

Cite this: *Chem. Sci.*, 2021, 12, 5938

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The *in situ* study of surface species and structures of oxide-derived copper catalysts for electrochemical CO₂ reduction†

Chunjun Chen,^{ab} Xupeng Yan,^{ab} Yahui Wu,^{ab} Shoujie Liu,^c Xiaofu Sun,^{ab} Qinggong Zhu,^a Rongjuan Feng,^a Tianbin Wu,^a Qingli Qian,^a Huizhen Liu,^a Lirong Zheng,^f Jing Zhang^f and Buxing Han^{abde}

Oxide-derived copper (OD-Cu) has been discovered to be an effective catalyst for the electroreduction of CO₂ to C₂+ products. The structure of OD-Cu and its surface species during the reaction process are interesting topics, which have not yet been clearly discussed. Herein, *in situ* surface-enhanced Raman spectroscopy (SERS), operando X-ray absorption spectroscopy (XAS), and ¹⁸O isotope labeling experiments were employed to investigate the surface species and structures of OD-Cu catalysts during CO₂ electroreduction. It was found that the OD-Cu catalysts were reduced to metallic Cu(0) in the reaction. CuO_x species existed on the catalyst surfaces during the CO₂RR, which resulted from the adsorption of preliminary intermediates (such as *CO₂ and *OCO[−]) on Cu instead of on the active sites of the catalyst. It was also found that abundant interfaces can be produced on OD-Cu, which can provide heterogeneous CO adsorption sites (strong binding sites and weak binding sites), leading to outstanding performance for obtaining C₂+ products. The Faradaic efficiency (FE) for C₂+ products reached as high as 83.8% with a current density of 341.5 mA cm^{−2} at −0.9 V vs. RHE.

Received 5th January 2021

Accepted 15th March 2021

DOI: 10.1039/d1sc00042j

rsc.li/chemical-science

Introduction

The electrochemical CO₂ reduction reaction (CO₂RR) has received wide attention, as it can not only reduce CO₂ into chemical fuels and feedstocks but it can also provide an energy storage solution for renewable energy sources.^{1–5} In particular, multi-carbon (C₂+) products are much more attractive due to their high energy densities and high economic values.^{6–10} However, the activity and selectivity for C₂+ products can be severely limited due to the slow kinetics of the C–C coupling step, which involves intricate multiple proton and electron transfer.^{11–15} Designing highly active catalysts for C₂+ products

and confirming the active sites are crucial for promoting the development of this area.

Cu-based catalysts are the most promising electrocatalysts for converting CO₂ to C₂+ products,^{16–20} especially oxide-derived Cu (OD-Cu).^{21–23} Studies using ambient-pressure X-ray photoelectron spectroscopy, electron energy-loss spectroscopy, and *ex situ* energy-dispersive X-ray spectroscopy have shown that oxide species in OD-Cu play a crucial role in activating CO₂ and in C–C coupling.^{24–26} However, it is well-known that Cu oxide and hydroxide species are unstable at negative potentials during the CO₂RR. In addition, most of the characterization studies of CuO_x species are based on *ex situ* methods, and it is difficult to study the structures and valences of catalysts under reaction conditions because reduced Cu can be oxidized very rapidly, even under an atmosphere with trace amounts of O₂. Recently, studies using *in situ* X-ray characterization have shown that the surface oxide layer can be fully reduced to metallic Cu.^{27,28} Thus, the real role of the CuO_x species of OD-Cu in promoting the CO₂RR is controversial. In previous reports,^{29,30} massive efforts have been focused on the presence or absence of Cu oxides; however, how Cu oxides are retained or formed during the CO₂RR is not clear. Also, comprehensive knowledge of the CuO_x species is crucial for understanding their role in promoting the CO₂RR.

Moreover, the surface structure of a catalyst often plays a crucial role in the production of C₂+ products.^{31–33} Experimental studies suggested that grain boundaries, low

^aBeijing National Laboratory for Molecular Sciences, CAS Key Laboratory of Colloid and Interface and Thermodynamics, CAS Research/Education Center for Excellence in Molecular Sciences, Institute of Chemistry, Chinese Academy of Sciences, Beijing 100190, P. R. China. E-mail: hanbx@iccas.ac.cn

^bUniversity of Chinese Academy of Sciences, Beijing 100049, China

^cChemistry and Chemical Engineering of Guangdong Laboratory, Shantou 515063, China

^dPhysical Science Laboratory, Huairou National Comprehensive Science Center, Beijing 101400, China

^eShanghai Key Laboratory of Green Chemistry and Chemical Processes, School of Chemistry and Molecular Engineering, East China Normal University, Shanghai 200062, China

^fInstitute of High Energy Physics, Chinese Academy of Sciences, Beijing 100049, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1sc00042j

coordination environments, and active crystal facets can alter the CO adsorption and C–C coupling steps during the CO₂RR.^{34–37} These factors are often associated rather than independent. For example, the active facets tend to be exposed on the surface in regions near grain boundaries, as the grain boundaries could stabilize facets with high surface energies, according to solid-state mechanical studies.^{38,39} Thus, the surface structure should also be comprehensively studied under the reaction conditions.

In this work, the surface species of OD-Cu catalysts were systematically studied *via in situ* surface-enhanced Raman spectroscopy (SERS), operando X-ray absorption spectroscopy (XAS), and isotope labeling experiments. It was found that CuO_x species existed on Cu surfaces during the CO₂RR, and they arose from the adsorption of preliminary intermediates (such as *CO₂ or *OCO[−]) on metallic Cu(0) at negative potentials rather than residual oxides. The presence of CuO_x species was unlikely to be the factor for enhancing C₂+ product formation. Detailed experimental studies indicate that the high density of interfaces between facets played a key role in the highly efficient C₂+ product production.

Results and discussion

Two different OD-Cu catalysts with similar morphologies were prepared in this work. Firstly, an oxide-derived Cu nanorod sample (Cu-nr) was prepared *via* the electroreduction of CuO nanorods. Another OD-Cu catalyst was prepared *via* the simple redox cycling treatment of Cu-nr. Cu-nr oxide (Cu-nr-O) was prepared *via* the oxidation of Cu-nr in 1.0 M KOH solution, then, reduced Cu-nr-O (Cu-nr-OR) was obtained *via* the electroreduction of Cu-nr-O (Fig. S1†). The obtained Cu-nr had a diameter of about 40 nm (Fig. S2†), and exhibited a cross-linked architecture. Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) studies suggested that Cu-nr-OR also exhibited nanorod morphology, which was similar to Cu-nr (Fig. 1A and B). The diameters of Cu-nr and Cu-nr-OR were about 40 nm. From high-resolution TEM images (Fig. 1C and S3†), the corresponding lattice spacings of CuO (001), Cu₂O (111), Cu₂O (110), and Cu (111) were observed for Cu-nr-OR, which may be because Cu can be oxidized by air before measurements. In addition, a small amount of Cu₂O can also be observed based on powder X-ray diffraction (XRD) and X-ray photoelectron spectroscopy (XPS) analysis (Fig. S4†). Furthermore, we can find that Cu-nr-OR comprised a lot of nanocrystals, and abundant interfaces between the nanocrystals were produced. Similar interfaces were found in Cu-nr (Fig. S5†), albeit the density of interfaces is significantly lower than that in Cu-nr-OR. This may be due to the size of the nanocrystals in Cu-nr being about twice that in Cu-nr-OR. Also, Cu nanoparticles (Cu-np) with a diameter of 50 nm were selected as a contrasting material. We can observe that Cu-np mostly exhibited single facets with few interfaces (Fig. S6†). These results indicated that OD-Cu exhibited abundant interfaces, which was consistent with previous reports.^{21–23}

In order to obtain detailed information regarding Cu speciation, X-ray absorption spectroscopy (XAS) was used to explore

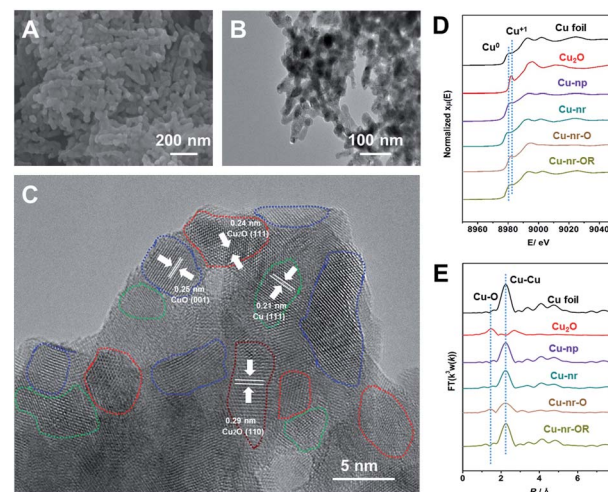


Fig. 1 Characterization of the Cu-nr-OR catalyst. (A) An SEM image of Cu-nr-OR. (B and C) TEM and HR-TEM images of Cu-nr-OR. The areas delineated by blue, red, green, and brown dashed lines denote the CuO (100), Cu₂O (111), Cu (111), and Cu₂O (110) facets, respectively. (D) XANES spectra at the Cu K-edge for different catalysts. (E) The corresponding Fourier transform (FT($k^3w(k)$)) EXAFS spectra.

the electronic structures of the catalysts. The X-ray absorption near-edge structure (XANES) spectra (Fig. 1D) showed that the pre-edge peak of Cu-nr-O was close to Cu₂O. The oscillation $k^3\chi(k)$ functions of Cu-nr-O (Fig. S7†) indicated that the low- k region was similar to Cu₂O and the high- k region was similar to Cu, indicating that both Cu₂O and Cu existed in Cu-nr-O. According to the extended X-ray absorption fine structure (EXAFS) spectra (Fig. 1E), Cu–O and Cu–Cu coordination peaks were observed in Cu-nr-O, implying that some Cu was oxidized to Cu₂O. In contrast, for Cu-nr and Cu-nr-OR, only peaks corresponding to metallic Cu were observed, indicating that CuO and Cu-nr-O were fully reduced to metallic Cu after electroreduction.

The electrocatalytic performances of the catalysts were evaluated in a flow cell, as reported in our previous work.⁴⁰ Gas chromatography (GC) and nuclear magnetic resonance (NMR) spectroscopy were used to analyze the gaseous and liquid products, respectively (Fig. S8 and S9†). For Cu-nr, the FE for C₂+ products (FE_{C₂+}) was 64.6% at −0.9 V *vs.* RHE (Fig. 2A). However, for Cu-np, FE_{C₂+} was only 45.2%. It is interesting to note that Cu-nr-OR exhibited the highest FE_{C₂+} value among the three catalysts, and FE_{C₂+} could reach up to 83.8% with a current density of 341.5 mA cm^{−2} at −0.9 V *vs.* RHE, which was among the best values reported to date (Table S1†). The partial current density for C₂+ products over Cu-nr-OR could reach 286.2 mA cm^{−2} at −0.9 V *vs.* RHE, which is about 1.6 and 2.3 times that over Cu-nr and Cu-np, respectively (Fig. 2B). Although FE_{C₂+} over Cu-np was significantly lower than that over Cu-nr-OR, the FE for H₂ was slightly higher than that over Cu-nr-OR (Fig. S10†), indicating that the increase in C₂+ products for OD-Cu did not mainly arise from the suppression of the hydrogen evolution reaction (HER). We can also observe that the FE for CO over Cu-np was significantly higher than those



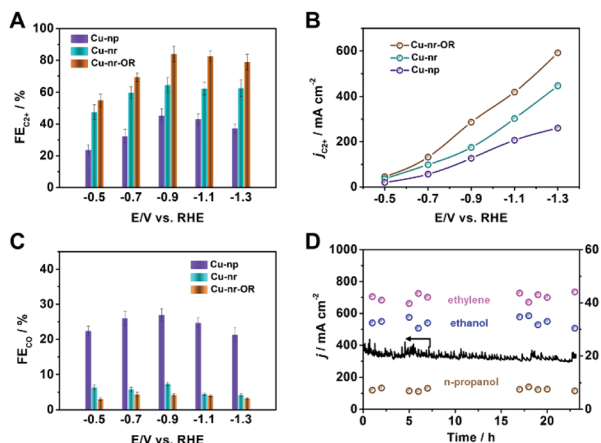


Fig. 2 (A) The FEs for C2+ products over different catalysts in 1 M KOH solution. (B) The partial current densities for C2+ products over different catalysts in 1 M KOH solution. (C) The FEs for CO over different catalysts in 1 M KOH solution. (D) The long-term stability of Cu-nr-OR at -0.9 V vs. RHE for 24 h.

over Cu-nr and Cu-nr-OR (Fig. 2C), indicating that generated CO can easily desorb from Cu-np, and C-C coupling can be hindered. Thus, we can assume that the low selectivity for C2+ products over Cu-np was mainly due to weak CO adsorption. The distributions of C2+ products are listed in the ESI (Fig. S11–S13†). In addition, the sums of the FEs for the different catalysts were close to 100% (Fig. S14†). Long-term operation was conducted at -0.9 V vs. RHE over Cu-nr-OR to elucidate the electrode stability, and the electrolyte was refreshed every hour to overcome the issue of salt accumulation.⁹ There was no obvious decay in either the current density or the FEs of the products during 24 h (Fig. 2D).

In order to verify that the products were derived from CO₂, electrolysis experiments were conducted using isotope-labeled ¹³CO₂ over Cu-nr-OR. From ¹H NMR spectra (Fig. S15†), we can see that the H signals of the products split into two groups of peaks, resulting from H-¹³C coupling effects.^{7,41} In addition, we can observe that the intensities of the product signals increased with reaction time. These results indicate that CO₂ is the only source of carbon in the products.

The intrinsic reasons for the enhanced activity and selectivity during CO₂ reduction to C2+ products over Cu-nr-OR were further investigated. The electrochemical active surface areas (ECSAs) of Cu-np, Cu-nr, and Cu-nr-OR were estimated *via* measuring the double-layer capacitance and using the Pb underpotential deposition (Pb_{UPD}) method (Fig. S16 and S17†). We can observe that the ECSA of Cu-nr-OR was similar to that of Cu-nr, and it was higher than that of Cu-np. Similar to the geometric current density trend, Cu-nr-OR also exhibited the highest normalized current density among the three samples (Fig. S18 and Table S2†). Thus, the improved C2+ product generation did not mainly result from a change in the ECSA. Furthermore, electrochemical impedance spectroscopy (EIS) was carried out to measure the charge transfer resistance (R_{ct}) values of Cu-nr and Cu-nr-OR. Cu-nr-OR showed a similar interfacial R_{ct} value as Cu-np and Cu-nr (Fig. S19†).

Due to the C2+ product selectivity and activity being sensitive to the surface species of catalysts, the role of surface species in enhancing the generation of C2+ products was explored *via in situ* SERS (Fig. S20†). For Cu-np, Cu-nr, and Cu-nr-OR, no bands corresponding to Cu oxides (CuO or Cu₂O) can be observed at -0.2 V vs. RHE or below, indicating the full reduction of Cu oxides to metallic Cu (Fig. 3A–C); this was consistent with the results of XAFS. In addition, from cyclic voltammetry curves (Fig. S21†), no oxide peaks can be observed at low potentials. New bands appeared at -0.2 V vs. RHE or below during the CO₂RR, and the bands will be analyzed according to the different potentials below.

At -0.2 V vs. RHE or below, well-defined bands appeared at 1064 cm^{-1} for Cu-np, Cu-nr, and Cu-nr-OR during the CO₂RR, which were attributed to CO₃^{2−},⁴² because CO₂ can dissolve in KOH electrolyte, forming a neutral-pH carbonate mixture. The intensity of the band became weaker as the potential decreased, which is because CO₂ can be reduced at negative potential and the formation of CO₃^{2−} becomes slow.

At -0.3 V vs. RHE or below, an additional band appeared at 524 cm^{-1} for Cu-np, Cu-nr, and Cu-nr-OR during the CO₂RR. According to a previous report,⁴² the band can be assigned to CuO_x or CuO_xOH_y species. When analysis was conducted in D₂O instead of H₂O, the band at 524 cm^{-1} displayed a negligible shift (Fig. S22†), indicating that the band could be attributed to CuO_x species rather than CuO_xOH_y.⁴² There are three possible reasons for the formation of CuO_x species (Fig. 3E): (1) the CuO_x species were from the original Cu₂O; (2) they arose from

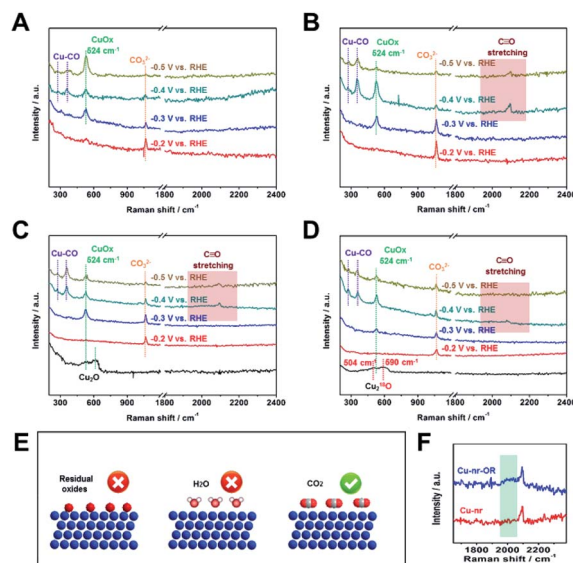


Fig. 3 The *in situ* surface-enhanced Raman spectra of Cu-np (A), Cu-nr (B), and Cu-nr-OR (C) at different potentials during the CO₂RR. (D) The *in situ* surface-enhanced Raman spectra of ¹⁸O-enriched Cu-nr-OR at different potentials during the CO₂RR. (E) A schematic illustration of the possible reasons for the formation of CuO_x species during the CO₂RR. (F) A local enlarged view of the *in situ* surface-enhanced Raman spectra of Cu-nr and Cu-nr-OR at -0.4 V vs. RHE during the CO₂RR.

a reaction between reduced Cu and H₂O; and (3) they arose from a reaction between reduced Cu and CO₂.

To answer this question, ¹⁸O isotope labeling experiments were carried out to confirm the source of oxygen in CuO_x. First, ¹⁸O-enriched Cu-nr-O catalysts were synthesized *via* the oxidation of Cu-nr in H₂¹⁸O electrolyte. Cu₂¹⁸O was formed in Cu-nr-O; the bands were at 504 and 590 cm⁻¹ (Fig. 3D), exhibiting significant red shifts compared with Cu₂¹⁶O.²⁹ This result indicates that the bands of CuO_x species showed significant red shifts when ¹⁶O was replaced with ¹⁸O, which can be an important indicator when exploring the oxygen source in the CuO_x species. For ¹⁸O-enriched Cu-nr-OR, CuO_x was also formed at -0.3 V *vs.* RHE, and the band remained at 524 cm⁻¹, displaying a negligible shift compared with that of Cu-nr-OR. The results suggest that the CuO_x species were not from the original Cu₂¹⁸O. Thus, it can be deduced that CuO_x was produced during the CO₂RR. Because both H₂O and CO₂ can react with reduced Cu to form CuO_x species, the oxygen source of the CuO_x species should be further studied. Furthermore, Cu-nr-OR was tested in H₂¹⁸O during the CO₂RR; we observed that the band of CuO_x was still at 524 cm⁻¹ (Fig. S23†), which indicated that CuO_x did not arise from a reaction between Cu and H₂¹⁸O. Thus, the last possible reason is reasonable, *i.e.*, CuO_x arose from a reaction between Cu and CO₂. To further verify this argument, electrolysis experiments using Cu-nr-OR were carried out under a N₂ atmosphere; no bands were observed at a negative potential (Fig. S24†), indicating that CO₂ played a key role in the formation of CuO_x species. In addition, the CuO_x band cannot be observed at -0.2 V *vs.* RHE, and it appeared at -0.3 V *vs.* RHE or below. Thus, we can assume that this peak cannot be attributed to the simple adsorption of CO₂ on Cu. CO₂ can be reduced to preliminary intermediates at low potentials; thus, we can assume that the CuO_x species were from the adsorption of preliminary intermediates (such as *CO₂ and *OCO⁻) on Cu, which has been observed on Ag metal.⁴³

From the above results, we can deduce that CuO_x that existed during the CO₂RR was just a signal of the chemisorption of CO₂ on Cu, which was not the main factor for facilitating C2+ product formation during the CO₂RR. And the CuO_x species were not specific to the Cu-oxide derived catalysts.

At -0.4 V *vs.* RHE or below, for Cu-np, Cu-nr, and Cu-nr-OR, the presence of adsorbed *CO on Cu was demonstrated by the appearance of Raman peaks located at 276, 360, and 2000–2100 cm⁻¹, which correspond to the restricted rotation of adsorbed *CO on Cu, Cu–CO stretching, and C≡O stretching, respectively.⁴⁴ We can observe that these peaks were significantly weaker for Cu-np than for Cu-nr and Cu-nr-OR, which can result from the weak adsorption of CO or rapid reactions on Cu-np. From the results of the CO₂RR (Fig. 2A and C), we know that the FE for CO over Cu-np was higher than those over Cu-nr and Cu-nr-OR, and the FE for C2+ products over Cu-np was lower than those over Cu-nr and Cu-nr-OR; thus, we can assume that the weak peaks on Cu-np could be attributed to the weak adsorption of CO on Cu-np. It is interesting to note that the C≡O stretching band of Cu-nr-OR is different from that of Cu-nr (Fig. 3F). A new peak appeared at about 2000 cm⁻² in the case of Cu-nr-OR compared with Cu-nr. Specifically, the stretching

mode of surface-adsorbed CO can serve as a molecular probe of the surface structure due to its sensitivity to the structures of adsorption sites.^{45,46} It is reasonable to analyze the active sites using surface-adsorbed CO at -0.4 V *vs.* RHE, because the FE for C2+ products is low (Fig. S25†), indicating that the C–C coupling step is slow at this potential. Thus, we can assume that new adsorption sites for CO were produced on Cu-nr-OR. Also, the new sites exhibited slightly weaker adsorption for CO than Cu-nr, due to the band having a lower wavenumber than that of Cu-nr. Therefore, Cu-nr-OR exhibited heterogeneous CO adsorption sites, which included stronger binding sites and weaker binding sites. According to previous reports,^{47,48} the energy barrier of C–C coupling can be decreased on stronger and weaker CO binding sites. Thus, we can suppose that the enhanced generation of C2+ products over Cu-nr-OR originated mainly from the formation of heterogeneous CO adsorption sites. Combining the HR-TEM results, we can assume that the heterogeneous CO adsorption sites could be attributed to the abundant interfaces of Cu-nr-OR.

After the potential was removed, we could observe that Cu₂O was formed rapidly (Fig. S26†), indicating that the reduced Cu can be oxidized in electrolyte. Cu₂O can be formed *via* the oxidation of Cu by O₂ in the electrolyte.²⁹ However, the O₂ content is very low in the cathodic electrolyte. It is interesting to note that Cu₂¹⁸O was formed when using H₂¹⁸O as the electrolyte after the potential was removed (Fig. S27†). Thus, we can assume that Cu₂O arose from the reaction between Cu and H₂O. It is known that the oxidation of Cu by water is a thermodynamically unfavorable process. Thus, only a small proportion of Cu can be oxidized. The results indicate that the *in situ* method has obvious advantages and is necessary when exploring the real state of a catalyst during a reaction.

Because the coordination environment of Cu can alter the adsorption of CO and the energy barrier of the C–C coupling step, we used operando XAS to monitor the local structures of Cu-np, Cu-nr, and Cu-nr-OR during the CO₂RR (Fig. S28†). For

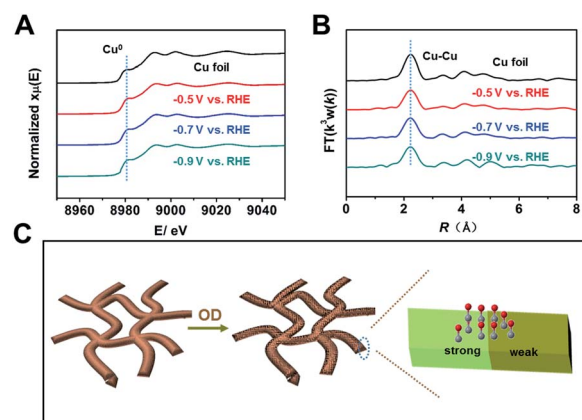


Fig. 4 (A) Operando XANES spectra at the Cu K-edge of Cu-nr-OR at different potentials during the CO₂RR. (B) The corresponding Fourier transform (FT($k^3w(k)$)) EXAFS spectra of Cu-nr-OR at different potentials during the CO₂RR. (C) A schematic illustration of the reason for the outstanding C2+ production performance over Cu-nr-OR.



both catalysts, only peaks corresponding to metallic Cu were observed at negative potentials during the CO₂RR (Fig. 4A, B and S29–S31†), indicating that CuO or Cu₂O was reduced to metallic Cu in the CO₂RR. Moreover, the quantified Cu–Cu coordination numbers of Cu-nr-OR, Cu-nr, and Cu-np were fit using the ARTEMIS programs of IFEFFIT during the CO₂RR (Fig. S32–S34 and Table S3†). No obvious differences in the Cu–Cu coordination numbers were observed for Cu-nr-OR, Cu-nr, and Cu-np during the CO₂RR, indicating that the enhanced C₂+ product formation over Cu-nr-OR did not mainly arise from the slight change in the coordination environment.

From the above results, we can make the conclusion that abundant interfaces can be produced on Cu-nr-OR, and stronger and weaker binding sites for CO appeared at the abundant interfaces, which can decrease the energy barrier of C–C coupling (Fig. 4C). Also, the coordination environment of Cu-nr-OR does not show obvious changes compared with Cu-nr. Thus, the increased interfaces can be considered as the main factor for the enhanced generation of C₂+ products over Cu-nr-OR.

Conclusions

In summary, the surface species and structures of OD-Cu catalysts were systematically investigated *via in situ* SERS, operando XAS, and ¹⁸O isotope labeling experiments. It was shown that Cu oxides indeed existed on the surfaces of the catalysts during the CO₂RR. However, they were formed *via* the adsorption of preliminary intermediates (such as *CO₂ and *OCO[−]) on Cu instead of on the active sites of the catalyst. Abundant interfaces can be produced on Cu-nr-OR, which can provide heterogeneous CO adsorption sites, enhancing the C₂+ product selectivity. In addition, this work also shows that *in situ* techniques have obvious advantages and are sometimes necessary for exploring the structures of catalysts and the surface species during the CO₂RR. We believe that the findings of this work provide useful knowledge for designing other efficient electrocatalysts for CO₂ reduction.

Author contributions

Chunjun Chen and Buxing Han proposed the project, designed the experiments and wrote the manuscript; Chunjun Chen performed the whole experiments; Xupeng Yan, Yahui Wu, Huizhen Liu, Xiaofu Sun, Qinggong Zhu, Tianbin Wu and Qingli Qian assisted in analyzing the experimental data; Fengrong Juan assisted in analyzing the experimental data of *in situ*-SERS; Shoujie Liu, Jing Zhang and Lirong Zheng assisted in analyzing the experimental data of XAS; Buxing Han supervised the whole project.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors thank the National Key Research and Development Program of China (2017YFA0403101, 2017YFA0403003 and 2017YFA0403102), National Natural Science Foundation of China (22003070, 21890761, 21733011), Beijing Municipal Science & Technology Commission (Z191100007219009), and Chinese Academy of Sciences (QYZDY-SSW-SLH013). The operando X-ray adsorption spectroscopy (XAS) measurements were performed using a flow cell at the 1W1B and 1W2B beamlines at the Beijing Synchrotron Radiation Facility (BSRF), China.

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