



Cite this: *Mater. Adv.*, 2022, 3, 3218

Alkylthio-based asymmetric liquid crystals: unravelling the substituent effects and intercalated cybotactic nematic and smectic phases†

Yuki Arakawa, ^a Yuko Ishida, ^a Yukito Sasaki, ^a Shunsuke Sasaki, ^b Masatoshi Tokita ^c and Hideto Tsuji ^a

To reveal not only the effects of alkylthio groups but also an unknown layered smectic (Sm) phase formed by short-chain-based liquid crystals (LCs), a novel series of asymmetric 1,4-phenylene bis(benzoate)s possessing 4,4'-alkylthio and alkoxy groups with various chain lengths (SO_n, $n = 1-8$) was established. All members formed well-defined enantiotropic LC phases. The SO_n homologs with $n = 2-7$ exhibited only a nematic (N) phase, and the longest SO₈ showed an inclined smectic C (SmC) phase in addition to the N phase. Notably, the shortest SO₁ formed an enantiotropic orthogonal smectic A (SmA) phase as well as the N phase. The primary effects of the alkylthio and alkoxy groups were identified for each phase transition over varied chain lengths and through comparison with those of their symmetric alkoxy- and alkylthio-based counterparts (OOn and SSn, respectively). In addition, X-ray diffractometry unraveled surprising results, wherein the N and SmA phases of SO₁ were found to possess a layer period that was half of the molecular length, thereby suggesting half-intercalation of the intermolecular cores. These phases are distinct from the conventional N and Sm phases without such core intercalation. The remarkable formation of half-intercalated LC phase was discussed based on plausible intermolecular interactions (e.g., multiple hydrogen bonds due to the presence of esters) implied by single-crystal structural analyses. Furthermore, the birefringence increased when replacing the alkoxy group with the more polarizable alkylthio group.

Received 18th January 2022,
Accepted 23rd February 2022

DOI: 10.1039/d2ma00050d

rsc.li/materials-advances

1. Introduction

Liquid crystals (LCs), which are anisotropic and fluidic mesophases with properties between three-dimensionally ordered solid crystals and isotropic liquids, have enabled the widespread development of LC displays (LCDs) in the last few decades. Moreover, the self-assembly and stimuli-responsive properties of LCs have attracted much attention from various interdisciplinary fields beyond those involving LCDs. The fundamental properties of LCs, including their phase transition behaviors as well as their optical and electronic properties,

primarily depend on their mesogenic core structures. Typically, chain-containing substituents such as alkyl and alkoxy groups are appended to the mesogenic core structures to effectively induce LC phases and confer flexibility onto the LC molecules, thereby improving their molecular affinity and miscibility while also lowering their phase transition temperatures. In addition, the elements present in these substituents and their bonds connecting to the mesogenic cores (such as methylene and oxygen for alkyl and ether, respectively), as well as their alkyl chain lengths, also significantly affect the properties of LCs and are therefore crucial in related molecular design.

In addition to conventional alkyl and alkoxy groups, alkylthio (*i.e.*, alkylsulfanyl or thioether: R-S-R) groups are also useful in LC materials. Because the sulfur atom is heavier and more electron rich than the commonly used carbon and oxygen atoms, it endows LC materials with a higher polarizability, larger dispersion forces, and elevated highest occupied molecular orbitals (HOMO). Thus, alkylthio-based LCs are used as high-birefringence materials¹⁻⁸ and organic LC semiconductors with high charge mobilities.⁹⁻¹⁴ In particular, the alkylthio group is often employed in the molecular design of discotic

^a Department of Applied Chemistry and Life Science, Graduate School of Engineering, Toyohashi University of Technology, 1-1 Hibarigaoka, Tempaku-cho, Toyohashi, Aichi 441-8580, Japan. E-mail: arakawa@tut.jp

^b Université de Nantes, CNRS, Institut des Matériaux Jean Rouxel, IMN, F-44000 Nantes, France

^c Department of Chemical Science and Engineering, School of Materials and Chemical Technology, Tokyo Institute of Technology, 2-12-1, Ookayama, Meguro-ku, Tokyo 152-8550, Japan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d2ma00050d



LCs,^{9,12–26} since the intermolecular interactions of the more polarizable sulfur atoms help induce and stabilize the columnar (Col) phases in discotic molecules.^{22,23} In addition, the alkylthio group has greater flexibility and bulkiness, in addition to a lower structural anisotropy, compared with the alkyl and alkoxy groups. These properties arise from the longer C–S bond lengths (~ 1.7 Å) and the smaller C–S–C bond angles ($\sim 100^\circ$) than those of alkyl and alkoxy groups (e.g., ~ 1.5 and ~ 1.4 Å for the C–C and C–O bond lengths; and $\sim 110^\circ$ and 118° for the C–CH₂–C and C–O–C bond angles, respectively). Such steric effects of the alkylthio groups have benefits in inducing not only the above-mentioned Col phase,^{22,23} but also the cubic phase,²⁷ which could effectively fill large spaces around the columns and gyroids in terms of free volume effects.²⁸

In contrast, the greater flexibility and steric bulkiness imparted by the C–S–C bonds prevent rod-like molecules from rearranging into LC phases and reduce the LC temperature ranges. As a result, compared to their alkyl- and alkoxy-based counterparts, alkylthio-containing rod-like molecules usually exhibit LC phases monotropically, only across a narrow temperature range, or not at all.^{1–5,29,34–36} On the other hand, enantiotropic LC phases were observed from some types of alkylthio-based calamitic molecules.^{4–8,29–40} However, there has been no report of covalent-bonded homologous series in which all members exhibit enantiotropic LC phases, while only one such hydrogen-bonded homologous series is known.³⁵ The lack of dataset of more versatile covalent-bonded homologous series showing the enantiotropic LC phases allowed us only to infer the effects of alkylthio groups on LC behaviours based on few of its homologous series. For example, as typical LC motifs, 1,4-phenylene bis(benzoate) (PBB) homologs^{41–48} possessing symmetric alkoxy-groups (OOn, Fig. 1) are known to display

well-defined enantiotropic LC phases.^{41–46} In addition, we reported the phase transition behavior of PBB homologs with symmetric alkylthio groups (SSn, Fig. 1).³⁴ SSn with $n = 1, 2$, or 6 exhibit enantiotropic LC phases with narrower temperature ranges, whereas those with $n = 3–5, 7$, or 8 show monotropic N phases. Thus, SSn typically exhibits poor mesogenic ability compared to OOn. In the paper,³⁴ we noted that the shortest chain SS1 homolog forms layered Sm phase, which is unusual and counter to the general trend where longer chains facilitate the formation of layered Sm phases due to enhanced microphase separation. However, because of its elusive monotropic nature with a narrow temperature range, the structure of this abnormal Sm phase for SS1 could not be further investigated. Moreover, although such Sm phases were also reported for other molecules with short chains,^{44,45,49} their structures have also remained unknown to date.

To realize a homologous series in which all members exhibit well-defined enantiotropic LC phases, an alkylthio-appended highly anisotropic structure such as PBB should be combined with an asymmetric substituent such as an alkoxy and alkyl group at the other end. The asymmetric substitution of alkylthio and alkoxy groups is expected to expand the LC temperature ranges beyond those found in the symmetric SSn homologs, enabling us to investigate not only the properties of alkylthio-based LCs but also the enigmatic Sm phases formed by the short-chain molecules. In addition, this raises the questions how the alkylthio and alkoxy (or alkyl) groups affect the overall phase transition temperature, and which structure has a dominant effect when multiple structures coexist in molecular homologs. Currently, there is only one published report on alkylthio- and alkyl-substituted asymmetric diphenyl-acetylene-based LC homologs, some of which show only monotropic LC phases.³³ To the best of our knowledge, however, no asymmetric LC homologous series bearing alkylthio and alkoxy (or alkyl) groups in which all members exhibit LC phases have been established. Therefore, the competing substituent effects have yet to be revealed for such homologous series.

In these contexts, we herein report a novel homologous series of asymmetric PBB homologs possessing alkylthio and alkoxy groups with various alkyl chain lengths (C_nH_{2n+1}) at the 4,4'-positions (SO_n, $n = 1–8$, Fig. 1). Their phase transition behavior, mesophase structures, and birefringence properties are investigated⁵⁰ and compared with those of previously reported symmetric alkoxy- and alkylthio-based homologs (OOn and SSn, respectively, Fig. 1). In addition, competing substituent effects on the LC properties between the alkylthio and alkoxy groups are discussed in detail upon varying the chain lengths. Furthermore, the structure of the unknown Sm phase is revealed for short-chain molecules.

2. Experimental

2.1. Synthesis

All SO_n homologs were prepared as outlined in Scheme S1, using the general procedures described in the ESI.† The

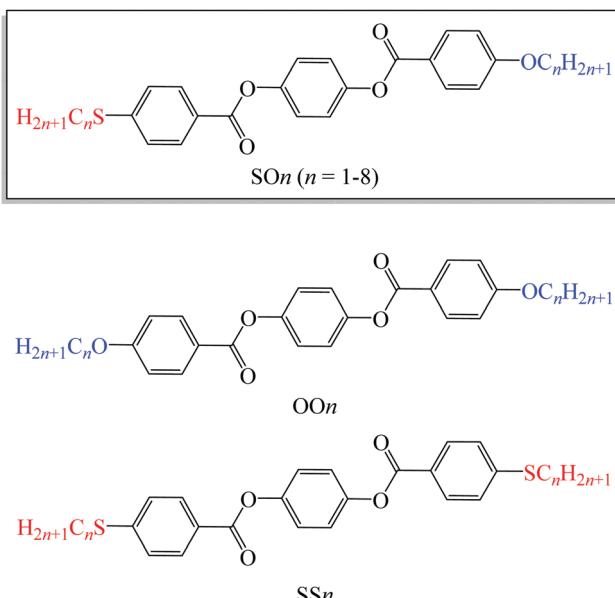


Fig. 1 Molecular structures of LC homologs bearing alkylthio and alkoxy groups (SO_n) and previously reported alkoxy- (OOn) and alkylthio-based (SSn) LC homologs.



molecular structures were determined using ^1H and ^{13}C nuclear magnetic resonance (NMR) spectroscopy on a JNM-ECX500 (500 MHz for ^1H and 126 MHz for ^{13}C) or a JNM-ECS400 (400 MHz for ^1H and 100 MHz for ^{13}C) spectrometer (JEOL Ltd., Tokyo, Japan). All NMR spectra were obtained using solutions of the samples in deuterated chloroform with a tetramethylsilane internal standard (Kanto Chemical Co., Tokyo, Japan). High-resolution mass spectrometry (HRMS) was also conducted using high-performance liquid chromatography (HPLC) (Agilent 1200 HPLC-Chip and 6520 Accurate-Mass Q-TOF). The molecular characterization data are provided in the ESI.†

2.2. Phase identification and thermal properties

Phase identification and analysis of the phase transition behavior were performed using polarized-light optical microscopy (POM; Olympus BX50) with a Linkam LK-600PM hot stage. The phase transition temperatures and the associated enthalpy changes were determined using differential scanning calorimetry (DSC, Shimadzu DSC-60 instrument). The heating and cooling cycles were performed at a rate of $3\text{ }^\circ\text{C min}^{-1}$ under an N_2 flow (50 mL min^{-1}). In addition, thermal stability of SO6 was evaluated by thermogravimetry analysis (TGA, Shimadzu, TGA-50 instrument) at a rate of $3\text{ }^\circ\text{C min}^{-1}$ under an N_2 flow (50 mL min^{-1}). All POM images and DSC and TGA curves, except those presented in the main text, are shown in the ESI.†

2.3. X-ray diffraction measurements

The mesophase structures of SO_n ($n = 1, 6$, and 8) were investigated using X-ray diffractometry (XRD, Bruker D8 DISCOVER diffractometer) with a $\text{Cu K}\alpha$ X-ray source and a Vantec-500 detector. The specimens were stored in glass capillary tubes (diameter: 1.5 mm; WJM-Glass Müller GmbH) and sandwiched between a pair of permanent magnets in a folder. After heating to the isotropic (Iso) phase temperature to erase the thermal histories of the samples, the measurements were carried out at different mesophase temperatures upon cooling.

2.4. Birefringence measurements

The birefringence (Δn) of SO6 was measured according to our previous study⁵¹ for comparison with those of OO6 and SS6.³⁴ SO6 was inserted into a uniaxially rubbed planer alignment cell (with a polyimide surface and a thickness of $5\text{ }\mu\text{m}$, EHC Co., Tokyo, Japan) by capillary action at its Iso temperature (controlled by a Mettler FP82HT hot stage). After filling the cell with SO6, the uniaxial alignment state of the N phase was confirmed by POM using a Nikon LV100 Pol optical microscope (Tokyo, Japan). Transmittance through the nematic SO6 medium with the N director at 45° under the cross-Nicol condition was measured using a USB4000 spectrometer (Ocean Insight, Orlando, FL, USA) through the Nikon LV100 Pol optical microscope. The transmittance was obtained upon cooling from the Iso phase temperature, while the temperature was controlled using a Mettler FP82HT hot stage. The transmittance (T) was fitted to eqn (1) along with Cauchy's equation, eqn (2):

$$T = A \sin^2\left(\frac{\pi d \Delta n}{\lambda}\right), \quad (1)$$

$$\Delta n = a + \frac{b}{\lambda^2} + \frac{c}{\lambda^4} + \frac{d}{\lambda^6}, \quad (2)$$

where d and λ denote the cell gap and the wavelength of light, respectively. The experimental transmittance spectra were well fitted to eqn (1) with Cauchy's coefficients (a , b , c , and d), thereby allowing the wavelength dispersion of the birefringence to be estimated.

3. Results and discussion

3.1. Phase transition behavior

3.1.1. Phase transitions of the SO_n homologs. The phase transition data for SO_n and their thermal phase sequence as a function of n are summarized in Table 1 and Fig. 2, respectively. All asymmetric SO_n homologs exhibited well-defined enantiotropic LC phases. This is the first homologous series of alkylthio-containing calamitic LC molecules with all members exhibiting enantiotropic LC phases. Considering that

Table 1 Phase transition of SO_n upon the 2nd heating (upper line) and 1st cooling (bottom line) at a rate of $3\text{ }^\circ\text{C min}^{-1}$

| n | Phase | $T/^\circ\text{C}$ | $\Delta H/\text{kJ mol}^{-1}$ | Phase | $T/^\circ\text{C}$ | $\Delta H/\text{kJ mol}^{-1}$ | Phase | $T/^\circ\text{C}$ | $\Delta H/\text{kJ mol}^{-1}$ | Phase |
|-----|-------|--------------------|-------------------------------|-------|--------------------|-------------------------------|-------|--------------------|-------------------------------|-------|
| 1 | Cr | 170.3 | 18.1 | SmA | 194.6 | 4.2 | N | 283.0 | 1.5 | Iso |
| | | 153.7 | 12.7 | | 193.6 | 4.2 | | 281.5 | 1.6 | |
| 2 | Cr | 216.8 | 40.2 | — | — | — | N | 259.3 | 1.7 | Iso |
| | | 207.2 | 40.7 | | | | | 257.0 | 1.8 | |
| 3 | Cr | 181.9 | 31.2 | — | — | — | N | 214.6 | 1.4 | Iso |
| | | 177.3 | 30.6 | | | | | 213.9 | 1.4 | |
| 4 | Cr | 158.4 | 32.9 | — | — | — | N | 210.0 | 1.9 | Iso |
| | | 153.9 | 33.0 | | | | | 209.0 | 1.9 | |
| 5 | Cr | 146.6 | 33.6 | — | — | — | N | 189.0 | 1.3 | Iso |
| | | 140.2 | 33.7 | | | | | 188.2 | 1.3 | |
| 6 | Cr | 130.3 | 30.2 | — | — | — | N | 185.1 | 1.6 | Iso |
| | | 124.2 | 29.4 | | | | | 184.1 | 1.6 | |
| 7 | Cr | 129.0 | 34.6 | — | — | — | N | 174.2 | 1.3 | Iso |
| | | 123.3 | 34.2 | | | | | 173.4 | 1.4 | |
| 8 | Cr | 122.9 | 31.3 | SmC | 127.5 | 2.4 | N | 171.6 | 1.7 | Iso |
| | | 113.5 | 30.0 | | 126.9 | 2.4 | | 170.6 | 1.7 | |



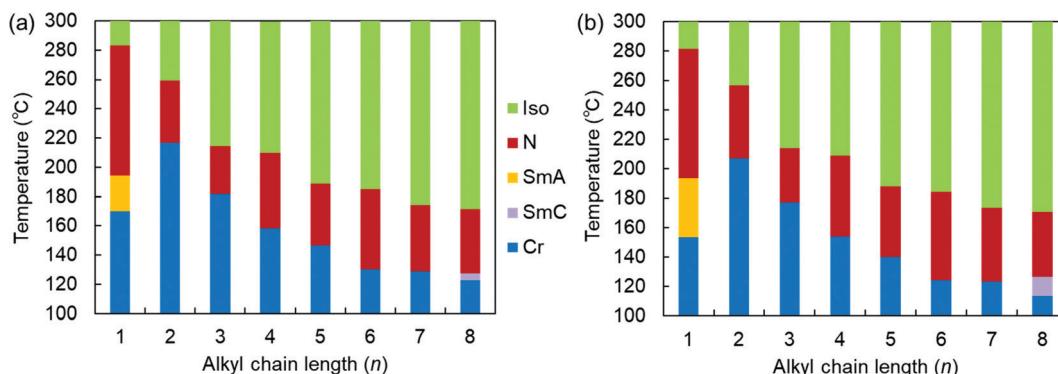


Fig. 2 Phase sequences of the SOn homologs as a function of n upon (a) heating and (b) cooling at a rate of $3\text{ }^{\circ}\text{C min}^{-1}$.

symmetric SS_n homologs³⁴ and other alkylthio-based molecules with relatively small cores exhibited either no LC phase or narrow LC phases,^{1–5,29} the excellent LC tendency of the SOn homologs is ascribed to the synergistic effects of the asymmetric substituents and the highly anisotropic core structure of 1,4-phenylene bis(benzoate).

More specifically, SOn with $n = 2\text{--}7$ exhibited only enantiotropic N phases, whereas the longest-chain homolog, SO8, exhibited an enantiotropic tilted layered smectic C (SmC) phase as well as the enantiotropic N phase. The formation of SmC phase is a consequence of microphase separation between the aliphatic chains and the aromatic cores, which is facilitated by longer alkyl chains. POM observations of SO8 upon cooling revealed a streak-like second schlieren texture in the N phase [Fig. S8(b) in the ESI[†]], which is typical of an N_{Cyc} phase containing SmC-like clusters (*i.e.*, a $\text{N}_{\text{Cyc}}\text{C}$ phase).^{52–54} These phases were also found for the SS6 and OO6 homologs.³⁴ During the transition from the N to the SmC phase upon cooling, a transition bar was observed and a less fluid schlieren texture with four-brush singularity appeared, indicating a homeotropic alignment in the tilted SmC phase [Fig. S8(c) and (d), ESI[†]]. This schlieren texture in the SmC phase partly appeared along the streak lines in the N_{Cyc} phase. The enantiotropic SmC-N phase transition of SO8 is considered a first-order transition, as shown in the DSC curves (Fig. S15, ESI[†]).

Remarkably, the shortest-chain homolog (SO1) enantiotropically exhibited not only an N phase but also a layered orthogonal SmA phase [Fig. 3(a) and (b), respectively]. Fig. 3(b) shows a fan-shaped texture typical to Sm phases, and this texture grew from a bâtonnet texture. The formation of a layered SmA phase in SO1 is unusual, because the formation of Sm layered structures is typically facilitated by long alkyl chains due to enhanced microphase separation, as exemplified by SO8. Nevertheless, similar abnormal Sm phase formation was previously reported for other PBB analogs bearing short methyl and ethyl chains (abbreviated as PBB-1 and PBB-2 in the present study)^{44,45} and PBB-S1 (namely SS1 in the present paper).³⁴ These phases were reported to be monotropic SmA phases in very narrow temperature ranges according to POM observations. In contrast, the present SO1 homolog exhibited

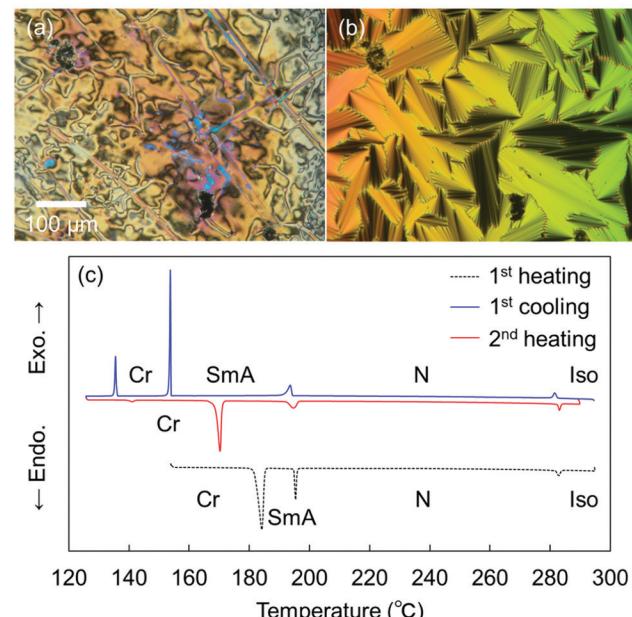


Fig. 3 POM images of SO1 (a) in the nematic (N) phase at $195\text{ }^{\circ}\text{C}$ and (b) in the SmA phase at $190\text{ }^{\circ}\text{C}$, upon cooling, and (c) DSC curves of SO1 collected at a rate of $3\text{ }^{\circ}\text{C min}^{-1}$.

an enantiotropic SmA phase within a practical temperature range, and so it was further investigated by XRD (see Section 3.2.2). The DSC curves [Fig. 3(c)] show that the SmA-N transition of SO1 is strongly first order, with an enthalpy change of $\Delta H = 4.2\text{ kJ mol}^{-1}$, which is similar to those of previously reported analogues PBB-1 and SS1.^{34,44} This first-order transition likely differs from the typical second-order nature of SmA-N transitions. SO1 exhibited different melting temperature (T_m) values during the 1st and 2nd heating [184.1 and $170.3\text{ }^{\circ}\text{C}$, respectively, Fig. 3(c)] due to the presence of crystalline polymorphs; the latter is listed in Table 1. Furthermore, because of the high-temperature LC phase over $250\text{ }^{\circ}\text{C}$, thermal stability of SO1 was examined by TGA, and the obtained TGA curve is shown in Fig. S16 (ESI[†]). Slight weight loss or thermal decomposition started over $250\text{ }^{\circ}\text{C}$ and the weight loss at $290\text{ }^{\circ}\text{C}$ was within 5%. Slight thermal



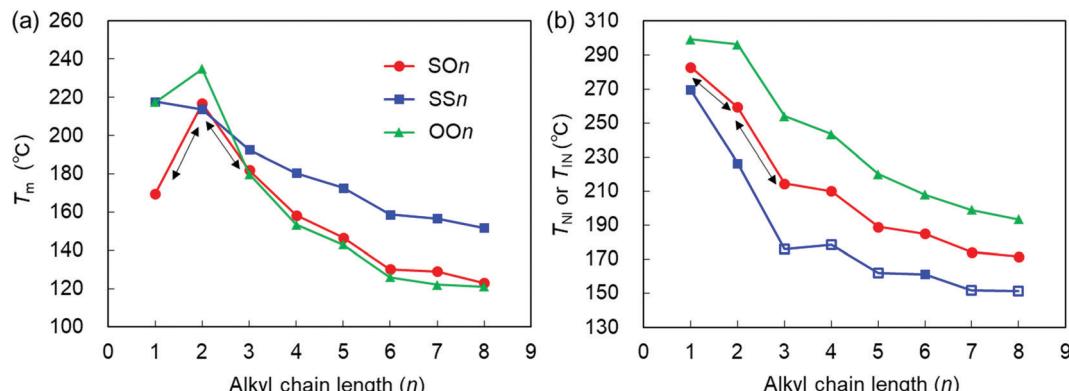


Fig. 4 (a) T_m and (b) T_{NI} (filled symbols) or T_{IN} (open symbols) of the $SOOn$, SSn , and $OOOn$ homologs ($n = 1-8$) upon heating ($10\text{ }^{\circ}\text{C min}^{-1}$) as a function of n . The data for $OOOn^{45}$ and SSn^{34} are drawn from the literature.

decomposition of the sample must have influenced the phase transition behavior and related analyses but these could not be critical.

3.1.2. Comparison of the $SOOn$, $OOOn$, and SSn homologs.

We compared the phase transition behavior of $SOOn$ with those of previously reported symmetric alkoxy-based $OOOn$ and alkylthio-based SSn homologs.^{34,45} The T_m and N-Iso phase transition temperatures (T_{NI} or T_{IN} upon heating and cooling, respectively) were plotted as a function of n (Fig. 4). As can be seen in Fig. 4(a), the T_m values of the alkoxy-containing $OOOn$ and $SOOn$ homologs initially increased upon increasing n from 1 to 2, before decreasing as n further increased from 2 to 8. This increase of T_m from $n = 1$ to 2 is associated with the conformationally different molecular geometries, both of which depend on the parity of n . An alkoxy group with an even n value is mostly aligned along the long axis of the molecular core, whereas that with an odd n value tends to deviate from this axis.^{55,56} Thus, the molecular anisotropy is higher for even- n than odd- n members. This odd–even effect on the conformational anisotropy can render T_m higher for $n = 2$ than for $n = 1$. A gradual reduction in the T_m values of $OOOn$ and $SOOn$ upon increasing n from 2 to 8 is attributed to dilution of the polarizable aromatic core moieties by the longer alkyl chains. However, it is interesting to note that the T_m of SSn decreased upon increasing n from 1 to 2, which is a different trend from those observed in $OOOn$ and $SOOn$. The SSn homologs possess two symmetric alkylthio groups, which bring larger London dispersion forces to enhance the T_m of $SS1$. However, this effect could be cancelled out by the smaller anisotropy, larger flexibility, bulkiness effects and, at least in part, the weaker $S\cdots H$ hydrogen bonds (due to the less electronegativity of S than O) and weaker electron donor property of the methylthio group compared to the methoxy group with strong electron donor property, for example. When n is increased to 2, the enhanced and significant conformational flexibility and steric bulkiness effects of the alkylthio groups could facilitate crystal melting for $SS2$ and lower its T_m below that of $SS1$, which in turn masks the steric odd–even n effect on the phase transition temperatures for short alkylthio-containing homologs ($n = 1-3$).³⁶ Consequently, the T_m of SSn gradually and monotonically decreases when n is increased from 1 to 8.

Here, it should be emphasized that the T_m values of the asymmetric $SOOn$ homologs ($n = 1-3$) exhibit a typical odd–even oscillation, as represented by the black arrows in Fig. 4(a). Interestingly, this odd–even oscillation of $SOOn$ is ascribed to the parity of the alkoxy chains at the other end. This provides an insight into the competing effects between the alkylthio and alkoxy groups on the value of T_m . In short-chain homologs ($n = 1-3$), the conformational flexibility and steric effects of the alkylthio group at one end did not outweigh the effect of the alkoxy group at the other end of the molecule, which indicates that molecular shape anisotropy also significantly affected the T_m and elevated that of $SO2$. The T_m values of the homologs with $n > 3$ decreased in the order $SSn > SOOn > OOOn$. In other words, T_m increases upon replacing the alkoxy group with an alkylthio group. More specifically, the symmetric alkylthio-based SSn homologs possess significantly higher T_m values than both the $SOOn$ and $OOOn$ homologs, which is ascribed to the stronger dispersion forces brought by the two symmetric, highly polarizable alkylthio groups. This interaction effect becomes more pronounced for longer alkyl chains, which could be assisted by enhanced microphase separation or molecular affinities [Fig. 4(a)],^{36,40} resulting in significantly higher T_m values for molecules containing alkylthio groups than their alkoxy counterparts. Nonetheless, the T_m values of $SOOn$ are only marginally higher than those of $OOOn$, which is ascribable to two competing effects. Although the dispersion forces provided by the alkylthio group elevates T_m , this effect is canceled out by asymmetric substitution or molecular inhomogeneity that tends to suppress it.

The T_{NI} or T_{IN} values of the three homologous series decrease in the order $OOOn > SOOn > SSn$, which implies that they decrease upon substituting the alkoxy group with an alkylthio group [Fig. 4(b)]. This trend seems contradictory to the stronger dispersion force expected for alkylthio groups. The stronger dispersion force of the more polarizable sulfur atom makes a larger contribution to increase the phase transition temperatures than that from the oxygen atom, as observed in the T_m values. Therefore, there must be other factors that explain lowering of T_{NI} or T_{IN} upon replacement the alkoxy with alkylthio groups. One of the possible rationales is the



steric effects of alkylthio groups such as their greater conformational flexibility, bulkiness, and lower molecular shape anisotropy.

Upon increasing n from 1 to 2, the T_{NI} values of SSn , SO_n , and OOn decrease by approximately 43, 22, and 3 °C, respectively, indicating that the magnitude of reduction in T_{NI} increases upon replacing the alkoxy group with the alkylthio group. This is due to the greater influence from the increased intrinsic conformational and steric effects of the alkylthio groups when lengthening the alkyl chains by one carbon atom, compared to the case of alkoxy groups. Overall, the T_{NI} or T_{IN} values of all three series of homologs decreased with increasing n , and their LC phase temperature ranges decreased in the order $OOn > SO_n > SSn$ upon substituting the alkoxy group with the alkylthio group.

Interestingly, the trend of T_{NI} or T_{IN} ($OOn > SO_n > SSn$) is the reverse of that of their T_m values ($SSn > SO_n > OOn$ for $n \geq 3$), as shown in Fig. 4. This result supports the view that the melting point of the solids is more likely to be influenced by the relative polarizability or attractive interactions imparted by the substituents, whereas the T_{NI} or T_{IN} values are primarily determined by the conformational and steric effects of the substituents rather than their attractive interactions. In addition, it should be emphasized that for the short chains ($n = 1-3$), in contrast to the case where the T_m of the asymmetric SO_n homologs exhibit an odd–even oscillation, their T_{NI} values show monotonous changes (*i.e.*, a decrease or increase), as represented by the black arrows in Fig. 4. This suggests that the conformational flexibility and bulkiness effects greatly impact the value of T_{NI} . For example, when alkylthio bonds are under steric constraints (*e.g.*, by incorporation into ring structures), sulfur-based rings lead to higher T_m and T_{NI} or T_{IN} values, in addition to enhanced LC phase stability compared with the oxygen-based counterparts, although they likely exhibit a lesser degree of molecular anisotropy than their ether-based counterparts.⁵⁷ On the other hand, in the case of side-chain LC polymers, the presence of a longer alkylthio spacer renders an alkoxy-based smectogenic azobenzene polymer exclusively nematogenic and drastically decreases its phase transition temperatures.⁴⁰ These facts can also be interpreted by the significant roles of the flexibility and bulkiness imparted by the C–S–C bonds.

Among the homologs bearing longer alkyl chains, the SmC phase was observed in $OO7$ ^{42,45,46} and $SO8$, but not in $SS8$,³⁴ thereby indicating that the alkylthio group hindered the formation of an Sm phase in SSn . The alkylthio group can either induce the formation of a highly ordered phase due to the enhanced intermolecular interactions,^{4,27,40} or hinder the Sm phase formation through the greater conformational and steric effects.^{30,34} These trends are counterbalanced by the anisotropy of the mesogenic core structures and the alkyl chains. For larger anisotropic mesogenic cores such as the present PBBs that tend to form N phases, formation of the Sm phase was found to be inhibited at the tested chain lengths ($n = 1-8$). For relatively small mesogenic core structures, the formation of a highly ordered Sm phase is facilitated by the alkylthio groups.^{4,30,33,40,58,59}

3.2. Liquid crystal phase structures

3.2.1. Longer-chain homologs. We carried out variable-temperature XRD measurements of the N phase in $SO6$ under a magnetic field and compared the results to those of $SS6$ and $OO6$ (using XRD data from our previous work).³⁴ The 2D XRD patterns of the N phases of $SO6$, $SS6$, and $OO6$ (Fig. 5) confirm that the long molecular axes are aligned with the magnetic field [represented by a white arrow in Fig. 5(a)]. The N phase of $SO6$ exhibited a broad wide-angle diffraction ($2\theta \approx 19^\circ$) and dumbbell-shaped split small-angle diffractions ($2\theta \approx 4^\circ$) [Fig. 5(a)]. The latter indicates that the N phase includes short-range tilted SmC-like pseudo-layer clusters (namely the N_{CyBC} phase).^{52,53} Similar N_{CyBC} phases have been reported for $SS6$ and $OO6$ [Fig. 5(b) and (c), respectively].³⁴

Fig. 5(d) plots the d -spacing values from the wide-angle diffractions (d_{WAX} , representing the average lateral intermolecular distance) of $SO6$, $SS6$, and $OO6$ as a function of the shifted temperature, $\Delta T = T_{IN} - T$, where T is the measurement temperature upon cooling from the Iso phase. The d_{WAX} values of all homologs monotonically decreased with decreasing temperature. In both their Iso and N phases, $SO6$ and $SS6$ exhibited similar d_{WAX} values, which were smaller than those of $OO6$. This is consistent with the typical trend that alkylthio-based molecules in the LC and Iso phases usually possess smaller d_{WAX} values than their alkoxy counterparts.^{4,30,34} This trend is ascribable to two factors, namely the larger dispersion forces of the alkylthio groups and the stronger electron-donating

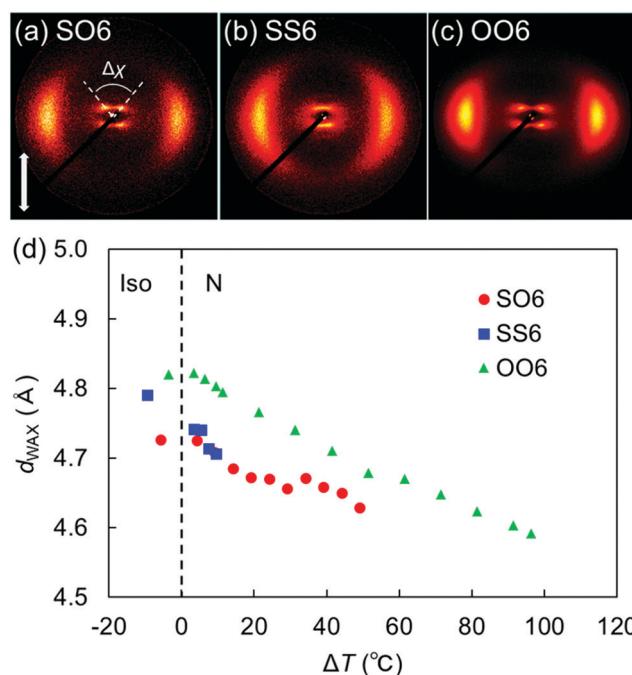


Fig. 5 2D XRD patterns of the N phases of (a) $SO6$, (b) $SS6$, and (c) $OO6$. Panels (b) and (c) were adapted from our previous work.³⁴ (d) Lateral intermolecular distances (d_{WAX}) in $SO6$, $SS6$, and $OO6$ as a function of the shifted temperature ($\Delta T = T_{IN} - T$). The data for $SS6$ and $OO6$ were taken from our previous work.³⁴ Plots for $\Delta T < 0$ denote temperatures in each Iso phase.



properties of the alkoxy groups. The former could lead to smaller lateral molecular distances, whereas the latter might generate repulsive forces between the mesogenic cores, which might in turn lead to larger distances between the alkoxy-based molecules (*i.e.*, OO6) than between the alkylthio-containing molecules (*i.e.*, SO6 and SS6). The dispersion forces from the sulfur atoms in SO6 contribute to its intermolecular packing in random head-to-tail and tail-to-head orientations.³⁶ In addition, the orientational order parameters (S) were estimated for SO6, OO6, and SS6 from their wide-angle diffractions (Fig. S17, ESI[†]) according to the literature.⁶⁰ When compared for a similar value of ΔT , the S values of SO6 were estimated to be higher than those of SS6 and OO6, and that of OO6 was found to be the lowest. The lower S values for SS6 than for SO6 could be due to its significantly narrower N phase temperature range, which is close to its Iso phase region.

The split peak maxima of the small-angle diffraction in the N_{CybC} phase of SO6 were located at $3.7\text{--}4.0^\circ$, and the corresponding d -spacing (d_{SAX}) was ~ 23 Å. In addition, the half values of the angular splitting of the small-angle scattering maxima ($\Delta\chi/2$) were estimated as shown in Fig. 5(a). The $\Delta\chi/2$ value of SO6 in the N_{CybC} phase was estimated to be $\sim 46\text{--}48^\circ$. The apparent molecular length (L) in the N_{Cyb} layers of SO6 was found to be $\sim 33\text{--}34$ Å by $d = L \cos \beta$, assuming d (layer spacing) = d_{SAX} and β (tilt angle) = $\Delta\chi/2$. The estimated L values are similar to the molecular length of SO6 (*i.e.*, ~ 32 Å). This suggests that the $\Delta\chi/2$ of SO6 may reflect the tilt angle in the SmC-like layer clusters, as well as that SO6 molecules would not intercalate among the neighboring SmC-like layer clusters [Fig. 6(d)]. The $\Delta\chi/2$ values in the N_{CybC} phase of SO6 ($\sim 46\text{--}48^\circ$) are between those of SS6 and OO6 ($\sim 40^\circ$ and $\sim 53^\circ$, respectively).³⁴ This means that the $\Delta\chi/2$ values decrease upon

displacing the alkoxy groups with alkylthio groups. This trend can be interpreted by molecular tilting in the SmC-like layer clusters. The larger dispersion forces of the sulfur atoms at the edge of mesogenic cores enhance lateral packing of the mesogenic core moieties, in combination with the dominant $\text{CH}\cdots\pi$ and $\pi\cdots\pi$ interactions among the cores to decrease the tilt angles and be toward non-tilted SmA-like clusters.^{4,30,34,36,40}

Subsequently, the 2D XRD patterns of the N and SmC phases of SO8 were obtained (Fig. 6). In this case, broad wide-angle peaks at $2\theta = 19.2^\circ$ ($d_{\text{WAX}} = 4.6$ Å) and dumbbell-like split small-angle diffractions at 3.3° ($d_{\text{SAX}} = 26.6$ Å) were observed at 140°C [Fig. 6(a)], indicating the tilted N_{CybC} phase. The $\Delta\chi/2$ value or tilt angle was estimated to be 43° , similar to that of SO6. The N_{CybC} –SmC phase transition occurred during X-ray irradiation of the N_{CybC} phase at 130°C [Fig. 6(b)]. As a result, two small-angle diffractions from the SmC layers appeared on the meridional line, in addition to the dumbbell-shaped diffractions from the N_{CybC} layers. The small-angle diffraction of the SmC layers at $2\theta = 3.3^\circ$ corresponds to $d_{\text{SAX}} = 26.7$ Å, which equals that of the pure SmC phase at 120°C [Fig. 6(c)]. The SmC layer was mostly parallel to the equatorial line, and the molecules inclined from the layer normal or the meridional line. The tilt angle of the SmC phase at 120°C was $\sim 20^\circ$, almost half the value of the SmC-like cybotactic clusters in the N_{CybC} phase at 140°C (43°). In addition, the d_{SAX} of the SmC phase (~ 27 Å) was less than the molecular length of 35 Å. Therefore, it was assumed that the alkyl chains intercalated into neighboring layers of the SmC phase [Fig. 6(d)]. Although the detailed layer structures of the N_{CybC} and SmC phases differ somewhat from each other in terms of the tilt angles and intercalation, the structural similarities of their inclined LC phases indicate that Sm-like clusters in the upper N_{CybC} phases likely influence the development of subsequent SmC phase structures.^{4,61}

3.2.2. Shortest-chain homologs. The 2D and 1D XRD patterns of SO1 obtained using a magnetically aligned specimen are shown in Fig. 7. In the N phase at 210°C , a broad wide-angle diffraction at $2\theta = 18.8^\circ$ ($d_{\text{WAX}} = \sim 4.7$ Å) and a weak diffraction peak at $2\theta = 8.3^\circ$ were detected without split on the meridional line [Fig. 7(a) and (c)]. In addition, the d_{WAX} value at $2\theta = 8.3^\circ$ is 10.7 Å, which is almost half of the SO1 molecular length (~ 20 Å), indicating that the molecular cores likely intercalated to form cybotactic Sm clusters. This intercalated structure is remarkably different from the N_{CybC} phases of SO6 and SO8, which show no intercalation among the lateral mesogenic cores. This N phase of SO1 is an intercalated N_{Cyb} phase containing intercalated SmA clusters and denoted as an N_{CybAint} phase [Fig. 7(d)].

For the Sm phase, broadened wide-angle and sharp small-angle diffractions were detected at $2\theta = 19.3^\circ$ ($d_{\text{WAX}} = 4.6$ Å) and 8.0° ($d_{\text{SAX}} = 11$ Å) [Fig. 7(b) and (c)]. Surprisingly, the value $d_{\text{SAX}} = 11$ Å is also almost half the molecular length, suggesting the presence of intercalated layer structures similar to those of the upper N_{CybAint} phase. In addition, the second-order diffraction peak of the layer structure was observed at $2\theta = 16.0^\circ$. The relatively broad wide-angle diffractions indicate the fluidity of the Sm phase and the lateral liquid-like correlations of

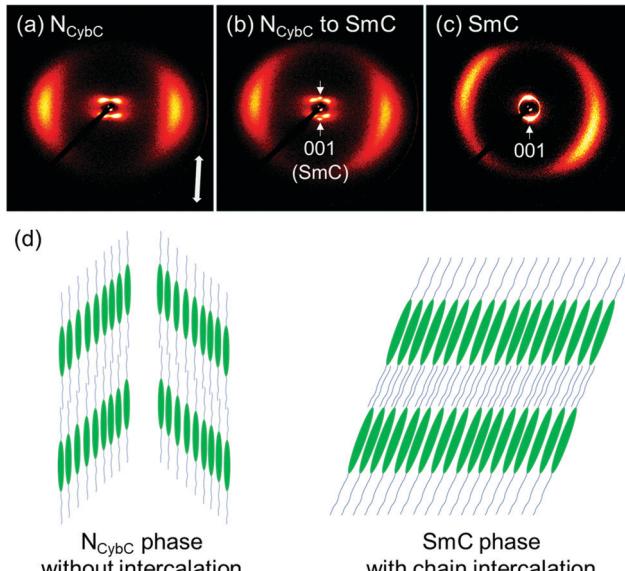


Fig. 6 2D XRD patterns of SO8 (a) in the cybotactic N (N_{Cyb}) phase with a tilted SmC cluster (N_{CybC} phase) at 140°C , (b) during the N–SmC phase transition at 130°C , and (c) in the SmC phase at 120°C . (d) Schematic packing models of the N_{CybC} and SmC phases.



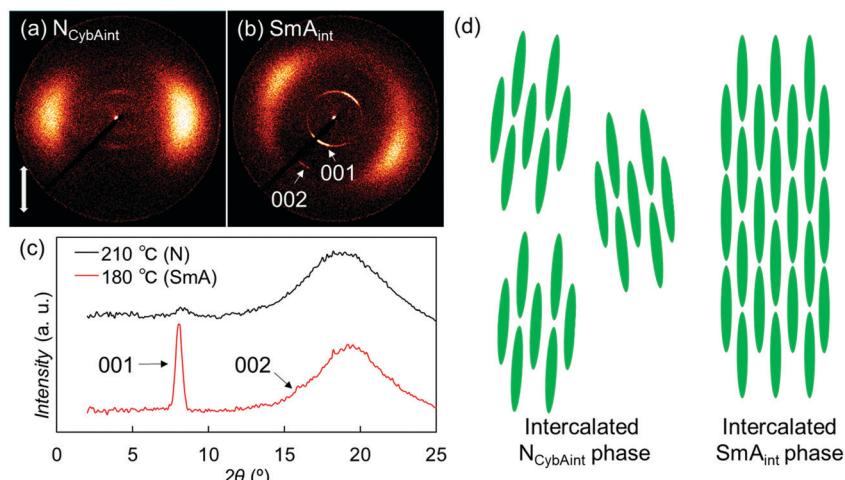


Fig. 7 2D XRD patterns of SO1 (a) in the N phase at 210 °C and (b) in the SmA phase at 180 °C. (c) The corresponding 1D XRD profiles. (d) Schematic representation of the possible packing models of the N_{Cyb} phase with an SmA-like cluster (N_{CybAint} phase) and the intercalated SmA (SmA_{int}) phase of SO1.

molecules in the Sm layers, which identify this phase of SO1 as an intercalated SmA (SmA_{int}) phase [Fig. 7(d)].⁶² The small-angle diffractions at $2\theta = 8.0^\circ$ and 16.0° were indexed as 001 and 002 reflections [Fig. 7(b) and (c)]. Considering the similarities in the molecular structures and the phase transition behaviors among OO1, SS1, and the other analogues, their monotropic SmA phases^{34,44,45} could also be classified as the same SmA_{int} phase with half-intercalated layer structures. The methyl chain is generally too short to facilitate microphase separation to form layered Sm phases. Therefore, the intercalated cybotactic clusters and Sm layers in the N_{CybAint} and SmA_{int} phases are likely ascribable to the strong core–core interactions arising from the highly anisotropic molecular structures that do not depend on the microphase separation driven by long alkyl chains. In addition, as described earlier in Section 3.2.1, PBB homologs with longer alkyl chains exhibited the N_{CybC} phase

with tilted SmC-like clusters and SmC phase.³⁴ Lengthening the alkyl chains led to a change from the non-tilted and intercalated phases of SO1 to the non-intercalated and tilted ones, which are ascribed to the enhanced microphase separation and free volume effect to cause the molecular tilts.²⁸

The LC phases of SO1 indicated strong tendency of PBB mesogenic core to form intercalated packing. If it really arises from core–core interactions, the pristine PBB molecules may exhibit similar packing manner even at low temperature limit: *i.e.*, in its crystalline state. Indeed, it was evidenced by the single-crystal XRD analysis (Fig. 8). The intercalated LC structures were then considered in terms of molecular packing in the single-crystal structures of non-substituted PBB⁶³ and 4,4'-dimethyl-substituted PBB-1.⁶⁴ PBB exhibits a half-shifted packing arrangement between fully overlapped packing domains with an apparent edge alignment [Fig. 8(a)].⁶³ The short interatomic contacts of C=O···H in the packing structures are indicated by O(1)···H(9) and O(2)···H(2) (2.48 and 2.89 Å) between fully overlapped domains and half-shifted packing moieties, respectively. These contacts are ascribable to the C=O···H hydrogen bonds, which are typically observed in phenyl benzoates.^{36,65,66} A similar half-slide packing structure was also observed in PBB-1 [Fig. 8(b)].⁶⁴ The intermolecular C=O···H contacts of O(2)···H(6) and O(2)···H(4) between the bis(benzoate) cores are 2.43 and 2.63 Å in length, respectively, suggesting the presence of C=O···H hydrogen bonds. These intrinsic half-shifted packing structures of the PBB cores in the crystal structure may be associated with the intercalated structures of the N_{CybAint} and SmA_{int} phases of the homologs bearing short alkyl chains [Fig. 7(d)]. Nonetheless, in the case of SS1 which also show a similar SmA phase,³⁴ PBB-core and methylthio moieties are clearly segregated into each domain in its crystalline state rather than such half-shifted packing.³⁶ This microphase separation was interpreted by the large dispersion forces and enhanced affinities between more polarizable core π -systems, afforded by the sulfur atoms. This is also

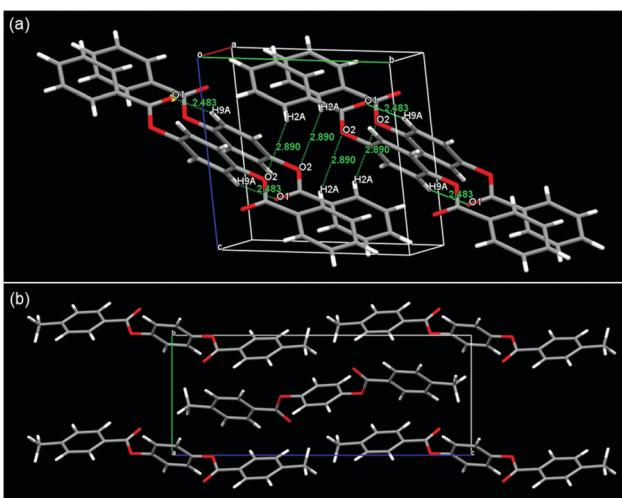


Fig. 8 Single-crystal structures of (a) non-substituted 1,4-phenylene bis(benzoate) (PBB)⁶³ and (b) 4,4'-dimethyl PBB (PBB-1).⁶⁴



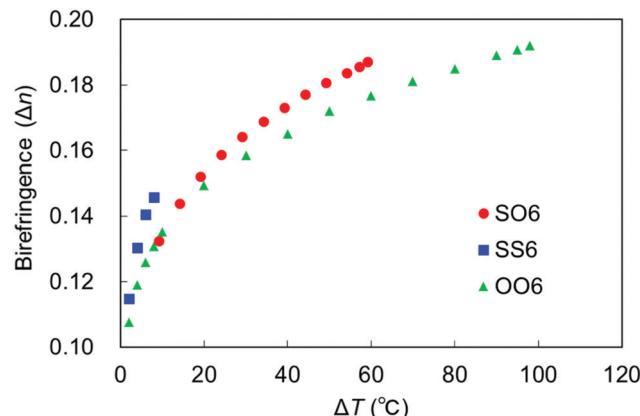


Fig. 9 Birefringence of SO6 (red circles), SS6 (blue squares), and OO6 (green triangles) as a function of ΔT .

likely to be linked to that the influences of those attractive interactions from the sulfur atoms become more pronounced in denser states such as Sm phases.^{36,40} It is interesting to note that the formation of such an unexpected Sm phase by short tail homologs was also observed in other 1,4-bis(ethynyl)benzenes (the so-called bistolanes) bearing methoxy or methylthio groups.⁴⁹ Their LC structures have not ever been reported due to their monotropic nature. This indicates that formation of the Sm phase by short-chain molecules could be characteristic of highly anisotropic molecular systems due to their strong core–core interactions.

3.3. Birefringence

The temperature dependence of birefringence in the N phase of a uniaxially aligned sample of SO6 was measured to characterize its optical properties. The results were compared with those of SS6 and OO6. Thus, the values of Δn obtained at 550 nm upon cooling were plotted as a function of ΔT (Fig. 9), and the Δn values of SS6 and OO6 were taken from our previous work.³⁴ The Δn values of all three molecules increase with decreasing T (*i.e.*, increasing ΔT), which is related to their elevated orientational order parameters upon cooling. Replacing alkoxy groups with alkylthio groups increases the birefringence in the order SS6 > SO6 > OO6, which is consistent with the results of our previous work.^{4,32,34} More specifically, the Δn values of SS6, which contains two alkylthio groups, were found to be significantly higher than those of SO6 and OO6. Moreover, SO6 exhibited higher Δn values than OO6 from the middle range of ΔT upon cooling. These are thought to be due to the high molecular polarizability and anisotropy provided by the alkylthio groups. Birefringence is known to be related

Table 2 Estimated values of Δn_0 , T_i , and β for SO6, OO6, and SS6. The values for OO6 and SS6 are taken from our previous work³⁴

| Code | Δn_0 | T_i | β |
|------|--------------|-------|---------|
| SO6 | 0.27 | 183.7 | 0.18 |
| OO6 | 0.25 | 210.7 | 0.16 |
| SS6 | 0.32 | 158.5 | 0.19 |

to the order parameter (S) by the equations $\Delta n = \Delta n_0 S$ and $S = (1 - T/T_i)^\beta$,⁶⁷ where Δn_0 , T , T_i , and β are the Δn value at $S = 1$, the temperature, the approximate isotropic temperature, and the fitting constant, respectively. The Δn values of the three compounds fit this equation well, and the obtained Δn_0 , T_i , and β values are summarized in Table 2. The obtained T_i values are similar to the experimental T_{IN} values upon cooling, while the β values are within the usual range for nematic phases. In addition, the Δn_0 value decreases in the order SS6 (0.32) > SO6 (0.27) > OO6 (0.25), *i.e.*, it increases upon replacing an alkoxy group with an alkylthio group. The S values of the sulfur-containing SO6 and SS6 homologs obtained from the Δn_0 values tend to be smaller compared to those of the alkoxy-containing OO6 homolog. The narrower LC phase temperature ranges of alkylthio-based materials and their strongly temperature-dependent birefringence properties lead to a higher Δn_0 than those of the alkyl- and alkoxy-containing counterparts.

4. Conclusions

We developed a homologous series of alkylthio- and alkoxy-based asymmetric 1,4-phenylene bis(benzoate) LCs possessing different alkyl chain lengths (SO_n , $n = 1$ –8). Their phase transitions were evaluated and compared with those of the already known symmetric alkoxy and alkylthio analogues (OO_n and SS_n). It was found that all SO_n homologs exhibited well-defined enantiotropic LC phases. The SO_2 – SO_7 homologs only exhibited the N phase, but the longest homolog (SO_8) showed a tilted SmC phase as well as an N phase. Interestingly, the shortest homolog (SO_1) formed an enantiotropic orthogonal SmA phase and an N phase. Upon substituting an alkoxy group with an alkylthio group, the N–Iso phase transition temperature decreased for all n values, whereas the melting points of homologs with $n \geq 3$ increased. The XRD results revealed that the N phases of longer SO_n homologs ($n = 6$ and 8) were N_{Cycb} phases with SmC-like clusters and without alkyl chain intercalation, wherein the alkyl chains of SO_8 intercalated into the neighboring layers in the SmC phase. However, the N and SmA phases of SO_1 were identified as intercalated $N_{CybAint}$ and SmA_{int} phases, respectively, which contained half-intercalated structures. The unexpected appearance of the SmA phase, in addition to the intercalated structures of SO_1 , was considered in terms of the molecular contacts and packings in the crystal structures. Furthermore, the birefringence of the 1,4-phenylene bis(benzoate) LCs could be increased by substituting their alkoxy groups with alkylthio groups. Overall, the present study provides not only a good understanding of the effects of alkylthio and alkoxy groups on the phase transitions and LC properties of such systems, but also insight into the remarkably different N_{Cycb} and Sm phase structures observed for the short- and long-chain homologs of highly anisotropic calamitic LCs.

Conflicts of interest

There are no conflicts to declare.



Acknowledgements

This work was financially supported by JSPS KAKENHI (Grant Numbers 15H06285 and 17K14493) and research grants from Nitto Foundation, Toukai Foundation for Technology, and Toyohashi University of Technology. We would like to thank Ms Tsugumi Shiokawa and Dr Hiroko Tada in the Division of Instrumental Analysis of Okayama University for carrying out the mass spectrometry measurements.

References

- 1 A. J. Seed, K. J. Toyne, J. W. Goodby and D. G. McDonnell, *J. Mater. Chem.*, 1995, **5**, 1.
- 2 A. J. Seed, K. J. Toyne and J. W. Goodby, *J. Mater. Chem.*, 1995, **5**, 2201.
- 3 G. J. Cross, A. J. Seed, K. J. Toyne, J. W. Goodby, M. Hird and M. C. Artal, *J. Mater. Chem.*, 2000, **10**, 1555.
- 4 Y. Arakawa, S. Kang, J. Watanabe and G. Konishi, *RSC Adv.*, 2015, **5**, 8056.
- 5 Y. Arakawa, S. Kang, H. Tsuji, J. Watanabe and G. Konishi, *RSC Adv.*, 2016, **6**, 16568.
- 6 D. Węglowska, P. Kula and J. Herman, *RSC Adv.*, 2016, **6**, 403.
- 7 Y. Arakawa, S. Kang, H. Tsuji, J. Watanabe and G. Konishi, *RSC Adv.*, 2016, **6**, 92845.
- 8 J. Herman and P. Kula, Design of new super-high birefringent isothiocyanato bistolanes—synthesis and properties, *Liq. Cryst.*, 2017, **44**, 1462.
- 9 D. Adam, P. Schumacher, J. Simmerer, L. Haussling, K. Siemensmeyer, K. H. Etzbach, H. Ringsdorf and D. Haarer, *Nature*, 1994, **371**, 141.
- 10 M. Funahashi and J. Hanna, *Jpn. J. Appl. Phys.*, 1996, **35**, L703.
- 11 M. Funahashi and J. Hanna, *Phys. Rev. Lett.*, 1997, **78**, 2184.
- 12 K. Ban, K. Nishizawa, K. Ohta, A. M. van de Craats, J. M. Warman, I. Yamamoto and H. Shirai, *J. Mater. Chem.*, 2001, **11**, 321.
- 13 K. Ohta, K. Hatsusaka, M. Sugibayashi, M. Ariyoshi, K. Ban, F. Maeda, R. Naito, K. Nishizawa, A. M. Van de Craats and J. M. Warman, *Mol. Cryst. Liq. Cryst.*, 2003, **397**, 25.
- 14 F. Nekelson, H. Monobe, M. Shiro and Y. Shimizu, *J. Mater. Chem.*, 2007, **17**, 2607.
- 15 E. F. Gramsbergen, H. J. Hoving, W. H. De Jeu, K. Praefcke and B. Kohne, *Liq. Cryst.*, 1986, **1**, 397.
- 16 P. Doppelt and S. Huille, *New J. Chem.*, 1990, **14**, 607.
- 17 G. Morelli, G. Ricciardi and A. Roviello, *Chem. Phys. Lett.*, 1991, **185**, 468.
- 18 E. Keinan, S. Kumar, S. P. Singh, R. Ghirlando and E. J. Wachtel, *Liq. Cryst.*, 1992, **11**, 157.
- 19 S. Kumar, E. J. Wachtel and E. Keinan, *J. Org. Chem.*, 1993, **58**, 3821.
- 20 S. H. Kang, Y.-S. Kang, W.-C. Zin, G. Olbrechts, K. Wostyn, K. Clays, A. Persoons and K. Kim, *Chem. Commun.*, 1999, 1661.
- 21 G. Kestemont, V. de Halleux, M. Lehmann, D. A. Ivanov, M. Watson and Y. H. Geerts, *Chem. Commun.*, 2001, 2074.
- 22 P. H. J. Kouwer, W. F. Jager, W. J. Mijs and S. J. Picken, *J. Mater. Chem.*, 2003, **13**, 458.
- 23 P. Kouwer, W. Jager, W. Mijs and S. Picken, *Mol. Cryst. Liq. Cryst.*, 2004, **411**, 305.
- 24 S. Kumar, *Liq. Cryst.*, 2004, **31**, 1037.
- 25 S. K. Varshney, H. Nagayama, V. Prasad and H. Takezoe, *Liq. Cryst.*, 2011, **38**, 1321.
- 26 H. Eichhorn, D. W. Bruce and D. Wöhrle, *Adv. Mater.*, 1998, **10**, 419.
- 27 M. Alaasar, A. F. Darweesh, X. Cai, F. Liu and C. Tschierske, *Chem. – Eur. J.*, 2021, **27**, 14921.
- 28 J. W. Goodby, R. J. Mandle, E. J. Davis, T. Zhong and S. J. Cowling, *Liq. Cryst.*, 2015, **42**, 593.
- 29 E. Kleinpeter, H. Köhler, A. Lunow, C. Tschierske and H. Zaschke, *Tetrahedron*, 1988, **44**, 1609.
- 30 Y. Arakawa, Y. Sasaki, K. Igawa and H. Tsuji, *New J. Chem.*, 2017, **41**, 6514.
- 31 Y. Arakawa, Y. Sasaki and H. Tsuji, *Chem. Lett.*, 2017, **46**, 1657.
- 32 Y. Arakawa and H. Tsuji, *Mol. Cryst. Liq. Cryst.*, 2017, **647**, 422.
- 33 Y. Arakawa, S. Inui and H. Tsuji, *Liq. Cryst.*, 2018, **45**, 811.
- 34 Y. Arakawa, Y. Sasaki, N. Haraguchi, S. Itsuno and H. Tsuji, *Liq. Cryst.*, 2018, **45**, 821.
- 35 Y. Arakawa, Y. Sasaki and H. Tsuji, *J. Mol. Liq.*, 2019, **280**, 153.
- 36 Y. Arakawa, Y. Ishida, Y. Sasaki, K. Igawa, S. Sasaki and H. Tsuji, *CrystEngComm*, 2022, DOI: 10.1039/D1CE01470F.
- 37 X. Zong, Z. Fang and C. Wu, *Liq. Cryst.*, 2018, **45**, 1844.
- 38 Y. Arakawa, K. Komatsu and H. Tsuji, *New J. Chem.*, 2019, **43**, 6786.
- 39 I. C. Pintre, J. L. Serrano, M. B. Ros, J. Ortega, I. Alonso, J. Martínez-Perdiguer, C. L. Folcia, J. Etxebarria, F. Goc, D. B. Amabilino, J. Puigmarti-Luis and E. Gomar-Nadal, *Chem. Commun.*, 2008, 2523.
- 40 Y. Arakawa, S. Inui and H. Tsuji, *Polymer*, 2021, **233**, 124194.
- 41 M. J. S. Dewar and J. P. Schroeder, *J. Org. Chem.*, 1965, **30**, 2296.
- 42 M. J. S. Dewar and R. S. Goldberg, *J. Org. Chem.*, 1970, **35**, 2711.
- 43 M. J. S. Dewar and R. S. Goldberg, *J. Am. Chem. Soc.*, 1970, **92**, 1582.
- 44 M. J. S. Dewar and A. C. Griffin, *J. Am. Chem. Soc.*, 1975, **97**, 6662.
- 45 J. P. Schroeder, *Mol. Cryst. Liq. Cryst.*, 1980, **61**, 229.
- 46 W. Weissflog and D. Demus, *Cryst. Res. Technol.*, 1984, **19**, 55.
- 47 M. M. Naoum, R. I. Nessim, G. R. Saad and T. Y. Labeeb, *Liq. Cryst.*, 2003, **30**, 165.
- 48 T. Tasaka, H. Okamoto, Y. Morita, K. Kasatani and S. Takenaka, *Liq. Cryst.*, 2003, **30**, 961.
- 49 P. Nguyen, G. Lesley, C. Dai, N. J. Taylor, T. B. Marder, V. Chu, C. Viney, I. Ledoux and J. Zyss, Well-Defined Conjugated Rigid-Rods as Multifunctional Materials: Linear



and Nonlinear Optical Properties and Liquid Crystalline Behavior, in *Applications of Organometallic Chemistry in the Preparation and Processing of Advanced Materials*, ed. J. F. Harrod and R. M. Laine, NATO ASI Series (Series E: Applied Sciences), Springer, Dordrecht, 1995, vol. 297, p. 333, DOI: 10.1007/978-94-011-0337-4_19.

50 Y. Ishida, Y. Sasaki, Y. Arakawa and H. Tsuji, *Polym. Prepr.*, 2018, **67**, 1.

51 Y. Arakawa, S. Nakajima, R. Ishige, M. Uchimura, S. Kang, G. Konishi and J. Watanabe, *J. Mater. Chem.*, 2012, **22**, 8394.

52 A. D. Vries, *Mol. Cryst. Liq. Cryst.*, 1970, **10**, 219.

53 C. Tschierske and D. J. Photinos, *J. Mater. Chem.*, 2010, **20**, 4263.

54 S. J. Picken, T. J. Dingemans, L. A. Madsen, O. Francescangeli and E. T. Samulski, *Liq. Cryst.*, 2012, **39**, 19.

55 S. Marčelja, *J. Chem. Phys.*, 1974, **60**, 3599.

56 A. P. J. Emerson and G. R. Luckhurst, *Liq. Cryst.*, 1991, **10**, 861.

57 C. Tschierske, D. Joachimi, H. M. Vorbrot, H. Zaschke, A. Wiegeleben, A. Hauser and D. Demus, *Liq. Cryst.*, 1989, **5**, 177.

58 Y. Arakawa, S. Inui, K. Igawa and H. Tsuji, *Liq. Cryst.*, 2019, **46**, 1621.

59 E. Cruickshank, G. J. Strachan, J. M. Storey and C. T. Imrie, *J. Mol. Liq.*, 2021, 117094.

60 S. Kang, S. Nakajima, Y. Arakawa, M. Tokita, J. Watanabe and G. Konishi, *Polym. Chem.*, 2014, **5**, 2253.

61 N. Vaupotič, J. Szydłowska, M. Salamonczyk, A. Kovarova, J. Svoboda, M. Osipov, D. Pociecha and E. Gorecka, *Phys. Rev. E: Stat., Nonlinear, Soft Matter Phys.*, 2009, **80**, 030701.

62 K. Fodor-Csorba, A. Jákli, A. Vajda, S. K. Prasad, D. S. Shankar Rao, R. Y. Dong, J. Xu and G. Galli, *Chem-PhysChem*, 2006, **7**, 2184.

63 J. A. Ganaie, J. Kumar, R. J. Butcher, J. P. Jasinski and S. K. Gupta, *J. Chem. Crystallogr.*, 2016, **46**, 93.

64 M. R. Ciajolo, A. Sirigu and A. Tuzi, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 1991, **C47**, 106.

65 J. M. Adams and S. E. Morsi, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1976, **B32**, 1345.

66 L. G. Kuz'mina, I. I. Konstantinov and A. V. Churakov, *Mol. Cryst. Liq. Cryst.*, 2018, **664**, 95.

67 I. Haller, *Prog. Solid State Chem.*, 1975, **10**, 103.

