



Cite this: *Mater. Adv.*, 2022,
3, 1101

Facet control of manganese oxides with diverse redox abilities and acidities for catalytically removing hazardous 1,2-dichloroethane†

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The preparation of four kinds of MnO₂, namely, α -, β -, γ -, and δ -type MnO₂, with distinct crystal phases and tunnel structures was achieved and they were applied for catalytic combustion of 1,2-dichloroethane (1,2-DCE). The redox ability and acidity of MnO₂ as well as the corresponding reaction mechanism were studied by means of various surface-sensitive techniques, including TPR, TPD, OIE, XPS, and *in situ* DRIFTS together with DFT calculations. The catalytic activities for 1,2-DCE combustion demonstrated that γ -MnO₂ displayed the most superior activity with the maximum HCl yield of 95% and CO₂ yield of 92% due to its abundant oxygen vacancies on the surface, easy formation of reactive oxygen species, and strong acidity to readily react with the adsorbed reactant to form HCl, which reflects the synergistic effect of its redox properties and acidity. However, the strong Mn–Cl bonding associated with high valence Mn⁴⁺ and the scarcity of acidic sites in β - and δ -MnO₂ hinder the HCl elimination process, even leading to the undesirable chlorine deposition associated with a series of polychlorinated byproducts including 1,1,2-C₂H₃Cl₃ and CCl₄ via a Cl substitution process. A mechanism for 1,2-DCE combustion involving dehydrochlorination via C–Cl bond cleavage at acidic sites as well as excess polychlorination by Cl₂ over a relatively oxidative sample was accordingly proposed.

Received 11th October 2021,
Accepted 9th November 2021

DOI: 10.1039/d1ma00943e

rsc.li/materials-advances

1 Introduction

In recent years, the situation of outdoor air pollution has become more and more serious. As major air pollutants, volatile organic compounds (VOCs) have attracted plenty of attention because of their contribution to the increase of ground-level ozone concentration and the formation of photochemical smog.^{1,2} Moreover, chlorinated volatile organic compounds (CVOCs) have high toxicity, volatility, low biodegradability and stability,³ and exhibit deleterious effect on the environmental ecosystem and human health.⁴ In order to eliminate CVOCs, many control technologies such as adsorption, condensation, biodegradation, non-thermal plasma, biological degradation, and catalytic combustion have been adopted.^{5–7} Among them, catalytic combustion is the most promising strategy owing to its high purification efficiency and product selectivity, energy-saving nature and less

secondary pollution. The challenge is the development of desirable catalysts for practical applications.^{8,9}

Noble-metal catalysts,¹⁰ zeolite catalysts,^{11,12} and transition-metal-oxide catalysts^{8,13} have previously been applied for the catalytic combustion of CVOCs. Noble-metal catalysts commonly show high activity at low temperatures, but their large-scale application is limited by their high cost and sensitivity to chlorine poisoning.¹⁴ The abundant acid sites of zeolite catalysts can promote the dechlorination of CVOCs, but it can easily cause carbon deposition.¹⁵ Transition-metal-oxide catalysts are also employed for catalytic combustion of CVOCs due to their considerable activity, easily-controllable nanostructure and attractive resistance towards chlorine poisoning.^{16,17} A yolk-shell-like mesoporous CoCrO_x was prepared for deep destruction of dichloromethane and it was found that the strong interaction of Co and Cr elements can generate active Cr⁶⁺ and oxygen species, which enhanced the catalytic activity.¹⁸ Although the Cr component shows good catalytic activity for purifying CVOCs, its application could cause potential harm to animals and humans.¹⁹ Among many transition-metal catalysts, MnO_x possess excellent redox properties and adjustable valence states, and have attracted much attention in CVOC combustion.²⁰ Two α -Mn₂O₃ single crystals exposed with the (1 1 1) facet and exposed with both (1 1 1) and (0 0 1) facets were prepared by Fan *et al.* and

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ma00943e



applied for the selective catalytic reduction (SCR) of NO_x with NH_3 , and it was found that the charge redistribution on the surface of the (0 0 1) facet, which produced more adsorption sites, benefited further adsorption and reaction of NH_3 and NO molecules and improved the catalytic activity.²¹ Wang *et al.* prepared a series of MnO_x including $\gamma\text{-MnO}_2$, Mn_2O_3 and Mn_3O_4 , which were used for catalytic oxidation of benzene and 1,2-DCE, and noted that $\gamma\text{-MnO}_2$ exhibited remarkable activity ascribed to the abundant surface adsorbed oxygen species and ideal low-temperature reducibility.²² According to the different tunnel structures of MnO_2 formed by MnO_6 octahedra connected by common vertices or edges linked into a single chain or double chain structure,²³ the crystal phase of MnO_2 can be divided into a one-dimensional (α -, β - and γ -type) tunnel structure and a two-dimensional (δ -type) layer structure. Li *et al.* improved the MnO_2 activity via *in situ* fabrication of a three-dimensional (3D) macroporous $\gamma\text{-MnO}_2/\text{LaMnO}_{3.15}$ and revealed that the outstanding activity may be ascribed to the increased BET surface area and the enhanced active reactive oxygen species quantity after $\gamma\text{-MnO}_2/\text{LaMnO}_{3.15}$ modification.²⁴ Zhang *et al.* elaborated the mechanism of catalytic combustion of 1,2-dichloropropane over the LaMnO_3 perovskite through experimental and theoretical studies, indicating that further decomposition was strongly dependent on the active reactive oxygen species and the rate-determining step for catalytic oxidation of 1,2-dichloropropane was decided by the reaction conditions.²⁵ Wang *et al.* introduced CeO_2 and HF etching to modify LaMnO_3 and found that it could resist the catalyst deactivation, which showed that the loaded CeO_2 provided enriched oxygen vacancies to retain the redox ability and the fluoride ions replaced the surface lattice oxygens to promote the mobility of oxygen species.²⁶ Weng *et al.* explored the influence of phosphoric acid etching and water vapor on dichloromethane elimination over $\text{La}_3\text{Mn}_2\text{O}_7$ and found that the former enhanced dichloromethane hydrolysis and the latter promoted desorption of accumulative chlorine with elevated temperatures.²⁷ However, the study on CVOC combustion over MnO_2 with different crystal planes has rarely been reported. What is more, how to improve the anti-poisoning ability of catalysts and avoid deactivation and production of more harmful by-products is still a challenge in industrial applications.

As a kind of typical CVOCs, 1,2-dichloroethane (1,2-DCE) is often used as a solvent, extractant and raw material for organic synthesis and its catalytic combustion over MnO_2 with diverse crystal structures has not been explored. The preparation of α -, β -, γ -, and δ -type MnO_2 with different crystal phases and tunnel structures was achieved and used for catalytic combustion of 1,2-DCE in this work. What is more, a lot of characterization studies were carried out to reveal the relationship between their physicochemical properties and catalytic activity and the reaction mechanism was also studied.

2 Experimental

2.1 Preparation of MnO_2 catalysts

Four kinds of MnO_2 catalysts with distinct crystal phases and tunnel structures were successfully prepared. All reagents used

in this work were analytically pure and used without any treatment. The typical synthesis methods are described in the ESI.†

2.2 Catalyst characterization

The crystallographic information was investigated by XRD using a diffractometer equipped with $\text{Cu K}\alpha$ radiation (Bruker, Germany). The specific surface areas (S_{BET}) were determined using N_2 sorption isothermal analysis (Quantachrome, USA) based on Brunauer–Emmett–Teller (BET) theory. TPR and TPD experiments were carried out on a device with a thermal conductivity detector (TCD, Huasi, China). SEM (S-4800, Hitachi, Japan) was used to observe the morphologies of the prepared MnO_2 solids under a 5.0 kV accelerating voltage, and HRTEM (JEM-2100, JEOL, Japan) was employed to observe the microstructures under a 200 kV operating voltage. XPS instrument (Shimadzu, Japan) equipped with a monochromatic $\text{Mg/Al K}\alpha$ X-ray source was used to analyze the superficial elemental properties. The *in situ* DRIFTS experiments were conducted with a spectrometer (Bruker, Germany) equipped with a mercury cadmium telluride (MCT) detector. The details are given in the ESI.†

2.3 Temperature programmed oxygen isotopic exchange (TPOIE) reaction

TPOIE has been used to evaluate oxygen mobility in a wide temperature range.^{28–30} In this study, the simple heterogeneous oxygen isotopic exchange (OIE) mechanism was involved over the as-prepared samples, and the experimental device, operation method and schematic diagram are depicted in the ESI.†

2.4 Density functional theory (DFT) calculations

During the theoretical simulation, the slab models were reconstructed using a $[2 \times 2 \times 1]$ supercell, and the slabs were periodically repeated with a vacuum spacing of 15 Å for oxygen vacancy formation. All slab calculations were performed using a Monkhorst–Pack grid ($3 \times 3 \times 1$ mesh) with a cutoff energy of 400 eV and the smearing parameter set to 0.1 eV. The constraint value was set to 5×10^{-6} energy charge between two consecutive iterations to achieve SCF convergence. The computational details are given in the ESI.†

2.5 Catalytic activity evaluation

The catalytic performance was investigated on a catalyst bed with a quartz tube (i.d. = 4 mm). 0.14 g of catalyst (40–60 mesh) and 1000 ppm of 1,2-DCE under a 100 mL min^{-1} airstream (20 vol% O_2 ; N_2 balanced) with a gas hourly space velocity (GHSV) of $42857 \text{ mL g}^{-1} \text{ h}^{-1}$ were used in each test. The reaction was performed from 100 to 500 °C in a step mode with a 30 min plateau at each temperature. The outlet gases were measured with an online gas chromatograph (Agilent 7890B) equipped with a TCD and a mass spectrometer (Agilent 5977A).



3 Results and discussion

3.1 Structural and morphological characterization

As illustrated in Fig. 1, it is clearly observed that MnO₂ catalysts are well indexed to the crystal phases corresponding to JCPDS Card Numbers 44-0141 (α -), 24-0735 (β -), 14-0644 (γ -) and 80-1098 (δ -), respectively. This confirms that the expected MnO₂ phases were successfully obtained. β -MnO₂ had better crystallinity originating from the narrow peak width and high intensity. As reported, β -MnO₂ is favorable for grain growth due to its regular arrangement of [MnO₆] octahedra.²³ The peaks of γ -MnO₂ were relatively wide, suggesting it possessed lower crystallinity. γ -MnO₂ has been reported to be ordinarily composed of hybrid phases in which pyrolusite and ramsdellite intergrow irregularly.³¹

The type and size of tunnels, BET surface areas (S_{BET}), and pore volumes of the four catalysts are shown in Table 1 and Fig. 2. It is reported that the different tunnel structures assembled by MnO₆ octahedra in MnO₂ will affect the related surface areas and pore volumes.³² α - and β -MnO₂ presented relatively low specific surface areas (35.9 and 29.9 m² g⁻¹, respectively) and pore volumes (0.207 and 0.191 cm³ g⁻¹, respectively), while γ - and δ -MnO₂ showed higher specific surface areas (87.4 and 86.5 m² g⁻¹, respectively) and pore volumes (0.313 and 0.271 cm³ g⁻¹, respectively). Among them, γ -MnO₂ has the largest specific surface area and pore volume. What is more, the nitrogen adsorption-desorption isotherms of γ - and δ -MnO₂ display a type V curve with H3 type hysteresis loops, which implies the existence of mesoporous structures in these two samples. However, the wide relative pressure range of hysteresis loops of δ -MnO₂ suggests that pseudopores are formed by the secondary construction of lamellar particles.

The SEM and TEM images of the as-prepared samples are presented in Fig. 3. α -MnO₂ presents a linear nanoneedle morphology with a length of 2–2.5 μm and width of 25–30 nm, which are interspersed and staggered with each other. The well-identified periodic lattice fringes of 6.90 \AA size in

Table 1 Tunnel structures and sizes, BET surface areas, and pore volumes of the samples

Samples	Tunnel ³³	Size (\AA) ³³	S_{BET} (m ² g ⁻¹)	Pore volume (V) (cm ³ g ⁻¹)
α -MnO ₂	(1 \times 1), (2 \times 2)	1.89, 4.6	35.9	0.207
β -MnO ₂	(1 \times 1)	1.89	29.9	0.191
γ -MnO ₂	(1 \times 1), (1 \times 2)	1.89, 2.3	87.4	0.313
δ -MnO ₂	Interlayer distance	7.0	86.5	0.271

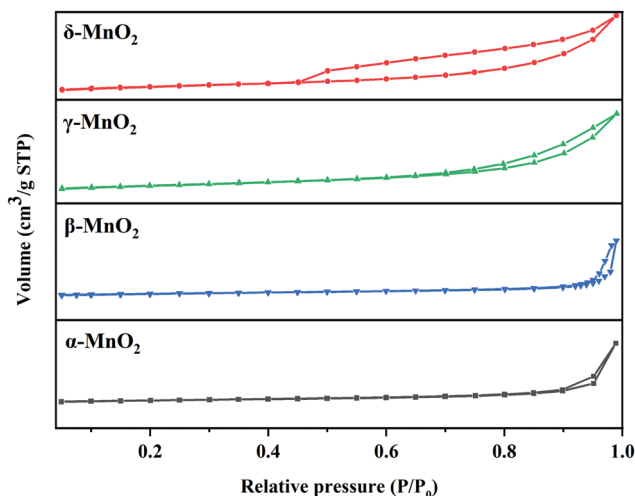


Fig. 2 The nitrogen adsorption-desorption isotherms of MnO₂.

Fig. 3(a4) correspond to the (1 1 0) facet of α -MnO₂. β -MnO₂ shows a nanorod-shaped structure with a similar length to that of α -MnO₂ and a width 2–5 times that of α -MnO₂. The well-identified periodic lattice fringes of 2.39 \AA size correspond to the (1 0 1) facet of β -MnO₂ in Fig. 3(b4). γ -MnO₂ presents a spherical nanostructure, which is assembled by interlaced sharp nanoneedles. The nanoneedles are \sim 30 nm in diameter and \sim 1 μm in length. The lattice fringes with distances of 3.98 \AA and 2.42 \AA match well the (1 3 1) and (1 2 0) facets in Fig. 3(c4), respectively. δ -MnO₂ shows a flower-globular nanostructure, which is different from the spherical structure of γ -MnO₂. It is made up of a large number of small nanosheets intertwined with each other, and the width of these nanosheets is about 15–25 nm. A typical interlayer spacing of 3.47 \AA in δ -MnO₂ can be observed in Fig. 3(d4).

3.2 H₂-TPR and O₂-TPD

The H₂-TPR profiles can be seen in Fig. 4a with H₂ consumption depicted in Table 2. Two reduction peaks of each sample emerge from the reduction of MnO₂ to Mn₂O₃/Mn₃O₄ and successive reduction to MnO, respectively.³⁴ The reduction peaks of α -MnO₂ are overlapped and centered at 307 and 354 $^{\circ}\text{C}$. The ratio of the H₂ consumption amount between the former reduction peak and the latter peak of α -MnO₂ was calculated to be about 1 : 1, which reveals that the reduction of α -MnO₂ might proceed as follows: MnO₂ \rightarrow Mn₂O₃ \rightarrow MnO. The reduction peaks of β -MnO₂ are centered at 318 and 417 $^{\circ}\text{C}$. The H₂ consumption

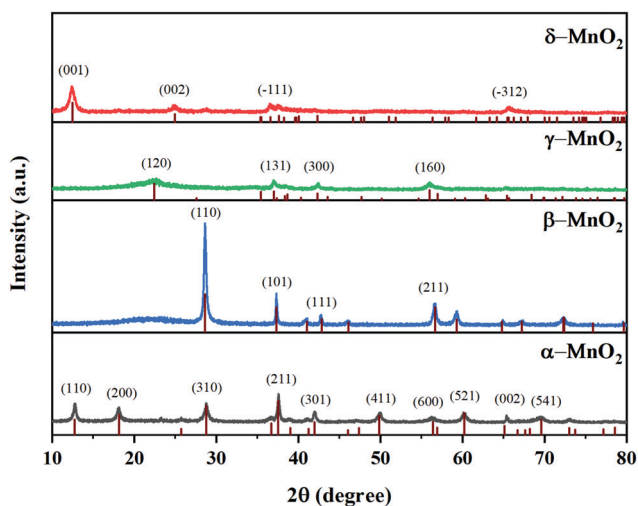


Fig. 1 The XRD patterns of MnO₂.



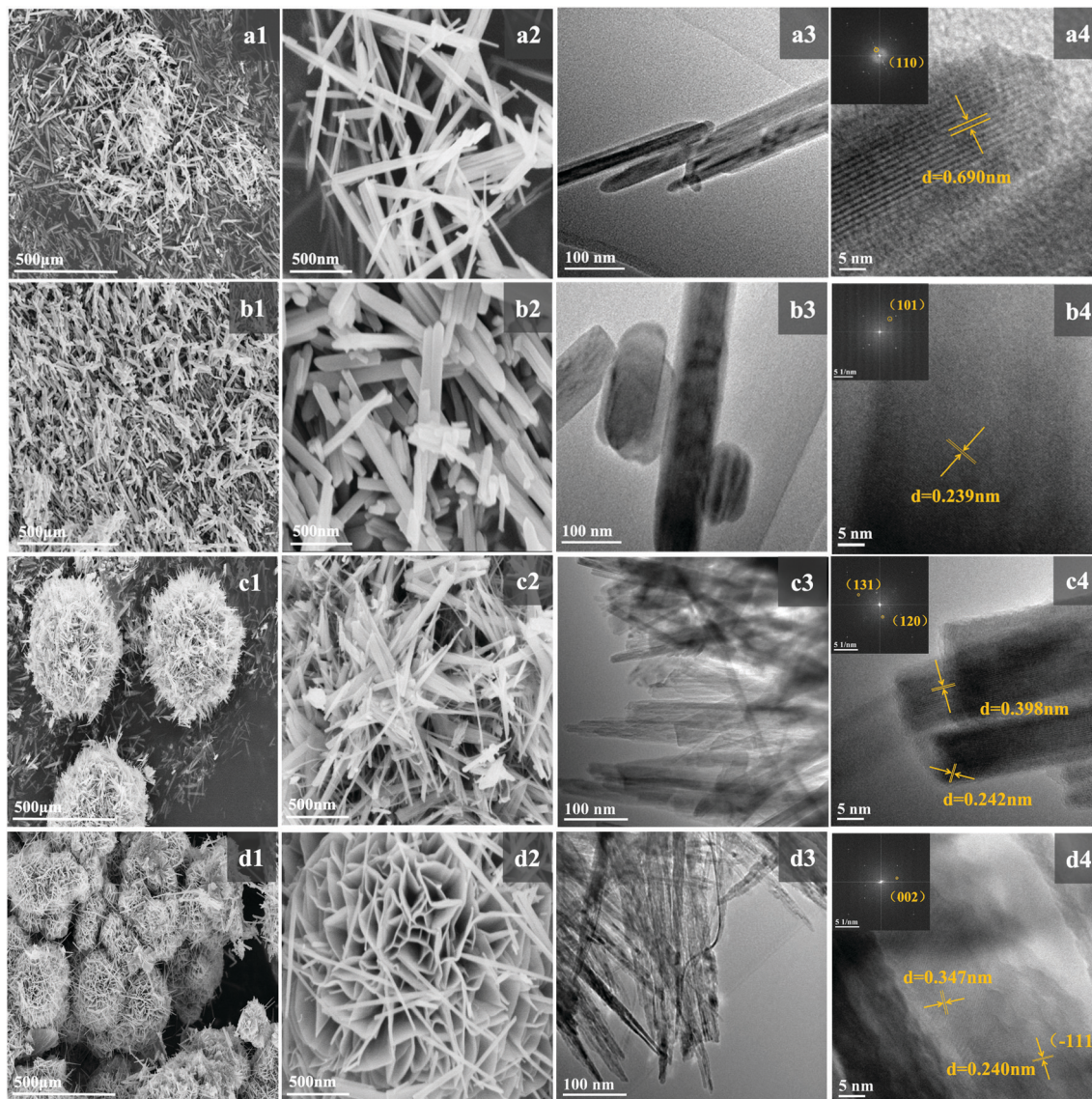


Fig. 3 The SEM and TEM images of (a) α -, (b) β -, (c) γ - and (d) δ - MnO_2 .

amount for the former peak is almost twice as large as that for the latter peak, suggesting that the reduction process of β - MnO_2 follows the $\text{MnO}_2 \rightarrow \text{Mn}_3\text{O}_4 \rightarrow \text{MnO}$ order. The H_2 -TPR profile of γ - MnO_2 exhibits a similar shape to the profile of β - MnO_2 , in which a H_2 consumption peak is centered at 284 °C with a hump at 396 °C. The ratio of H_2 consumption of the low-temperature peak to the high-temperature peak of γ - MnO_2 is about 2:1, which indicates that the reduction follows the $\text{MnO}_2 \rightarrow \text{Mn}_3\text{O}_4 \rightarrow \text{MnO}$ order. For δ - MnO_2 , two overlapped reduction peaks located at 288 and 331 °C were observed. The area proportion of the former peak to the latter peak is approximately 1:1, suggesting that δ - MnO_2 is initially reduced to Mn_2O_3 and then to MnO . According to the location of the initial reduction peak of the MnO_2 samples, the reducibility sequence is γ - $\text{MnO}_2 \approx \delta$ - $\text{MnO}_2 > \alpha$ - $\text{MnO}_2 \approx \beta$ - MnO_2 , which indicates that γ - and δ - MnO_2 have better redox properties. What stands out is that H_2 -TPR could reveal the migration of oxygen in the four catalysts because

γ - MnO_2 displays the lowest temperature reduction peak, and its lattice oxygen species are more mobile with respect to the other three MnO_2 samples. This possibly results in more oxygen being absorbed and their further excitation as surface reactive species to participate in the reaction, which is also reflected in the TPOIE reaction below (section 3.4).

The O_2 -TPD profiles are presented in Fig. 4b. It was widely accepted that active physically adsorbed oxygen can be eliminated from the surface at a low temperature, and thus we can confirm that the oxygen desorption peaks at the tested temperatures are categorized as chemisorbed oxygen species.³⁵ In the process of the reaction, molecular oxygen will be firstly absorbed on the surface of the catalysts and then converted into atomic oxygen or other surface reactive oxygen species, such as O^- , O^{2-} , etc. There are two peaks for α - and γ - MnO_2 , while there is only one peak for β - and δ - MnO_2 . MnO_2 surface was mainly covered by Mn^{3+} species.⁹ Considering γ - MnO_2 has



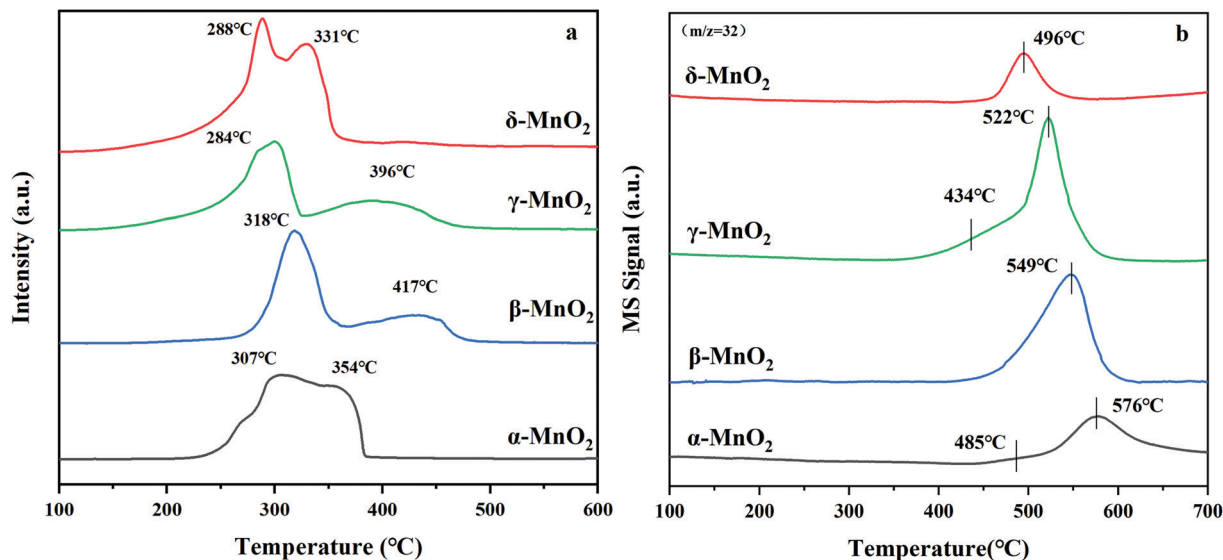


Fig. 4 The (a) H₂-TPR and (b) O₂-TPD profiles of MnO₂.

Table 2 The total H₂ consumption of MnO₂

Samples	Reduction peak (°C)		Total H ₂ consumption (mmol g ⁻¹)
	Position (°C)	Position (°C)	
α-MnO ₂	314	337	10.75
β-MnO ₂	318	417	10.82
γ-MnO ₂	284	396	10.73
δ-MnO ₂	288	331	10.94

the lowest temperature of the first desorption peak amongst all samples, it is deduced that γ-MnO₂ possesses superior oxygen mobility. That is to say, multitudinous surface oxygen vacancies on the surface of γ-MnO₂ may tend to form atomic oxygen or other surface reactive oxygen species, which contributes to further oxidation of 1,2-DCE.

3.3 XPS

The chemical status and surface composition were acquired by XPS. Fig. 5a displays the O 1s spectrum of MnO₂. Generally, the O 1s pattern can be divided into three regions. The binding energy (BE) at 529.2–529.6 eV is ascribed to lattice oxygen (O_{latt}),²² the BE at 531.6–531.8 eV belongs to surface adsorbed oxygen (O_{ads}), which is naturally chemisorbed in the surrounding of the surface oxygen vacancies,^{36,37} and the BE at 533.2–533.6 eV is ascribed to the surface hydroxyl (OH) groups (OH_{surf}).²² Generally gaseous O₂ tends to be adsorbed on the surface oxygen vacancies and these adsorbed O₂ could be further dissociated into atomic oxygen at a certain temperature.³⁸ As indicated in Table 3, γ-MnO₂ contained the most abundant surface adsorbed oxygen (O_{ads}) with a ratio of O_{ads} to total oxygen species of 49.5%, which suggests that, among all the samples, γ-MnO₂ has abundant oxygen vacancies on its surface.

Fig. 5b illustrates the Mn 2p_{3/2} spectrum of MnO₂. The peaks at 642.6 and 641.8 eV can be assigned to surface Mn⁴⁺ and Mn³⁺

species.³⁹ From Table 3, the quantitative calculations of the Mn 2p^{3/2} XPS were conducted and the Mn⁴⁺/Mn³⁺ molar ratio was determined. Obviously, the molar ratio of surface elements Mn⁴⁺/Mn³⁺ is distinct in these four kinds of MnO₂. The δ-MnO₂ sample exhibits the maximum Mn⁴⁺/Mn³⁺ molar ratio (3.8). Besides, Mn⁴⁺–O bonds are more stable than Mn³⁺–O bonds.⁴⁰ This suggests that γ-MnO₂ contains more unstable Mn–O bonds on the surface, which indicates that the atomic oxygen species adsorbed on its surface are more probably released to take part in catalytic combustion. In the same way, the presence of low valence manganese on the surface could facilitate the dissociation and activation of the surrounding atomic oxygen species.⁹

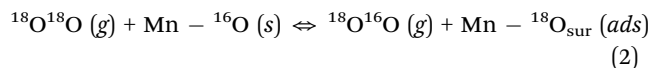
The average oxidation state (AOS) of manganese on the surface was assessed based on Mn 3s doublet splitting (ΔE_s) according to the following relationship:⁴¹

$$\text{AOS} = 8.956 - 1.126\Delta E_s \quad (1)$$

The AOS follows the order of δ-MnO₂ (3.82) > β-MnO₂ (3.80) > α-MnO₂ (3.78) > γ-MnO₂ (3.76), as presented in Fig. 5c. This rule accords well with the result of H₂-TPR, which indicates that the content of Mn³⁺ in γ-MnO₂ is the maximum.

3.4 Temperature programmed oxygen isotopic exchange (TPOIE) reaction

As mentioned above, the four samples followed the sample heterogeneous exchange mechanism, that is, exchange of one oxygen atom at a time. Here, the surface reactivity of the four samples at lower temperatures involving route (2) is accordingly investigated. The heterogeneous exchange reaction involving bulk lattice oxygens is used to evaluate their transfer ability at relatively high temperatures as formulated by route (3).



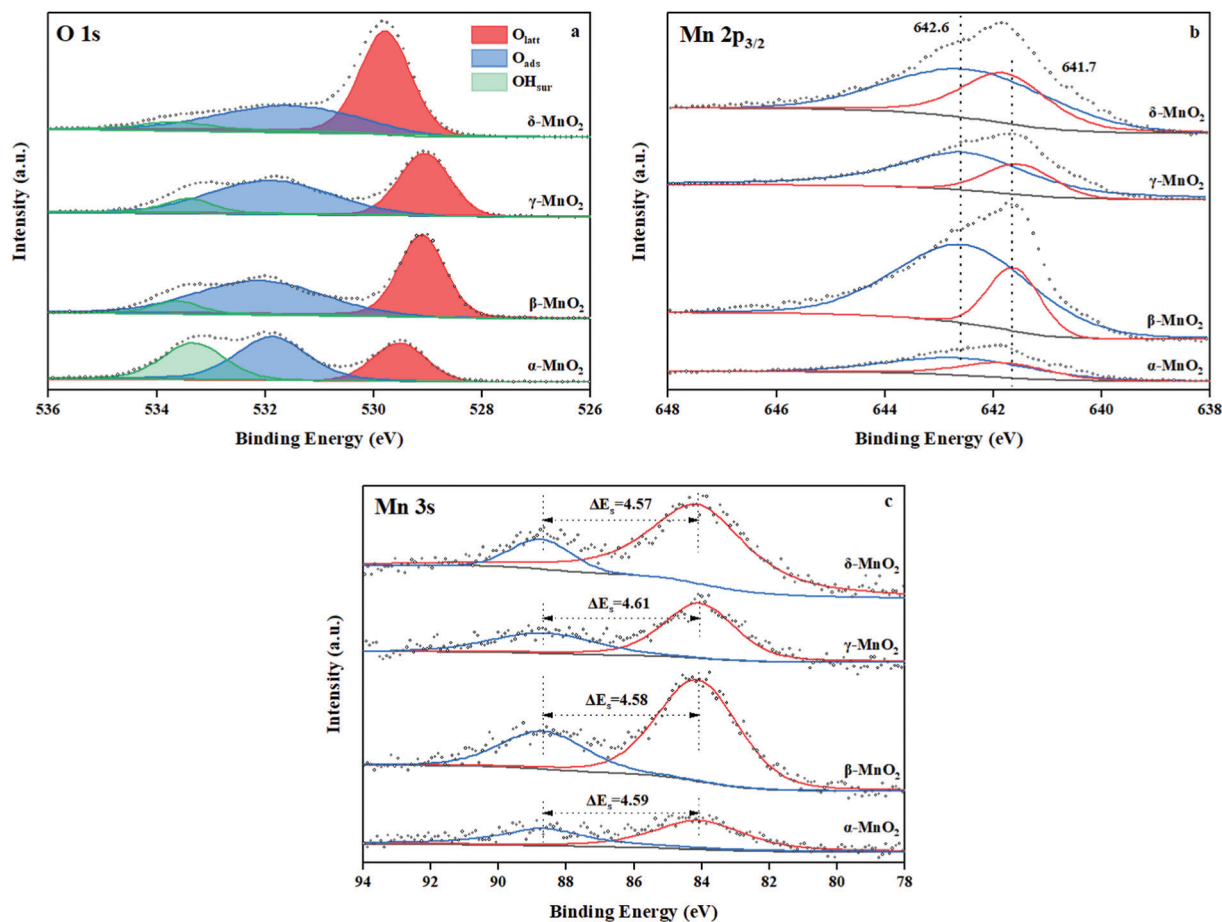
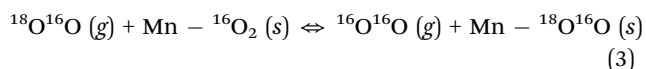


Fig. 5 The XPS of (a) O 1s, (b) Mn 2p_{3/2}, and (c) Mn 3s of MnO₂.

Table 3 The XPS results of MnO₂ and the average oxidation state (AOS) calculated by Mn 3s doublet splitting (ΔE_s)

Samples	Binding energy (eV) and ratio (%)			Mn ⁴⁺	Mn ³⁺	Average oxidation state (AOS)
	O _{latt}	O _{ads}	OH _{sur}			
α -MnO ₂	529.5/26.4	531.7/43.3	533.3/31.3	642.6/75.9	641.8/24.1	3.78
β -MnO ₂	529.3/47.0	531.8/47.3	533.5/5.7	642.6/77.9	641.7/22.1	3.80
γ -MnO ₂	529.2/40.6	531.7/49.5	533.4/9.9	642.6/75.7	641.6/24.3	3.76
δ -MnO ₂	529.6/57.8	531.6/36.4	533.5/5.8	642.5/79.0	641.8/21.0	3.82



where (s), (g) and (ads) refer to the lattice oxygen in the solid phase, gaseous oxygen in the gas phase and the adsorbed oxygen, respectively. As illustrated in Fig. 6, the ¹⁸O¹⁶O curve increases in the range of 200 to 400 °C, but it can hardly be observed in the ¹⁶O₂ formation, which is conducive to the predominant participation of oxygen atoms, corresponding to route (2). Further increase of the temperature leads to an enhancement of the ¹⁸O¹⁶O uptake, accompanied by ¹⁶O₂ production (400 to 600 °C). The ¹⁶O¹⁶O curve corresponds to route (3), which can only be observed over γ -MnO₂. The adsorbed molecular oxygen is mainly generated by route (2),

and the lattice oxygen is generated by route (3) at the same time.

From the experiment, the cumulative exchange oxygen atom amounts (N_e) and oxygen exchange rates (R_e) with temperature are obtained (Fig. 7a and b). As shown in Fig. 7a, the sum of exchanged oxygen atom amount is distinct in these four kinds of MnO₂ samples. The γ -MnO₂ sample contains the maximum amount of exchanged ¹⁸O atoms and the α -MnO₂ sample shows the opposite result. γ -MnO₂ has an excellent exchange ability throughout the whole temperature range (200–600 °C) since it has bigger specific surface area for releasing and making use of the surface and lattice oxygen atoms well. What is more, a huge difference among the four MnO₂ catalysts can be seen at the initial activated temperature. The temperature of exchange



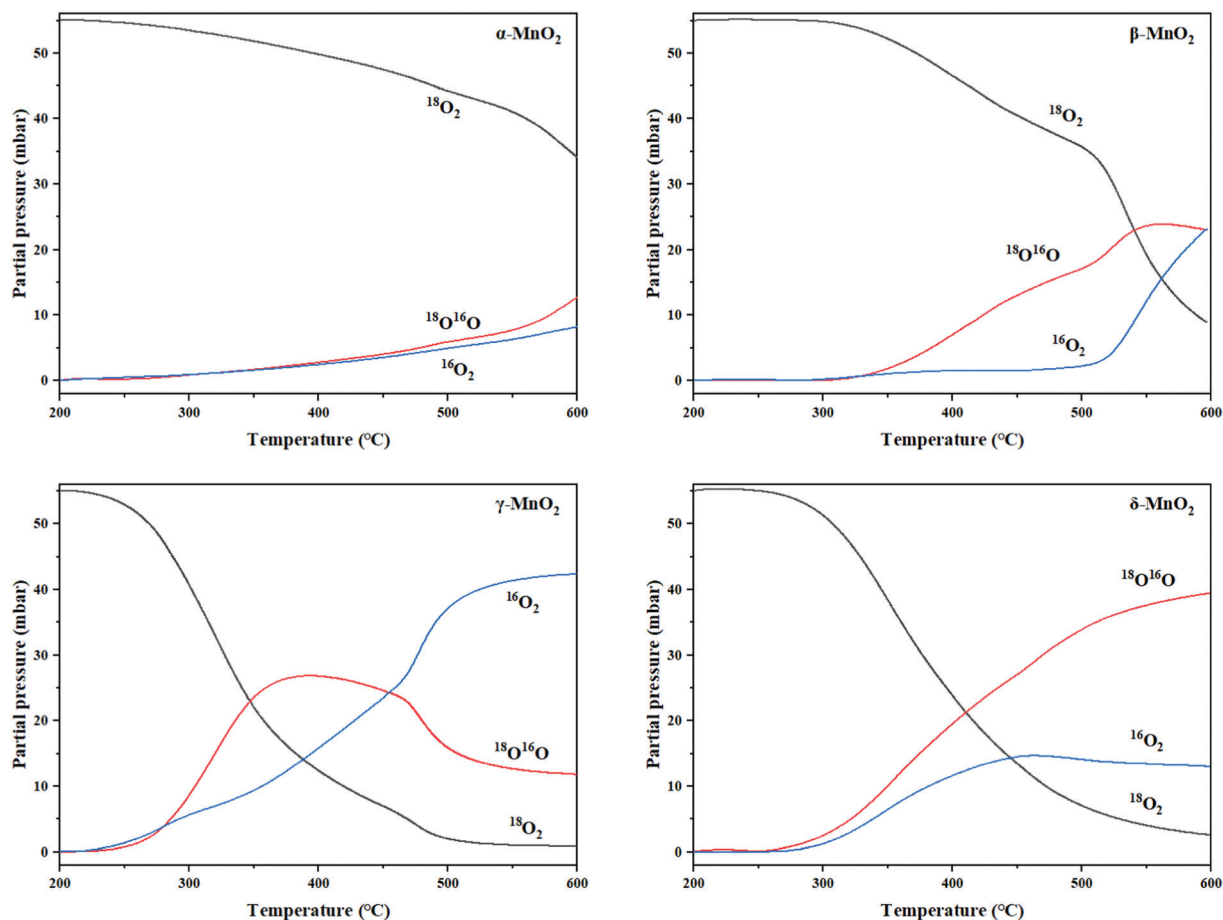


Fig. 6 Temperature programmed reaction of the $^{18}\text{O}_2/^{16}\text{O}_2$ isotopic exchange of MnO_2 .

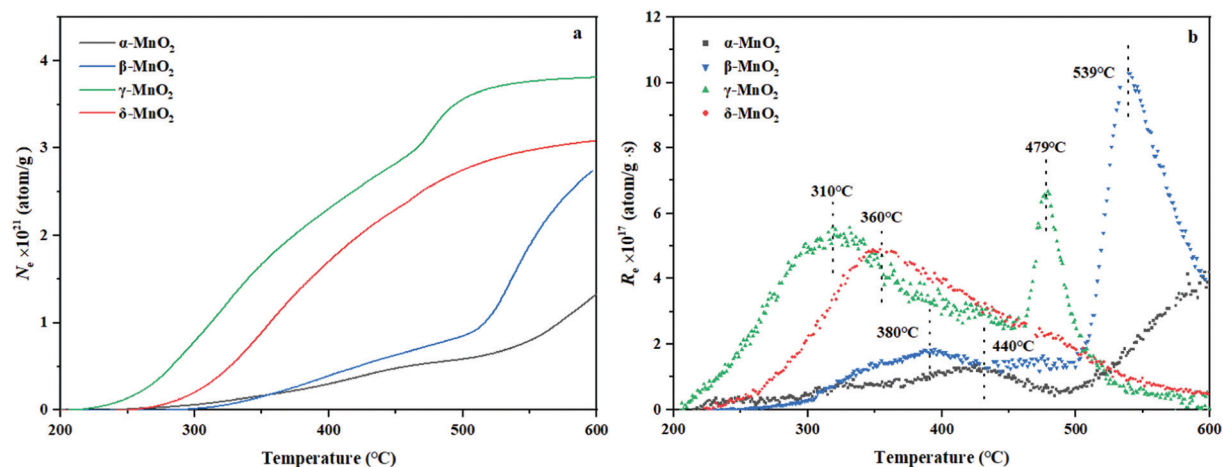


Fig. 7 (a) The curves of the total amount of exchanged ^{18}O (N_e) and (b) the exchange rate (R_e) with temperature of MnO_2 .

starts at 200 °C for $\gamma\text{-MnO}_2$, 220 °C for $\delta\text{-MnO}_2$, 240 °C for $\alpha\text{-MnO}_2$, and 280 °C for $\beta\text{-MnO}_2$ in an increasing order.

According to the R_e curves, activated temperature for the four MnO_2 catalysts (Fig. 7b) indicated the discrepancy in oxygen mobility ability over these catalysts further. In the

experimental temperature window, the maximum exchange rate ($R_{e\text{max}}$) value emerges when the gas–solid phase balance of ^{18}O amounts is achieved, and thus the behavior of each catalyst could be distinguished by the starting exchange temperature at which the value of $R_{e\text{max}}$ is constant, that is,



the lower the starting temperature is, the faster the migration of oxygen is. As displayed in Fig. 7b, the mobility behavior of reactive oxygen species is ranked as follows: γ -MnO₂ (310 °C) > δ -MnO₂ (360 °C) > β -MnO₂ (380 °C) > α -MnO₂ (440 °C). It is verified that γ -MnO₂ shows the most excellent activity toward the mobility of oxygen species which could be related to its abundant oxygen vacancies to yield the corresponding reactive oxygens. Actually, the formation of reactive oxygen species demands unequal activation energies in the process of isotopic ¹⁸O exchange. Interestingly, γ -MnO₂ and β -MnO₂ have the second extreme points, which likely presented exchange with some lattice oxygens in bulk. In other words, the exchange of lattice oxygen becomes the highest with an increase in temperature. In conclusion, the oxygen mobility of γ -MnO₂ is superior at both low and high temperatures, consistent with the O₂-TPD result.

3.5 NH₃-TPD

The surface acidity of the catalyst was thought to play a key role in the cleavage of the C–Cl bonding of CVOCs in our previous work.¹² Herein, the surface acidity of the catalyst was analyzed by NH₃-TPD and the acid amounts of these as-prepared MnO₂ catalysts were estimated according to the integral area of the desorption curve. The curve can be divided into three peaks for temperatures < 200 °C, 200–400 °C, and > 400 °C, corresponding to the weakly, moderately and strongly acidic sites for adsorbed ammonia, respectively.²⁷

For α -MnO₂, there are two peaks centered at 401 and 467 °C, both of which belong to desorption at strong acid sites. There is only one peak centered at 510 °C in the NH₃-TPD profile of β -MnO₂, corresponding to the strong acid sites. γ -MnO₂ shows four NH₃ desorption peaks at 122, 176, 401 and 550 °C; the former two peaks are assigned to desorption occurring at weak acid sites and the latter two are attributed to those occurring at strong acid sites. The NH₃-TPD curve of δ -MnO₂ contains two desorption peaks centered at 402 and 463 °C, both of which are due to desorption at strong acid sites. As displayed in Table 4, among the four samples, the existence of weak acid sites was found only in γ -MnO₂, which contained the highest amount of total acid, whereas the lowest acid amount was observed in the case of β -MnO₂. CVOCs were reported to be absorbed on acid sites or oxygen vacancies for dechlorination.²⁷ Thus, γ -MnO₂ containing relatively abundant acid sites is likely beneficial for C–Cl cleavage and HCl formation to achieve CVOC degradation by selective catalytic combustion.

3.6 Catalyst performance

The 1,2-DCE conversion with reaction temperature over the four kinds of MnO₂ catalysts is depicted in Fig. 9a. In detail,

Table 4 The results of NH₃-TPD

Sample	Weak acid (μmol g ⁻¹)	Moderate acid (μmol g ⁻¹)	Strong acid (μmol g ⁻¹)
α -MnO ₂	—	—	67.4
β -MnO ₂	—	—	4.4
γ -MnO ₂	7.8	—	65.0
δ -MnO ₂	—	—	31.1

these catalysts follow the T₅₀ order of γ -MnO₂ (236 °C) < α -MnO₂ (331 °C) < β -MnO₂ (352 °C) < δ -MnO₂ (397 °C). Compared to β -, γ -, and δ -MnO₂, γ -MnO₂ exhibits the most superior catalytic activity in the whole temperature window. The CO₂ yield as the desired oxidation products during the 1,2-DCE combustion is shown in Fig. 9b. This shows that γ -MnO₂ exhibited the most outstanding CO₂ selectivity in the combustion of 1,2-DCE. What is more, it can be observed that the CO₂ yield of the catalysts in the whole temperature window follows the order of γ -MnO₂ > α -MnO₂ > β -MnO₂ > δ -MnO₂, which indicates that β - and δ -MnO₂ would generate a larger number of chlorinated byproducts than α - and γ -MnO₂ during the 1,2-DCE combustion. Besides, the yield of HCl is also shown in Fig. 9c, which is another key product that should be considered. The order of HCl yield is as follows: γ -MnO₂ > α -MnO₂ > β -MnO₂ > δ -MnO₂.

As displayed in Fig. S2 (ESI[†]), the chlorinated organic byproducts, including C₂H₃Cl, C₂H₂Cl₂, 1,1,2-C₂H₃Cl₃ and CCl₄, are hardly observed on α - and γ -MnO₂, but these chlorinated byproducts are readily formed over β - and δ -MnO₂, especially on δ -MnO₂, resulting in the HCl yield presenting an inflection point at 400 °C, which indicates the occurrence of not only hydrochlorination but also polychlorination during the overall combustion process. Commonly, it is noticeable that all these chlorine-containing byproducts initially increased and then decreased with increasing temperature, which might vanish if the temperature was high enough. The further fatal HCl oxidation into Cl₂ would readily induce the formation of polychlorinated byproducts such as 1,1,2-C₂H₃Cl₃ and CCl₄, which are enhanced with increasing temperature.⁴² With further increase of temperature, these chlorinated byproducts will be gradually transformed into final target products. Although α -MnO₂ is not found to generate the chlorinated byproducts, the small amount of surface reactive oxygen species resulted in a CO₂ yield lower than that achieved over γ -MnO₂ (Fig. 5 and 7). Owing to the small amount of acidity, β - and δ -MnO₂ cannot easily break the C–Cl bond to produce HCl for dechlorination which also makes further transformation of the chlorinated byproducts difficult (Fig. 8).

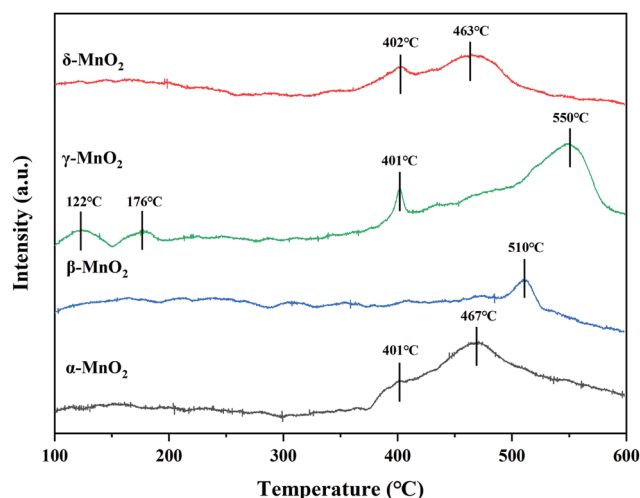


Fig. 8 The NH₃-TPD profile of MnO₂.



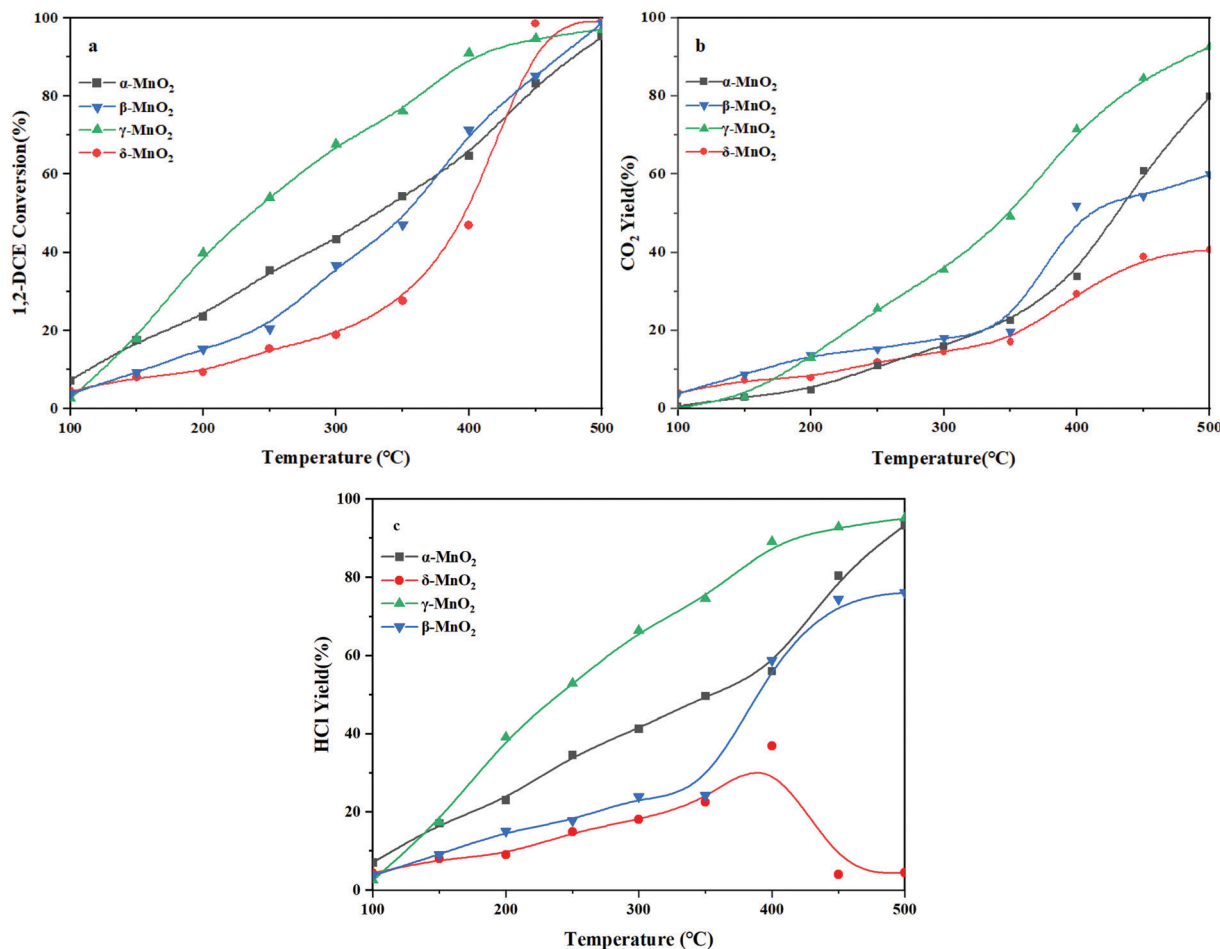


Fig. 9 (a) The conversion of 1,2-DCE, (b) CO₂ yield and (c) HCl yield over MnO₂ samples.

In spite of the reasonable redox properties of δ-MnO₂, it is unable to obtain a satisfactory HCl yield due to its enriched Mn⁴⁺, which results in strong Mn–Cl bonding to induce the catalyst inactivation and promote the formation of a large number of polychlorinated compounds. As mentioned above, if the catalyst does not possess both acidic sites and redox properties, 1,2-DCE will not dissolve or break bonds thoroughly, which suppresses the HCl formation at low temperatures. That is to say, with further increase of temperature, the formation of HCl will be very hard unless the temperature is relatively high. From this, we can see that γ-MnO₂ is the best catalyst for 1,2-DCE removal, which provides the highest yield of the final target product of HCl with a synergistic effect of acidity and redox properties. Its outstanding CO₂ selectivity can be probably attributed to the superior reducibility and abundant surface reactive oxygens, as demonstrated by H₂-TPR and OIE (Fig. 4 and 7), which achieves superior combustion of 1,2-DCE at relatively low temperatures. Moreover, the average oxidation state of the γ-MnO₂ sample is the lowest among the four samples, which indicates that it possesses more surface oxygen vacancies and relatively few Mn⁴⁺ (Fig. 4 and 5). Combined with the NH₃-TPD analysis mentioned above (Fig. 8), 1,2-DCE is firstly adsorbed and dissociated on the oxygen vacancies or the acid sites,

preventing the formation of a rather stable Mn–Cl bond and undesirable Cl₂.

3.7 *In situ* DRIFTS and reaction mechanism

In situ DRIFTS of 1,2-DCE oxidation over the four kinds of MnO₂ at different temperatures was conducted as shown in Fig. 10. Besides, the bands at 831 and 933 cm⁻¹ over the β-, γ- and δ-MnO₂ catalysts correspond to C–Cl bond vibration.^{43,44} The C–Cl bond vibration (Fig. 10b) appears on β-MnO₂ when the temperature is higher than 200 °C, which is related to the generation of 1,1,2-C₂H₃Cl₃ and CCl₄ byproducts. The vibrational bands of this bonding for the γ-MnO₂ sample are only visible at temperatures between 50 and 100 °C (Fig. 10c), which is supposed to be the remaining 1,2-DCE in the reaction. The C–Cl bond vibration on δ-MnO₂ appears at 100–200 °C, and the stretching vibration peak of C=C on the ethylene substituent is also found at 1621 cm⁻¹ (Fig. 10c),⁴³ so it is inferred that C₂H₃Cl and *cis*-C₂H₂Cl₂ are formed. Interestingly, the C–Cl bond (Fig. 10a) was marginally observed on α-MnO₂, which is consistent with the activity data. The bands in the 1300–1500 cm⁻¹ region over β-, γ- and δ-MnO₂ are assigned to the deformed vibration peaks of –CH₂– and –CH– groups,⁴⁴ which are related to 1,2-DCE and the produced chlorinated



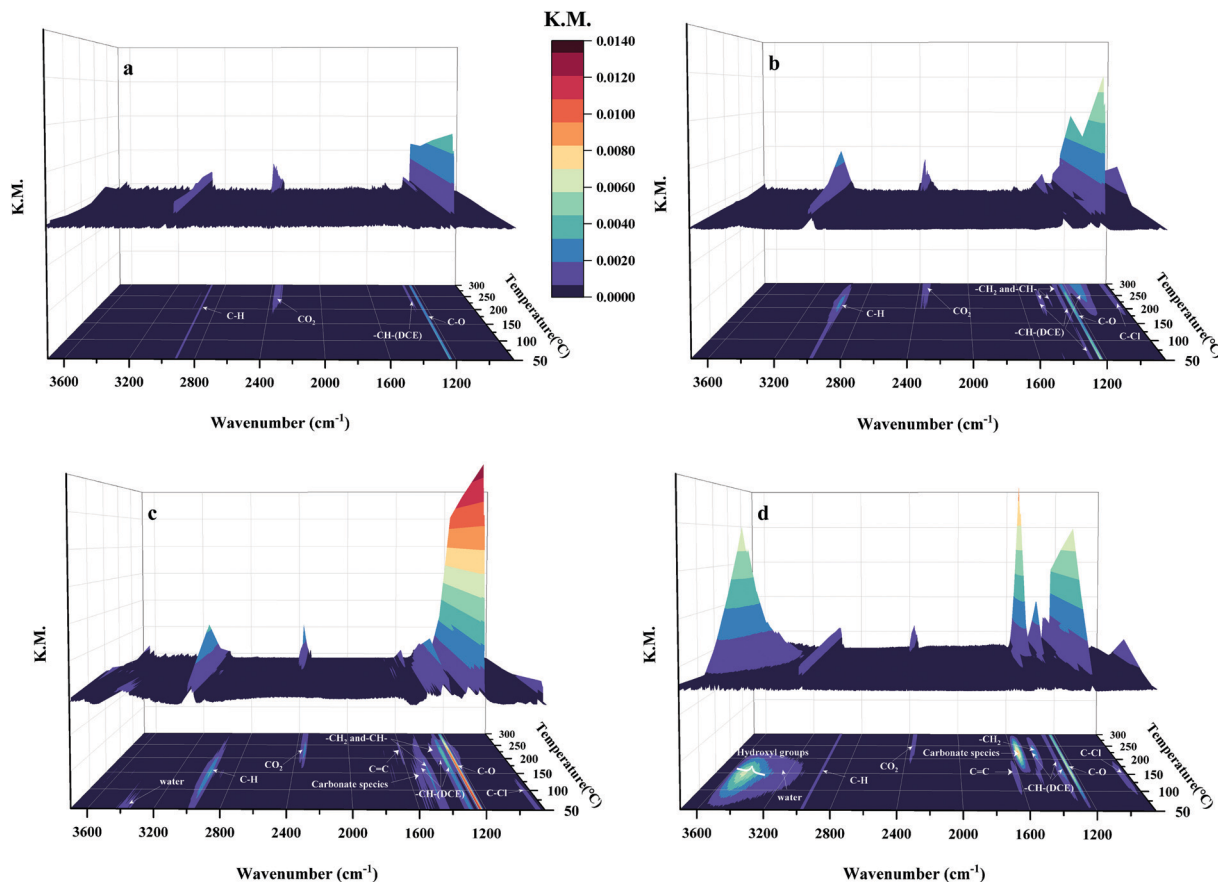
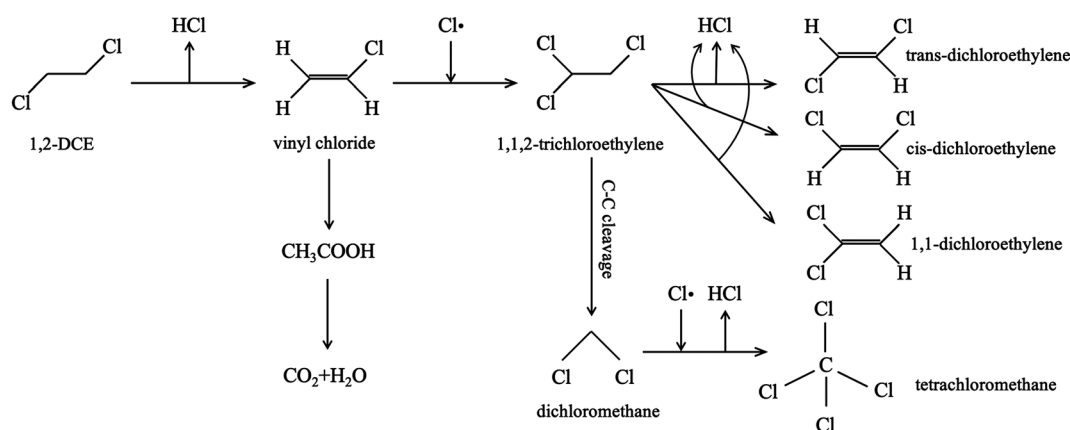


Fig. 10 The *in situ* DRIFTS results for oxidation of 1,2-DCE on (a) α -, (b) β -, (c) γ - and (d) δ -MnO₂ at different temperatures.

hydrocarbons. It is reported that the bands at 1226 and 1237 cm⁻¹ are ascribed to the C-O bond stretching vibration of acidic species,⁴⁵ which indicates that the acetic acid species were produced by the oxidation of 1,2-DCE and then deeply oxidized to H₂O and CO₂. The bands at 1284, 1297 and 1464 cm⁻¹ correspond to the vibration in the -CH-group in 1,2-DCE.⁴⁶ The bands in the range of 3200–3600 cm⁻¹ are associated with the OH group and water stretching vibration, which are observed over γ - and δ -MnO₂.⁴⁷ As reported, the

bands at 1555 and 1577 cm⁻¹ over γ - and δ -MnO₂ are assignable to the adsorbed carbonate species,¹² which are the sources of generation of CO₂ and CO products. Combining the results of previous characterization and Fig. 10, the probable combustion mechanisms of 1,2-DCE over the four kinds of MnO₂ are summarized in Scheme 1. As reported, the chlorinated compounds caught by the oxygen vacancies are more unstable than those bonded with Mn⁴⁺ ions⁴⁸ and tend to react to form HCl, which facilitates Cl desorption. Hence, it is



Scheme 1 The possible mechanisms of 1,2-DCE catalyst combustion over MnO₂.



deduced that 1,2-DCE is firstly activated at surface oxygen vacancies and successively destructed by acid sites *via* reaction with sufficient H atoms to generate HCl which is finally desorbed. C_2H_3Cl (by HCl elimination reaction), $C_2H_2Cl_2$ (by H_2 elimination reaction), 1,1,2- $C_2H_3Cl_3$ (by HCl addition or Cl substitution reaction) and CCl_4 (Cl substitution reaction) byproducts are produced, which are thereafter oxidized to acetic acid species. CO_2 and CO products are generated by

the deep oxidation of acetic acid species and carbonate species. Based on the above analysis, we can find that γ - MnO_2 with outstanding C-Cl/C-H cleavage ability displays a better catalytic performance during the 1,2-DCE combustion process, which is attributed to the abundant oxygen vacancies, excellent surface oxygen migration ability, more acid sites and less Mn^{4+} content on the surface. Furthermore, there are relatively abundant acid sites and few Mn^{4+} in α - MnO_2 , and thus we almost cannot

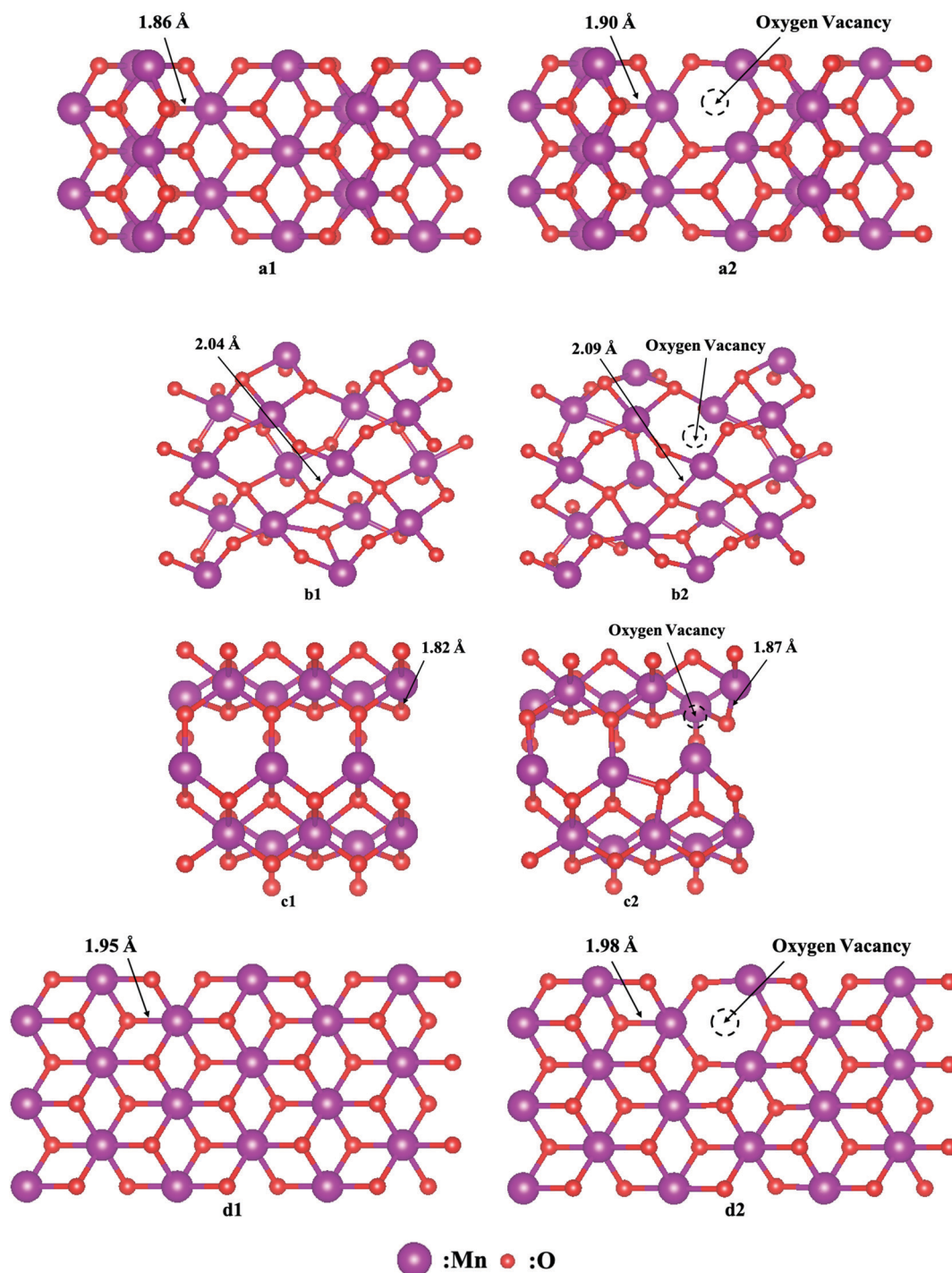


Fig. 11 Model structures of (a) α -, (b) β -, (c) γ - and (d) δ - MnO_2 (a1–d1) without and (a2–d2) with oxygen vacancies.



observe any chlorinated byproducts, but the yield of CO₂ is essentially low because of the poor exchange capacity of the surface oxygen of α -MnO₂. The AOS of the Mn element on the surface of β -MnO₂ and δ -MnO₂ is relatively high, and there are many Mn⁴⁺ species. 1,2-DCE mainly dissociates on Mn⁴⁺ to form a stable Mn–Cl bond, leading to serious catalyst deactivation.

3.8 DFT calculations

As discussed in the preceding characterization, the surface reactive oxygen species of MnO₂ play an important role in 1,2-DCE catalytic combustion reactions. It is believed that the formation of oxygen vacancies is due to an easy loss of the lattice oxygen, and thereafter a large number of surface reactive oxygens can be generated upon these ionic vacancies through a redox cycle, which is conducive to the fast migration of surface oxygen species.³⁶ Therefore, the easy formation of oxygen vacancies theoretically supports the high amount of surface reactive oxygens. γ -MnO₂ exhibited the highest specific surface area of 87.4 m² g⁻¹ associated with the best catalytic performance, while δ -MnO₂ exhibited the second largest specific surface area of 86.5 m² g⁻¹; however, it showed the worst catalytic performance. Thus, the specific surface area should not be the only factor affecting the catalytic performances of the four samples although a higher surface area usually leads to a better activity. In addition, the morphology, especially for the exposed facet, was considered and correlated to catalytic performances in the present work. The diverse morphologies observed by high-resolution transmission electron microscopy (Fig. 3) indicate the distinct crystal planes formed over the as-prepared samples. As illustrated in Fig. 11, a theoretical calculation on the exposed facets was conducted. Herein, the formation energy of oxygen vacancies by removing a lattice oxygen atom from the four kinds of MnO₂ catalysts was calculated based on eqn (4):

$$E_{\text{Ov}} = E_{\text{sys-O}} - E_{\text{sys}} + E_{\text{O}_2}/2 \quad (4)$$

Here, $E_{\text{sys-O}}$ represents the energy of the system with oxygen vacancies, E_{sys} represents the energy of the system, and E_{O_2} represents the energy of the O₂ molecules.

It is noticed that the lengths of the Mn–O bonds around the oxygen vacancies were all extended because of the Jahn–Teller effect,⁴⁹ causing more activated lattice oxygen to take part in the reaction readily. The E_{Ov} of α -, β -, γ - and δ -MnO₂ is 1.79 eV, 3.23 eV, 1.47 eV and 2.59 eV, respectively. It is clear that the energy for oxygen vacancy formation follows the sequence: β -MnO₂ > δ -MnO₂ > α -MnO₂ > γ -MnO₂, which suggests that the oxygen vacancy is likely to be generated over the γ -MnO₂ sample, resulting in adsorption of more reactive oxygen species, which makes it easier to be adsorbed as molecular oxygen and further dissociated on these surface vacancies, in accordance with the results of the present OIE experiment. To sum up the above discussion, we think that morphology is a more crucial factor than specific surface area that can significantly impact the redox ability and thereby improve the corresponding activity.

4 Conclusions

In summary, the preparation of α -, β -, γ -, and δ -type MnO₂ was successful for 1,2-DCE catalyst combustion. γ -MnO₂ displayed the most excellent catalytic activity to achieve an almost complete oxidation of 1,2-DCE with a synergistic effect of excellent redox properties and strong acidity during the 1,2-DCE oxidation process ($T_{90} = 406$ °C, $y_{\text{CO}_2} = 92\%$ and $y_{\text{HCl}} = 95\%$) at 42 857 mL g⁻¹ h⁻¹ GHSV. It is confirmed that oxygen vacancies tend to form on the surface of γ -MnO₂ based on DFT calculations. What is more, combined with O₂-TPD, NH₃-TPD and TPOIE technique analyses, γ -MnO₂ was confirmed to exhibit the most superior surface oxygen migration owing to the enriched surface reactive oxygens on the ionic vacancies. Based on the H₂-TPR and XPS analyses, having less Mn⁴⁺ on the surface significantly reduced the formation of strong Mn–Cl bonding along with the adsorption of 1,2-DCE, which makes it easy to facilitate the subsequent reaction yielding polychlorinated byproducts. Moreover, the corresponding intermediates could be transformed into CO₂ and H₂O over catalysts with good redox properties. Besides, it is proved that the HCl formation and C–Cl bond cleavage ability are related to both the relatively abundant acid sites and the surface mobile oxygens, which accelerates the formation of HCl and finally enables conversion into CO₂ and H₂O. Both the dehydrochlorination by HCl elimination and the polychlorination by Cl substitution contribute to the formation of chlorinated byproducts including C₂H₃Cl, C₂H₂Cl₂, 1,1,2-C₂H₃Cl₃ and CCl₄.

Author contributions

Baicheng Shi: resources, writing – original draft, data curation, formal analysis, writing – review & editing. Runduo Zhang: supervision, funding acquisition, validation. Zhaoying Di: investigation, validation. Xiaonan Guo: investigation, visualization. Ying Wei: project administration, visualization. Jingbo Jia: project administration, visualization.

Conflicts of interest

The authors declare that they have no competing financial interests.

Acknowledgements

The financial support from the National Natural Science Foundation of China (No. U1862102, 22176010 and 21976012) and the Fundamental Research Funds for the Chinese Central Universities (No. XK1802-1 and JD2117) is gratefully acknowledged.

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