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Guest size limitation in metal–organic framework crystal–glass composites†

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Metal–organic framework crystal–glass composites (MOF CGCs) have previously been formed by embedding crystalline MIL-53(Al) within a ZIF-62 glass (a_9 ZIF-62) matrix. Here we highlight thermal stability considerations in the formation of MOF CGCs, and subsequently report the synthesis of two novel MOF CGCs, by incorporating MIL-118 and UL-MOF-1 within a_9 ZIF-62. These new materials, alongside the prototypical MOF CGC, formed using MIL-53(Al), were studied using scanning electron microscopy, powder X-ray diffraction, and gas sorption techniques. The gas uptake in composites formed from MIL-118C and UL-MOF-1 is largely dominated by the a_9 ZIF-62 matrix, suggesting that to improve the porosity of the MOF CGC, the matrix porosity must be improved, or a percolation threshold must be overcome.

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Introduction

Metal–organic frameworks (MOFs) are hybrid materials containing inorganic nodes, or clusters, linked together by organic units in continuous networks.¹ These hybrid materials can be constructed using numerous combinations of metal ions—typically d-block elements—and polydentate ligands. Consequently, over 99 000 unique MOFs have been reported in the literature, many of which possess exceptionally high surface areas.² The hypothetical potential of such surface areas in MOFs has been predicted to reach $\sim 14\,600\text{ m}^2\text{ g}^{-1}$,^{3,4} though currently, the largest reported Brunauer–Emmett–Teller (BET) surface area is $7839\text{ m}^2\text{ g}^{-1}$. Accordingly, they are of interest for several applications, such as sensing, gas storage, and water harvesting.^{1,5–7}

A wide range of stimuli-responsive behaviour is observed by MOFs as a result of their chemical and physical diversity.⁸ Some frameworks, such as UiO-66, are commonly referred to as ‘rigid’ and display little, or no, structural change upon application of (non-extreme) external stimuli.⁹ Others, however, undergo reversible structural transformations under relatively mild pressure or temperature changes; these are referred to as flexible MOFs.¹⁰ Perhaps the most commonly illustrated example of

flexibility in MOFs is the phenomenon of “breathing”. An example of this occurs in the MOF $\text{Al}_2(\text{bdc})_3$ (bdc = benzene-1,4-dicarboxylate, $\text{C}_8\text{H}_4\text{O}_4^{2-}$)—referred to as MIL-53 from hereon—which undergoes a solid–solid phase transition from a narrow-pore phase at room temperature to a metastable large-pore phase upon desolvation at high temperature.¹¹ The breathing effect in MIL-53 dynamically changes the pore volume and, by extension, alters the guest uptake capacity. Understanding such transitions is important for designing materials for molecular separations, catalysis, and gas storage since the interaction strength between host and guest is dependent upon pore size.

Reversible breathing behaviour represents just a small portion of the structural rearrangements exhibited by MOFs; further examples include recrystallisation, interpenetrated lattice movements and irreversible collapse upon solvent/guest removal.¹² The most extensive of the observed rearrangements occur in a subset of MOFs known as zeolitic imidazolate frameworks (ZIFs), which contain an imidazolate-derived linker (Im , $\text{C}_3\text{H}_3\text{N}_2^-$), and adopt network topologies similar to those of inorganic zeolites. Heating under ambient pressure results in several ZIF structures undergoing melting to highly viscous liquid phases.¹³ These liquids then solidify upon cooling to form glasses in which tetrahedral Zn^{2+} ions are linked by imidazolate ligands in a continuous random network, analogous to that of amorphous silica.

The glass formed from ZIF-62, $\text{Zn}(\text{Im})_{1.75}(\text{bIm})_{0.25}$ (bIm, benzimidazolate, $\text{C}_7\text{H}_5\text{N}_2^-$)¹⁴ (melting temperature, $T_m = 437^\circ\text{C}$),¹⁵ has been identified as a suitable host matrix to secure crystalline MIL-53 particles within a bulk MOF structure. Specifically, a physical mixture of ZIF-62 and MIL-53 was heated to 450°C for 15 min to melt the ZIF-62 and allow sufficient

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liquid flow around the crystalline component before cooling to room temperature; the resultant material is referred to as a MOF crystal-glass composite (MOF CGC). Interestingly, within this MOF CGC, the large-pore phase of MIL-53 was stabilised at room temperature, inducing a significant increase in the amount of CO₂ adsorbed when compared to the uptake by a combination of the parent materials.¹⁶ The loading capacity of MIL-53 within the ZIF-62 glass (*a*_gZIF-62) matrix was also investigated, showing that 60–70 wt% MIL-53 could be encapsulated in the composite whilst retaining the MIL-53-lp state. As a result, a composite was fabricated with a CO₂ capacity considerably greater than the phase-pure MIL-53 parent material.¹⁷

To date, the MOF CGC formed from MIL-53 and *a*_gZIF-62 is the first of only three known examples of this class of materials.^{16–18} The extension of this approach to other crystalline MOFs is dependent upon their structural integrity at the temperatures required for the fabrication of the MOF CGC (*e.g.* 437 °C).¹⁵ This is further complicated by the lack of thermal stability data using standardised atmospheric conditions and heating rates, and furthermore by the lack of data on stability when held isothermally at elevated temperatures.¹⁹

In turn, the lack of additional MOF CGC chemistries limits any further conclusions on the gas sorption and separation behaviour of these materials, which is confined to N₂ and CO₂ adsorption on MIL-53 based samples only. Such considerations motivated us to provide further examples of MOF CGCs from different crystalline chemistries and architectures and evaluate their gas adsorption behaviour.

Results and discussion

Several MOF candidates for the crystalline filler were selected according to their reported thermogravimetric analyses (TGA) indicating the required stability to 450 °C (Table 1) and were synthesised using the published procedures (see Experimental, Fig. S1–S10†). Each of the eight frameworks selected was analysed using TGA conducted at a standard heating rate of 10 °C min^{−1} under inert nitrogen (N₂), to confirm thermal stability (Fig. S11†). The onset of thermal decomposition, *T*_d, below 450 °C was only observed for samples of DUT-6 and DUT-8. However, the *T*_ds of these two materials were greater than the reported values due to the faster heating rate used in our evaluation compared to those employed in the literature. As mass loss had already begun to occur at 450 °C, and given the need for isothermal heat treatment at this temperature for CGC fabrication, this precluded further study of these specific materials (Fig. S12†).

To provide a more accurate evaluation of thermal stability, the remaining MOFs were heated to 450 °C for 1 min, under N₂, and allowed to cool to room temperature *in situ*. Ambient temperature powder X-ray diffraction (PXRD) data were then recorded (Fig. S1–S10†). This relatively simple experiment highlighted the importance of performing thermal analysis of MOFs using an appropriate set of conditions, chosen according to the individual processing or application requirements. For example, Bragg peaks intensities in the PXRD pattern for MIL-

Table 1 Reported thermal stabilities of crystalline MOF component candidates^b

MOF	Composition	Decomposition temperature, <i>T</i> _d (°C)	Heating rate (°C min ^{−1})	Ref.
CUMOF-9	Yb ₂ (2,6-ndc) ₃ (H ₂ O)·(H ₂ O) ₂	550 ^a	5	20
DUT-6	Zn ₄ O(2,6-ndc)(btb) ₄ /3(def) ₁₆ (H ₂ O) _{9/2}	~400	5	21
DUT-8	Ni ₂ (2,6-ndc) ₂ (dabco)	~400	5	22
MIL-68	V(OH)(bdc)	~500	10	23
MIL-118	Al ₂ (OH) ₂ (C ₁₀ O ₈ H ₂)	~450	1	24
MIL-120	Al ₄ (OH) ₈ (C ₁₀ O ₈ H ₂)	~480	1	25
MIL-126(Sc)	Sc ₃ O(H ₂ O) ₂ (bpdc) ₃ X (X ~500 = OH or Cl)	~500	10	26
UL-MOF-1	Li ₂ (2,6-ndc)	610 ^a	10	27

^a Experiment performed under an inert atmosphere. ^b 2,6-ndc = naphthalene-2,6-dicarboxylate, bdc = benzene-1,4-dicarboxylate, bpdc = biphenyl-4,4'-dicarboxylate, btb = benzene-1,3,5-tribenzoate, def = *N,N*-diethylformamide, dabco = 1,4-diazabicyclo[2.2.2]octane.

68 (Fig. S2†) were reduced to near negligible levels, alongside a significant reduction in intensities, or changes in the patterns, for MIL-120 and MIL-126(Sc) (Fig. S3 and S4†). This is consistent with the thermal analysis performed by Volkringer *et al.*, confirming that the first step in decomposition for MIL-120 is due to ligand degradation.²⁵ An unreported recrystallisation of CUMOF-9 was also observed (Fig. S1†), though the high-temperature phase was not identified. MIL-68, MIL-120, MIL-126, and CUMOF-9 were, therefore, not studied further.

The PXRD patterns of the three remaining samples: MIL-118, UL-MOF-1, and MIL-53-np (Fig. 1a–d), displayed Bragg peaks in good agreement with their room temperature structures after heating to 450 °C (Fig. S5–S10†). UL-MOF-1 has been reported to display exceptional thermal stability (*T*_d = 610 °C); the structure consists of alternating two-dimensional antiferrotype type lithium oxide layers connected by 2,6-ndc struts (Fig. 1c).²⁷ No flexible behaviour has been reported for this material. Conversely, MIL-118, Al₂(OH)₂(C₁₀O₈H₂) (C₁₀O₈H₂^{4−} – benzene-1,2,4,5-tetracarboxylate), exhibits a water-driven reversible structural transition, analogous to that of MIL-53.²⁴ MIL-118 is typically synthesised with excess ligand in the pores of the framework and is named MIL-118A (*C2/c*). Upon heating MIL-118A, this excess ligand is removed, resulting in the open-pore framework, MIL-118B (*Pbam*), which is stable at high temperatures. Upon cooling, the framework adsorbs water to form the room temperature stable phase, MIL-118C (*Pnam*). The transition causes a shift from the rectangular 1-D tunnels in MIL-118B to lozenge-shaped channels, with water molecules occupying the pores in MIL-118C. Temperature-induced breathing is observed between the MIL-118C and MIL-118B phases (Fig. 1a and b).

MOF CGCs containing 50 wt% crystalline UL-MOF-1 and MIL-118 were synthesised using the method previously



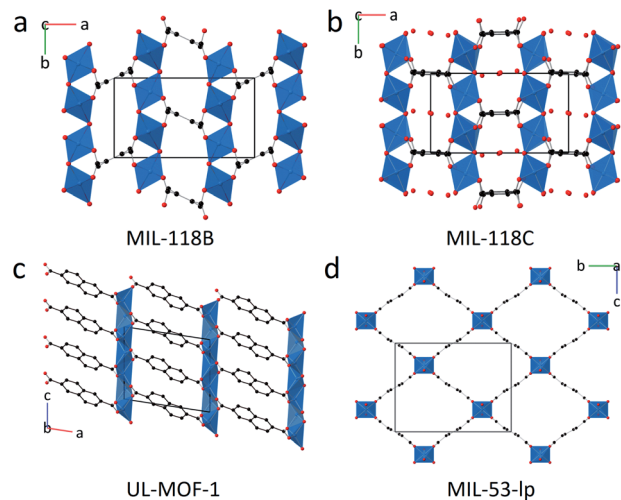


Fig. 1 Representations of (a) MIL-118B, (b) MIL-118C, (c) UL-MOF-1, (d) MIL-53-lp where; carbon – black, oxygen – red and M–O nodes replaced by blue polyhedra. Hydrogen atoms have been omitted for clarity. The unit cell is indicated by a black box with an overlay of crystallographic axes.

published in the literature for a CGC containing MIL-53. Briefly, the synthesis involved mixing the two crystalline components by ball-milling, pelletising the mixture, and heating to 450 °C under argon (for full procedure, see Experimental). These are referred to as (crystal)_x(glass)_{1–x} where x = weight fraction of the crystalline material in the composite, consistent with prior nomenclature. The MOF CGC behaviour was compared to a sample of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75}, synthesised according to previous methods.¹⁶

Ambient temperature PXRD data were recorded for the three MOF CGCs. As expected, Bragg peaks present in the crystalline sample of UL-MOF-1 were also present in (UL-MOF-1)_{0.5}(*a*_gZIF-62)_{0.5}. PXRD data for (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} demonstrated the presence of the open-pore MIL-53 phase, per previous results.^{16,17} Interestingly, the PXRD data for (MIL-118)_{0.5}(*a*_gZIF-62)_{0.5} indicated little change in lattice parameters from the synthesised MIL-118 sample, which remained in the low-temperature MIL-118C (*Pnam*) phase (Fig. 2 and Table S1†). This shows that, unlike the MIL-53 in (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75}, the MIL-118 encapsulated within the glass does not retain the high-temperature MIL-118B (*Pbam*) phase (Fig. 1a).

In our previous work, we demonstrated that the ingress of water did not drive a transition from MIL-53-lp back to MIL-53-np in a room temperature MOF CGC.¹⁶ It was, therefore, suggested that “the stabilisation of MIL-53-lp does not arise because it is excluded from entering the composite CGC material. Instead, we suggest that the polymeric phase is not soft enough to accommodate a large-scale change of the crystal phase structure, while the interfacial contact between the two phases is maintained”.¹⁶

Here, it is perhaps the case that either (i) interfacial contact between MIL-118 and the glass is less than that of MIL-53 and the glass, and/or (ii) that the volume changes upon the

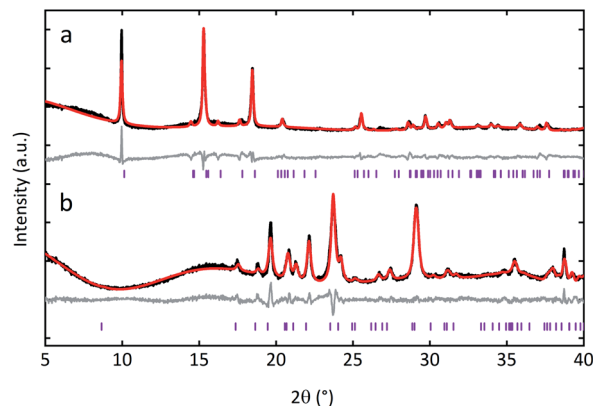


Fig. 2 PXRD data of (a) (MIL-118)_{0.5}(*a*_gZIF-62)_{0.5}, rwp: 11.152, and (b) (UL-MOF-1)_{0.5}(*a*_gZIF-62)_{0.5}, rwp: 7.677. Recorded data in black, Pawley refinements in red, difference patterns in grey, and *hkl* indices in purple underneath the respective patterns.

transition from MIL-118B to MIL-118C can be accommodated by the glass phase.

Scanning electron microscopy (SEM) images of MIL-53-np, MIL-118, and UL-MOF-1 display micrometre-sized crystals (Fig. S13–S15†). Samples of MIL-118 and MIL-53 display reasonable size and shape uniformity in comparison to UL-MOF-1, which comprises a range of morphologies from 10 μm cubic structures to 200 μm sheets (Fig. S15†). This occurs despite (i) X-ray diffraction phase purity which matches the reported crystallographic information, and (ii) a single thermal decomposition event at the reported temperature.

The MOF CGCs formed in each case demonstrated self-supporting, contiguous, bulk morphologies. SEM performed upon deliberately fractured pieces of these materials did not contain distinguishable remnants of the respective parent crystalline phases at the surface of the composite (Fig. 3 and S16–S18†). Their self-supporting nature, coupled with the smooth surface of the MOF CGCs, provides evidence that the liquid ZIF-62 permeated through the pellet. However, a region of coarse material matching the morphology of MIL-118 was observed at a macroscale surface defect in the sample of (MIL-118)_{0.5}(*a*_gZIF-62)_{0.5} (Fig. S17†).

Prior work on MOF CGCs has investigated only the gas uptake behaviour of a series of (MIL-53)_x(*a*_gZIF-62)_{1–x} CGCs with carbon dioxide (CO₂) and nitrogen (N₂).^{16,17} A range of analyte gases was therefore employed here to investigate the factors underpinning gas uptake in MOF CGCs. Gas sorption isotherms were performed predominantly at 273 K, with further experiments performed for some gases at 77 K and 293 K. For full experimental methodology, see Experimental.

A sample of pure *a*_gZIF-62 was prepared (see Experimental) and observed to exhibit porosity towards H₂ (1.05 mmol g^{–1} at 77 K, and 0.18 mmol g^{–1} at 273 K) and CO₂ (0.95 mmol g^{–1} at 273 K) (Fig. 4 and S19†), both significantly lower than crystalline ZIF-62 which adsorbs ~6.03 mmol g^{–1} H₂ at 77 K and ~1.79 mmol g^{–1} CO₂ at 273 K.²⁸ These results are in accordance with previous studies which demonstrate porosity in ZIF-glass materials.²⁹



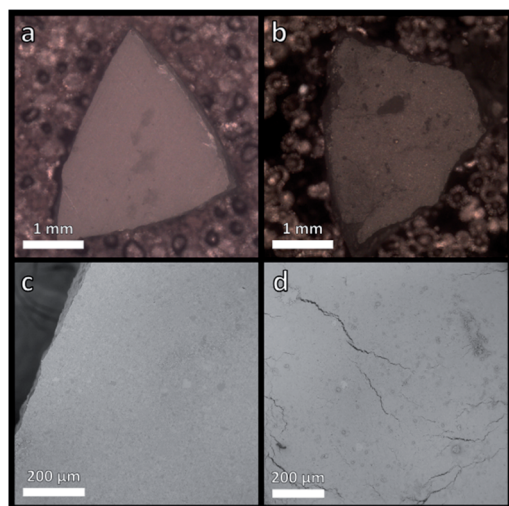


Fig. 3 Optical images of (a) $(\text{MIL-118})_{0.5}(\text{agZIF-62})_{0.5}$, and (b) $(\text{UL-MOF-1})_{0.5}(\text{agZIF-62})_{0.5}$, and scanning electron microscopy images taken of a fragment of (c) $(\text{MIL-118})_{0.5}(\text{agZIF-62})_{0.5}$, and (d) $(\text{UL-MOF-1})_{0.5}(\text{agZIF-62})_{0.5}$.

Methane (CH_4) adsorption for agZIF-62 was observed in this work at 0.21 mmol g^{-1} at 273 K (Fig. S19†), in agreement with the reported measurement of 0.18 mmol g^{-1} and representing a decrease from the reported value of 1.21 mmol g^{-1} in the crystalline ZIF-62.³⁰ Hysteresis is observed in the agZIF-62 isotherms as a result of the diffusion limitations through the amorphous structure, increasing in magnitude with the kinetic diameter of the adsorbent, also consistent with prior work.^{31,32} This is especially evident in the propene isotherm, with the uptake amount appearing to diminish in the region 70–100 kPa. This is an artefact of slow diffusion kinetics. The recorded adsorption/desorption points are not at equilibrium and are imperfectly estimated in the higher-pressure region.

Remarkably, in addition to the reported selectivity towards propane (C_3H_8) and propene (C_3H_6) by agZIF-62 by Frentzel-Beyme *et al.*,³¹ some selectivity towards ethene (C_2H_4) over ethane (C_2H_6) is observed here with uptakes of 0.46 and 0.29 mmol g^{-1} , respectively (Fig. 4). Industrially, the separation

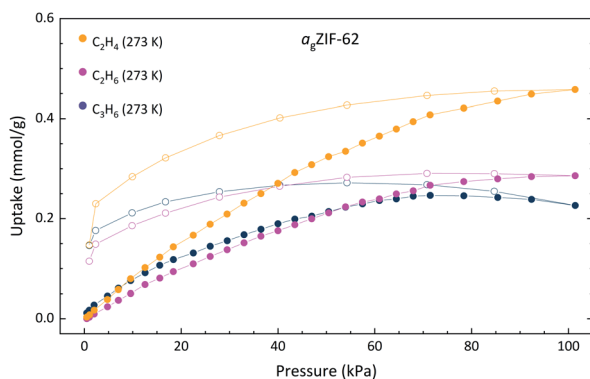


Fig. 4 Gas adsorption isotherms of agZIF-62 . Solid and open circles of the same colour indicate the adsorption and desorption respectively.

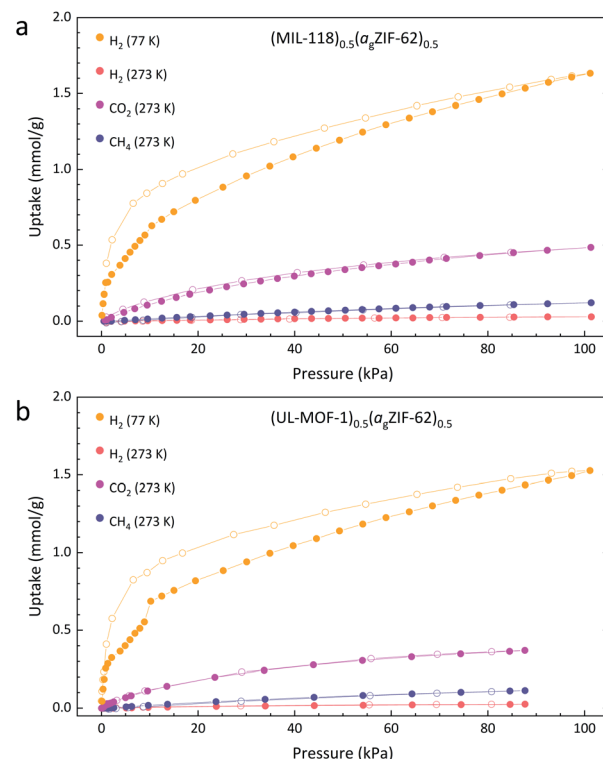


Fig. 5 Gas adsorption isotherms of samples of (a) $(\text{MIL-118})_{0.5}(\text{agZIF-62})_{0.5}$ and (b) $(\text{UL-MOF-1})_{0.5}(\text{agZIF-62})_{0.5}$. Solid and open circles of the same colour indicate the adsorption and desorption respectively.

of small hydrocarbons is currently performed using cryogenic high-pressure distillation processes and accounts for a large portion of global energy expenditure. Membrane-based separation of these materials is reportedly tenfold less energy-intensive, making development in this area both environmentally and economically desirable.³³

To understand how incorporation within agZIF-62 affects the gas sorption properties of the crystalline MOF, we compare the results for the pure crystalline material to that of the MOF CGC. Samples of MIL-118C and $(\text{MIL-118})_{0.5}(\text{agZIF-62})_{0.5}$ display identical trends in their adsorption isotherms of H_2 , CH_4 , and CO_2 , indicating no significant change in the chemical environment of MIL-118 upon inclusion in the MOF CGC (Fig. 5a and S20†). Since MIL-118C displays porosity to H_2 at 77 K, the negligible uptake of H_2 at 273 K is ascribed to temperature effects. However, the generally poor uptake capacity of MIL-118C is consistent with a dense atomic arrangement, in accordance with the published crystallographic information.²⁴

The adsorption of gases at 273 K by the 50 wt% composite material, $(\text{MIL-118})_{0.5}(\text{agZIF-62})_{0.5}$, is broadly comparable to a linear combination of the parent materials (Table 2). This, however, is not the case for H_2 adsorption at 77 K. Here, the recorded composite uptake is greater than either of its parent components and more than double the weighted average of the parent materials; this may suggest the presence of macroporous interfacial regions.

Identical experiments were performed on UL-MOF-1 and $(\text{UL-MOF-1})_{0.5}(\text{agZIF-62})_{0.5}$, and again, the generally poor



Table 2 Predicted and experimentally measured gas uptake of MOF crystal–glass composites based on the combination of their parent materials. Experimental temperatures of 273 K unless otherwise stated

Sample/gas	Crystalline MOF (mmol g ⁻¹)	<i>a</i> _g ZIF-62 (mmol g ⁻¹)	Weighted average (mmol g ⁻¹)	Measured MOF CGC (mmol g ⁻¹)
MIL-118				
Methane	0.016	0.212	0.114	0.122
Carbon dioxide	0.152	0.947	0.550	0.485
Hydrogen (77 K)	0.440	1.051	0.746	1.632
UL-MOF-1				
Methane	0.015	0.212	0.113	0.113
Carbon dioxide	0.051	0.947	0.499	0.370
Hydrogen (77 K)	0.320	1.051	0.686	1.527

adsorption to UL-MOF-1 is consistent with a dense structure (Fig. 5b and S20†). The same thermal effects observed in the MIL-118 H₂ isotherms at 77 K and 273 K were also observed for UL-MOF-1. The gas uptakes displayed by (UL-MOF-1)_{0.5}(*a*_gZIF-62)_{0.5} are broadly consistent with a linear combination of parent UL-MOF-1 and *a*_gZIF-62 materials (Table 2). Interestingly, the 77 K H₂ sorption isotherm for (UL-MOF-1)_{0.5}(*a*_gZIF-62)_{0.5} also exhibits a twofold increase in gas uptake over the weighted average of its parent materials.

The prediction of gas uptake through the weighted average of its components calculated here assume that the quantity of gas adsorbed in the MOF CGCs is due to the combined individual contributions from the two composited materials, where their adsorption is identical to the isolated components. It is important to note, however, that the hysteresis observed in the CO₂ isotherms for both *a*_gZIF-62 and the isolated crystalline MOFs is no longer present in the respective MOF CGCs. This suggests that the MOF CGC reaches equilibrium more rapidly than its isolated components, indicating that there may be new pathways for the gas to diffuse in the MOF CGC.

A CO₂ adsorption isotherm was also recorded at 273 K for the pure crystalline sample of MIL-53-np and is consistent with previously reported data (2.25 mmol g⁻¹ (recorded), and 2.13 mmol g⁻¹ (reported) (273 K), Fig. S21†).¹⁷ As expected, MIL-53-np displays a significantly higher uptake of C₃H₆ than C₂H₄, C₂H₆, or xenon (Xe), reaching 3.47 mmol g⁻¹ (273 K). The observation of two-step isotherms when using larger hydrocarbons is consistent with pore-opening behaviour where MIL-53-np expands to the MIL-53-lp phase, as comprehensively illustrated in the literature.¹¹

Consistent with our prior work, the MOF CGC exhibits appreciable porosity. However, due to fixation of the MIL-53-lp phase within the (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} composite, the pore opening behaviour of pure phase MIL-53 is no longer observed, and all recorded isotherms of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} display

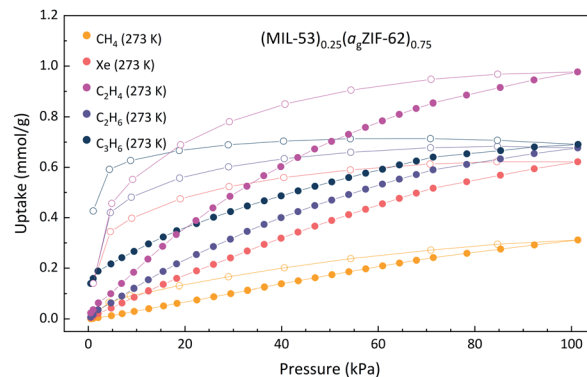


Fig. 6 Gas adsorption isotherms of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75}. Solid and open circles of the same colour indicate the adsorption and desorption respectively.

Langmuir-type behaviour, often with hysteresis in the desorption branch. The CO₂ sorption of a sample of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} displayed similar CO₂ uptake behaviour to that recorded previously (1.33 mmol g⁻¹ (recorded), and 1.14 mmol g⁻¹ (reported)).¹⁷ The sample of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} also displayed poor adsorption of both N₂ and CH₄ (<0.13 and <0.36 mmol g⁻¹ at 273 K, respectively), since neither component of the CGC strongly adsorbs these gases (Fig. S22†).

Since this sample, unlike (MIL-118)_{0.5}(*a*_gZIF-62)_{0.75} and (UL-MOF-1)_{0.5}(*a*_gZIF-62)_{0.75}, demonstrates a permeable crystalline component, further gas sorption experiments using larger gases were performed (Fig. 6). A considerable difference between the uptake of C₂H₄ and C₂H₆ is observed in a sample of (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75}, which is typical of the *a*_gZIF-62 component but is not observed in MIL-53. This result evidences the contribution of the adsorption properties of *a*_gZIF-62 to the overall composite characteristics.

Of the gases employed, the most striking result was that of C₃H₆, where the composite adsorbs far less than the combination of the parent materials. A weighted average C₃H₆ adsorption of MIL-53 (3.47 mmol g⁻¹) and *a*_gZIF-62 (0.23 mmol g⁻¹) for a 25 wt% MIL-53 mixture would be 1.04 mmol g⁻¹, where an uptake of only 0.69 mmol g⁻¹ is observed for the composite. However, the adsorption isotherm for C₃H₆ to (MIL-53)_{0.25}(*a*_gZIF-62)_{0.75} shows an increase in uptake on desorption, strongly indicating that the system had not reached equilibrium. This is likely due to poor diffusion of C₃H₆ through the *a*_gZIF-62 matrix as a result of the comparatively large kinetic diameter of the gas. Though the MIL-53 component in the composite would be the dominant contributor to the gas uptake, adsorption to this component is greatly hindered whilst encapsulated within a matrix that is poorly permeable to the analyte gas.

Conclusions

Thermal analysis of a selection of crystalline MOFs revealed that two: UL-MOF-1, and MIL-118, are suitable for inclusion with a MOF glass (*a*_gZIF-62) matrix. Despite reported TGA evidence to the contrary, several other crystalline MOFs were observed to



undergo partial or complete collapse at the processing temperatures required for composite formation. This highlights the necessity of detailed thermal characterisation and the avoidance of an over-reliance on constant-rate TGA experiments.¹⁹

It is clear from this research that to expand the scope of MOF CGCs to include frameworks with more diverse properties, new synthetic pathways must be explored. Two of such avenues may lie in (i) the utilisation of coordination polymers or MOFs with lower melting temperatures or (ii) exchanging the crystalline glass-forming MOF for a premade glass MOF. In doing so, the temperatures required for the fabrication of the MOF CGC may be reduced as the glass transition temperature (T_g) occurs at a lower temperature than the melting temperature (T_m) used here.^{34–37}

The adsorption behaviour of the resultant two new MOF CGCs is dominated by contributions from the a_g -ZIF-62 matrix, which prevents the diffusion of molecules with a kinetic diameter larger than that of C_3H_6 . The development of more permeable a_g -MOFs may aid in the expansion of MOF CGC applicability. Alternatively, work on overcoming percolation thresholds in MOF CGCs would mean guest diffusion is primarily controlled by the crystalline MOF component. This would allow further control of interactions leading to multi-functional materials with the ability to act as both molecular sieves and separators.

Experimental

Material synthesis

Cumof-9. Ytterbium(III) chloride hexahydrate (1.08 g, 2.79 mmol) and water (6.13 mL) was added to a Teflon lined autoclave and stirred for 2 min. 2,6-Naphthalenedicarboxylic acid (0.103 g, 0.476 mmol) and trimethylamine (0.111 mL) were added to the solution and the mixture was stirred for a further 5 min. The suspension was then sealed and placed in a 145 °C preheated oven for 4 h. The reaction vessel was allowed to cool to room temperature (RT) and the product was isolated by filtration under vacuum and washed with ethanol (30 mL) before drying in the oven at 70 °C overnight.

DUT-6. Zinc nitrate tetrahydrate (0.142 g, 0.543 mmol), 2,6-naphthalenedicarboxylic acid (0.017 g, 0.079 mmol), and benzene-1,3,5-tribenzoic acid (0.054 g, 0.123 mmol) were dissolved in *N,N*-diethylformamide (DEF, 10 mL). The mixture was sonicated for 5 min and heated to 100 °C for 24 h in a Pyrex tube. After cooling to RT the product was isolated by decanting the mother liquid and was washed with DEF. Yield: 49%. The resulting solid was immersed in dichloromethane (DCM) for 2 days. During this time the DCM was replaced with fresh DCM three times. Elemental analysis calculated (%) for $C_{128}H_{211}O_{33.5}N_{16}Zn_4$: C 55.5, H 7.67, O 19.34, N 8.09, Zn 9.44; found: C 56.2 ± 0.3 , H 7.66 ± 0.07 , O 19.95 ± 0.09 , N 8.01 ± 0.08 , Zn 9.22 ± 0.05 .

DUT-8(Ni). Nickel nitrate hexahydrate (0.407 g, 1.40 mmol) in 6 mL DMF, 2,6-naphthalenedicarboxylic acid (0.303 g, 1.40 mmol) in 15 mL DMF and 1,4-diazabicyclo[2.2.2]octane (0.100 g, 0.89 mmol) in 9 mL methanol (MeOH) were mixed.

Subsequently the mixture was transferred into a Teflon vessel (50 mL) and heated in an autoclave to 120 °C at a heating rate of 4 °C min⁻¹ and held at that temperature for 48 h. After cooling to RT over 3 h, the sample was washed three times, first with 30 mL *N,N*-dimethylformamide (DMF) and then with 30 mL ethanol and finally with 30 mL DCM. Afterwards, a washing step with 150 mL of DCM was continued for 3 days. The resulting solid was filtered in an argon flow and activated under dynamic vacuum at 393 K for 4 h.

MIL-53. Aluminum nitrate nonahydrate (26 g, 6.93×10^{-2} mol) and terephthalic acid (5.76 g, 4.96×10^{-2} mol) were dissolved in water (100 mL) and placed into a Teflon-lined autoclave, and then placed in an oven at 220 °C for 72 h. The resulting powder was washed with deionised water (3 × 30 mL) and dried in a vacuum oven at 150 °C for 24 h. MIL-53 was activated by heating at 330 °C for 72 h, and then to 450 °C for 6 min before cooling to RT.

MIL-68. Aluminum nitrate nonahydrate (1.4 g, 6.57 mmol), benzene-1,4-dicarboxylic acid (0.6 g, 3.61 mmol), and tetrahydrofuran (THF, 16 mL) were added to a 50 mL round bottom flask and refluxed for 3 days at 70 °C. The resultant white powder was washed with THF (3 × 15 mL) and dried at RT overnight. To remove excess linker from the pores, the product was activated by heating to 300 °C at 1 °C min⁻¹ and held for 8 h.

MIL-118. Aluminum nitrate nonahydrate (150 mg, 7.04×10^{-4} mol) and benzene-1,2,4,5-tetracarboxylic acid (50 mg, 1.97×10^{-4} mol) were added to a Teflon lined autoclave before adding water (5 mL). The autoclave was sealed and placed into a 210 °C preheated oven for 24 h. The product of this was isolated by replacing the liquid with water (20 mL) and centrifuging (2500 rpm, 10 min) twice. The resultant white powder was placed in a 70 °C preheated oven overnight.

MIL-120. Aluminium nitrate nonahydrate (1.44 g, 6.76 mmol), 1,2,4,5-benzenetetracarboxylic acid (0.225 g, 0.885 mmol), sodium hydroxide (4 M, 1.53 mL), and water (9 mL) were added to a Teflon lined autoclave and placed in a 210 °C preheated oven and held for 24 h. The resulting white powder was filtered and washed with water (40 mL) before being transferred to a 100 mL round-bottom flask with water (60 mL), and refluxed at 100 °C for 10 h. The product was centrifuged (2500 rpm, 15 min) and dried in the oven at 70 °C overnight. Further to the reported method, MIL-120 was heated to 280 °C to produce a white powder with a PXRD pattern matching the predicted pattern.

MIL-126(Sc). Scandium nitrate hydrate (85 mg, 0.368 mmol), biphenyl-4,4'-dicarboxylic acid (68 mg, 0.281 mmol), DMF (6.25 mL), conc. hydrochloric acid (37%, 25 μL) were added to a 25 mL screw-top jar and sonicated until dissolution (~15 min). The clear solution was placed into a 120 °C preheated oven for 24 h. The product was transferred to a centrifuge tube and centrifuged (2500 rpm, 15 min) to collect a white powder. This powder was stood in DMF (10 mL) overnight and centrifuged again (2500 rpm, 10 min) and acetone exchanged three times over three days before collecting the product by centrifuging (2500 rpm, 5 min), decanting the liquid and drying in a 60 °C preheated oven overnight.



ULMOF-1. Lithium nitrate (0.345 g , $5.00 \times 10^{-3}\text{ mol}$), naphthalene-2,6-dicarboxylic acid (0.565 g , $2.61 \times 10^{-3}\text{ mol}$), ammonium fluoride (38 mg), and DMF (15 mL) were added to Teflon lined autoclave and placed in a $180\text{ }^{\circ}\text{C}$ preheated oven and held for 5 days. Upon cooling, the reaction mixture was transferred to a centrifuge tube and the liquid was replaced with ethanol (20 mL). The sample was stirred for 5 min before centrifuging (3000 rpm , 5 min) to collect a white powder which was dried in an oven at $60\text{ }^{\circ}\text{C}$ overnight.

ZIF-62. Zinc nitrate hexahydrate (1.65 g , $5.54 \times 10^{-3}\text{ mol}$) and imidazole (8.91 g , 0.13 mol) were added to a 200 mL screw-top jar, dissolved in DMF (75 mL) and stirred for 1 h. Once complete dissolution was achieved, benzimidazole (1.55 g , $1.31 \times 10^{-2}\text{ mol}$) was added and heated to $130\text{ }^{\circ}\text{C}$ for 48 h. The product was allowed to cool to RT and crystals were separated by vacuum-assisted filtration and washed with DMF (40 mL) and DCM (40 mL) before being dried in an oven at $60\text{ }^{\circ}\text{C}$ overnight.

***a*_gZIF-62.** 150 mg of ZIF-62 was heated to $450\text{ }^{\circ}\text{C}$ in a furnace at a rate of $10\text{ }^{\circ}\text{C min}^{-1}$ under an argon (Ar) atmosphere, and held for 15 min before being allowed to cool to RT.

MOF CGC materials. ZIF-62 and the corresponding crystal-line material were ball-milled together using a Retsch MM400 instrument, in appropriate wt% ratios using a 7 mm -diameter stainless steel ball for 15 min, at a frequency of 30 Hz . The mixed powder was pressed in a 13 mm -diameter die at 0.74 GPa for 1 min. The pellet was then clamped between glass slides, heated to $450\text{ }^{\circ}\text{C}$ in a tube furnace at a rate of $20\text{ }^{\circ}\text{C min}^{-1}$ under an Ar atmosphere, and held for 15 min before being allowed to cool to RT.

Characterisation

Powder X-ray diffraction. Data were collected on ground samples of the composite materials with a Bruker D8 Advance powder diffractometer using $\text{Cu K}\alpha$ radiation ($\lambda = 1.5418\text{ \AA}$) and a LynxEye position-sensitive detector in Bragg–Brentano (θ – θ) parafocusing geometry at RT. Diffraction patterns were recorded at 2θ values of 5 – 40° with a time/step of 0.75 s over 1724 steps through a 0.012 mm Ni filter.

Scanning electron microscopy. SEM was performed using a Thermo Scientific™ Phenom ProX scanning electron microscope. Powder and monolithic samples were prepared for SEM by securing to aluminum SEM pin stubs using carbon tape. Samples were coated in gold using an Emtech K575 sputter coater.

Thermogravimetric analysis. Thermogravimetric analysis was performed using a TA Q500 TGA. All scans were performed at $10\text{ }^{\circ}\text{C min}^{-1}$ with a nitrogen protective gas and allowed to cool to RT with air.

Gas sorption. Samples were degassed overnight at the specified temperature for 12 h on before transferring to the analysis port of a Quantachrome iQ2 instrument. Sample weight was measured post-degas activation. Sample temperature was accurately equilibrated at 273 K and 293 K with a temperature-controlled water bath and at 77 K with liquid N_2 . Gas adsorption measurements were performed using ultra-high purity (99.99%) gases.

Author contributions

C. W. A. and T. D. B. designed the project with assistance from S. G. T. Materials were synthesised by C. W. A. with assistance from Y. Z. (*a*_gZIF-62), and S. K. (DUT-6 and DUT-8). C. W. A. recorded PXRD and TGA data and performed Pawley refinements. T. J. F. S. and C. W. A. prepared samples and recorded SEM images. Gas sorption experiments were performed by L. K. M., S. G. T., C. W. A., and Y. Z. The manuscript was written by all authors.

Conflicts of interest

There are no conflicts to declare.

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