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3, 1263**Mixed-dimensional organic–inorganic metal halide perovskite (OIMHP) based gas sensors with superior stability for NO₂ detection†**The Duong, ^{‡*a} Alishba T. John, ^{‡b} Hongjun Chen,^{*c} Huyen Pham,^d Krishnan Murugappan, ^b Thanh Tran-Phu,^{be} Antonio Tricoli ^{be} and Kylie Catchpole^a

Organic–inorganic metal halide perovskite (OIMHP) materials have been intensively investigated for optoelectronic applications, especially in solar cells. Here, we investigate OIMHP materials for gas sensor application to detect NO₂ gas, a harmful gas for humans and the environment. As a control, although the three-dimensional (3D) OIMHP NO₂ gas sensors exhibit good performance toward NO₂ gas when operating at room temperature, ambient stability remains the bottleneck for practical application. Using dimensional engineering of the OIMHP materials, not only the ambient stability but also the performance of the perovskite sensors has been greatly improved. The mixed-dimensional 2D/3D OIMHP gas sensors show an excellent sensitivity of 6.3 ± 0.83 times per ppm NO₂, good selectivity toward NO₂ gas, and quick response and recovery times of 5.7 s and 12.7 s, respectively. More importantly, the sensor performance can be stable under ambient conditions for almost two months, much longer than the control 3D-only device (13 days). The work paves the way for the development of perovskite based NO₂ gas sensors with superior stability and excellent performance at room temperature.

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rsc.li/materials-advances**Introduction**

Nitrogen dioxide (NO₂) gas is deadly to humans as the gas can cause serious health issues even in extremely low concentrations (<10 parts per million (ppm)).^{1,2} In addition, the gas is also harmful to the environment since it can react with other pollutants or water in the environment to generate ozone and acid rain.^{3,4} Therefore, it is critical to reliably detect trace amounts of NO₂ gas. Over the past few decades, extensive research work on NO₂ gas sensors has been carried out. Various materials have been used to fabricate NO₂ gas sensors, each with distinctive performance and stability depending on the material properties, deposition techniques and sensor designs.

These include polymers,^{5–8} carbon-based nanomaterials (*i.e.* carbon black,⁹ carbon nanotubes,^{10,11} and graphene^{12,13}), metal oxides,^{14–18} and 2D transition metal di-chalcogenides.¹⁹ Despite considerable progress, it is still required to find a new material to develop NO₂ gas sensors which can operate at room temperature with excellent performance/stability and at the same time is easy to miniaturize for on-chip integration.

Organic–inorganic metal halide perovskite (OIMHP) materials have recently attracted tremendous interest for various optoelectronic applications especially in solar cells and light-emitting diodes (LEDs). This is due to the simple deposition process and excellent optoelectronic properties including the sharp optical absorption edge and low non-radiative recombination loss leading to remarkable photoluminescence quantum efficiency of the OIMHP materials.^{20–22} Although the sensitive nature of OIMHP materials creates challenges for the fabrication of stable solar cells and LEDs,^{23,24} this property makes OIMHP materials good candidates for chemical sensing applications.²⁵ As a result, OIMHP materials have recently been investigated for the detection of oxygen,²⁶ ozone,²⁷ multiple volatile organic compounds (VOCs) *i.e.* acetone, ethanol, and ammonia,^{26,28} and NO₂ gas.^{29–31} In the chemical sensing applications, the stability of the sensor devices in an ambient environment with the presence of moisture is the main challenge for their practical implementation given that the sensors

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cannot be encapsulated. Therefore, it is essential to develop OIMHP based gas sensors with superior ambient stability.

In this work, we employed a dimensional engineering strategy to tackle the ambient stability of OIMHP based gas sensors for the detection of NO₂. A large organic cation *n*-octylammonium bromide was applied to passivate the three-dimensional (3D) perovskite,^{32,33} resulting in a mixed-dimensional 2D/3D perovskites with greatly improved resistance to moisture degradation. By virtue of the heterojunction structure, this 2D/3D perovskite based gas sensor also shows much enhanced sensor response toward NO₂ gas compared to its 3D counterpart. As a result, the mixed-dimensional OIMHP based gas sensors not only show excellent performance for NO₂ detection but also demonstrate very good ambient stability, paving a way for practical applications.

Materials and methods

Materials

Formamidinium iodide (FAI), formamidinium bromide (FABr), methylammonium iodide (MAI), methylammonium bromide (MABr), and *n*-octylammonium bromide (*n*-OABr) were ordered from GreatCell Solar Materials. Other materials were ordered from Sigma Aldrich.

Methods

Fabrication of perovskite sensors. Indium-doped tin oxide (ITO) glass substrates were cut into 10 mm × 10 mm pieces. Isolation lines were created using a laser to separate two electrodes on a single substrate. The substrates were cleaned with detergent for 90 minutes and for 15 minutes in acetone, 2-propanol, and ethanol sequentially. The substrates were further cleaned using UV ozone for 30 minutes before being immediately transferred to a N₂ purging glove box. 1 ml of 3D perovskite precursor was prepared by mixing 0.75 ml of FAPbI₂Br (1.3 M in DMF/DMSO 4:1 volume ratio), 0.15 ml MAPbI₂Br (1.3 M in DMF/DMSO 4:1 volume ratio), 0.1 ml CsPbI₂Br (1.3 M in DMSO) and 0.05 ml RbI (1.3 M in DMSO). The 3D perovskite was deposited on the substrates by spinning 40 μL of the solution at 1000 rpm for 10 s (acceleration rate 100 rpm s⁻¹) and 4000 rpm (acceleration rate 1000 rpm s⁻¹) for 20 s. 5 s before the program ends, 150 μL of chlorobenzene was quickly dropped in the middle of the substrates. This was then followed by annealing at 100 °C on a hotplate for 30 minutes. 90 μL of *n*-octylammonium bromide diluted in 2-propanol with different concentrations (2.5 mg ml⁻¹, 5.0 mg ml⁻¹, and 10.0 mg ml⁻¹) was dynamically spun on top of the 3D perovskite substrates at 5000 rpm for 30 s. The substrates were further annealed at 100 °C for 5 minutes. Au metal contact was deposited on the substrates by thermal evaporation through a shadow mask.

Characterization

Material characterization. Scanning electron microscopy (SEM) imaging was performed on an FEI Verios system. Grazing

incidence X-ray diffraction (GIXRD) was measured with a high-resolution PANalytical X'Pert PRO MRD system with a step size of 0.005° and an integration time of 1 s per step. The absorbance was measured with a Lambda 1050 spectrophotometer (PerkinElmer) in an integrating sphere mode. Transmission electron microscopy (TEM) analyses were performed using a JEOL 2100F instrument. Steady state photoluminescence measurements and time-resolved photoluminescence measurements were performed using a LabRAM HR Evolution system with a time-correlated single photon counting (TCSPC) system (DeltaPro-DD, Horiba). A 508 nm diode laser (DD-510L, Horiba) with a pulse duration of 110 ps and a fluence of ~10 μJ cm² per pulse, and a repetition rate of 312.5 kHz was used for excitation. Photoluminescence images of perovskite films were taken following the procedure reported in our previous work.^{34,35}

Chemiresistive gas sensing characterization. NO₂ (10 ppm in N₂ (Coregas)), ethanol (10 ppm in N₂ (Coregas)), acetone (10 ppm in N₂ (Coregas)), propane (10 ppm in N₂ (Coregas)), ethyl benzene (10 ppm in N₂ (Coregas)), and carbon dioxide (10 ppm in N₂ (Coregas)) diluted in simulated air (2 L min⁻¹ O₂ + 0.8 L min⁻¹ N₂ (BOC Ltd)) were controlled using a mass flow controller (Bronkhorst) to obtain the desired concentration range (2–8 ppm) while the total flow rate of the gas was kept constant at 1 L min⁻¹. The temperature of the hotplate in the gas sensing chamber (Linkam) was kept constant at 25 °C using a temperature controller. The dynamic responses were recorded using an electrochemical workstation (CHI 660E, USA) at a constant applied voltage of 1 V. The sensor response is calculated based on the following formula:

$$S = \left[\frac{I_{\text{analyte}}}{I_{\text{air}}} \right] - 1$$

where I_{analyte} and I_{air} are the current of the sensor under target gas and simulated air conditions respectively. The sensor sensitivity is defined as the slope of the graph representing the change of the sensed current vs. the gas concentration. The response time and recovery time are the time required for the sensed current to change to 90% of its saturated value since the time the gas is on and off, respectively. Humidity tests were performed by supplying water vapour into the system *via* an air flow through a bubbler.

Results and discussion

Fig. 1a illustrates the steps to fabricate mixed-dimensional perovskite films. First, the 3D perovskite is deposited on the substrates by spin coating and the black 3D perovskite phase is formed after annealing at 100 °C for 30 minutes. Second, a passivation solution containing the long alkylammonium chain organic cation *n*-octylammonium bromide (*n*-OABr) C₈H₂₀NBr layer is deposited on top of the 3D perovskite film to form a mixed 2D/3D perovskite film after short annealing at 100 °C for 5 minutes. Fig. 1b shows a simple top-bottom electrode configuration of a perovskite sensor device. Given that all the layers can be deposited at temperature lower than 150 °C, this



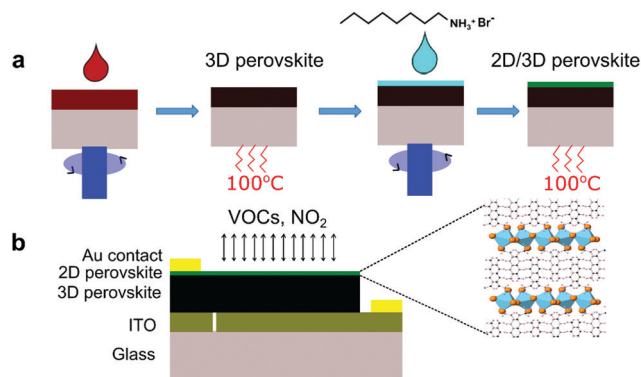


Fig. 1 (a) Scheme showing the fabrication steps for the mixed-dimensional perovskite films on ITO glass substrates. (b) Top-bottom electrode configuration of perovskite sensor devices for NO₂ detection.

structure can be potentially implemented on flexible substrates as wearable sensors.

Improvement in the performance of mixed-dimensional perovskite-based gas sensors

The gas sensing performance of the 3D and 2D/3D perovskite sensors was tested toward various analyte gases namely NO₂, ethanol, acetone, propane, and ethyl benzene. As shown in Fig. 2a, the 3D perovskite sensor device exhibits a good sensor response to NO₂ accompanied by the increase in the current during the injection of 2–8 ppm of NO₂ gas at room temperature. The corresponding response and recovery times were measured to be 11.9 s and 12.9 s, respectively (Fig. S1a, ESI†). In the 3D perovskite composition, the primary and secondary amines are present in the formamidinium cation (CH(NH₂)₂) and primary amine in the methylammonium cation (CH₃NH₂). Therefore, it is expected that the electron donor interaction among NO₂ gas with the amine groups in the 3D perovskite is responsible for the working mechanism as reported previously.³⁶ In comparison, the 2D/3D perovskite sensor device shows greatly enhanced sensor response and the current increases by up to 50 times while being exposed to 8 ppm of NO₂ (Fig. 2b). Based on the sensor response curves shown in Fig. 2c and d, the sensitivity of the 2D/3D perovskite sensor was calculated to be 6.3 ± 0.83 times per ppm as compared to the sensitivity of 1.25 ± 0.16 times per ppm in the control 3D-only device. The response time and recovery time in the 2D/3D perovskite sensor device are reduced to 5.7 s and 12.7 s, respectively (Fig. S1b, ESI†). The 2D/3D perovskite device was tested toward sub-ppm NO₂ concentrations. The device shows the response of 0.04 and 0.19 toward 0.2 ppm and 0.4 ppm NO₂, respectively (Fig. S2, ESI†). As summarized in Table S1 (ESI†), the 2D/3D perovskite sensor in this work has comparable responsivity and slightly faster response/recovery time compared to room temperature operation metal oxide-based NO₂ gas sensors. Although metal oxide based NO₂ sensors generally show a very low detection limit, we note that many of them require light activation. We tested the 2D/3D perovskite devices toward 8 ppm NO₂ in the presence of light and directly compared that with the sensor response in the dark. The device was

illuminated through a quartz window using a solar simulator (NewSpec, LCS-100) with an intensity of 100 mW cm⁻². We found that the responsivity of the device was significantly reduced in the presence of light (Fig. S3, ESI†). Previous reports have demonstrated that when chemiresistive sensors with light sensitive active layers operate under light, the light-activated adsorption processes compete with a light-activated desorption around the oxygen adsorbates.^{37,38} This resulted in complex changes in the response of the sensors depending on many factors such as the properties of the active layer, adsorption and photodesorption rates, etc. In this case, the presence of light significantly reduces the responsivity of the perovskite-based NO₂ sensor. The exact mechanism underlying this phenomenon requires further investigation beyond the scope of this work. It is worth noting that the response in the 2D/3D sample gradually reduces after reaching saturation, when the gas is on, which would require further investigation. Nevertheless, 2D/3D perovskite sensors demonstrate much better sensor response than the 3D perovskite sensor device. In addition, both 3D and 2D/3D perovskite sensors show very good selectivity toward NO₂ gas after being tested with other analyte gases such as ethanol, acetone, propane, ethyl benzene, and carbon dioxide (Fig. 2e and f).

Investigation on the surface morphology, phase formation and electronic properties

To investigate the enhanced sensor performance, the morphology changes of the 3D perovskite layer upon passivation with the long alkyl chain organic cation *n*-OABr to form 2D/3D perovskites was characterized using scanning electron microscopy (SEM). Fig. 3a presents the surface morphology of the 3D perovskite showing a compact layer with the grain size ranging from 100 nm to 500 nm. It is worth noticing that the bright features on the surface of the films are related to a Rb-rich phase as in previous reports.^{39,40} With a low concentration of 2.5 mg ml⁻¹ for the passivation precursor, the morphology of the perovskite film remains unchanged except that a very thin layer of new substance at the top of the film is evidenced by the changes in the SEM image contrast (Fig. 3b). As the concentration of the passivation increases to 5.0 mg ml⁻¹ and 10.0 mg ml⁻¹, the top layer becomes more obvious, indicating the formation of a new layer on the surface of the 3D perovskite (Fig. 3c and d). Energy dispersive X-ray spectroscopy (EDS) measurements were performed on the 3D and 2D/3D (with 5.0 mg ml⁻¹) perovskite films. As shown in Note S1 (ESI†), the 2D/3D perovskite film surface shows a higher atomic ratio of C and N compared to the 3D perovskite film surface. In addition, the atomic ratio of Pb in the 2D/3D perovskite film surface is significantly lower than in the 3D perovskite film surface. Grazing incidence X-ray diffraction (GIXRD) in Fig. 3e shows the formation of a 2D phase at $2\theta = 3.55^\circ, 7.1^\circ, 10.65^\circ$ and 12.6° . These phases are related to the Ruddlesden–Popper 2D perovskite *n*-OA₂Pb(I_xBr_{1-x})₄, which forms due to the reaction between the *n*-OA cation and the under-coordinated Pb on the 3D perovskite surface.⁴¹ When the concentration increases, the 2D perovskite peaks become more prominent indicating that a thicker 2D layer is formed. To determine the thickness of the 2D layer when the concentration of the



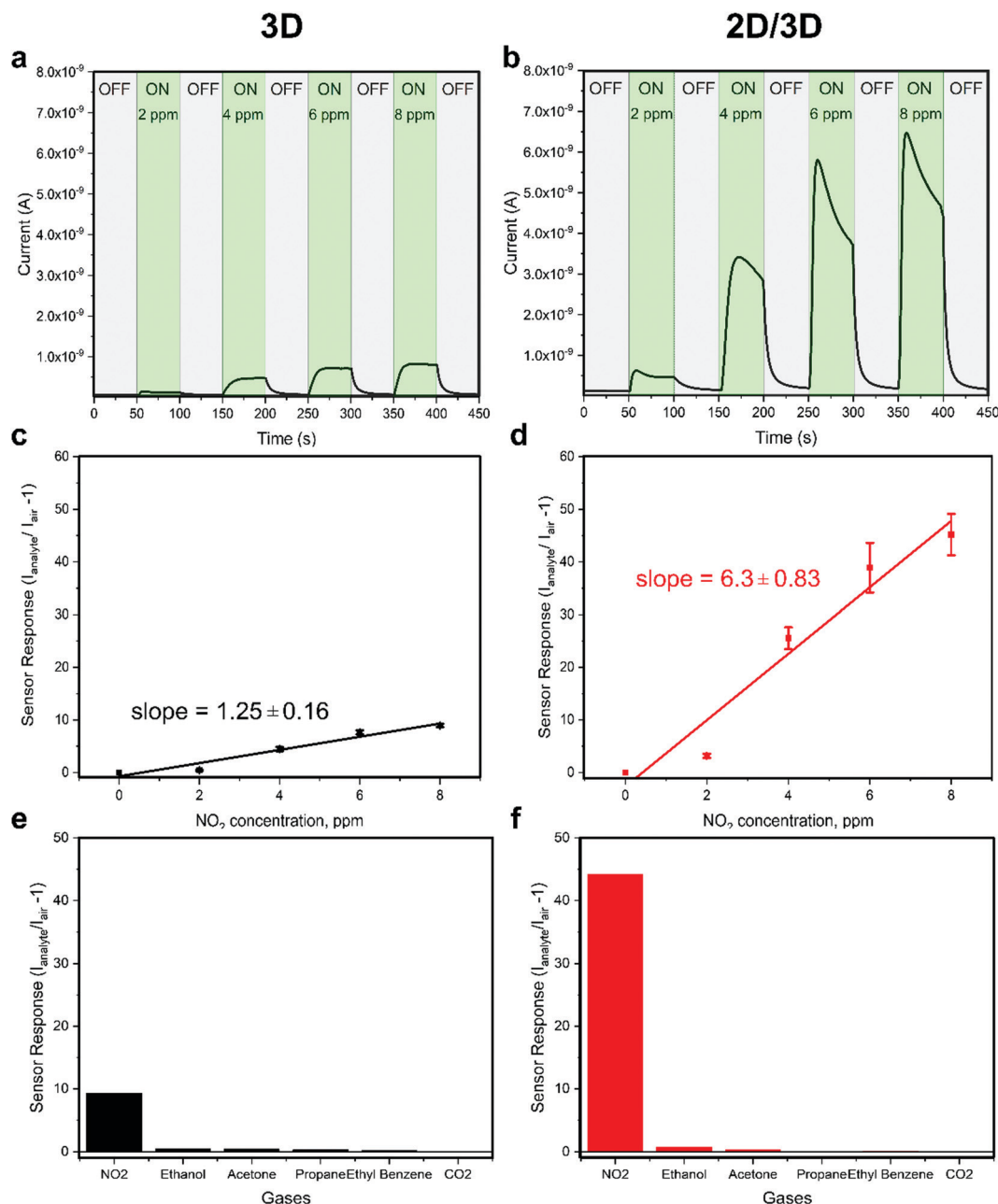


Fig. 2 Current responses of sensor devices with (a) 3D perovskite and (b) 2D/3D perovskite gas sensors under consecutive injections of NO_2 gas with concentrations ranging from 2 to 8 ppm. Sensor responses of (c) 3D perovskite and (d) 2D/3D perovskite sensors as a function of the NO_2 concentration. Selectivity testing of (e) 3D perovskite and (f) 2D/3D perovskite-based sensors under a group of gases including NO_2 , ethanol, acetone, propane, ethyl benzene, and carbon dioxide (with a concentration of 8 ppm).

passivation solution is 2.5 mg ml^{-1} , we used transmission electron microscopy (TEM) to examine the cross-section of the 2D/3D sensor device. Although the 2D layer is ultrathin and it cannot be noticed in the low-magnification bright-field TEM (BF-TEM) image in Fig. 3f, it is evident that multiple phases with different interplanar spacings are present in the high-resolution TEM (HR-TEM) image in Fig. 3g. One phase with an interplanar spacing of 6.834 \AA is at the top surface layer. This phase is related to the new 2D peak at $2\theta = 12.6^\circ$. The location and size of this phase indicate that the thickness of the 2D layer is around

20 nm. With such a thin layer, we have shown that the 2D perovskite does not completely cover the 3D perovskite underneath, which leads to local passivation and enhanced performance in perovskite solar cells.⁴² Another phase with a much smaller interplanar spacing of 4.428 \AA in the bulk of the film comes from the 3D perovskite. The SEM, GIXRD and TEM results confirm that a thin layer of pure 2D perovskite $n\text{-OABr}$ is formed after passivating the 3D perovskite with $n\text{-OABr}$. As illustrated in Fig. 1b, the crystal structure of the pure 2D perovskite is comprised of bulky organic cation $n\text{-OA}$ sandwiched

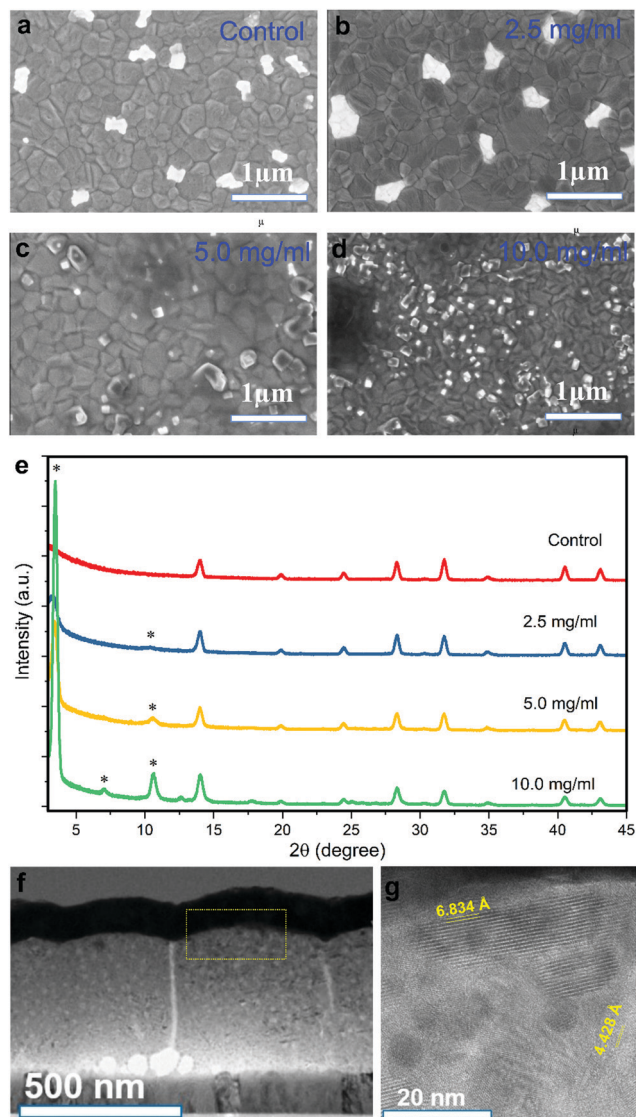


Fig. 3 Top-view scanning electron microscopy (SEM) images of perovskite films (a) control 3D perovskite film, film passivated with (b) 2.5 mg ml⁻¹, (c) 5.0 mg ml⁻¹ and (d) 10.0 mg ml⁻¹ of passivation solution. (e) Grazing incidence X-ray diffraction (GIXRD) of perovskite films with and without passivation layers. The * symbol represents the 2D perovskite phase. (f) Cross-sectional low-magnification bright-field transmission electron microscopy (BF-TEM) image of a passivated device with a concentration of 2.5 mg ml⁻¹. (g) High-resolution TEM (HRTEM) images showing the characteristic interplanar spacing of 2D and 3D perovskites.

between two inorganic sheets and they are held together by Coulombic forces.⁴³ This 2D structure may allow faster diffusion of NO₂ molecules into the perovskite to better interact with primary amine groups in the *n*-OA cation compared with the 3D-only perovskite device. Therefore, we hypothesize that the formation of this 2D perovskite layer may contribute to the enhancement in the sensor response toward NO₂ gas. We note that while the response time of the 2D/3D device is significantly shorter than the response time of the 3D device, the recovery time of the 2D/3D device is only slightly lower than the recovery time of the 3D device. We speculate that the desorption kinetics

of NO₂ molecules in the 3D and 2D/3D perovskites are similar. Since the recovery of the device mainly comes from the desorption process when NO₂ gas is pumped out,⁴⁴ a very close recovery time on the 2D/3D and 3D devices is obtained. Because this phenomenon is quite complicated, it requires further research work especially in the molecular simulation to deeply investigate it. When the concentration of the passivation layer is 10.0 mg ml⁻¹, the 2D layer thickness increases to approximately 50 nm as determined by the cross-sectional SEM results in Fig S4 (ESI[†]), and the 2D layer fully covers the 3D layer underneath. We tested the performance of 2D/3D perovskite sensor devices with high concentration of the passivation solution (5.0 mg ml⁻¹ and 10.0 mg ml⁻¹). As shown in Fig. S5 (ESI[†]), the response becomes smaller than the control device without any passivation solution. This is due to the non-conductive property of the 2D layer coming from its horizontal crystal plane arrangement, which hinders the current flow in the device when the 2D layer becomes excessively thick and fully covers the 3D layer underneath.^{45,46} Thus, the 2D/3D perovskite sensor reaches the optimal performance with a low concentration of 2.5 mg ml⁻¹ for the passivation precursor.

To further investigate the enhancement performance in the 2D/3D perovskite sensor device, steady state photoluminescence (PL) and time-resolved PL measurements were used to examine the electronic properties of the mixed-dimensional perovskite films. As shown in Fig. 4a and b, the PL intensity is strongly enhanced in the 2D/3D perovskite as compared to the 3D perovskite. In addition, the charge carrier lifetime is also greatly improved in the 2D/3D perovskite. We also examined the PL images of the 3D and 2D/3D perovskite films as presented in Fig. 4c and d, respectively. The 2D/3D perovskite shows much stronger and more uniform PL intensity than the 3D perovskite. This indicates that the 2D layer has passivated the defects on the 3D perovskite surface, hence reducing the recombination loss due to those defects. As the recombination centres might reduce the current flowing through the device in the presence of NO₂ gas. It is highly possible that the reduction in recombination loss in the 2D/3D perovskite would contribute to the enhanced gas sensing current in the 2D/3D perovskite sensor as compared to that of the 3D counterpart (Fig. 2a and b).

Improvement in the moisture stability of mixed-dimensional perovskite-based gas sensors

One of the major issues related to the practical use of perovskite-based sensors is their ambient stability. We first tested the short-term operation of both the 3D and 2D/3D perovskite sensors toward 8 ppm of NO₂ gas in a 40% relative humidity (RH) environment. As shown in Fig. S6 (ESI[†]), both devices operate reliably in the 40% RH environment for over 30 minutes. However, the response of the 2D/3D device to 8 ppm NO₂ in the 40% RH environment is similar to the response in the dry environment, while the response of the 3D device to 8 ppm NO₂ in the 40% RH environment is significantly lower than the response in the dry environment. This indicates that the 2D/3D perovskite NO₂ sensor is less

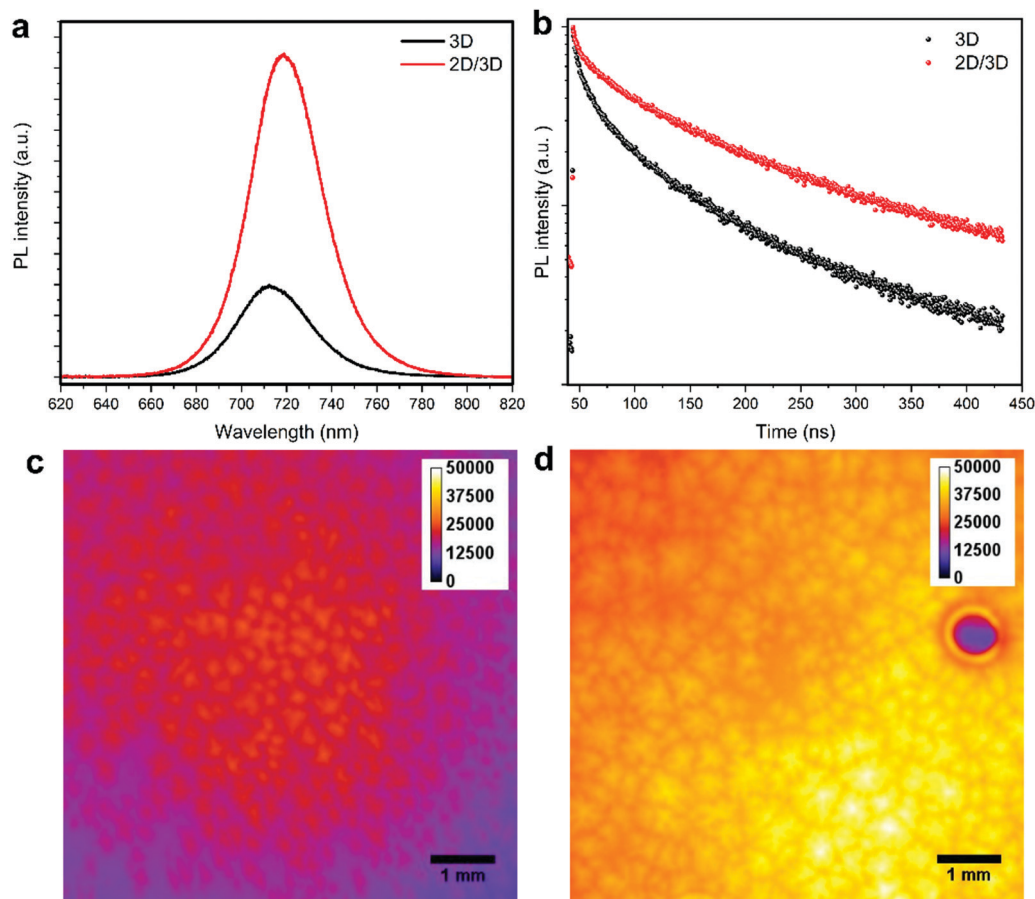


Fig. 4 (a) Steady state photoluminescence (PL) of perovskite films. (b) Time-resolved PL of perovskite films. PL images of a (c) control perovskite film and (d) passivated perovskite film.

affected by the presence of moisture compared to the 3D perovskite device. The long-term stability of the fabricated sensor devices by storing them under ambient conditions with relative humidity (RH) ranging from 30–80% was examined. As shown in Fig. 5a, the absorbance of the 3D perovskite film is quickly decreased after 6, 14 and 22 days of storage. In addition, the absorbance spectrum develops a sharp absorbance peak at ~ 450 nm, which indicates the formation of the PbI_2 phase. This irreversible degradation is due to hydrolysis reactions between water molecules with the organic species in the perovskite composition, releasing the volatile organic compounds (methylammonium CH_3NH_3 , formamidinium $\text{CH}(\text{NH}_2)_2$, HI, and HBr) and finally converting the perovskite black phase to the PbI_2 yellow phase.^{47,48} It is essential to tackle this degradation issue in perovskite-based sensors since the OIMHP perovskite is the active material for the detection of NO_2 and the devices cannot be encapsulated. In comparison, the absorbance spectra of the 2D/3D perovskite film only show minor changes over the same period. Fig. 5c shows the normalized absorbance of the perovskite films at a wavelength of 660 nm over time, which clearly exemplifies the superior ambient stability of the 2D/3D perovskite film over the 3D-only perovskite film. The inset of Fig. 5a and b displays the contact angle of the 3D and 2D/3D perovskite films

to water. The contact angle of the 3D perovskite film is 45.8° , while the contact angle of the 2D/3D perovskite film greatly increases to 90.6° . This indicates that the perovskite film becomes hydrophobic after the deposition of a top 2D layer, and this explains the great improvement in the ambient stability of the 2D/3D perovskite film. The hydrophobicity of the 2D layer comes from the large organic cations with a hydrocarbon chain surface which separate the water from the underlying 3D perovskite layer.⁴⁹ The long-term stability of the 3D and 2D/3D perovskite sensors was tested, and the normalized sensor response of the devices is shown in Fig. 5d. It can be observed that in the first 4 days, both devices degrade greatly although the 2D/3D sensor is slightly more stable than the 3D counterpart. The initial degradation in the sensor response of both devices could be ascribed to the early-stage decomposition of perovskite into PbI_2 as demonstrated in Fig. 5a and b. The release of organic compounds (methylammonium and formamidinium) with amine groups reduces the interaction between NO_2 gas and the active perovskite layer. After 5 days, the response of the 2D/3D device become stable until day 22 while the response of the 3D sensor device kept decreasing and the device stopped working after 13 days under ambient conditions. The 2D/3D sensor device was continuously tested, and the device



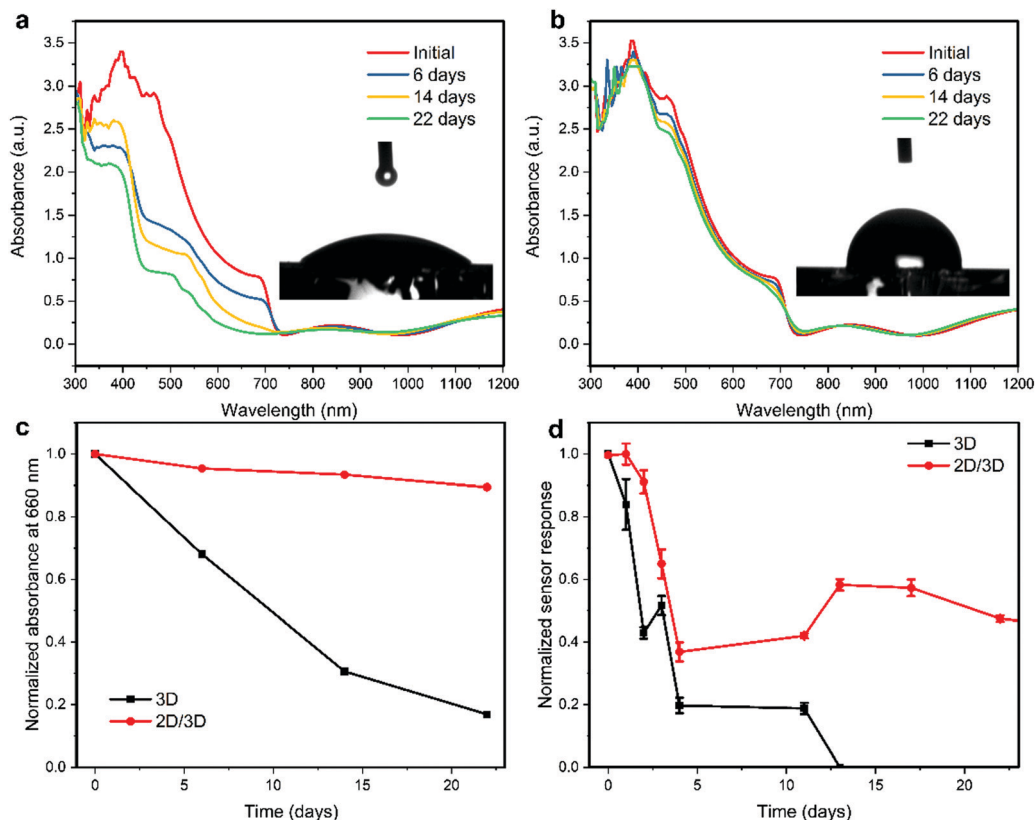


Fig. 5 The absorbance spectra of the (a) 3D perovskite film and (b) 2D/3D perovskite film under ambient conditions with the relative humidity ranging from 30% to 80% for up to 22 days under ambient conditions. (c) Normalized absorbance of the perovskite films at 660 nm over 22 days. (d) Normalized sensor response of 3D and 2D/3D perovskite NO₂ gas sensors under ambient conditions for 22 days.

still shows good response after 57 days, retaining almost 25% of the initial response (Fig. S7, ESI†). The results demonstrate that the mixed-dimensional 2D/3D perovskites greatly improve the ambient stability of perovskite-based gas sensors.

Conclusions

In summary, we have presented a strategy to improve both the performance and stability of perovskite NO₂ gas sensors. By passivating 3D perovskite films with a long alkyl chain organic cation n-octylammonium bromide to form mixed-dimensional 2D/3D perovskite films, the performance and ambient stability of perovskite-based NO₂ gas sensors are greatly enhanced. This can be ascribed to the formation of the ultrathin 2D layer on the surface of the 3D perovskite, which increase the interaction between the NO₂ gas and the amine groups in both the 3D and 2D perovskites. In addition, the 2D layer passivates the defects on the 3D perovskite surface acting as recombination centres. The 2D/3D perovskite sensor shows good selectivity toward NO₂, excellent sensitivity of 6.3 ± 0.83 times per ppm and quick response/recovery times of 5.7 s and 12.7 s, respectively. Most importantly, the device demonstrates superior stability under ambient conditions for almost two months, shedding light on the design of stable perovskite based gas sensors for practical applications.

Author contributions

The Duong: conceptualization, methodology, investigation, writing – original draft; Alishba T. John: methodology, validation, formal analysis, writing – review & editing; Hongjun Chen: conceptualization, methodology, writing – review & editing; Huyen Pham: investigation, writing – review & editing; Krishnan Murugappan: investigation, writing – review & editing; Thanh Tran-Phu: investigation, writing – review & editing; Antonio Tricoli: resources, supervision, funding acquisition, writing – review & editing; Kylie Catchpole: resources, supervision, funding acquisition, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

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