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Aqueous zinc ion batteries based on sodium vanadate electrode materials with long lifespan and high energy density

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With the excessive consumption of non-renewable energy sources and subsequent environmental pollution, research focuses have shifted to the development of emerging energy storage devices with desired performance. Aqueous zinc-ion batteries (AZIBs) are expected to become potential alternatives due to their high safety, low cost and environmental friendliness. In this work, we have prepared $\text{Na}_7\text{V}_{7.6}\text{O}_{20}\cdot 4\text{H}_2\text{O}$ (NVO) nanobelts by a facile hydrothermal route. NVO/Zn batteries have been assembled with the as-obtained product as cathodes, which deliver a capacity of $309.4 \text{ mA h g}^{-1}$ at 0.3 A g^{-1} . They maintain excellent cycling stability after 10 000 cycles at 10 A g^{-1} .

1. Introduction

It is well known that traditional fossil energy is non-renewable and its overuse will inevitably lead to the quick depletion of energy resources.^{1–4} Therefore, developing emerging energy storage devices is imperative to effectively alleviate the energy crisis. For instance, commercial lithium-ion batteries (LIBs) have been widely used in portable electronic devices such as mobile phones, computers and wearable watches.^{5,6} However, the application of LIBs is largely restricted due to the scarcity of lithium reserves, its high price and poor safety.^{7,8} Therefore, metal ion batteries with high energy densities have attracted widespread attention.⁹ Among the various multivalent ion systems (Al^{3+} ,¹⁰ Mg^{2+} ,^{11,12} Ca^{2+} ,^{13,14} and Zn^{2+}), AZIBs show high theoretical capacity (820 mA h g^{-1}), low redox potential (-0.76 V vs. SHE) and excellent cycle stability.¹⁶ Recently, some progress has been achieved in developing cathode materials for ZIBs, such as manganese oxide,^{17,18} vanadium compounds,^{19,20} Prussian blue analogs,²¹ and polyanionic compounds.²²

Therefore, it is important to further study vanadium-based cathodes because of their high capacity and rich oxidation states.^{23–25} They mainly include V_2O_5 ,²⁶ VO_2 ,²⁷ and vanadates.^{28–31} However, their poor energy density and sluggish ion transfer speed have not been well resolved. For these reasons, some efforts have been made. In the literature, Wang *et al.* prepared a composite of V_2O_5 and carbon nanotube film (CNF@ V_2O_5). The as-assembled device delivers a specific

capacity of $356.6 \text{ mA h g}^{-1}$ at 0.4 A g^{-1} and maintains 80.1% of the initial capacity after 500 cycles at 2 A g^{-1} .³² Jia and coworkers fixed $\text{VO}_2\cdot 0.2\text{H}_2\text{O}$ on graphene sheets using a microwave-assisted strategy. The assembled battery shows a capacity of 423 mA h g^{-1} at 0.25 A g^{-1} and a discharge capacity retention rate of 87% after 1000 cycles at 8 A g^{-1} .³³ In addition, $\text{Na}_{1.1}\text{V}_3\text{O}_{7.9}$ nanorods were obtained by Zhou's group. Using the obtained product as the cathode, the cells possess a capacity of 134 mA h g^{-1} after 1000 cycles at 5 A g^{-1} .³⁴

Herein, we prepared NVO nanobelts using $\text{C}_6\text{H}_8\text{O}_7\cdot \text{H}_2\text{O}$ as the binding agent. The obtained NVO as an electrode material not only maintains a stable structure during charging and discharging but also shows an energy density of $433.02 \text{ W h kg}^{-1}$ at a power density of 420 W kg^{-1} at 0.3 A g^{-1} . Furthermore, the assembled NVO/Zn batteries deliver a high capacity of $309.4 \text{ mA h g}^{-1}$ at 0.3 A g^{-1} and a capacity retention rate of 98.56% after 10 000 cycles at 10 A g^{-1} , suggesting their excellent rate performance.

2. Experimental section

2.1 Material preparation

All chemicals were used as purchased without further purification. Firstly, 10 mmol NaVO_3 was added into 60 mL deionized water and stirred at room temperature for 40 min. Subsequently, 8 mmol $\text{C}_6\text{H}_8\text{O}_7\cdot \text{H}_2\text{O}$ and 20 mmol $\text{Na}_2\text{S}_2\text{O}_8$ were poured into the above solution and stirred for 1 h. Finally, the clear solution was transferred into a 100 mL Teflon-lined autoclave and heated at $180 \text{ }^\circ\text{C}$ for 6 h. After cooling to room temperature, the obtained precipitate was washed many times with deionized water. Then it was dried in an oven.

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2.2 Structural characterization

The morphology and element analysis of the products were studied by field emission scanning electron microscopy (FESEM, Gemini 300) and transmission electron microscopy (TEM, JEM-F200) equipped with energy dispersive spectroscopy (EDS). The crystal structure of the sample was analyzed using X-ray diffraction (XRD, Bruker D2) and X-ray photoelectron spectroscopy (XPS, Thermo Fisher ESCALAB XI+).

2.3 Electrochemical measurement

The cathode, zinc anode and separator were assembled in 2032-type coin cells in air. The cathode was prepared by mixing the NVO, Ketjen black (KB) and polyvinylidene fluoride (PVDF) with a weight ratio of 7:2:1 in N-methyl-2-pyrrolidone (NMP) solvent. The obtained slurry was subjected to magnetic stirring for 8 h to mix it uniformly. The slurry was coated evenly on a carbon paper as the cathode, and then placed it in a 60 °C vacuum drying oven for a night. A zinc sheet with a thickness of 100 μm was used as the anode, and glass fiber was used as the intermediate separator. The devices were assembled using 4 M Zn(CF₃SO₃)₂ aqueous electrolyte. The cyclic voltammetry (CV) performance curves were tested in a CHI760E electrochemical workstation. The galvanostatic charge–discharge (GCD) curves, long-cycle stability and galvanostatic intermittent titration

technique (GITT) were characterized by a LAND-CT3001 automatic battery tester with the voltage range of 0.2–1.6 V. The average loading of the cathode active material is about 1.3 mg.

3. Results and discussion

The crystal structure of the sample was first studied by XRD. The diffraction peaks (Fig. 1a) locate at $2\theta = 8.25^\circ$ (001), 24.7° (003), 25.64° (110), 33.13° (004), 33.21° (−204), 34.38° (401), 14.77° (005), 46.87° (−602), 50.02° (020), 61.1° (−712) and 64.32° (−407). These peaks are consistent with NVO (PDF#85-1407). Except these, there are no other impurity peaks, demonstrating the high purity of the prepared sample. Then, XPS was used to further investigate the oxidation states of the Na, V, and O elements. The high resolution XPS of Na 1s (Fig. 1b) only shows a sharp peak (1071.35 eV), indicating that Na⁺ is the only existing form of the Na element. For V 2p, as shown in Fig. 1c, the diffraction peaks of V 2p_{3/2} and V 2p_{1/2} are all fitted into V⁺⁵ and V⁺⁴. There are three diffraction peaks that appear in the XPS spectra of O 1s (529.6 eV, 530 eV, and 530.5 eV), as shown in Fig. 1d. They can be ascribed to V–O, OH[−] and low-coordination oxygen ions, respectively.³⁵

Subsequently, we used SEM to observe the morphology of the product. It was found that the prepared NVO product shows

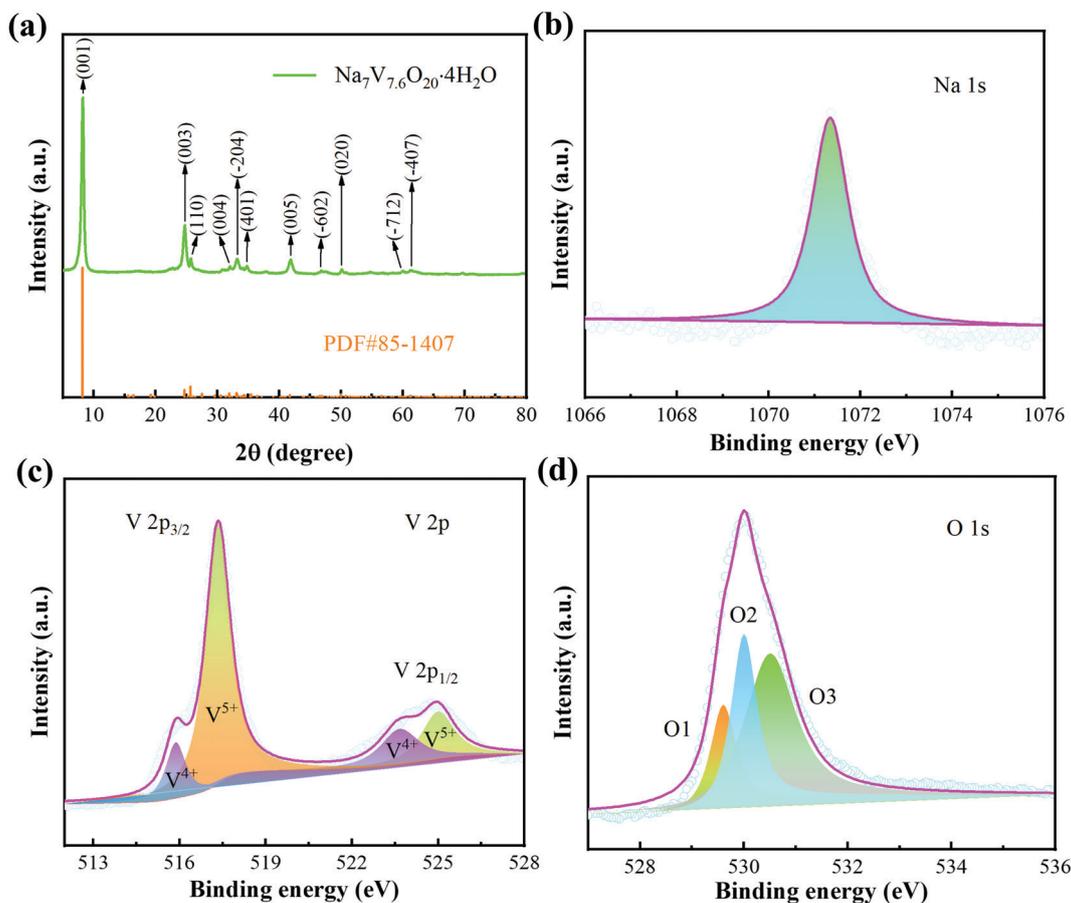


Fig. 1 Structural characterization of the samples. (a) XRD patterns. (b–d) XPS spectra of Na 1s, V 2p and O 1s.



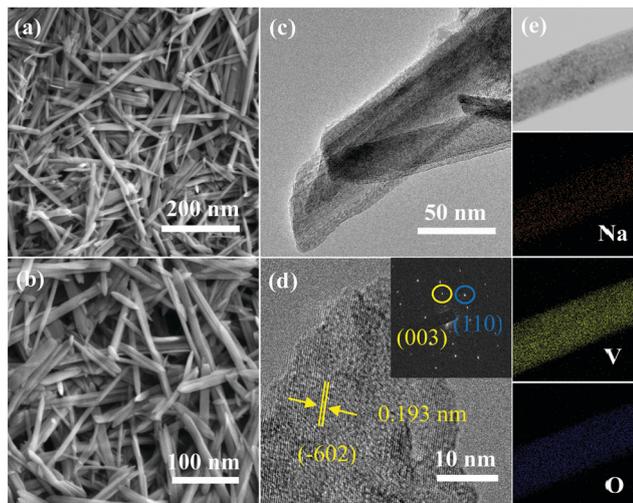


Fig. 2 Morphology and structure characterization of the samples. (a and b) SEM images. (c) TEM image. (d) HRTEM image. (e) The corresponding elemental mappings of Na, V and O.

a belt-like shape (Fig. 2a and b). TEM was used to further study the microstructures of the sample. Fig. 2c is a low magnification TEM image, revealing that the products possess a rectangular feature. The HRTEM image (Fig. 2d) show that the interplanar spacing is 0.193 nm, which can be indexed to the (-602) plane (PDF#85-1407). The SAED image (the inset in Fig. 2d) suggests the single crystalline structure of the sample. The element mapping in Fig. 2e proves the uniform distribution of the Na, V, and O elements in the product.

In order to test the electrochemical performance characteristics of the sample, the batteries were assembled using obtained NVO as the cathode in a 4 M $\text{Zn}(\text{CF}_3\text{SO}_3)_2$ electrolyte. Fig. 3a shows the first five cycles of the CV curves with a potential of 0.2–1.6 V at 0.5 mV s^{-1} . The corresponding oxidation peaks locate at 0.548 V, 0.735 V, 0.988 V, 1.061 V, 1.362 V and 1.432 V. The reduction peaks appear at 1.328 V, 0.961 V, 0.82 V, 0.59 V and 0.397 V. It is found that the shapes of the curves in the first cycle and the last four cycles are different, revealing that the oxidation-reduction reaction in the first circle is irreversible. The rate performance of the device is shown in Fig. 3b. As the current density varies from 0.3 to 10 A g^{-1} , the discharge specific capacity of the device gradually decreases from 309.4 to $240.4 \text{ mA h g}^{-1}$. The small capacity change during whole reaction process indicates that the device shows a little capacity loss at different current densities. The GCD curves of the battery at different current densities (Fig. 3c) are consistent with the rate curves. Fig. 3d gives the cycling time dependent on the GCD curves at 1 A g^{-1} . The results show that the capacity of the device increases with the increase in cycling time, which can be attributed to the activation behavior of NVO in the electrolyte. Moreover, the shapes of the GCD curves remain unchanged, suggesting the highly reversible redox reaction during the entire charge-discharge cycling.

The cycling stability of the NVO/Zn batteries was also tested at different current densities. The cycling tests at 1 A g^{-1} (Fig. 3e) show that the initial specific capacity is $298.5 \text{ mA h g}^{-1}$, and the

largest specific capacity of $301.5 \text{ mA h g}^{-1}$ is obtained after 100 cycles. The devices can still show a specific capacity of $231.8 \text{ mA h g}^{-1}$ after 300 cycles, and the capacity retention rate is 77.6%. Fig. 3f shows the cycle performance of the battery at 5 A g^{-1} . After 1400 cycles, the specific discharge capacity of the cell does not decrease significantly. We then studied the long-cycle performance of the NVO/Zn battery at 10 A g^{-1} . The initial discharge capacity is $195.1 \text{ mA h g}^{-1}$. The capacity of the device reaches a maximum capacity of $243.5 \text{ mA h g}^{-1}$ at 3000 cycles and then decreases gradually to $192.3 \text{ mA h g}^{-1}$ at 10 000 cycles, as shown in Fig. 3g. The capacity retention rate is 98.56% during this process. The Coulombic efficiency always remains above 99%. The rate performance and cycling stability is better than several reported vanadium-based materials (Table 1).^{36–42}

To evaluate the electrochemical reaction kinetics of the NVO/Zn battery, the CV curves were tested at different scan rates (from 0.6 to 1 mV s^{-1}) (Fig. 4a). The integral areas of the CV curves increase as the scanning speed rises, demonstrating the multi-step process of the deintercalation/intercalation of Zn^{2+} . In addition, the oxidation and reduction peaks shift to the high voltage and low voltage regions, respectively. Fig. 4b is obtained by fitting the currents at different redox peaks, where peaks 1, 2, 3, and 4 correspond to their respective ones in Fig. 4a. The correlation between the peak current (i) and scan rate (v) abide by the following formula:

$$i = av^b \quad (1)$$

When the b value is 0.5, the charge storage can be expressed as diffusion control behavior. When $b = 1$, it is thought to be a capacitive process. The calculated b values of peaks 1, 2, 3, and 4 are 0.6574, 0.9529, 0.5976, and 0.8396, respectively. The value of b is between 0.5 and 1, indicating that pseudo-capacitance is dominant during the charge storage process. The ratio of surface and diffusion control contribution can be obtained according to the formula:

$$i = k_1v + k_2v^{1/2} \quad (2)$$

where k_1 and k_2 are constants and k_1v and $k_2v^{1/2}$ are the surface control and diffusion control contribution, respectively.⁴³ Fig. 4c shows a histogram of the surface and diffusion controlled ratios. When the scan rate increases from 0.6 to 1 mV s^{-1} , the capacitance contribution gradually increases from 65.1% to 71.6%. The results show that the devices possess high reaction kinetics and rate performance.

The diffusion coefficient of Zn^{2+} ($D_{\text{Zn}^{2+}}$) of the NVO electrode was then calculated from the obtained GITT curves. Firstly, the NVO/Zn batteries are charged or discharged at 0.3 A g^{-1} for 3 min, and then rested for 7 min to make the voltage reach steady. This process is repeated to obtain the GITT curves during battery operation (Fig. 4d). The calculated $D_{\text{Zn}^{2+}}$ is shown in Fig. 4e according to the formula:

$$D_{\text{Zn}^{2+}} = \frac{4}{\pi\tau} \left(\frac{\text{mV}}{\text{MS}} \right)^2 \left(\frac{\Delta E_s}{\Delta E_\tau} \right)^2 \quad (3)$$



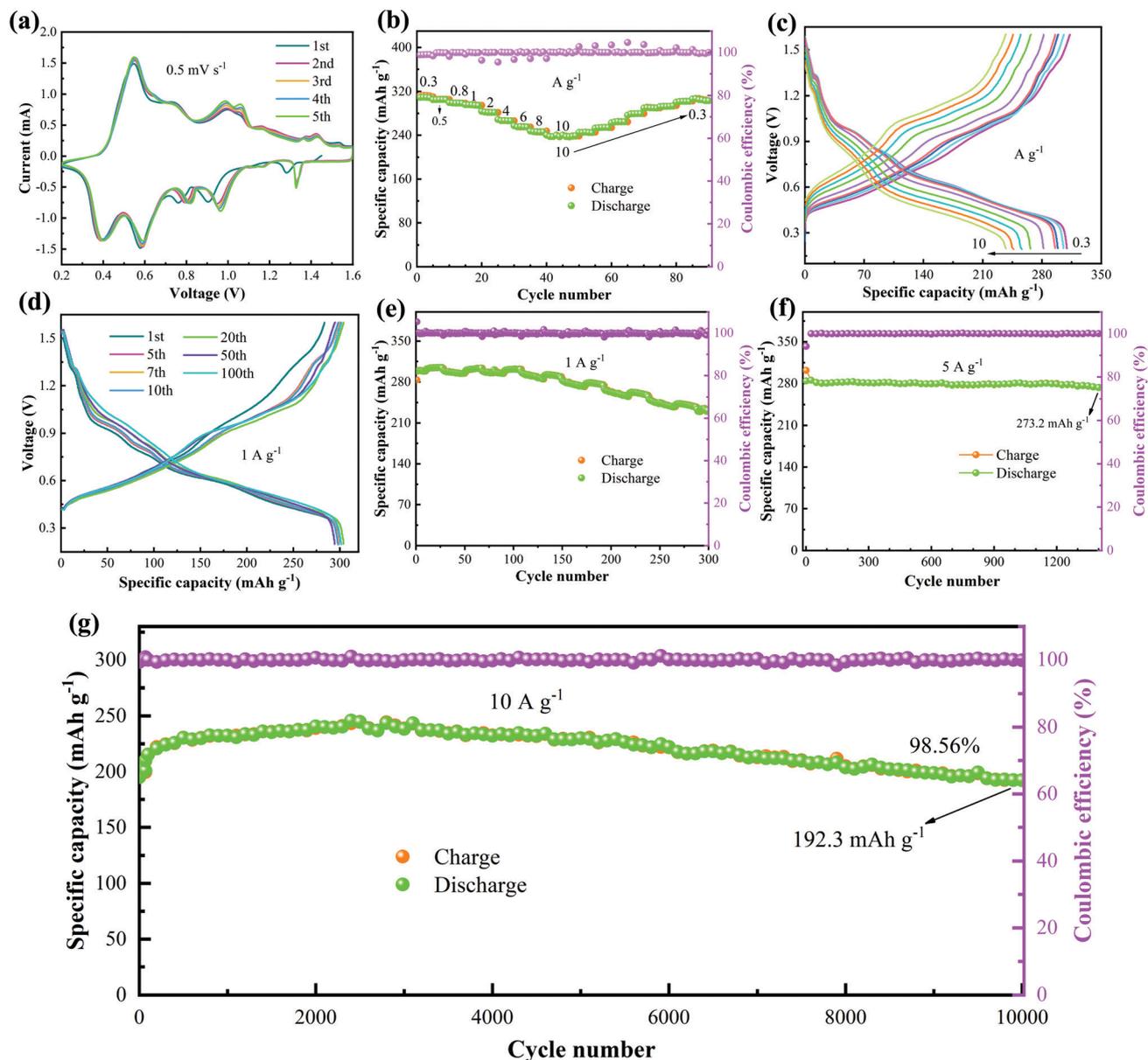


Fig. 3 Electrochemical performance of the electrodes. (a) CV curves of the first 5 cycles at 0.5 mV s^{-1} , (b and c) rate capability and GCD curves at different current densities, (d) GCD curves of various cycles at 1 A g^{-1} , (e) cycling performance at 1 A g^{-1} , (f) cycling performance at 5 A g^{-1} , and (g) long-term cycles at 10 A g^{-1} .

Table 1 Rate performance and cycle stability of different cathodes

Cathode material	Rate performance (mA h g^{-1})	Retention, %/cycle	Ref.
$\text{Na}_7\text{V}_{7.6}\text{O}_{20} \cdot 4\text{H}_2\text{O}$	240.4 (10 A g^{-1})	98.56/10 000 (10 A g^{-1})	This work
$\text{Zn}_3\text{V}_3\text{O}_8$	141 (5 A g^{-1})	72.6/2000 (5 A g^{-1})	36
$\text{Al}_{0.2}\text{V}_2\text{O}_5$	143.7 (10 A g^{-1})	61.4/5000 (6 A g^{-1})	37
$\text{Cu}_{0.26}\text{V}_2\text{O}_5@C$	163.8 (2 A g^{-1})	93.5/500 (2 A g^{-1})	38
$\text{Na}_{1.25}\text{V}_3\text{O}_8$	162 (10 A g^{-1})	88.2/2000 (10 A g^{-1})	39
$r\text{GO}/\delta\text{-Na}_x\text{V}_2\text{O}_5 \cdot n\text{H}_2\text{O}$	244.1 (5 A g^{-1})	70.5/1000 (5 A g^{-1})	40
$\text{V}_2\text{O}_5 \cdot 4\text{VO}_2 \cdot 2.72\text{H}_2\text{O}$	215 (10 A g^{-1})	94.0/1000 (10 A g^{-1})	41
$\text{K}_2\text{V}_8\text{O}_{20}$	92 (4 A g^{-1})	83/300 (6 A g^{-1})	42

where τ is the duration time (s) of charging or discharging, m is the loading of the active material on the cathode (g), V and M

are the molar volume ($\text{cm}^3 \text{ mol}^{-1}$) and molecular weight (g mol^{-1}) of the active material, respectively, S is the contact



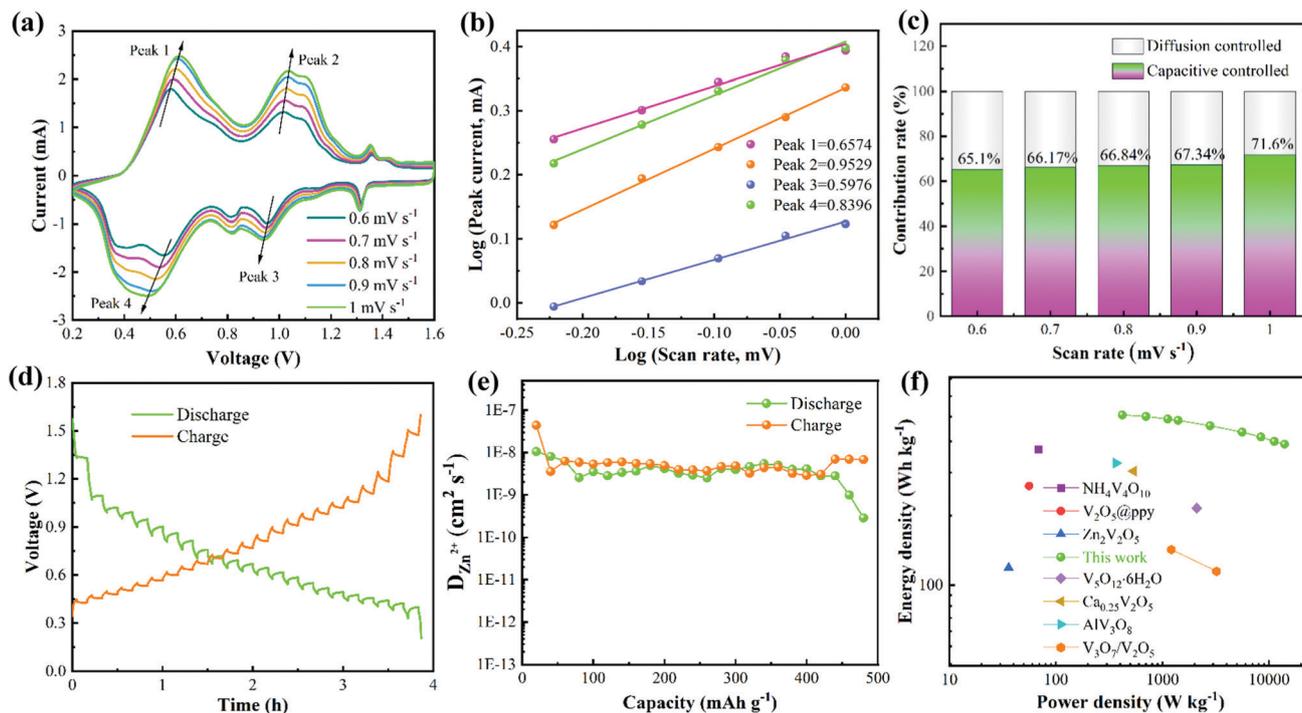


Fig. 4 The electrochemical reaction kinetics of the batteries. (a) CV curves at various scan rates, (b) the fitting plots of log (*i*) and log (*v*), (c) the capacitive contribution ratio at various scan rates, (d) GITT curves, (e) Zn²⁺ diffusion coefficient, and (f) Ragone plot.

area between the electrode and the electrolyte, ΔE_s is the variation between the initial voltage and the steady-state voltage, and $\Delta E\tau$ is the variation of the battery voltage during

titration.⁴⁴ In our work, the range of $D_{\text{Zn}^{2+}}$ during discharge is between 10^{-8} and 10^{-10} cm² s⁻¹, and the value of $D_{\text{Zn}^{2+}}$ drops to the lowest when the discharge reaches 0.2 V. It shows that

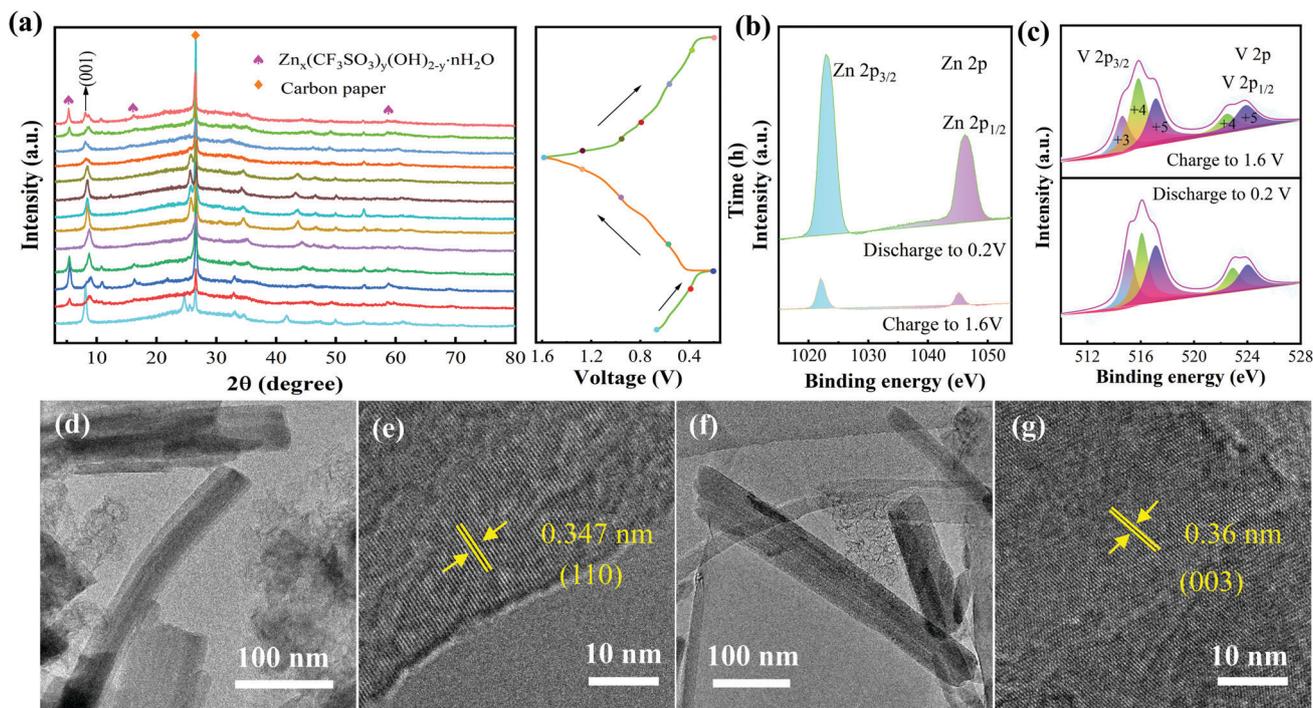


Fig. 5 The reaction mechanism of the batteries. (a) *Ex situ* XRD patterns at various charge/discharge states. (b) and (c) XPS of Zn 2p and V 2p at the charge–discharge states. (d–g) TEM images of the sample discharging at 0.2 V and charging at 1.6 V.



the intercalation of Zn^{2+} is beneficial to the enhancement of electrostatic repulsion inside the NVO.⁴⁵ The value of $D_{\text{Zn}^{2+}}$ during charging is between 10^{-9} and 10^{-10} $\text{cm}^2 \text{s}^{-1}$, which is better than the reported vanadium-based cathode.^{46,47} The battery presents an energy density of 433.02 W h kg^{-1} at a power density of 420 W kg^{-1} at 0.3 A g^{-1} . When the power density reaches 14 000 W kg^{-1} at 10 A g^{-1} , it still maintains an energy density of 336.56 W h kg^{-1} . The Ragone plot (Fig. 4f) lists the energy densities of several vanadium-based materials ($\text{NH}_4\text{V}_4\text{O}_{10}$,⁴⁸ $\text{V}_2\text{O}_5@ppy$,⁴⁹ $\text{Zn}_2\text{V}_2\text{O}_7$,⁵⁰ $\text{V}_5\text{O}_{12}\cdot 6\text{H}_2\text{O}$,⁵¹ $\text{Ca}_{0.25}\text{V}_2\text{O}_5$,⁵² AlV_3O_9 ,⁵³ and $\text{V}_3\text{O}_7/\text{V}_2\text{O}_5$,⁵⁴). It is obvious that the energy density and power density of the obtained NVO are better than those cathodes as well as some other materials.^{55–57}

Finally, we investigated the dynamic migration process of Zn^{2+} . The *ex situ* XRD patterns (Fig. 5a) indicate the structural evolution of the NVO electrode during charging and discharging. It can be seen that the diffraction peaks of the (001) crystal plane show an angle red shift after the first discharge to 0.2 V. It reveals the strong electrostatic force between the intercalation of Zn^{2+} and the cathode during the first discharge. The (001) plane returns to its original state gradually with the process of charging to 1.6 V. The angles of the (001) plane only slightly shift when the second discharge is performed, indicating that the intercalation of Zn^{2+} does not significantly change the crystal plane spacing. It further proves the excellent structural stability of the electrode materials during the charge–discharge process. Besides this, there are several new diffraction peaks during the first discharge process, revealing the formation of interphase $\text{Zn}_x(\text{CF}_3\text{SO}_3)_y(\text{OH})_{2-y}$.⁵⁸ During the subsequent charging to 1.6 V, the diffraction peaks of the interphase gradually disappear, indicating the high reversibility of the interphase during the entire cycle.

To study the valence state and surface composition of the NVO electrode, *ex situ* XPS tests were performed on the cathode in different charging/discharging states. As shown in Fig. 5b, the strong signals of Zn 2p_{3/2} (1022.9 eV) and Zn 2p_{1/2} (1045.9 eV) appear during the discharge to 0.2 V. The peak intensity decreased significantly during charging to 1.6 V, demonstrating that Zn ions are successfully inserted into the cathode. Fig. 5c shows the high-resolution XPS spectra of V 2p at the charging and discharging states. The peaks at V 2p_{3/2} can be fitted by V⁺³ (515.5 eV), V⁺⁴ (516 eV) and V⁺⁵ (516.8 eV), while the peaks at V 2p_{1/2} are only fitted by V⁺⁴ (522.15 eV) and V⁺⁵ (523.75 eV). In the fully charged and discharged states, the peak areas change at different valence states. In discharging to 0.2 V, the V⁺⁴ signals weaken and the signals of V⁺³ and V⁺⁵ strengthen. The signals of V⁺³ and V⁺⁵ gradually increase and the signal of V⁺⁴ decreases when charging to 1.6 V. It shows that the oxidation–reduction reactions of the V element are highly reversible during the cycle.⁵⁹ The structural characteristics of the electrode were tested again using TEM. Fig. 5d and e and Fig. 5f and g show the TEM images of the electrode discharged to 0.2 V and charged to 1.6 V, respectively. During the discharge to 0.2 V, the structure of the NVO is maintained well and the interplanar spacing of the (110) plane is 0.347 nm, suggesting that the intercalation of Zn^{2+} does not obviously change the interplanar spacing. The interplanar

spacing of the (003) plane is 0.36 nm after the release of Zn^{2+} , proving the excellent structural stability of the NVO products.

4. Conclusion

In summary, we have prepared NVO nanobelts using a simple one-step hydrothermal strategy. The obtained NVO product can maintain superior structural stability after the intercalation of zinc ions. It also enhances the transfer rate of Zn^{2+} inside the cathode. The assembled batteries show high rate performance, fast electrochemical reaction kinetics and a long-term cycle lifespan. The subsequent TEM observation of the electrode materials proves that the NVO electrode maintains unchanged structural characteristics during the charge–discharge process. It suggests that the battery assembled by the as-prepared product can be utilized in future portable micro/nanodevices.

Author contributions

Chengkang Luo: conceptualization, methodology, software, data processing, writing – original draft preparation. Chengkang Luo and Li Xiao: visualization, investigation. Xiang Wu: validation, supervision, writing – review & editing.

Conflicts of interest

The authors declare that they have no known competing financial interests.

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