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Enhanced 1.54 μm luminescence of a perfluorinated erbium complex sensitized by perfluorinated Pt(II) and Zn(II) phthalocyanines with 980 nm emission†

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By the sensitization effect of metallophthalocyanines showing ~980 nm emission to an erbium complex, a remarkably long average lifetime of 1.05 ms and an optimal PLQY of 13% with a sensitization efficiency of 81 for the Er³⁺ 1.54 μm emission are obtained in a perfluorinated organic erbium co-doped system.

Introduction

The erbium-doped optical fiber amplifier (EDFA) was reported in 1987 as an important invention for global fiber-optic telecommunication.¹ Since then, research on Er³⁺ doped waveguide amplifiers (EDWA) operating at 1.5 μm has attracted considerable attention^{2–8} because multiple optical elements can be integrated on compact photonic chips to achieve optical amplification. Er³⁺ ions have weak optical absorption coefficients originating from 4f–4f electronic transitions forbidden by the Laporte rule.³ Combining Er³⁺ ions with organic ligands can significantly enhance the near infrared (NIR) emission because the indirect excitation of Er³⁺ ions can occur by energy transfer (ET) from organic ligands *via* an antenna effect or sensitization effect.^{9–13} However, chemical bonds containing H atoms such as C–H and O–H in most organic components can cause severe vibrational quenching of the 1.54 μm emission,

resulting in a weak NIR emission and short lifetimes with tens of microseconds and consequently preventing Er³⁺ ions from attaining population inversion.^{14–17} One effective strategy to eliminate C–H bonds is to fully fluorinate organic compounds to reduce the vibrational quenching.^{18–22} The sensitized Er³⁺ energy levels of the fully fluorinated Er³⁺ complexes centre on the ²H_{11/2}, ⁴F_{7/2}, and ⁴S_{3/2} energy levels, corresponding to photons of green light.^{8,22,23} Such a sensitization of tris[tetrakis(pentafluorophenyl)imidodiphosphinate] Er³⁺ (Er(F-TPIP)₃)¹⁸ occurs in the green region regardless of whether a 266 nm laser is used to excite Er(F-TPIP)₃ with sensitizing of Er³⁺ ions or the perfluorinated zinc complex Zn(F-BTZ)₂ or a six-coordinate iridium complex (Ir-tBuPBI) excited at 405 nm to sensitize the Er³⁺ ions in Er(F-TPIP)₃.^{8,23} There are large energy differences between the sensitized energy levels (~2.38 eV) and the ⁴I_{13/2} energy level (~0.81 eV) in the NIR region, and consequently, part of the energy is lost in the form of heat during the process of relaxation preceding NIR emission, leading to a decreased sensitization efficiency. Finding a sensitizer with luminescence close to the ⁴I_{13/2} state can increase the sensitization efficiency because of the corresponding shift to a longer wavelength, and for a given photon flux, lower exciton energy means lower exciton power density (which is equal to the product of photon-flux and photon energy). With one exciton sensitizing one Er ion numerically, this approach can reduce the power needed to induce population inversion.

Phthalocyanines (Pcs) have highly delocalized 18 π -electron conjugated rings leading to extreme stability,^{24,25} bright photoluminescence (PL),²⁶ and electroluminescence (EL) properties in the NIR region.^{27,28} Metallophthalocyanines (MPcs) have contributed greatly in many fields, for example, organic dye-sensitized solar cells²⁹ and as photosensitizers in photodynamic therapy.^{30,31} However, no available perfluorinated matrix material that can avoid self-quenching of ErPc is known. In this article, we demonstrate enhanced 1.54 μm luminescence of Er³⁺ in Er(F-TPIP)₃ by using M(II) hexadecafluorophthalocyanines (F₁₆MPc; M = Zn, Pt) as sensitizers. Both these compounds

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possess luminescence at ~ 980 nm: $F_{16}PtPc$ exhibits dominant phosphorescence, and the wavelength is consistent with the 980 nm excitation used for EDFA, providing high pump rate efficiency. $Er(F-TPIP)_3$ and $F_{16}MPc$ are co-doped at the molecular level using a co-evaporation technique. Photoexcitation gives optimized sensitization efficiencies of 36 and 81 for the Er^{3+} 1.5 μm emission with $F_{16}ZnPc$ and $F_{16}PtPc$, respectively. The average lifetimes for the Er^{3+} 1.5 μm emission are $\sim 0.89 \pm 0.011$ ms (doped with $F_{16}ZnPc$) and $\sim 1.04 \pm 0.012$ ms (doped with $F_{16}PtPc$). This is so far the longest lifetime of organic sensitized Er^{3+} 1.5 μm emission, which can even match those of the inorganic systems.³² A maximum photoluminescence quantum yield (PLQY) of 13% is obtained for the sensitized Er^{3+} 1.5 μm emission in an organic co-doped system.

Results and discussion

The chemical structures of $F_{16}MPc$ and $Er(F-TPIP)_3$ are depicted in Fig. 1. $F_{16}PtPc$ was synthesized and purified by modifying a published procedure for the preparation of $F_{16}CuPc$.³³ $F_{16}PtPc$ of good purity was obtained following a simple reprecipitation from 98% sulfuric acid by dilution with water. This approach avoids the need for preparative chromatography, which can be difficult to scale up for Pc derivatives due to their generally low solubility in organic solvents. The purity and identity of $F_{16}PtPc$ were confirmed by MALDI-TOF mass spectrometry, as shown in Fig. S1 (ESI[†]). The most intense peak (m/z 996) corresponded to the protonation of $F_{16}PtPc$ by the MALDI matrix and aromatic carboxylic acids to form an $[F_{16}PtPc + H]^+$ pseudo molecular ion containing the abundant ^{195}Pt nuclide. This was part of an isotopic cluster, as expected, based on the natural occurrence of other stable isotopes. There were no detectable impurities, apart from a group of very weak peaks, 16 atomic mass units heavier than the main cluster; these had an intensity $< 1\%$ of the major component and were attributed to halogen exchange with platinum(II) chloride having occurred during the Pc ring synthesis, forming traces of $ClF_{15}PtPc$. The ^{19}F NMR spectrum of $F_{16}PtPc$ in D_2SO_4 solution showed just two aromatic fluorine environments: these had equal integrals, which is consistent with the expected four-fold symmetry of the highly delocalised phthalocyanine system. The energy dispersive X-ray (EDS) spectrum shown in Fig. S2 (ESI[†]) confirmed the presence of C, N, F, and Pt; Cl is too weak to be detected. The successful synthesis of $F_{16}PtPc$ was also supported by Fourier transform infrared (FTIR) spectroscopy. As shown in Fig. S3 (ESI[†]), the synthesized $F_{16}PtPc$ has seven groups of peaks similar to the IR spectrum of $F_{16}CuPc$,³⁴ consistent with the successful formation and metalation of the $F_{16}Pc$ macrocycle.

Since the $F_{16}MPc$ s have a large coplanar conjugated system of 18 π electrons, they could form molecular aggregates and quench the luminescence, which would disfavor energy transfer from $F_{16}MPc$ to $Er(F-TPIP)_3$.³⁵ Absorption spectra provide a method to study the aggregation. As shown in Fig. 2a, the electronic absorption spectra of $F_{16}ZnPc$ in acetone solution and in a film deposited onto a glass substrate show two distinct

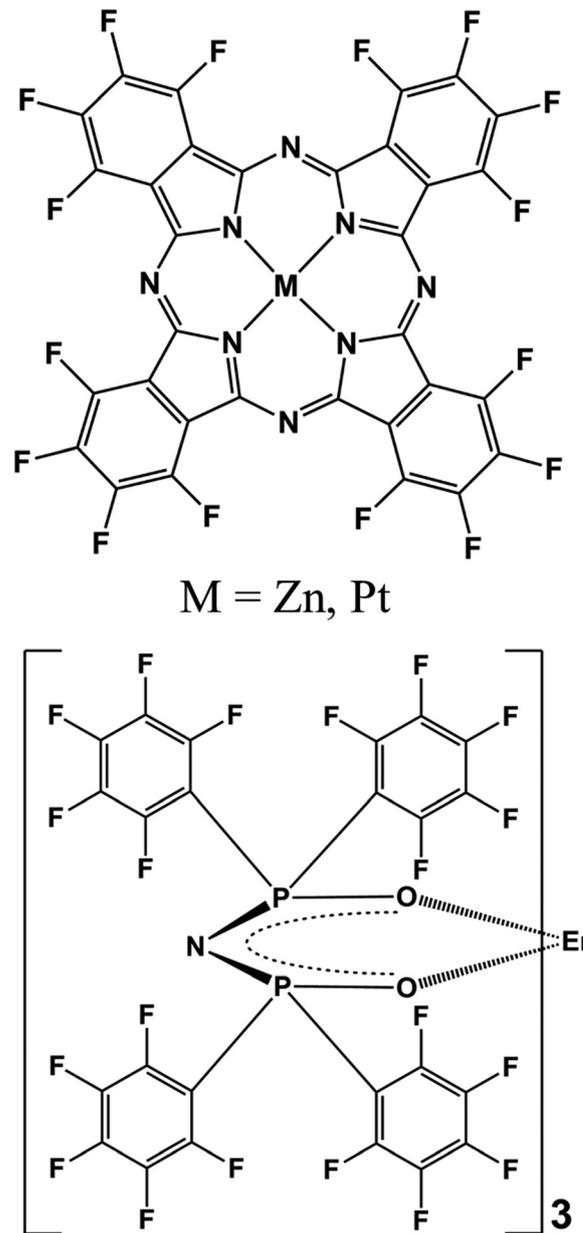


Fig. 1 The chemical structures of $F_{16}MPc$ and $Er(F-TPIP)_3$.

absorption bands from the UV to the NIR spectral region. The high-energy absorption band in the region of < 450 nm is ascribed to the B-band originating from a direct electronic transition from a_{2u} orbital to e_g orbital. The low energy absorption band in the region from 500 to 850 nm is ascribed as a Q-band originating from an a_{1u} to e_g transition. The Q band of the $F_{16}ZnPc$ solution in acetone shows two absorption bands at 635 and 667 nm with a shoulder at 580 nm that indicates the dimer behavior of $F_{16}ZnPc$ in solution.³⁴ In comparison, it is found that the spectrum of a neat $F_{16}ZnPc$ film shows both significant broadening and a bathochromic shift; these are typical effects of aggregation behavior.³⁶ However, the absorption spectrum of the Q band of $F_{16}ZnPc$ in a 50% $Er(F-TPIP)_3$:50% $F_{16}ZnPc$ co-doped film appears as one band at 624 nm with a weak shoulder



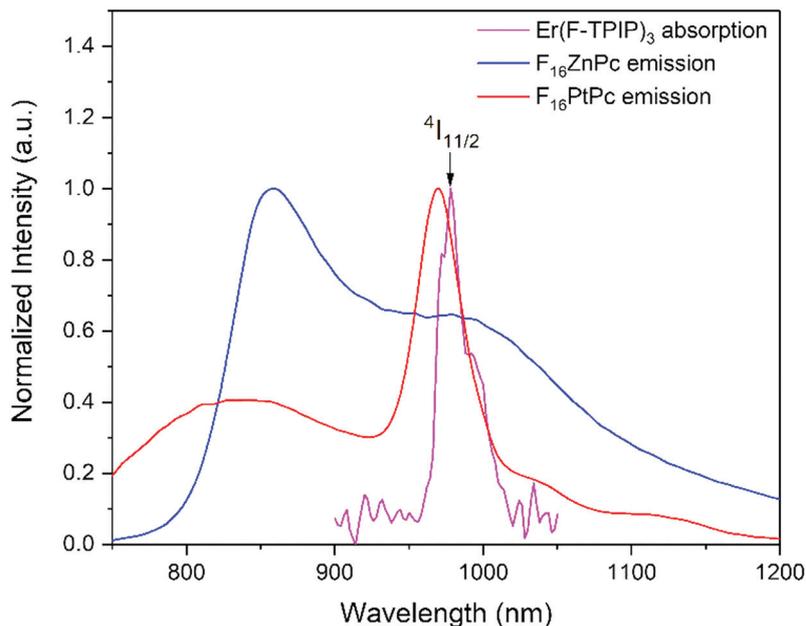


Fig. 2 The absorption spectra of (a) $F_{16}ZnPc$ and (b) $F_{16}PtPc$. (A) Dissolved in acetone. (B) Neat film. (C) Co-doped film with 50% $Er(F-TPIP)_3$.

at 675 nm. Compared to that of neat $F_{16}ZnPc$ film, this absorption band is dramatically narrowed and its full width at half maximum (FWHM) is only slightly wider than that in solution. The comparison indicates that the aggregation of the $F_{16}ZnPc$ molecules in the co-doped film is effectively suppressed, and they are likely to be dispersed homogeneously.

The Q band of $F_{16}PtPc$ in the solution, neat and co-doped films are similar, and the FWHM of the Q band of the 50% $Er(F-TPIP)_3$:50% $F_{16}PtPc$ co-doped film is even slightly narrower than that in solution (Fig. 2b). This contrasts markedly with $F_{16}ZnPc$ (Fig. 2a) considered above, where the Q band of $F_{16}ZnPc$ in the neat film is significantly different from those in the solution and the co-doped film. Especially, $F_{16}PtPc$ in the co-doped film shows

one Q band, which indicates that the $F_{16}PtPc$ molecules in the co-doped film exist homogeneously as monomers.

The emission spectra of neat $F_{16}ZnPc$ and $F_{16}PtPc$ films are shown in Fig. 3. Zn^{2+} has a relatively low atomic number and a d^{10} closed shell structure without any unpaired electrons, leading to a low probability of intersystem crossing (ISC).³⁷ Its very weak phosphorescence possibly comes from the $\pi \rightarrow \pi^*$ transition of the $F_{16}Pc$ ligands. Because no long-lived emission is detected at room temperature, we can attribute the luminescence of $F_{16}ZnPc$ at about 980 nm to fluorescence. On the other hand, non-fluorinated PtPc exhibits phosphorescence in the NIR region, and the perfluorination of PtPc can further increase the ISC, so $F_{16}PtPc$ is regarded as a phosphorescent material.^{38–40}

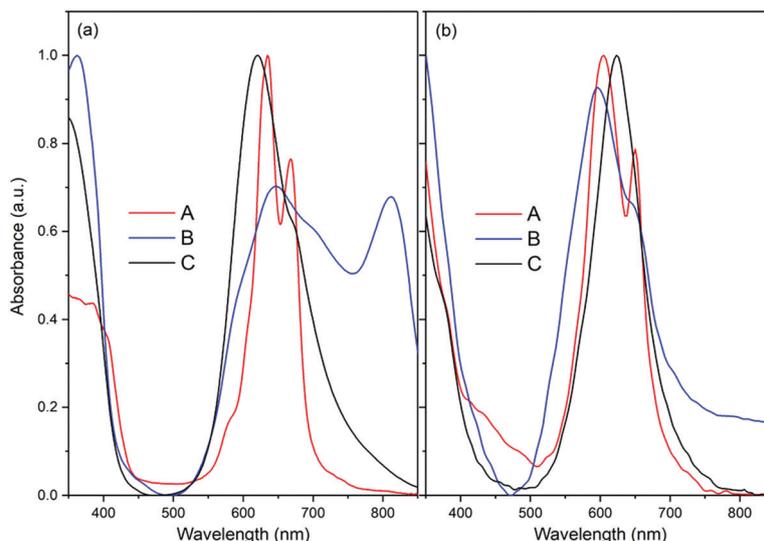


Fig. 3 Absorption spectrum of ${}^4I_{15/2} \rightarrow {}^4I_{11/2}$ of Er^{3+} ion in an $Er(F-TPIP)_3$ crystal; emission spectra of neat $F_{16}ZnPc$ and $F_{16}PtPc$ films.



Table 1 Calculated emission energies (eV), oscillator strength (*f*) and radiative lifetime τ (s) by SOC-TDDFT at PBE/TZ2P level of theory

State	ΔE [eV]	λ [nm]	<i>f</i>	Transition contribution	Assignment	T [s]
$^3A_{1u}$	1.3173	941.3	0.000	$61e_{1/2g} \rightarrow 61e_{1/2u}$ (0.9986)	$^3LC/^3MLCT$	—
$^3A_{2u}$	1.3173	941.3	0.2438×10^{-7}	$61e_{1/2g} \rightarrow 61e_{1/2u}$ (0.9986)	$^3LC/^3MLCT$	0.5447
3E_u	1.3346	929.1	0.1015×10^{-3}	$61e_{1/2g} \rightarrow 61e_{1/2u}$ (0.5199)	$^3LC/^3MLCT$	1.275×10^{-4}
3E_u	1.3346	929.1	0.1015×10^{-3}	$58e_{3/2g} \rightarrow 61e_{1/2u}$ (0.4789)	$^3LC/^3MLCT$	1.275×10^{-4}
$^3B_{1u}$	1.3519	917.2	0.000	$61e_{1/2g} \rightarrow 61e_{1/2u}$ (0.9987)	$^3LC/^3MLCT$	—
$^3B_{2u}$	1.3519	917.2	0.000	$58e_{3/2g} \rightarrow 61e_{1/2u}$ (0.9987)	$^3LC/^3MLCT$	—

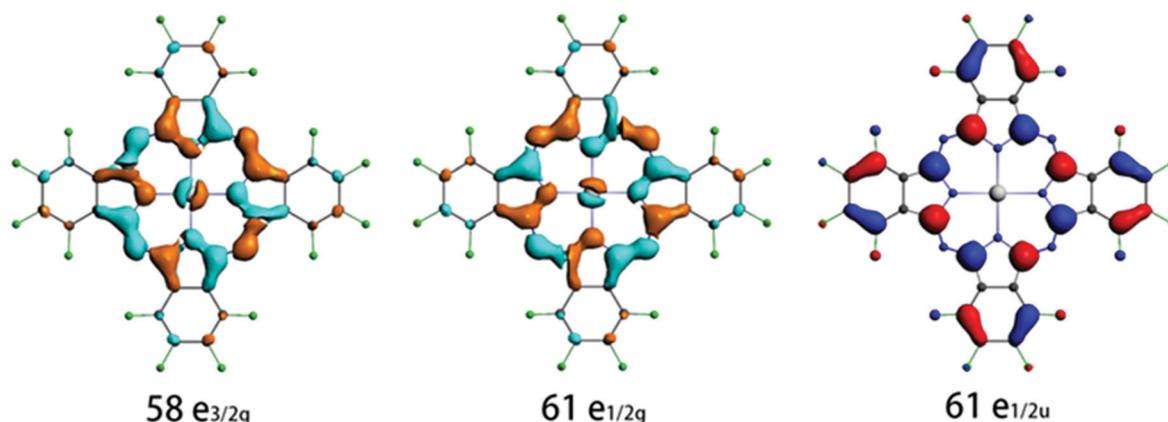
To gain insight into the origin of the intense emission peak at *ca.* 980 nm of the neat F₁₆PtPc film, we have performed SOC-TDDFT calculations based on the T₁ excited state structure. The calculated emission energies in terms of wavelength and dominant orbital transitions for the low-lying triplet excited states are collected in Table 1, and the important frontier molecular orbitals are shown in Fig. 4. The calculated emission wavelength in the Franck–Condon state (T₁ equilibrium geometry) for F₁₆PtPc is 940 nm, which relates well to the experimental phosphorescence peak of 980 nm. As shown in Table 1, the lowest energy transitions for these emission bands are LUMO → HOMO transitions. From Fig. 4, the LUMO (61e_{1/2g}) is largely a π orbital of F₁₆Pc mixed with the d π of Pt, and the HOMO (61e_{1/2u}) is largely a π^* orbital on the F₁₆Pc. We attribute the origin of the 980 nm emission in the experiment to mixed metal-to-ligand charge transfer (3MLCT) and ligand-centered (3LC) characteristics. The calculated long excited-state lifetime of 544.7 μ s confirms that F₁₆PtPc emits phosphorescence.

The absorption spectrum of the Er³⁺ ion in the Er(F-TPIP)₃ crystal in the NIR region is shown in Fig. 3. Both the broad emission bands of F₁₆ZnPc and especially F₁₆PtPc have good overlap of the absorption band of $^4I_{15/2} \rightarrow ^4I_{11/2}$ of the Er³⁺ ion in Er(F-TPIP)₃, which means that an efficient sensitization may be realized through energy transfer from F₁₆MPC to Er(F-TPIP)₃.

In order to study the sensitization effect of F₁₆ZnPc on Er(F-TPIP)₃, a series of Er(F-TPIP)₃:F₁₆ZnPc co-doped films with the concentrations of Er(F-TPIP)₃ varying from 20% to 80% were prepared; their emission spectra are shown in Fig. 5a. It is found that the emission intensity of F₁₆ZnPc at \sim 980 nm tends to decrease with increasing Er(F-TPIP)₃ concentration, implying that energy is transferred to Er(F-TPIP)₃ quenching the excitation

of F₁₆ZnPc. The emission spectra of the Er(F-TPIP)₃:F₁₆PtPc co-doped films under the excitation of a 655 nm laser are shown in Fig. 5b. It is found that the centre of the emission spectrum is shifted from 970 nm in the neat F₁₆PtPc film to \sim 980 nm in the co-doped film. Being different from the Er(F-TPIP)₃:F₁₆ZnPc co-doped films, the dips in the mixed $^3LC/^3MLCT$ emission spectra of F₁₆PtPc in co-doped films appeared at \sim 980 nm, which was not found in the spectrum of the neat F₁₆PtPc film as shown in Fig. 3. Their shapes resemble those of the inverted absorption bands of Er(F-TPIP)₃ at \sim 980 nm, as shown in Fig. 5b. This is a clear evidence that there is an energy transfer mechanism based on the emission–reabsorption process where the Er³⁺ $^4I_{11/2}$ state absorbs part of the F₁₆PtPc emission.^{41,42} This mechanism is different from the Förster energy transfer, wherein the whole emission intensity of F₁₆PtPc should be decreased with no change in spectral features.

It is interesting that there is no dip in the emission spectrum of F₁₆PtPc in the 20% Er(F-TPIP)₃:80% F₁₆PtPc co-doped film. This means that the re-absorption process does not occur at this concentration, and there may be another energy transfer path from F₁₆PtPc to Er(F-TPIP)₃. Förster transfer is a resonant dipole coupling process that is dependent on the overlap between the donor emission spectrum and the acceptor absorption spectrum. Here the overlap between the PL spectrum of F₁₆PtPc and the absorption of Er(F-TPIP)₃ is obvious, as shown in Fig. 3. It indicates that the Förster energy transfer may also be an efficient mechanism and suggests that the energy transfer process from F₁₆PtPc to Er(F-TPIP)₃ comprises Förster transfer and emission–reabsorption. The energy transfer paths are shown in Fig. 6. The 20% Er(F-TPIP)₃:80% F₁₆PtPc co-doped film shows

**Fig. 4** Calculated electronic density contours of the frontier molecular orbitals involved in the main electronic transitions.

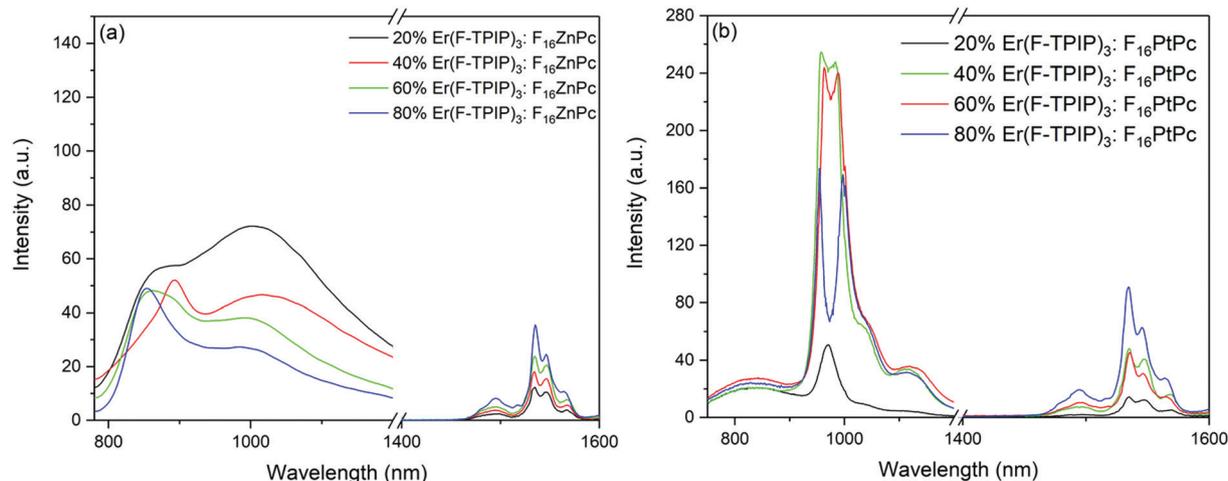


Fig. 5 Emission spectra of (a) Er(F-TPIP)₃:F₁₆ZnPc and (b) Er(F-TPIP)₃:F₁₆PtPc co-doped films with different Er(F-TPIP)₃ concentrations. $\lambda_{\text{ex}} = 655$ nm.

the weakest 1.54 μm emission from Er(F-TPIP)₃ due to the low concentration of Er(F-TPIP)₃. Meanwhile, given the weak emission from F₁₆PtPc, there could be a triplet-triplet annihilation (TTA) due to highly concentrated 80% F₁₆PtPc, which reduces the energy transfer. As the concentration of Er(F-TPIP)₃ is increased, the 1.54 μm emission of Er(F-TPIP)₃ increases gradually while the emission of F₁₆PtPc decreases. The strongest 1.54 μm emission is obtained based on 80% Er(F-TPIP)₃ doped concentration, due to the increase in the Er(F-TPIP)₃ molecules and possibly a decrease in the TTA effect in F₁₆PtPc.

To quantify the sensitization efficiency, the Er(F-TPIP)₃:Y(F-TPIP)₃ co-doped films with different diluted Er³⁺ concentrations were used as reference samples, where Y³⁺ is optically inactive, resulting in the exclusion of energy transfer from Y(F-TPIP)₃ to Er(F-TPIP)₃. The measurement set up is identical for all co-doped films. The power-dependent Er³⁺ PL intensities of these co-doped films with Er³⁺ concentrations of 20%, 40%, 60%, and 80% recorded at 1532 nm are shown in Fig. 7, where the Er³⁺ ions are directly excited to the ⁴F_{9/2} level. In the co-doped films containing the sensitizers, 655 nm photoexcitation of F₁₆ZnPc and F₁₆PtPc was again used to induce sensitization. Hence, the 1532 nm emission intensities from both the F₁₆ZnPc and F₁₆PtPc co-doped films are larger than those of the

Y(F-TPIP)₃ doped films. The sensitization effect can be simply calculated by fitting the emission intensity of the Y(F-TPIP)₃ doped films *versus* power density, which shows a linear correlation. Subsequently, the power densities used in F₁₆MPC can be used in the correlation to calculate the corresponding 1532 nm emission intensity, and these values can be used to divide the measured 1532 nm intensity to get the sensitization efficiency. The values of the sensitization efficiency for F₁₆ZnPc/F₁₆PtPc in different Er(F-TPIP)₃ doped films are listed in Table 2. The sensitization efficiency decreased with the increased percentage of Er(F-TPIP)₃ in the F₁₆ZnPc doped films. F₁₆PtPc behaves differently, the reason being that the TTA effect of F₁₆PtPc is quite strong. However, under a high concentration of Er(F-TPIP)₃, the F₁₆PtPc molecules get separated and the TTA effect is dramatically decreased, and the emission from F₁₆PtPc gets stronger so that more energy can be transferred to the Er³⁺ ions through the re-absorption effect. In addition, long-lived triplet excitons of F₁₆PtPc travel over a long distance to sensitize Er³⁺ ions, along with the perfect overlap between the absorption band of Er(F-TPIP)₃ and the emission peak of F₁₆PtPc. So, F₁₆PtPc shows a more efficient sensitization effect than F₁₆ZnPc with high Er(F-TPIP)₃ doped concentration.

The time-resolved PL (TRPL) decay curves for the F₁₆PtPc emission recorded at 980 nm with different Er(F-TPIP)₃ doped concentration films are shown in Fig. 8a. The values of their short component (τ_s), long component (τ_L), and average lifetimes (τ_{Ave}) are listed in Table 3a. The lifetime of the F₁₆ZnPc is too short to be measured, lying below the detection limit of the equipment (~ 5 ns) due to the strong self-quenching. In F₁₆PtPc doped films, the changing concentrations affect the τ_L of F₁₆PtPc through two opposing processes: (I) with increased Er(F-TPIP)₃ concentrations, F₁₆PtPc is diluted, and the reduced TTA effect leads to increased τ_L . (II) At the same time, each F₁₆PtPc molecule is surrounded by more Er(F-TPIP)₃ molecules; their excitons can more easily transfer their energy to the central Er³⁺ ions in Er(F-TPIP)₃, thus tending to reduce τ_L . As the Er(F-TPIP)₃ concentration increases from 20% to 60%, τ_L gradually increases from 1.00 ± 0.015 ms to 1.89 ± 0.017 ms

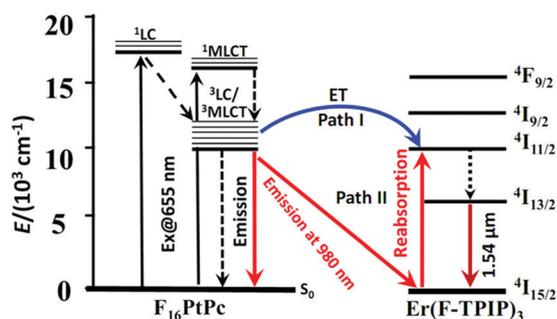


Fig. 6 Schematic energy diagram describing the two sensitization mechanisms of NIR luminescence of Er(F-TPIP)₃ via F₁₆PtPc energy transfer (path I) and emission-reabsorption (path II).



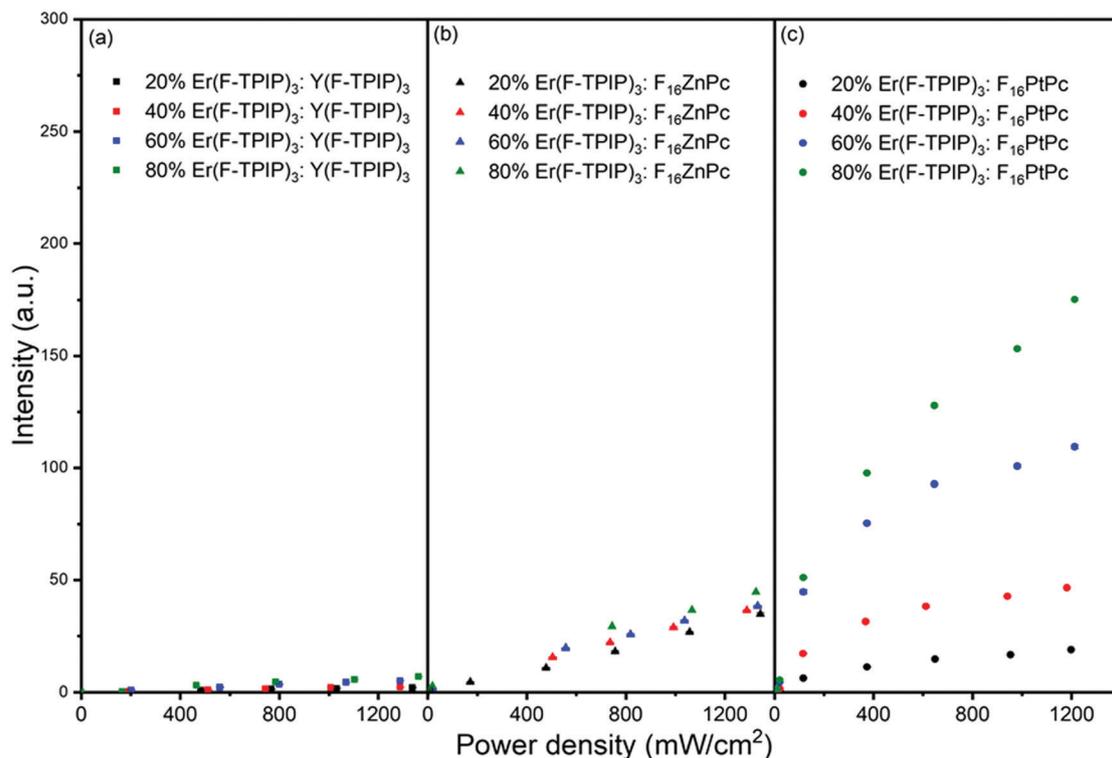


Fig. 7 The emission intensity at 1532 nm as a function of laser power density for co-doped films with different Er(F-TPIP)₃ concentrations diluted by (a) Y(F-TPIP)₃, (b) F₁₆ZnPc, and (c) F₁₆PtPc, respectively. $\lambda_{\text{ex}} = 655$ nm.

Table 2 The sensitization efficiency of F₁₆ZnPc and F₁₆PtPc in co-doped films with different concentrations of Er(F-TPIP)₃

Sample	Concentration [%]	Sensitization efficiency
Er(F-TPIP) ₃ :F ₁₆ ZnPc	20	36 ± 1.6
Er(F-TPIP) ₃ :F ₁₆ ZnPc	40	32 ± 1.3
Er(F-TPIP) ₃ :F ₁₆ ZnPc	60	30 ± 1.2
Er(F-TPIP) ₃ :F ₁₆ ZnPc	80	28 ± 1.2
Er(F-TPIP) ₃ :F ₁₆ PtPc	20	22 ± 1.1
Er(F-TPIP) ₃ :F ₁₆ PtPc	40	53 ± 2.3
Er(F-TPIP) ₃ :F ₁₆ PtPc	60	70 ± 3.2
Er(F-TPIP) ₃ :F ₁₆ PtPc	80	81 ± 3.9

because factor (I) plays the dominant role. However, in the 80% Er(F-TPIP)₃ doped film, τ_L decreases to 1.30 ± 0.014 ms as factor (II) is predominant. Basically, the energy transfer from F₁₆PtPc to Er(F-TPIP)₃ is more efficient with an increase in Er(F-TPIP)₃ because more energy is transferred to Er(F-TPIP)₃, which causes the lifetime of F₁₆PtPc to decrease from 484.64 ± 7.09 μ s to 384.48 ± 6.39 when the Er(F-TPIP)₃ concentration is increased from 20% to 80%.

The results of TRPL measurements for the Er³⁺ emission lifetimes ($\lambda_{\text{mon}} = 1532$ nm) with different concentrations of Er(F-TPIP)₃ films doped with F₁₆ZnPc and F₁₆PtPc are shown in Fig. 8b, c and Table 3b and c, respectively. The decay can be described by a biexponential process, with two lifetime components for Er³⁺ in each co-doped film. Being sensitized by F₁₆ZnPc and F₁₆PtPc, the rise in the Er(F-TPIP)₃ concentrations makes their τ_S and τ_L to increase, which shows the same trend of the emission intensity of Er(F-TPIP)₃ as shown in Fig. 5, except

for the 80% Er(F-TPIP)₃:F₁₆PtPc co-doped film. Compared to the 60% Er(F-TPIP)₃ doped film, the τ_L of the 80% Er(F-TPIP)₃ doped film is slightly decreased. With the increase in Er(F-TPIP)₃ doped concentration, although their τ_S and τ_L are increased, the erbium ion-ion interaction is also increased. This leads the percentage contribution of τ_L to be reduced, while τ_S is increased, which makes τ_{Ave} almost unaffected by the doped concentrations of sensitizers. τ_S and τ_{Ave} of Er³⁺ in the F₁₆PtPc co-doped films are longer than those of the F₁₆ZnPc co-doped films because of the long phosphorescence lifetimes of F₁₆PtPc and the emission-reabsorption process from F₁₆PtPc to Er(F-TPIP)₃ in doped films. The longest average lifetime of 0.91 ± 0.015 and 1.05 ± 0.041 ms for F₁₆ZnPc and F₁₆PtPc co-doped films are obtained, and they are 4.55 and 5.25 times longer than the neat Er(F-TPIP)₃ film (0.2 ms), respectively. This prolonged Er³⁺ lifetime is caused by the long-lived organic triplet excitons;⁴³ so the largest PLQY of 13% is achieved in the F₁₆PtPc co-doped film, which is the highest ever reported for erbium in an organic complex.

Conclusions

We demonstrate that two fully fluorinated MPCs: F₁₆ZnPc and F₁₆PtPc can be used as sensitizers to enhance the Er 1.54 μ m emission in the co-doped organic Er³⁺ complex systems. This significantly reduces the energy difference between the sensitized Er³⁺ energy levels and the ⁴I_{13/2} level by utilizing perfluorinated 980 nm luminescent chromophores and consequently decreases the energy loss caused by the Er³⁺ nonradiative transitions from the



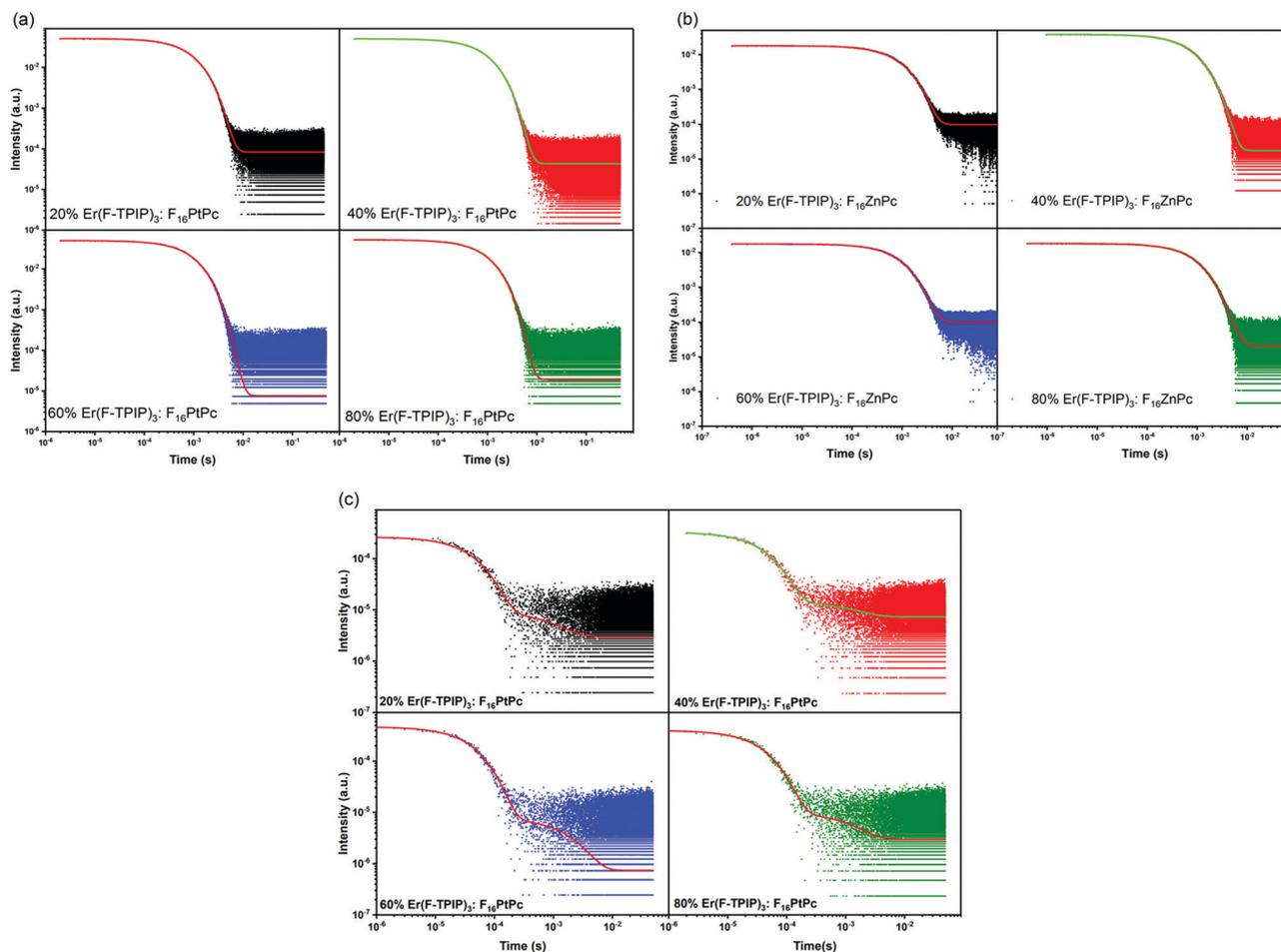


Fig. 8 (a) The TRPL decay of the $F_{16}PtPc$ doped films with $Er(F-TPIP)_3$ concentration varying from 20% to 80%. $\lambda_{mon} = 980$ nm. The TRPL decay of $Er(F-TPIP)_3$ in (b) $F_{16}ZnPc$ doped films and (c) $F_{16}PtPc$ doped films. $\lambda_{mon} = 1532$ nm.

higher sensitized energy levels to the $^4I_{13/2}$ level. Optimized sensitization efficiencies of 36 and 81 are obtained for the Er^{3+} complex doped with $F_{16}ZnPc$ and $F_{16}PtPc$, respectively. An average lifetime for the Er^{3+} 1.5 μm emission in an $Er(F-TPIP)_3$ - $F_{16}PtPc$ system reaches ~ 1 ms, which is by far the longest ever reported. 13% PLQY is achieved in the $F_{16}PtPc$ doped film that is the highest ever reported for erbium in an organic system. Our research opens the door for improving power efficiency to reach population inversion of Er ions based on the perfluorinated organic system.

Experimental

Chemical reagents

$F_{16}ZnPc$ (90%) was purchased from Sigma Aldrich and purified by sublimation under 280 $^{\circ}C$ at a pressure of $\sim 5 \times 10^{-6}$ mbar. 3,4,5,6-Tetrafluorophthalo-1,2-dinitrile and anhydrous platinum(II) chloride were obtained from Energy Chemical and Mascal companies, respectively.

Synthesis of $F_{16}PtPc$

3,4,5,6-Tetrafluorophthalo-1,2-dinitrile (0.40 g, 2 mmol) and anhydrous platinum(II) chloride (0.13 g, 0.5 mmol) were mixed

well and placed in a glass vessel. The vessel was sealed under vacuum (0.13 Pa) and then slowly heated to 210 $^{\circ}C$ for 6 hours. After cooling to room temperature, the crude purple product was isolated and washed with petroleum ether and ethanol. The crude $F_{16}PtPc$ product was dissolved in concentrated sulfuric acid (~ 2 ml) with intense stirring in an ice-water bath for 1 hour at a solution temperature not exceeding 5 $^{\circ}C$. Then it was filtered by a sintered glass filter, and the filtrate was slowly added to the ice water mixture under stirring and at a temperature lower than 10 $^{\circ}C$. Finally, the suspension was filtered and the cake was washed to neutral with deionized water. After drying, the $F_{16}PtPc$ product (0.21 g, 43%) was obtained as a purple solid. ^{19}F NMR (377 MHz, D_2SO_4 solution) δ_F : -129.6 (8F, br s) and -137.2 (8F, br s); IR (KBr, cm^{-1}): C-F (1496 s), C=C (1530 s), C=N (1647 w), Pc ring (750 w, 1160 s and 1469 s), N-Pt-N (845 w). UV/Vis: $\lambda_{max} = 605$ nm. MS (MALDI-TOF): m/z calcd for $[M + H]^+ [C_{32}HF_{16}N_8^{195}Pt]^+$, 996.0; found: 996.1.

Theoretical calculations

All calculations were performed with the ADF 2013.01 program package.⁴⁴ The geometry optimization of the ground state (S_0) and lowest triplet excited state (T_1) was carried out using gradient



Table 3 (a) The lifetime components and average lifetime of $F_{16}PtPc$ emission ($\lambda_{mon} = 980$ nm) in the $Er(F-TPIP)_3:F_{16}PtPc$ co-doped films with different $F_{16}PtPc$ concentrations. (b) The Er^{3+} 1.5 μm emission lifetime components and average lifetimes in the $Er(F-TPIP)_3:F_{16}ZnPc$ co-doped films with different $F_{16}ZnPc$ concentrations. (c) The Er^{3+} 1.5 μm emission lifetime components and average lifetimes in the $Er(F-TPIP)_3:F_{16}PtPc$ co-doped films with different $Er(F-TPIP)_3$ concentrations

(a)			
Concentration [%]	τ_s [μs]	τ_L [ms]	τ_{Ave} [μs]
20	43.50 \pm 0.047 (64%)	1.30 \pm 0.014 (36%)	484.64 \pm 7.09
40	45.75 \pm 0.039 (61%)	1.89 \pm 0.017 (39%)	754.29 \pm 7.00
60	43.44 \pm 0.068 (52%)	1.45 \pm 0.010 (48%)	708.62 \pm 8.56
80	44.89 \pm 0.076 (64%)	1.00 \pm 0.015 (36%)	384.48 \pm 6.39
(b)			
$F_{16}ZnPc$ concentrations [%]	Short component τ_s [ms]	Long component τ_L [ms]	Average lifetime τ_{Ave} [ms]
80	0.67 \pm 0.0025 (49%)	1.08 \pm 0.0047 (51%)	0.88 \pm 0.015
60	0.68 \pm 0.0013 (62%)	1.22 \pm 0.0048 (38%)	0.89 \pm 0.003
40	0.76 \pm 0.0009 (81%)	1.45 \pm 0.0116 (19%)	0.89 \pm 0.014
20	0.77 \pm 0.0006 (86%)	1.72 \pm 0.0213 (14%)	0.91 \pm 0.015
(c)			
$Er(F-TPIP)_3$ concentrations [%]	Short component τ_s [ms]	Long component τ_L [ms]	Average lifetime τ_{Ave} [ms]
20	0.85 \pm 0.0025 (48%)	1.22 \pm 0.0012 (52%)	1.04 \pm 0.025
40	0.90 \pm 0.0013 (65%)	1.30 \pm 0.0014 (35%)	1.04 \pm 0.034
60	0.95 \pm 0.0009 (82%)	1.50 \pm 0.0016 (18%)	1.05 \pm 0.041
80	0.98 \pm 0.0006 (91%)	1.46 \pm 0.0010 (9%)	1.02 \pm 0.047

approximation (GGA) using the PBE exchange–correlation functional.⁴⁵ All electron TZ2P basis set was employed for all atoms, and the relativistic effect was also taken into account *via* the zeroth-order regular approximation (ZORA).^{46,47} On the basis of the optimized T_1 geometry, relativistic calculations were performed to obtain electron excitations and radiative rate. In these calculations, spin–orbit coupling (SOC) was included in one-component time dependent density functional theory (TDDFT) utilizing the ZORA.^{48,49} The radiative rate (k_r) of an excited state was then calculated as⁵⁰

$$k_r = 2\varepsilon^2 f f_c^3$$

where ε is the excitation energy and f is the corresponding oscillator strength for the electronic excitation from the ground state to the excited state.

Sample preparation

Co-doped films of $Er(F-TPIP)_3:F_{16}MPC$ and $Er(F-TPIP)_3:Y(F-TPIP)_3$ were deposited by vacuum sublimation at a vacuum pressure of $\sim 10^{-7}$ mbar. 120 nm thick aluminium was evaporated onto the organic layer to protect the material from atmospheric degradation. Samples were prepared with $Er(F-TPIP)_3$ at a molecular concentration of 20%, 40%, 60%, and 80%. In $F_{16}MPC$ doped films, they had an identical amount of $F_{16}MPC$ chromophore (75 nm thick). In the $Y(F-TPIP)_3$ doped film, the thickness of $Y(F-TPIP)_3$ was kept constant as 163 nm.

Instruments and measurements

The UV-vis electronic absorption measurements were performed using a Shimadzu UV-2600 spectrophotometer. The Fourier-transform infrared (FT-IR) spectrum was collected in the range

400–4000 cm^{-1} on a Nicolet Fourier spectrophotometer using KBr pellets. The mass spectrum was measured with a MALDI-TOF/TOF Mass spectrometer, Ultraflex Treme TOF/TOF (Bruker). The analysis of chemical composition was performed on a Zeiss Merlin Compact scanning electron microscope with an energy dispersive X-ray spectrometer.

Emission spectra and TRPL measurements

For the PL measurements, lasers of different wavelengths were used to excite the samples; the emission from the samples were guided into a spectrometer (Jobin Yvon Horiba Triax 550), and the reflected laser light was removed by putting high-pass filters in front of the spectrometer. The spectrometer was connected to a photomultiplier (PMT), and the signal from the PMT was transmitted to an oscilloscope or a lock-in-amplifier for time-resolved or emission spectrum measurements, respectively. A Q-switch Nd:YAG laser was used to produce pulsed laser, and the laser wavelength was tuned by using an optical parametric oscillator (OPO). Pulsed laser beams were made to be incident on co-doped films to give TRPL data. The data were fitted by using exponential functions: $I(t) = I_0 + \sum_i A_i \exp[-t/\tau_i]$. A lifetime component percentage is obtained by an expression of $A_i \tau_i / \sum_i A_i \tau_i$. An average lifetime is obtained by an expression of $\langle \tau \rangle = \sum_i \left[\tau_i A_i \tau_i / \left(\sum_i A_i \tau_i \right) \right]$.

Sensitization measurement

Two apertures were used to ensure the consistency of the alignment of the optical path. A mirror, made by growing an



aluminium circle with 2 mm diameter on the 20 mm × 20 mm glass substrate, was set to an angle of 45° to reflect light normally onto the sample and allow the PL to be collected. An aperture with a 1 mm diameter hole was placed directly in front of the sample, and the laser was defocused to ensure uniform illumination over the whole sample. Once the optical path was set up, the measurement was started with the Er(F-TPIP)₃:Y(F-TPIP)₃ co-doped films; the 655 nm laser was used to directly excite the Er³⁺ ions, and the emission was monitored at 1532 nm. The laser power was increased from low to high. At each power, ten measurements were taken to determine the statistical errors. After measuring with the Y(F-TPIP)₃ doped films, the experiment was repeated using the Er(F-TPIP)₃:F₁₆ZnPc/F₁₆PtPc co-doped films. After the data were collected, the sample was replaced with a calibrated silicon detector to measure the excitation power density on the sample.

Conflicts of interest

There are no conflicts to declare.

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