



Cite this: *Mater. Adv.*, 2022, 3, 3961

# Confirmation of pore formation mechanisms in biochars and activated carbons by dual isotherm analysis†

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In this study biochars and activated carbons were synthesized either directly *via* the pyrolysis of sodium carboxymethyl cellulose (NC) or *via* hydrothermal carbonization of sawdust (SD) in an aqueous solution of KOH. The amount of porogen was varied by modulating the degree of sodium carboxymethyl substitution on NC or the amount of KOH mixed in solution with SD. Pore size distributions (PSDs) of these carbons were determined from the dual fit of kernels based on the two-dimensional version of the nonlocal density functional theory (2D-NLDFT) heterogeneous surface models to either N<sub>2</sub> and H<sub>2</sub> or O<sub>2</sub> and H<sub>2</sub> isotherms measured at −196 °C. By comparing PSDs of carbons from the same starting material at increasing degrees of activation, we show that those derived using O<sub>2</sub> and H<sub>2</sub> isotherms not only give more detail of variations in pore size but that the results also fit better with current understandings of porosity development in carbons derived through oxidative activation. This is likely a result of superior diffusion of O<sub>2</sub> into ultramicropores at low pressure relative to N<sub>2</sub>.

Received 9th February 2022,  
Accepted 29th March 2022

DOI: 10.1039/d2ma00141a

rsc.li/materials-advances

## 1. Introduction

Porosimetry *via* the measurement of gas adsorption isotherms has provided a great depth of insight into the nature of pore channels in solids, particularly on disordered-like materials for which other techniques are not suitable for pore sizing. While N<sub>2</sub> at −196 °C remains the dominant adsorptive used in such experiments due to its availability, studies have concluded that it is lacking in some properties required for good isotherm measurement and thus accurate determination of textural parameters.<sup>1–5</sup> In the case of porous carbons as well as other porous materials such as conjugated microporous polymers (CMPs), it is well established that pores of width similar to the size of the N<sub>2</sub> molecule are common. This can lead to inadequate diffusion of N<sub>2</sub> into such pores when it is used as an adsorptive.<sup>5–7</sup> Further, polar groups on the surface of many carbons may interact with the N<sub>2</sub> molecule's large quadrupole moment, which obfuscates the molecular cross-sectional area used in the calculation of various measures of porosity.<sup>2,8,9</sup> As such, alternative adsorptives have enjoyed some use in recent years, including Ar, O<sub>2</sub>, H<sub>2</sub>, CO<sub>2</sub>, CH<sub>4</sub>, CF<sub>4</sub>, and SF<sub>6</sub> as they are

either smaller in size or have a lower quadrupole moment than N<sub>2</sub> – or both (see Table 1).<sup>10–15</sup> Simultaneously the development of advanced DFT kernels for an array of adsorptives on carbons, in particular the 2D-NLDFT heterogeneous surface model mean that quantities such as pore size, volume, and surface area can now be more accurately determined than in the past. This model accounts for surface roughness present in the graphene sheets which make up turbostratic porous carbons by assuming a corrugated slit pore model.<sup>15,16</sup> The nanoporous carbon structure based on curved graphene building blocks was demonstrated by using atomic resolution images taken within nanoporous carbon samples.<sup>17</sup>

Nonetheless, a single isotherm cannot typically be used to accurately determine pore sizes across the entire micropore

Table 1 Physical properties of relevant molecules for porosimetry measurements<sup>20–23</sup>

Species	$d_k/\text{\AA}$	$T_b/^\circ\text{C}$	$\mu$
H <sub>2</sub>	2.89	−252	0.260
CO <sub>2</sub>	3.30	−78 <sup>a</sup>	2.139
CH <sub>4</sub>	3.80	−161	0.000
N <sub>2</sub>	3.64	−196	0.697
O <sub>2</sub>	3.46	−183	0.155
Ar	3.40	−186	—
CF <sub>4</sub>	4.70	−128	—
SF <sub>6</sub>	5.50	−51	—

<sup>a</sup> Sublimes.  $k_d$  = kinetic diameter,  $T_b$  = boiling point,  $\mu$  = quadrupole moment

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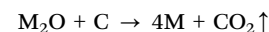
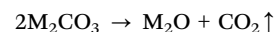
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† Electronic supplementary information (ESI) available: One table, five figures, all isotherms in human- and machine-readable.aif files. See DOI: 10.1039/d2ma00141a

region. For example, while  $\text{H}_2$  sorption at  $-196^\circ\text{C}$  can probe pores as small as  $3\text{ \AA}$  its upper limit is  $\sim 10\text{ \AA}$ . Conversely,  $\text{O}_2$  adsorption at  $-196^\circ\text{C}$  is best utilized for probing pores larger than  $7\text{ \AA}$ , although it has been shown to also penetrate pores smaller than this limit to some degree.<sup>18</sup> In order to take advantage of the information gained from measuring two isotherms, Jagiello *et al.* developed a method whereby appropriate kernels could be simultaneously fit to two isotherms to yield a single pore size distribution (PSD).<sup>2,5</sup> This method was tested by using a variety of combinations of isotherms, including  $\text{N}_2/\text{CO}_2$ ,  $\text{O}_2/\text{H}_2$ , and  $\text{N}_2/\text{H}_2$ . Generally speaking, the results derived from each of these combinations yields similar PSDs.<sup>2,4,5</sup> However, as reported by Beda *et al.*,  $\text{N}_2$ 's poor diffusion into ultramicropores means that  $\text{N}_2$  and  $\text{O}_2$  isotherms on so-called hard carbons can be quite different, and as a result PSDs may differ significantly with  $\text{O}_2$  showing much more porosity in the ultramicropore region.<sup>18</sup> In addition, the high quadrupole moment of  $\text{CO}_2$  can lead to incorrect results due to its interaction with polar moieties which are common on activated carbons.<sup>19</sup> The combination of  $\text{O}_2$  and  $\text{H}_2$  may provide more accurate information than the  $\text{N}_2$  and  $\text{CO}_2$  isotherms due to the former pair's lower quadrupolarity.<sup>2,4</sup> Despite being fairly reactive gases at room temperature, under cryogenic conditions the reactivity is extremely low, so is unlikely to affect carbon surface chemistry.

Within the group of turbostratic porous carbons, there are two closely related subtypes. The first is biochar, which is produced upon the pyrolysis of biomass or other carbon-rich material.<sup>24,25</sup> So-called activated carbon, on the other hand, differs from biochar due to the addition of an activating agent/porogen to aid in the development of porosity.<sup>25–27</sup> This activating agent can produce porosity either through physical or chemical means. The former proceeds *via* gasification using an oxidizing gas such as  $\text{CO}_2$ , which burns away volatile matter.<sup>26,28,29</sup> Gasification also occurs during biomass pyrolysis as  $\text{CO}_2$  is produced from volatile carbon moieties, as such biochar can be thought of as a form of physically activated carbon.<sup>24–26</sup> At sufficiently low temperatures ( $< 900^\circ\text{C}$ ), gasification produces very narrow pores of width  $< 7\text{ \AA}$ , referred to as ultramicropores.<sup>5</sup>

For chemical activation, while there is a large variety of porogens, amongst the most popular are the alkali metal hydroxides KOH and NaOH due to the fact that they facilitate a high degree of control over porosity, this being especially the case for KOH.<sup>26</sup> Additionally, the majority of pores generated by alkali metal hydroxides have widths typically  $< 20\text{ \AA}$ , *i.e.* micropores. Such pore structures are desirable for their utility in applications such as small molecule capture and storage as well as in supercapacitors.<sup>26,30–32</sup> In alkali metal hydroxide (MOH) activation, pores are developed *via* the oxidation of C to  $\text{CO}_2$ , which in turn leads to further gasification. Finally, residual alkali metal atoms (M) intercalate into voids between graphitic layers resulting in further micropore development.<sup>27,33–36</sup>



The use of MOH increases the overall pore volume and broadens the PSD relative to that for carbons derived in the absence of a porogen.<sup>5,25,26</sup> Thus so-called biochars are highly ultramicroporous, while MOH-activated carbons possess a higher degree of supermicroporosity (micropores of width  $> 7\text{ \AA}$ ). As mentioned above, there are a variety of reasons that  $\text{N}_2$  may be unsuitable as a probe for these smaller pores present in biochars. Therefore, given the critical importance of being able to accurately determine textural properties with respect to any targeted applications, herein we investigate the use of combinations of  $\text{N}_2$ ,  $\text{O}_2$ , and  $\text{H}_2$  isotherms and their effectiveness in assessing the porosity of biochars and activated carbons.

## 2. Experimental

### 2.1. Carbon synthesis

Carbons were synthesized with and without the use of a porogen, but keeping all other variables constant. In the case of carbons derived from sawdust, this was done by addition of KOH at the hydrothermal carbonization step. For carbons from sodium carboxymethyl cellulose the degree of substitution, *i.e.* the average number of carboxymethyl sodium groups per monomer was varied from 0.0 to 0.9.

**2.1.1. *Via* hydrothermal carbonisation of sawdust.** Sawdust was treated by hydrothermal carbonization at 250 and  $300^\circ\text{C}$  for 2 h at a ramp rate of  $5^\circ\text{C min}^{-1}$  in an aqueous solution of KOH. The concentration of sawdust was  $0.32\text{ g cm}^{-3}$  in each experiment, whereas KOH:sawdust (w/w) was varied from 0.0, 0.5 to 1.0 at each hydrothermal carbonization temperature. This produced a total of six black slurries which were dried overnight at  $100^\circ\text{C}$  in an alumina boat, and then activated by heating under nitrogen at  $800^\circ\text{C}$  for 1 h (ramp rate  $3^\circ\text{C min}^{-1}$ ) and allowed to cool. In all cases the resultant carbons were stirred in HCl (10% v/v,  $600\text{ cm}^3$ ) thoroughly prior to filtration and washing in water ( $\sim 5\text{ L}$ ) to remove any chloride salts, until washings were neutral (pH 7). Finally, samples were dried at  $100^\circ\text{C}$ , yielding six carbons designated SAx-*T*, where *x* indicates KOH:sawdust ratio and *T* is the hydrothermal carbonization temperature.

**2.1.2. From Sodium carboxymethyl cellulose.** Sodium carboxymethyl cellulose (Sigma) with a degree of substitution of 0.0, 0.7, and 0.9 (the first being pure cellulose) was heated under nitrogen to  $800^\circ\text{C}$  for 1 h (ramp rate  $3^\circ\text{C min}^{-1}$ ) then cooled, washed, and dried in the same manner as the SAx-*T* samples. This yielded three samples designated NCx where *x* indicates the degree of substitution.

### 2.2. Characterisation

CHN Elemental analysis of the raw sawdust precursor was performed using an Exeter Analytical CE-440 Elemental Analyser. O content is assumed to be the remaining percentage when



C, H, and N are accounted for, as other elements do not occur in significant quantities.

Thermogravimetric analysis was performed using a TA Q500 Thermogravimetric Analyser, in order to determine if any residual metals remained in the carbon samples (see Table S1, ESI†). All samples were analysed using a platinum pan and in the presence of air/argon. The parameters for all experiments were: Ramp  $10\text{ }^{\circ}\text{C minute}^{-1}$  from  $20\text{--}1000\text{ }^{\circ}\text{C}$  with an isotherm for 10 minutes at  $1000\text{ }^{\circ}\text{C}$ , gas flow:  $60\text{ mL minute}^{-1}$ .

$\text{H}_2$ ,  $\text{O}_2$  and  $\text{N}_2$  isotherms were measured at  $-196\text{ }^{\circ}\text{C}$  using a Micromeritics 3flex sorptometer. In the case of  $\text{O}_2$  and  $\text{N}_2$  isotherms were measured up to  $P_0$  ( $\sim 220$  and  $1013\text{ mbar}$ , respectively), which is measured as each equilibrated point in the isotherm is recorded to account for small variations in temperature. However for  $\text{H}_2$  the isotherm  $P_0$  is taken as a constant of  $1013\text{ mbar}$  for convenience because  $\text{H}_2$  is supercritical at  $-196\text{ }^{\circ}\text{C}$ . All samples were degassed at  $300\text{ }^{\circ}\text{C}$  for 16 h prior to isotherm measurement. Porosity calculations were performed *via* the dual-fitting of 2D-NLDFT heterogeneous surface kernels on either  $\text{O}_2$  and  $\text{H}_2$  or  $\text{N}_2$  and  $\text{H}_2$  isotherms using SAIEUS software.<sup>16,37</sup> While 2D-NLDFT kernels account for surface roughness *via* the use of a corrugated slit-pore model,<sup>16</sup> network connectivity is accounted for by the use of two differently sized probe molecules.<sup>15</sup> All dual-fit calculations were performed using an overall pore width minimum of  $3.00\text{ }\text{\AA}$ , a minimum for the  $\text{O}_2$  or  $\text{N}_2$  kernel of  $3.60\text{ }\text{\AA}$  and the pore width maximum was set at  $500.00\text{ }\text{\AA}$ . For consistency the fitting parameter,  $\lambda^{38-40}$  is fixed at 4. For comparison, conventional determinations of pore volume (*via*  $t$ -plot) were obtained from  $\text{N}_2$  isotherms alone. Raw isotherms are available as human- and machine-readable AIF files in the SI.<sup>41</sup>

**2.2.1. Equilibration interval and rate of change.** The instrument operation software allows for designation of a so-called equilibration interval,  $t_E$ , as a parameter for the isotherm. This is used in order to estimate when equilibration of the adsorption system has occurred. In the measurement of a single point on the isotherm, the pressure is measured at intervals of  $t_E$  and a rate of change in pressure,  $R$ , is determined over the case of ten intervals. For example, for a  $t_E$  of  $10\text{ s}$ , pressure will be measured ten times over  $100\text{ s}$  and  $R$  is calculated over this period. When  $R < 0.010\%$ , the system is considered to be at equilibrium and the isothermal point,  $V(P/P_0)$  is determined. As such,  $t_E$  can be decreased in order to speed up analysis but this risks reducing precision of measurement, *i.e.* the system is less likely to actually be at equilibrium when the isothermal point is recorded.

### 3. Results and discussion

#### 3.1. Diffusion of adsorptives

During the measurement of isotherms, it became obvious that the duration of measurement of  $\text{N}_2$  isotherms on the SA0.0- $T$  and NC0.0 (biochar) samples was impractical. Using the initial  $t_E$  of  $45\text{ s}$  for  $P/P_0$  up to  $10^{-3}$ , only one or two points were achieved after  $> 72\text{ h}$ . To avoid this problem,  $t_E$  were iteratively

reduced to  $15\text{ s}$  for determination of  $\text{N}_2$  isotherms on SA0.0- $T$ . Even with this modification, overall analysis time was up to  $50\text{ h}$ ; twice the time taken for a similarly determined  $\text{O}_2$  isotherm. This implies that the pore structure of biochars is such that  $\text{N}_2$  has difficulty diffusing into the smallest pores, possibly indicating that the pore width is similar to the kinetic diameter of the adsorptive. Alternatively, this may be a result of the interaction between polar moieties outside of the ultramicropores and  $\text{N}_2$ 's quadrupole moment. The attraction between  $\text{N}_2$  and such moieties may be competitive with  $\text{N}_2$ 's diffusion into ultramicropores. The less activated materials are also likely to have a higher oxygen content and thus more polar functional groups, meaning it is difficult to distinguish which of these phenomena is creating this effect. However, similarly slow diffusion of  $\text{N}_2$  has been observed for ultramicroporous CMPs with no polar groups.<sup>7</sup> Regardless, the poor diffusion onto the biochar/less activated samples calls into question the accuracy of textural parameters derived from  $\text{N}_2$  isotherms.

This discrepancy is most evident when comparing the time taken between isothermal points,  $t$ , as shown in Fig. 1a. While  $t$  is broadly similar for  $\text{O}_2$  and  $\text{N}_2$  isotherms of SA1.0-250 as well as the  $\text{O}_2$  isotherm of SA0.0-250,  $t$  is much higher in the low relative pressure region for the  $\text{N}_2$  isotherm of SA0.0-250, with a

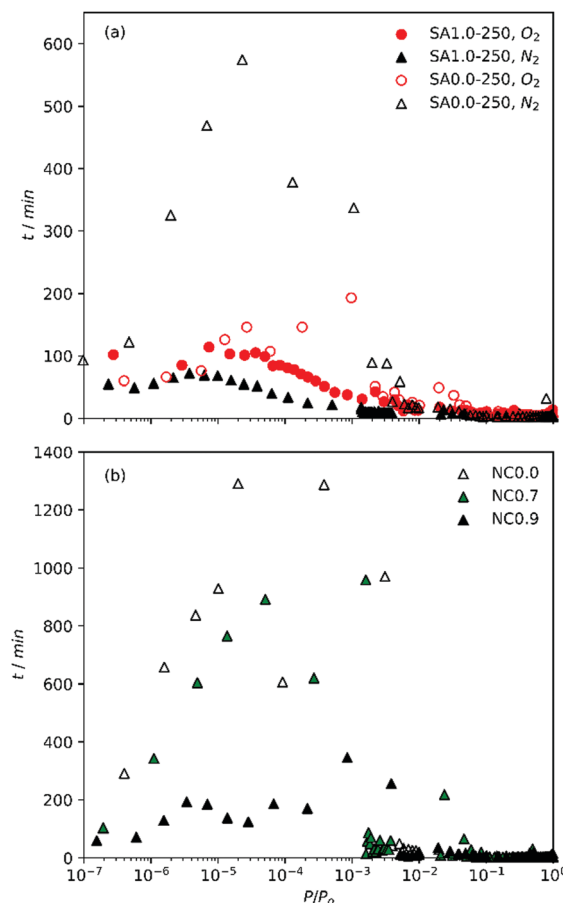


Fig. 1 Time between points with increasing relative pressure of (a)  $\text{O}_2$  and  $\text{N}_2$  isotherms on SA1.0-250 and SA0.0-250 and (b)  $\text{N}_2$  isotherms of NCx carbons.



single point taking up to 10 h to measure. This suggests that the slow diffusion may be related to compatibility of N<sub>2</sub> with some feature of the biochar's porosity. Similarly, when comparing the time taken between points for the NCx samples synthesized with increasing amounts of Na within the cellulosic structure (Fig. 1b) it is apparent that expansion of the porosity by use of a porogen improves the ability of N<sub>2</sub> to diffuse into the pores, and that this improvement is accelerated at some point between a degree of substitution of 0.7 and 0.9. That is, while NC0.7 allows faster diffusion than NC0.0, the improvement is only small but there is a far more marked change in *t* when *x* rises to 0.9. Indeed, while *t* can be extremely large (approaching 24 h) for NC0.0 and NC0.7, the maximum *t* for NC0.9 is closer (~6 h) to that for the activated carbons derived from sawdust. Our observations of the lengthy measurement of N<sub>2</sub> isotherms relative to O<sub>2</sub> are counter to work recently published by Beda *et al.*, wherein measurement of N<sub>2</sub> isotherms on so-called hard-carbons was fast relative to O<sub>2</sub>.<sup>18</sup>

This may be an indication of relative pore sizes *i.e.* that ultramicropores in biochars reported in this work are partially penetrable by N<sub>2</sub>, unlike those in hard carbons reported by Beda *et al.*

Table 2 shows the atomic ratios of porogen to carbon. As the principle pore development process in activated carbons begins with oxidation of carbon by the porogen, pore broadening therefore increases with increased porogen/C ratio. These ratios are much lower in the NCx samples than for SAx-T. While variables such as precursor structure and composition as well as differences in reaction kinetics of Na *vs.* K also play a factor,<sup>33–35</sup> it is also clear that the higher amounts of K used in the synthesis of SAx-T samples results in porosity that is much more amenable to N<sub>2</sub> diffusion into the pores.

Gravimetric yields of samples described in this work can be found in Table S1 (ESI†). In general, yield decreases with increasing amount of activating agent. The yield of SA0.0-300 is much higher (26%) than that of SA0.0-250 (15%). This is likely a result of greater stability of hydrochar synthesised at 300 °C on account of the lower amounts of oxygen rich-moieties.<sup>42</sup>

### 3.2. The isotherms

Inspection of isotherms further reveals some uncertainty in the accuracy of N<sub>2</sub> measurements on SA0.0-T carbons (see Fig. S1, ESI†). While a characteristic plateau exists for N<sub>2</sub> isotherms of SA0.5-T (Fig. S1(2a, b), ESI†) and SA1.0-T (Fig. S1(3a, b), ESI†)

samples, indicating micropore filling, this is not the case for SA0.0-T samples (Fig. S1(1a, b), ESI†), where there is a linear increase in adsorption in the mesopore region. This is not observed in O<sub>2</sub> isotherms, wherein monolayer completion is much more obvious. All samples also have a high H<sub>2</sub> uptake indicating the presence of pores in the so-called ultramicropore region, which may hinder adsorption of N<sub>2</sub>.

Furthermore, isotherms for NCx carbons (Fig. S2, ESI†) lend further credence to the notion that significant pore broadening does not occur until a critical ratio of porogen/C (in the range 0.09–0.12) is achieved. For both NC0.0 (Fig. S2(a), ESI†) and NC0.7 (Fig. S2(b), ESI†), the maximum H<sub>2</sub> adsorption is similar to or greater than that of N<sub>2</sub>, indicating the presence of a high proportion of ultramicropores. For NCx samples where porogen/C ratio is increased to 0.12 (NC0.9), H<sub>2</sub> adsorption falls significantly below the maximum for N<sub>2</sub>, indicating that porosity has broadened into the supermicropore region. Samples NC0.7 and NC0.9 also exhibit a rapid increase in adsorption for O<sub>2</sub> and N<sub>2</sub> as *P/P*<sub>0</sub> approaches 1 such that the resulting isotherms exhibit some type II character. Such shape is likely a morphological effect, either due to condensation of sorptive in macropores or large interparticle voids.<sup>1</sup> As the materials contain no non-carbonaceous matter (within experimental error, see Table S1, ESI†) this cannot be ascribed to some chemical effect. Type II character is not observed for NC0.0, thus it can be assumed that the morphological phenomenon is a result of some action of Na<sup>+</sup>.

In the case of both NC0.0 and SA0.0-300 (Fig. S1(1b), ESI†), the H<sub>2</sub> isotherm actually crosses the N<sub>2</sub> isotherm which is anomalous since under the conditions used H<sub>2</sub> is supercritical, meaning that multilayer adsorption cannot occur. As such this high level of adsorption must be due to monolayer adsorption of H<sub>2</sub> in pores impenetrable to N<sub>2</sub>. Such pores however do appear to be penetrable by O<sub>2</sub>, which is interesting as while O<sub>2</sub> is slightly smaller than N<sub>2</sub> the kinetic diameters of the two molecules are quite similar – 3.46 and 3.64 Å for O<sub>2</sub> and N<sub>2</sub> respectively.<sup>20,21</sup>

### 3.3. Dual-fit analysis

Dual fits to O<sub>2</sub>/H<sub>2</sub> and N<sub>2</sub>/H<sub>2</sub> isotherms for SAx-T and NCx carbons are shown in Fig. S3 and S4 (ESI†), respectively. As indicated from the isotherms themselves, PSDs for SAx-T carbons (Fig. 2(1P, 2P), S5(1P, 2P), ESI†) derived using O<sub>2</sub>/H<sub>2</sub> isotherms indicate greater porosity than when using N<sub>2</sub>/H<sub>2</sub>, and the same is true for PSDs of NCx carbons (Fig. 3(1P, 2P)). The position of the first maximum (~4 Å) is consistent across all derived PSDs regardless of the subset of isotherms is used as this is mainly derived from the H<sub>2</sub> isotherm. However, with regards to the second maximum, the two sets of PSDs for both SAx-T and NCx offer slightly different pictures. While N<sub>2</sub>/H<sub>2</sub> derived PSDs show either only small differences or indeed a small decrease in pore size for the activated carbons, O<sub>2</sub>/H<sub>2</sub> PSDs show that porosity expands out of the ultramicropore and into the supermicropore region when KOH is used as porogen. Furthermore, there is some indication of PSD broadening in the second maximum. The latter narrative fits much better with

**Table 2** Atomic ratios of components in precursors, determined from molecular compositions for NCx samples and from weight ratios and elemental (CHN) analysis in the case of SAx samples

Sample	Porogen/C	Porogen/O	O/C
NC0.0	0.00	0.00	0.83
NC0.7	0.09	0.11	0.86
NC0.9	0.12	0.13	0.87
SA0.0	0.00	0.00	0.72
SA0.5	0.21	0.72	0.72
SA1.0	0.43	1.45	0.72





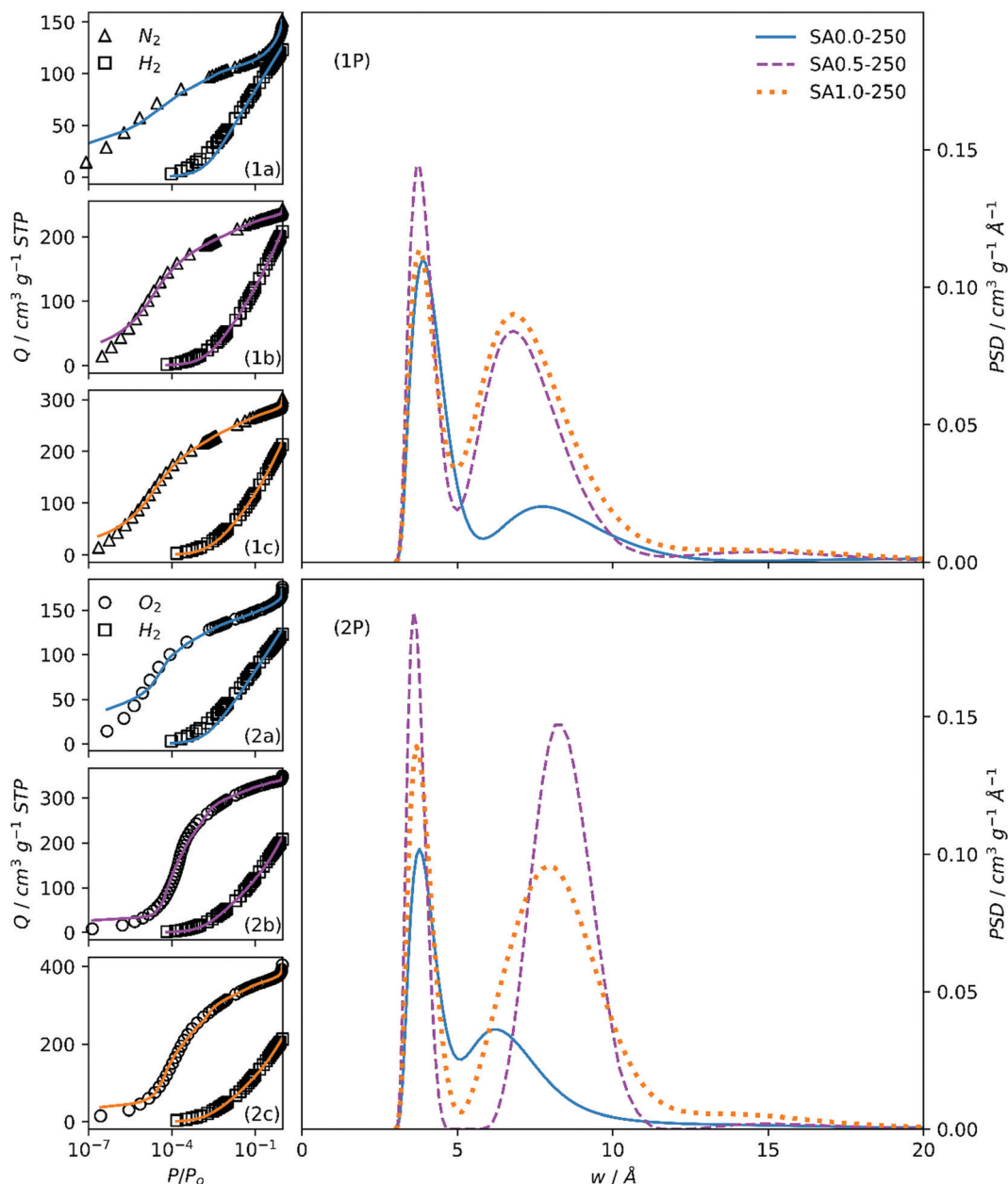


Fig. 2 Fits to  $N_2/H_2$  (1a, 1b, 1c) and  $O_2/H_2$  (2a, 2b, 2c) isotherms of samples  $SAx-250$  and resultant PSDs (1P, 2P).

current understanding of pore formation in carbons, in that the use of alkali metal hydroxide porogens during pyrolysis leads to increases in pore size relative to those formed *via* gasification of biomass (*i.e.*, pyrolysis with no added porogen). Again, this discrepancy in trends in pore size with amount of porogen according to simultaneous fitting of kernels the two isotherm pairs is likely a failing of the  $N_2$  isotherm in that ultramicro-pores are not adequately probed during analysis.

While results from dual  $O_2/H_2$  fitting to their respective 2D-NLDFT kernels indicate greater compliance with current understanding of porosity development in carbons, this does not alone confirm its superiority over  $N_2/H_2$  for this application. On examining the fits, it is obvious that the respective 2D-NLDFT

kernels struggle to account for the isotherm shape at low pressures in the case of both  $O_2$  and  $N_2$ . This is particularly evident for the  $SA0.0-T$  samples, suggesting either incomplete equilibration of the system at such low pressures or inability of the  $O_2$  and  $N_2$  kernels to account for ultramicro-pore geometry and internal texture of in biochars, or a combination of both limitations.

$O_2/H_2$  analysis also gives far more information on differences in porosity between  $SA0.5-T$  and  $SA1.0-T$  samples. While the PSD calculated *via*  $N_2/H_2$  indicates that the porosity of these two activated carbons is practically identical except for a slight increase in overall pore volume,  $O_2/H_2$  analysis, on the other hand, shows a broadening in the second maxima, which results



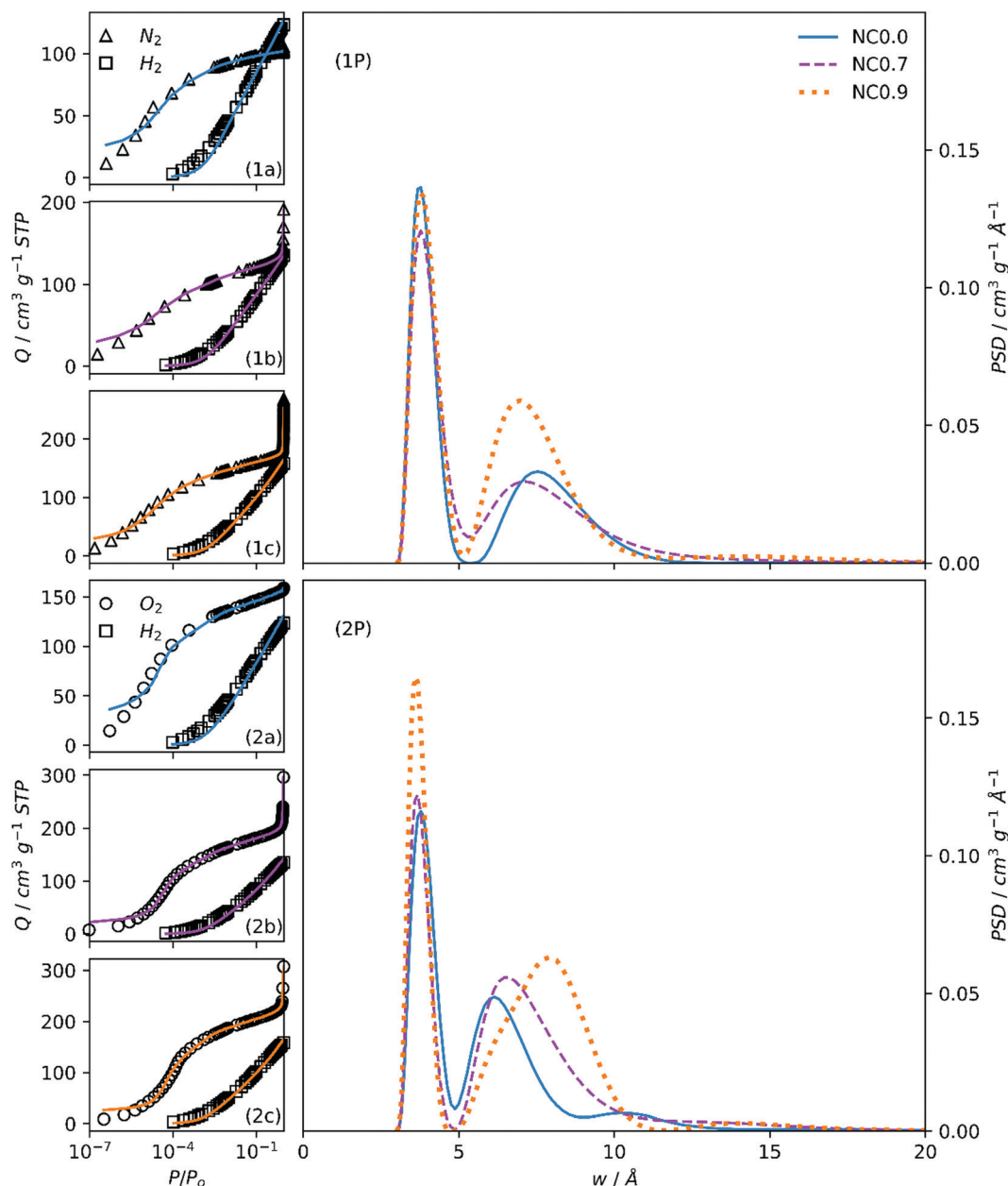


Fig. 3 Fits to  $N_2/H_2$  (1a, 1b, 1c) and  $O_2/H_2$  (2a, 2b, 2c) isotherms of samples  $NC_x$  and resultant PSDs (1P, 2P).

in the obfuscation of the two, initially distinct pore regions. A mechanistic explanation for this is that in the case of  $SA0.5-T$ , the principal route of porosity expansion comes *via* the formation of new pores by simple oxidation of as yet non-porous carbon, resulting in pores centred at *ca.* 8 Å but without significantly disturbing the ultramicropores previously formed by carbon gasification. Due to the relative increase in the quantity of K in the synthesis of  $SA1.0-T$ , the cation is able to infiltrate ultramicropores and *via* this intercalation expands them, resulting in the merging of the two previously distinct pore size regions.

The variation in porogen/C ratio for  $SAX-T$  carbons with increasing  $x$  is relatively large. Comparing PSDs of  $NC_x$  carbons

(Fig. 3(1P, 2P)) facilitates the observation of changes in porosity when the changes in porogen/C ratio are relatively small (see Table 2). Unsurprisingly, the  $N_2/H_2$  calculations again show little difference between each of the three samples apart from an indication of higher overall pore volume for  $NC0.9$ .

On the other hand, there is a gradual migration of the second PSD maxima into the supermicropore region as porogen/C is increased, accompanied by a broadening of this peak. This again is a result of the oxidative effect of the  $(Na^+)$  cation leading to PSD broadening. While the maximum  $O_2$  adsorption for  $NC0.9$  and  $NC0.7$  carbons is almost twice that of  $NC0.0$ , this is not indicated in the PSDs in the range displayed. This is because the increased porosity that causes this feature is in the macropore region.<sup>1</sup>



**Table 3** Total ( $V_t$ ), micropore ( $V_{mic}$ ), and ultramicropore ( $V_{umic}$ ) pore volumes of SAx-*T* and NCx carbons determined using classical and dual-isotherm NLDFT calculations. Classical volumes determined using the single-point and t-plot methods for  $V_t$  and  $V_{mic}$ , respectively. NLDFT calculations employed dual-fitting of isotherms to respective 2D-NLDFT, heterogeneous surface kernels. Numbers in brackets indicate percentage of pore volume made up by pores in the given pore width subregion

Sample	Pore volume/cm <sup>3</sup> g <sup>-1</sup>							
	N <sub>2</sub> classical		Dual N <sub>2</sub> /H <sub>2</sub> NLDFT			Dual O <sub>2</sub> /H <sub>2</sub> NLDFT		
	$V_t$	$V_{mic}$	$V_t$	$V_{mic}$	$V_{umic}$	$V_t$	$V_{mic}$	$V_{umic}$
SA0.0-250	0.25	0.13 (50%)	0.29	0.22 (78%)	0.16 (54%)	0.26	0.23 (90%)	0.18 (68%)
SA0.0-300	0.17	0.10 (57%)	0.25	0.20 (79%)	0.14 (56%)	0.24	0.22 (90%)	0.17 (67%)
SA0.5-250	0.37	0.32 (87%)	0.44	0.42 (97%)	0.26 (60%)	0.50	0.49 (98%)	0.15 (30%)
SA0.5-300	0.44	0.38 (85%)	0.50	0.47 (95%)	0.29 (58%)	0.53	0.48 (91%)	0.21 (39%)
SA1.0-250	0.45	0.37 (83%)	0.51	0.47 (92%)	0.26 (52%)	0.54	0.50 (92%)	0.20 (37%)
SA1.0-300	0.54	0.44 (81%)	0.63	0.58 (92%)	0.32 (50%)	0.70	0.63 (89%)	0.29 (41%)
NC0.0	0.16	0.14 (90%)	0.23	0.22 (99%)	0.15 (65%)	0.22	0.21 (95%)	0.16 (76%)
NC0.7	0.21	0.16 (77%)	0.35	0.25 (70%)	0.16 (47%)	0.38	0.24 (65%)	0.14 (58%)
NC0.9	0.28	0.22 (80%)	0.45	0.31 (69%)	0.20 (44%)	0.40	0.28 (71%)	0.13 (44%)

These inconsistencies in porosity development according to the two dual isotherm methods are further illustrated by the calculated total pore volumes, as well as the pore volume arising from micropores and ultramicropores (Table 3). Classical calculations from N<sub>2</sub> isotherms alone reveal much smaller overall pore volumes for all samples as well as lower proportions of micropores compared to their dual-fit 2D-NLDFT counterparts for NCx and SAx-*T* carbons. This is because, firstly N<sub>2</sub> alone cannot probe the smallest of ultramicropores,<sup>4,5</sup> thus up to 32% of total pore volume (relative to N<sub>2</sub>/H<sub>2</sub> analysis) is unaccounted for. Secondly, single point and t-plot calculations do not account for chemical and energetic heterogeneity of the internal pore surface unlike the 2D-NLDFT models.<sup>16</sup> Additionally, determination of total- and micropore volume by classical methods is notoriously unreliable. For example, the single-point pore volume determination relies on the isotherm possessing an obvious plateau.<sup>1</sup> As discussed above, these plateaus do not strictly exist for the N<sub>2</sub> isotherms of the SA0.0-*T* samples. As such, significant errors may arise using single-point methods.

On the other hand, the dual O<sub>2</sub>/H<sub>2</sub> fit predicts a much higher proportion of both micropores and ultramicropores than N<sub>2</sub>/H<sub>2</sub> for SA0.0-*T* carbons. Again, this is likely a result of poor diffusion of N<sub>2</sub> into ultramicropores, which limits their detection. For the so-called activated carbon samples, N<sub>2</sub>/H<sub>2</sub> shows roughly similar proportions of ultramicroporosity compared to the aforementioned biochars, while O<sub>2</sub>/H<sub>2</sub> shows drastic decreases. The latter picture from O<sub>2</sub>/H<sub>2</sub> seems more likely due to known mechanisms of porosity development by KOH. Interestingly, estimates of percent microporosity are fairly consistent for calculations of both sets of isotherms for SA0.5-*T* and SA1.0-*T* samples indicating that N<sub>2</sub> diffusion limitations are not so prevalent in supermicropores.

For NCx carbons, both overall pore volumes and absolute and percentage micropore volumes are relatively consistent across the two dual isotherm techniques. The most significant difference is in the absolute and percentage micropore volume, especially for samples NC0.0 and NC0.7. This again is an indication of the diffusion limitations encountered during

measurement of N<sub>2</sub> isotherms on these samples. While this difference in percent ultramicropore volume is 11% for these two samples, NC0.9 shows identical proportions of ultramicropores for both methods. In fact, the percentage of *micropores*, is the most similar here too – within 2%. While the broadening of the PSD with increasing porogen/C ratio from principally ultramicroporous into the supermicropore and larger regions is evident when using both techniques, this is far more obvious according to analysis of O<sub>2</sub>/H<sub>2</sub> isotherms.

Interestingly, all three measures of porosity agree on one matter – that is, when comparing the relative contribution of the different pore width regions between SAx-*T* samples with the same KOH:SD ratio (*x*), but varying hydrothermal carbonization temperature (*T*). The hydrothermal carbonization temperature does not seem to affect what percentage of pores are made up by micropores and ultramicropores. However, *total* pore volumes do vary with *T*; in the case of the activated carbons SA0.5-*T* and SA1.0-*T*, hydrothermal carbonization at 300 °C can increase  $V_t$  by up to 30% at KOH:SD ratio 1.0 according to O<sub>2</sub>/H<sub>2</sub> NLDFT calculations. The apparent difference decreases when classical techniques are used, and in the case of SA0.5-*T*, dual O<sub>2</sub>/H<sub>2</sub> analysis shows an increase in  $V_t$  of only 6% as *T* increases from 250 to 300 °C. The two dual isotherm 2D-NLDFT analyses nonetheless tend to show similar differences in  $V_t$ , so perhaps this discrepancy is a result of the relative robustness of DFT methods. On the other hand, the biochars show the opposite trend, with  $V_t$  decreasing when *T* is raised to 300 °C. This is only by a small amount (7–14%) compared to the value derived from dual isotherm calculations, but it is evident nevertheless. This decrease in porosity of biochars with hydrothermal carbonization temperature is likely a result of a correlated decreased O/C content in the hydrochar, as well as an increase in the diameter of the characteristic microspheres.<sup>42</sup> The latter phenomenon means that there is an increase in the proportion of stable hydrophobic aromatic moieties (in the core) relative to the more easily activated hydrophilic groups in the shell of the hydrochar microspheres. This compositional change is known to increase so-called activation resistance and therefore can result in a less porous



carbon upon pyrolysis.<sup>43–45</sup> Additionally, when KOH is included during hydrothermal carbonization it catalyses the breakdown of the activation-resistant lignin<sup>45–47</sup> which may further explain the improvements in porosity for  $T = 300\text{ }^{\circ}\text{C}$ .

**3.3.1. Comparison to single fit.** When PSDs are derived from two isotherms simultaneously, the SAIEUS software accounts for pore regions that are apparent in both isotherms. As an example, consider an adsorptive A which probes pores between 3 and 7 Å while adsorptive B can probe pores larger than 5 Å. For a sample which has some pores of width  $\sim 5\text{--}7\text{ Å}$ , SAIEUS will not simply add the PSDs from A and B within the full range of widths, but instead takes into account this overlap.

It is interesting therefore to compare PSDs derived through single and dual fits to isotherms as in Fig. 4. Further, the agreement between single and dual fits is interesting in terms of the degree to which a sample is activated, as such a low activation sample (SA0.0-250, Fig. 4(1a–c)) and highly activated sample (SA1.0-250, Fig. 4(2a–c)) were investigated. Single fit PSDs were calculated using identical parameters to those used in dual fitting, except that the maximum pore width for  $\text{H}_2$  was limited to 10 Å. Of the three individually fit adsorptives, only  $\text{H}_2$  produces a bimodal PSD similar to that found in dual fits, while  $\text{N}_2$  and  $\text{O}_2$  both show a single maximum at or above 7 Å. It appears that the position of the first maximum for  $\text{H}_2$  ( $\sim 4\text{ Å}$ ) corresponds almost exactly in all cases to that found in dual isotherm PSDs. On the other hand, the shape and position of the second maximum ( $> 5\text{ Å}$ ) appears to be some combination of the two single-isotherm PSDs. This is most apparent in the case of SA0.0-250 (Fig. 4(1b, 1c)). Conversely in the case of SA1.0-250 (Fig. 4(2b, 2c)) the position and shape of the second

maximum in the dual fit is approximately the same that of the sole  $\text{N}_2$  or  $\text{O}_2$  peak. These observations suggest that as PSD broadens with increasing degree of activation the fitting algorithm is able to rely more on individual isotherms for different sections of the PSD, *i.e.*  $\text{H}_2$  for pores 3–7 Å,  $\text{N}_2$  or  $\text{O}_2$  for pores  $> 7\text{ Å}$ . Whereas for so-called biochar samples, as the second maximum is lower than 7 Å

### 3.4. Three-way fit analysis

In order to further understand discrepancies between  $\text{O}_2/\text{H}_2$  and  $\text{N}_2/\text{H}_2$  analyses, the 2D-NLDFT heterogeneous surface kernels were simultaneously fit to all three ( $\text{O}_2$ ,  $\text{N}_2$ ,  $\text{H}_2$ ) isotherms for some samples. In particular, the aim here is to compare the quality of fit according to the degree of activation in each sample. The fits and PSDs achieved by this method are compared for SA0.0-300 and SA1.0-300 in Fig. 5. When the fits are examined using a logarithmic relative pressure axis (Fig. 5(1b, 2b)), it is very apparent that the fit is extremely poor for SA0.0-300, particularly in the low pressure region. While individual  $\text{O}_2/\text{H}_2$  and  $\text{N}_2/\text{H}_2$  analyses for SA0.0-300 had trouble with isothermal points below  $P/P_0 = 10^{-4}$  (see Fig. S3(1b, 2b), ESI†) this low quality of fit is very pronounced for the three-way fit in the entire low pressure, (*i.e.* micropore) region of the isotherm.

On the other hand, the three-way fit for SA1.0-300 is much more satisfactory. It is also of note that for both of these samples the fit to the  $\text{H}_2$  isotherm is always good, indicating that the uncertainties are a result of disagreements between the  $\text{N}_2$  and  $\text{O}_2$  isotherms rather than  $\text{H}_2$ . As for the PSDs, while peak positions for SA1.0-300 are essentially the same as those seen in

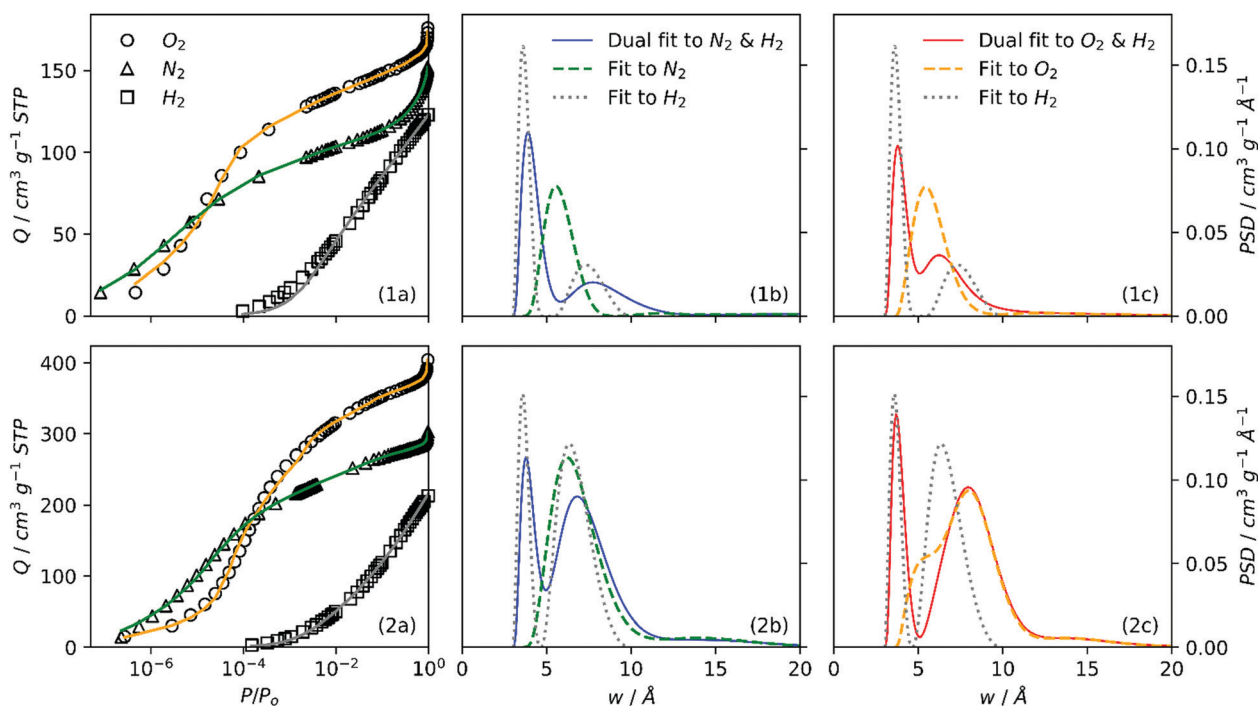


Fig. 4 Individual fits to  $\text{O}_2$ ,  $\text{N}_2$  and  $\text{H}_2$  isotherms for SA0.0-250 (1a) and SA1.0-250 (2a) and resultant PSDs from individual  $\text{N}_2$  and  $\text{H}_2$  (1b, 2b), and  $\text{O}_2$  and  $\text{H}_2$  (1c, 2c) isotherms. Corresponding dual fits overlaid for comparison (1b, 1c, 2b, 2c).





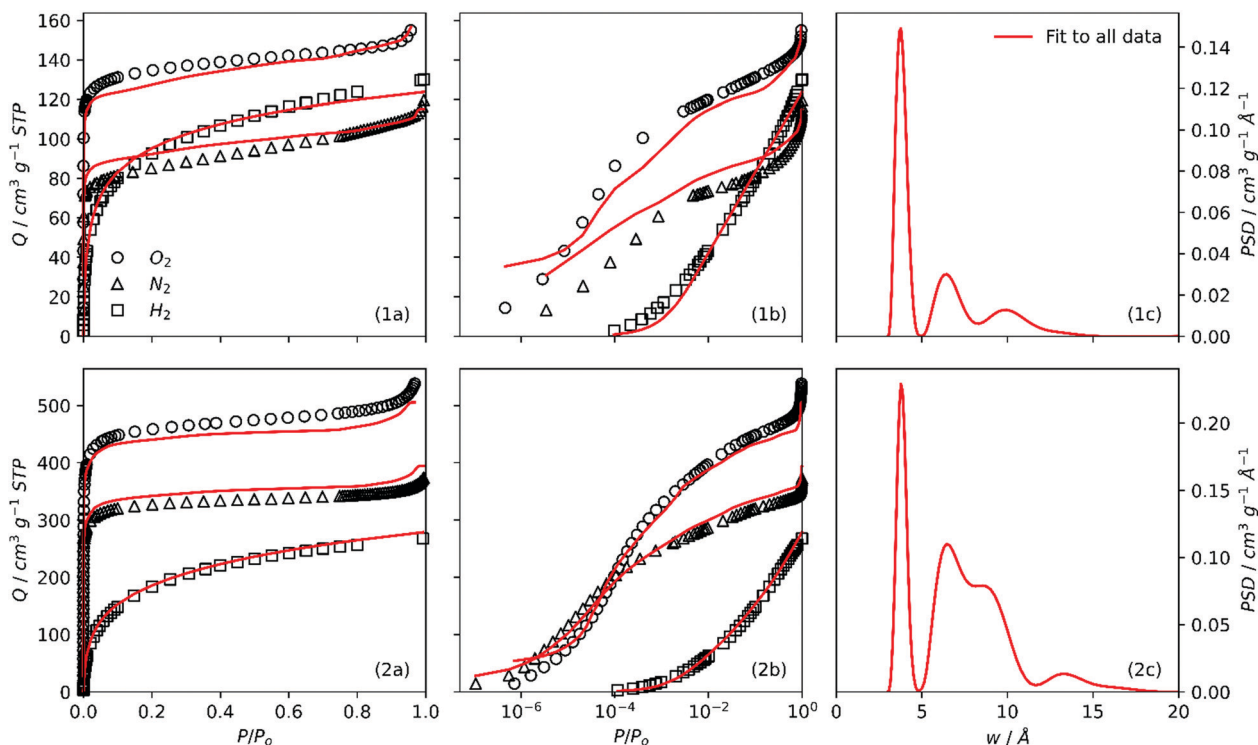


Fig. 5 Simultaneous fit of  $O_2$ ,  $N_2$ , and  $H_2$  isotherms of SA0.0-300 (1a, 1b) and SA1.0-300 (2a, 2b) using x-axis in linear (1a, 2a) and logarithmic (1b, 2b) scales, as well as PSDs determined from the three-way fit (1c, 2c). Fits determined using same parameters as in dual fits, i.e. identical upper and lower pore limits, and  $\lambda = 4$ .

either of the dual isotherm analyses (Fig. S3(1f, 2f), ESI†) the same is not apparent for SA0.0-300.

To summarise, while there is consistency between  $O_2$  and  $N_2$  fits for more highly activated samples, this is not true for less-activated, so-called biochar samples. In other words when probing ultramicropores,  $O_2$  and  $N_2$  produce different PSD results, which can be attributed to the poor diffusion of  $N_2$  into smaller pores. As porosity broadens into the supermicropore region, pores can accurately be measured by either adsorptive.

## 4. Conclusions

This study compares porosity in biochars and carbons activated using alkali metals as determined by two methods of dual isotherm analysis, namely using the adsorptive pairs  $N_2/H_2$  and  $O_2/H_2$ . In particular, the reliability of using  $N_2$  in such analyses is called into question due to evident diffusion limitations in carbon samples with a low degree of activation (biochars) or where the smallest pores (ultramicropores) are prevalent. As these diffusion limitations are not present for carbons activated to a higher degree using  $Na^+$  or  $KOH$ , we conclude that the problems are associated with the size of pores found in biochars. Indeed, when using  $O_2$  (which penetrates these pores with much less difficulty) PSDs derived using the dual isotherm 2D-NLDFT method are centred at widths below 7 Å, i.e. in the

ultramicropore region. Theoretically, improved ultramicropore penetration by  $N_2$  could be achieved either by allowing the system to equilibrate for longer times at low pressure, or by measuring the isotherm at a higher temperature. These modifications would however be impractical as in the first case the duration of isotherm measurements would be excessive. As for increasing temperature, this would require use of a different cryogen and/or higher pressures for 'total' isotherm measurement. This proves impractical on current, standard porosimetric sorptometers while the use of  $O_2$  and  $H_2$  requires the trivial task of attaching an additional cylinder to the sorptometer.

Furthermore, using  $O_2$  and  $H_2$  gives a more realistic picture of porosity development with increasing degree of oxidative chemical activation as carbons are activated with an increasing amount of porogen. That is, as the ratio of porogen/C is increased we consistently observe a broadening in the PSD for samples derived both from sodium carboxymethyl cellulose and sawdust. This is in contrast to results derived using  $N_2$  and  $H_2$  wherein despite increases in overall porosity with increasing porogen/C ratio, there is little change in the widths of pores at higher levels of activation. Indeed, for some sample sets a contraction in pore width is observed as the amount of activating agent is increased. The narrative given by PSDs derived using  $N_2$  and  $H_2$  cannot be accounted for by current understanding of mechanisms of porosity development during oxidative chemical activation. The subtle effects of increasing

amounts of porogen on the fine pore structure of these carbons is only detectable by the highly sensitive methods used in this work. Such sensitivity means that it is paramount to select the appropriate adsorptives for the material.

Finally, attempts at fitting the 2D-NLDFT heterogeneous surface kernels simultaneously to all three isotherms gives very poor fits to the O<sub>2</sub> and N<sub>2</sub> isotherms for biochars. In theory, due to the similar size of these adsorptives, fitting these two isotherms together should be facile. This lends further credence to the notion that there is in fact an inconsistency in the porosity which is probed by these two adsorptives. As these problems are not present for more highly activated samples, it is clear that above a certain pore width, O<sub>2</sub> and N<sub>2</sub> are equally able to probe the porosity of such carbons. In other words O<sub>2</sub> penetrates pores in the ultramicropore region to a much greater extent.

Thus, we suggest that for carbons with a low degree of activation, N<sub>2</sub> is an inadequate probing gas for determination of porosity as it does not adequately probe ultramicropores. When determining PSDs and other textural quantities for such carbons it is therefore advisable to look to non-traditional adsorptives such as O<sub>2</sub>. This logic should also be extended to the use of adsorptives when attempting to understand the development of porosity in carbons as the amount of activating agent is increased. Further, the understanding of porosity in non-carbonaceous ultramicroporous materials (such as CMPs) or carbonaceous materials with richer surface functionality may benefit from a similar analysis.

## Author contributions

L. Scott Blankenship: conceptualisation, methodology, formal analysis, investigation, writing – original draft, visualization. Jacek Jagiello: methodology, software, validation, formal analysis, writing – review & editing. Robert Mokaya: resources, writing – review & editing, supervision, funding acquisition.

## Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

## Acknowledgements

We are thankful to the EPSRC (Low-Dimensional Materials & Interfaces DTP) for a studentship for LSB. RM thanks the Royal Society for a Royal Society Wolfson Research Merit Award.

## Notes and references

- 1 M. Thommes, K. Kaneko, A. V. Neimark, J. P. Olivier, F. Rodriguez-Reinoso, J. Rouquerol and K. S. Sing, *Pure Appl. Chem.*, 2015, **87**, 1051–1069.
- 2 J. Jagiello, J. Kenvin, C. O. Ania, J. B. Parra, A. Celzard and V. Fierro, *Carbon*, 2020, **160**, 164–175.
- 3 J. Jagiello and M. Thommes, *Carbon*, 2004, **42**, 1227–1232.
- 4 J. Jagiello, C. Ania, J. B. Parra and C. Cook, *Carbon*, 2015, **91**, 330–337.
- 5 J. Jagiello, J. Kenvin, A. Celzard and V. Fierro, *Carbon*, 2019, **144**, 206–215.
- 6 F. Rodriguez-Reinoso, J. de, D. Lopez-Gonzalez and C. Berenguer, *Carbon*, 1982, **20**, 513–518.
- 7 L. Qin, G.-j. Xu, C. Yao and Y.-h. Xu, *Chem. Commun.*, 2016, **52**, 12602–12605.
- 8 J. Garrido, A. Linares-Solano, J. Martin-Martinez, M. Molina-Sabio, F. Rodriguez-Reinoso and R. Torregrosa, *Langmuir*, 1987, **3**, 76–81.
- 9 A. L. McClellan and H. F. Harnsberger, *J. Colloid Interface Sci.*, 1967, **23**, 577–599.
- 10 A. D. Lueking, H.-Y. Kim, J. Jagiello, K. Bancroft, J. K. Johnson and M. W. Cole, *J. Low Temp. Phys.*, 2009, **157**, 410–428.
- 11 J. Jagiello and W. Betz, *Microporous Mesoporous Mater.*, 2008, **108**, 117–122.
- 12 M. Konstantakou, A. Gotzias, M. Kainourgiakis, A. K. Stubos and T. A. Steriotis, in *Applications of Monte Carlo Method in Science and Engineering*, ed. S. Mordechai, IntechOpen, 2011.
- 13 A. D. Buckingham, R. L. Disch and D. A. Dunmur, *J. Am. Chem. Soc.*, 1968, **90**, 3104–3107.
- 14 C. Nguyen and D. D. Do, *J. Phys. Chem. B*, 1999, **103**, 6900–6908.
- 15 M. Lopez-Ramon, J. Jagiello, T. Bandoz and N. Seaton, *Langmuir*, 1997, **13**, 4435–4445.
- 16 J. Jagiello and J. P. Olivier, *Carbon*, 2013, **55**, 70–80.
- 17 J. Guo, J. R. Morris, Y. Ihm, C. I. Contescu, N. C. Gallego, G. Duscher, S. J. Pennycook and M. F. Chisholm, *Small*, 2012, **8**, 3283–3288.
- 18 A. Beda, C. Vaultot and C. M. Ghimbeu, *J. Mater. Chem. A*, 2021, **9**, 937–943.
- 19 M. Seredych, J. Jagiello and T. J. Bandoz, *Carbon*, 2014, **74**, 207–217.
- 20 D. Lide, *Handbook of Chemistry and Physics*, CRC Press, 88th edn, 2017.
- 21 D. W. Breck, *Zeolite molecular sieves*, Wiley, New York, 1974.
- 22 B. E. Poling, J. M. Prausnitz and J. P. O'connell, *Properties of gases and liquids*, McGraw-Hill Education, 2001.
- 23 C. Graham, D. A. Imrie and R. E. Raab, *Mol. Phys.*, 1998, **93**, 49–56.
- 24 K. Weber and P. Quicker, *Fuel*, 2018, **217**, 240–261.
- 25 A. Downie, A. Crosky and P. Munroe, in *Biochar for Environmental Management: Science and Technology*, ed. J. Lehman and S. Joseph, 2009, ch. 2, p. 14.
- 26 M. Sevilla and R. Mokaya, *Energy Environ. Sci.*, 2014, **7**, 1250–1280.
- 27 Y. Yamashita and K. Ouchi, *Carbon*, 1982, **20**, 41–45.
- 28 J. F. González, J. M. Encinar, C. M. González-García, E. Sabio, A. Ramiro, J. L. Canito and J. Gañán, *Appl. Surf. Sci.*, 2006, **252**, 5999–6004.



- 29 J. Ganan, J. Gonzalez, C. Gonzalez-Garcia, A. Ramiro, E. Sabio and S. Roman, *Appl. Surf. Sci.*, 2006, **252**, 5988–5992.
- 30 I. Susanti and N. Widiastuti, *Mal. J. Fundam. Appl. Sci.*, 2019, **15**, 249–253.
- 31 X. Fang, G. Li, J. Li, H. Jin, J. Li, V. Jegatheesan, S. Li, H. Wang and M. Yang, *Desalin. Water Treat.*, 2017, **96**, 120–127.
- 32 J. E. Park, G. B. Lee, S. Y. Hwang, J. H. Kim, B. U. Hong, H. Kim and S. Kim, *Appl. Sci.*, 2018, **8**, 1596.
- 33 M. Lillo-Ródenas, J. Juan-Juan, D. Cazorla-Amorós and A. Linares-Solano, *Carbon*, 2004, **42**, 1371–1375.
- 34 M. Lillo-Ródenas, D. Cazorla-Amorós and A. Linares-Solano, *Carbon*, 2003, **41**, 267–275.
- 35 E. Raymundo-Pinero, P. Azaïs, T. Cacciaguerra, D. Cazorla-Amorós, A. Linares-Solano and F. Béguin, *Carbon*, 2005, **43**, 786–795.
- 36 R. Xue and Z. Shen, *Carbon*, 2003, **41**, 1862–1864.
- 37 J. Jagiello, *Langmuir*, 1994, **10**, 2778–2785.
- 38 P. Hansen, in *Computational Inverse Problems in Electrocardiology*, ed. P. Johnston, WIT Press, Southampton, 2001, pp. 119–142.
- 39 P. C. Hansen and D. P. O'Leary, *SIAM J. Sci. Comput.*, 1993, **14**, 1487–1503.
- 40 P. C. Hansen, *SIAM Rev.*, 1992, **34**, 561–580.
- 41 J. D. Evans, V. Bon, I. Senkovska and S. Kaskel, *Langmuir*, 2021, **37**, 4222–4226.
- 42 M. Sevilla and A. B. Fuertes, *Chem. – Eur. J.*, 2009, **15**, 4195–4203.
- 43 A. E. Ogungbenro, D. V. Quang, K. A. Al-Ali, L. F. Vega and M. R. M. Abu-Zahra, *J. Environ. Chem. Eng.*, 2018, **6**, 4245–4252.
- 44 A. M. Aljumaly and R. Mokaya, *Mater. Adv.*, 2020, **1**, 3267–3280.
- 45 N. Balahmar and R. Mokaya, *J. Mater. Chem. A*, 2019, **7**, 17466–17479.
- 46 E. Evstigneev, *Russ. J. Appl. Chem.*, 2011, **84**, 1040–1045.
- 47 K. Latham, G. Jambu, S. Joseph and S. Donne, *ACS Sustainable Chem. Eng.*, 2014, **2**, 755–764.

