CORRECTION

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Correction: An aqueous hybrid electrolyte for low-temperature zinc-based energy storage devices

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Correction for 'An aqueous hybrid electrolyte for low-temperature zinc-based energy storage devices' by Nana Chang *et al.*, *Energy Environ. Sci.*, 2020, **13**, 3527–3535, https://doi.org/10.1039/D0EE01538E.

In Fig. 2c on page 3530 of this article, the previous notation of "10^{-5}" should be corrected to "10^{-8}" in the *y*-axis label on the right, due to a mistake in the process of converting "$\text{Å}^2\ nS^{-1}$" (unit of theoretical calculation) into "$cm^2\ S^{-1}$" (unit of picture presentation). There is no description of the value of the diffusion coefficient, only the discussion about the trend.

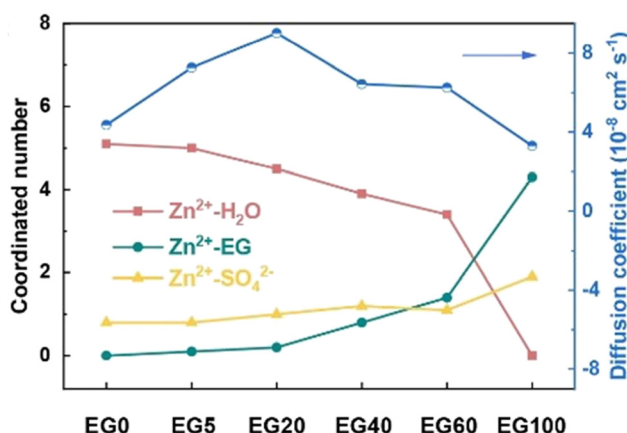Fig. 2c should appear as follows:

Fig. 2   (c) Coordination numbers of H_2O, EG, and SO_4^{2-} around Zn^{2+} and the diffusion coefficient of Zn^{2+} in different electrolytes calculated from MD simulations.

The discussion of the diffusion coefficient in Fig. 2c should be as follows:

"Moreover, the diffusion rate of Zn^{2+} in these electrolytes can be qualitatively evaluated by the mean-squared displacement (MSD) as a function of time (Fig. S13, ESI†). Surprisingly, the diffusion coefficient of Zn^{2+} rises and then declines as the EG content increases (Fig. 2c). When the EG-to-water ratio is lower than 60%, the diffusion coefficient of the hybrid electrolyte is still higher than that of EG0. This result shows that the solvation structure of Zn^{2+} coordinated with EG and H_2O contributes to the fast transport of Zn^{2+}. Nevertheless, the calculated fast ion diffusion coefficient of Zn^{2+} is not quite consistent with the slight decrease in the ionic conductivities of the hybrid electrolytes at 20 °C (Fig. 1b), which probably results from the gradually increasing viscosities of the hybrid electrolytes (Fig. S3, ESI†). Even so, fast Zn-ion transport in the hybrid electrolyte is achieved at low temperatures, when the solvation interaction of EG with Zn^{2+} becomes more prominent."

This correction does not have any impact on the core idea and conclusion of this work.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

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