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Phase evolution of electrochemically potassium intercalated graphite†

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Graphite is commonly known as a layered material to accommodate alkali metal ions between graphene layers and is used as a negative electrode material of most Li-ion batteries and for developing K-ion batteries. Phase evolution of graphite during chemical K-vapor intercalation was well studied in the 1920s–1980s, but that during electrochemical (de)potassiation is different and not fully clarified yet. Herein, we elucidate the phase evolution behaviors of electrochemically potassium intercalated graphite by *operando* X-ray diffraction in comparison to those of a Li system based upon optimizing the graphite electrode and electrolyte conditions for precise analysis. The *operando* diffraction data and first-principles calculations results reveal staging transformations from graphite to stage-1 KC_8 through disorderly stacked high stage, stage 4L, stage 3L, stage 2L, and stage 1 phases and reversal transitions including hysteresis. Based on the experimental and theoretical data, we propose structural change mechanisms of graphite during electrochemical K-intercalation and deintercalation on the basis of Daumas–Hérolde defects and the defect disappearance by complete K-intercalation.

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Introduction

From the 1990s onward, graphite has been broadly used as a negative electrode material in most Li-ion batteries owing to high gravimetric and volumetric capacities, low operation potential, and high electronic conductivity.^{1–3} Electrochemical lithium intercalation (lithiation) into highly crystalline graphite proceeds through formation of staged lithium graphite intercalation compounds (Li-GICs). In the 1950s–1970s, Li-GICs were chemically prepared by lithium-vapor synthesis⁴ or annealing techniques⁵ by Hérolde *et al.*, and reversible electrochemical Li-intercalation was demonstrated later.⁶ Similar to Li-GICs, potassium GICs (K-GICs) were chemically synthesized by potassium-vapor synthetic techniques in the 1920s–1980s.^{7,8} As K-GICs are stable at room temperature similar to Li-GICs,

graphite is expected to be a promising candidate for use as a negative electrode material for K-ion batteries. Actually, in 2015, Jian,⁹ Luo,¹⁰ and our group.^{11–13} demonstrated reversible electrochemical potassium intercalation into graphite in a non-aqueous K cell at room temperature and reported reversible formation of K-GICs through electrochemical potassium intercalation.

Generally, GICs have various periodical stage structures which are described as “stage *n*” where the stage index *n* refers to the number of graphene layers stacked between intercalant (guest) layers; *e.g.* the structure of stage-3 Li-GIC consists of Li intercalant layers stacked between every three graphene layers. Staging structural evolution of Li-GICs during lithium (de) intercalation was studied with *in situ* and *ex situ* X-ray diffraction (XRD) by Dahn¹⁴ and Ohzuku *et al.*¹⁵ Lithium intercalation into graphite proceeds along with phase transition in the following sequence; graphite → dilute stage 1 → stage 4 → stage 3 → “liquid-like” stage 2 (denoted as 2L) → stage 2 (LiC_{12}) → stage 1 (LiC_6) at room temperature.

In the potassium case, Jian's⁹ and our studies¹¹ on electrochemical potassium intercalation into graphite revealed phase evolution in a similar but different manner from the lithium case, that is, graphite → stage 3 (KC_{36}) → stage 2 (KC_{24}) → stage-1 KC_8 , which was proved from *ex situ* XRD measurements of potassiated graphite electrodes. This electrochemical process was compared with chemical potassium intercalation; graphite → stage 4 → stage 3 → stage 2 → stage 1.¹⁶ Also, Luo *et al.* proposed a different mechanism of potassium intercalation:

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graphite \rightarrow stage 3 (KC_{24}) \rightarrow stage 2 (KC_{16}) \rightarrow stage-1 KC_8 , which was deduced from density functional theory (DFT) calculations although the *ex situ* XRD pattern of only stage-1 KC_8 was proved in the literature.¹⁰

Pramudita,¹⁷ Fan,¹⁸ and our group¹³ reported *operando* XRD patterns during electrochemical potassium (de)intercalation into graphite, but did not mention the possible dilute stage-2 KC_{16} . Liu *et al.* proved electrochemical formation of several types of stage-2 phases by using *in situ* Raman spectroscopy in a K cell while any evidence of formation of the different stage-2 phases was hardly found in the *operando* XRD patterns.¹⁹ Phase evolution studies employing laboratory-scale *operando* XRD generally require slow-rate operation of charging and discharging processes to earn data acquisition time for enhancing intensities and to obtain the phase close to the equilibrium state due to the relatively low intensity of the X-rays. Thus, suppression of side reactions such as electrolyte decomposition is highly required for *operando* XRD measurements. Insufficient passivation of graphite and counter K metal electrodes with the solid electrolyte interphase (SEI)²⁰ always results in larger irreversible capacity during usual slow-rate *in situ* operation of a K cell.

In this paper, we examined the influence of the graphite particle size and binders in a composite electrode and electrolyte salts and solvents in a K cell on the electrochemical intercalation/deintercalation properties of K||graphite cells to avoid the issue of passivation and the irreversible reaction. Then, the open-circuit voltage (OCV) of the K-GIC electrodes was measured and compared with that estimated by DFT calculations. Based on this optimization and observation, we conducted *operando* XRD measurements of the graphite electrode in a K cell in comparison to those in a Li||graphite cell in order to understand the phase transition mechanism of K-GICs and their difference from Li-GICs.

Experimental

Electrodes and electrochemical cells

Electrochemical measurements were carried out by using R2032-type coin cells (Hosen Corp.) and three electrode cells (Toyo System Co., Ltd.) assembled in an argon-filled glove box. The working electrode, consisting of natural graphite (SNO3, SNO15, and SNO30, where the number represents the median size (μm) of graphite particles, SEC Carbon Ltd.) and a binder in 90 : 10 weight ratio was prepared by mixing with *N*-methyl pyrrolidone (NMP) or deionized water, coating on Al foil, and drying at 80 °C under vacuum overnight. The binders used were poly(vinylidene fluoride) (PVdF, PolySciences) and sodium polyacrylate (PANa, Kishida chemical Co., Ltd.). The mass loading of graphite was *ca.* 1.5 mg cm^{-2} for electrochemical evaluation and *ca.* 3.3 mg cm^{-2} for *in situ* XRD measurements. The counter electrode was potassium metal (Sigma-Aldrich Co., LLC) or lithium foil (THE HONJO CHEMICAL CO.) and a glass fiber filter (GB-100R, ADVANTEC Co.) was used as a separator. Electrolyte solution used in this study was 1.0 mol dm^{-3} potassium bis(fluorosulfonyl)amide (KFSA, Solvionic S.A.) and lithium bis(fluorosulfonyl)amide (LiFSA, Kanto Chemical Co., Inc.)

dissolved in mixed solvent of ethylene carbonate (EC, Kishida Chemical Co., Ltd.) and diethyl carbonate (DEC, Kishida Chemical Co., Ltd.) at 1 : 1 v/v. Potassium bis(trifluoromethylsulfonyl)amide (KTFSa, Kanto Chemical Co., Inc.) and KPF_6 (Tokyo Chemical Industry Co., Ltd.) salts as well as propylene carbonate (PC) and dimethyl carbonate (DMC) solvents purchased from Kishida chemical Co., Ltd., Japan, were used.

Electrochemical tests

OCV curves were collected by using the galvanostatic intermittent titration technique (GITT) for Li||graphite and K||graphite cells applying a constant current at C/30 (= 12.4 mA g^{-1} for Li and 9.3 mA g^{-1} for K system) rate for 30 min followed by a relaxation process for 3 h. Charge–discharge (corresponding to intercalation–deintercalation) measurements were carried out in the voltage range of 0.0–2.0 V vs. K^+/K or Li^+/Li in a constant current (CC) mode and 0.002–2.0 V vs. K^+/K or Li^+/Li in a CC–CV (constant voltage) mode. A constant voltage of 0.002 V was applied after reaching 0.002 V only during the charging process (intercalation process). For *operando* XRD measurements, a two-electrode *in situ* XRD cell equipped with an Al-coated Be window (Rigaku Corporation)²¹ was used in the voltage range of 0.0–2.0 V vs. K^+/K or Li^+/Li at a current rate of C/30 for the K||graphite cell and C/40 rate for the Li||graphite cell.

Structural and surface analysis

Operando XRD patterns were collected with a laboratory XRD diffractometer (MultiFlex, Rigaku Corporation) equipped with a high-speed position sensitive detector (D/teX Ultra, Rigaku Corporation) by setting the above-mentioned *in situ* XRD cell on the specific sample stage and using Ni-filtered $\text{Cu K}\alpha$ radiation (weighted average $\lambda = 1.5418 \text{ \AA}$, $\lambda(\text{K}\alpha_1) = 1.5406 \text{ \AA}$, $\lambda(\text{K}\alpha_2) = 1.5444 \text{ \AA}$, $I(\text{K}\alpha_2)/I(\text{K}\alpha_1) = 0.5$). The morphology of the graphite electrodes was observed with a scanning electron microscope (SEM, JCM-6000, JEOL). Hard X-ray photoelectron spectroscopy (HAXPES) was employed for the tested graphite electrodes by using high excitation energy of hard X-rays, 7939 eV, and a photoelectron energy analyzer of R-4000 (Scienta Omicron) at BL46XU at SPring-8, Japan. The photoelectron detection angle and pass energy of the analyzer were 80° and 200 eV, respectively. Electrochemically tested graphite electrodes were carefully taken out from cycled coin cells, rinsed with PC and then DEC, dried at room temperature in an Ar-filled glovebox under ambient pressure, and transferred using a transfer vessel to avoid air exposure. The detailed setup and conditions of the HAXPES measurements are described in our previous paper.²² The binding energy of the obtained spectrum was calibrated with the binding energy of the sp^2 carbon of graphite being 284.6 eV. Photoelectron peaks were deconvoluted with the peak-fit program, Fityk, with Pseudo-Voigt functions.

First-principles calculations

First-principles calculations were performed within the framework of DFT in Perdew–Burke–Ernzerhof (PBE) parametrization as implemented in the Vienna *ab initio* simulation package



(VASP). We selected the vdW-optPBE functional²³ to account for the van der Waals interactions. The projector augmented wave (PAW) method was used with the plane-wave basis cutoff of 1000 eV. We used the following PAW potentials provided by VASP: C_h, Li_sv, and K_sv. For the structural optimization and energy calculations, the Brillouin zone was sampled with a Γ -centered $20 \times 20 \times 10$ k -point mesh. Atomic positions and the cell volume were optimized until the forces on each atom were converged to less than 10^{-2} eV \AA^{-1} .

Results and discussion

Electrochemical properties

First, the influence of the particle size of graphite on the electrochemical K-intercalation properties was examined using non-aqueous coin-type K cells at a current rate of 25 mA g^{-1} in the voltage range of 0.0–2.0 V vs. K in a CC mode as shown in the ESI, Fig. S1.† Three different types of natural graphite having different median sizes (3, 15, and 30 μm) were selected, and their composite electrodes with a PVdF binder deliver almost the same discharge (deintercalation of potassium) capacities of ca. 240 mA h g^{-1} at the initial cycle. As expected from the difference in the surface area, the smallest graphite sample (3 μm) exhibits the lowest coulombic efficiency of ca. 60% at the initial cycle, which is likely due to electrolyte decomposition to form the SEI layer on the particle surface. Nevertheless, under our test conditions, the smallest particles of graphite are beneficial to achieve larger reversible capacity with good capacity retention during cycles, though graphite particles having ca. 10–15 μm in diameter are generally suitable for practical use in Li-ion batteries to suppress irreversible side reactions. For *in situ* measurements, we used 3 μm graphite powder for further tests.

According to Zhang's²⁴ and our reports,^{13,25} selection of potassium electrolyte salts significantly affects the reversible capacities, coulombic efficiency, and cycle stability of the graphite electrode in a K cell. Using PVdF as a binder for the graphite electrodes, we confirmed that the highest coulombic efficiency and the superior cycle stability are obtained in 1.0 mol dm^{-3} KFSA/EC : DEC (1 : 1 v/v) electrolyte compared to those in 0.8 mol dm^{-3} KPF₆/EC : DEC or 1.0 mol dm^{-3} KTFSA/EC : DEC as shown in Fig. S2.† Although we considered anodic decomposition of the SEI²⁶ at a relatively high potential ($>1 \text{ V vs. K}^{\dagger}/\text{K}$) and examined the influence of the upper cutoff voltage, the graphite cell cycled at 0.002–2.0 V demonstrates slightly better cycle retention compared to those tested at 0.002–1.5 V and 0.002–1.0 V as shown in Fig. S3.†

Because of the 1.6-fold volume expansion from graphite to KC₈, the binder content and polymers should be another influencing electrode component on the electrochemical performance, as we have systematically studied.^{27,28} Dependence of the PVdF binder content in the graphite composite electrodes was checked in K||graphite cells compared with that for Li||graphite cells as shown in Fig. S4.† In the Li case, increasing the PVdF content from 2 wt% to 10 wt% suppresses capacity degradation during cycles. The K cells, indeed, exhibit better capacity retention with an increase in the PVdF content,

but even in the 10 wt% case the reversible capacity obviously decays during cycles. Surface analysis using HAXPES reveals a larger amount of alkyl carbonate-like species on the electrode for the K case than that of the Li case (Fig. S5†), indicating severe electrolyte decomposition due to insufficient passivation function of the SEI formed on the surface and/or insufficient mechanical strength against large volume changes of K-GICs during K-(de)intercalation.¹¹ We can solve these issues by substituting a functional binder of PANa or sodium carboxymethyl cellulose²⁹ for PVdF.

Fig. 1a and b display the galvanostatic charge–discharge curves and cycle stability of the graphite electrode with the PANa binder in K cells. The graphite–PANa electrode demonstrates a higher initial coulombic efficiency of 88% compared to the PVdF case (ca. 60%) and excellent long-term cycle stability with no capacity decay over 300 cycles^{11,13} (see the comparison in Fig. S6†). When PANa is used as a binder, the deposition of potassium alkyl carbonate-like species on the graphite surface is suppressed compared to the PVdF binder from HAXPES data (Fig. S7†). Compared to the PVdF case (Fig. S4†), an obvious influence is not found for the PANa content dependence (see Fig. S8†). Furthermore, when the PANa binder is used, the influence of different electrolyte salts of KFSA, KTFSA, and KPF₆ appears to be insignificant (Fig. S9†). We note that a solution of 1.0 mol dm^{-3} KFSA/EC : DEC still offers higher initial coulombic efficiency and cycle stability compared to 0.8 mol dm^{-3} KPF₆/EC : DEC.²⁵ Additionally, the influence of EC-based binary solvents of EC : DEC, EC : DMC, and EC : PC on the reversibility is negligible for the PANa electrode during 100 cycles as confirmed by Fig. S10,† but 1.0 mol dm^{-3} KFSA/PC

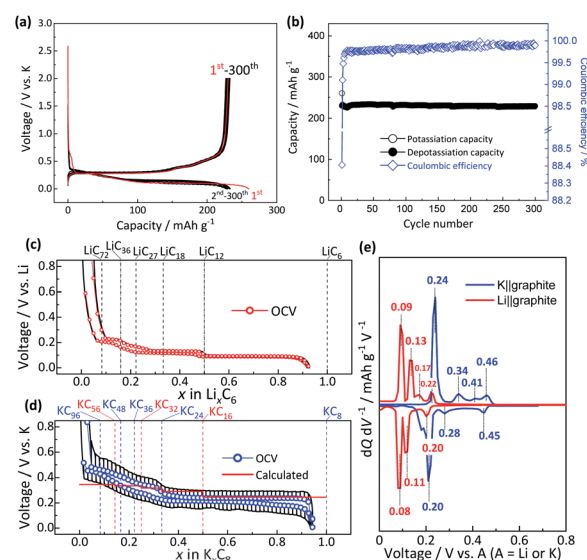


Fig. 1 (a) Galvanostatic charge–discharge curves and (b) long-term cycling stability of the graphite electrode with the PANa binder in a K cell. OCV profiles of the graphite electrode in a (c) Li cell and (d) K cell at C/30 rate after the 2nd galvanostatic charge–discharge cycle. The calculated profiles of K_xC_8 (red solid line) are shown as the inset in panel 'd'. (e) dQ/dV curves of the OCV plots of Li (red solid lines) and K (blue solid lines) cells.



results in continuous electrolyte decomposition at 0.7 V vs. K (equal to 0.8 V vs. Li), which is likely due to co-intercalation of PC-solvated K^+ ions and/or dissolution of SEI components, leading to graphite exfoliation and electrolyte decomposition²⁴ similar to a Li|LiClO₄-PC|graphite cell.^{30–32} Thus, we selected an electrolyte of 1.0 mol dm⁻³ KFSa/EC : DEC for the subsequent studies. Even at a high current rate of 1C (= 279 mA g⁻¹), the graphite electrode exhibits quite stable reversible capacities with high coulombic efficiency over 300 cycles in a three-electrode K||graphite cell (see Fig. S11†).

By using the three-electrode cell and the composite electrode of 3 μm graphite particles containing 10 wt% PANa binder, we further confirmed the good rate performance of the K||graphite cell. Charge–discharge rate capability of Li and K cells was examined and compared (see Fig. S12†), and the K||graphite cell demonstrates better rate performance than the Li cell owing to the larger potential margin between the charging plateau potential and the metal-plating potential. When we fixed the charging (intercalation) rate to 25 mA g⁻¹ and varied the discharging (deintercalation) rate of the Li and K cells, both the cells demonstrate comparable rate capability up to ca. 100C whereas larger polarization is observed for the K cell (Fig. S13†). The polarization in the Li cell depends on the lithium concentration in graphite, but that in the K cell simply and linearly increases with the current values applied to the K cell. This result might represent larger resistance at the electrode/electrolyte interface rather than K^+ diffusion in K-GICs. Surface and charge transfer resistances in a K||graphite cell are often higher than those of the Li cell.³⁴ Looking back at the influence of electrolyte and additive development on improving kinetics of Li-ion and Na-ion chemistry,^{35–37} future development of electrolyte and additives for K-ion batteries is rationally predicted to enhance the kinetics of the potassium intercalation process especially at the interface of a graphite electrode, which is ongoing in our laboratory.

Fig. 1c and d display the OCV profiles as a function of the intercalant concentration for the graphite electrodes in coin-type Li and K cells measured at the 6th cycle at 25 °C. The intercalant concentrations of Li and K, which were calculated from the charge passed, nearly reach $x \approx 1$ in Li_xC₆ and K_xC₈, respectively, at the end of the charging (intercalation) process while the values reversibly become almost zero at the end of the discharging (deintercalation) process. These results indicate the successful optimization of the electrode and appropriate selection of electrolyte components for the sufficient passivation of graphite. Actually, natural graphite used in this study has high crystallinity³⁸ and exhibits excellent reversibility of the lithium and potassium intercalation reactions without apparent electrolyte decomposition owing to formation of a suitable SEI on the graphite surface. The Li cell shows several distinctive plateaus related to staging reactions, and the formula of Li_xC₆, which can be described as LiC_{6/x} at the end of each voltage plateau corresponds to LiC₁₂, LiC₁₈, LiC₂₇, LiC₃₆, and LiC₇₂ as guided by dotted lines in Fig. 1c, whose compositions are well consistent with those reported previously.^{14,15,39,40} The OCV curve confirms the reversible feature, but a little voltage hysteresis between lithium intercalation and deintercalation

processes is also observed in the limited compositional range of $1/6 \leq x \leq 1/3$ in Li_xC₆ similar to data in previous literature.^{15,41} On the other hand, the K cell also exhibits several voltage plateaus, but these are with vague voltage-changes especially in the low K amount region (ca. $x < 0.33$ in K_xC₈). As seen in Fig. 1d, the formula of K-GICs K_xC₈, which can be also described as KC_{8/x}, at the end of each plateau is close to KC₂₄, KC₃₆, KC₄₈, and KC₉₆, of which compositions are in agreement with those of chemically synthesized K-GICs.⁴² A little voltage hysteresis is also observed in $1/6 \leq x \leq 1/3$ in K_xC₈. The differential curves of the OCV profiles in Fig. 1e obviously display reduction and oxidation peaks, whose voltage values are consistent with those in the literature.^{14,15,43} In both Li and K cells, all the peaks are attributed to redox couples except for the oxidation peaks at 0.17 V and 0.41 V in Li and K cells, respectively, which correspond to the voltage hysteresis between intercalation and deintercalation processes.

To understand the theoretically and thermodynamically stable phases of K-GICs and compare the calculated voltage curve to the experimental ones, we performed DFT calculations on K-GICs. Fundamentally, graphite is composed of stacked graphene layers where the carbon atoms are planarly bonded with sp² hybrid bonds in a honeycomb six-ring configuration. The graphene layers are bound by weak van der Waals interactions and are stacked in an ABAB sequence along the *c*-axis with an interlayer distance of 3.354 Å as shown in Fig. 2a and b. When alkali metal ions are intercalated into the interlayer space in graphite to form stage *n* GICs, the stacking sequence of the

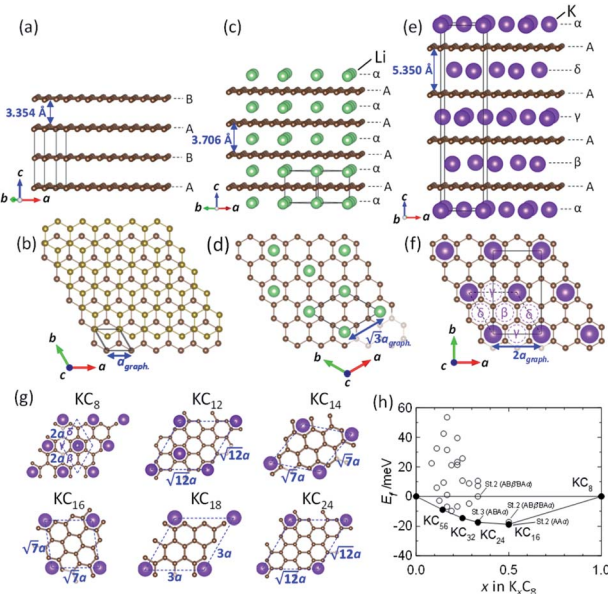


Fig. 2 (a, c and e) Schematic illustrations of crystal structures and (b, d and f) projected in-plane structures along the *c*-axis for (a and b) graphite, (c and d) stage-1 LiC₆, and (e and f) stage-1 KC₈. (g) In-plane unit cells for $i \times j$ ($2 \leq i, j \leq \sqrt{12}$) superlattices for which the carbon atom amount per potassium atom is denoted as labeled KC_{*m*} compositions. (h) Formation energies E_{form} of stage-*n* K-GICs as a function of total potassium concentration *x* in K_{*x*}C₈ which can be described as KC_{8/*x*}.



graphene layers changes into A|A, where the vertical bar denotes an intercalated layer and a letter refers to graphite layers.¹⁶ Stage-1 GICs usually have two dimensional (2D) in-plane orderings of the guest atoms in the interlayers. Stage-1 LiC₆ has the 2D in-plane unit cell of LiC₆ with the $p(\sqrt{3} \times \sqrt{3})R30^\circ$ superlattice with the stacking sequence of AαAαAα in the temperature range of $T < 715$ K⁴⁴ (Fig. 2c and d). On the other hand, stage-1 K-GIC, KC₈, has the in-plane unit cell of KC₈ with the $p(2 \times 2)R0^\circ$ superlattice and a three dimensional (3D) stacking sequence of AαAβAγAδ below 610 K (ref. 45) as shown in Fig. 2e and f.⁴⁶ The in-plane and interplanar (stacking) orderings are known to be dependent on the in-plane density of intercalant ions.⁴² Thus, we considered six types of in-plane unit cells with KC_m compositions; *i.e.* KC₈, KC₁₂, KC₁₄, KC₁₆, KC₁₈, and KC₂₄ (Fig. 2g) and several types of stacking sequences such as AAαAAα and ABβ'BAα (Table S1†). When we considered the stacking sequences including three types of graphene layers, A, B, and C types, the energies of AB|BA|AB| and AB|BC|CA| or ABA|ABA| and ABA|ACA| stacking sequences are expected to be close because the relative positional relationship between the two graphene layers, AB, BC, and CA types, is the same. Combining the in-plane unit cells and stacking sequences mentioned above, 30 different structures of K_xC₈, corresponding to stage-*n* KC_m×*n*, are obtained, and their formation energies were calculated as plotted in Fig. 2h. From the structures having the lowest formation energies in the convex hull, DFT calculations predict stable staging structures (see their in-plane structures and stacking sequences in Table S2†) and the phase evolution sequence upon potassium intercalation as follows; graphite → stage 4 (KC₅₆ and KC₃₂) → stage 3 (KC₂₄) → stage 2 (KC₁₆) → stage 1 (KC₈). Based on the calculation results, the average voltage between the compositions, *x*₁ and *x*₂, is computed as

$$\bar{V} = - \frac{E_{K_{x_2}C_8} - E_{K_{x_1}C_8} - (x_2 - x_1)E_K}{e(x_2 - x_1)} \quad (1)$$

where *x*₁ and *x*₂ are the concentration of K in stable K_xC₈ listed in Table S2,† $E_{K_{x_1}C_8}$ and $E_{K_{x_2}C_8}$ are the energies of K_{x₁}C₈ and K_{x₂}C₈, respectively, E_K is the energy of metallic potassium, and *e* is the charge of an electron. A red solid line in Fig. 1d presents the calculated voltage profiles for the stable structures of K_xC₈ as a comparison with experimental voltage curves. The calculated voltage profile agrees with the measured OCV curves. The fact suggests that stage-2 KC₁₆ is more stable than stage-2 KC₂₄ and is expected to be formed during potassium (de)intercalation. The stable structure of stage-2 KC₁₆ is slightly different from that in Luo's results¹⁰ and is rather similar to Lenchuk's ones⁴⁷ (see Table S3†). Consequently, our computational results similarly predict thermodynamically stable formation of stage-2 KC₁₆.

Operando XRD

To elucidate the phase evolution behaviors of graphite, *operando* XRD was conducted for the Li||graphite and K||graphite cells during galvanostatic charge and discharge at the third cycle (to ensure SEI formation) at a quite slow rate of C/40 for Li

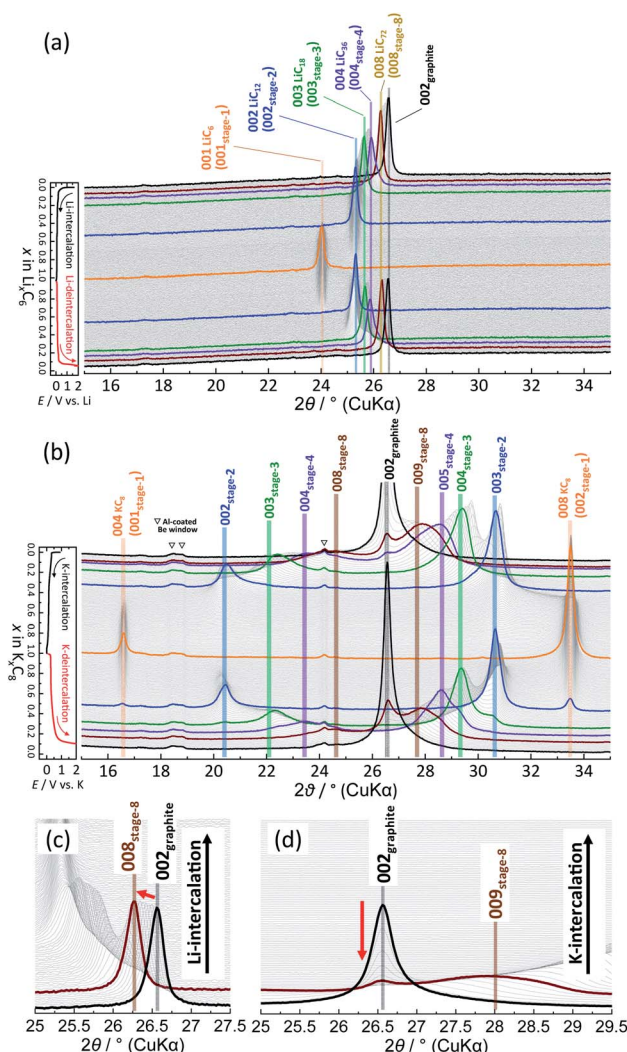


Fig. 3 *Operando* XRD patterns of graphite electrodes in (a and c) Li||graphite and (b and d) K||graphite cells at a current density of 9.3 mA g⁻¹; in the diffraction angle regions of (a and b) 15–35°, (c) 25–27.5°, and (d) 25–29.5°. The CV mode at 2 mV for 5 h was employed only in the K system at the end of a K-intercalation process.

and C/30 for K systems. Fig. 3a and b show the *operando* XRD patterns of graphite electrodes in a Li cell and a K one, respectively; an enlarged view for the Li cell is shown in Fig. S14.† Due to flake-like shape of graphite particles (Fig. S7†) and preferred orientation, only 00*l* Bragg diffraction peaks are observed.

The 00*l* reflection at the diffraction angle, 2θ of 24°–26° for the Li system and 20°–26° for the K system gives the repeat distance I_c corresponding to the minimum distance between intercalant layers.⁴² Considering the structure of a stage *n* compound, I_c can be calculated with the following equation assuming that the distance of the empty interlayer spacing (3.35 Å)⁴⁸ and that of the interlayer spacing filled with Li (3.706 Å) as in LiC₆ (ref. 5) or K (5.35 Å) as in KC₈ (ref. 16) remain constant across different stages.

$$I_c = (n - 1)d_0 + d_i, \quad (2)$$



where d_0 and d_i represent the above-mentioned distances of the empty interlayer spacing and the interlayer spacing filled with the intercalant, respectively. These distances of the empty and filled interlayer spacings are, indeed, known to be different in the experimental results for GICs: e.g., 3.33 Å and 3.748 Å for stage-3 LiC₃₀, 3.33 Å and 3.735 Å for stage-2L LiC₁₈, and 3.24 Å and 3.785 Å for stage-2 LiC₁₂, respectively.⁴⁹ However, the differences from the standard values of 3.35 Å and 3.706 Å in the Li-GICs counteract each other, and the I_c values given by eqn (2) are almost identical to the experimental ones. The average interlayer distance d_{00n} is then calculated from the repeated distance I_c as follows:

$$d_{00n} = \frac{I_c}{n} \quad (3)$$

The calculated d_{00n} values are shown in Table S3† and these were compared to the observed d -spacings of the $00n$ reflections to identify the stage number at each state-of-charge. From the indexation of the $00n$ reflection with the d_{00n} value in the d -spacing, the d -spacing of the $00n+1$ reflection is given by the following equation:

$$d_{00n+1} = d_{00n} \times \frac{n}{n+1} \quad (4)$$

In the KC_x system, the reflections at 20°–26° and 26°–34° correspond to $00n$ and $00n+1$, respectively, in this hypothetical structure, confirming the presence of stage-ordered potassium-intercalated graphite. We note that indexation may differ for the structures with an interlayer ordering along the stacking direction, as in stage 1 KC₈ where the c unit-cell parameter is quadrupled and potassium layers have a specific stacking ordering with the AαAβAγAδ sequence⁸; the indexation based on the KC₈ structure was used in this case instead.

In the Li case, a 002 Bragg diffraction peak of graphite located at a 2θ of 26.55° shifts toward a lower diffraction angle during intercalation and shifts back to a high angle during deintercalation, which indicates an enlarged average interlayer distance by lithium intercalation into the graphite host.⁵ The peak shift is not fully continuous, and multiple and staging phase-transitions are observed from graphite to LiC₆ during an intercalation process, and a deintercalation process from LiC₆ to graphite is accompanied with slight structural change hysteresis. The structural changes of graphite during lithium intercalation occur in the transition sequence: graphite → dilute stage 1 → stage 4L–3L → (stage 2L →) stage 2 (LiC₁₂) → stage 1 (LiC₆). During the lithium deintercalation process, the phase evolution proceeds in the sequence; stage 1 (LiC₆) → stage 2 (LiC₁₂) → stage 2L → stage 4L–3L → dilute stage 1 → graphite, where L refers to the “liquid-like” distribution of lithium atoms in the interlayer space. For instance, dilute stage-2 LiC₁₈ has a lower in-plane lithium density of a LiC₉ in-plane unit than a LiC₆ unit in stage-2 LiC₁₂ and has a “liquid-like” in-plane distribution of lithium atoms at room temperature.³⁹ Thus, a primitive in-plane unit of $p(1 \times 1)R0^\circ$ is often employed as an averaged in-plane unit for the structural refinement of

stage 2L^{50,51} without consideration of a possible short-range ordering of lithium atoms although stage-2L LiC₁₈ might have a “liquid-like” local in-plane ordering as a $p(3 \times 3)R0^\circ$ superlattice which was proposed with XRD data⁵² and predicted by DFT calculations.⁵³ Despite the “liquid-like” in-plane distribution of lithium atoms, stage-2L LiC₁₈ is known to have an interplanar ordering of the AB|BA| stacking sequence at room temperature.^{39,54} Furthermore, stage-2L LiC₁₈ is a metastable phase, of which the pure phase is prepared by an electrochemical method,^{39,52} and transforms into dense stage-2 LiC₁₂ and stage-3 LiC₁₈ below ca. 290 K.⁵⁵ Thus, the operation temperature for the Li cell tests is an important factor, and Dahn and coworkers predicted the disappearance of the stage 2L phase below 10 °C based on the voltage curves.¹⁴ Similar to previous literature,^{51,56,57} a Bragg diffraction peak of stage 2L is clearly observed at a 2θ of 25.2° upon the lithium deintercalation process at room temperature in Fig. 3a. Although the stage 2L might be formed during the Li-intercalation process, distinct Bragg peaks for this phase are not obviously observed, which is consistent with the previous results of *operando* neutron diffraction (ND) measurements.^{51,56,57}

For the K cell in Fig. 3b, the 002 Bragg diffraction peak of graphite first splits into two peaks, and they shift gradually to lower and higher angles upon potassium intercalation and reversibly go back to the initial position during the deintercalation process, indicating that a reversible change in average interlayer distances increased/decreased by potassium intercalation/deintercalation, respectively. Despite 1.6 times lattice expansion of graphite by complete K-intercalation and formation of stage-1 KC₈, the reversible potassium intercalation/deintercalation behaviors are confirmed with obvious multiple staging reactions.

A remarkable difference in phase transitions for Li and K systems is found at the early intercalation process. The magnified *operando* XRD patterns of Li and K cells in the high stage regions are shown in Fig. 3c and d, respectively. In the Li system, the graphite 002 peak shifts continuously to a lower angle without a significant change in the peak width as lithium intercalates into graphite. This indicates expansion of the interstitial space with homogeneously increasing in-plane lithium concentration (see Fig. S15a†). Thus, this phase is called dilute stage 1 (ref. 14) having a “gas-like” in-plane lithium distribution,⁵⁸ because the low lithium concentration of the intercalant layer has a negligibly weak interplanar interaction, and the graphite host retains the ABAB stacking manner without gliding of graphene sheets.⁵⁸ On the other hand, in the K cell, the graphite 002 peak remains at the initial peak position, but its intensity rapidly decreases as potassium intercalation proceeds, and a pretty broad peak simultaneously appears with peak shift toward a higher angle as shown in Fig. 3d. The broadening of $00l$ diffraction corresponds to reduction of coherence along the stacking axis, indicating increased stacking disorder⁵¹ as drawn in the schematic model in Fig. S15b.† It is most likely that potassium atoms randomly and heterogeneously intercalate into the graphite host to form locally staged K-GICs layers,⁵⁹ resulting in the stage disordering along the c axis in the high stage region ($n > 8$) at room



temperature. This transition behavior is consistent with that in the chemically prepared K-GICs.⁶⁰ Furthermore, the staging difference can be explained by Liu's⁶¹ and Lenchuk's⁴⁷ computational results. In both calculations, stabilizing GICs mainly require large negative binding energies contributed from the electrostatic and van der Waals interactions between layers and the ionization and decohesion energies of alkali metals. The binding energy is generally high when the concentration of intercalant ions is high. In contrast to the binding energy, the energy cost required to change the stacking of graphite layers from AB to AA and expand the interlayer spacing (so called graphite-deformation energy) is positive and increases with increasing the atomic size of the intercalant ions.⁴⁷ Although the graphite-deformation energies of Li-GICs are smaller than those of K-GICs due to the smaller size of the lithium ion, energy contribution from the AB-AA transition is large because Li-GICs have a narrower interlayer spacing and stronger van der Waals interactions between the graphene layers as well as stronger electrostatic and covalent interactions between alkali metal ions and the nearest graphene layers. The low concentration of Li intercalant does not generate enough binding energy to overcome the AB-AA transition, forming the dilute stage-1 Li-GIC with ABAB stacking remaining. On the other hand, in K-GICs, energy contribution from the AB-AA transition is relatively small, and the expansion of interslab spacing contributes to the deformation energy predominantly.⁴⁷ The binding energy is also large in K-GICs and overcomes the graphite-deformation energy. However, the interlayer interactions in K-GICs are weaker than those in Li-GICs and further weaker in lower potassium concentrations. These factors probably lead to a disordered stacking transition of the graphite layers as the disorderly stacked high-stage K-GIC is experimentally observed at room temperature in this study.

The operation temperature is an important factor in staging evolution of GICs as mentioned above for the stability of stage-2L LiC_{18} . In this study, we measured an OCV curve for a $\text{K}||\text{graphite}$ cell at 0 °C only in the potassium deintercalation process. However, no significant difference in the OCV from that examined at room temperature is confirmed from the differential curve as shown in Fig. S16.† This is because stage n K-GICs ($n \geq 2$) have a "liquid-like" in-plane structure without strong interplanar interactions with the facing graphene layers above 250 K.^{62,63} Furthermore, the dilute stage-2L KC_{24} is stable in the wide temperature range and represents no phase-separation reaction at 12–300 K (ref. 62 and 64) unlike the phase transition of the dilute stage-2L LiC_{18} into stage-2 LiC_{12} and -3 LiC_{18} in *ca.* 290 K.^{14,55}

The continuous and discrete changes in d -spacings as well as phase evolutions are effectively visualized in the color contour plots with the d -spacing and diffraction intensity for both the Li and K cells as shown in Fig. 4a and b where the d values were converted from the diffraction angles 2θ in Fig. 3a and b, respectively.

In the Li cell, the d -spacing corresponding to dilute stage 1 and stage 4L–3L continuously changes during charge and discharge. In contrast to continuous variations, a constant d -spacing is observed for stage 2 (3.52 Å) and stage 1 (3.71 Å) Li-

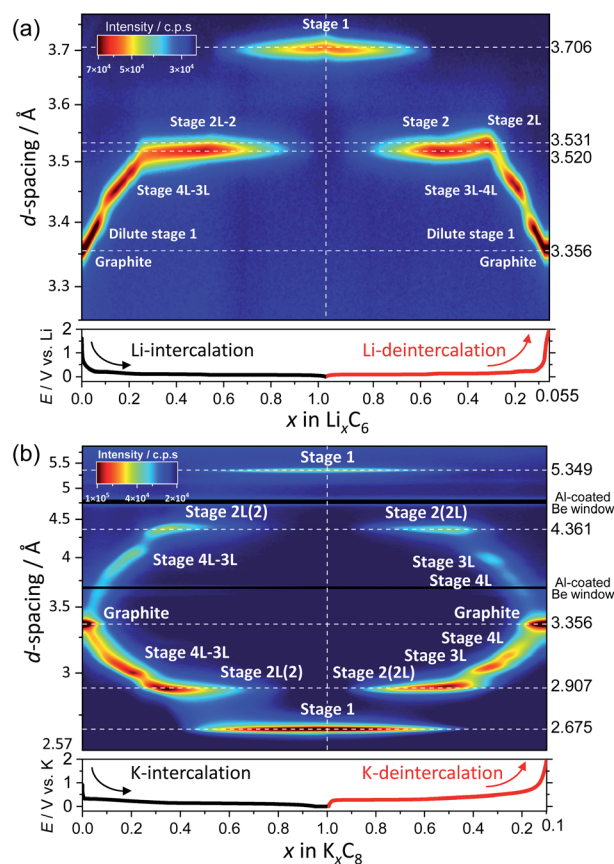


Fig. 4 Contour plots of the *operando* XRD patterns of graphite electrodes in (a) Li and (b) K cells at a current density of 9.3 mA g⁻¹.

GICs as a two-phase reaction corresponding to the flat voltage plateau is observed in the charge–discharge curves in Fig. 1c. In the deintercalation process, a wider d -spacing of stage 2L (3.53 Å) than that of stage 2 (3.52 Å) is also consistent with those in the literature.^{39,51,56,57}

On the other hand, the contour plots for the K cell (Fig. 4b) represent weak and broad diffraction peaks during the early intercalation process until the formation of stage 3L, which is indicative of the formation of a disorderly stacked high stage. Without clear observation of stage 4L in the contour plots, the d -spacing corresponding to stage 3L appears in the relatively early intercalation process and continuously changes like stage 4L–3L in Li-GICs. Upon the further potassium intercalation, a constant d -spacing for stage 2 (including 2L) (4.36 Å) and 1 (5.35 Å) appears along with the increase in (stage 1)/(stage 2) peak ratio, suggesting a two-phase reaction similar to the Li system. These results agree with a flat voltage plateau in $0.5 \leq x \leq 1$ in K_xC_8 (Fig. 1d). In contrast to the intercalation process, the d -spacing corresponding to stage 4L is clearly observed and isolated from that of stage 3L during the subsequent deintercalation process.

To understand more details of the phase evolution, we further analyzed the diffraction patterns by peak fitting with the pseudo-Voigt function (see the selected fitting results in Fig. S17†). The peak fitting enables precise determination of the



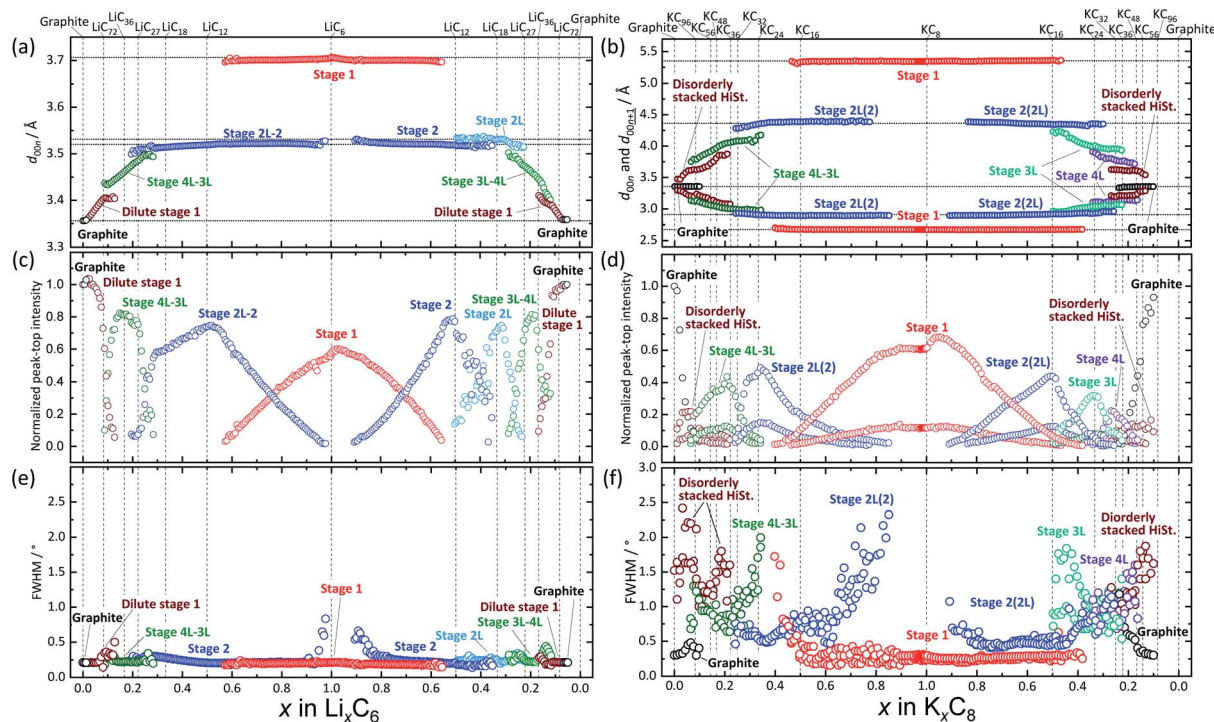


Fig. 5 Variation of (a and b) d_{00l} , (c and d) peak-top intensity, and (e and f) FWHM as a function of x amounts of (a, c and e) Li_xC_6 and (b, d and f) K_xC_8 .

d -spacing and the peak height and width of each phase component. Fig. 5 shows variation in the peak height and full width half maximum (FWHM) of d_{00l} diffraction as a function of x in Li_xC_6 and K_xC_8 . In the Li system (Fig. 5a, c, and e), stage-1 LiC_6 , -2 LiC_{12} , and -2L LiC_{18} phases are found to have averaged specific interlayer distances of 3.71 Å, 3.52 Å, and 3.53 Å, respectively. The different interlayer distances between stage 2 and 2L indicate different in-plane lithium densities sandwiched between neighbor graphene sheets.³⁹ In the other stages, d_{00l} values continuously vary as a function of lithium concentration, showing a variable in-plane lithium density although typical compositions and structures of some stages such as stage-3L LiC_{29-40} were reported.⁴⁹ Consequently, lithium intercalation proceeds in the following manner; graphite is transformed into dilute stage 1 in $x \leq 0.1$ in $\text{Li}_x\text{C}_6 \rightarrow$ stage 4L-3L in $0.1 \leq x \leq 0.3 \rightarrow$ stage 2 (including 2L) in $0.2 \leq x \leq 0.5 \rightarrow$ stage 2 + stage 1 in $0.5 < x < 1.0 \rightarrow$ stage 1 LiC_6 . During following lithium deintercalation from LiC_6 , phase evolution proceeds in the following sequence; stage-1 $\text{LiC}_6 \rightarrow$ stage 1 + stage 2 in $0.5 < x < 1.0 \rightarrow$ stage-2 $\text{LiC}_{12} \rightarrow$ stage 2 + stage 2L in $0.3 < x < 0.5 \rightarrow$ stage-2L $\text{LiC}_{18} \rightarrow$ stage 2L + stage 3L in $0.17 \leq x < 0.3 \rightarrow$ stage-3L-4L $\text{LiC}_{36} \rightarrow$ dilute stage 1 in $0.1 < x < 0.16 \rightarrow$ graphite.

It should be noted that peak-top intensity of stage 2 in the lithium intercalation process has two inflection points at LiC_{18} and LiC_{12} as seen in Fig. 5c. The fact might represent formation of stage 2L during the lithium intercalation process as a phase mixture between stage 2L and stage 2 as suggested by Ohzuku¹⁵ and Billaud.⁴⁹

Similar to the Li system, we discuss the K system (Fig. 5b, d, and f) with consideration of continuous and discrete changes in

the potassium concentration in graphite. In previous literature, chemically prepared K-GIC phases and the staging phase evolution are as follows: stage-1 KC_8 , stage-2L KC_{24} and KC_{28} ,⁶⁵ stage-3L KC_{36} , stage-4L KC_{48} , and stage-5L KC_{60} .⁸ From the *operando* XRD data of Fig. 5d, electrochemical potassium intercalation is found to proceed in the transition sequence of graphite \rightarrow graphite + disorderly stacked high stage in $x < 0.07$ in $\text{K}_x\text{C}_8 \rightarrow$ the high stage + stage 4L-3L in $0.07 \leq x < 0.22 \rightarrow$ stage 4L-3L at $x \sim 0.22 \rightarrow$ stage 4L-3L + stage 2L in $0.22 < x < 0.33 \rightarrow$ stage 2L in $0.33 \leq x \leq 0.38 \rightarrow$ stage 2L + stage 1 in $0.38 < x \leq 0.83 \rightarrow$ stage-1 KC_8 . In the potassium deintercalation process, the phase transitions proceed through the reversal process from stage-1 KC_8 to graphite. However, the potassium concentration range for each stage and the phase transition between stage 3L and 4L are different from those in the intercalation process, which is reasonably related to the reaction and voltage hysteresis between charge and discharge. The transition sequence and compositional ranges in the potassium deintercalation are as follows: stage-1 $\text{KC}_8 \rightarrow$ stage 1 + stage 2(L) in $0.48 < x < 0.90$ in $\text{K}_x\text{C}_8 \rightarrow$ stage 2(L) at $x \sim 0.48 \rightarrow$ stage 2(L) + stage 3L in $0.31 < x < 0.48 \rightarrow$ stage 3L $x \sim 0.31 \rightarrow$ stage 3L + stage 4L $0.20 < x < 0.31 \rightarrow$ stage 4L + disorderly stacked high stage in $0.13 < x < 0.20 \rightarrow$ disorderly stacked high stage + graphite in $0.05 < x < 0.13 \rightarrow$ graphite. Interestingly, obvious peak separation among stage-3L, -4L, and the high stage is found in the deintercalation process, while the continuous variation of the interlayer distance from the disordered high stage to the stage 3L is observed in the intercalation process as shown in Fig. 5d.

The transition hysteresis between the potassium intercalation and deintercalation is clearly found in the offset-free



operando XRD patterns (see Fig. S18†). In the deintercalation, three inflection points in intensities are observed at 2θ values of 29.40° , 28.68° , and 27.86° , and the reflections can be assigned to stage 3L, 4L, and the high stage, respectively. Reduction and gain of the intensities between the phases are indicative of two-phase reactions. On the other hand, on intercalation, the broad peak region at 2θ values of 26.56° – 29.76° can be divided into mainly three regions; a two-phase region of graphite and high-stage at 26.56° – 28° , a two-phase region of high-stage and stage 4L–3L at 28° – 29.08° , and a single-phase region of stage 4L–3L at 29.08° – 29.76° . In the former two two-phase regions of 26.56° – 28° and 28° – 29.08° , reduction and gain in intensity are found, and a continuous peak-shift and monotonous intensity gain are seen in the latter region of 29.08° – 29.76° . This hysteresis of the phase evolutions is consistent with the voltage hysteresis in the charge and discharge profiles around 0.41 V in Fig. 1d and e.

Furthermore, the maximum peak-top intensity of stage 2(L) is obtained at $x = ca. 0.5$ in K_xC_8 upon the potassium deintercalation process (Fig. 5d), indicating the formation of stage-2 KC_{16} . In contrast, the maximum intensity of stage 2(L) in the intercalation process is located at $x = 0.33$ in K_xC_8 corresponding to the KC_{24} composition. As mentioned in a previous report,⁵¹ lithium intercalation into a stage-3 phase produces a stage-2 phase having stacking faults and/or modulated domains between the stage-3 and -2 Li-GICs whereas almost fault-free stage-2 and -2L Li-GICs form during the reversal deintercalation process. The detailed phase transition mechanism related to crystallite domains is discussed in the later section.

It is worth mentioning that another notable feature of the K-GIC staging transition appears in the FWHM plots of Fig. 5f which is distinctly different from that of the Li-GIC plots in Fig. 5e. In the Li cell, larger values of the FWHM are observed only for the diffraction of stage-2 around $x = ca. 0.9$ in Li_xC_6 . Considering Scherrer's equation,⁶⁶ this broadening should originate from nanocrystal domains just after nucleation or before the disappearance of the stage-2 phase at compositions close to the terminal phases in the two-phase reaction of stage-1 and -2. Additionally, slight but certainly large FWHM values are seen at the boundary compositions for dilute stage 1/stage 4L–3L and stage 4L–3L/stage 2 as previously reported.⁵¹

In contrast to the Li system, the K system exhibits much larger FWHM values and pronounced variation as seen in Fig. 5e and f. This indicates drastic reduction in the stacking coherence of graphene and potassium layers along the c axis, probably related to variation of long-range ordering including stacking faults and micro-strain at the phase boundary between the K-GICs phases. Owing to a larger atomic size and a lower Lewis acidity of the potassium ion compared to the lithium one, in-plane distribution and concentration of alkali metal ions as well as stacking manners of graphene layers in the graphite host are different between K-GICs and Li-GICs.

Structures and the intercalation mechanism

From the above discussion including the stacking faults and/or modulated domains as well as in-plane and interlayer orderings

of Li- and K-GICs, we herein propose structural evolution of Li- and K-intercalated graphite schematically shown in Fig. 6.

In the Li system, at the beginning of lithium intercalation into graphite, “gas-like” dilute stage 1 is formed with retaining the ABAB-stacked graphite host structure with approximate compositions of graphite– LiC_{72} . When the in-plane lithium density with the same total composition as in the dilute stage-1 reaches LiC_{72} , further lithium intercalation produces A|A stacking around lithium by gliding the graphene layers to form stage 4L with ABAB|BABA| stacking⁵¹ and then stage 3L with ABA|ACA| stacking.⁴⁹ These phases, however, usually have stacking faults in the structures at room temperature^{49,67} because of the “liquid-like” in-plane lithium distribution and the weak interaction of the intercalant lithium ions with neighbor graphene layers at a low lithium concentration.

Stage-4L is generated as a two-phase reaction with the dilute stage-1, and its initial composition is close to LiC_{36} in total. Then, a continuous phase transition occurs between stage 4L and 3L according to the *operando* XRD patterns in Fig. 4a and 5a, c and e, which is consistent with those reported in the literature.^{14,51,57} In terms of a structural change mechanism, the stage transition between the even and odd n indices of stage- n GICs cannot be explained by using a typical flat layered model (so called Rüdorff–Hofmann model⁶⁸) but can be explained by a domain model with Daumas–Hérolde defects,⁶⁹ which was proposed by N. Daumas and A. Hérolde. As shown in Fig. 6a, stage 4L and 3L domains coexist and are separated by a Daumas–Hérolde type defect (so-called Daumas–Hérolde wall) as a domain boundary. Thus, phase transition proceeds like a single-phase reaction *via* migration of the Daumas–Hérolde defect boundary. As further lithium intercalation leads to reaching LiC_{36} in total composition, stage 2 domains are generated and co-exist with stage 4L–3L as well as the Daumas–Hérolde defects. The stage-2 LiC_{12} has the 3D-ordered structure with the in-plane unit cell of LiC_6 with the $p(\sqrt{3} \times \sqrt{3})R30^\circ$ superlattice similar to that of stage-1 LiC_6 and the unique stacking sequence of $AA\alpha AA\alpha$ below 500 K.^{5,54} In this staging region of LiC_{36} – LiC_{18} , dilute stage 2L having the “liquid-like” LiC_9 in-plane unit could be co-generated with stage 2 at room temperature as previous *operando* X-ray and neutron diffraction data also proved the low but trace amount of the stage 2L phase during the intercalation process.^{14,51,70} In our *operando* XRD data in Fig. 5c and as mentioned above, two inflection points are observed at LiC_{18} and LiC_{12} in the peak-top intensities of the stage-2 phase. The former point probably corresponds to the maximum phase fraction of stage-2L. When the total composition reaches the later compositional point of LiC_{12} by further lithium intercalation, a single and 3D ordered phase of stage-2 LiC_{12} is expected to be obtained according to our *operando* data in Fig. 5c. However, on the basis of the analogy to the potassium intercalation and the differences between intercalation and deintercalation described in a later section, it is speculated that the LiC_{12} phase formed in the intercalation process has Daumas–Hérolde defects that prevent the integration of the LiC_{12} domains.

Once stage-1 LiC_6 with 3D $A\alpha A\alpha A\alpha$ stacking is formed by complete lithium filling, Daumas–Hérolde defects disappear. So,



deintercalation.⁵¹ We employed this reaction model to the K-GIC system.

In the K system of Fig. 6b, at the beginning of the potassium intercalation process in the approximate composition range of graphite-KC₉₆, graphite partially and gradually transforms into a disorderly stacked high stage accompanied by gliding of graphene layers to form A|A or B|B stacking unlike “gas-like” lattice formation of dilute stage-1 Li-GIC retaining the ABAB stacked graphite host structure as mentioned above in Fig. 3c and d. Although the structure of the disorderly stacked high stage is schematically illustrated as the Rüdorff–Hofmann model⁶⁸ representing stacked empty and filled interlayers (Fig. 6), the Daumas–Hérolde model with island domains might be suitable for depicting the high stage K-GIC structure. Due to intercalation of large potassium ions and formation of large interlayer spacings, significant coherency strain between the graphite and high stage K-GIC domains as well as a large in-plane size of graphite domains contribute to the random distribution of the stage domains as proposed by Safran.⁷¹

Further potassium intercalation into the high-stage K-GIC produces stage 4L–3L, and the two-phase reaction proceeds in the composition range of KC_{96–36} (see Fig. 5b, d, f, and S18†). Although the stage 4L–3L transition looks like a single-phase reaction during the potassium intercalation process and is a two-phase reaction during the deintercalation process as mentioned above, the stage 4L and -3L K-GICs are reported to have a “liquid-like” in-plane potassium distribution at room temperature⁶³ similar to the stage *n* Li-GICs (*n* > 2). Furthermore, variable peak positions were reported for the *in situ* XRD patterns of stage-*n* K-GICs (*n* > 2) under chemical potassium vapor pressure.⁷² Thus, both the stage-4L and -3L phases are expected to have variable potassium concentration of the in-plane unit as plausible composition ranges are shown in Fig. 6b. Furthermore, despite the “liquid-like” in-plane potassium distribution at room temperature, an interplanar ordering (stacking ordering) for stage-*n* K-GICs (*n* ≥ 2) is known as follows:^{16,73}

$$|(AB)_{n/2}|(BC)_{n/2}|(CA)_{n/2}| \text{ for even } n, \quad (5)$$

$$|A(BA)_{(n-1)/2}|A(CA)_{(n-1)/2}| \text{ for odd } n. \quad (6)$$

The difference in stacking sequences between even- and odd-numbered stages requires the Daumas–Hérolde defects as domain boundaries in staging evolution of K-GICs like the Li-GICs system. Thus, the stage 4L–3L transition proceeds *via* formation of multiple domains separated by Daumas–Hérolde defects as shown in Fig. 6b even though the variation of intensities and peak-shift looks like a single-phase reaction in the *operando* XRD patterns (see Fig. S18†). When the potassium concentration increases and the total composition of K-GICs reaches KC₃₆, the stage 2L (including stage 2) phase appears and coexists with stage 3L as a two-phase reaction in the compositional range of KC₃₆–KC₂₄ (Fig. 5d and f). The stage 3L–2L transition also requires Daumas–Hérolde defects Fig. 6. The subsequently formed stage-2L K-GIC also retains the defects within the phase as the stage 2L phase exhibits quite broad and

asymmetric diffraction peaks in Fig. 3b, 4b, and 5d. The broad and asymmetric peak profiles probably correspond to disordered staging domains including Daumas–Hérolde defects and/or stacking faults. The asymmetric peak profile is also reported for Li-GICs during Li (de)intercalation, and structural models including stacking faults are proposed for stage-2 LiC₁₂ (ref. 67) and stage-2L LiC₁₈.⁵¹

A distinct asymmetric peak profile for the $00n+1$ reflections of stage 2 and stage 1 phases and a high intensity baseline between the peaks are observed during the intercalation process (see Fig. S19†). However, in the subsequent deintercalation process from stage-1 KC₈, asymmetric peak-broadening is not obvious for both stage-1 and stage-2 peaks. The fact indicates that most of the Daumas–Hérolde defects disappear by complete filling of potassium atoms and formation of a single phase of 3D-ordered stage-1 KC₈ at the end of the intercalation process similar to Li-GICs in Fig. 6a. Stage-1 KC₈ has the ordered KC₈ in-plane unit of the $p(2 \times 2)R0^\circ$ superlattice and would transform into stage-2 KC₁₆ with retaining the in-plane ordering with the ordered KC₈ unit⁶⁵ as discussed in Fig. 5b, d, and f. Although the stacking sequence of stage-2 KC₁₆ could not be determined due to too low intensities of *hk0* Bragg reflections from the oriented graphite particles, our DFT calculations predict that the most stable stage-2 KC₁₆ has AAαAAα stacking, which is consistent with the previous calculated results.^{10,47} However, the formation energy is found to be quite close to that of KC₁₆ with ABβ'BAα stacking, and the energy difference is 1.54 meV (see Fig. 2h), indicating a possible mixture of AααAAα and ABβ'BAα stacking sequences as stacking faults. Indeed, a slight but certain peak-shift is observed for stage-1 and -2(L) K-GICs in our *operando* XRD data even during the deintercalation process (see the *operando* XRD pattern enlarged regions for stage-2(L) and -1 without intensity offset in Fig. S20 and S21,† respectively, and those in Li-GICs as a reference in Fig. S22†). The peak-shift for stage-2(L) during the deintercalation process is smaller than that during the intercalation process (see Fig. S20†), indicating the lower degree of Daumas–Hérolde defects during the deintercalation process.

Although the peak-shift or different peak-positions for stage-2(L) K-GICs were expected to originate from the transition between stage-2 KC₁₆ and stage-2L KC_{~24} phases as proposed by Liu,¹⁹ our calculated results predict negligibly small differences in peak position just by a $\Delta 2\theta$ value of 0.01° for both the $00n$ and $00n+1$ peaks ($l_c = 8.748 \text{ \AA}$, $2\theta_{00n} = 20.30^\circ$, and $2\theta_{00n+1} = 30.66^\circ$ for KC₁₆ and 8.744 \AA , 20.31° , and 30.67° for KC₂₄). The average interlayer distance of stage-2 KC₁₆ is about the same as but slightly larger than that of stage-2 KC₂₄, which is consistent with Liu's computational results.¹⁹ However, the simulated XRD pattern of stage-2 KC₁₆ presents lower-angle $00l$ reflections than those of stage-2 KC₂₄ and cannot explain the peak shift in the opposite direction of $00n$ and $00n+1$ reflections for the stage-2(L) K-GICs in Fig. S20.† Interestingly, this is different from the stage-2 Li-GICs, in which only a lower-angle shift of $00l$ reflections is experimentally confirmed for the transition from stage-2 LiC₁₂ to stage-2L LiC₁₈ as a two-phase reaction.^{39,50,51} Since the FWHM of $003_{\text{stage-2(L)}}$ reflection for stage-2(L) K-GIC becomes smaller through the peak shift from



30.84° to 30.66° during the potassium deintercalation process (see Fig. S20†), the peak-shift and -broadening are probably due to stacking faults such as unusual AB|BA|, AC|CA|, and BC|CB| sequences proposed for chemically prepared stage-2L $KC_{\sim 24}$.⁷³

Similar to the case of stage-2 K-GIC, stage-1 KC_8 reflections show a slight peak shift and intensity-gain and -reduction at the beginning of the deintercalation process in our *operando* XRD data (Fig. S21†) as well as in the reported data.^{18,19} Such a phenomenon is also confirmed for stage-1 Li-GICs during the beginning of lithium deintercalation in Fig. 5a and S21† as well as in previous results.^{51,74} As variation of long-range ordering along the stacking direction was proposed for the stage-1 Li-GIC domain,⁵¹ formation of stacking faults is expected for the stage-1 K-GIC domain during potassium (de)intercalation similar to the above mentioned stage-2(L). Furthermore, stage-1 KC_8 with $A\alpha A\beta A\alpha A\beta$ stacking was experimentally observed as an intermediate phase between the $A\alpha A\beta A\gamma A\delta$ stacked stage-1 KC_8 and the AB|BC|CA| stacked stage-2L in the *in situ* high-temperature XRD patterns of K-GICs.⁷⁵ Stacking faults in a potassium stacking sequence might contribute to the peak shift and intensity variation.

Although a slight peak shift and intensity change are found for stage-2(L) and -1 as mentioned above, a distinct difference between potassium intercalation and deintercalation processes is confirmed in FWHM values (Fig. 5). The FWHM values of stage-2(L) during potassium deintercalation are smaller than those during intercalation. The fact reveals no or less Daumas-Hérolde defects in the 2D or 3D ordered stage-2 KC_{16} phase during the potassium deintercalation process like in a 3D ordered stage-2 LiC_{12} during a lithium deintercalation process.⁵¹ Furthermore, the maximum peak-top intensity is obtained at the KC_{16} composition during the defect-less early deintercalation process (Fig. 5d), suggesting electrochemical formation of thermodynamically stable stage-2 KC_{16} at room temperature.

The subsequent potassium deintercalation from stage-2 KC_{16} causes critical phase transitions between even and odd indices of stage-*n*; *i.e.* stage 2 to stage 3L, again. Daumas-Hérolde defects appear between the stage domains as confirmed by large FWHM values in Fig. 5f. Formation of dilute stage-2L KC_{24} - KC_{28} was expected from the literature on chemically prepared phases,⁶⁵ but the peak intensity of the stage 2 phase has an inflection point at the KC_{16} composition and linearly decreases by potassium deintercalation (Fig. 5d), resulting in the formation of stage-3L KC_{24} as a two-phasic reaction. This transition sequence is consistent with Luo's¹⁰ and our DFT calculations results. Further potassium deintercalation from stage-3L KC_{24} leads to the transition from stage 3L to 4L with Daumas-Hérolde defects. Although these phases have a "liquid-like" in-plane potassium distribution, a commensurate ($\sqrt{7} \times \sqrt{7}$) $R19.11^\circ$ superlattice and/or incommensurate in-plane structures, depending on the in-plane potassium densities, might be locally formed (Fig. 6b) as reported for chemically prepared K-GICs at low temperature.^{62,63,76-81} The presence of stage 4L is evident in the *operando* XRD patterns in which the diffraction peaks were deconvoluted into stage 3L and 4L during the deintercalation process but not during the intercalation process

(see Fig. S17†). During the potassium intercalation process, stage-4L domains are formed inside the dilute stage 1 phase, which is island-domain-free and has large crystallites of the same size as graphite. On the other hand, during the deintercalation process, the pre-formed stage 3L phase has Daumas-Hérolde type island domains with smaller in-plane crystallite size. This might enhance the kinetics of the stage-3L to -4L transition during the deintercalation process. Finally, highly crystalline graphite is formed back through the formation of disorderly stacked high stage K-GICs as observed during the potassium intercalation process and discussed in Fig. 3d.

Electrochemical intercalation of the large ionic radius and low Lewis acidity potassium ion provides significantly different interlayer distances of graphene layers between empty and intercalated cases. Electrochemically formed K-GICs have unique characteristics such as staging disorder at the high stage, obvious reaction hysteresis during charge-discharge related to the phase-integration/-isolation of stage 4L with/from a disorderly stacked high stage, strong evidence of Daumas-Hérolde defects as a domain boundary between the even- and odd-index of stage-*n* K-GICs domains, and formation of stage-2 KC_{16} , which are successfully confirmed by *operando* XRD during potassium (de)intercalation into graphite and by DFT calculations in comparison with those of Li-GICs. Indeed, our *operando* XRD measurements were carried out under current flow (closed circuit) conditions, therefore, the phase evolution observed here is not under ideal equilibrium conditions. Also, the phase evolution of graphite during electrochemical intercalation would be influenced by materials and conditions, as discussed for the Li-GIC data⁵¹ with factors such as graphite crystallinity,³⁸ electrode compression,⁸² the current rate,⁵⁷ and temperature.¹⁴ In this study, the phase evolution for K-GICs as well as Li-GICs was investigated by using highly crystalline natural graphite and comprehensive analysis such as electrochemical methods, *operando* XRD at very slow current rates, and DFT calculations with consideration of the previous reports on chemically prepared K-GICs. Our findings on electrochemical potassium intercalation and corresponding phase evolutions are expected to contribute to the fundamental understanding of GICs as well as the further development of high-energy K-ion batteries.

Conclusions

The influence of the graphite particle size, binder, and electrolyte salts and solvents on the electrochemical potassium intercalation properties is systematically examined in non-aqueous K cells. Because we optimized the graphite electrode and electrolyte for higher reversibility intercalation, we successfully studied structural evolutions of graphite and K-GICs during electrochemical potassium intercalation into graphite by *operando* XRD and compared them with those of conventional Li-GICs ones. Both the results of *operando* XRD and DFT calculations reveal staging transformations from graphite to KC_8 through a disorderly stacked high stage, (stage 4L), stage 3L, 2L(2), and 1 phase without formation of dilute stage 1, which is observed in Li-GIC, and reversal transitions from stage 1 to graphite through stage 2, (2L), 3L, 4L, and



a disorderly stacked high stage in the K deintercalation process. The phase transition hysteresis corresponds to the voltage hysteresis in the galvanostatic charge–discharge curves of a K||graphite cell. Furthermore, the formation of stage-2 KC₁₆ was confirmed by both experimental and computational results. Based on the results, we propose structural change mechanisms of graphite during electrochemical potassium intercalation and deintercalation based on Daumas–Hérolde defects and their disappearance.

Author contributions

H. Onuma and S. Muratsubaki conducted the electrochemical and structural characterization. H. Onuma and K. Kubota analyzed the experimental data and prepared the manuscript. H. Onuma, S. Muratsubaki, and S. Yasuno conducted the HAXPES measurements. W. Ota, M. Shishkin, H. Sato, and K. Yamashita conducted the first-principles calculations. S. Komaba supervised the project and co-wrote the manuscript. All the authors discussed the results and contributed to writing the manuscript.

Conflicts of interest

There are no conflicts to declare.

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