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Microporous MOF as nanogen facilitating diffusion-coupled charge transfer near the percolation threshold in a polyaniline pseudo-supercapacitor

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Several approaches have recently been investigated with the aim of enhancing the specific capacitance of polyaniline (PANI). We herein report a novel pathway to boost the specific capacitance of PANI, which is highly tunable and versatile, based on wet chemistry techniques. In this approach, utilization of a specific metal–organic framework (MOF), UiO-66-NH$_2$, as a nanopore generator (*Nanogen*) is demonstrated. This systematic study demonstrates a tight interplay between the enhanced electrolyte accessibility to the PANI chains, through the incorporation of MOF nanogens, and the optimal relative amount of the conductive PANI, necessary to attain a charge percolation threshold in the binary system (MOF@PANI). Satisfying the charge percolation while maintaining porosity is necessary to arrive at a maximized specific capacitance of the PANI as the active phase. The enhancement in the specific capacitance of the PANI was fully exploited through a systematic investigation that helped to pinpoint the saddle point at which the two orthogonal properties, namely porosity (electrolyte diffusivity) and charge mobility (PANI inter-chain electronic conductivity) can be fine-tuned *via* controlling the materials composition. Of the different compositions investigated, the composite containing 23 wt% PANI of the total weight, doped with MOF nanogens, resulted in an enhanced specific capacitance of 872 F g^{-1} for the PANI, in comparison to only 469 F g^{-1} for the pristine PANI investigated under identical conditions.

Introduction

The distinctive high surface area, presence of chemically functionalized nanocages, and uniform pore systems are all common attributes of the large family of microporous solids. Among these, zeolites, metal–organic frameworks (MOFs),[1] and porous organic polymers (POPs),[2] as well as covalent-organic frameworks (COFs)[3] have attracted considerable interest due to their proven uses in several applications. Implementation of microporous solids in water treatment,[4,5] catalysis,[6–8] gas sorption and separation,[9,10] and battery technology,[11] among others, are largely facilitated due to their intrinsic properties; namely the dimensionality and functionality of their cages and/or pores.

Conductive polymers have emerged as a novel class of materials that can bridge the gap between metals, with their superior electronic conductivity, and organic matter, as well as

their ease of synthesis and functionalization.[12] PANI is arguably one of the most explored conductive polymers.[13–15] Due to its enhanced electronic conductivity and redox behavior, it was recently investigated for its potential as a pseudosupercapacitor material.[16–19] Recently, the construction of microporous materials on conductive supports, including graphene and multi-walled carbon nanotubes (MWCNTs), has provided an efficient pathway to merge the microporosity and electrical conductivity of such composites.[20–23]

In recent studies, MOF–PANI hybrid materials were recently explored with demonstrated performance as electrode materials for several electrochemical applications. Recent examples have included the utilization of MOF@Se@PANI composites in batteries,[24] MOF@PANI for the hydrogen evolution reaction (HER) in electrocatalysis,[25] Ni–MOF/PANI-derived materials for electrochemical sensors,[26] as well as Ni–MOF–PANI nanosheets used in supercapacitors.[27] The aforementioned applications were accessed through the synthesis of various MOF–PANI composites, each tailored to a specific application by judicial choice of the different components in the composite.

In other previous reports, it was demonstrated that the presence of several types of MOFs,[28] including UiO-66[29] and

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ZIF-67,[30,31] resulted in enhanced electrochemical performance of MOF–PANI hybrids. However, we opted here to explore the potential of the amine-functionalized MOF UiO-66-NH$_2$ due to its excellent chemical stability and presence of ionizable amine groups that can facilitate ion mobility within the material. A recent study also demonstrated the efficient confinement of PANI chains within the pores of UiO-66-NH$_2$, with a positive impact on the stability of the PANI chains observed upon extended cycling.[32] Attempts were also successfully made to link the observed enhancement to the intrinsic microporosity of the MOFs and the electrical conductivity of the PANI.[33] However, a deeper understanding of the underlying mechanism for this observed synergy is still lacking.
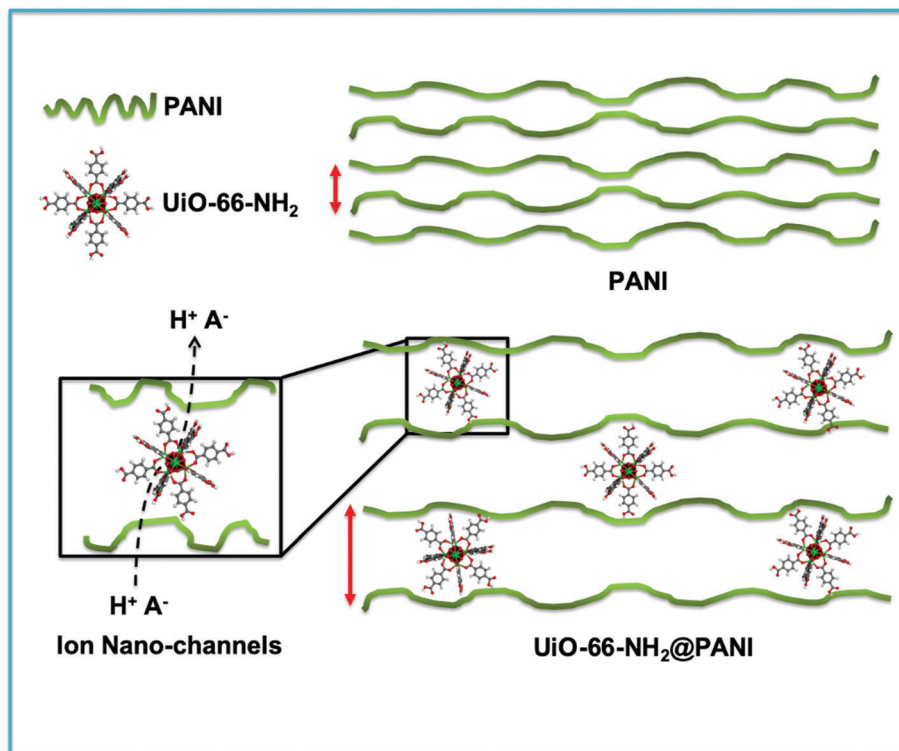
In this report, a systematic investigation was conducted, by varying the UiO-66-NH$_2$ to PANI composition in a well-controlled *in situ* PANI polymerization setup, in order to decipher the relative significance that the intrinsic properties of the two components of the composite have on the overall observed performance of the materials as an electrode material for supercapacitor applications. PANI utilization as an electrode material in supercapacitors stems from its high electrical conductivity ($\sim 10^1$ S cm^{-1})[34] and pronounced electrochemical activity due to its ability to undergo doping/dedoping states. Although PANI has a theoretical specific capacitance of 964 F g^{-1}, calculated assuming a potential window of 1.1 V,[35] without taking into consideration capacitance from the electrical double layer, this is greatly affected by the surface area of the PANI. However, under practical conditions, such maximal specific capacitance is difficult to attain. Specifically, care must be taken when conducting electrochemical measurements on PANI, by cycling within the potential range that can result in reversible oxidation from leucoemeraldine (L) to emeraldine (E), avoiding irreversible damage to the PANI when measurements were extended to the fully oxidized state, pernigraniline (P).[35] A large number of reports have demonstrated enhancement in the specific capacitance of PANI through different methodologies, where in the vast majority the target was to enhance the surface-to-volume ratio of PANI molecular wires.[36,37] This is of interest when designing a supercapacitor material in order to enhance the surface charge density and therefore increase the specific capacitance of the same material. Earlier attempts in that direction targeted the formation of PANI nanowires,[38] their deposition on flat supports,[39] increasing their surface roughness and decreasing the particle size,[40] *etc.*
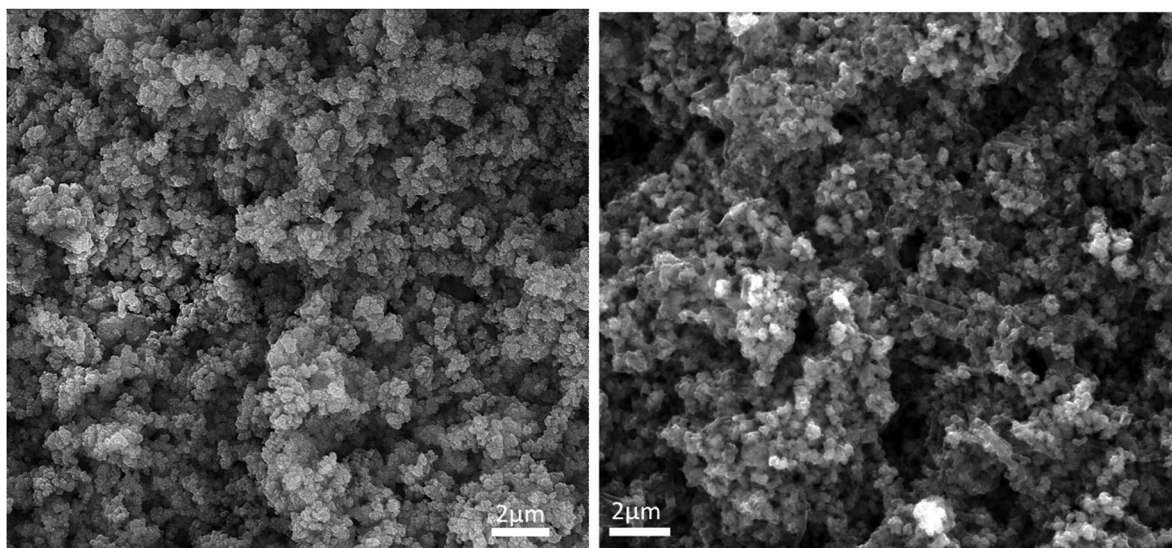
Results and discussion

In the approach presented herein, we opted to utilize a specific MOF to be incorporated as a filler within the matrix of PANI. More specifically, the enhanced microporosity, chemical stability, and high surface area of the Zr-based MOF (UiO-66-NH$_2$)[41] were key to its selection for this study. In this design, the incorporation of homogeneously dispersed nanocrystallites of the MOF within the matrix of PANI was targeted to provide

ion-conducting nano-channels imbedded in-between the PANI polymer chains (Scheme 1). By providing such action to facilitate ion migration from the electrolyte to reach deeper into the PANI chains and their subsequent electrosorption onto the polymer chains, the MOF can effectively be regarded as a filler for generating nanocavities, which we refer to here as *nanogens*. This is in analogy to the well-established concept of utilizing porogens in the fabrication of porous polymer membranes, where mostly nanoparticles of soluble salts are physically blended within a viscous solution of the polymer, which can subsequently then be leached out from the casted membrane to leave behind a uniform array of pores.[42–44] In this design, we envisioned that two orthogonal properties, namely the microporosity of the MOF filler with its fairly low electrical conductivity, and the low porosity of the PANI with its high electrical conductivity, need to be optimized in order to realize the full potential of the nanogens. Therefore, to capture the saddle point at which the favorable impact of the MOF on the overall MOF–PANI composite is maximized, a systematic investigation was conducted to probe the optimal MOF composition that can result in enhanced ion mobility within the matrix of the composite, yet without drastically decreasing the electrical conductivity of the overall composite. As the former is a viable target to enhance diffusivity of electrolyte species to reach the PANI–electrolyte interface, the latter is the essential element to maintain, in order to retain good electronic properties for the supercapacitor electrode. The synthesis we report here is highly controllable, allowing for the isolation of different compositions with increasing concentration of the nanogen filler. This strategy is especially suited for the systematic study of the global variance in the MOF–PANI composition needed to result in the optimized performance of the PANI as a pseudosupercapacitor. The PANI wt% in the end product can readily be tuned by controlling the ratio between the MOF filler and the aniline–ammonium persulfate mixture at the beginning of the reaction. Suspending the MOF nanocrystals into the PANI precursor solution while maintaining mechanical stirring resulted in the nucleated growth and deposition of the PANI chains onto the MOF crystals, resulting in a textured morphology of the obtained composite, as shown in the scanning electron microscopy (SEM) images in Fig. 1. Five different MOF@PANI compositions were synthesized, namely those containing 14.5 wt%, 23 wt%, 30 wt%, 47 wt%, and 79 wt% of PANI within the composite. To further characterize the samples prepared, Fourier-transform infrared (FTIR) spectra of PANI, UiO-66-NH$_2$, and the different MOF@PANI compositions were recorded and are shown in Fig. 2. The major characteristic peaks of the UiO-66-NH$_2$ include the Zr–O stretching at 765 cm^{-1}, the carboxylate C–O symmetric and asymmetric stretching (1380 cm^{-1} and 1574 cm^{-1}, respectively), as well as the C=C stretching observed at 1432 cm^{-1}.[45] The FTIR spectrum of PANI exhibits characteristic peaks at 1560, 1472 and 1297 cm^{-1}. The peaks at 1560 and 1478 cm^{-1} can be assigned to the C=C stretching of quinoid and benzenoid rings, respectively, while the band observed at 1283 cm^{-1} can be assigned to C–N stretching vibrations.[46] Interestingly, the composites with

Scheme 1 Schematic representation of the effect of inclusion of the UiO-66-NH$_2$ MOF to facilitate ion diffusion through the nanopores of the MOF acting as a nanogen between the PANI chains.

Fig. 1 SEM images of the UiO-66-NH$_2$ MOF (left) and the MOF@PANI (right), showing the textured morphology of the MOF–PANI composite.

lower PANI wt% demonstrate mostly the characteristic peaks of the MOF, with a decrease in the intensity of the MOF peaks observed upon an increase in the mass% of the PANI in the composite, until the MOF peaks become barely noticeable in the spectrum of the composite containing 79 wt% PANI. Furthermore, Fig. 3 shows the powder X-ray diffraction (PXRD) patterns of PANI, UiO-66-NH$_2$, and the different MOF@PANI

composites. The PANI diffraction pattern features a broad peak at $2\theta = 25.4°$ signifying its highly doped emeraldine salt form.[47] All of the MOF@PANI composites show the characteristic peaks of UiO-66-NH$_2$, whereas as the PANI content increases, the characteristic peaks for the MOF become less visible. The changes in the pore size distribution histograms, Fig. 3b, demonstrate that the pore systems centered at ∼1.6 and
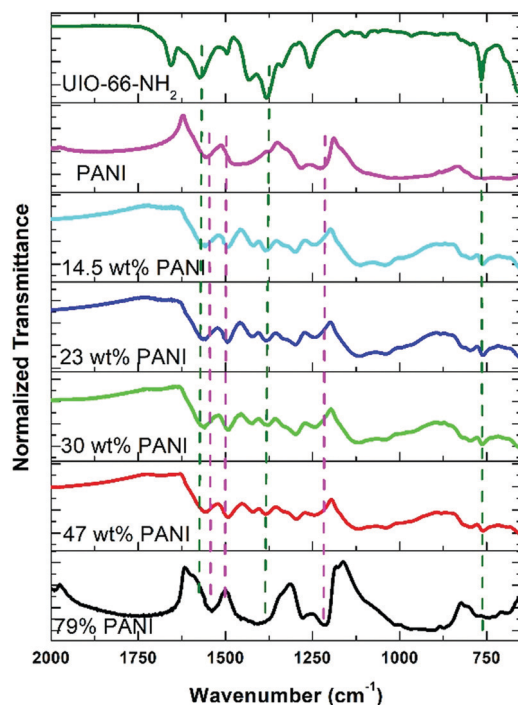
Fig. 2 FTIR spectra of the UiO-66-NH$_2$, PANI, and the composites with different PANI compositions. The dotted lines show the characteristic peaks of UiO-66-NH$_2$ (green) and PANI (violet).
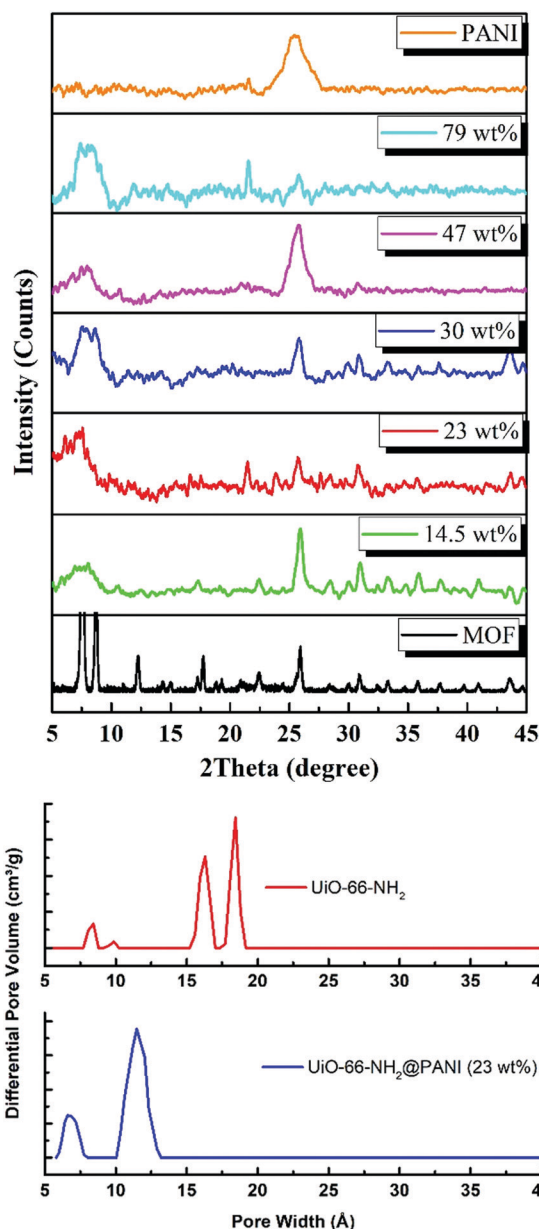

Fig. 3 (Top) XRD patterns of MOF, PANI, and their reported composites (the wt% refers to that of the PANI), and (below) the pore size distribution histogram of the MOF and PANI@MOF 23 wt% composite.

1.8 nm in the UiO-66-NH$_2$ MOF are lost in the PANI@MOF composite, indicating a buildup of the PANI molecules within the cages of the MOF.

Overall, five different samples of varying PANI–MOF compositions were tested for their electrochemical response *via* cyclic voltammetry (CV) measurements, in addition to those of the pristine MOF and PANI, using a three electrode system in 1 M H$_2$SO$_4$ aqueous electrolyte, Fig. 4, to determine the specific capacitance of the PANI in each sample. The CV scans in Fig. 4 clearly demonstrate the initial enhancement in the PANI pseudo electrochemical capacitance upon introducing the MOF into the composite, where the composite with 79 wt% PANI demonstrated slightly enhanced specific capacitance (537 F g^{-1}) compared to that of the pristine PANI (469 F g^{-1}). Increasing the MOF content resulted in a further enhancement in the PANI specific capacitance, where the sample containing 47 wt% PANI showed a specific capacitance of 718 F g^{-1}, increasing in the sample containing 30 wt% PANI (782 F g^{-1}), and further increasing in the sample containing 23 wt% (872 F g^{-1}), before dropping in the sample containing 14.5 wt% PANI (643 F g^{-1}). Moreover, noticeable changes in the peak shape, comparing the sample of pristine PANI to those containing the MOF fillers, were noticeable. The CV scans of the samples containing the nanogens demonstrated a more pronounced diffusion-controlled redox process of the PANI, evident in the well-resolved double hump shape of the CV scans. This observation points towards facilitated ion diffusion from the electrolyte to reach deeper into the PANI chains in the presence of the nanogens. Plots of the specific capacitance and

surface area of the obtained solids *versus* the PANI wt% are shown in Fig. 5. It is clear from this data that the increase in surface area as a result of the increased MOF wt% is accompanied by an increase in the specific capacitance of the PANI, reaching an inflection point beyond the sample containing 23 wt% PANI. Interestingly, this behavior is in agreement with a previous investigation on the charge percolation threshold of a graphene–MOF composite system, where a sudden increase in the overall electrical conductivity of the composite was attained at ~30 wt% conductive filler composition.[23] This charge percolation threshold, identified previously for a different MOF–graphene system seems to hold true in this case as
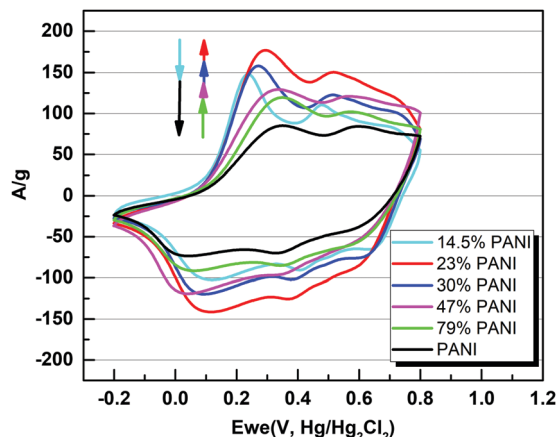
Fig. 4 Cyclic voltammetry (CV) plots of PANI and its composites deposited on glassy carbon electrodes at a scan rate of 10 mV s^{-1} in 1 M H$_2$SO$_4$.
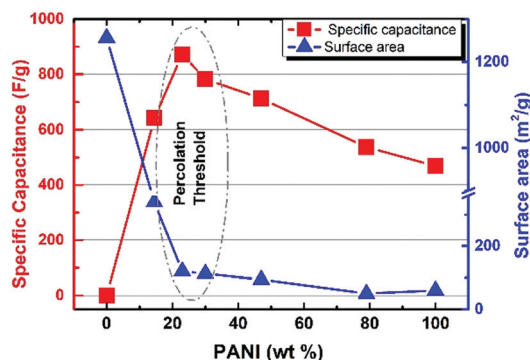

Fig. 5 Composite surface area and specific capacitance of PANI dependence on the wt% of the PANI in the MOF@PANI, showing the highest attained specific capacitance at an optimal composition of 23 wt% PANI.

well, taking into consideration the different nature of the conductive element in the composite.

To gain deeper insight into the operating mechanism(s) resulting in the enhanced specific capacitance of the PANI in the prepared composites, electrochemical impedance spectroscopy (EIS) was utilized, with the results shown in Fig. 6. The EIS technique was employed in order to dissect the complex nature of the time-dependent processes using a complex frequency domain study. The Nyquist/Bode EIS data recorded within the frequency range of 50 kHz to 10 mHz at the open circuit potential with an AC perturbing signal of 10 mV were nicely fitted to an equivalent circuit with two time constants, using a modified Randle's circuit, as shown in Fig. 6. At relatively high frequency, the impedance of the system was found to be mostly dominated by the internal and interfacial resistances of the electrode along with the ohmic resistance of the electrolyte, represented by the resistor component (R_{Sol}) in the fitting parameters. In the medium to high frequency range, a distorted semicircle shape of the real *versus* imaginary components (Z' and Z'') of the impedance was noted, and was assigned to the charge-transfer resistance across the interface within the composite (R_{CT1}) and the constant phase element (CPE$_{DL}$) representing the double layer capacitance. Most significantly, the inter-chain charge transfer, represented by the R_{CT1} element, was found to drop significantly upon increasing the conducting PANI wt% (Table 1), matching the charge percolation threshold described earlier. This charge-transfer process, most likely representing charge hopping within the composite, is thought to dominate the overall charge transfer/storage mechanism within the composite and is in good agreement with the findings of the electrochemical CV section, *vide supra*. The two-sloped linear tail profile observed at relatively low frequency can be assigned to the short circuit terminus Warburg element, describing a finite length ion diffusion process across the electrode interface, accompanied with a second time constant. Such element is fairly utilized to describe the interlinked nature of the charge storage through an ion diffusion process within the voids of a porous film with finite thickness. The variations in the B parameter, which describes the diffusion time with relative concentration of the PANI within the composite matrix, reveals that the fastest diffusion process was recorded in the PANI concentration range of 30–47 wt%.
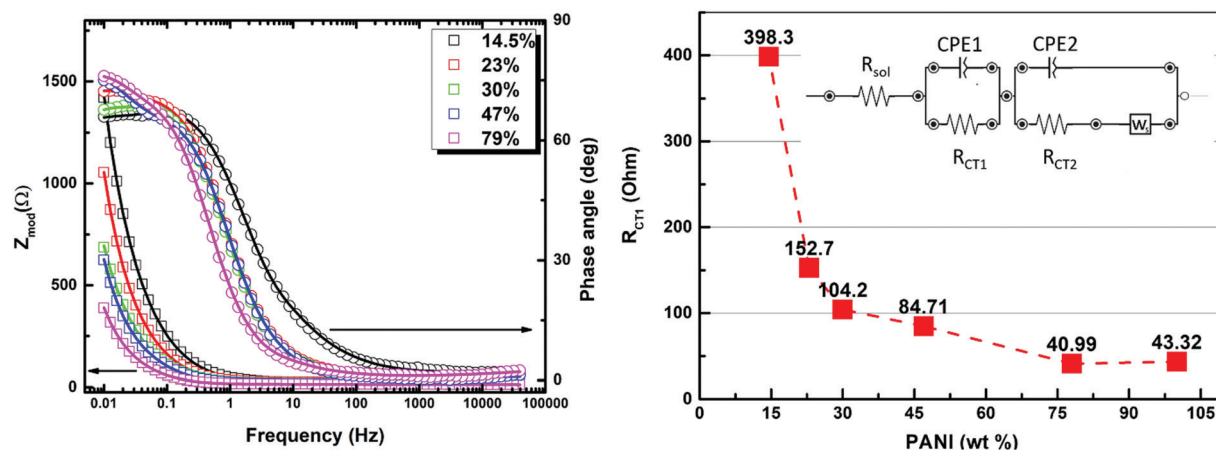

Fig. 6 EIS data and fitting of the samples containing different PANI wt% (left), and changes in the R_{CT2} element upon increased PANI wt% (right), where the inset shows the equivalent circuit used for fitting.

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Table 1 The fitting parameters of the different elements of the equivalent circuit used in modeling the EIS data

Composition (PANI wt%)	14.5%	23%	30%	47%	79%	100%	Unit
R_{soln}	10.63	17.81	16.46	15.24	9.666	10.69	Ohm
R_{CT1}	398.3	152.7	104.2	84.71	40.99	43.32	Ohm
C_{CPE1}	0.01047	0.02329	0.0353	0.02995	0.07901	0.07433	S*s^a
n_{CPE1}	0.763	0.809	0.7858	0.8276	0.8498	0.7251	
C_{CPE2}	0.0006259	0.00055	0.001813	0.00124	0.001104	0.001934	S*s^a
n_{CPE2}	0.3714	0.3491	0.4446	0.5312	0.3702	0.4801	
R_{CT2}	6. 105	5.496	1.554	1.106	3.701	2.002	Ohm
$Y_{W|}$	0.0278	0.06563	0.1063	0.3411	0.1033	0.1047	S*s^(1/2)
$B_{W|}$	0.3842	0.2168	0.1764	0.06827	0.3775	0.429	Sec^(1/2)

The second charge transfer resistance (R_{CT2}), was assigned to represent the charge transfer within the PANI chains. Considering the aforementioned fitted parameters, it is clear that the samples containing 23–30 wt% PANI seem to exhibit the optimum combination of low PANI–MOF interface resistance, low diffusion time, and moderate intra-chain charge transfer resistance within the PANI chains, making them the best candidates for maximizing the specific capacitance of the PANI. Indeed, this result is in close agreement with the demonstrated highest specific capacitance recorded for the 23 wt% composite in the CV measurements. Considering all such factors and the sudden drop in the porosity of the composite beyond the 23 wt% composition, it becomes evident that a delicate balance between establishing charge percolation pathways within the composite while simultaneously retaining good porosity is key to achieving targeted improvement in the specific capacitance of PANI.

Overall, it is clearly demonstrated that: (i) the presence of nanogens results in enhanced electrochemical performance most likely linked to facilitated ion diffusion impacting the shape of the CV scans, (ii) increasing the nanogen concentration results in an increase in the specific capacitance of PANI, and (iii) beyond the optimal composition, a drop in the specific capacitance is observed. Coupled with the results from EIS study, it is reasonable then to assume that increasing the nanogen concentration enhances ion diffusivity within the PANI matrix, thus resulting in a more accessible surface of the PANI chains for the electrolyte. However, increasing the concentration of the microporous, electrically non-conductive MOF beyond a certain saddle point results in an inflection in the specific capacitance. This drop in the specific capacitance upon an increase in the MOF content beyond an optimal value can be ascribed to tipping the balance between generating more accessible voids and decreasing the overall electrical conductivity of the composite. The decrease in the electrical conductivity of the composite when the MOF wt% increases is in agreement with the poor electrical conductivity of the MOF. This composition-electrical conductivity relationship was recently reported for the MOF doping of PANI, where electrical conductivity was found to drastically drop upon increasing the MOF wt% in the composite.[29,53] It is worth noting that such previous attempts to prepare MOF–PANI composites were only concerned with the electrical conductivity of the composite or did not address the systematic variations in the MOF/PANI

Table 2 Selected values of the PANI specific capacitance reported recently using wet chemistry polymerization techniques

Sample	Discharge rate	Potential range (V)	Specific capacitance ($F\ g^{-1}$)	Ref.
PANI@UiO-66-NH$_2$	7 mV s^{-1}	$-0.1:0.8$	872	This work
PANI	10 mV s^{-1}	$-0.1:0.8$	503	40
PANI	10 mV s^{-1}	$-0.2:1.0$	283	48
PANI	5 mV s^{-1}	$0:0.85$	187	49
PANI–MnO$_2$			330	
PANI nanofibers	0.18 A g^{-1}	$0:0.8$	548	50
G/PANI nanoFiber	100 mV s^{-1}	$-0.2:0.8$	480	51
De-doped PANI nanofibers	2.5 A g^{-1}	$0–0.75$	593	52

ratios and their impact on the specific capacitance of the PANI. Unwinding this tight entanglement between the porosity and electrical conductivity of the MOF–PANI composite was made possible by the systematic approach devised herein to result in the designable and controllable composition of the binary system PANI@MOF for successful improvement of the PANI specific capacitance. Noticeably, the highest specific capacitance value of 872 F g^{-1} for the composite with 23 wt% PANI is higher than other PANI forms prepared via wet chemistry methods, as shown in Table 2, and therefore presents a facile and tunable approach to enhance the pseudosupercapacitor characteristics of PANI, and is potentially a transferrable model to be exploited for other types of MOFs towards enhancing such properties of PANI and also other conductive polymers.

To further investigate the electrochemical performance of the composite, a two-electrode setup was assembled and analyzed using galvanostatic charge–discharge (GCD) measurements. The sample containing 23 wt% PANI was utilized in this supercapacitor characterization setup, where the material was drop cast on Ti electrodes separated by a porous membrane and 1 M H$_2$SO$_4$ was used as an electrolyte. The charge–discharge curves of the prepared supercapacitor were probed by varying the current density in the range of 0.5–10 A g^{-1}, as shown in Fig. 7a, which demonstrates the nearly symmetric triangular shape of the GCD profile, indicating that the facilitated redox process is associated with the presence of nanogens in the matrix, facilitating ion diffusion from the bulk of the electrolyte to reach the solid–electrolyte interface. The specific capacitance of the electrode was calculated based on GCD data, where a normal dependence of the specific capacitance on the
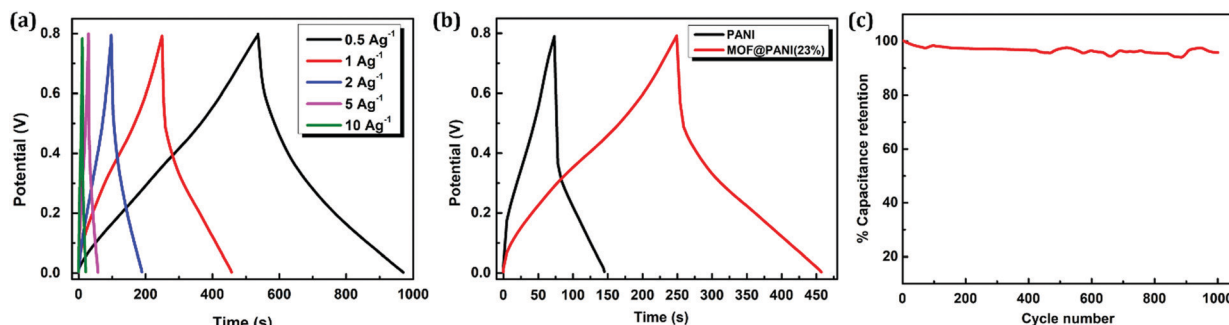
Fig. 7 Two-electrode setup characterization of the MOF@PANI (23 wt%) showing (a) the GCD profile at different charging current density for the MOF@PANI_23 wt%, (b) GCD response comparison with PANI at a current density of 1 A g^{-1}, and (c) capacitance retention over 1000 cycles.

current density was observed with an obtained value of 345 F g^{-1} at a current density of 0.5 A g^{-1}. Furthermore, comparison between the charge–discharge performance at the same current density between pristine PANI and the sample containing 23 wt% of the PANI clearly shows a longer time span for the GCD curve of the MOF@PANI composite, indicating a higher charge storage capacity for the MOF@PANI composite, as shown in Fig. 7b. Additionally, excellent retention of the specific capacitance over 1000 cycles at a current density of 1 A g^{-1} was demonstrated, as shown in Fig. 7c, where the MOF@PANI retained over 96% of its initial capacitance, even after 1000 cycles, indicating the robustness of the system, a performance that surpasses that of pristine PANI, which commonly shows quick decay in capacity over a relatively small number of cycles.[54–57]

Conclusion

In conclusion, we report a facile and potentially transferrable model to enhance the pseudocapacitor behavior of PANI through the incorporation of MOF nanocrystals, as a microporous filler (nanogens), into the PANI matrix, providing ion-diffusion channels within the matrix to enhance the electrolyte diffusion to reach the PANI chains, as compared to bulk non-porous PANI prepared *via* a wet chemistry synthesis. Through a systematic investigation to probe the effect(s) of varying the MOF content within the MOF@PANI composites, an optimal composition was attained at a weight ratio of 23 : 67 (PANI : MOF), which demonstrated the highest specific capacitance for PANI among the investigated composites. It is argued that this optimal composition provides for enhanced diffusivity of the electrolyte species towards the PANI–electrolyte interface through the microporous MOF filler (nanogens), while maintaining good electrical conductivity of the overall composite (provided by the PANI chains). This approach paves the way for a large family of MOFs that demonstrate a wide variety of pore dimensions and chemistry, as well as stability in different electrolytes, to be directly implemented as nanogens to boost the electrochemical characteristics of another family of readily available conductive polymers, addressing the limitations of both types of drastically different materials to prepare

well-designed composite materials that exhibit far-enhanced properties compared with their parents materials.

Experimental

All reagents were of commercial grade and used without further purification. Electrochemical measurements were carried out in a three-electrode system using a calomel electrode as the reference electrode (Hanna Instruments, HI 5412, 3.5 M KCl), glassy-carbon electrode as the working electrode, and Pt wire as the counter electrode. Measurements were recorded on a Bio-Logic SP-50 potentiostat/galvanostat and data were recorded using the EC-lab software. FTIR spectra were recorded on a ThermoScientific iS10 spectrophotometer. Activated carbon was purchased from Carlo Erba, aniline (99.5%, extra pure) was purchased from Acros Organics, ammonium persulfate (APS, acs reagent, >98%) was purchased from Sigma-Aldrich, ZrCl$_4$ (anhydrous, >98%) was purchased from Acros, 2-aminoterephthalic acid (99%) was purchased from Acros organics, HCl (37%) was purchased from Honeywell, N-methyl-2-pyrrolidone (NMP) was purchased from Sisco Research Laboratories, N,N'-dimethylformamide (DMF, analytical reagent grade, 99.99%) and acetonitrile (ACN, HPLC gradient grade) was purchased from Fisher Chemical. Aniline was purified by distillation under reduced pressure prior to use.

Synthesis of PANI

In a clean glass vial, 15 mL of APS (100 mM in 1 M HCl) was added into 10 mL of aniline (100 mM in 1 M HCl) while vigorously stirring and was left to stir at room temperature for 20 h. The formed precipitate was filtered and washed with 50 mL of deionized water twice and left to dry under vacuum at room temperature to give the product in a yield of 40 mg.

Synthesis of the UiO-66-NH$_2$ (MOF)

A previously published procedure for the synthesis of UiO-66-NH$_2$ using HCl as a modulator was utilized.[58]

Synthesis of MOF@PANI_X

For the UiO-66-NH$_2$@PANI composites, where X represents the PANI wt% in the composite, a similar procedure to the

Table 3 Amounts of precursor solutions and MOF filler used for the preparation of the MOF@PANI composites and their respective yields

Sample	Aniline (mL)	APS (mL)	UiO-66-NH$_2$ (mg)	Yield (mg)
MOF@PANI_14.5	10	10	122	142.8
MOF@PANI_23	10	10	82	107.1
MOF@PANI_30	10	10	62	88.7
MOF@PANI_47	10	10	40	75.6
MOF@PANI_79	10	10	10	47.1

synthesis of PANI described above was used, but varying the weight of UiO-66-NH$_2$ dry powder dispersed briefly in the aniline solution using ultrasonication for 15 minutes and the APS was added prior to leaving the mixture to stir at 550 rpm for 20 h at room temperature (Table 3).

General procedure for ink preparation

For each sample, $\sim$8.1 mg of the active material was well ground in an agate mortar, to which $\sim$2.6 mg of activated carbon was added and the resultant mixture was ground for 10 minutes. The solid mixture was then added to a solution of $\sim$0.8 mg of PVDF dissolved in 1 mL of NMP. The solution mixture was then sonicated for 15 minutes, then 10 μL was cast onto a polished glassy carbon electrode and left at 60 °C until dry. Another 10 μL was cast and dried before electrochemical testing. The reference electrode was a calomel electrode (3.5 M KCl), and a platinum wire was used as the counter electrode.

Specific capacitance calculations

CV measurements were performed from -0.2 to 0.8 V at scan rates ranging from 10 to 250 mV s^{-1}. The specific capacitance (C) (F g^{-1}) of the PANI in the different composites was calculated from the CV at a scan rate 10 mV s^{-1} and derived from the following equation:

$$C = (I/V)*A/(SC*m)$$

where I (mA cm^{-2} V^{-1}) is the integrated peak area, V is the potential window, A is the electrode surface area, SC is the scan rate, and m (mg) is the mass of PANI in the sample under investigation.

Two-electrode setup

Galvanostatic charge–discharge cycles were recorded on a Bio-Logic SP-50 potentiostat/galvanostat and the reported specific charge was calculated using the EC-lab software utilizing the mass of one electrode of the assembly. To prepare the ink, 8 mg of the MOF@PANI (23%) was ground with 2.6 mg of activated carbon and the solids were suspended in 1 mL of PVDF (0.79 mg) in NMP solution *via* ultrasonication. 200 μL of the ink was then cast on each of the Ti electrodes and left to dry at 60 °C for 2 h. 1 M H$_2$SO$_4$ was used as the electrolyte.

Electrochemical impedance spectroscopy (EIS) measurements

EIS measurements were conducted on similarly prepared ink as that used in the CV measurements using a PalmSens4 potentiostat galvanostat in a three-electrode system in 1 M HCl aqueous electrolyte, using GCE as the working electrode, a calomel reference electrode (3.5 M KCl), and a platinum wire as the counter electrode. All measurements were conducted at open circuit potential covering a frequency range from 10 mHz to 50 kHz, at an AC perturbing signal of 10 mV.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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