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## Protonation and orientation: a computational approach to cocaine diffusion through a model membrane†

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We study the diffusion of cocaine through a DMPC lipid bilayer as an example of a protonable, amphiphilic molecule passing a biological membrane. Using classical molecular dynamics simulations, the free energy surfaces are computed applying the umbrella sampling technique for the protonated and the neutral molecule. For the combined surface, we numerically solve the diffusion equation at constant flow and for time-dependent concentrations. We find a potential of mean force dominated by a barrier of 3.5 kcal mol<sup>-1</sup> within the membrane, and a pH-dependent entry and exit barrier of 2.0 kcal mol<sup>-1</sup> and 4.1 kcal mol<sup>-1</sup>, respectively. This behaviour can be rationalized chemically by the amphiphilic nature of the molecule and the change of its protonation state while passing the membrane. Diffusion through the barriers is 3.5 times slower than along the membrane, and the typical time scale of passage amounts to 0.1 ms. We discuss biochemical and medical implications of our findings, and comment on the mechanism of the drug passing the blood–brain barrier.

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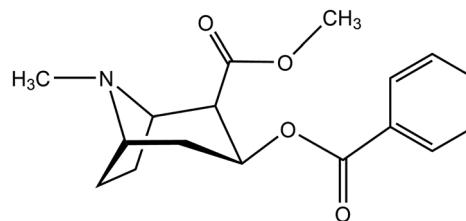
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### 1. Introduction

Lipid bilayers are a fundamental structure in biological systems, as they maintain cell functionality by compartmentalizing reaction areas and act as selective barriers.<sup>1</sup> Understanding these structures and their interaction with molecules or their aggregates – ranging from protons over the smallest signal molecules to protein complexes – is an important topic in biochemistry, biophysics and medicinal chemistry. In the past decade, transport through membranes has come into the focus of classical molecular dynamics computer simulations. This is due to an increase of computer power, the adaptation of program suites for GPU-based machines<sup>2,3</sup> and the effective implementation of algorithms for the computation of free energy surfaces.<sup>4</sup> As examples, we have the diffusion of medical drugs or drug-like compounds,<sup>5–7</sup> small polymers,<sup>8</sup> oligopeptides<sup>9</sup> or particles that reach into the nanoscale.<sup>10,11</sup> A recent review covers the state of the art in this field.<sup>12</sup>

In our work, we focus on a single molecule passing a membrane, cocaine (2-(methoxycarbonyl)-3-(benzyloxy)tropane).



It has been used as a local anesthetic,<sup>13</sup> but is also a psychoactive, highly addictive drug. Its abuse has severe medical, social and economical implications.<sup>14</sup> At synaptic clefts, it inhibits the reuptake of monoamine neurotransmitters, such as dopamine. In turn, the increasing concentration of dopamine in the synapses activates postsynaptic dopamine receptors, which is experienced as a constant reward while the drug is operative. To exert its effect, the drug has to cross the blood–brain barrier. This barrier is composed of endothelial cells that enable permeation either by passive diffusion through lipid membranes or by dedicated, substance-specific transport proteins. For cocaine, it has been assumed that the molecule can passively permeate the barrier in its unprotonated form due to its sufficient lipophilicity. This lipophilic character is attested by its solvation behaviour:<sup>15</sup> the molecule is slightly soluble in water (1 g/600 ml), well soluble in alcohol (1 g/6.5 ml),

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chloroform (1 g/0.7 ml), ether (1 g/3.5 ml) and olive oil (1 g/12 ml), but to a much lesser extent in petrolatum (1 g per 30–50 ml). A curious clinical case report has been published by Jakkala-Saibaba *et al.*,<sup>16</sup> who have successfully treated an overdose patient applying a lipid emulsion of soy bean oil (a total of 100 ml *via* 500 ml of 20% Fresenius Kabi Intralipid). This also highlights the lipophilic nature of the cocaine molecule.

However, at a  $pK_a$  value of 8.61, cocaine in neutral aqueous solution acts as a weak base and is predominantly protonated, similar to nicotine and many neurotransmitters.<sup>17</sup> It has been suggested that the molecule in its cationic form crosses the blood–brain barrier through active transport *via* a polyspecific  $H^+$ -antiporter, which also recognizes the psychoactive molecules clonidine and nicotine.<sup>18</sup> Effects hinting at active transport *via* an antiporter have been found in mouse hCMEC/D3 brain endothelial cells, which serve as a model system of the human blood–brain barrier. Here, a passage rate larger than anticipated by simple diffusion corresponds to a high synaptic activity. While an outline of possible features of the antiporter has been made recently<sup>19</sup> the suggested protein complex or parts thereof have to the best of our knowledge neither been isolated nor characterized any further.

The question whether cocaine and related molecules cross the blood–brain barrier assisted by a protein or protein complex or by diffusion along a concentration gradient has an impact on addiction liability and treatment:<sup>18</sup> a protein mediating the transport would present an additional target for medication *via* its inhibition. Furthermore, a genetic disposition towards an overexpression of the antiporter or mutations in this biochemical machine may be related to the sensitivity towards the drug.

As the pharmacological impact of cocaine is related to its molecular structure, the molecule has been approached by theoretical methods. An early attempt to describe the cocaine geometry and its energetics has been performed by Villar and Loew using molecular mechanics and semi-empirical calculations.<sup>20</sup> Recently, Fagan *et al.* have combined chiroptical methods and DFT calculations to access the structure of cocaine in solution,<sup>21</sup> and three of the present authors have studied the naturally occurring molecule and its diastereomers in detail.<sup>22</sup>

The remaining part of the article is organized as follows. In the next section, we describe the umbrella sampling setup and the resulting free energy surfaces. These are followed by the description of the unconstrained molecular dynamics simulations. The emerging results permit a chemical interpretation of the passage of the molecule through the membrane, and they provide the input parameters for the following two sections. In there, the diffusion equation is solved numerically at constant flow and for time-dependent concentrations, giving rise to an effective diffusion coefficient and the relevant time scales of the process. All results are subject to a final discussion and concluding remarks.

## 2. Umbrella sampling

### 2.1. Methods

The diffusion of molecules through liquid bilayers is usually outside the time scale of straightforward classical molecular

dynamics simulations. In addition, to obtain a proper statistical basis, several hundred events have to be analyzed. Alternative setups at cyclic boundaries comprise two lipid bilayers and two water slabs with different initial concentrations of the diffusing molecule, they are subject to the same restrictions on the time scale.

To simulate rare events such as membrane permeation, a series of harmonic restraints can be introduced, which are operative on the distance between the centers of mass of a molecule and selected atoms of the bilayer.<sup>4</sup> As we center the membrane at the origin of the Cartesian coordinate system and arrange it in the  $xy$ -plane, the reaction coordinate simply equals  $z$ . The probability  $P_k(z)$  of finding the center of mass difference at  $z$  within a window  $k$  can be related to the free energy  $\Delta G_k$  of the system,

$$\Delta G_k(z) \simeq \Delta A_k(z) = -k_B T \ln P_k(z) - w_k(z) + F_k, \quad (1)$$

where  $w_k(z)$  is the bias potential and  $F_k$  a constant yet to be determined. This is performed by matching the distributions of different windows by the weighted histogram analysis method (WHAM)<sup>23,24</sup> to obtain a smooth free energy curve. Within each window, the system has to be equilibrated before sampling, and the probability distributions have to show sufficient overlap between neighboring windows. This method is usually referred to as umbrella sampling, and we note that it is not limited to harmonic constraint potentials.<sup>25</sup>

Force field parameters for cocaine in its neutral and its protonated form have been generated using the Amber antechamber utility.<sup>2</sup> Input geometries stem from high-level quantum chemical computations,<sup>22,26</sup> and we have also checked the resulting charge parameters stemming from the Austin model 1 calculations (with bond charge corrections, AM1-BCC) against this reference. The TIP3P force field has been used to describe the water molecules,<sup>27</sup> and the lipid17 force field has been applied on the dimyristoylphosphatidylcholine (DMPC) molecules.<sup>28</sup> Three system sizes have been studied, each containing a single cocaine molecule and (i) 50 lipid molecules and 1866 water molecules, (ii) 76 lipid molecules and 2918 water molecules and (iii) 196 lipid molecules and 7470 water molecules. The largest systems comprise 45 581 atoms and have a dimension of  $\sim 80 \times 80 \times 80 \text{ \AA}^3$ . The systems have been generated using the PACKMOL-Memgen utility.<sup>29</sup>

The energy of the system was minimized for a total of 16 000 cycles, switching from steepest descent to conjugate gradient minimization after 8000 cycles without the use of the SHAKE algorithm. The system was first brought to a temperature of 100 K, and in the second step to the production temperature (300 K) while the lipid molecules were restrained by a harmonic potential with a force constant of  $10.0 \text{ kcal mol}^{-1} \text{ \AA}^{-2}$ . The simulation was run at constant volume. Bonds containing hydrogen atoms were restrained by the SHAKE algorithm. The externally coupled temperature was held constant by the Langevin thermostat with a collisional frequency of 1 ps. The system was equilibrated for 710 ps at constant pressure and with periodic boundary conditions. Constant pressure was enforced by applying semiisotropic pressure scaling and using



a constant surface tension with the interfaces in the  $xy$  plane. The externally coupled temperature (300 K) was again held constant by the Langevin thermostat with a collisional frequency of 1 ps.

Finally, the simulation was run under constant pressure and periodic boundary conditions. The cocaine molecule was pulled from the center of the lipid bilayer towards the water phase over a simulation time of 64 ns and with a pulling rate of 0.5 Å per ns. This trajectory was then split into 69 time windows, which were simulated for 100 ns each by applying a parabolic potential with a force constant of 2.5 kcal mol<sup>-1</sup> Å<sup>-2</sup> to the cocaine molecule. The output was analysed using the weighted histogram analysis method<sup>23,24</sup> (WHAM), for which a program is included in the Amber 20 simulation package.<sup>2</sup> The histograms are shown as ESI,† they exhibit a broad overlap, as essential for the resulting analysis. From the WHAM analysis, 680 data points were obtained, which are also presented in the ESI.†

All the simulations within the umbrella sampling windows were run under periodic boundary conditions and the particle-mesh Ewald method was used for electrostatic interactions. The cutoff distance was 10.0 Å. The temperature was held constant by the Langevin thermostat and the pressure by the Berendsen barostat. A simulation snapshot is shown in Fig. 1.

## 2.2. Free energy surface

Utilizing the umbrella sampling method, we obtain two curves for the potential of mean force as a function of the coordinate,  $z$ , which points perpendicular to the lipid bilayer, and has its zero at the center of the bilayer. One set of data represents the protonated, the other the unprotonated form, the corresponding raw data are given as ESI.† As the membrane is symmetric w.r.t. an inversion of  $z$ , the curves have been recorded for positive values and mirrored. Using an empirical force field, it is not possible to determine the offset of the free energy curve between the protonated and the neutral form of the cocaine molecule. Nevertheless, they can be calibrated making use of the experimental  $pK_a$  value of cocaine, 8.61. In the aqueous phase, we have

$$\Delta G = 2.3k_B T(pK_a - \text{pH}) = 2.2 \text{ kcal mol}^{-1}, \quad (2)$$

which we apply at pH = 7 and for  $|z| > 30$  Å, *i.e.* in a range where both free energy curves become constant.

In Fig. 2, we show the free energy as a Boltzmann-weighted average of the calibrated cationic and the neutral curves:

$$\Delta G(z) = \frac{\Delta G^{(0)}(z) \exp(-\beta \Delta G^{(0)}(z)) + \Delta G^{(+)}(z) \exp(-\beta \Delta G^{(+)}(z))}{\exp(-\beta \Delta G^{(0)}(z)) + \exp(-\beta \Delta G^{(+)}(z))}, \quad (3)$$

where the (0) and (+) superscripts refer to the neutral and to the positively charged species, respectively. The resulting potential of mean force (POMF) as a function of  $z$  is dominated by the cationic curve for  $|z| > 16.3$  Å, and by the neutral POMF otherwise, the crossovers are indicated in the Fig. 2 by vertical dotted lines. Rather than computing the energy contribution of

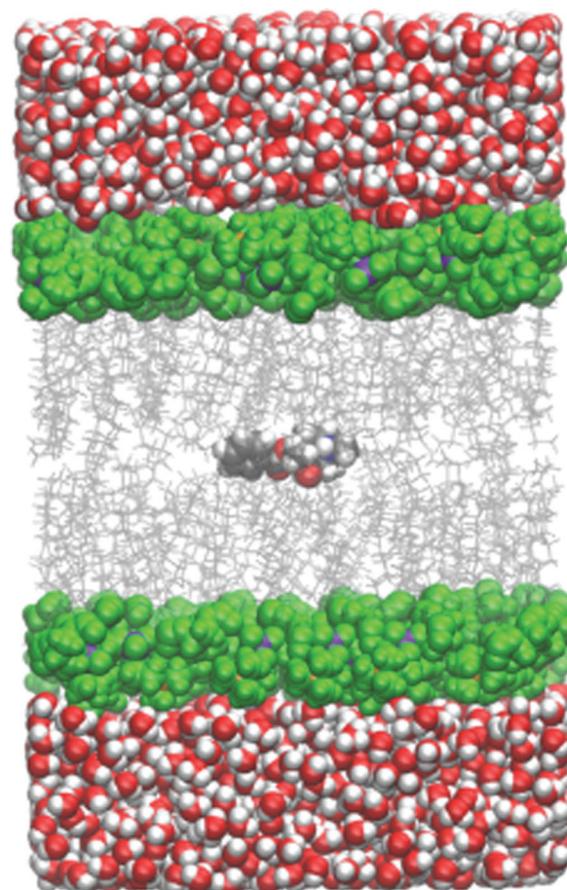


Fig. 1 Simulation snapshot with cocaine originally placed in the lipid phase. The DMPC head group carbon atoms are shown as green spheres, side chains are depicted as a grey wireframe model. All other atoms follow the CPK coloring scheme.

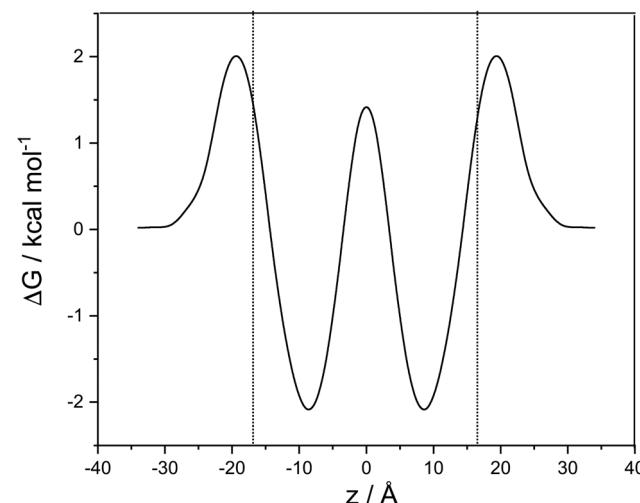


Fig. 2 Potential of mean force as a function  $z$ , the coordinate perpendicular to the membrane. Between the dotted lines cocaine exists as a predominantly neutral species, left and right of the lines the protonated form is dominant.  $\Delta G$  is in kcal mol<sup>-1</sup>,  $z$  in Ångstroms.

the dominant species using eqn (3), the most abundant form can also be detected from the calibrated neutral and charged



free energy curves, as shown as Fig. S3 in the ESI.<sup>†</sup> In the aqueous phase, the free energy of the cation (red curve) is smaller than that of the neutral molecule (black curve). The curves coincide at  $z = 16.7 \text{ \AA}$ , and up to the center of the membrane, the neutral molecule free energy is smaller than that of the cation. With decreasing pH, the free energy of the protonated species is shifted downwards (as indicated by the red arrow in Fig. S3, ESI<sup>†</sup>), while the free energy of the neutral species is unaltered. Hence, the point of intersection between the two curves is shifted to the left, and the spatial region in which the cation is the dominant species is extended.

A cubic regression spline with 28 grid points has been applied to obtain a smooth curve. We observe two global minima at  $z = \pm 8.8 \text{ \AA}$ , which correspond to a localization of the neutral cocaine molecule within the membrane, *cf.* Fig. 2. The two minima are separated by a barrier of  $3.5 \text{ kcal mol}^{-1}$ . Our conjecture about the nature of the molecule is a statistical one, the charged form is dominant in the aqueous phase, and the neutral molecule constitutes the majority form within the membrane.

Furthermore, we find an entry barrier of  $2.0 \text{ kcal mol}^{-1}$  and an exit counterpart of  $4.1 \text{ kcal mol}^{-1}$ . These barriers depend on the pH value in the two aqueous phases: increasing the pH value shifts the protonated curve towards the unprotonated one on the free energy axis, and the point of crossover towards higher values of  $|z|$ , *i.e.* in the direction of the aqueous phase. The height of both the entry and the exit barrier increase. Lowering the pH value, this behaviour is reversed, and we have lower activation barriers for exit and entry. In a similar sense, this also holds for aqueous phases with different pH values. Partition coefficients stemming from experiments<sup>30</sup> or reaction field computations<sup>31</sup> can be used in a thermodynamic model to compute the pH-dependent distribution of anesthetics in detail.<sup>32</sup> While we qualitatively discuss pH effects, the main focus of our work is on the transfer dynamics and its time scale. Below, we make use of the potential of mean force computed in this section and the diffusion coefficient parallel to the lipid layer – as estimated below – to address the time scales of transport in the following two sections.

Close to the center of the lipid membrane ( $< 2.5 \text{ \AA}$ ), we find a spurious effect in the potential of mean force of the protonated molecule. Its free energy is close to that of the neutral form, and the slope of the curve at  $z = 0$  is not vanishing, as expected for a symmetric system. Inspecting snapshots of the simulations (see ESI<sup>†</sup>), we can attribute this behaviour to a cation that is globally, but not locally centered in the membrane, and locally deforms the bilayer and is in contact with the water phase. This does not change the thermodynamics, and – if not a simulation artefact – is relevant only on a minute fraction of the diffusion path.

### 3. Unconstrained molecular dynamics

#### 3.1. Model and methods

Two series of simulations have been performed, one with an initial placement of the cocaine molecule in the aqueous phase,

and one with an initial localization in the hydrophobic part of the liquid bilayer. System sizes were identical to those used for the umbrella sampling simulations.

The energy of each system was minimized for a total of 5000 cycles, switching from a steepest descent to conjugate gradient procedure after 2500 cycles. Here, bonds involving hydrogen were not constrained. During the heating process, the system was brought to production temperature (300 K) at constant volume. Bonds with hydrogen were constrained applying the SHAKE algorithm. The system was equilibrated for 800 ps at constant pressure. A constant pressure was enforced by applying isotropic position scaling. In the production run, the system was simulated for 1000 ns at constant pressure. The coordinates were saved every 5000 steps for further analysis, the restart file was written every 5000 steps.

To determine the diffusion coefficient of cocaine in water, the *R*-cocaine molecule was solvated in 1000 water molecules in a simulation box of  $40 \times 40 \times 40 \text{ \AA}^3$ , and a molecular dynamics simulation was carried out. First, the energy of the system was minimized for a total of 5000 cycles, switching from steepest descent to conjugate gradient after 2500 cycles and without the use of the SHAKE algorithm. Then the system was brought to a production temperature of 300 K while being simulated under constant volume. Bonds including hydrogen were restrained by the SHAKE algorithm. The system was equilibrated for 800 ps at constant pressure by applying isotropic position scaling. Finally, the system was simulated for 2 ns at constant temperature and constant pressure.

If not stated otherwise, all the simulations described above were run under periodic boundary conditions and the particle-mesh Ewald method was used for electrostatic interactions. The cutoff distance was  $10.0 \text{ \AA}$ . The temperature was held constant by the Langevin thermostat and the pressure by the Berendsen barostat. The diffusion coefficient of cocaine was determined by using the diffusion command included in the Amber 20 simulation package.<sup>2</sup>

#### 3.2. Results

With the help of the concentration profile displayed in Fig. 3, the origin of the barrier finds a straightforward chemical interpretation. At a minimum, the neutral cocaine molecule is oriented with its hydrophobic phenyl ring embedded by likewise hydrophobic side chains. The tropane part of the molecule – with a stronger hydrophilic character – points towards the charged head groups of DMPC. Dragging the molecule deeper into the bilayer, this structural element now also experiences a purely hydrophobic environment, increasing the free energy. With due care and lacking a more appropriate terminology, we refer to this behaviour as amphiphilic, in line with recent literature.<sup>33</sup> While oriented, the neutral cocaine molecule does not function as a detergent due to its small size. We will further discuss these orientational effects in connection with the findings from the diffusion computations in the final section of the paper.

Trajectories from the simulation of a DMPC bilayer incorporating a cocaine molecule have been used to compute the



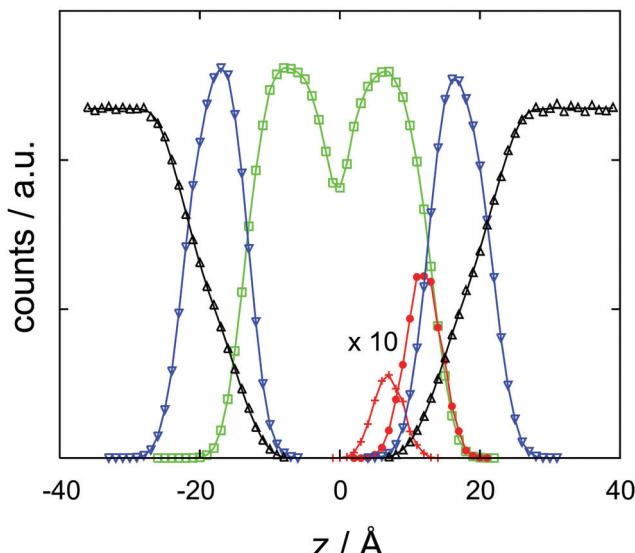


Fig. 3 Atom count histograms as a function of  $z$ , the coordinate orthogonal to the membrane, for a cocaine molecule initially localized in the lipid phase. Symbols and colors refer to the cocaine phenyl ring (+, red), cocaine tropane ring (●, red), DMPC head group (▽, blue), DMPC side chains, (□, green) and water (Δ, black). Note that the cocaine profiles have been multiplied by ten for convenience. Counts are dimensionless,  $z$  is in Ångstroms.

diffusion coefficient parallel to the bilayer. Care has been taken to analyse a time window that shows a lower boundary that is sufficiently large to exclude cage effects, and an upper boundary that is small enough to be unaffected by the diffusion or undulations of the membrane. In two dimensions, we expect a dependence of the root mean square displacement as a function of time that follows  $r_{\text{rmsd}}^2 = 4Dt$ . To look for a potentially anomalous diffusion and spurious effects induced by the membrane, we have also interpreted the exponent as a variable and analyzed  $\ln r_{\text{rmsd}}^2 = \ln(4D)/2 + n \ln t$ . We find  $n = 0.53 \pm 0.03$ , still within the range of classical diffusion, and  $D = 2.7 \pm 0.6 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$ . This should be compared to a self-diffusion coefficient of  $D = 1 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$  obtained for DMPC by NMR spectroscopy.<sup>34</sup> We note that the viscosity and diffusion coefficients for lipid bilayers comprise a range of several orders of magnitude, depending on the membrane composition and on the type of measurement.<sup>34–36</sup>

We have also computed the diffusion constants for cocaine in a pure water environment. Minimization, equilibration and relaxation have been performed as described above. The diffusion coefficient amounts to  $\sim 11 \times 10^{-9} \text{ m}^2 \text{ s}^{-1}$ , with a correlation coefficient larger than or equal to 0.999.

#### 4. Diffusion at stationary flow

Stationary flow between two reservoirs with fixed concentrations can be described by Fick's first law, relating the flow,  $\vec{j}$ , to the concentration gradient,  $\nabla c$ , with the diffusion coefficient,  $D$ , as a proportionality constant,

$$\vec{j} = D \nabla c \quad (4)$$

Above, we have computed the diffusion coefficient for a molecule confined to one of the minima of the free energy surface. Here, the molecule essentially performs a random walk parallel to the membrane. Perpendicular to the bilayer, the random walk experiences the free energy profile shown in Fig. 2, and  $D$  depends on the  $z$  coordinate,  $D(z)$ . We discretize the concentrations along the  $z$  coordinate using  $n = 101$  equi-distant steps  $\Delta z = 0.654 \text{ Å}$ , and arrive at

$$\vec{j}_{i,i+1} = D_{i,i+1} \frac{c_i - c_{i+1}}{\Delta z}. \quad (5)$$

At stationarity, the flow into each node has to be equal to the flow out of each node,  $\vec{j}_{i-1} = \vec{j}_{i+1}$ ,  $i$ , leading to a tridiagonal system of equations for the unknown concentrations

$$D_{i-1,i}(c_i - c_{i-1}) = D_{i,i+1}(c_{i+1} - c_i). \quad (6)$$

At the boundaries, we assume a constant diffusion coefficient. Without any loss of generality, we set  $c_{n+1} = 0$  and use a fixed  $c_0$ . On the left side of the system, we have  $2c_1 - c_2 = c_0$ , on the right side  $2c_n - c_{n-1} = 0$  holds. We assume the diffusion coefficient to be equal to that parallel to the bilayer if it moves downhill along the potential of mean force, and have it modified by a factor of  $\exp(-\beta \Delta G_{i,i+1})$  otherwise. The linear system of equations is solved using subroutines from the LINPACK suite,<sup>37</sup> and the resulting concentration profile is shown in Fig. 4. Within the lipid bilayer, it is slightly structured reflecting the positions of the free energy barriers, but its main trend is one not strongly deviating from a straight line.

Between two nodes, we can now readily compute the effective flow through the system,  $\vec{j} = D_{i,i+1}(c_i - c_{i+1})/\Delta z$  and derive an effective diffusion coefficient for transfer along a system with a length  $L$ ,  $D_{\text{eff}} = c_0 |\vec{j}|/L$ . For a linear arrangement of sites, as encountered here, computing the effective diffusion coefficient is equivalent to calculating the effective resistivity for a series of resistors. Thus,  $L/D_{\text{eff}} =$

$\Delta z \sum_i D_{i,i+1}^{-1}$  provides an alternative bypassing the solution of the

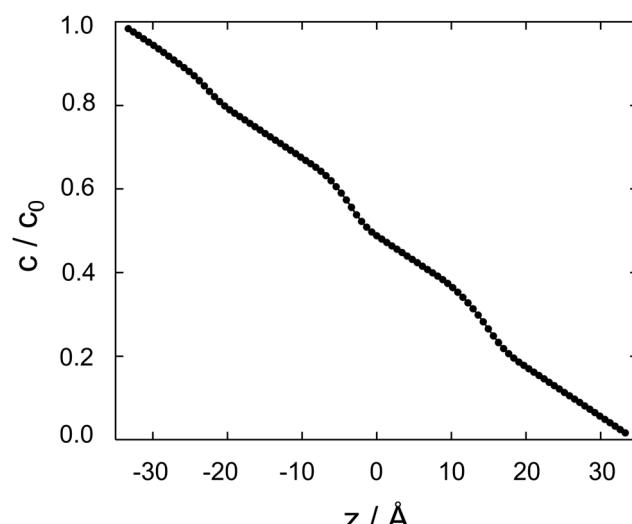


Fig. 4 Cocaine concentration profile,  $c/c_0$  as a function  $z$ , the coordinate perpendicular to the membrane. The normalized concentration is dimensionless,  $z$  is in Ångstroms.

system of equations formulated above. Numerically, we find  $D_{\text{eff}} = 7.2 \times 10^{-12} \text{ m}^2 \text{ s}^{-1}$ , a value that is considerably smaller than free diffusion parallel to the membrane, yet the two diffusion processes are not separated by an order of magnitude. We note that diffusion in spatially varying potentials has also been discussed by Livshits,<sup>38</sup> who provided diffusion equations corrected for thermally activated processes.

## 5. Time-dependent diffusion and rate theory

To explore the time scales involved in passing the membrane, we turn to time-dependent diffusion along the  $z$  coordinate, *i.e.* perpendicular to the lipid bilayer. We consider Fick's second law,

$$\frac{\partial c(z, t)}{\partial t} = D(z) \frac{\partial^2 c(z, t)}{\partial z^2} \quad (7)$$

and again discretize the system in its spatial coordinate using the abbreviation  $c(z, t) = c_i$ . We arrive at equations of the type

$$\frac{dc_i}{dt} = \frac{D_{i-1,i}c_{i-1} - (D_{i,i-1} + D_{i,i+1})c_i + D_{i+1,i}c_{i+1}}{\Delta z^2}. \quad (8)$$

They constitute a system of coupled linear, ordinary, first-order differential equations, or Pauli–Master equations, with a solution  $\vec{c}(t) = (c_1(t) \dots c_n(t))^T$ . As the dynamics of the system may cover several orders of magnitude, we set  $\vec{c}(t) = \vec{q} \exp(\lambda t)$ , where  $\vec{q}$  is independent of time. This approach is equivalent to a Laplace transform of the concentration in the time domain, and has *e.g.* been applied to problems in chemical kinetics by Cramer *et al.*<sup>39</sup> or by Gellrich and co-workers.<sup>40</sup>

In this way, eqn (8) is transformed into a row of a real unsymmetric eigenvalue problem,

$$\frac{D_{i-1,i}}{\Delta z^2} q_{i-1} + \left( -\frac{D_{i,i-1} + D_{i,i+1}}{\Delta z^2} - \lambda \right) q_i + \frac{D_{i+1,i}}{\Delta z^2} q_{i+1} = 0. \quad (9)$$

or

$$(D(\Delta z)^{-2}) - \lambda E \vec{q} = \vec{0}. \quad (10)$$

Using the eigenvalues  $\lambda_a$  and the eigenvectors  $\vec{q}_a$ , the concentrations as a function of time can be computed *via*

$$\vec{c}(t) = \sum_a A_a \vec{q}_a \exp(\lambda_a t), \quad (11)$$

where the  $A_a$  are constructed to meet an initial concentration profile,

$$\vec{c}(t=0) = \sum_a A_a \vec{q}_a. \quad (12)$$

Here, the  $A_a$  are the unknowns in a linear system of equations. The eigenvalue problem has been solved making use of the EISPACK system,<sup>41</sup> the linear system of equations has again been tackled utilizing LINPACK subroutines.<sup>37</sup> Note that for a real general eigenvalue problem, the eigenvectors  $\vec{q}_a$  are not necessarily orthogonal, so  $\langle \vec{c}(t=0) | \vec{q} \rangle \neq A_a$  holds in general.

We construct a three-layer system with the region for which the free energy curve displayed above (Fig. 2) lies in the center.

To the left and to the right, it is extended by an aqueous layer containing 100 grid points each; we have a total of 301 grid points. Initially, the concentration in the leftmost part of the system ( $1 \leq i \leq 100$ ) is set to  $c_0 = 0.01$  in arbitrary units. All eigenvalues are real, and the matrix is negative semidefinite with  $\lambda_1 = 0$  and  $\lambda_a < 0$  for  $a > 1$ . The spectrum of nonzero eigenvalues ranges from  $-6 \times 10^{-4} \text{ ns}^{-1}$  to  $-23 \text{ ns}^{-1}$ .

The results of the semi-analytical integration of the discretized diffusion equation is shown in Fig. 5. Here, the concentration of the molecule – either in its majority protonated or unprotonated form – is shown as a function of time for the aqueous layer, the central (membrane) part and the right aqueous layer, integrated over  $z$  for each of the regions. We denote these concentrations as  $c_l$ ,  $c_c$  and  $c_r$ , respectively. The diffusion process covers a time range from  $\sim 10^{-11}$  to  $\sim 10^{-4} \text{ s}$ , or seven orders of magnitude. Retrospectively, this justifies the comparatively elaborate approach to the time evolution of the concentrations by eqn (9)–(11). For each of the three concentrations, we have typical time scales when  $c$  has been diminished to half of its initial value in the left region, or equals half its equilibrium value in the central or in the right region. Diffusion into the membrane takes slightly longer than  $3 \mu\text{s}$ , and we find a time scale of  $1 \times 10^{-5} \text{ s}$  for the penetration of the membrane. The latter is consistent with the eigenvalue spectrum discussed above, *i.e.* the smallest nonzero eigenvalue,  $(|\lambda|_{\min})^{-1}$ , gives rise to the largest time scale.

As discussed above, the point of the crossover between the protonated and the neutral molecule depends on the pH value. Here, we briefly discuss kinetic effects associated with protonation. Entering the membrane, a protonated cocaine molecule experiences a sufficient number of neighbour water molecules accepting the proton where the protonated and unprotonated free energy curves intersect, and we assume

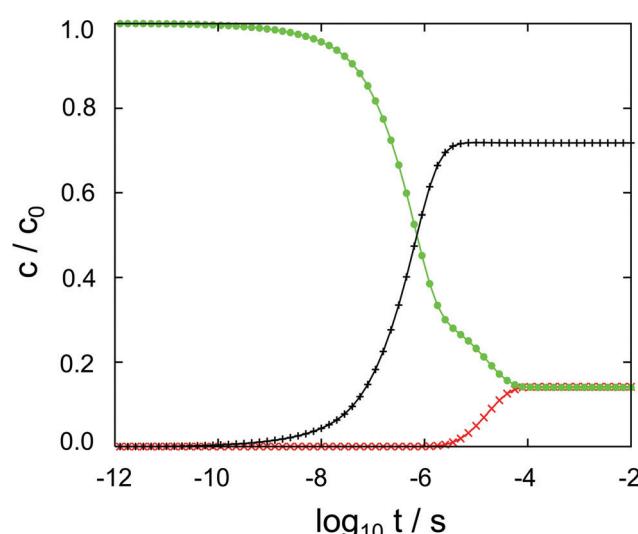


Fig. 5 Cocaine concentrations in the left aqueous phase (● symbol, green), within the membrane (+ symbol, black) and in the right aqueous phase (× symbol, red) as a function of the decadic logarithm of time. Dimensionless concentrations are normalized to unity, the time is given in seconds.



instant deprotonation. Leaving the membrane, however, the neutral molecule has to find a  $\text{H}_3\text{O}^+$  cation in its vicinity as a proton donor, or continue its path on the POMF corresponding to the neutral molecule. As this is higher in free energy than the protonated curve once the aqueous phase is approached, it may slow down transfer from the membrane into the water layer.

In our numerical setup, equilibrium is reached at  $10^{-4}$  s by the latest. Located at the interface between the polar head groups and the hydrophobic tails of the DMPC molecules, the neutral molecule is strongly localized in one of the two minima of the free energy surface. For the geometry used here, we find  $c_e/c_1 = c_e/c_r \simeq 6$  as the liquid–water partition coefficient.

Focussing on the rate-limiting exit of a molecule from the lipid to an aqueous phase with a small concentration and assuming fast, irreversible diffusion away from the membrane within the water phase, rate theory can be used. In the strongly damped, viscous regime that applies here, the exit rate can be estimated as (eqn (4.54) of ref. 42)

$$k = \frac{m\bar{\omega}^2\beta D}{2\pi} \exp(-\beta\Delta G_f), \quad (13)$$

where we make use of variables that can easily be extracted from the simulation results:  $m\bar{\omega}^2 = 0.2 \text{ kcal mol}^{-1} \text{ \AA}^{-2}$  is geometric mean of the curvature of the POMF in its minimum and in the exit transition state, numerically calculated for the smoothed curve,  $D$  the diffusion constant parallel to the membrane, and  $\Delta G_f$  the height of the exit barrier. The obtained escape rate,  $k = 4 \times 10^5 \text{ s}^{-1}$ , comes close to the numerical solution of the kinetics described above, which, however, also includes reversible back diffusion from the aqueous phase into the membrane into account.

In aqueous solution, substrate diffusion is believed to set a rate limit of  $\sim 10^9 \text{ s}^{-1} \text{ mol}^{-1}$  to reactions catalyzed by an enzyme.<sup>43–45</sup> For cocaine diffusion in water, we have computed a diffusion coefficient of the order of  $10^{-10} \text{ m}^{-2} \text{ s}^{-1}$ , a factor of ten smaller for diffusion parallel to the lipid bilayer, and an effective  $D$  that is again roughly four times smaller. Reactions of cocaine with a channel protein within the membrane or a receptor beyond the membrane are likely to be slowed down accordingly, if diffusion-controlled. In our calculations, the only process that is close to the water limit is the entry of cocaine into the membrane, *cf.* Fig. 5. Transfer across the membrane is about five orders of magnitude slower.

## 6. Discussion and conclusions

We have applied classical molecular dynamics simulations to address the question whether the cocaine molecule can permeate a model lipid bilayer. As the time scale of a crossing event exceeds that of typical simulations, we have computed the potential of mean force using the umbrella sampling technique, and addressed transport *via* a numerical solution of Fick's laws and rate theory based upon the free energy data and diffusion coefficients computed in the aqueous phase and in the bilayer. In addition, unconstrained simulations have been

used to obtain a quantitative view of the molecule's orientation within the membrane.

Umbrella sampling simulations have been performed both for the protonated and the neutral form of the molecule, and the free energy curves have been calibrated making use of the cocaine  $\text{p}K_a$  value in the aqueous phase. From our computations, the following scenario emerges. Cocaine in its protonated form enters the hydrophilic part of the lipid bilayer by crossing a barrier of  $2.0 \text{ kcal mol}^{-1}$ . In turn, it loses its proton and is located in a local minimum of the free energy surface. Here, it shows an orientation with the phenyl ring pointing towards the hydrophobic DMPC tails, and the polar part interacting with the heads of the lipid molecules. Moving to the second lipid layer, a barrier of  $3.5 \text{ kcal mol}^{-1}$  has to be overcome, and the orientation of the neutral molecule is reversed. Finally, release into the aqueous phase implies crossing a final barrier of  $4.1 \text{ kcal mol}^{-1}$  and the uptake of a proton. We note that orientational effects while crossing a membrane have been observed both in computer simulations and by NMR spectroscopy<sup>6</sup> for clozapine and amitriptyline and that ROESY NMR provides an excellent experimental method to verify or to falsify the mechanism just outlined. Different penetration depths and orientational effects for the charged and the uncharged form of ibuprofen have also been reported,<sup>5</sup> stemming from a combination of molecular dynamics and neutron diffraction studies. Like cocaine, the anesthetics lidocaine and articaine have a preference for the region close to the water–membrane interface.<sup>46,47</sup>

We acknowledge the complexity of the blood–brain barrier from the scale of cells down to the composition and organization of the endothelial cell membranes.<sup>48</sup> At present, our computer models are limited to simulating what we assume to be the basic constituents of transfer. Taking into account membrane proteins, different lipid compositions or membrane asymmetry is likely to be in reach soon with increasing computer performance and progress in simulation algorithms.

The amphiphilic nature of the neutral cocaine molecule and its ability to further adapt to hydrophobic or hydrophilic environments by changing its protonation state leads to barriers of a few  $k_{\text{B}}T$  for crossing a membrane, which is also manifested in a time scale of less than a millisecond. The process can be driven by a pH gradient, which lowers the rate-limiting exit barrier. As a consequence, a pH-dependent transport rate does not necessarily point at transfer mediated by a proton antiporter. It is interesting to note that the blockade of sodium channels – as relevant for cocaine as an anesthetic – also depends on the pH.<sup>49</sup>

As pointed out by Martinotti *et al.*,<sup>12</sup> small molecules are ideal candidates for membrane permeation and fulfill one of Veber's empirical rules:<sup>50</sup> the fewer rotatable bonds in a molecule, the higher its bioavailability. With two stereocenters, cocaine has 16 diastereomers, of which eight do not exhibit a serious sterical overlap or deformation. These stereoisomers exhibit a remarkable spectrum of dipole moments, ranging from 0.65 to 4.60 Debye.<sup>22</sup> One may speculate whether these molecules show a different membrane permeation behaviour,



and whether the effect of the naturally occurring *R*-cocaine may not only depend on the drug–receptor interaction, but also on the kinetics of transport. We are confident that future work will incorporate polarizable force fields<sup>51,52</sup> and will benefit from recent developments in nonequilibrium methods suited to simultaneously compute the POMF and the friction while pulling a molecule.<sup>53,54</sup>

## Conflicts of interest

There are no conflicts to declare.

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