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# Is electrochemical CO<sub>2</sub> reduction the future technology for power-to-chemicals? An environmental comparison with H<sub>2</sub>-based pathways†

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Power-to-chemical pathways are gaining attention as a promising approach to reduce greenhouse gas (GHG) emissions of the chemical industry. Most power-to-chemical pathways use water electrolysis and hydrogenation of carbon dioxide (H<sub>2</sub>-based pathways). An emerging power-to-chemical technology is electrochemical CO<sub>2</sub> reduction (eCO<sub>2</sub>R) that converts carbon dioxide (CO<sub>2</sub>) and water (H<sub>2</sub>O) directly into chemical products. An early assessment of the environmental potential of eCO<sub>2</sub>R can accelerate the eCO<sub>2</sub>R development toward more sustainable technology. So far, only a few environmental assessments have been conducted using fixed assumptions of eCO<sub>2</sub>R operating parameters. To overcome this limitation and reflect all possibilities of future eCO<sub>2</sub>R development, we design a parameter-dependent approach using a variable eCO<sub>2</sub>R model. The considered eCO<sub>2</sub>R model can produce carbon monoxide (CO), methanol, or ethylene. Together with the conversion of syngas (CO and hydrogen mixture) to methanol and with the methanol-to-olefins (MtO) conversion to ethylene, five different eCO<sub>2</sub>R pathways to produce methanol and ethylene are possible. By applying our parameter-dependent approach, we determine the minimum development requirements for the eCO<sub>2</sub>R pathways to achieve climate benefits over the H<sub>2</sub>-based pathways. By comparing the obtained minimum development requirements to data from experimental studies, we evaluate potential development gaps. The direct eCO<sub>2</sub>R to ethylene pathway shows a moderate development gap and could reduce GHG emissions over the H<sub>2</sub>-based pathway by up to 44%. The eCO<sub>2</sub>R *via* CO to methanol and eCO<sub>2</sub>R *via* CO to ethylene pathways can already fulfill the minimum development requirements; however, the GHG reduction potential is limited to 12%. The direct eCO<sub>2</sub>R to methanol pathway shows a significant development gap. Compared to the fossil fuel-based technologies, both the H<sub>2</sub>-based and the eCO<sub>2</sub>R pathways offer large climate benefits of up to 88% and up to 93%, respectively. The presented parameter-dependent assessment offers an in-depth understanding of the eCO<sub>2</sub>R operating parameters and provides a valuable approach for future assessments of the emerging eCO<sub>2</sub>R technology.

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## 1. Introduction

The urgency of reducing anthropogenic greenhouse gas (GHG) emissions has intensified in all sectors. The chemical industry is one contributor with about 6% of worldwide GHG emissions.<sup>1</sup> Hence, more low-carbon technologies to produce chemicals need to be developed. Many researchers focus on bulk chemicals due to their quantitative significance.<sup>2</sup> Two bulk chemicals with large market volumes are methanol (145 mmt a<sup>-1</sup> global annual capacity<sup>3</sup>) and ethylene (185 mmt a<sup>-1</sup> global annual capacity<sup>3</sup>). Methanol (CH<sub>3</sub>OH) serves as a universal feedstock

for chemical syntheses and as an intermediate energy carrier for synthetic fuels, while ethylene (C<sub>2</sub>H<sub>4</sub>) is an essential material for most plastics.

In addition to biomass-based and photocatalytic syntheses, power-to-chemical is a promising technology in a future low-carbon energy system.<sup>4</sup> All power-to-chemical technologies rely on a highly decarbonized electricity sector to be able to reduce GHG emissions.<sup>4</sup> Most power-to-chemical technologies use hydrogen as an intermediate (H<sub>2</sub>-based). Hydrogen (H<sub>2</sub>) can be synthesized electrochemically *via* proton exchange membrane (PEM) water electrolysis. H<sub>2</sub> is subsequently reacted to synthesize chemicals by hydrogenation of carbon dioxide (CO<sub>2</sub>).<sup>5-7</sup> In recent years, a new power-to-chemical technology has drawn attention: electrochemical CO<sub>2</sub> reduction (eCO<sub>2</sub>R). The eCO<sub>2</sub>R converts CO<sub>2</sub> and water (H<sub>2</sub>O) directly into valuable products in an electrochemical cell. Four key operating

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parameters can measure the performance of the eCO<sub>2</sub>R cell: cell voltage, faradaic efficiency, single-pass CO<sub>2</sub> conversion, and current density. The cell voltage provides the electrochemical potential for the chemical conversions. The lower limit of the cell voltage is the equilibrium voltage; however, excess voltage is needed to enforce the electrochemical conversion.<sup>2</sup> The faradaic efficiency defines the ratio of electric charge converted into the desired product to the total applied electric charge. Hence, the higher the faradaic efficiency, the better is the selectivity toward the desired product. The single-pass CO<sub>2</sub> conversion reflects that part of the inlet CO<sub>2</sub> leaves the cell unreacted. The current density states the electrical current per cell area and can indicate economic feasibility since increasing current density reduces the specific investment costs of an electrochemical cell.<sup>8</sup>

Theoretically, both eCO<sub>2</sub>R and H<sub>2</sub>-based technologies could produce chemicals independent from fossil resources using only electricity, CO<sub>2</sub>, and H<sub>2</sub>O. Low-carbon electricity is needed for both technologies to be able to reduce GHG emissions of the chemical industry; however, low-carbon electricity is currently limited and is expected to remain limited in the future due to an increasing demand by electrification ambitions in all sectors.<sup>4</sup> With low-carbon electricity being limited, eCO<sub>2</sub>R and H<sub>2</sub>-based technologies are competing power-to-chemical technologies. While PEM electrolyzers are already available at industrial scale (1 MW),<sup>5</sup> practical realization of eCO<sub>2</sub>R still requires development as its technology readiness level ranges from 2–4.<sup>9</sup> Laboratory eCO<sub>2</sub>R setups have investigated catalyst behavior,<sup>10</sup> electrolyte interactions,<sup>11</sup> and general setup optimization.<sup>12</sup> However, several technical challenges have to be overcome to

reach industrial feasibility, such as lowering cell voltages, increasing faradaic efficiency, enhancing CO<sub>2</sub> conversion, and increasing current densities.<sup>2,10</sup>

To accelerate the development of eCO<sub>2</sub>R, it is desirable to evaluate the technological prospect of eCO<sub>2</sub>R as soon as possible to identify the most promising approaches and potential development gaps. For this purpose, previous studies<sup>13–15</sup> analyzed future eCO<sub>2</sub>R technologies using estimations for the operating parameters based on other alkaline electrolyzers. The varying estimated parameters and future developments are summarized in Table 1.

Adnan and Kibria<sup>13</sup> conducted a comparative techno-economic and environmental assessment of three power-to-methanol technologies, including an H<sub>2</sub>-based pathway, a direct eCO<sub>2</sub>R pathway to methanol, and an eCO<sub>2</sub>R pathway to carbon monoxide (CO) with subsequent methanol synthesis from syngas. They estimated the electrochemical cells to operate at cell voltages from 3.5 V to 7.4 V and faradaic efficiencies from 80% to 100%. A cradle-to-gate life cycle analysis showed that the H<sub>2</sub>-based pathway leads to the lowest GHG emissions for methanol production, followed by the eCO<sub>2</sub>R to CO pathway.

Guzmán *et al.*<sup>14</sup> compared the GHG emissions of eCO<sub>2</sub>R to methanol with an H<sub>2</sub>-based pathway. The estimated faradaic efficiency to methanol ranges from 32.7% to 89.3%. Their comparative life cycle assessment concluded that eCO<sub>2</sub>R and H<sub>2</sub>-based pathways emit equal GHG emissions in the most optimistic scenario, while the H<sub>2</sub>-based pathway is preferable in all other scenarios. Furthermore, Guzmán *et al.* reported a considerable electricity demand for compression and

**Table 1** Development estimations of comparative studies on eCO<sub>2</sub>R and H<sub>2</sub>-based power-to-chemical pathways. The considered eCO<sub>2</sub>R products were carbon monoxide (CO), methanol, and ethylene

| Author                                    | Power-to-chemical | eCO <sub>2</sub> R product | Estimated $U_{\text{cell}}$               | Estimated faradaic efficiency             | Estimated CO <sub>2</sub> conversion (single-pass) | Separation equipment                             | Lowest GHG emissions: eCO <sub>2</sub> R vs. H <sub>2</sub> -based? |
|---|-------------------|----------------------------|---|---|--|--|---|
| Adnan and Kibria (2020) <sup>13</sup>     | Power-to-methanol | CO                         | 7.4 V                                     | 100%                                      | 50%  | None   | H <sub>2</sub> -based   |
| Adnan and Kibria (2020) <sup>13</sup>     | Power-to-methanol | CO                         | 3.7 V                                     | 100%                                      | 50%  | None   | H <sub>2</sub> -based   |
| Adnan and Kibria (2020) <sup>13</sup>     | Power-to-methanol | Methanol                   | 4.4 V                                     | 80%                                       | 50%  | Electrolyte recirculation, distillation, and PSA | H <sub>2</sub> -based   |
| Adnan and Kibria (2020) <sup>13</sup>     | Power-to-methanol | Methanol                   | 2.5 V                                     | 90%                                       | 50%  | Electrolyte recirculation, distillation, and PSA | H <sub>2</sub> -based   |
| Guzmán <i>et al.</i> (2020) <sup>14</sup> | Power-to-methanol | Methanol                   | n.a.                                      | 32.7%                                     | n.a.   | Flash and distillation                           | H <sub>2</sub> -based   |
| Guzmán <i>et al.</i> (2020) <sup>14</sup> | Power-to-methanol | Methanol                   | n.a.                                      | 89.3%                                     | n.a.   | Flash and distillation                           | Equal   |
| Nabil <i>et al.</i> (2021) <sup>15</sup>  | Power-to-methanol | CO                         | n.a. (total 4.8 kW h kg <sup>-1</sup> CO) | n.a. (total 4.8 kW h kg <sup>-1</sup> CO) | 50%  | Pressure swing adsorption                        | eCO <sub>2</sub> R  |
| Nabil <i>et al.</i> (2021) <sup>15</sup>  | Power-to-methanol | Methanol                   | 2.6 V (1.4 V overpotential)               | 90%                                       | 50%  | Pressure swing adsorption and distillation       | H <sub>2</sub> -based   |
| Nabil <i>et al.</i> (2021) <sup>15</sup>  | Power-to-ethylene | Ethylene                   | 2.55 V (1.4 V overpotential)              | 90%                                       | 50%  | Pressure swing adsorption                        | eCO <sub>2</sub> R  |



pumping with 28% of the total electricity consumption for a low faradaic efficiency of 32.7%.

The most current study published by Nabil *et al.*<sup>15</sup> evaluated the global warming impact of eight eCO<sub>2</sub>R products, including CO, methanol, and ethylene. Methanol and ethylene were produced with a membrane electrode assembly cell operating at a faradaic efficiency of 90% and cell voltages of 2.6 V and 2.55 V, respectively. For eCO<sub>2</sub>R to CO, a solid-oxide electrolysis cell was considered with an electricity demand of 4.8 kW h kg<sup>-1</sup> CO; the corresponding cell voltage and faradaic efficiency were not reported. The energy demand of liquid product separation by distillation contributes to 62% of the total energy demand of eCO<sub>2</sub>R to methanol. In contrast, gaseous product separation (*i.e.*, CO and ethylene) has a substantially low impact due to an efficient pressure swing adsorption. Consistent with previous studies, Nabil *et al.*'s assessment reported that eCO<sub>2</sub>R to methanol is not environmentally favorable compared to the H<sub>2</sub>-based pathway in terms of GHG emissions. For eCO<sub>2</sub>R to CO and ethylene, they concluded climate benefits of the eCO<sub>2</sub>R pathway over the H<sub>2</sub>-based pathway.

All three studies rely on fixed estimations regarding eCO<sub>2</sub>R technology development and the future operating parameters.<sup>13–15</sup> Hence, the comparative assessments of eCO<sub>2</sub>R and H<sub>2</sub>-based pathways are only valid for the specific parameter estimations that differ between the studies (*cf.* Table 1). To provide a more comprehensive assessment, a comparative study including all possible parameters is desirable. For this purpose, we consider a variable process inventory for eCO<sub>2</sub>R based on an electrochemical model. Thereby, we can evaluate the entire range of eCO<sub>2</sub>R operating parameters in a parameter-dependent assessment (Section 2.2). This parameter-dependent assessment allows us to derive minimum development requirements for eCO<sub>2</sub>R technologies to achieve equal or lower GHG emissions than H<sub>2</sub>-based pathways (Sections 4.2 and 4.3). We compare the minimum development requirements with the state-of-the-art experimental data to examine the development gaps. By applying our parameter-dependent assessment, we provide a deeper understanding of the technical development requirements for eCO<sub>2</sub>R to achieve climate benefits compared to H<sub>2</sub>-based power-to-chemicals pathways.

This article is structured in four sections. Section 2 introduces the applied methods, followed by presenting the eCO<sub>2</sub>R model and the chemical system in Section 3. In Section 4, we illustrate the results of the eCO<sub>2</sub>R model and chemical system. Section 5 discusses our results and their limitations. In Section 6, we conclude and summarize the findings of this study.

## 2. Methods

In this section, we introduce the methodological framework to identify optimal production pathways in a parameter-dependent assessment.

### 2.1. Technology choice optimization

In the present study, we investigate technological pathways within a chemical system that consists of many individual

production processes. We use a technology choice optimization to find the production pathway with the lowest global warming impact.<sup>16</sup> The technology choice optimization solves a linear programming problem according to eqn (1)–(3).

$$\min_s \text{GWI} = q_{\text{GWP}} B s \quad (1)$$

$$\text{s.t. } A s = f \quad (2)$$

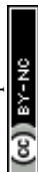
$$s_j \geq 0 \quad (3)$$

The notation of process inventory follows the generalized calculus for life cycle assessment.<sup>17,18</sup> The technology matrix *A* describes process exchanges within the chemical system, *e.g.*, material and energy balances of each process. Each column of the technology matrix *A* contains the process inputs that are defined as negative values, and the process outputs that are positive (*cf.* ESI 3†). The desired outcome of the chemical system is set by the functional unit *f*; here, 1 kg of methanol or 1 kg of ethylene. The scaling vector *s* expresses the usage of each process. The scaling vector *s* can only be positive since the processes cannot be inverted (eqn (3)). All solutions of eqn (2) define a valid technology choice of the chemical system to supply the functional unit *f*. The elementary flow matrix *B* contains information about interactions with the environment, *e.g.*, direct process emissions (*cf.* ESI 3†). The characterization vector *q*<sub>GWP</sub> weights each elementary flow by its global warming potential using ReCiPe 2016.<sup>19</sup> Thus, the objective function (eqn (1)) finds the technology choice with the lowest global warming impact, *i.e.*, the most climate friendly production pathway.

### 2.2. Parameter-dependent assessment

Most life cycle assessment studies using technology choice optimization use static process inventories. However, for our investigation of the emerging eCO<sub>2</sub>R technology, a static process inventory cannot reflect the process development. Therefore, we complement a static process inventory for the state-of-the-art chemical system with a variable process inventory for the eCO<sub>2</sub>R process. First, we identify the key operating parameters of eCO<sub>2</sub>R. We discretize the range of eCO<sub>2</sub>R development with a resolution of 50 steps per eCO<sub>2</sub>R parameter within a reasonable range. Second, we develop an eCO<sub>2</sub>R model to gain a variable process inventory. The non-linear eCO<sub>2</sub>R model determines the eCO<sub>2</sub>R process exchanges for each parameter combination. Third, each parameter combination of the variable eCO<sub>2</sub>R process inventory is combined with the static inventory of the chemical system and fed to the technology choice optimization (*cf.* Fig. 1). The linear technology choice optimization finds the production pathway with the lowest GHG emissions for the entire pattern of discretized eCO<sub>2</sub>R parameters (*i.e.*, 50 × 50 resolution refers to 2.500 technology choice optimization runs per pathway). Thereby, we conduct a parameter-dependent assessment that evaluates the entire range of eCO<sub>2</sub>R development.

The production pathway is either static using only state-of-the-art processes from the static chemical system or an eCO<sub>2</sub>R



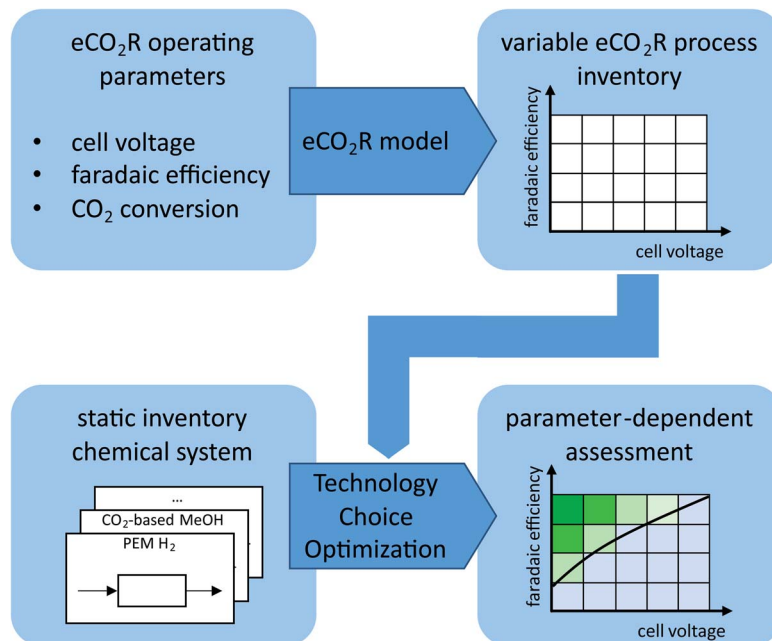


Fig. 1 Methodological approach of the parameter-dependent assessment. The static process inventory contains the state-of-the-art chemical processes, while the variable process inventory reflects the eCO<sub>2</sub>R.

pathway using the variable eCO<sub>2</sub>R process inventory. If a static production pathway is favorable compared to the variable eCO<sub>2</sub>R pathway, GHG emissions of the optimal production pathway (which does not include any variable eCO<sub>2</sub>R processes) are independent from the variable eCO<sub>2</sub>R operating parameters (*cf.* Fig. 1: parameter-dependent assessment at high cell voltage and low faradaic efficiency). If the eCO<sub>2</sub>R pathway is favorable, the parameter-dependent assessment shows the GHG emissions depending on the eCO<sub>2</sub>R operating parameters (*cf.* Fig. 1: parameter-dependent assessment at low cell voltage and high faradaic efficiency). With our methodological approach, the minimum development requirements for eCO<sub>2</sub>R can be derived that are necessary to compete against state-of-the-art H<sub>2</sub>-based power-to-chemical pathways using water electrolysis and CO<sub>2</sub> hydrogenation.

### 3. Model

The model of this study consists of two parts with different levels of detail: the chemical system and the more detailed eCO<sub>2</sub>R model. The chemical system uses static process inventories and represents the state-of-the-art chemical processes. The eCO<sub>2</sub>R model describes the eCO<sub>2</sub>R process and provides a variable process inventory depending on the operating parameters.

#### 3.1. eCO<sub>2</sub>R model

Fig. 2 presents the status of the current eCO<sub>2</sub>R development (*cf.* ESI 2†).<sup>11,12,20–24</sup> We focus on two of the most common eCO<sub>2</sub>R products, CO and ethylene. Additionally, we evaluate eCO<sub>2</sub>R to methanol due to its universality as energy carrier and basic chemical. The laboratory experiments are sorted by cell voltage

on the horizontal axes and faradaic efficiency on the vertical axes. If no anode potential was reported, we assumed the ideal oxygen evolution reaction potential of 1.23 V. Therefore, the applied cell voltage during the experiment might be higher than illustrated in Fig. 2. Many experiments have been conducted on eCO<sub>2</sub>R to carbon monoxide (CO) up to high faradaic efficiencies of over 90% and cell voltages down to 2 V. eCO<sub>2</sub>R to ethylene is commonly reported at similar cell voltages; yet, faradaic efficiencies do not regularly reach more than 50%. Methanol is only detected with very low faradaic efficiencies. The current density indicates that economically feasible operation<sup>8</sup> at  $\geq 300$  mA cm<sup>-2</sup> is within reach for eCO<sub>2</sub>R to CO and eCO<sub>2</sub>R to ethylene. The single-pass CO<sub>2</sub> conversion reported reaches only 4% on average and 28% at maximum. However, our review shows that increasing current density correlates with increasing CO<sub>2</sub> conversion (*cf.* ESI 1 Section 2†). Thus, by further technology development toward industrial operation at high current density, we expect the CO<sub>2</sub> conversion to increase.

For our study, we developed an eCO<sub>2</sub>R model (*cf.* Fig. 3) based on the following operating parameters: cell voltage, faradaic efficiency, and single-pass CO<sub>2</sub> conversion. The current density is not considered here since it mainly affects capital investments, and previous LCA studies on electrolyzers have shown a negligible ecological impact from capital construction.<sup>5</sup> The model is solely based on energy, mass, and entropy balances and is valid for all construction types, *i.e.*, alkaline electrolyzers or solid oxide electrolyzers. In the electrochemical cell, CO<sub>2</sub> and H<sub>2</sub>O are converted into chemical products with the single-pass CO<sub>2</sub> conversion defining the ratio of reacted CO<sub>2</sub> at a single pass.

The eCO<sub>2</sub>R's electricity demand per kilogram product includes the energy demand of the cell and the separation unit (eqn (4)).



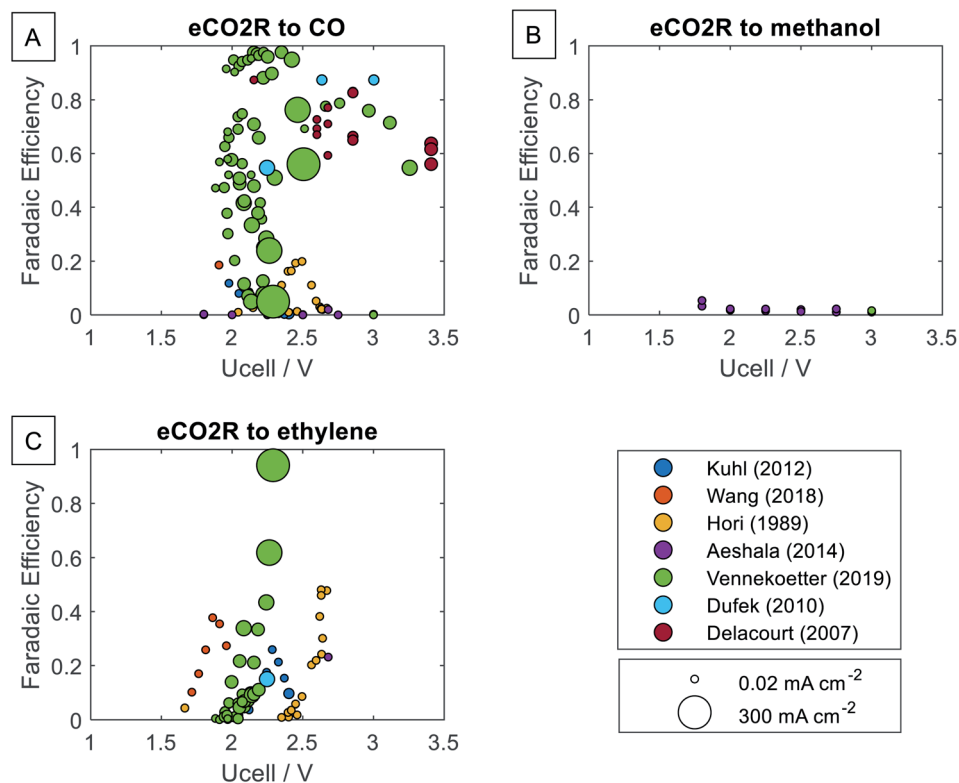


Fig. 2 Status of laboratory eCO<sub>2</sub>R operating parameters sorted by product: (A) carbon monoxide (CO), (B) methanol, and (C) ethylene. High faradaic efficiency and low cell voltage yield energy-efficient operation; high current density indicates economic feasibility.

$$w_{\text{el}} = w_{\text{el,cell}} + w_{\text{el,SU}} \quad (4)$$

The electricity demand of the cell  $w_{\text{el,cell}}$  depends on the cell voltage  $U_{\text{cell}}$  and the faradaic efficiency  $\eta_{\text{FE}}$  (eqn (5)).

$$w_{\text{el,cell}} = f(U_{\text{cell}}, \eta_{\text{FE}}) = \frac{U_{\text{cell}} z_p F}{\eta_{\text{FE}} M_p} \quad (5)$$

The cell voltage cannot be lower than the equilibrium voltage of (a) 1.33 V for CO, (b) 1.20 V for methanol, and (c) 1.15 V for ethylene.<sup>2</sup> The number of exchanged electrons  $z_p$  and the molecular weight  $M_p$  depend on the product. The Faraday constant  $F$  expresses the electric charge per mole of electrons. The generated oxygen at the anode is released into the environment without further utilization. All cathode products are assumed to be gaseous under the industrial operating temperature of 70–80 °C. Possible electrolyte and liquid product fractions are assumed to be recycled without further treatment or separation. The outlet vapor is a mixture of unreacted CO<sub>2</sub>, the main product, and by-products. The chemical composition of outlet vapor is determined by the faradaic efficiency, the single-pass CO<sub>2</sub> conversion, and a by-product distribution factor. We assume a commonly reported distribution of by-products with H<sub>2</sub> (40%), CO (30%), formate (10%), methane (10%), and ethylene (10%).<sup>2</sup>

In the subsequent separation unit, unreacted CO<sub>2</sub> is separated and recycled, and the product is purified from by-products. An exact modeling of the separation unit for eCO<sub>2</sub>R

post-processing is difficult, given the variety of eCO<sub>2</sub>R products<sup>2</sup> and the amount of unreacted CO<sub>2</sub> in the cathode stream.<sup>12</sup> The separation of eCO<sub>2</sub>R is more complex than for water electrolysis<sup>5</sup> due to the higher number of chemical compounds involved at similar vapor pressures.<sup>2</sup> Regarding the separation unit, several technologies are applicable: from energy-intensive distillation with good numerical models to energy-efficient membrane technologies with poor numerical models being available. Previous studies used generalized process calculations for standard separation processes, *i.e.*, pressure swing adsorption and distillation (*cf.* Table 1).<sup>13–15</sup> While they generated a more accurate separation duty of each process, the dependency on input compositions and superposition of separation processes<sup>25</sup> could not be considered. To avoid complex process design but consider the dependency on input compositions, we use an ideal separation unit that estimates the separation duty based on the minimum thermodynamic energy for separation (eqn (6)).

$$w_{\text{el,SU}} = f(x_i(\eta_{\text{FE}}, C_{\text{conv}})) = \frac{RT}{\dot{m}_p} \cdot \left[ - \sum_i n_i^{\text{vap}} \ln(x_i^{\text{vap}}) + \sum_i n_i^{\text{sep}} \ln(x_i^{\text{sep}}) \right] \quad (6)$$

The difference between the mixing entropy of the incoming vapor (vap) and the separated outlet streams (sep) determines the required electricity. The index  $i$  represents each component in the streams, while  $T$  refers to ambient temperature, and  $R$



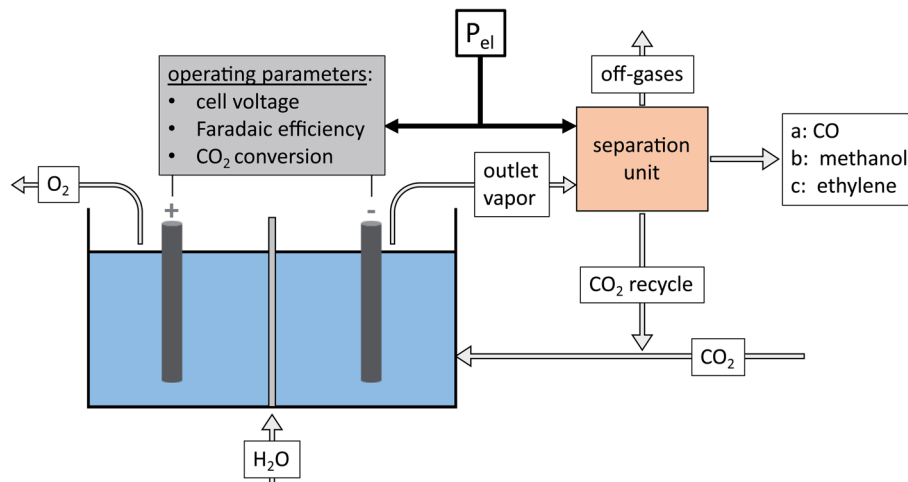


Fig. 3 Schematic illustration of the eCO<sub>2</sub>R model that allows calculating the electricity demand  $P_{el}$ , the required CO<sub>2</sub>, and the required water per kilogram product (i.e., carbon monoxide (CO), methanol, or ethylene) depending on the operating parameters cell voltage, faradaic efficiency, and single-pass CO<sub>2</sub> conversion.

represents the universal gas constant. We analyze the effect of non-ideal separation in Section 4.1 and discuss the impact on the results in Section 5.2.

All by-products are regarded as waste and are released into the environment as off-gases at fully oxidized state without heat integration. We do not consider multi-product eCO<sub>2</sub>R pathways that synthesize several valuable eCO<sub>2</sub>R products simultaneously. This assumption is in line with the findings of Larrazábal *et al.*,<sup>10</sup> who concluded the need for high selectivity toward a single eCO<sub>2</sub>R product. In Section 5.1, we discuss the effect of multi-product eCO<sub>2</sub>R on our results.

Our eCO<sub>2</sub>R model allows calculating the amount of required electricity, carbon dioxide, and water per kilogram desired chemical. Due to the equation-based approach, a variable eCO<sub>2</sub>R process inventory is available to conduct a parameter-

dependent assessment. More details regarding the eCO<sub>2</sub>R model can be found in the ESI 1 Section 3.†

### 3.2. Chemical system

The chemical system complements the variable (non-linear) eCO<sub>2</sub>R model with a static (linear) model of the state-of-the-art chemical processes. The goal of the chemical system is to produce methanol (1) or ethylene (2), as illustrated in Fig. 4. All energy requirements, i.e., electricity, heat, and cooling energy, are provided by electricity using electric heating and compression chillers (*cf.* ESI 3†). We consider low-carbon electricity from wind power with 0.017 kg CO<sub>2</sub>-eq. per kW h,<sup>26</sup> while the impact of other electricity sources is evaluated in Section 4.4. A direct air capture (DAC) plant adsorbs carbon dioxide (CO<sub>2</sub>) from the atmosphere to supply the chemical's carbon demand.<sup>27</sup> The

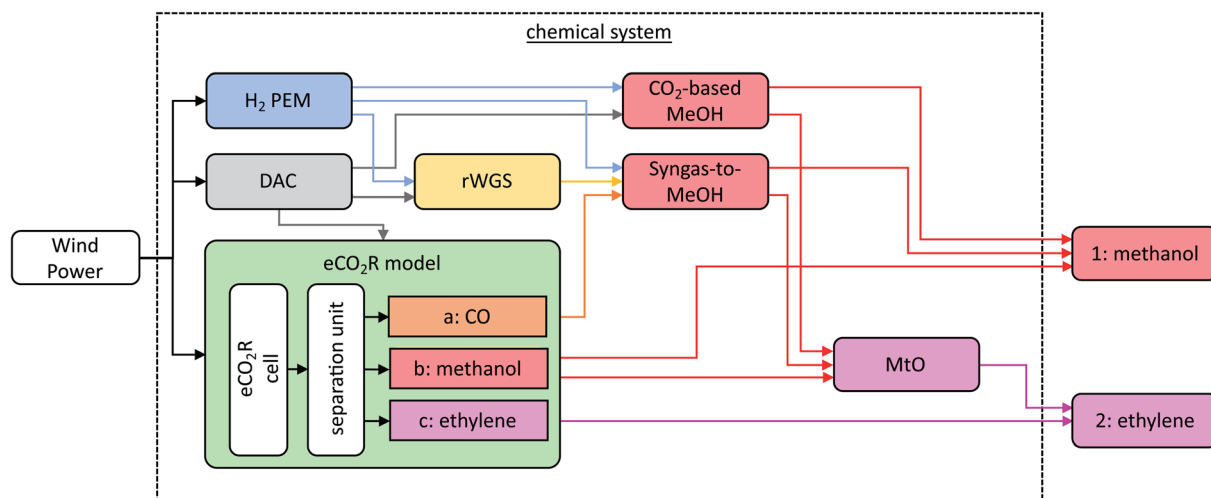


Fig. 4 Illustration of the chemical system. The color indicates the product of each process. H<sub>2</sub> PEM: proton exchange membrane water electrolysis, DAC: direct air capture, rWGS: reverse water gas shift, eCO<sub>2</sub>R: electrochemical CO<sub>2</sub> reduction, MeOH: methanol, MtO: methanol to olefins.



Table 2 Technological power-to-methanol and power-to-ethylene pathways (cf. Fig. 4)

| (1) Methanol          |      | (2) Ethylene   |      |  |
|-----------------------|------|--|------|--|
| H <sub>2</sub> -based | Ref1 | <b>H<sub>2</sub>-based MeOH pathway</b><br>(H <sub>2</sub> PEM + DAC + CO <sub>2</sub> -based MeOH)                      | Ref2 | <b>H<sub>2</sub>-based ethylene pathway</b><br>(H <sub>2</sub> PEM + DAC + CO <sub>2</sub> -based MeOH + MtO)                      |
| eCO <sub>2</sub> R    | 1a   | <b>eCO<sub>2</sub>R via CO to MeOH pathway</b><br>(DAC + eCO <sub>2</sub> R to CO + H <sub>2</sub> PEM + syngas-to-MeOH) | 2a   | <b>eCO<sub>2</sub>R via CO to ethylene pathway</b><br>(DAC + eCO <sub>2</sub> R to CO + H <sub>2</sub> PEM + syngas-to-MeOH + MtO) |
|                       | 1b   | <b>Direct eCO<sub>2</sub>R to MeOH pathway</b><br>(DAC + eCO <sub>2</sub> R to MeOH)                                     | 2b   | <b>eCO<sub>2</sub>R via MeOH to ethylene pathway</b><br>(DAC + eCO <sub>2</sub> R to MeOH + MtO)                                   |
|                       |      |  | 2c   | <b>Direct eCO<sub>2</sub>R to ethylene pathway</b><br>(DAC + eCO <sub>2</sub> R to ethylene)                                       |

desorption energy of the DAC is provided by a heat pump system.<sup>27</sup> An industrial-scale proton exchange membrane (PEM) electrolyzer generates hydrogen (H<sub>2</sub>) from electricity and water (H<sub>2</sub>O).<sup>5</sup> The eCO<sub>2</sub>R system reduces CO<sub>2</sub> and H<sub>2</sub>O to three single products: (a) CO, (b) methanol, or (c) ethylene. The reverse water gas shift (rWGS) reaction<sup>28</sup> tunes the ratio of H<sub>2</sub> to CO to produce syngas with the suitable composition for the syngas-to-methanol<sup>13</sup> process. The CO<sub>2</sub>-based methanol<sup>29</sup> process converts H<sub>2</sub> and CO<sub>2</sub> to methanol. The methanol-to-olefins (MtO) process produces ethylene with a high energy efficiency of 98%.<sup>30</sup> For both methanol and ethylene, we assume these products to be incinerated with stoichiometric combustion to reflect the carbon release at their end-of-life. Thereby, we consider the entire life cycle from cradle-to-grave. The process inventories are obtained from scientific studies that conducted detailed process design and are provided (cf. ESI 3†). Further background data like the supply of deionized water are taken from the database ecoinvent v3.7 (cf. ESI 1 Section 1†).<sup>26</sup>

In summary, the chemical system has three technological options for power-to-methanol production (cf. Table 2). The H<sub>2</sub>-based pathway using PEM electrolysis, DAC, and CO<sub>2</sub>-based methanol synthesis defines the state-of-the-art reference (Ref1). Regarding eCO<sub>2</sub>R, two pathways are possible: (a) the eCO<sub>2</sub>R via CO to MeOH pathway that reacts CO<sub>2</sub> from DAC to CO with H<sub>2</sub> from PEM electrolysis to produce methanol via the syngas-to-methanol process (1a), and (b) the direct eCO<sub>2</sub>R to MeOH pathway, where the eCO<sub>2</sub>R cell generates methanol directly from CO<sub>2</sub> and H<sub>2</sub>O (1b).

For power-to-ethylene production, a total of four pathways are possible (cf. Table 2). The reference pathway (Ref2) uses the H<sub>2</sub>-based pathway to methanol with subsequent conversion to ethylene via the MtO process. Pathways 2a and 2b express the two eCO<sub>2</sub>R pathways to methanol with subsequent MtO conversion to ethylene. The direct eCO<sub>2</sub>R to ethylene pathway (2c) directly converts CO<sub>2</sub> and H<sub>2</sub>O to ethylene. All technological pathways and their abbreviations are summarized in Table 2.

## 4. Results

We present the results starting with the analysis of the eCO<sub>2</sub>R model in Section 4.1. In Sections 4.2 and 4.3, we show the results of the parameter-dependent assessment and the minimum development requirements for eCO<sub>2</sub>R for the two products, methanol (Section 4.2) and ethylene (Section 4.3). We continue by presenting the sensitivity analysis on the electricity source and the comparison to fossil technologies in Sections 4.4 and 4.5, respectively.

### 4.1. eCO<sub>2</sub>R system analysis

Fig. 5A shows the electricity demand per kilogram of carbon monoxide (CO) for a constant single-pass CO<sub>2</sub> conversion of 50%. The electricity demand rises progressively with decreasing faradaic efficiency. Decreasing faradaic efficiency increases the electric charge that does not contribute to the production of the desired product. The high dependency on the faradaic efficiency underlines the importance of selective catalysts, as already

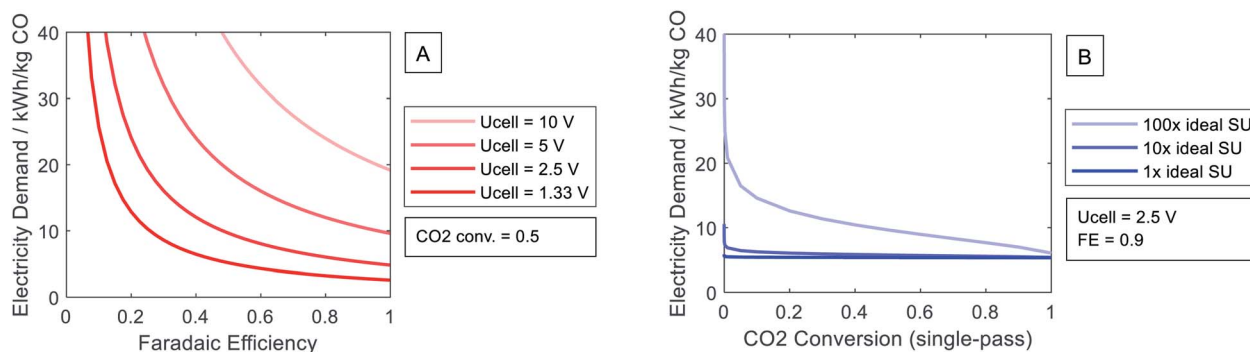


Fig. 5 Electricity demand of eCO<sub>2</sub>R depending on the key operating parameters: (A) the influence of faradaic efficiency (FE) and cell voltage, (B) the impact of the single-pass CO<sub>2</sub> conversion and non-ideal separation units (SU).



highlighted by Larrazábal *et al.*<sup>10</sup> The cell voltage  $U_{\text{cell}}$  increases the electricity demand linearly. In practice, the cell voltage cannot be decreased down to the equilibrium potential because excess voltage is required to enforce the electrochemical conversion. Thus, minimal difference to the equilibrium potential, *i.e.*, minimal overpotential, is necessary for efficient electrochemical processes.

To illustrate the effect of the single-pass  $\text{CO}_2$  conversion on the electricity demand (*cf.* Fig. 5B), we chose commonly estimated future values of the cell voltage  $U_{\text{cell}} = 2.5$  V and the faradaic efficiency  $\eta_{\text{FE}} = 90\%$ .<sup>13,15</sup> The ideal separation unit has a limited contribution of 0.8% to the overall electricity demand at a  $\text{CO}_2$  conversion of 50%. Fig. 5B also shows the effect of the separation unit requiring 10 and 100 times the ideal separation duty. For comparison purposes, a study on  $\text{CO}_2$  sources reported that the separation units require between 1.5 and 9 times the ideal separation duty.<sup>31</sup> The illustrated non-ideal separation units of 10 and 100 times the ideal separation duty refer to a contribution of 7.5% and 45% of the overall electricity demand at a  $\text{CO}_2$  conversion of 50%. Thus, efficient separation processes are essential to achieve climate benefits. Yet, the precise energy demand depends on the design of the separation unit that varies for different  $\text{eCO}_2\text{R}$  products and mass flows.<sup>15</sup> To avoid complex process design, we continue with an ideal separation unit to analyze the minimum development requirements for  $\text{eCO}_2\text{R}$ . We further discuss the effect of non-ideal separation in Section 5.2 and ESI 1 Section 4.† As shown in Fig. 5, the impact of the single-pass  $\text{CO}_2$  conversion is limited compared to cell voltage and faradaic efficiency. For further investigations, the single-pass  $\text{CO}_2$  conversion is set to the commonly estimated level of 50%.<sup>13–15</sup>

#### 4.2. Power-to-methanol

Fig. 6 presents the parameter-dependent assessment of power-to-methanol technologies. It shows the cradle-to-grave global warming impact (indicated by colors) of the chemical system for 1 kg of methanol (MeOH) as a function of the  $\text{eCO}_2\text{R}$  operating parameters (cell voltage on the x-axis and faradaic efficiency on the y-axis). If the  $\text{eCO}_2\text{R}$  pathways (1a and 1b) emit more GHG

emissions than the competing  $\text{H}_2$ -based pathway (Ref1), Ref1 is chosen; thus, the GHG emission of methanol production remains independent from the  $\text{eCO}_2\text{R}$  operation parameters. At the breakeven line, the GHG emissions of 1 kg methanol production are identical for the  $\text{eCO}_2\text{R}$  pathway (1a or 1b) and the  $\text{H}_2$ -based pathway (Ref1). Thus, the  $\text{eCO}_2\text{R}$  operating parameters at the breakeven line represent minimum development requirements for  $\text{eCO}_2\text{R}$  to achieve climate change competitiveness. From the scatter of experimental studies, we can identify the development gap. To investigate the GHG reduction potential of the  $\text{eCO}_2\text{R}$  pathways, we define the best-case as operation at equilibrium voltage and 100% faradaic efficiency (top-left corners in Fig. 6).

Regarding the  $\text{eCO}_2\text{R}$  *via* CO to MeOH pathway (1a), the  $\text{eCO}_2\text{R}$  has to reach a faradaic efficiency higher than 46% and a cell voltage lower than 3.3 V to reduce GHG emissions (*cf.* Fig. 6A). Several laboratory studies investigating the cathodic behavior already fulfill these requirements.<sup>12,23,24</sup> The  $\text{eCO}_2\text{R}$  *via* CO to MeOH pathway (1a) benefits from the higher energy efficiency of the syngas-to-methanol process compared to the  $\text{CO}_2$ -based methanol process. However, most of the electricity demand is still due to the PEM electrolysis to generate hydrogen for the syngas. Therefore, the GHG reduction potential per kilogram of methanol is limited to 12% for the best-case  $\text{eCO}_2\text{R}$  (1a best-case: 0.210 kg  $\text{CO}_2$ -eq. per kg MeOH, Ref1: 0.238 kg  $\text{CO}_2$ -eq. per kg MeOH).

Fig. 6B shows the parameter-dependent assessment of the direct  $\text{eCO}_2\text{R}$  to MeOH pathway (1b). In contrast to the  $\text{eCO}_2\text{R}$  *via* CO to MeOH pathway, all chemical conversions occur inside the  $\text{eCO}_2\text{R}$  cell, and no PEM electrolysis is needed. The direct  $\text{eCO}_2\text{R}$  to MeOH pathway (1b) requires a faradaic efficiency higher than 64% and a cell voltage below 2.3 V for climate benefits over the  $\text{H}_2$ -based pathway. The minimum development requirements are more ambitious than those for the  $\text{eCO}_2\text{R}$  *via* CO to MeOH pathway (1a) since the highly efficient PEM electrolysis is replaced. However, the GHG reduction potential is more significant once the minimum development requirements are fulfilled, up to 38% for the best-case direct  $\text{eCO}_2\text{R}$  to MeOH pathway (1b best-case: 0.147 kg  $\text{CO}_2$ -eq. per kg

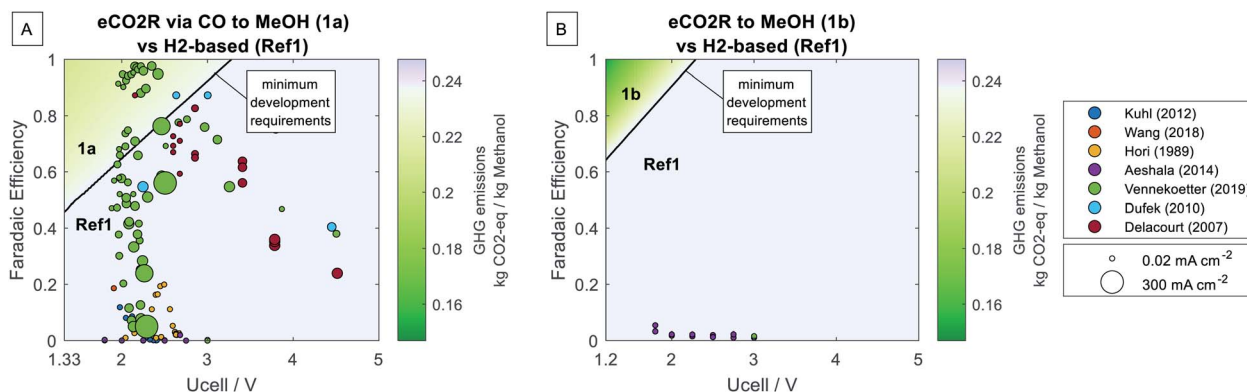


Fig. 6 Parameter-dependent assessment of (A) the  $\text{eCO}_2\text{R}$  *via* CO to MeOH pathway 1a and (B) the direct  $\text{eCO}_2\text{R}$  to MeOH pathway 1b using wind power with 0.017 kg  $\text{CO}_2$ -eq. per kW h. The single-pass  $\text{CO}_2$  conversion is set to 50%. The black line illustrates the minimum development requirements to achieve climate benefits over the  $\text{H}_2$ -based pathway to methanol (Ref1). The scatter indicates the status of laboratory experiments.



MeOH, Ref1: 0.238 kg CO<sub>2</sub>-eq. per kg MeOH). Unfortunately, methanol proves very difficult to be produced directly in electrochemical cells despite recent efforts.<sup>13,21</sup> Thus, the development gap to achieve climate benefits with the direct eCO<sub>2</sub>R to MeOH pathway is large, as illustrated by the scatter in Fig. 6B.

### 4.3. Power-to-ethylene

Besides power-to-ethylene pathways using methanol as an intermediate (2a, 2b, and Ref2), direct eCO<sub>2</sub>R to ethylene (2c) is possible. The minimum development requirements for the eCO<sub>2</sub>R *via* CO to ethylene pathway (2a) and the eCO<sub>2</sub>R *via* MeOH to ethylene pathway (2b) are identical to the previous assessment (Section 4.2). The reason for the identical minimum development requirements is that the same methanol-to-olefins process is required for the 2a, 2b, and Ref2 pathways. The GHG emissions per kilogram ethylene using the H<sub>2</sub>-based pathway Ref2 refer to 0.589 kg CO<sub>2</sub>-eq. per kg ethylene (*cf.* Fig. 7). Assuming the best-case for the eCO<sub>2</sub>R *via* CO to ethylene pathway (2a) and the eCO<sub>2</sub>R *via* MeOH to ethylene pathway (2b), the same relative reductions of 12% and 38% as for methanol alone would be achieved (2a best-case: 0.522 kg CO<sub>2</sub>-eq. per kg ethylene, 2b best-case: 0.367 kg CO<sub>2</sub>-eq. per kg ethylene; *cf.* Fig. 7A and B).

The direct eCO<sub>2</sub>R to ethylene pathway (2c) requires a faradaic efficiency higher than 58% and a cell voltage below 2.5 V to achieve climate benefits over the H<sub>2</sub>-based pathway Ref2 (*cf.* Fig. 7C). The minimum development requirements of the direct eCO<sub>2</sub>R to ethylene pathway (2c) are about 10% less ambitious than ones for the eCO<sub>2</sub>R *via* methanol to ethylene pathway (2b) since efficiency losses and direct process emissions of the methanol-to-olefins process are omitted. The best-case for the direct eCO<sub>2</sub>R to ethylene pathway (2c) can reduce the GHG emissions of ethylene production by 44% compared to the H<sub>2</sub>-based pathway (2c best-case: 0.328 kg CO<sub>2</sub>-eq. per kg ethylene).

While the relative reduction potential of the eCO<sub>2</sub>R *via* CO to ethylene pathway (2a) is limited to 12%, both eCO<sub>2</sub>R *via* MeOH to ethylene pathway (2b) and direct eCO<sub>2</sub>R to ethylene pathway (2c) have high reduction potentials once the minimum development requirements are fulfilled. In contrast to eCO<sub>2</sub>R to methanol, eCO<sub>2</sub>R to ethylene is commonly reported in experimental studies.<sup>12</sup> The more straightforward synthesis of ethylene in eCO<sub>2</sub>R cells implicates a smaller development gap to achieve climate benefits. Additionally, the omission of the methanol-to-olefins process contributes to 10% less ambitious minimum development requirements. Hence, the small development gap to reach significant climate benefits over the H<sub>2</sub>-

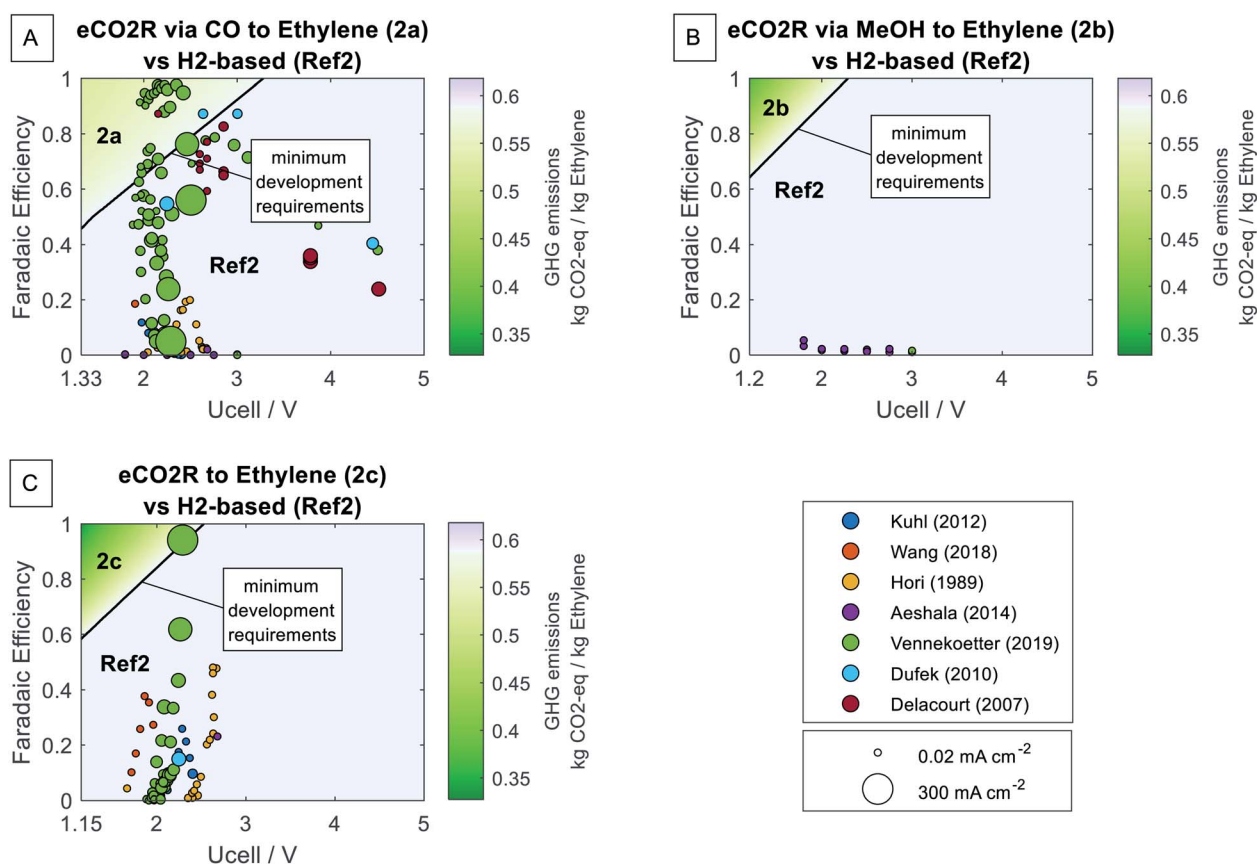


Fig. 7 Parameter-dependent assessment of (A) the eCO<sub>2</sub>R *via* CO to ethylene pathway 2a, (B) the eCO<sub>2</sub>R *via* MeOH to ethylene pathway 2b, and (C) the direct eCO<sub>2</sub>R to ethylene pathway 2c using wind power with 0.017 kg CO<sub>2</sub>-eq. per kW h. The single-pass CO<sub>2</sub> conversion is set to 50%. The black line illustrates the minimum development requirements to achieve climate benefits over the H<sub>2</sub>-based pathway to ethylene (Ref2). The scatter indicates the status of laboratory experiments.



based pathway is promising. Our findings are in line with Nabil *et al.*'s study, which predicted future benefits for eCO<sub>2</sub>R to ethylene over the H<sub>2</sub>-based pathways.<sup>15</sup> Our results show that the direct eCO<sub>2</sub>R to ethylene pathway has considerable potential for GHG emission reduction in the chemical industry.

#### 4.4. Sensitivity on electricity source

Since power-to-chemical pathways are inherently dependent on the electricity source, we evaluate the sensitivity of our findings with respect to the electricity source. Fig. 8 shows the parameter-dependent assessment of the direct eCO<sub>2</sub>R to ethylene pathway using wind power, photovoltaic, and the current German grid mix.<sup>26</sup> Wind power represents the lowest GHG emissions, and the German grid mix represents a considerable fossil fuel-based electricity generation. Photovoltaic and other low-carbon electricity sources like concentrated solar power are in-between and follow the same trends. The life cycle GHG emissions of the power-to-ethylene pathways increase with the increase of the electricity's carbon intensity. Nevertheless, both the minimum development requirements and the relative reduction potential of the direct eCO<sub>2</sub>R to ethylene pathway over the H<sub>2</sub>-based pathway remain constant. The same trend can be seen in all the five eCO<sub>2</sub>R pathways. Thus, our findings for minimum development requirements are independent of the electricity source.

The GHG emissions of all considered power-to-chemical pathways are dominated by the electricity source. However, not all GHG emissions are caused by the electricity supply since there are process emissions from eCO<sub>2</sub>R off-gases, the CO<sub>2</sub>-based methanol process, the syngas-to-methanol process, the methanol-to-olefins process, and the deionized water supply (*cf.* ESI 3†). Therefore, a minimization of the electricity demand instead of our chosen objective of GHG minimization would yield similar but not identical results regarding the minimum development requirements of eCO<sub>2</sub>R.

#### 4.5. Comparison to fossil technologies

Power-to-chemical pathways are usually not beneficial unless the electricity sector is largely decarbonized. The black line in the color bar of Fig. 8 indicates the fossil state-of-the-art ethylene production of 4.54 kg CO<sub>2</sub>-eq. per kg ethylene.<sup>26</sup> An implementation with the current German grid mix would increase the GHG emissions of ethylene production by at least 149% (2c best-case using German grid mix: 11.3 kg CO<sub>2</sub>-eq. per kg ethylene) up to 347% (Ref2 using German grid mix: 20.3 kg CO<sub>2</sub>-eq. per kg ethylene).

Fig. 8 also illustrates that the relative reduction potential of the direct eCO<sub>2</sub>R to ethylene pathway over the H<sub>2</sub>-based pathway is minor compared to the reduction potential by using low-carbon electricity. The same trend is observed for all eCO<sub>2</sub>R pathways. In Table 3, we compare the GHG reduction potential

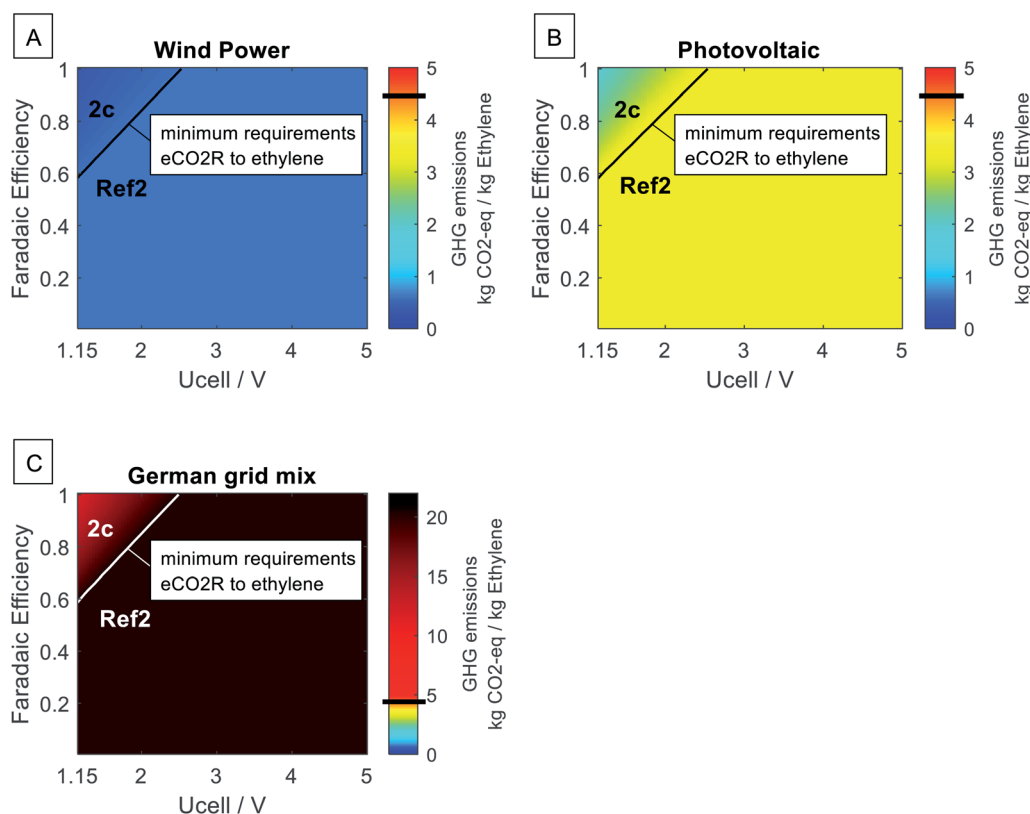


Fig. 8 Sensitivity analysis for the direct eCO<sub>2</sub>R to ethylene pathway (2c) vs. H<sub>2</sub>-based pathway to ethylene (Ref2) using (A) wind power 0.017 kg CO<sub>2</sub>-eq. per kW h, (B) electricity from photovoltaic 0.095 kg CO<sub>2</sub>-eq. per kW h, and (C) the current German grid mix 0.580 kg CO<sub>2</sub>-eq. per kW h.<sup>26</sup> The horizontal black line in the color bar indicates the fossil state-of-the-art technology of 4.54 kg CO<sub>2</sub>-eq. per kg ethylene.<sup>26</sup>



**Table 3** Comparison of GHG emissions per kilogram chemical of the fossil-based state-of-the-art technologies, the H<sub>2</sub>-based pathways, and the eCO<sub>2</sub>R pathways (in kg CO<sub>2</sub>-eq. per kg). Best-case implicates operation at equilibrium potential and 100% faradaic efficiency. The relative reduction potentials refer to the fossil-based state-of-the-art

|              | Fossil state of the art <sup>26</sup> | Wind-powered power-to-chemicals |                                   |  |  |  |
|--------------|---------------------------------------|---------------------------------|-----------------------------------|--|--|--|
|              |                                       | H <sub>2</sub> -based (Ref)     | H <sub>2</sub> -based (best-case) | Best-case eCO <sub>2</sub> R to CO (a) | Best-case eCO <sub>2</sub> R to methanol (b) | Best-case eCO <sub>2</sub> R to ethylene (c) |
| Methanol (1) | 2.00                                  | 0.238 (−88%)                    | 0.168 (−92%)                      | 0.210 (−90%)                           | 0.147 (−93%)                                 |  |
| Ethylene (2) | 4.54                                  | 0.589 (−87%)                    | 0.419 (−91%)                      | 0.522 (−89%)                           | 0.367 (−92%)                                 | 0.328 (−93%)                                 |

of all considered power-to-chemical pathways over the state-of-the-art fossil-based pathway.<sup>26</sup> By using low carbon wind power for H<sub>2</sub>-based pathways, the GHG emissions per kilogram chemical are reduced by 88% for methanol and by 87% for ethylene. The necessary technologies of PEM electrolysis, direct air capture, and heat pumps for the H<sub>2</sub>-based pathways are already available on an industrial scale. While PEM electrolysis currently reaches 60% energy efficiency,<sup>5</sup> further efficiency gains in water electrolysis are expected due to higher plant efficiency of PEM electrolyzers<sup>5</sup> and more advanced solid oxide water electrolyzers becoming industrial available. Best-case water electrolysis operating under the same assumptions of equilibrium voltage and 100% faradaic efficiency would reduce the GHG emissions by 92% for methanol and by 91% for ethylene (*cf.* Table 3).

The eCO<sub>2</sub>R pathways could further reduce the GHG emissions of chemical production. Assuming best-case eCO<sub>2</sub>R that operates at equilibrium voltage and 100% faradaic efficiency, the eCO<sub>2</sub>R pathways yield GHG emission reduction between 89% and 93% compared to the fossil state-of-the-art. The direct eCO<sub>2</sub>R to ethylene pathway reaches the largest relative reduction and the eCO<sub>2</sub>R *via* CO pathways the smallest relative reduction (*cf.* Table 3). The additional reduction potential of eCO<sub>2</sub>R pathways over the H<sub>2</sub>-based pathways is minor in relation to the fossil-based state-of-the-art technology.

Given the limited reduction potential over the industrial available H<sub>2</sub>-based pathway, the low technology readiness level of eCO<sub>2</sub>R,<sup>9</sup> the complex electrochemical phenomena,<sup>2</sup> the product separation energy demand,<sup>15</sup> and expected efficiency gains in water electrolysis,<sup>5</sup> we see major challenges for eCO<sub>2</sub>R to play a key role to reduce GHG emissions of the chemical industry in the near future. Nonetheless, efficient eCO<sub>2</sub>R can contribute to a reduction in GHG emissions. The direct eCO<sub>2</sub>R to ethylene pathway is the most promising candidate to achieve climate change competitiveness over the H<sub>2</sub>-based power-to-chemical pathway.

## 5. Discussion

### 5.1. Multi-product eCO<sub>2</sub>R pathways

In the present article, we investigated eCO<sub>2</sub>R to single products only and considered all by-products as waste. Nonetheless, the simultaneous production of several products by eCO<sub>2</sub>R is possible and might yield further GHG emission reduction. We differentiate three types of multi-product eCO<sub>2</sub>R: (a) co-electrolysis, where eCO<sub>2</sub>R by-products are used without

further separation, (b) heat recovery from eCO<sub>2</sub>R by-products incineration, and (c) stand-alone multi-product eCO<sub>2</sub>R, where by-products are purified and sold as stand-alone products.

The most common co-electrolysis is the simultaneous production of carbon monoxide (CO) and hydrogen (H<sub>2</sub>), where the by-product H<sub>2</sub> does not need to be separated from CO.<sup>32</sup> The usage of the by-product H<sub>2</sub> lowers the required H<sub>2</sub> from other sources and thus, yields further GHG emission reduction. In practice, the co-production of H<sub>2</sub> in eCO<sub>2</sub>R cells has to compete against the PEM electrolysis that is optimized on H<sub>2</sub>.<sup>5</sup> Therefore, co-electrolysis usually focuses on maximal CO yield<sup>33</sup> and the GHG reduction potential is limited.

Heat recovery of eCO<sub>2</sub>R by-products has a similar effect. The incineration of by-products for heat recovery and subsequent integration reduces the heat demand from other sources and yields further GHG emission reduction. The potential of eCO<sub>2</sub>R by-products as electricity-based heating fuels is limited since more efficient technology is available to generate heat from electricity.<sup>34</sup>

Multi-product eCO<sub>2</sub>R in terms of simultaneous production of several stand-alone products requires further separation and purification. On one side, the by-product usage decreases GHG emissions; on the other side, the increased energy demands for separation and purification increases GHG emissions. From an economic point of view, eCO<sub>2</sub>R with complex separation units is unlikely to be economically feasible due to the limited scale-up potential of electrolyzers.<sup>5,10</sup>

While a utilization of by-products might further reduce GHG emissions and the minimum development requirements of eCO<sub>2</sub>R, the above discussion shows that these reductions are expected to be limited in practice.

### 5.2. Separation duty of eCO<sub>2</sub>R

The minimum development requirements for eCO<sub>2</sub>R presented in this article refer to an ideal separation unit and serve as a lower bound to guide eCO<sub>2</sub>R development. The preliminary assumption of ideal separation is an appropriate approach to avoid complex design; however, the assumption underestimates the energy demand of real separation processes. Thus, the conducted minimum development requirements only illustrate the lower bound for eCO<sub>2</sub>R development and guide toward the most promising pathways. A detailed process design for the eCO<sub>2</sub>R separation needs to be performed to determine the development requirements for eCO<sub>2</sub>R practical implementation and is out of the scope of this study. A sensitivity analysis on the



separation duty (*cf.* ESI 1 Section 4†) has shown that the energy demand for separation can considerably increase the development requirements for eCO<sub>2</sub>R. When using energy-efficient separation processes requiring previously reported ratios<sup>31</sup> of up to 10 times the ideal separation duty, the development requirements are moderately increased by 6%. Before implementing an eCO<sub>2</sub>R pathway, we recommend conducting a detailed separation process design and determine the eCO<sub>2</sub>R process requirements for the specific design.

### 5.3. Further perspectives

We thoroughly examined the eCO<sub>2</sub>R pathways regarding their potential to reduce GHG emissions and the technological competition against the industrial available H<sub>2</sub>-based pathways. However, there are further perspectives on eCO<sub>2</sub>R.

The eCO<sub>2</sub>R pathways could yield economic benefits. The one-step conversion of the direct eCO<sub>2</sub>R to MeOH pathway and the direct eCO<sub>2</sub>R to ethylene pathway offer reduced plant complexity and geographical independence from chemical production sites for pre- or post-processing. Furthermore, the omission of PEM electrolysis provides potential capital investment savings by using less expensive catalysts for eCO<sub>2</sub>R than the noble metals for PEM electrolysis.<sup>5</sup> A previous study by Adnan *et al.*<sup>13</sup> has shown similar costs for both the eCO<sub>2</sub>R and H<sub>2</sub>-based pathways. A detailed techno-economic assessment including capital expenditure remains uncertain until higher technological readiness levels are reached, given the different catalysts used for eCO<sub>2</sub>R<sup>10</sup> and divers cell construction types.<sup>12</sup> A preliminary analysis of operating expenditure (*cf.* ESI 1 Section 5†) shows that the power-to-chemical pathways range from -4% cost reduction to 73% cost increase compared to fossil-based state-of-the-art technologies. A detailed assessment of the development requirements from an economic point will further clarify the prospect of eCO<sub>2</sub>R and can be the object of future work.

Flexible operation of eCO<sub>2</sub>R can be beneficial for the surrounding energy system. When interconnecting the chemical industry with the electricity sector, more flexible operation due to the volatile characteristics of renewable electricity sources is needed.<sup>4</sup> Flexible operation of eCO<sub>2</sub>R and the comparison to the flexibility provided by H<sub>2</sub>-based pathways are yet to be investigated.

Future development of other technologies may affect eCO<sub>2</sub>R. In the present article, we intensively assessed the future development of eCO<sub>2</sub>R but used a static process inventory for the state-of-the-art chemical system. The interacting chemical processes might yield efficiency gains in future, and new technologies may be developed. Future development of competing technologies will increase the minimum development requirements for eCO<sub>2</sub>R and has to be regarded over time.

## 6. Conclusions

Electrochemical CO<sub>2</sub> reduction is an emerging technology that is competing against H<sub>2</sub>-based power-to-chemical pathways. Comparative assessments using fixed estimations of future

eCO<sub>2</sub>R operating parameters cannot reflect the development entirely. Our approach of a parameter-dependent assessment provides an in-depth understanding of the key operating parameters of eCO<sub>2</sub>R. We determined the eCO<sub>2</sub>R minimum development requirements on a multi-dimensional scale to achieve climate benefits over the H<sub>2</sub>-based pathway for power-to-methanol and power-to-ethylene.

- Electrochemical CO<sub>2</sub> reduction to ethylene requires a cell voltage below 2.5 V and a faradaic efficiency higher than 58% to achieve climate benefits over the H<sub>2</sub>-based pathway. According to reported experimental studies of direct eCO<sub>2</sub>R to ethylene, the development gap to fulfill the minimum development requirements is moderate. Regarding GHG emissions, the direct eCO<sub>2</sub>R to ethylene pathway offers a reduction potential of 44% over the H<sub>2</sub>-based pathway.

- Electrochemical CO<sub>2</sub> reduction to carbon monoxide requires a faradaic efficiency higher than 46% and a cell voltage lower than 3.3 V. Experimental studies have already fulfilled these requirements; however, the reduction potential over the H<sub>2</sub>-based pathway is limited to 12%.

- Electrochemical CO<sub>2</sub> reduction to methanol shows the most significant development gap to fulfill the minimum development requirements since eCO<sub>2</sub>R to methanol is rarely reported in experimental studies. Yet, the reduction potential over the H<sub>2</sub>-based pathway is 38%.

Compared to the fossil state of the art, both the H<sub>2</sub>-based and the eCO<sub>2</sub>R power-to-chemical pathways offer large GHG reduction of up to 88% and up to 93%, respectively. The additional reduction potential of the eCO<sub>2</sub>R pathways over the industrial available H<sub>2</sub>-based pathways is minor compared to the reduction potential by using low-carbon electricity. While H<sub>2</sub>-based pathways excel with high technology readiness level and further efficiency gains in water electrolysis are expected, the eCO<sub>2</sub>R pathways show a lower technology readiness level and require a complex separation. For these reasons, we see major challenges for eCO<sub>2</sub>R to be a key technology to reduce GHG emissions of the chemical industry. Nonetheless, efficient eCO<sub>2</sub>R, operating at low cell voltages and high faradaic efficiencies, can reduce GHG emissions of chemical production even more than H<sub>2</sub>-based pathways. We recommend focusing further research on direct eCO<sub>2</sub>R to ethylene due to its reduction potential over the H<sub>2</sub>-based pathway and the moderate experimental development gap.

## Conflicts of interest

There are no conflicts to declare.

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