic field: in a classical direction, that is clockwise or diatropic, and in a nonclassical direction, that is anticlockwise or paratropic (with directions deter-

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†Electronic supplementary information (ESI) available: Calculated IR spectra, molecular graphs, BOMD bond lengths analyses, BOMD ¹H NMR chemical shifts. See DOI: <https://doi.org/10.1039/d2nr02415b>



mined by the right-hand screw rule). The strength of this current can then be related to the shielding/deshielding of the nuclei due to the magnetic field generated by these currents.

In the present work, we elucidate the nature of an unusual chemical shift observed in $\mu_n\text{-H}$ ($n = 3\text{--}5$) coordinated hydrides in $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ ³³ and $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_{12}$ ⁵² complexes. The $\mu_3\text{-H}$ is found capping the outer part of the clusters with a chemical shift of -0.99 (CDCl_3 , room T.), -0.55 ppm (Tol- d_8 , room T.) for $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ and -0.17 (CDCl_3 , room T.) -0.28 (CDCl_3 , -60 C) for $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_{12}$ compared with similar $\mu_3\text{-H}$ positions with chemical shifts of 1.3 ,¹⁴ 2.10 ,⁴¹ 2.74 ,⁴³ 3.51 ,⁴ and 3.58 ¹⁰ ppm. We perform asystematic analysis of the chemical bonding in these clusters, ^1H NMR chemical shifts, dynamic effects on the Cu- $\mu_3\text{-H}$ distances, and thermal contributions to the NMR peaks. Finally, we use our recent implementation of gauge-including magnetically induced currents (GIMIC)^{53,54} to analyse the diatropic and paratropic currents generated by the external magnetic field in the NMR probe. We observe a global diatropic current around a chain of six Cu atoms that form (111)-like facets but also a highly localized diatropic current around the $\mu_3\text{-H}$ hydrides shielding for the nuclei and thus causing an upfield negative shift in the NMR spectrum.

2. Results and discussion

The systems studied in this work are based on the structures by Liu and coworkers.^{33,52} The $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ (1, Fig. 1a) and $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_{12}$ (2, Fig. 1b) clusters were modeled from crystallographic data of $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{P}(\text{O}^i\text{Pr})_2)_9$ and $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{P}(\text{O}^i\text{Pr})_2)_{12}$ where the ligands were simplified to improve computational efficiency. Cluster 1 can be geometrically described as an elongated triangular orthobicupola of 18 Cu atoms enclosing 2 central vertical Cu atoms. Cluster 2 can be described as a hexacapped rhombohedron of 14 Cu atoms enclosed by two twisted triangular cupolas of 9 Cu atoms similar to cluster 1.

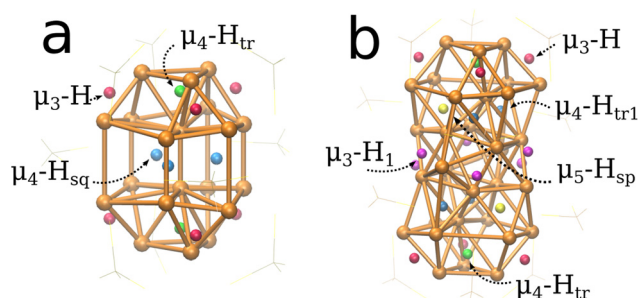


Fig. 1 (a) For the $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ cluster, the eleven hydrides are divided into three groups based on their positions; six $\mu_3\text{-H}$ hydrides, two $\mu_4\text{-H}_{\text{tr}}$ hydrides and three $\mu_4\text{-H}_{\text{sq}}$ hydrides. (b) For the $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_{12}$ cluster, the twenty hydrides are divided into five groups; six $\mu_3\text{-H}$ hydrides, two $\mu_4\text{-H}_{\text{tr}}$ hydrides, four $\mu_4\text{-H}_{\text{tr1}}$ hydrides, six $\mu_3\text{-H}_1$ hydrides, and two $\mu_5\text{-H}_{\text{sp}}$ hydrides. The S_2PH_2 ligands are omitted for clarity.

Both model systems were optimized at the B88-P86/def2-TZVP level of theory as discussed in the computational details. To verify the local minimum, a vibrational analysis was performed for each system and the simulated infra-red absorption spectra for clusters 1 and 2 are shown in Fig. SI-1.[†] The normal modes for the hydrides in cluster 1 were observed in a narrow range between $905\text{--}968\text{ cm}^{-1}$ due to low coordination of the hydrides with the Cu atoms. Cluster 2 absorbs in a broader range of frequencies with vibrational normal modes between $831\text{--}1283\text{ cm}^{-1}$. The wider range is due to the wider variety of hydride coordination present in this cluster. The topological analysis of the density⁵⁵ gives insight into the chemical bonding in the two copper hydrides (see Fig. SI-2[†]). Details of these calculations are given in the Computational section. The calculated bond paths shown in the density molecular graphs confirm the hydrides' formal assigned coordination. This analysis also reveals that there are almost no copper-copper bond paths and, thus, no copper-copper bonding in these clusters. A remarkable exception is the central copper dimer bond in 1 (see Fig. SI-2[†]). To gain more qualitative insight into the chemical bonding in these systems, we calculated the density, $\rho(\mathbf{r}_{\text{CP}})$, and Laplacian of the density, $\nabla^2\rho(\mathbf{r}_{\text{CP}})$, at the Cu-H and Cu-Cu bond critical points, \mathbf{r}_{CP} . For all of these bonds, $\nabla^2\rho(\mathbf{r}_{\text{CP}}) > 0$ holds, which excludes covalent bonding.⁵⁶ The corresponding densities at these critical points are between 0.05 and 0.1 a.u., indicating weak metallic or ionic bonding. Of particular interest is the fact that in cluster 1, the $\rho(\mathbf{r}_{\text{CP}})$ for the central copper dimer is with 0.51 a.u. smaller than the $\rho(\mathbf{r}_{\text{CP}})$ for all Cu-H bonds. Thus, the weak metallic bonding between the copper atoms is the reason for the missing Cu-Cu bond paths. The copper hydride bonds are therefore essential for the cluster structures observed. This is experimentally confirmed by the fact that these clusters decompose under thermal hydrogen release.^{33,52}

^1H NMR chemical shifts

To identify the different types of hydride positions for clusters 1 and 2, the magnetic shielding tensor was computed as described in the Computational details section. The proton chemical shifts for cluster 1 are summarized in Table 1.

In agreement with previous experimental studies,³³ we have identified three different types of hydrides for cluster 1 in a 3:2:6 ratio. Six $\mu_3\text{-H}$ hydrides on triangular facets were obtained with an unusual upfield shift of -1.00 ppm ($-1.52/\text{PBE}$) with a deviation of $0.01\text{--}0.45$ ppm ($0.53\text{--}0.97$ ppm/PBE)

Table 1 Experimental and calculated ^1H NMR chemical shifts in ppm for the $\text{Cu}_{20}\text{H}_{11}$ (1) cluster at B88-P86/def2-TZVP and PBE/aug-cc-pVDZ level of theory

Position	CDCl_3 ^a	Tol- d_8 ^a	B88-P86/def2-TZVP	PBE/aug-cc-pVDZ
$\mu_3\text{-H}$	-0.99	-0.55	-1.00	-1.52
$\mu_4\text{-H}_{\text{tr}}$	1.46	1.88	1.98	1.83
$\mu_4\text{-H}_{\text{sq}}$	2.80	3.32	4.25	4.24

^a Ref. 33.



with respect to experimental values.³³ These deviations depend on the functional and the basis set as well as the solvent used in experiments. Two hydrides $\mu_4\text{-H}_{\text{tr}}$, each in a tetrahedral position inside the triangular cupola at the top and bottom, were identified with a downfield shift of 1.98 ppm (1.83 ppm/PBE) within 0.1–0.52 ppm (0.05–0.37 ppm/PBE) from experimental values reported.³³ Three additional hydrides are shifted downfield with a $\mu_4\text{-H}_{\text{sq}}$ position in the center of a slightly distorted square polygon formed by the two central Cu atoms and the vertices of the Cu shell. These hydrides exhibit a calculated downfield shift of 4.25 ppm (4.24 ppm/PBE) differing by 0.93–1.45 ppm (0.92–1.44 ppm/PBE) from experimental measurements. They are also more sensitive to the solvent where Tol- d_8 produces a larger downfield shift than CDCl_3 . These shifts also show the largest difference between the experimental and DFT values. The larger difference is due to the calculations being performed in gas phase. We expect that further studies including the solvent effect, either explicitly or through an implicit solvent model, would improve the agreement between the computed and experimental ^1H proton shifts but this is beyond the scope of this work.

For the copper hydride cluster 2, we identified the position of the twenty hydrides divided in five groups summarized in Table 2 in agreement with previous experimental reports.⁵² As in cluster 1, there are six $\mu_3\text{-H}$ hydrides on triangular (111)-like facets, located on the triangular cupola at the top and bottom of the cluster, with similar unusual upfield chemical shifts of 0.20, 0.34 ppm (−0.15, −0.3 ppm/PBE) in a 4 : 2 ratio with a deviation of 0.37–0.62 ppm (0.02–0.13 ppm/PBE) from experimental values. To compare two calculated chemical shifts with one experimental value, the deviation is given as the minimum and maximum difference between the two data sets. Both B88-P86 and PBE functionals showed upfield shifting, but only the PBE results were shifted to the negative part of the spectrum as observed experimentally. It is worth mentioning that both chemical shifts were measured shifted upfield independently of the temperature in CDCl_3 solvent. We want to highlight this unusual behaviour and bring later the discussion on this $\mu_3\text{-H}$ hydrides on 1 and 2 clusters.

A second set of six $\mu_3\text{-H}_1$ hydrides was observed on triangular facets around the central rhombohedron of 14 Cu atoms. At room temperature, the experimental NMR showed only one

peak at 0.92 ppm but when the temperature was decreased to −60 °C, two peaks at 1.32 and 0.13 ppm were reported. We also observed two different values for these $\mu_3\text{-H}_1$ positions, 2.27 and 1.44 ppm (2.12, 1.24 ppm/PBE) in a 4 : 2 ratio which deviate by 0.95–1.31 ppm (0.80–1.11 ppm/PBE) from experimental values at −60 °C and 0.52–1.35 ppm (0.32, 1.20 ppm/PBE) at room temperature. The effects of temperature on the chemical shifts for these hydrides is addressed later in the discussion. Each $\mu_3\text{-H}_1$ hydride appears symmetric on alternating triangular facets around the rhombohedron pointing towards the top or bottom triangular cupola but always outside the cluster.

The tetrahedrally coordinated hydrides are divided into two groups, $\mu_4\text{-H}_{\text{tr}}$ and $\mu_4\text{-H}_{\text{tr1}}$, based on their positions inside the cluster. The former are located in positions similar to those observed in cluster 1 with Cu tetrahedra in the triangular cupolas at the top and bottom of the cluster each containing one $\mu_4\text{-H}_{\text{tr}}$ hydride. The latter are located in four tetrahedra formed by the triangular cupolas and the rhombohedron and each tetrahedron contains one $\mu_4\text{-H}_{\text{tr1}}$ hydride. The calculated chemical shifts for $\mu_4\text{-H}_{\text{tr}}$ hydrides are 5.27 ppm (5.11 ppm/PBE) with a deviation of 2.15–2.18 ppm (1.99–2.02 ppm/PBE) from experimental values, while the calculated chemical shifts for the $\mu_4\text{-H}_{\text{tr1}}$ hydrides are 6.03 ppm (5.95 ppm/PBE) with a deviation of 1.33–2.04 ppm (1.25–1.96/PBE) from the experimental values.

The last two highly coordinated $\mu_5\text{-H}_{\text{sp}}$ hydrides are each located in distorted square pyramids formed by Cu atoms from both triangular cupolas and the rhombohedron. Each hydride may be found near the center of the distorted square base. The $\mu_5\text{-H}_{\text{sp}}$ hydrides have the largest downfield chemical shift at 8.76 ppm (8.68 ppm/PBE) with a deviation of 2.83–4.06 ppm (2.75–3.98 ppm/PBE) with respect to experimental values. A 2 : 4 : 2 : 6 : 6 ratio was observed between the hydrides of cluster 2, in good agreement with that previously reported.⁵²

Structural dynamic effects

To address the dynamic effects on the chemical shifts of the clusters 1 and 2, Born–Oppenheimer Molecular Dynamics (BOMD) were performed for 15 ps at different temperatures as described in the Computational details section. Each system was thermally equilibrated for 10 ps. An additional 5 ps were simulated and the trajectories taken to analyze structural changes. As shown above, the chemical shifts for the $\mu_3\text{-H}$ hydrides in both clusters 1 and 2, were experimentally and computationally observed as shifted upfield to negative values. Experimentally, these shifts are independent of solvent or probe temperature, therefore we examined the bond lengths with respect to the closest Cu atoms.

As the ^1H NMR measurements were done at 20 and 30 °C for cluster 1, we performed the BOMD simulations at 300 K for this cluster. We followed the three characteristic distances, \bar{d}_1 , \bar{d}_2 and \bar{d}_3 , of one of the $\mu_3\text{-H}$ hydrides with respect to each of the Cu atoms to which it is coordinated, as shown in Fig. 2. The 5 ps trajectory analysis showed longer $\bar{d}_1 = 1.80 \text{ \AA}$ ($\sigma_1 =$

Table 2 Experimental and calculated ^1H NMR chemical shifts in ppm for $\text{Cu}_{32}\text{H}_{20}$ (2) cluster at B88-P86/def2TZVP and PBE/aug-cc-pVDZ level of theory

Position	Room T. ^{a,b}	−60 °C ^{a,b}	B88-P86/def2-TZVP	PBE/aug-cc-pVDZ
$\mu_3\text{-H}$	−0.17	−0.28	0.20, 0.34 (4 : 2)	−0.15, −0.3 (4 : 2)
$\mu_3\text{-H}_1$	0.92	1.32, 0.13	2.27, 1.44 (4 : 2)	2.12, 1.24 (4 : 2)
$\mu_4\text{-H}_{\text{tr}}$	3.12	3.09	5.27	5.11
$\mu_4\text{-H}_{\text{tr1}}$	4.70	3.99	6.03	5.95
$\mu_5\text{-H}_{\text{sp}}$	4.70	5.93	8.76	8.68

^a Ref. 52. ^b CDCl_3 solvent.



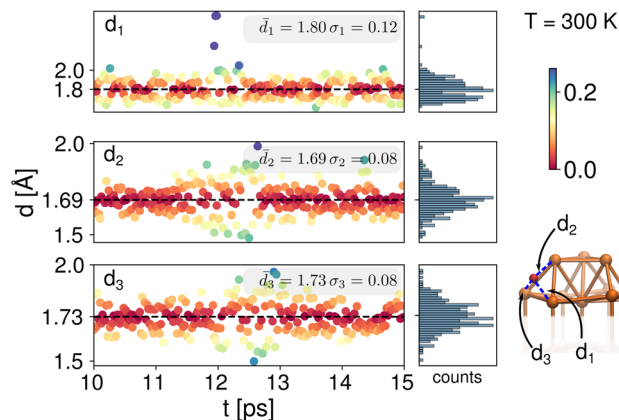


Fig. 2 Bond lengths analysis between a μ_3 -H hydride and its three coordinating neighbor Cu atoms for the BOMD trajectory during 10–15 ps for the cluster 1 at 300 K.

0.12 Å) and shorter $\bar{d}_2 = 1.69$ Å ($\sigma_2 = 0.08$ Å) bond lengths. The third characteristic distance, \bar{d}_3 , has a value of 1.73 Å ($\sigma_1 = 0.08$ Å). As can be seen from the color bar and histograms in Fig. 2, despite thermal fluctuations in the three different bond lengths, the deviation from the mean values is small, approximately 0.1 Å. All of these values are longer compared with previously reported experimental bond lengths³³ for cluster 1 ($\bar{d}_1 = 1.60 \pm 0.04$, $\bar{d}_2 = 1.72 \pm 0.03$ and $\bar{d}_3 = 1.76 \pm 0.04$, Table SI-1†). The experimental data was collected at 100 K and our longer values are expected due to the higher dynamics at 300 K. Our values are also in good agreement with typical experimental bond lengths for μ_3 -H which are 1–7–1.9 Å.^{4,14,28}

For cluster 2, previous experimental ^1H NMR measurements⁵² were done at -60 °C and room temperature. We selected three different temperatures (100, 200, 300 K) to perform the BOMD simulations. The simulations at 300 K showed severe structural distortions due to the elongated shape of the cluster 2 and the small S_2PH_2 model ligands. We therefore chose to focus our analysis on simulations at 100 and 200 K. The trajectory analysis for the μ_3 -H hydride at 100 K (Fig. SI-3†) showed two nearly equivalent bond lengths, $\bar{d}_2 = 1.74$ Å ($\sigma_2 = 0.03$ Å) and $\bar{d}_3 = 1.74$ Å ($\sigma_3 = 0.02$ Å), and a third shorter bond length, $\bar{d}_1 = 1.67$ Å ($\sigma_1 = 0.01$ Å). At this temperature, the thermal fluctuations are almost negligible as shown in the color bar and histograms of Fig. SI-3.† Although the BOMD trajectories at 200 K (Fig. SI-6†) showed broader deviations from the mean values, the average bond lengths are consistent with their respective values at 100 K (Fig. SI-5†). Again, two equal bond lengths, $\bar{d}_2 = 1.73$ Å ($\sigma_2 = 0.1$ Å) and $\bar{d}_3 = 1.73$ Å ($\sigma_3 = 0.09$ Å), and a third shorter bond length, $\bar{d}_1 = 1.69$ Å ($\sigma_1 = 0.08$ Å) were observed. We observe that the two equivalent distances are shorter and the third one longer compared to the corresponding distances at 100 K. For cluster 2, the experimental values reported at 100 K (ref. 52) are $\bar{d}_1 = 1.63 \pm 0.04$, $\bar{d}_2 = 1.72 \pm 0.03$ and $\bar{d}_3 = 1.78 \pm 0.07$ (Table SI-1†). Our values obtained from BOMD at 100 K are thus in good agreement with experiment.

Thermal contributions to ^1H NMR chemical shifts

We computed the shielding tensor during the BOMD trajectories from 10–15 ps and analyzed the respective chemical shifts for one hydride sample of every individual type as shown in Fig. 3, SI-4 and SI-5,† for both clusters 1 and 2 at the respective temperatures.

For cluster 1, at 300 K, the small fluctuations in the bond lengths (0.08–0.12 Å) show a direct correlation between the shortening and lengthening of the three bond distances with the standard deviation in the chemical shifts of 0.96–1.16 ppm over time that can also occur in the experimental NMR setup. The average chemical shielding values are always shifted downfield with the μ_4 -H_{sq} hydride showing the highest downfield shift ($\bar{\delta} = 5.66 \pm 1.16$ ppm) followed by the μ_4 -H_{tr} hydride ($\bar{\delta} = 3.65 \pm 0.96$ ppm) and the μ_3 -H ($\bar{\delta} = -0.54 \pm 1.02$ ppm). These results show that the ordering of the hydride chemical shifts is not altered by temperature effects and μ_3 -H is still observed with an average negative chemical shift.

For cluster 2, the standard deviations in the chemical shifts are smaller at lower temperatures, between 0.26–0.51 ppm (100 K), and larger, between 0.84–1.26 (200 K), at higher temperatures in good correspondence with the fluctuations in the three bond lengths observed in the trajectory analysis for both clusters. While the calculations at 0 K did not reproduce the negative chemical shift value observed for the μ_3 -H hydrides at the B88-P86/def2TZVP level of theory, once the thermal contributions are considered, the ^1H chemical shift is moved further upfield to values of $\bar{\delta} = -0.09 \pm 0.26$ ppm at 100 K and $\bar{\delta} = -0.16 \pm 0.95$ ppm at 200 K.

Interestingly, the chemical shift for the μ_3 -H₁ hydride shows an average upfield chemical shift of $\bar{\delta} = 1.33 \pm 0.29$ ppm at 100 K and a downfield chemical shift of $\bar{\delta} = 2.37 \pm 0.84$ ppm at 200 K compared with the calculated results

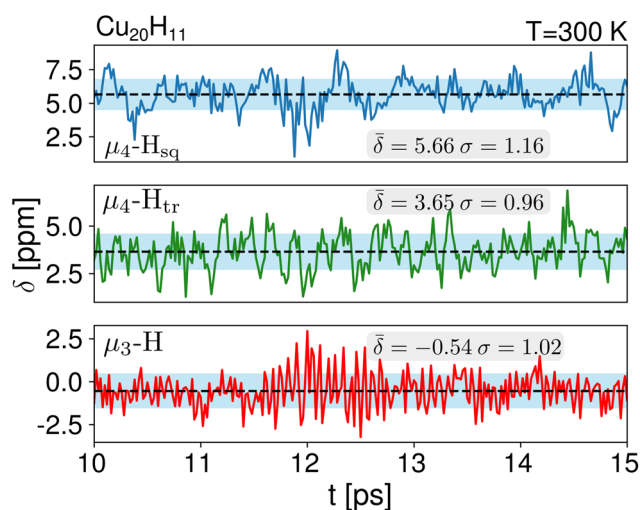


Fig. 3 ^1H NMR dynamics of one sample of each at the μ_3 -H, μ_4 -H_{tr} and μ_4 -H_{sq} hydride types for cluster 1 during 10–15 ps at 300 K. The average shifts are given by the dashed lines while the standard deviations are shown with the blue shading.



at 0 K. As previously discussed, at 0 K, the $\mu_3\text{-H}_1$ hydrides give a signal in a ratio of 4:2 with a significant difference of 0.83 ppm between the peaks. This raises the possibility that thermal contributions average both upfield and downfield shifts during the trajectory.

The highly coordinated hydride positions did not show on average any significant change compared with the 0 K computations after including the thermal contributions, with $\mu_4\text{-H}_{\text{tr}}$ hydride $\bar{\delta} = 5.48 \pm 0.29$ ppm, $\mu_4\text{-H}_{\text{tr1}}$ hydride $\bar{\delta} = 6.42 \pm 0.30$ ppm, $\mu_5\text{-H}_{\text{sp}}$ hydride $\bar{\delta} = 8.67 \pm 0.51$ ppm at 100 K and with $\mu_4\text{-H}_{\text{tr}}$ $\bar{\delta} = 5.10 \pm 1.26$ ppm, $\mu_4\text{-H}_{\text{tr1}}$

$\bar{\delta} = 5.96 \pm 1.03$ ppm, $\mu_5\text{-H}_{\text{sp}}$ $\bar{\delta} = 7.41 \pm 1.07$ ppm at 200 K (Table 3).

Magnetically induced current densities

Understanding the local and global magnetic environment in the clusters is crucial to rationalize the different responses of proton chemical shifts to the external magnetic field imposed in NMR measurements. Therefore, we studied the magnetically induced current density $\mathbf{J}^{(1)}(\mathbf{r})$ obtained by the gauge-including magnetically induced currents (GIMIC).^{53,54,57} Within the computations a magnetic field in the z-direction $\mathbf{B} = (0, 0, 1 \text{ T})$ was imposed for both clusters 1 and 2. The magnetically induced current density at $\mathbf{J}^{(1)}(\mathbf{r}) = 0.01 \text{ nA T}^{-1}$ is shown in Fig. 4 for cluster 1 (a) and cluster 2 (b). It is observed that the diatropic contributions appear on top and bottom of the orthobicupola creating a ring with these six Cu atoms that encompass the $\mu_3\text{-H}$ hydrides both also the S atoms in the ligands, creating a global current density in this region of the clusters. Then to analyze the strength of the current density the projected current density fields are shown in three x-y planes at different z-coordinates (Fig. 5). Each plane contains a distinct group of hydrides. The current density fields are circularly integrated to show the contributions of different currents as a

Table 3 Average ^1H NMR chemical shifts (standard deviation) in ppm from 10 to 15 ps for cluster 1 and 2 at B88-P86/def2TZVP level of theory and different temperatures

Position	1		2	
	300 K	Position	100 K	200 K
$\mu_3\text{-H}$	−0.54 (1.02)	$\mu_3\text{-H}$	−0.09 (0.26)	−0.16 (0.95)
$\mu_4\text{-H}_{\text{tr}}$	3.65 (0.96)	$\mu_3\text{-H}_1$	1.33 (0.29)	2.37 (0.84)
$\mu_4\text{-H}_{\text{sq}}$	5.66 (1.16)	$\mu_4\text{-H}_{\text{tr}}$	5.48 (0.29)	5.10 (1.26)
		$\mu_4\text{-H}_{\text{tr1}}$	6.42 (0.30)	5.96 (1.03)
		$\mu_5\text{-H}_{\text{sp}}$	8.67 (0.51)	7.41 (1.07)

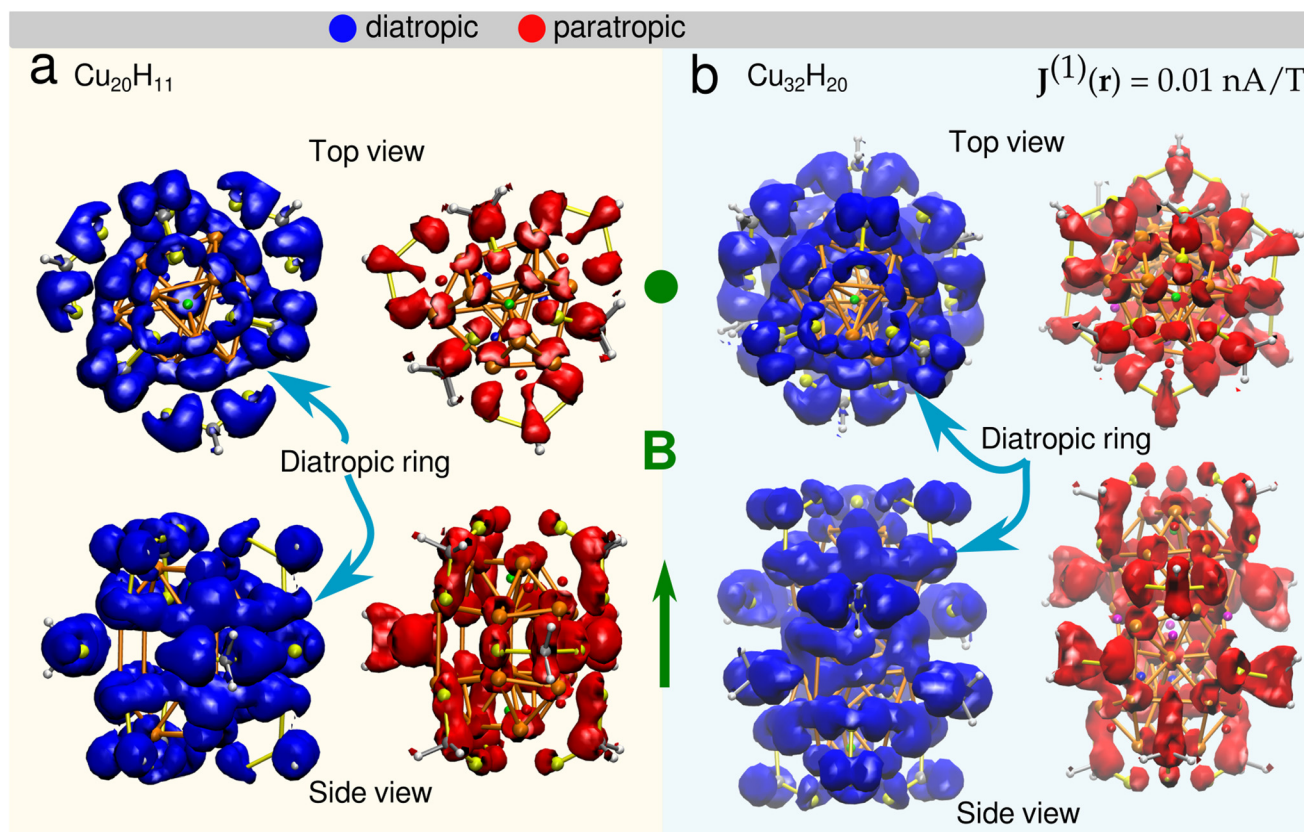


Fig. 4 Magnetically induced current density at $\mathbf{J}^{(1)}(\mathbf{r}) = 0.01 \text{ nA T}^{-1}$ for cluster 1 (a) ($\text{Cu}_{20}\text{H}_{11}$) and cluster 2 (b) ($\text{Cu}_{32}\text{H}_{20}$). Diatropic (blue) and paratropic (red) contributions are shown. A diatropic ring is observed in the top and bottom triangular cupolas for both systems encompassing each one three $\mu_3\text{-H}$ hydrides, six Cu atoms and three S atoms from the closest ligands. The external magnetic field \mathbf{B} (green) is pointing toward the reader (top) and in z-direction (bottom).



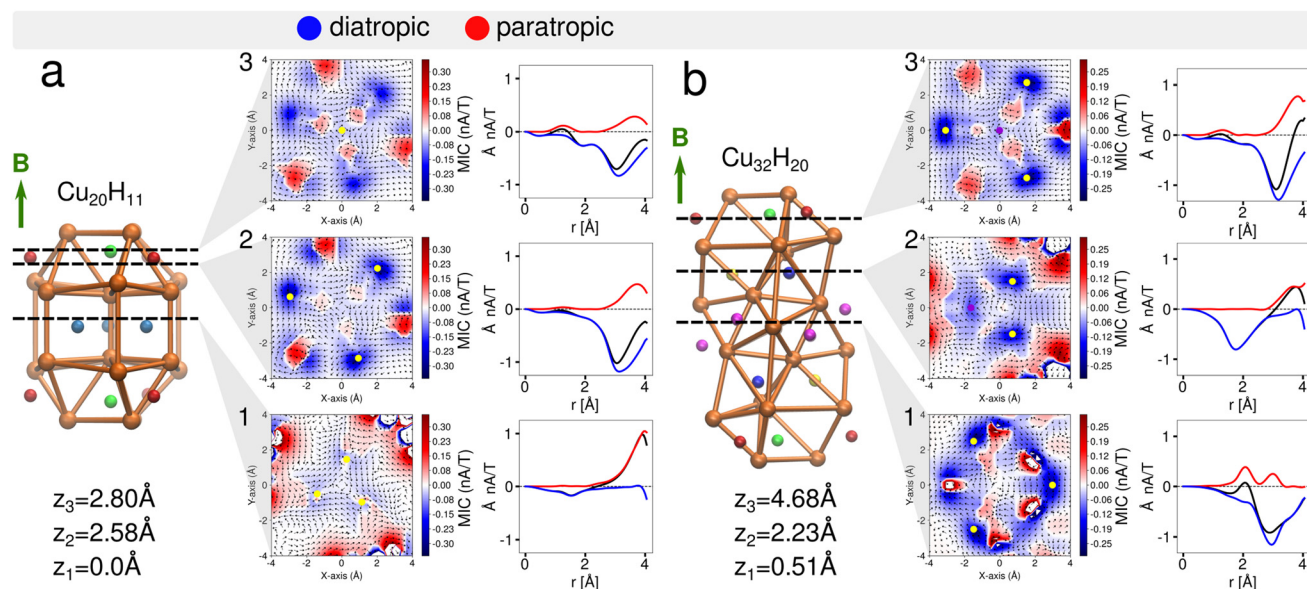


Fig. 5 Vectorial fields of the magnetically induced current density and the circularly integrated current density for the cluster 1 (a) ($\text{Cu}_{20}\text{H}_{11}$) and the cluster 2 (b) ($\text{Cu}_{32}\text{H}_{20}$) at the z -coordinate planes. Diatropic (blue) and paratropic (red) contributions are shown. The total integrated current is shown in black. The external magnetic field B (green is pointing toward the z -direction of the clusters and toward the reader in the planes.)

function of the radius r from the center of the plane located at $(0, 0, z)$. The diatropic, paratropic and total current contributions on the plane are distinguished as summarized in Table 4.

For the cluster 1 (Fig. 5a1), in the plane containing the three $\mu_4\text{-H}_{\text{sq}}$ hydrides ($z_1 = 0.0 \text{ Å}$), they are embedded in global diatropic currents of $-0.18 \text{ Å}^2 \text{ nA T}^{-1}$ ($r = 2.0 \text{ Å}$) and $-0.22 \text{ Å}^2 \text{ nA T}^{-1}$ ($r = 3.0 \text{ Å}$) spread inside the core of the cluster thereby causing the high deshielding and thus the downfield shift in the NMR spectrum.

In the plane containing a $\mu_4\text{-H}_{\text{tr}}$ hydride ($z_3 = 2.8 \text{ Å}$), there are three localized paratropic contributions from the Cu atoms surrounding the hydride (Fig. 5a3). At $r = 2.0 \text{ Å}$ a global diatropic current of $-0.20 \text{ Å}^2 \text{ nA T}^{-1}$ is observed but a small current of $0.07 \text{ Å}^2 \text{ nA T}^{-1}$ from the paratropic counterpart decreases

the total current strength (Table 4). The downfield shift of this $\mu_4\text{-H}_{\text{tr}}$ inside the cluster compared with the $\mu_3\text{-H}$ hydride is therefore expected due to the proton's deshielding.

Looking at the $\mu_3\text{-H}$ hydrides in the $z_1 = 2.58 \text{ Å}$ plane, we observed highly localized diatropic current contributions around each of them (Fig. 5a2) and a global total current of -0.59 ($r = 3.0 \text{ Å}$) and $-1.23 \text{ Å}^2 \text{ nA T}^{-1}$ ($r = 4.0 \text{ Å}$). This difference between the two radii is due to the high contributions of the localized diatropic currents around the hydrides. The localized diatropic areas are easily identified from their spacial representation on the plane containing the $\mu_3\text{-H}$ hydride. The $\mu_3\text{-H}$ hydrides are highly shielded because of the highly localized diatropic currents around them thus giving origin to the negative peak value in the NMR spectrum.

Table 4 Radial integrals of the circularly integrated current strength (Fig. 5) (in $\text{Å}^2 \text{ nA T}^{-1}$) at radius r from the center of an x - y plane located at $(0, 0, z)$ in clusters 1 and 2

1					2				
z	r	Tot	Para	Dia	z	r	Tot	Para	Dia
2.8	1.0	-0.04	0.01	-0.05	4.68	1.0	-0.05	0.01	-0.06
	2.0	-0.13	0.07	-0.20		2.0	-0.11	0.06	-0.18
	3.0	-0.52	0.09	-0.61		3.0	-0.54	0.08	-0.63
	4.0	-0.93	0.32	-1.25		4.0	-0.96	0.60	-1.56
2.58	1.0	-0.04	0.00	-0.05	2.23	1.0	-0.05	0.00	-0.05
	2.0	-0.12	0.02	-0.14		2.0	-0.61	0.00	-0.61
	3.0	-0.59	0.05	-0.63		3.0	-0.97	0.00	-0.97
	4.0	-1.23	0.40	-1.62		4.0	-0.71	0.35	-1.06
0	1.0	-0.05	0.00	-0.05	0.51	1.0	-0.01	0.01	-0.02
	2.0	-0.17	0.01	-0.18		2.0	-0.10	0.10	-0.20
	3.0	-0.15	0.07	-0.22		3.0	-0.58	0.30	-0.88
	4.0	0.42	0.65	-0.24		4.0	-1.22	0.35	-1.56

In cluster 2, we observed a global diatropic current surrounding the central part of the Cu core with highly localized paratropic current contributions (Fig. 5b1). In the plane ($z_1 = 0.51 \text{ \AA}$) containing three of the six $\mu_3\text{-H}_1$ hydrides, there is a local paratropic current effect closer to the two symmetrically equivalent $\mu_3\text{-H}_1$ hydrides compared to the third $\mu_3\text{-H}_1$ hydride. The location of the paratropic current explains the 4 : 2 ratio observed in the calculated NMR at 0 K. Although the total global contributions are very similar (-0.58 at $r = 3.0 \text{ \AA}$ and $-1.22 \text{ \AA}^2 \text{ nA T}^{-1}$ at $r = 4.0 \text{ \AA}$) to those observed for the $\mu_3\text{-H}$ hydrides at $z_2 = 2.58 \text{ \AA}$ in cluster 1, the paratropic currents reduce the total shielding effect around these hydrides. The chemical shifts for these $\mu_3\text{-H}_1$ positions are thus distinguished from the $\mu_3\text{-H}$ at the cluster's extremities.

The $z_2 = 2.23 \text{ \AA}$ plane contains one $\mu_5\text{-H}_{\text{sp}}$ and two $\mu_4\text{-H}_{\text{tr1}}$ hydrides (Fig. 5b2). These hydrides inside the core of the Cu cluster have higher local diatropic contribution for $\mu_4\text{-H}_{\text{tr1}}$ hydrides than for $\mu_5\text{-H}_{\text{sp}}$, with global diatropic currents of $-0.61 \text{ \AA}^2 \text{ nA T}^{-1}$ spread in the plane within the area enclosed by $r = 2.0 \text{ \AA}$ with completely negligible paratropic contributions. Then, the higher downfield shift of $\mu_5\text{-H}_{\text{sp}}$ is understood due to the higher deshielding of these protons inside the cluster compared with $\mu_4\text{-H}_{\text{tr1}}$ hydrides.

The last plane at $z_1 = 4.68 \text{ \AA}$ (Fig. 5b3) contains both $\mu_4\text{-H}_{\text{tr}}$ and $\mu_3\text{-H}$ hydrides. The $\mu_4\text{-H}_{\text{tr}}$ hydride at the center is surrounded by localized paratropic currents causing a similar deshielding effect to that observed in cluster 1 thus producing an equivalent downfield shift in the ^1H NMR peak. The highly localized diatropic currents around the $\mu_3\text{-H}$ hydrides and the global diatropic currents contributions of -0.63 ($r = 3.0 \text{ \AA}$) and $-1.56 \text{ \AA}^2 \text{ nA T}^{-1}$ ($r = 4.0 \text{ \AA}$) are also similar to the corresponding quantities in cluster 1. These $\mu_3\text{-H}$ hydrides are therefore highly shielded as in cluster 1 because of the highly localized diatropic currents around them and thus show similar negative chemical shifts in the ^1H NMR spectrum. The analysis of the magnetic induced currents is supported by the separation of the ^1H chemical shieldings into diamagnetic and paramagnetic contributions. Most of the hydrides in clusters 1 and 2 have negative diamagnetic shifts that are smaller in absolute value than the corresponding positive paramagnetic shifts. For the $\mu_3\text{-H}$ hydrides with overall negative chemical shifts, this relation is reversed due to the strong shielding from their local environment.

3. Conclusions

We studied the magnetic response of two copper hydride nanoclusters, $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ and $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_9$, by means of DFT and BOMD calculations to obtain shielding tensors and gauge-including magnetically induced currents. We validated the hydrides' ^1H NMR chemical shifts and positions by comparing with previous experimental reports. Our 0 K ground state and finite temperature BOMD simulations reproduced the unusual negative chemical shift observed for the $\mu_3\text{-H}$, capping the outer part of (111)-facets in the triangular

cupolas of each cluster. By analyzing the dynamic thermal effects on these ^1H NMR chemical shifts, we established that small fluctuations in the $\mu_3\text{-H-Cu}$ characteristic bond distances have an immediate upfield/downfield shift with respect to the average value. The correct order assignment of the peaks and the $\mu_3\text{-H}$ hydrides' negative shift, however, is recovered over time. Finally, we used our implementation of gauge-including magnetically induced currents to elucidate the nature of the shielding around the $\mu_3\text{-H}$ hydrides. A global diatropic current ring, formed by a chain of six Cu atoms including the $\mu_3\text{-H}$ hydrides but also the sulfur atoms from the closest ligands on the clusters, was observed. Additionally, a highly localized diatropic current in the region closest to the $\mu_3\text{-H}$ hydrides was shown causing a higher shielding for these nuclei and thus their upfield negative shift in the NMR spectrum. We expect that this investigation gives insights in the importance of considering the currents within these challenging nanoparticles when elucidating NMR spectra.

4. Computational details

Auxiliary density functional theory⁵⁸ (ADFT) ground state calculations were performed on the two model clusters, $\text{Cu}_{20}\text{H}_{11}(\text{S}_2\text{PH}_2)_9$ ³³ and $\text{Cu}_{32}\text{H}_{20}(\text{S}_2\text{PH}_2)_{12}$,⁵² employing the B88 exchange⁵⁹ and P86 correlation⁶⁰ functionals. The non-interacting Kohn–Sham ADFT wave function was expanded with the linear combination of Gaussian type orbital (LCGTO) approximation. For H, P and S the all-electron LCGTO def2-TZVP basis sets⁶¹ were used, whereas for Cu the 19 valence electron Stuttgart–Dresden effective core potential (ECP)⁶² in combination with the corresponding valence basis set was employed. The cluster structures were fully optimized without symmetry constraints with a quasi-Newton restricted step method and the obtained minima were characterized by analytic frequency analyses. To investigate structural dynamic effects, we also performed canonical BOMD simulations with the same electronic structure method. However, to ease computational demand, we substituted the all-electron def2-TZVP basis by the corresponding DZVP-GGA basis⁶³ in these runs. The temperature was controlled by a three chain Nosé–Hoover thermostat^{64,65} with a coupling frequency of 260 cm^{-1} . The BOMD trajectories were recorded for 15 ps with a time step of 2.0 fs. For all calculations the GEN-A2* auxiliary function set was employed.⁶³ For the topological analysis of the density, the density critical points are located with a Newton–Raphson type method.⁶⁶ The bond paths are calculated starting from the (3,−1) bond critical points by a constrained restricted step algorithm as implemented in deMon2k.

For the magnetic shielding tensor calculations, either of the optimized ground state structures or along the BOMD trajectories, the B88-P86/def2-TZVP/GEN-A2* or PBE⁶⁷/aug-cc-pVDZ^{68,69}/GEN-A2* methodologies were used. The latter has proven particularly accurate for hydrogen shift calculations.⁷⁰ The ^1H chemical shifts for the hydrides were referenced against the ^1H chemical shift of TMS computed at the same



level of theory. The magnetically induced current densities were computed employing our recent implementation^{53,54} of the electronic structure method for the gauge-including magnetically induced currents (GIMIC) theory.⁵⁷ All computations were performed with the quantum chemistry code deMon2k.⁷¹

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Academy of Finland (grants 294217, 319208, 315549, 317739), and through H. H.'s Academy Professorship. The authors wish to acknowledge CSC – IT Center for Science, Finland, for computational resources and the Barcelona Supercomputing Center as part of PRACE project no. 2018194723. B. Z. G. acknowledges funding from CONACyT project CB-2015-258647. At Cinvestav this work was supported by the SENER-CONACyT project 280158 and the CONACyT project A1-S-11929. The authors are grateful to Dr Laura Laverdure for her gracious help providing valuable comments and criticism on the manuscript.

References

- W. S. Mahoney, D. M. Brestensky and J. M. Stryker, *J. Am. Chem. Soc.*, 1988, **110**, 291–293.
- Z. Mao, J. S. Huang, C. M. Che, N. Zhu, S. K. Y. Leung and Z. Y. Zhou, *J. Am. Chem. Soc.*, 2005, **127**, 4562–4563.
- P.-k. Liao, C.-s. Fang, A. J. Edwards, S. Kahlal, J.-y. Saillard and C. W. Liu, *Inorg. Chem.*, 2012, **51**, 6577–6591.
- E. L. Bennett, P. J. Murphy, S. Imberti and S. F. Parker, *Inorg. Chem.*, 2014, **53**, 2963–2967.
- M. A. Huertos, I. Cano, N. A. Bandeira, J. Benet-Buchholz, C. Bo and P. W. Van Leeuwen, *Chem. – Eur. J.*, 2014, **20**, 16121–16127.
- R. S. Dhayal, J. H. Liao, X. Wang, Y. C. Liu, M. H. Chiang, S. Kahlal, J. Y. Saillard and C. W. Liu, *Angew. Chem., Int. Ed.*, 2015, **54**, 13604–13608.
- T. A. D. Nguyen, Z. R. Jones, B. R. Goldsmith, W. R. Buratto, G. Wu, S. L. Scott and T. W. Hayton, *J. Am. Chem. Soc.*, 2015, **137**, 13319–13324.
- J.-H. Liao, R. S. Dhayal, X. Wang, S. Kahlal, J.-Y. Saillard and C. W. Liu, *Inorg. Chem.*, 2014, **53**, 11140–11145.
- R. S. Dhayal, J. H. Liao, H. N. Hou, R. Ervilita, P. K. Liao and C. W. Liu, *Dalton Trans.*, 2015, **44**, 5898–5908.
- T. A. D. Nguyen, B. R. Goldsmith, H. T. Zaman, G. Wu, B. Peters and T. W. Hayton, *Chem. – Eur. J.*, 2015, **21**, 5341–5344.
- P. V. Kishore, J. H. Liao, H. N. Hou, Y. R. Lin and C. W. Liu, *Inorg. Chem.*, 2016, **55**, 3663–3673.
- T. A. D. Nguyen, Z. R. Jones, D. F. Leto, G. Wu, S. L. Scott and T. W. Hayton, *Chem. Mater.*, 2016, **28**, 8385–8390.
- S. Liu, M. S. Eberhart, J. R. Norton, X. Yin, M. C. Neary and D. W. Paley, *J. Am. Chem. Soc.*, 2017, **139**, 7685–7688.
- H. Z. Ma, J. Li, A. J. Canty and R. A. O'Hair, *Dalton Trans.*, 2017, **46**, 14995–15003.
- H. H. Nie, Y. Z. Han, Z. Tang, S. Y. Yang and B. K. Teo, *J. Cluster Sci.*, 2018, **29**, 837–846.
- A. W. Cook and T. W. Hayton, *Acc. Chem. Res.*, 2018, **51**, 2456–2464.
- T. Nakajima, Y. Kamiryo, K. Hachiken, K. Nakamae, Y. Ura and T. Tanase, *Inorg. Chem.*, 2018, **57**, 11005–11018.
- K. K. Chakrahari, R. P. Silalahi, J. H. Liao, S. Kahlal, Y. C. Liu, J. F. Lee, M. H. Chiang, J. Y. Saillard and C. W. Liu, *Chem. Sci.*, 2018, **9**, 6785–6795.
- P. V. Kishore, D. R. Shi, J. H. Liao, A. K. Gupta and C. W. Liu, *Inorg. Chim. Acta*, 2019, **496**, 119068.
- A. Ghosh, R. W. Huang, B. Alamer, E. Abou-Hamad, M. N. Hedhili, O. F. Mohammed and O. M. Bakr, *ACS Mater. Lett.*, 2019, **1**, 297–302.
- A. Chen, X. Kang, S. Jin, W. Du, S. Wang and M. Zhu, *J. Phys. Chem. Lett.*, 2019, **10**, 6124–6128.
- H. Shen, L. Wang, O. López-Estrada, C. Hu, Q. Wu, D. Cao, S. Malola, B. K. Teo, H. Häkkinen and N. Zheng, *Nano Res.*, 2021, **14**, 3303–3308.
- H. Xu, Y. Z. Han, J. OuYang, Z. C. Chen, H. J. Chen, H. H. Nie, Z. Tang, S. Y. Yang, R. B. Huang, L. S. Zheng and B. K. Teo, *Dalton Trans.*, 2021, **50**, 4028–4035.
- C. Sun, B. K. Teo, C. Deng, J. Lin, G. G. Luo, C. H. Tung and D. Sun, *Coord. Chem. Rev.*, 2021, **427**, 213576.
- R. P. Brocha Silalahi, G. R. Huang, J. H. Liao, T. H. Chiu, K. K. Chakrahari, X. Wang, J. Cartron, S. Kahlal, J. Y. Saillard and C. W. Liu, *Inorg. Chem.*, 2020, **59**, 2536–2547.
- R.-W. Huang, J. Yin, C. Dong, A. Ghosh, M. J. Alhilaly, X. Dong, M. N. Hedhili, E. Abou-Hamad, B. Alamer, S. Nematulloev, Y. Han, O. F. Mohammed and O. M. Bakr, *J. Am. Chem. Soc.*, 2020, **142**, 8696–8705.
- S. K. Barik, S. C. Huo, C. Y. Wu, T. H. Chiu, J. H. Liao, X. Wang, S. Kahlal, J. Y. Saillard and C. W. Liu, *Chem. – Eur. J.*, 2020, **26**, 10471–10479.
- D. A. Ekanayake, A. Chakraborty, J. A. Krause and H. Guan, *Inorg. Chem.*, 2020, **59**, 12817–12828.
- T. Nakajima, K. Nakamae, R. Hatano, K. Imai, M. Harada, Y. Ura and T. Tanase, *Dalton Trans.*, 2019, **48**, 12050–12059.
- K. K. Chakrahari, R. P. Silalahi, T. H. Chiu, X. Wang, N. Azrou, S. Kahlal, Y. C. Liu, M. H. Chiang, J. Y. Saillard and C. W. Liu, *Angew. Chem., Int. Ed.*, 2019, **58**, 4943–4947.
- A. Aloisi, É. Crochet, E. Nicolas, J. C. Berthet, C. Lescot, P. Thuéry and T. Cantat, *Organometallics*, 2021, **40**, 2064–2069.
- R. S. Dhayal, W. E. Van Zyl and C. W. Liu, *Dalton Trans.*, 2019, **48**, 3531–3538.
- R. S. Dhayal, J.-H. Liao, Y.-R. Lin, P.-K. Liao, S. Kahlal, J.-Y. Saillard and C. W. Liu, *J. Am. Chem. Soc.*, 2013, **135**, 4704–4707.
- J. Li, H. Z. Ma, G. E. Reid, A. J. Edwards, Y. Hong, J. M. White, R. J. Mulder and R. A. O'Hair, *Chem. – Eur. J.*, 2018, **24**, 2070–2074.



- 35 J. Li, J. M. White, R. J. Mulder, G. E. Reid, P. S. Donnelly and R. A. O'Hair, *Inorg. Chem.*, 2016, **55**, 9858–9868.
- 36 A. J. Edwards, R. S. Dhayal, P. K. Liao, J. H. Liao, M. H. Chiang, R. O. Piltz, S. Kahlal, J. Y. Saillard and C. W. Liu, *Angew. Chem., Int. Ed.*, 2014, **53**, 7214–7218.
- 37 A. J. Jordan, G. Lalic and J. P. Sadighi, *Chem. Rev.*, 2016, **116**, 8318–8372.
- 38 T. Nakajima, K. Nakamae, Y. Ura and T. Tanase, *Eur. J. Inorg. Chem.*, 2020, **2020**, 2211–2226.
- 39 C. Sun, N. Mammen, S. Kaappa, P. Yuan, G. Deng, C. Zhao, J. Yan, S. Malola, K. Honkala, H. Häkkinen, B. K. Teo and N. Zheng, *ACS Nano*, 2019, **13**, 5975–5986.
- 40 T. Nakajima, Y. Kamiryo, M. Kishimoto, K. Imai, K. Nakamae, Y. Ura and T. Tanase, *J. Am. Chem. Soc.*, 2019, **141**, 8732–8736.
- 41 A. W. Cook, T. A. D. Nguyen, W. R. Buratto, G. Wu and T. W. Hayton, *Inorg. Chem.*, 2016, **55**, 12435–12440.
- 42 C. M. Zall, J. C. Linehan and A. M. Appel, *J. Am. Chem. Soc.*, 2016, **138**, 9968–9977.
- 43 K. Nakamae, M. Tanaka, B. Kure, T. Nakajima, Y. Ura and T. Tanase, *Chem. – Eur. J.*, 2017, **23**, 9457–9461.
- 44 S. Lee, M. S. Bootharaju, G. Deng, S. Malola, W. Baek, H. Häkkinen, N. Zheng and T. Hyeon, *J. Am. Chem. Soc.*, 2020, **142**, 13974–13981.
- 45 K. Nakamae, T. Nakajima, Y. Ura, Y. Kitagawa and T. Tanase, *Angew. Chem., Int. Ed.*, 2020, **59**, 2262–2267.
- 46 P. Yuan, R. Chen, X. Zhang, F. Chen, J. Yan, C. Sun, D. Ou, J. Peng, S. Lin, Z. Tang, B. K. Teo, L. S. Zheng and N. Zheng, *Angew. Chem., Int. Ed.*, 2019, **58**, 835–839.
- 47 R. S. Dhayal, W. E. Van Zyl and C. W. Liu, *Acc. Chem. Res.*, 2016, **49**, 86–95.
- 48 S. Wang, T. Liu and D. E. Jiang, *ACS Appl. Mater. Interfaces*, 2021, **13**, 53468–53474.
- 49 S. Wang, Z. Wu, S. Dai and D. en Jiang, *Angew. Chem., Int. Ed.*, 2021, **60**, 12289–12292.
- 50 B. Reif, S. E. Ashbrook, L. Emsley and M. Hong, *Nat. Rev. Methods Primers*, 2021, **1**, 2.
- 51 T. Helgaker, M. Jaszuński and K. Ruud, *Chem. Rev.*, 1999, **99**, 293–352.
- 52 R. S. Dhayal, J.-H. Liao, S. Kahlal, X. Wang, Y.-C. Liu, M.-H. Chiang, W. E. vanZyl, J.-Y. Saillard and C. W. Liu, *Chem. – Eur. J.*, 2015, **21**, 8369–8374.
- 53 O. López-Estrada, E. Selenius, B. Zuniga-Gutierrez, S. Malola and H. Häkkinen, *J. Chem. Phys.*, 2021, **154**, 204303.
- 54 O. López-Estrada, B. Zuniga-Gutierrez, E. Selenius, S. Malola and H. Häkkinen, *Nat. Commun.*, 2021, **12**, 2477.
- 55 R. Bader, *Atoms in Molecules, A Quantum Theory*, Oxford University Press, 1990.
- 56 R. Bianchi, G. Gervasio and D. Marabello, *Inorg. Chem.*, 2000, **39**, 2360–2366.
- 57 D. Sundholm, H. Fliegl and R. J. Berger, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2016, **6**, 639–678.
- 58 P. Calaminici, A. Alvarez-Ibarra, D. Cruz-Olvera, V.-D. Dominguez-Soria, R. Flores-Moreno, G. U. Gamboa, G. Geudtner, A. Goursot, D. Mejia-Rodriguez, D. R. Salahub, B. Zuniga-Gutierrez and A. M. Köster, in *Auxiliary Density Functional Theory: From Molecules to Nanostructures*, ed. J. Leszczynski, A. Kaczmarek-Kedziera, T. Puzyn, M. G. Papadopoulos, H. Reis and M. K. Shukla, Springer International Publishing, Cham, 2017, pp. 795–860.
- 59 A. D. Becke, *Phys. Rev. A*, 1988, **38**, 3098–3100.
- 60 J. P. Perdew, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **33**, 8822–8824.
- 61 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 62 B. P. Pritchard, D. Altarawy, B. Didier, T. D. Gibson and T. L. Windus, *J. Chem. Inf. Model.*, 2019, **59**, 4814–4820.
- 63 P. Calaminici, F. Janetzko, A. M. Köster, R. Mejia-Olvera and B. Zuniga-Gutierrez, *J. Chem. Phys.*, 2007, **126**, 044108.
- 64 S. Nosé, *J. Chem. Phys.*, 1984, **81**, 511–519.
- 65 W. G. Hoover, *Phys. Rev. A*, 1985, **31**, 1695–1697.
- 66 M. Leboeuf, A. M. Köster, K. Jug and D. R. Salahub, *J. Chem. Phys.*, 1999, **111**, 4893–4905.
- 67 J. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 68 D. E. Woon and T. H. Dunning, *J. Chem. Phys.*, 1993, **98**, 1358–1371.
- 69 N. B. Balabanov and K. A. Peterson, *J. Chem. Phys.*, 2005, **123**, 064107.
- 70 B. Zuniga-Gutierrez, G. Geudtner and A. M. Köster, *J. Chem. Phys.*, 2011, **134**, 124108.
- 71 A. Köster, G. Geudtner, A. Alvarez-Ibarra, P. Calaminici, M. Casida, J. Carmona-Espindola, V. Dominguez, R. Flores-Moreno, G. Gamboa, A. Goursot, T. Heine, A. Ipatov, A. de la Lande, F. Janetzko, J. del Campo, D. Mejia-Rodriguez, J. U. Reveles, J. Vasquez-Perez, A. Vela, B. Zuniga-Gutierrez and D. Salahub, *deMon2k, Version 6*, The deMon developers, Cinvestav, Mexico City, 2018, <https://www.demon-software.com/>.

