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5D total scattering computed tomography reveals the full reaction mechanism of a bismuth vanadate lithium ion battery anode†

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We have used *operando* 5D synchrotron total scattering computed tomography (TSCT) to understand the cycling and possible long term deactivation mechanisms of the lithium-ion battery anode bismuth vanadate. This anode material functions *via* a combined conversion/alloying mechanism in which nanocrystals of lithium–bismuth alloy are protected by an amorphous matrix of lithium vanadate. This composite is formed *in situ* during the first lithiation of the anode. The *operando* TSCT data were analyzed and mapped using both pair distribution function and Rietveld methods. We can follow the lithium–bismuth alloying reaction at all stages, gaining real structural insight including variations in nanoparticle sizes, lattice parameters and bond lengths, even when the material is completely amorphous. We also observe for the first time structural changes related to the cycling of lithium ions in the lithium vanadate matrix, which displays no interactions beyond the first shell of V–O bonds. The first 3D *operando* mapping of the distribution of different materials in an amorphous anode reveals a decline in coverage caused by either agglomeration or partial dissolution of the active material, hinting at the mechanism of long term deactivation. The observations from the *operando* experiment are backed up by *post mortem* transmission electron microscope (TEM) studies and theoretical calculations to provide a complete picture of an exceptionally complex cycling mechanism across a range of length scales.

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Introduction

Lithium-ion batteries (LIBs) are the most successful portable and reusable energy storage devices since the petrol tank.¹ They are a vital part of the ongoing transition from a fossil fuel based economy to one based on renewable energy sources, but they remain far from perfect: improvements in capacity, charge/discharge speed, cycle lifetime and cost are all demanded by users and manufacturers. Finding the best materials to provide this combination of properties requires a detailed understanding

of the chemistry and physics of the components, especially the anode and cathode, which store and release energy as lithium ions shuttle between them. *Operando* methods-structural studies of batteries during charge and discharge – provide one of the best sources of information to help us explore this. They capture chemical reactions and structures as they occur in real time and avoid damage and structure degradation/relaxation which can occur while dismantling cells and preparing samples for *ex situ* experiments, even in a glovebox.²

Several of the most interesting new battery materials lack long range order in some states of charge and discharge, making the otherwise powerful structural tools of diffraction useless.^{3–5} In the last decade total scattering methods (often using the atomic pair distribution function, PDF) have become popular for studying disordered materials thanks to the increasing availability of high energy X-ray and neutron sources and rapid development of data treatment tools.^{6,7} We previously reported *operando* total scattering computed tomography (TSCT) as a method for obtaining high quality PDF data on specific battery components, avoiding the problem of removing signals from battery components in which we are not interested (*e.g.* cell casing, current collectors, separator, electrolyte). The

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experiment revealed the full cycling mechanism of a phosphorus anode in a sodium ion battery, despite the anode being weakly scattering and amorphous during most of the cycling process.⁸ This method, initially termed pair distribution function computed tomography (PDFCT) and developed for studying catalysts,⁹ is derived from X-ray diffraction computed tomography (XRDCT),¹⁰ which has been applied to batteries in several studies.¹¹ TSCT on the phosphorus anode gave us high quality area averaged PDFs for a difficult system, but not spatially resolved PDF mapping. Throughout this manuscript we will refer to the experiment in general as TSCT, the data extracted from Bragg diffraction as XRDCT and that obtained from the pair distribution function as PDFCT.

Bismuth vanadate (BiVO_4) is a promising LIB and sodium ion battery (NIB) anode material whose high and stable capacity (up to 1000 cycles) make up for a rather high voltage plateau.^{4,12} Its cycling mechanism combines conversion and alloying reactions. The long cycle life is believed to be due to the lithium bismuth alloy (Li_xBi) particles (which give high capacity) being protected by a matrix of lithium vanadate (Li_yVO_4) formed on decomposition of the initial bismuth vanadate. The $\text{Li}_x\text{Bi}/\text{Li}_y\text{VO}_4$ composite that appears in the first lithiation is amorphous, and its structure is still not fully understood despite XAS, diffraction and theoretical studies. After the initial cycles the Li_xBi formed from BiVO_4 becomes more crystalline and its cycling can be tracked with *operando* XRD, but the structure and behavior of the amorphous and weakly scattering Li_yVO_4 remain unclear and vanadium XAS data have proved practically impossible to collect in an operating battery due to strong absorption. The combination of amorphous structure, good X-ray scattering power and an intriguing, poorly understood cycling mechanism involving multiple phases make this system a tempting target for *operando* TSCT studies.

We can now present data collected at the upgraded ESRF beamline ID15A,¹³ that use the full power of TSCT to reveal comprehensive chemical and microscopic information on a working BiVO_4 lithium ion battery (LIB) anode in 5 dimensions: time, space and chemical structure information. XRD and PDF

based maps for individual chemical components observed during cycling confirm and add detail to our understanding of BiVO_4 anode cycling. Furthermore, they show changes in the microscopic distribution and crystallinity of the active material that hint at the longer-term deactivation mechanism. In addition, using the area summed TSCT data, we can follow the cycling behavior of the amorphous Li_yVO_4 , as well as tracking the mechanism and crystallite growth in the Li_xBi throughout the entire cycling process. This information can be obtained either by fitting the PDF and XRD data or directly from the radial extent of the PDF. Finally, we support our observations with TEM studies on a lithiated anode material and DFT calculations. The former show the distribution of Li_xBi and Li_yVO_4 in the composite created by the initial conversion of BiVO_4 , while the latter help to explain some of the fine details observed in the cycling mechanism using TSCT.

Results

Coin cell tests at a cycling rate of C/6.6 on the BiVO_4 /super P carbon/PAA binder composition used for the *operando* experiments showed that it performs as expected from our earlier studies.¹² The cells showed stable retained capacity of around 350 mA h g^{-1} between 250 and 500 cycles (Fig. S3, ESI†). The dQ/dV plot (Fig. S4, ESI†) derived from the electrochemical data collected in the *operando* experiment also matches with earlier reports.

1 TSCT mapping and the Li_xBi cycling mechanism

An initial inspection of the TSCT maps produced from XRD and PDF data for the pristine BiVO_4 at the start of cycling shows that they are slightly different (Fig. 1), revealing that the distributions of crystalline and amorphous material are slightly different (see also Fig. S8, ESI†). The maps show the poor coverage obtained by painting the active material onto the carbon piston, with significantly less material present as we move away from the piston surface from slice 0 to slice 2.

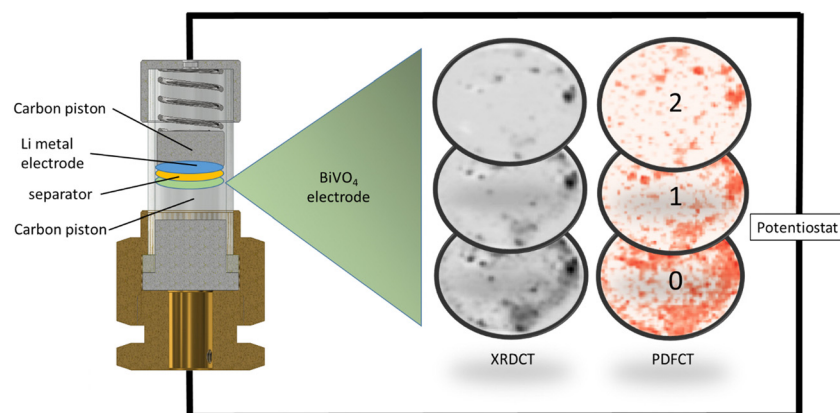


Fig. 1 A sketch of the electrochemical TSCT cell with examples of tomographic slices for the three different axial positions (0, 1, 2) in the BiVO_4 electrode measured at the start of the *operando* TSCT experiment reconstructed in XRDCT (total diffracted intensity) and PDFCT (BiVO_4 distribution) modes.



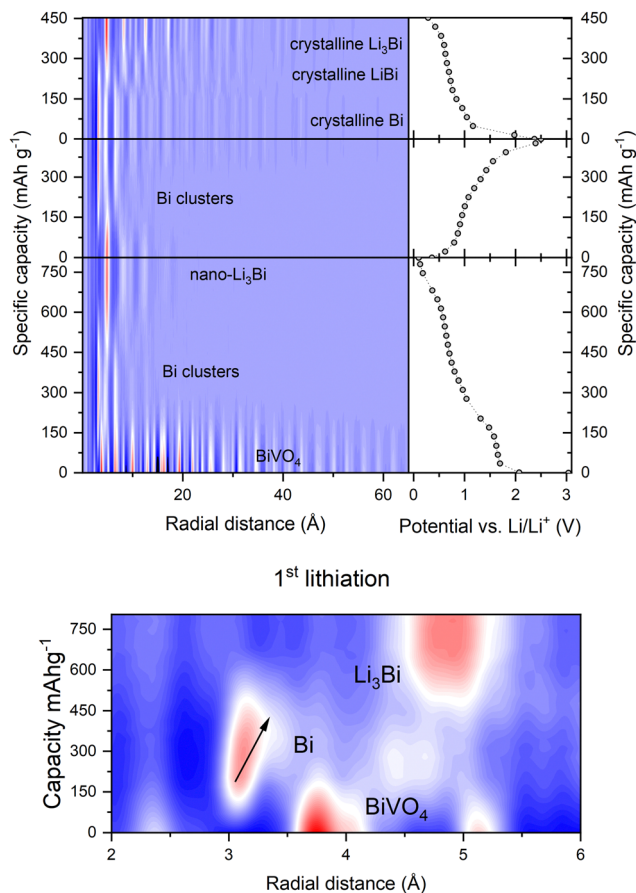


Fig. 2 Contour plot view of the $G(r)$ data averaged from all voxels in the tomograms containing active material in slice 0, with cycling data plotted alongside to indicate the progress of the electrochemistry. The radial distance to which peaks are observed in the PDF indicates the size of the particles. A zoomed view of the low r region for the first lithiation is shown below. Phases are labelled to indicate where they appear in the cycling and an arrow indicates the slight shift of the Bi–Bi cluster peak to longer radial distance in the first lithiation.

Examination of PDFs obtained by summing the active area in each slice gives us a quick picture of the initial BiVO_4 and Li/Bi phases that appear and disappear during cycling (ESI†, Fig. S9, ESI†), confirming the progression observed with XAS/XRD and theoretical calculations:¹²

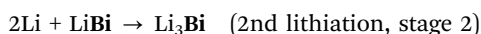
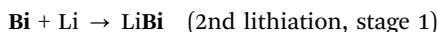
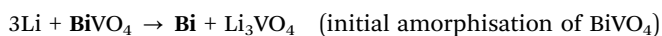


Fig. 2 shows the PDFs vs. specific capacity in a contour plot. The radial distance (r) to which peaks are observed in the PDF gives a simple indication of the varying particle size and

crystallinity of the phases formed during cycling. The maximum r observed corresponds to the longest atom–atom distance in the sample. After the initial crystalline BiVO_4 ($r > 60 \text{ \AA}$) breaks down, bismuth clusters with $r < 15 \text{ \AA}$ form. The Li_3Bi formed after the first lithiation has a maximum r of around 20 \AA and breaks down on delithiation into Bi clusters ($r < 15 \text{ \AA}$) before crystalline Bi ($r > 40 \text{ \AA}$) appears as the first delithiation is completed. The second lithiation goes *via* crystalline LiBi ($r > 30 \text{ \AA}$) to crystalline Li_3Bi ($r > 45 \text{ \AA}$). The appearance of the crystalline structures also leads to the appearance of Bragg peaks in the area averaged XRD data from the first delithiation onwards (ESI†, Fig. S9). By zooming in to the low r region for the first lithiation (Fig. 2, bottom) we can see a slight increase in the r -position of the peak for the direct Bi–Bi bond in the initial Bi metal clusters (shown by an arrow in the figure), indicating extension of this bond. The change could be explained by one or more of: lithiation around the Bi clusters, increasing amounts of surface Bi with a relaxed bonding environment as the clusters grow, or the beginning of interactions between the nano-sized clusters.

We were also able to fit the area-summed *operando* PDF and XRD data against structural models using TOPAS V6.¹⁴ This further clarifies the progress of the reactions involving BiVO_4 and Li_xBi (Fig. 3) and shows some more subtle structural variations. Details of the fits and structural models are given in the (ESI†, Fig. S5–S7, section: PDF and Rietveld fitting). From the disappearance of BiVO_4 to the end of the first lithiation the Bragg peaks in the XRD data are too broad to give a meaningful fit. The PDF data, by contrast, can be fitted throughout. After the crystalline BiVO_4 breaks down the first Bi phase to appear can be fitted with a single cluster model consisting of 7 Bi atoms with no long range interactions (achieved by placing the clusters in the centre of a much larger unit cell in the structure model). The fits do not reveal any further evidence of extension of the Bi–Bi bond observed in Fig. 2, probably because the model is not a completely accurate description of the various XRD amorphous Bi clusters present in the material. When Li_3Bi appears it is also in the form of a nano-sized clusters, this time fitted using the crystal structure (ICSD 58797)¹⁵ with a refined spherical damping parameter to describe cluster size. The fit used two Li_3Bi phases with different damping values, “bulk” ($r = \sim 10 \text{ \AA}$) and Li_3Bi “nano” ($r = \sim 5 \text{ \AA}$). The delithiation begins with Li_3Bi disappearing and being replaced by the Bi-metal in the nanocluster form only, but at around 350 mA h g^{-1} we see a new Bi metal phase appearing. This phase is fitted in the PDF with the crystal structure (ICSD 64703)¹⁶ with a refined radial damping parameter ($r = \sim 45 \text{ \AA}$). The crystal structure can also be used to fit the sharp Bragg peaks which appear in the XRD data at this stage (note that the Rietveld fit treats the nanosized phases as background). A crystalline LiBi phase ($r = \sim 20 \text{ \AA}$) appears in the second lithiation and can be fitted in the same way, both to the PDF and XRD data (ICSD 58796).¹⁵ This phase coexists with Bi metal, suggesting that the mechanism is split between direct lithiation of nano Bi to Li_3Bi and lithiation of bulk Bi *via* LiBi to Li_3Bi (the thermodynamic equilibrium pathway, see supporting figure [DFT], ESI†).

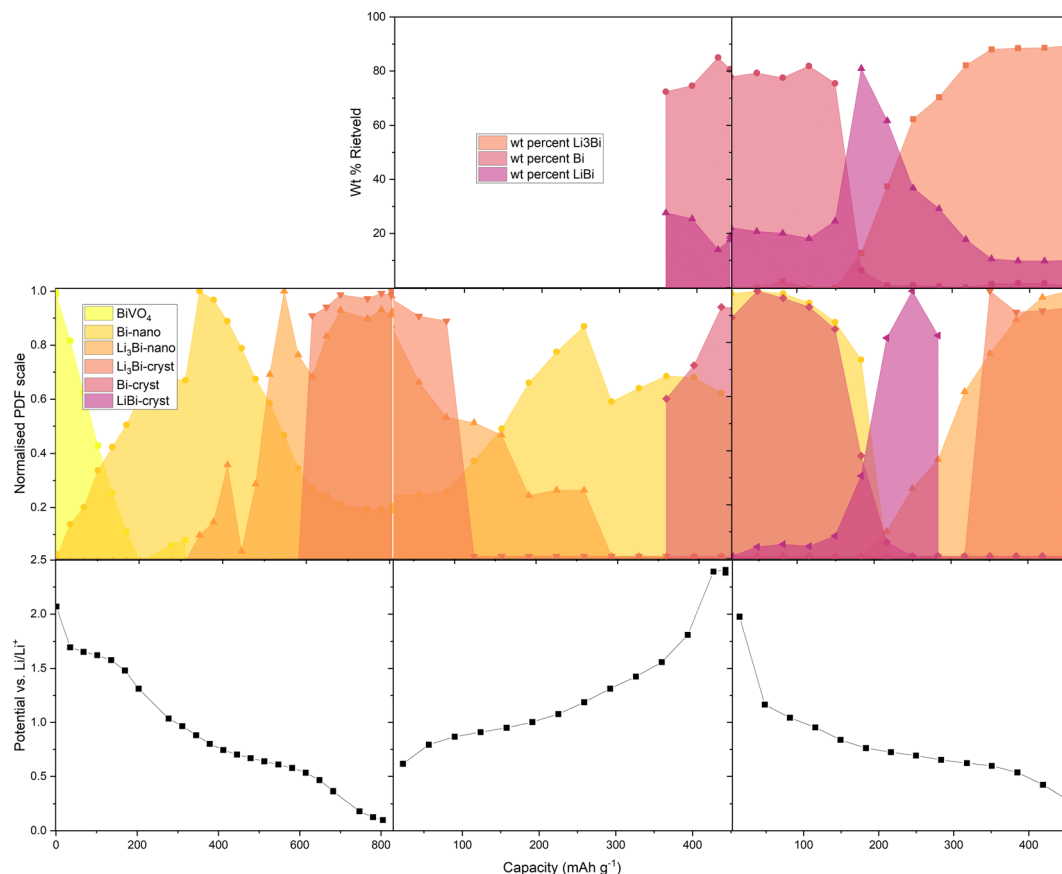


Fig. 3 Phase changes in BiVO_4 and Li_3Bi during half-cell cycling in the *operando* TSCT experiment from fitting of PDF (normalised PDF scale) and XRD data (weight % Rietveld).

The Li_3Bi which appears in the second lithiation is very clearly split into bulk ($r = \sim 30 \text{ \AA}$) and nano size ($r = \sim 3 \text{ \AA}$) regimes in the PDF fit. The bulk crystal structure fits the XRD data, in which we see the Li_3Bi appearing in a very similar pattern to nano Li_3Bi in the PDF fit. Radial damping variations throughout the PDF fit are shown in Fig. S10 (ESI†).

We also see variations in the refined lattice parameters for the bulk Li_3Bi (ESI†, Fig. S11). A decrease in the lattice parameters as Li_3Bi forms during the first lithiation is followed by a slight increase in the cubic lattice parameter after the phase has become the dominant component of the electrode (above 600 mA h g^{-1}). The reverse occurs in the delithiation (Li_3Bi is dominant below 100 mA h g^{-1}). Before the Li_3Bi becomes the dominant phase in the second lithiation (above approx. 200 mA h g^{-1}), we see large estimated standard deviations (ESDs) in the a -axis values, meaning the results are less reliable at this stage, while the Li_3Bi phase fraction is small. The lattice parameter is slightly larger in the second lithiation than in the first and once the phase is established as the dominant electrode component there is an excellent agreement between the PDF and XRD lattice parameters, with very small ESDs. DFT calculations suggest that as the Li content in non-stoichiometric $\text{Li}_{3-\delta}\text{Bi}$ is reduced, the unit cell volume contracts (see ESI† Table ST1). Sub-stoichiometric $\text{Li}_{3-\delta}\text{Bi}$ is probably responsible for the changes in Li_3Bi lattice parameter during

cycling and it may also be that the small clusters of Li_3Bi which dominate in the first lithiation are actually $\text{Li}_{3-\delta}\text{Bi}$ with a reduced lattice parameter. The other phases show essentially constant lattice parameters in both PDF and XRD fitting.

2 Spatial distribution of the active material

From the area summed TSCT data we selected characteristic peaks that allowed us to map the main Li_xBi phases in PDF and XRD. The lack of Bragg peaks in the initial stages of Li_xBi cycling means that the XRDCT does not provide a clear picture of the phase changes (ESI†, Fig. S12), but in PDFCT the appearance and disappearance of the phases is easy to follow (Fig. 4), this progression can also be seen in Movie S1 (ESI†). A small amount of intensity in the maps of LiBi during the first lithiation agrees with the evidence of lithiation from Bi–Bi bond extension from the area-summed PDFs (Fig. 2).

From the maps we see no clear differences between the charge states at which different phases appear and disappear in the 3 TSCT slices through the thickness of the electrode (*i.e.* along the axis of the cell, ESI†, Fig. S13) at the rate of cycling used. Increases in the intensity of the PDFCT signals in the maps also fits with increasing crystallinity as the cycling process proceeds.

Comparing the PDFCT maps of Li_3Bi after the first and second lithiations, or the total intensity XRDCT maps at the



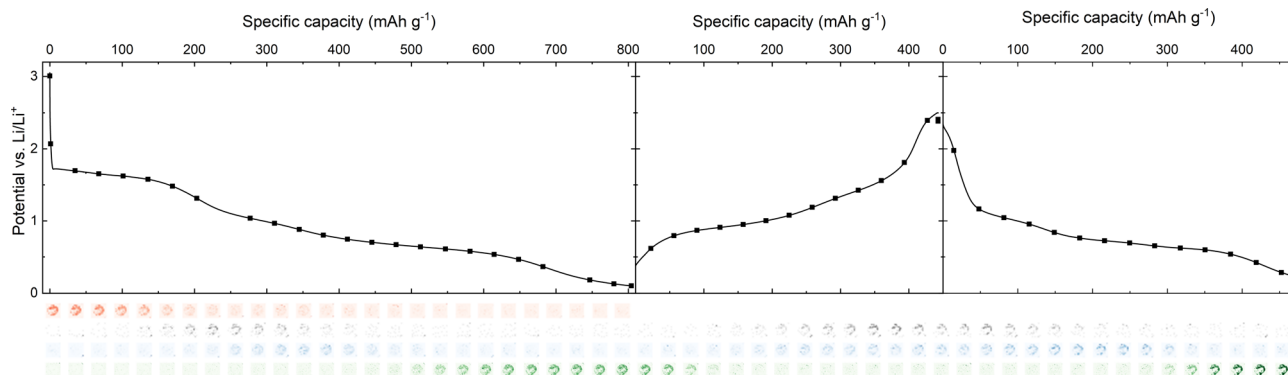


Fig. 4 PDFCT tomograms (bottom) from anode slice 0 showing the progression of the phases during *operando* cycling (capacity plot, top). Red = BiVO_4 , grey = Bi, blue = LiBi , green = Li_3Bi .

same stages, we see an apparent reduction of the area covered by the active material (Fig. 5). Difference maps of the changes in total XRDCT intensity compared to the start point show this even more clearly. We believe this indicates that as the Li_xBi material becomes more crystalline after the first lithiation (see above) it is confined to smaller areas, either due to agglomeration of the nanoparticles into larger crystallites or dissolution of some of the nanoparticle material into the electrolyte.

By counting pixels above a set threshold in the phase specific PDFCT and total scattering XRDCT maps we can quantify the coverage of each phase and of all the scattering material. This gives us firstly, another measure of the progress of the Li_xBi cycling processes and secondly, a numerical quantification of how the coverage of the piston with active material decreases with cycling (Fig. 6). The progress of the $\text{BiVO}_4/\text{Li}_x\text{Bi}$ cycling reactions observed with this method is very similar to that indicated by looking at and/or fitting the area-summed

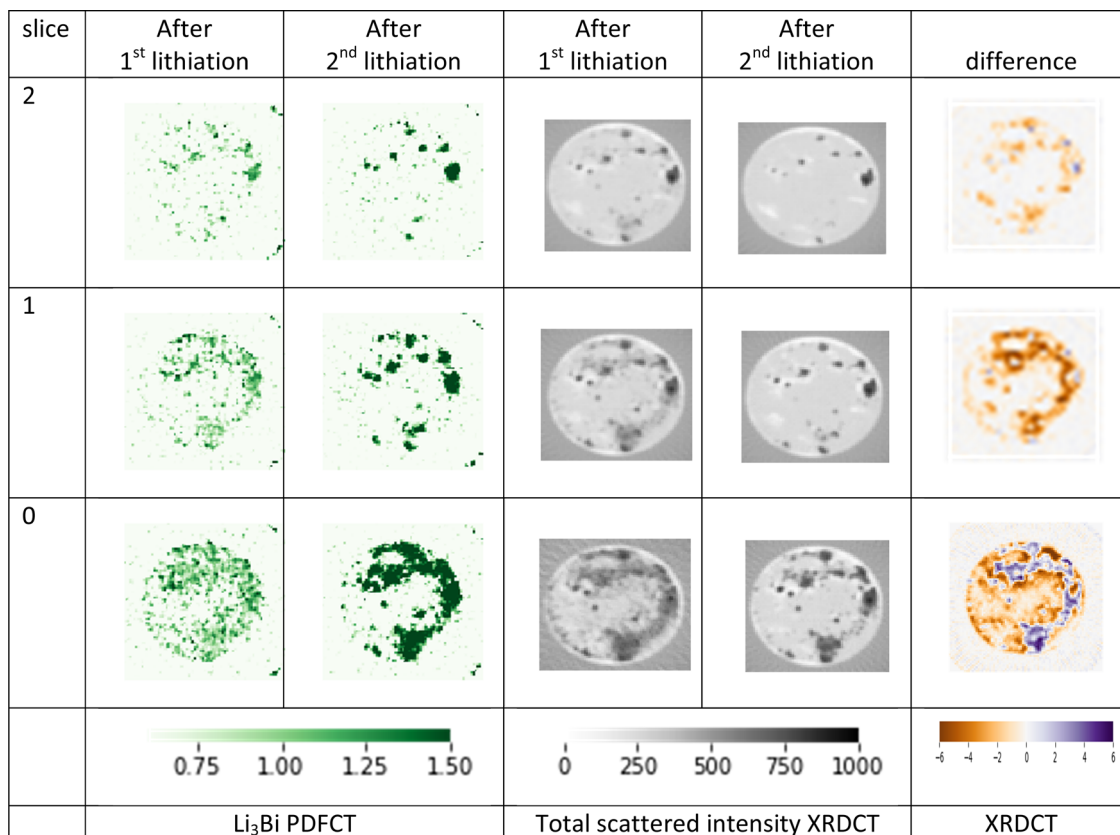


Fig. 5 Differences between the coverage of Li_3Bi (PDFCT) and total anode material (XRDCT for the total scattered intensity) at the end of the first and second lithiations. In the difference plot orange represents areas in which intensity decreases (material lost) and blue/purple areas where it increases (material gained).



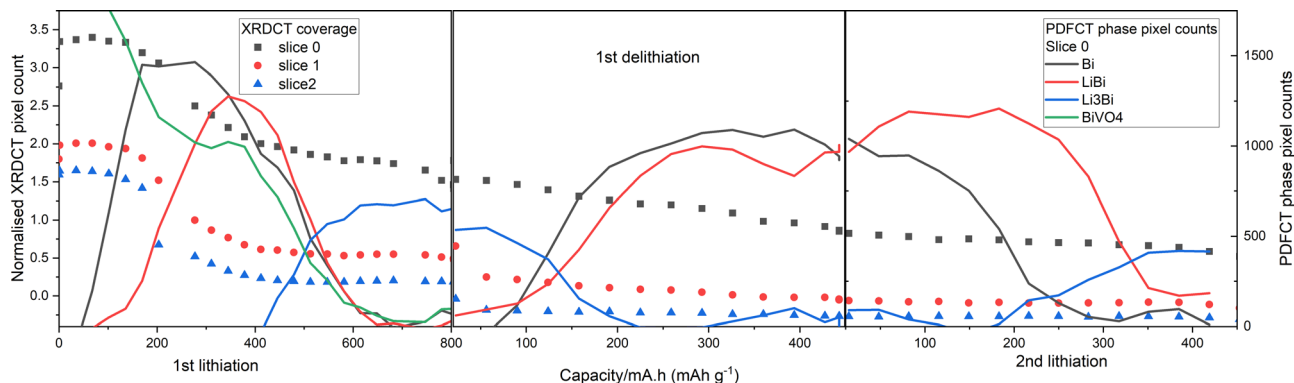


Fig. 6 Progression of the BiVO_4 and Li_3Bi phases in slice 0 during cycling in the *operando* TSCT experiment (solid lines), from pixel counting in the phase specific PDFCT images; and total coverage of active material from pixel counting in the total scattering XRDCT images (dotted lines). The PDFCT pixel count for BiVO_4 was not measured after the 1st lithiation as the phase was no longer present.

operando PDF data or looking at the series of phase-specific PDFCT maps, but the contribution for LiBi appears to be much more significant. This is probably due to significant overlap between the characteristic Bi and LiBi peaks in the radially rebinned PDFs used to make the PDFCT maps (see Methods section). The stronger signal from LiBi in the second lithiation is reproduced as in the other analyses. We can also see that the coverage of Li_3Bi is reduced in the second lithiation compared to the first. This follows the general trend of decline visible in the total coverage in the XRDCT tomograms for all three slices (dotted lines in Fig. 6).

3 Li_yVO_4 and carbon

The only sign of the Li_yVO_4 phase(s) in the area summed PDFCT data is a sharp peak at around 1.8 \AA from the V–O bond. No significant features beyond this can be fitted with any of the Li_yVO_4 phases which are stable according to DFT calculations (ESI† Table ST2), or with lithium/bismuth oxides. By fitting the 1.8 \AA peak position *vs.* charge state we can see a pattern that follows small features in the potential *vs.* capacity curve at the start of lithiation and the end of delithiation

(Fig. 7). This probably corresponds to cycling of Li in the Li_yVO_4 matrix, with oxidation and reduction of vanadate, occurring before lithiation of Bi and after delithiation of Li_xBi . A study of chemically lithiated $\text{Li}_z\text{V}_2\text{O}_5$ ($z = 0, 0.33, 0.66, 1.2, 0.48$, and 0.1) shows that the average V–O bond length increases up to about $z = 0.8$,¹⁷ but since we do not know the initial composition of the Li_yVO_4 phase (for possible candidates, see ESI† table ST2) it is difficult to speculate on the value of y from the V–O bond length. We can also see the peaks for 2 C–C interactions at 1.45 and 2.35 \AA , as observed previously for amorphous carbon.^{3,18}

Results 4 TEM/EDXS

To help us understand the TSCT results and how the structures observed fit together, we used a high-resolution transmission electron microscope (HRTEM) to collect images, electron diffraction (ED) and a high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) image of the sample. We also did energy-dispersive X-ray spectroscopy mapping in STEM mode (STEM-EDX) of material from a fully lithiated BiVO_4 anode (Fig. 8). In the low magnification TEM

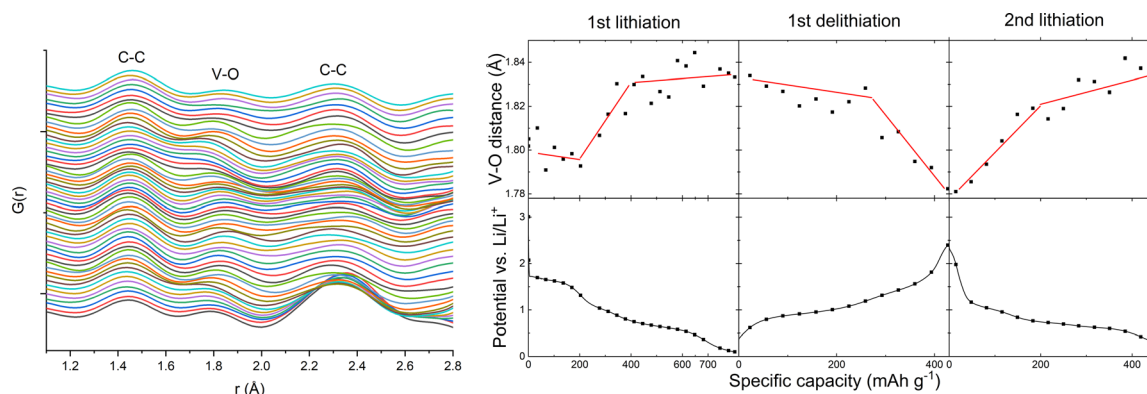


Fig. 7 Left, low- r region of the area summed PDF series for the entire *operando* experiment with V–O and C–C bond peaks labelled. Right, fitting of the V–O peak position plotted alongside the cycling data for the *operando* study. The red lines drawn over the data points are a guide (not linear regression) to indicate the steps in the peak position fits.

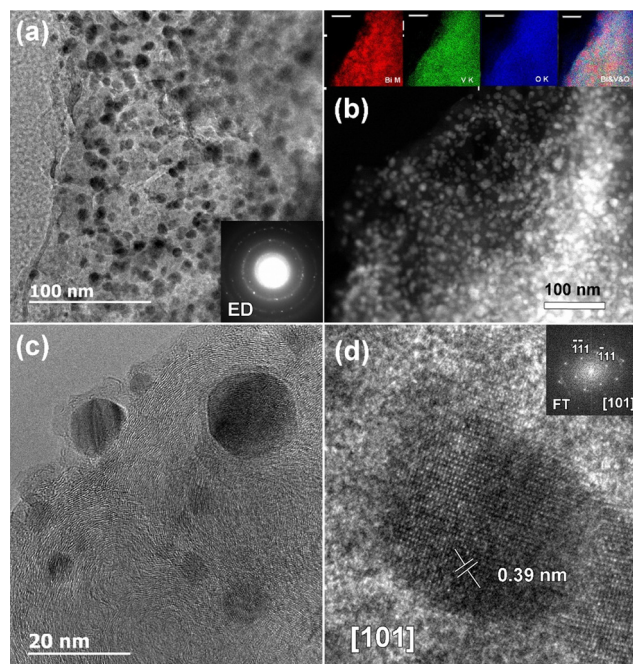


Fig. 8 (a) Bright field low magnification TEM image of material from a fully lithiated BiVO_4 anode and corresponding ring ED pattern, showing dark clusters of nanoparticles in a matrix of less dense material. (b) Low magnification HAADF-STEM image and corresponding EDX-STEM elemental maps of the same sample showing the different distributions of Bi (M-edge, red), V (K-edge, green), O (K-edge, blue) and overlaid color image. The scale bar is 50 nm. Note that the bright contrast regions in this image correspond to Li_3Bi nanoparticles. (c) Bright field overview HRTEM image and (d) selected Li_3Bi nanoparticle bright field HRTEM image along the [101] zone axis. The corresponding Fourier transform pattern is shown in the insert.

(Fig. 8a) and HAADF-STEM images (Fig. 8b) we can see 5–20 nm particles. The ring ED patterns allow us to identify these as Li_3Bi . The size range is rather larger than the average size of around 4.5 nm (45 Å) from the PDF fits, due to the wide distribution of particle sizes in the anode samples. They are distributed over a less organized matrix. The ED data (Fig. S14, ESI†) also show Li_3Bi diffraction patterns from different areas containing a mixture of single crystal and powder type diffraction data, supporting our use of different crystallite sizes to fit the area summed PDFCT data. Some of the particles appear to be covered by a layer of less dense material (Fig. 8c), presumably Li_yVO_4 , but possibly carbon or some kind of SEI as observed for Bi_2MoO_6 .¹⁹ In the lower magnification EDX-STEM maps (Fig. 8b), the vanadium and oxygen signals cover the whole sample area uniformly, supporting the theory that Li_yVO_4 forms an amorphous mass without long range order, while the Bi map is more mottled, showing a mixture of Li_xBi cluster sizes spread over the area. The overlaid image makes the clustering of Bi very clear. Fig. 8d shows a HRTEM image of single Li_3Bi nanoparticle imaged along the [101] zone axis, along with the Fourier transform of the image, showing sharp spots. These confirm the Li_3Bi structure and good crystallinity of the nanoparticles.

Discussion

The basic mechanism of lithium cycling in the BiVO_4 based anode is confirmed.¹² BiVO_4 converts irreversibly into Bi metal and Li_yVO_4 . The nano Bi formed is lithiated directly to Li_3Bi in the first lithiation, but there is a suggestion of the presence of LiBi (small intensity mapping signal: Fig. 5 and 6, slight extension of Bi–Bi bonds in the Bi cluster: Fig. 2). Delithiation proceeds directly from Li_3Bi to Bi metal. The second lithiation clearly includes crystalline intermediate LiBi . This is consistent with the dQ/dV cycling data (Fig. S4, ESI†) which show a single, broad anodic peak in the first lithiation which clearly splits into two sharper peaks in the second lithiation. There may be two parallel mechanisms: $\text{Li} \rightarrow \text{Li}_3\text{Bi}$ and $\text{Li} \rightarrow \text{LiBi} \rightarrow \text{Li}_3\text{Bi}$. Mechanisms with and without intermediates, occurring in parallel and depending on the rate of charge/discharge, have been reported for phosphorus sodium ion battery anodes.^{8,20}

Both area-summed data (increasing maximum r in the PDFs and appearance of Bragg peaks in conventional XRD) and maps from TSCT show that crystallite size increases with cycling. This has been observed in other bismuth metallate anodes (in both LIBs and NIBs),^{12,19,21} as well as for other conversion/alloying anode materials *e.g.* bismuth sulphide.²² TSCT maps also show that the area covered with active material decreases during this process. This could be caused by crystal growth fed by the amorphous material (which is used up and shifted to the crystalline areas) or dissolution of the smaller particles of active material into the electrolyte (leaving only the crystalline areas). Reactions with the electrolyte have been suggested as reasons for longer term degradation of electrode materials including Prussian blue analogues²³ and Bi_2MoO_6 .¹⁹

Post-mortem TEM shows that 5–20 nm Li_xBi particles sit in a Li_yVO_4 matrix. TSCT shows that this has no observable or fit-able order beyond the first V–O bond. Despite this lack of order, variations in the V–O bond length linked to specific features in the charge/discharge curves are visible in the area summed *operando* PDFCT data. Two PDF peaks are attributed to the carbon additive.

Based on the above, we propose the following mechanism:

1. BiVO_4 breaks down on lithiation to nano-clusters of Bi metal in a Li_yVO_4 matrix. This process cannot be reversed.
2. The Bi-nano-clusters are lithiated, possibly *via* small amounts of LiBi -like intermediates, before fully lithiated Li_3Bi is formed, now with particles in two size regimes.
3. Delithiation goes directly from Li_3Bi to Bi metal followed by (partial) delithiation of Li_yVO_4 . Crystalline Bi metal is observed alongside nanosized amorphous Bi.
4. The second lithiation begins with lithiation of Li_yVO_4 and proceeds from Bi (nanoclusters and crystalline) *via* LiBi (crystalline) to Li_3Bi (crystalline + amorphous). It seems likely that some of the amorphous component is directly lithiated to Li_3Bi .
5. As crystalline Li_xBi appears and the crystallites grow, the active area of the anode reduces. This may be a factor in longer-term loss of capacity and activity.

A convex energy hull diagram plotted from DFT minimized structures in the Li_xBi system (ESI† Fig. S15) shows a very



straight line from Li_3Bi (the lowest point of the hull) to Bi metal passing through LiBi and a theoretically stable Li_2Bi structure (not experimentally known). This means that there is very little difference in energy between the route that has the LiBi structure as an intermediate and the direct conversion of Bi to Li_3Bi . The hull thus helps to explain how steps 2 and 4 of the mechanism above can bypass LiBi (partly or completely) and go directly from Bi to Li_3Bi in the same way that the delithiation in step 3 goes directly from Li_3Bi to Bi. This appears to be dependent on the level of crystallinity, with crystalline Li_xBi more likely to follow the route *via* LiBi during lithiation, perhaps due to slower kinetics in the larger crystallites. The mechanism seems to fit somewhere between those observed in *operando* studies of NIBs with bismuth metal²⁴ and Bi_2MoO_6 ¹⁹ anodes, in which NaBi is formed as an intermediate during sodiation and desodiation; and phosphorus⁸ anodes, where NaP is formed during sodiation but not during desodiation and (dis)charge rate may also influence the route followed.

By collecting three TSCT slices through the anode during the *operando* experiment we obtained information in the axial dimension of the battery, which shows that within the time-scale of the experiment (at a cycling rate of approximately C/8) there is no observable difference in the extent of lithiation between adjacent slices along the axis of the cell collected within 15 minutes of one another. This suggests that despite the low electronic conductivity of all the theoretically possible phases that could exist in the Li_yVO_4 matrix (calculated by DFT to be either semimetals or indirect bandgap semiconductors, see Table S2, ESI†), the overall anode composition (including the conductive super-P carbon additive) remains capable of conducting both Li ions and electrons during the *operando* experiment.

Conclusion

The mechanism of BiVO_4 LIB anodes is extremely complex and the behavior of the various amorphous components has proved very difficult to track. Using a single TSCT dataset analyzed with multiple methods, backed up with *ex situ* TEM on a lithiated anode, we have obtained a very detailed, time resolved picture of how the metallate anode functions at a range of length scales. We have confirmed the basic mechanism of bismuth (de)lithiation, which explains the majority of the capacity, and furthermore, reveal previously hidden complexity in this mechanism. For the first time the atomic structure and activity of the Li_yVO_4 matrix are clearly visible in the area averaged *operando* PDFCT data, which also contain information on the Super-P carbon additive, helping us to understand sources of anode capacity beyond Li_xBi . The *operando* TSCT mapping and TEM data reveal larger scale changes in the distribution and structure of the active material and hint at how the longer-term deactivation of the anodes might be triggered by the growth of large crystallites within the amorphous Li_yVO_4 matrix and/or dissolution of the active material into the electrolyte.

The richness of information obtained from the *operando* TSCT is beyond anything we have previously encountered for a

single *operando* experiment. Not only does this dataset prove once again that this method is superior to typical transmission *operando* total scattering cells for measuring high quality PDF data on complex materials, it also shows that PDFCT imaging can be a highly effective tool for studying the microstructural changes in a working electrode. Never before has an amorphous electrode been imaged in real time with chemical selectivity. Alongside the *post-mortem* TEM, TSCT provides a tremendous insight into the entire cycling process at both nano and micro scales. Although the experimental methods and data processing are complex and a limited number of experimental facilities are currently available, we believe that the effort is worthwhile: our success in unravelling a very complex cycling process in a multi-phase, non-crystalline system points to tremendous possibilities for understanding materials and improving battery performance. The detail we have obtained on the amorphous and weakly scattering components of the anode (Li_yVO_4 , super-P carbon) suggests a bright future for TSCT when applied to difficult systems like silicon⁷ that have so far proved very difficult to study with *operando* total scattering.

Conflicts of interest

There are no conflicts to declare.

Methods

Sample preparation

The oxide-carbon composite $\text{BiVO}_4\text{:C}$ was prepared by milling commercial BiVO_4 (99.9%, Alfa Aesar) powder and conductive carbon black (C, Timcal Super P) in a 7:3 mass ratio under argon atmosphere (99.999% purity). A Fritsch Pulverisette 23 was used with a 10:1 ball-to-powder ratio at 50 Hz for 20 min. The grinding balls and bowls were made of steel.

Electrode preparation

The working electrode was prepared by spreading a slurry composed of 70 wt% of $\text{BiVO}_4\text{:C}$, 10 wt% of conductive carbon black (Super P, Timcal) and 20 wt% poly(acrylic acid) (PAA, Sigma Aldrich) as binder, dissolved in ethanol, directly on the current collector (copper foil or glassy carbon for *operando* TSCT measurements) with a paint brush. The electrodes were dried under vacuum at 60 °C overnight.

Battery assembly

Half cells were assembled in coin cells (type CR2032) or in specially constructed electrochemical cells for the *operando* TSCT measurements. For coin cell measurements the electrodes deposited on Cu current collector foil were then punched into discs with a mass loading of active material of about 1 mg cm^{-2} . For *operando* measurement slightly thicker electrodes with an active material mass loading of about 2.5 mg cm^{-2} were prepared (the accuracy of the mass loading calculation was limited by the balance available when the samples were prepared at the ESRF, which was only accurate to $\pm 1 \text{ mg}$, this



in turn leads to problems calculating the charge rate (C), potentially leading to reduced observed capacity due to higher than calculated C -rate. The working electrode was separated from the lithium metal disk as counter electrode by electrolyte soaked glass fibres (GF/C, Whatman). The electrolyte was a 1 M solution of LiPF_6 in ethylene carbonate/diethyl carbonate (EC/DEC, 1 : 1 by volume) solution (Sigma Aldrich). The battery was galvanostatically cycled in a voltage range of 0.01 V to 2.5 V vs. Li/Li^+ at 100 mA g^{-1} using a Biologic SP150 battery cycler. The specific capacity/current values are expressed based on the mass of BiVO_4 . The approximate C -rate for the *operando* experiment was $C/8$.

Operando TSCT

The *operando* electrochemical cell design is based on our previous design.⁸ In the new version (Fig. S[cell], ESI†) the sample container consists of a fully X-ray amorphous sealed quartz tube containing two glassy carbon pistons (3 mm in diameter) on which the working electrode and Li metal counter electrode are directly deposited. The pistons are pressed together by a spring to create electric and ionic contact within the electrochemical cell. The sample container is aligned such that the approximately 30 μm thick working electrode layer is oriented in plane with the X-ray beam and can be rotated by up to 360° around its vertical axis during the measurement.

TSCT data were collected at beamline ID15A of the European Synchrotron (ESRF). The cell was mounted on a fast translation/rotation stage capable of movement in the x , y and z planes of the diffractometer. Powder diffractograms were collected at an energy of 87 keV using a 40 μm wide and 13 μm high beam. The beam was focused using X-ray lenses. A Pilatus 3X 2M CdTe area detector was used to collect the diffractograms. Detector distance (28.1 cm) and other detector constants were calibrated from diffractograms of a capillary sample of NIST SRM674b CeO_2 using PyFAI.²⁵ The TSCT cell was rotated through 180° and the beam was scanned across the 5 mm wide cell with the 40 μm beam width as the step size. A line scan of 110 diffractograms across the sample was collected for every 3° rotation from 0 to 180° , giving 61 projections to reconstruct for each TSCT slice. Three slices (slice 0, 1 and 2) were collected in the anode at heights of 10, 25 and 40 μm above the piston surface. Each slice took approximately 7 minutes to collect. This resulted in a voxel (3D pixel) size of $40 \times 40 \times 13 \mu\text{m}^3$.

In-house ID15A software was used to remove air scattering contributions to the raw diffraction data based on diffractograms collected outside the sample. The data were integrated using PyFAI and reconstruction was carried out using the ID15A Matlab scripts. A trimmed mean filter was applied to remove artefacts due to larger single crystals. Integrated powder patterns for each voxel were converted to $G(r)$ using PDFGetX3 in batch mode. TSCT processing was carried out in two different modes:

1. For analysis of the reaction mechanism only diffraction and total scattering data from the active material in each tomographic slice was taken into account. The initial distribution (which was rather inhomogeneous due to brush painting

of the electrode slurry onto the piston) of crystalline BiVO_4 was determined using standard XRDCT methods (ESI† Fig. S1). Based on this distribution the powder patterns of all voxels containing active material (all other voxels were masked and not account for) were summed to a single high quality powder pattern for each tomographic slice and converted to $G(r)$ and $I(2\theta)$. The slice closest to the current collector (slice 0 in Fig. 1), contained the most active material and was used for analysis of the mechanism.

2. For TSCT mapping a further set of $G(r)$ data were produced in which the data from individual voxels were rebinned with a large radial step size (0.2 \AA) to reduce the noise levels. The difference between the PDFs obtained from an average of the entire active area and a single voxel after 0.2 \AA rebinning is shown in ESI† Fig. S2. XRDCT maps were produced using the untreated $I(q)$ data. The $G(r)$ maps were produced using a single characteristic r value for each of the phases ($\text{BiVO}_4 = 3.6 \text{ \AA}$, $\text{Bi} = 3.0 \text{ \AA}$, $\text{LiBi} = 3.4 \text{ \AA}$, $\text{Li}_3\text{Bi} = 4.8 \text{ \AA}$) with no background subtraction. $I(q)$ maps were produced by integrating the area under characteristic peaks for each phase ($\text{BiVO}_4 = 1.29\text{--}1.42 \text{ \AA}^{-1}$, $\text{Bi} = 2.715\text{--}2.805 \text{ \AA}^{-1}$, $\text{LiBi} = 2.345\text{--}2.445 \text{ \AA}^{-1}$, $\text{Li}_3\text{Bi} = 1.55\text{--}1.7 \text{ \AA}^{-1}$) after background subtraction.

The unprocessed $I(Q)$ data were also used to create maps equivalent to traditional X-ray adsorption tomography slices using the sum of all Q channels. These were used to study the variation in total volume of anode material during cycling.

Ex situ TEM

Samples for *ex situ* TEM were taken from an electrode fully lithiated in a coin cell. The cell was disassembled and the electrode particles were crushed, washed with DEC and transferred to a TEM grid in an argon filled glove box. The TEM grid was then carefully transferred to the TEM chamber to minimize contamination with atmospheric species.

The transmission electron microscopy (TEM) measurements were performed using a double aberration corrected JEM ARM200F cold FEG microscope equipped with CENTURIO large angle EDX detector, ORIUS CCD camera and Quantum GIF.

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