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Bifunctional activation of amine-boranes by the W/Pd bimetallic analogs of “frustrated Lewis pairs”†

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The reaction between basic [(PCP)Pd(H)] (PCP = 2,6-(CH₂P(t-C₄H₉)₂)₂C₆H₄) and acidic [LWH(CO)₃] (L = Cp (1a), Tp (1b); Cp = η⁵-cyclopentadienyl, Tp = κ³-hydridotris(pyrazolyl)borate) leads to the formation of bimolecular complexes [LW(CO)₂(μ-CO)⋯Pd(PCP)] (4a, 4b), which catalyze amine-borane (Me₂NHBH₃, ^tBuNH₂BH₃) dehydrogenation. The combination of variable-temperature (¹H, ³¹P{¹H}, ¹¹B NMR and IR) spectroscopies and computational (ωB97XD/def2-TZVP) studies reveal the formation of an η¹-borane complex [(PCP)Pd(Me₂NHBH₃)⁺[LW(CO)₃]⁻] (5) in the first step, where a BH bond strongly binds palladium and an amine group is hydrogen-bonded to tungsten. The subsequent intracomplex proton transfer is the rate-determining step, followed by an almost barrierless hydride transfer. Bimetallic species 4 are easily regenerated through hydrogen evolution in the reaction between two hydrides.

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Introduction

Catalytic processes occurring under the action of “frustrated Lewis pairs” (FLPs) have been intensively sought after during the past decade.^{1,2} Heterolytic cleavage of H₂ by main group intermolecular FLPs has been proposed to occur through an ‘encounter complex’ where the Lewis acid and Lewis base are in close proximity, but non-bonding, and H₂ is accommodated between the two centers prior to heterolysis. DFT metadynamics studies for the prototypical FLP system P(Mes)₃/B(C₆F₅)₃ gave evidence for the H₂ polarization followed by the rate-determining hydride transfer to B and proton transfer to P.³

The heterolytic cleavage of H₂ into a proton and a hydride is a crucial step in many chemical and biochemical processes such as, e.g., hydrogen oxidation by hydrogenases, transition metal-catalyzed hydrogenation of C=O bonds, or metal-catalyzed hydrogen oxidation in energy conversion reactions. From this point of view, the nomenclature of FLPs has many parallels in metal–ligand cooperation (bifunctional) catalysis, where the two sites of cooperation are typically metal- and ligand-based.^{4,5} Closely related to this concept is also an area of

bimetallic catalysis, wherein two metal sites demonstrate cooperativity in fundamental catalytic reactions.⁶

The species featuring an H₂ molecule poised between the Lewis acid and Lewis base centers have been found by us in previous work as the intermediate of dihydrogen evolution in the reaction of two neutral transition metal hydrides of different polarities. Dihydrogen bonding (DHB, M–H^{δ-}⋯^{δ+}H–M’; M–H = (PCP)NiH, (PCP)PdH; M’–H = CpWH(CO)₃ (1a); PCP = 2,6-C₆-H₃(CH₂P^tBu₂)₂) precedes the concerted proton and hydride transfer,⁷ eventually yielding the bimetallic isocarbonyl bridging M/W bimetallic complex 4 (Scheme 1).^{8,9} Up to date, there are only a few examples of the attempted use of such systems in catalysis. CpMoH(CO)₃ has been used as a proton source in Rh-catalyzed hydroformylation.¹⁰ More recently, the switch of selectivity between migratory insertion *versus* C–H activation has been shown for the reaction of [(ⁱPrPOCOP)Ni]⁺ (ⁱPrPOCOP = 2,6-C₆H₃(OPⁱPr₂)₂) with phenylacetylene in the presence of [CpW(CO)₃]⁻ presumably acting as a proton shuttle.¹¹ However, this reaction was performed only in a stoichiometric regime.

In the field of chemical hydrogen storage, lightweight B/N-containing inorganic hydrides (ammonia borane, amine-boranes, and hydrazine boranes) have gained increasing attention in the last few years as sources of carbon-free and high-purity H₂ for renewable energy applications.^{12–14}

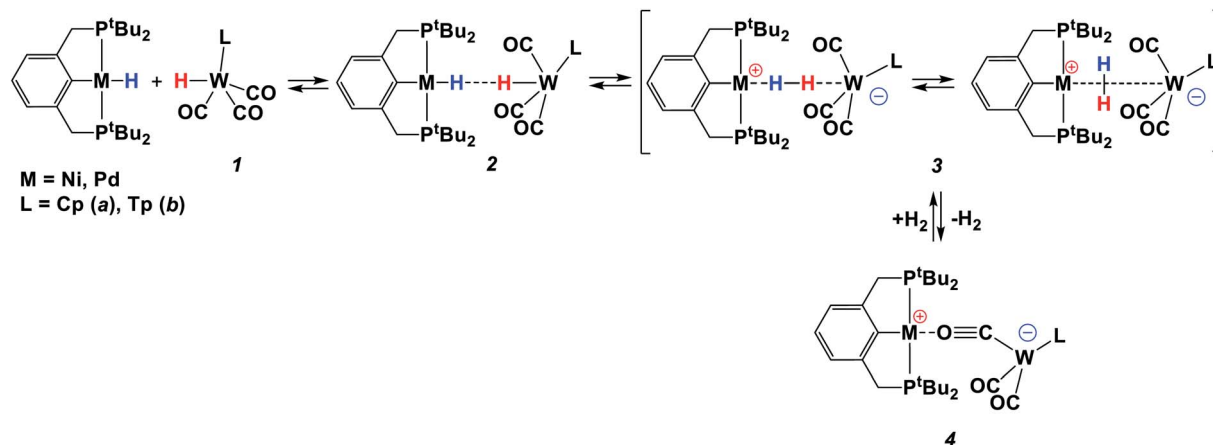
Metal-catalyzed dehydrogenation became a well-developed route for H₂ production from amine-boranes¹⁵ and new studies keep appearing in the literature aiming to optimise the reaction conditions and to control the selectivity in product distribution as dehydrogenative coupling not only produces H₂

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Scheme 1 Reaction between two metal hydrides, (PCP)PdH and 1.

as an energy source but also B–N-containing materials.^{16,17} The metal-free catalysts including FLPs are also known but far less numerous.¹⁸ Although direct comparison between different catalysts is somewhat ambiguous due to the difference in the reaction conditions, RuH(PMe₃)₂{N(CH₂CH₂PⁱPr₂)₂} can be considered as the best catalyst for BH₃NH₃ dehydrogenation (TOF_{max} 72 000 h⁻¹)¹⁹ and (POCOP)IrH₂ (TOF 2400 h⁻¹)¹² and

Rh(Xantphos-ⁱPr)H (1500 h⁻¹)²⁰ for BH₃NMe₂ dehydrogenation, whereas a cationic zirconocene–phosphinoaryloxy complex [Cp₂ZrOC₆H₄PⁱBu₂]⁺ which can be described as an early transition-metal-containing FLP gives TOF 600 h⁻¹ for BH₃NMe₂H dehydrogenation.²¹

The simultaneous presence of protic and hydridic H atoms within the same amine–borane molecule makes them

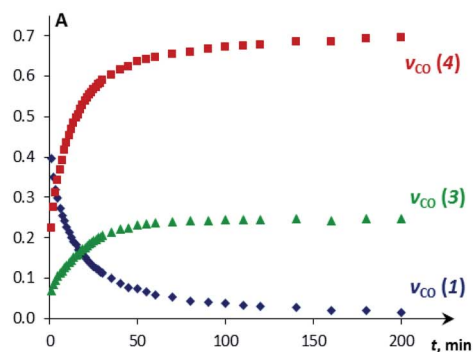
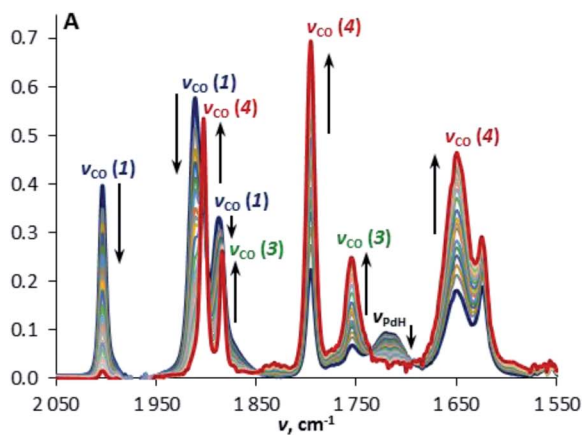


Fig. 1 IR spectra (top) and the corresponding kinetic curves (bottom) of a mixture of TpWH(CO)₃ (1a, c = 0.003 M) and (PCP)PdH (c = 0.045 M) at 240 K in THF.

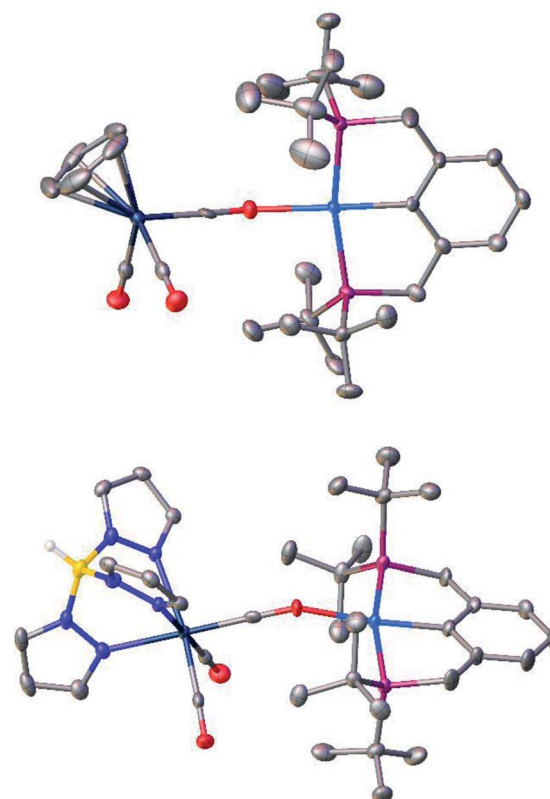


Fig. 2 Molecular structures of bimetallic products 4a (top) and 4b (bottom). Atom color code: white, H; gray, C; yellow, B; red, O; blue, N; navy, W; pink, P; and cyan, Pd. Thermal ellipsoids are given at 50% probability level; hydrogen atoms except for the BH group in 4b are omitted for clarity.



Table 1 ν_{CO} vibrations (cm^{-1}) of the metal complexes in THF and toluene at 298 K

Complex	ν_{CO} , cm^{-1} in THF	ν_{CO} , cm^{-1} in toluene
1a CpWH(CO) ₃	2018, 1924	2020, 1926
1b TpWH(CO) ₃	2003, 1910, 1888	2006, 1914, 1891
3a, 5a [CpW(CO) ₃] [−]	1891, 1775	1887, 1782
3b, 5b [TpW(CO) ₃] [−]	1884, 1754	1862, 1767
4a [CpW(CO) ₂ (μ-CO)⋯Pd(PCP)]	1910, 1819, 1660	1907, 1814, 1656
4b [TpW(CO) ₂ (μ-CO)⋯Pd(PCP)]	1901, 1796, 1650	1904, 1798, 1644

alternative reagents in transfer hydrogenation which attracts increased research interest.²² It also triggers an extended dihydrogen bonding (DHB) throughout their solid-state structure.²³ DHB is preliminary to H₂ evolution occurring after a simple thermal treatment or in a catalytic fashion.^{15,24} Given the analogy between protic and hydridic H atom co-existence and heterolytic H₂ cleavage, the latter is intimately related to B/N inorganic hydride activation. Herein, we show for the first time that the bimetallic systems reported above act as homogeneous amine-borane dehydrogenation catalysts and explore the impact of hydride and proton transfer steps in the reaction mechanism using **1a** in comparison with its more acidic analog TpWH(CO)₃ (**1b**; Tp = κ³-hydridotris(pyrazolyl)borate) as proton donating hydrides (Scheme 1).

Results and discussion

Pairwise interaction of two metal hydrides

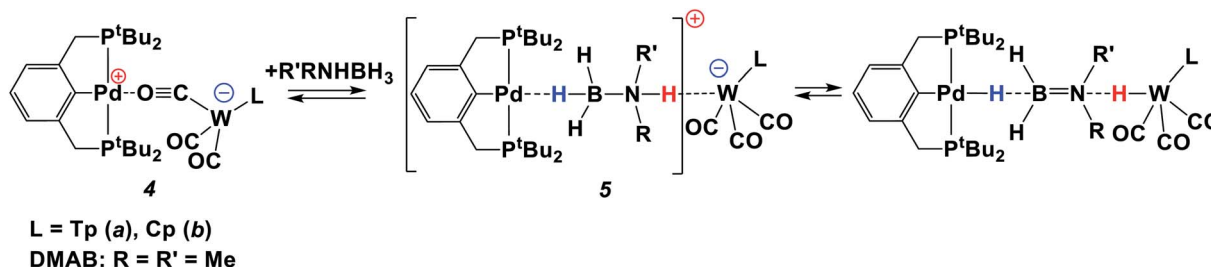
The reaction of TpWH(CO)₃ (**1b**) with (PCP)PdH follows the same mechanism (Scheme 1) as that of CpWH(CO)₃ (**1a**). The IR spectra obtained for the mixture of **1b** with a 1.5-fold excess of (PCP)PdH in pure THF at 200 K show two new ν_{CO} bands of the intermediate **3b** at 1884 and 1754 cm^{-1} , which disappear gradually upon temperature increase (Scheme 1 and Fig. S1†). At the same time, the bands at 1650, 1796, and 1901 cm^{-1} of the H₂ evolution product – isocarbonyl complex **4b** – increase in intensity. In a closed system, the reaction goes to equilibrium (Fig. 1). When the reaction mixture was allowed to reach equilibrium at, e.g., 240 K and then warmed to 290 K, the intensity decrease of the ν_{CO} bands of the ionic intermediate **3b** and a slight increase of the ν_{CO} bands of the initial hydride **1b** are observed. This indicates the shift of the proton transfer

equilibrium **1** ↔ **3** to the left. At the same time, the ν_{CO} intensity of complex **4b** increases due to the hydrogen evolution and right shift of the **3** ↔ **4** equilibrium (Fig. S2†).²⁵ The IR monitoring of the reaction kinetics showed that the proton transfer is the rate-determining step of the overall process and the rate constant values are larger for TpWH(CO)₃ (**1b**) than those with CpWH(CO)₃ (**1a**): at 240 K $k_{\text{obs}} = 0.202 \text{ mol}^{-1} \text{ s}^{-1}$ for **1b** and $0.140 \text{ mol}^{-1} \text{ s}^{-1}$ for **1a**. This gives Gibbs free energy values of $\Delta G_{298 \text{ K}}^{\ddagger} = 16.5 \pm 0.1 \text{ kcal mol}^{-1}$ for **1b** and $16.9 \pm 0.1 \text{ kcal mol}^{-1}$ for **1a**. Complexes **4** reversibly capture molecular hydrogen similarly to FLPs. In a closed system, the reverse cooling from 290 K to 240–200 K restores the original spectroscopic picture. The crystal structures of both bimetallic complexes **4** (Fig. 2) appeared to be very similar to those reported for the related Ni-containing complexes.^{8,9,11} The larger steric volume of the Tp ligand leads to a different orientation of the [TpW(CO)₃] fragment: rotation around the Pd–O bond places terminal CO ligands on one side of the [(PCP)Pd] plane.

Interaction with dimethylamine-borane (DMAB)

Since the bifunctional ionic pairs [LW(CO)₂(μ-CO)⋯M(PCP)] (**4**) reversibly bind hydrogen, we hypothesised that they could interact with amine-boranes which are typical bifunctional molecules (RR'NHBH₃; R = H, Me; R' = H, Me, ^tBu). We used substituted amine-boranes for these studies because they produce boron-containing reaction products and intermediates soluble in organic solvents.²⁶ Besides, Me₂NHBH₃ usually reacts slower than ammonia borane (NH₃BH₃).²⁷

The addition of excess Me₂NHBH₃ (DMAB, 3.3 equiv.) to the bimetallic complexes **4** in THF at 270 K disrupts the inter-ion interaction due to the simultaneous coordination of DMAB molecules between the two metals. The decrease of ν_{CO} bands of **4** and the appearance of two new ν_{CO} bands of a new ionic complex **5** (Table 1) are observed in the IR spectra (Fig. S3†). These changes are reversible; the equilibrium shifts toward **5** upon cooling. Changing the solvent to less polar toluene allows the reaction to move one step forward: the bands of LW(CO)₃ (**1**) and (PCP)PdH ($\nu_{\text{PdH}} 1717 \text{ cm}^{-1}$) appear in the IR spectrum (Fig. S4†), confirming the occurrence of proton and hydride transfer and the formation of neutral molecules. The hydride transfer from the boron atom to palladium is confirmed by the synchronous decrease of the DMAB band intensity ($\nu_{\text{BH}} 2364 \text{ cm}^{-1}$) and the ν_{PdH} growth on going from 260 K to 190 K (Fig. S5†). Warming the mixture from 190 to 260 K restores the



Scheme 2 Reaction between the ionic pair **4** and amine-borane.



original spectral picture, further confirming the existence of the equilibrium shown in Scheme 2. Definitely, the neutral trimolecular complex should be unfavorable due to the high entropy effect (see DFT calculations below), but it is conserved at low temperatures when $T\Delta S$ contribution to the Gibbs energy is diminished.

The behavior of two ionic complexes **4a** and **4b** is qualitatively the same. However, the proton transfer from DMAB to **4a** is easier; the tungsten hydride **1a** is formed in a larger quantity, in agreement with the higher basicity of the $[\text{CpW}(\text{CO})_3]^-$ anion: pK_a of the conjugated acid $\text{LWH}(\text{CO})_3$ in CH_3CN is 16.1 (ref. 28) and 14.4 (ref. 29) for **1a** and **1b**, respectively.

Thus, at low temperature (below 270 K), the DMAB molecule becomes “arrested” between the two metal ions of **4**. The bifunctional nature of **4** allows the proton transfer from the N–H moiety to the Lewis base center on tungsten and the hydride transfer from the B–H moiety to the Lewis acid center on palladium. The position of the equilibrium (Scheme 2) depends on the solvent, temperature, and basicity of the

$[\text{LW}(\text{CO})_3]^-$ anion. Interestingly, neither $(\text{PCP})\text{PdH}$, $\text{LWH}(\text{CO})_3$, $[\text{LW}(\text{CO})_3]^-$ (obtained by the reaction of $\text{LWH}(\text{CO})_3$ with NaHMDS) or $[(\text{PCP})\text{Pd}]^+$ (ref. 27) taken individually shows any activity in DMAB dehydrogenation. This confirms the crucial role played by the bimetallic complex in the process.

Catalytic DMAB dehydrogenation

At ambient temperature in the presence of up to 50-fold excess of DMAB over **4**, the catalytic dehydrogenation reaction occurs. To deeper understand the mechanism, the reaction was followed spectroscopically and volumetrically. IR monitoring shows the progressive decrease of ν_{BH} ($2360\text{--}2260\text{ cm}^{-1}$) and ν_{NH} (3200 cm^{-1}) bands of DMAB in the presence of **4a** or **4b** until their complete disappearance (Fig. S6†). This observation provides strong confirmatory evidence for the cleavage of B–H and N–H bonds. A decrease of the typical DMAB signal ($\delta_{\text{B}} -15.3$) and the accumulation of the dehydrogenation product – *i.e.* the cyclic dimer $(\text{Me}_2\text{NBH}_2)_2$ ($\delta_{\text{B}} 3.4$), as well as the appearance of the steady-state intermediate – aminoborane $\text{Me}_2\text{N}=\text{BH}_2$ ($\delta_{\text{B}} 35.8$, Fig. 3) – was observed in the $^{11}\text{B}\{^1\text{H}\}$ NMR spectra, evidencing the off-metal dimerization mechanism.³⁰

Under these conditions, ^1H NMR spectroscopy shows the initial appearance of metal hydride resonances (δ_{WH} and δ_{PdH} , Fig. S7 and S8†). Then, the tungsten hydride singlet ($\delta_{\text{WH}} -7.4$ and -2.3 for **1a** and **1b**, respectively) disappears, but the palladium hydride triplet ($\delta_{\text{PdH}} -4.2$) still accumulates. $^{31}\text{P}\{^1\text{H}\}$ NMR spectra show the corresponding growth of the $(\text{PCP})\text{PdH}$

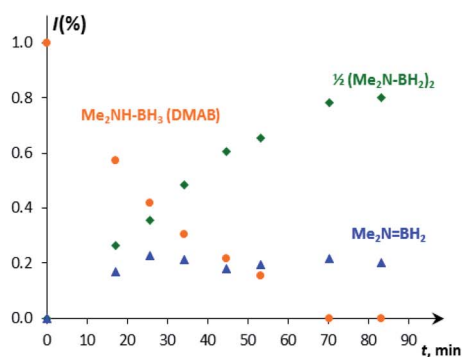
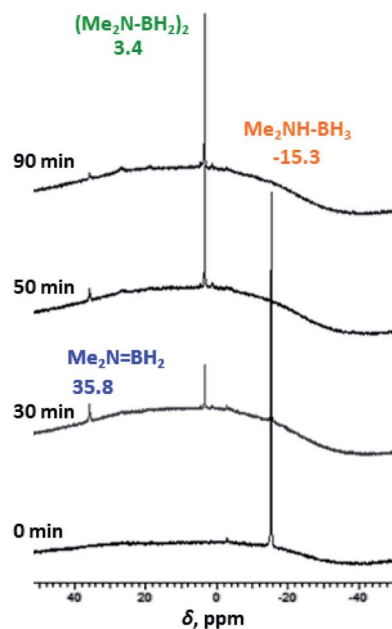


Fig. 3 $^{11}\text{B}\{^1\text{H}\}$ NMR spectra (128.3 MHz, THF-d_8) of DMAB (5 equiv.) in the presence of **4a** (top) and the corresponding changes in molar fractions in time (bottom).

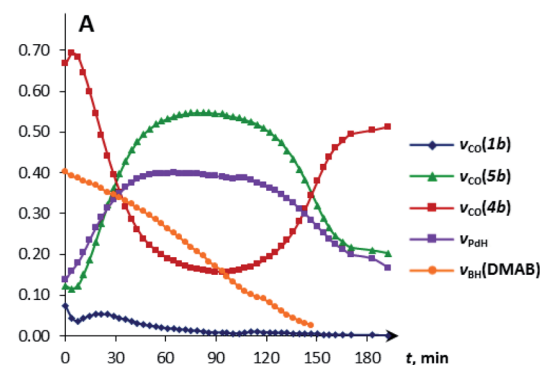
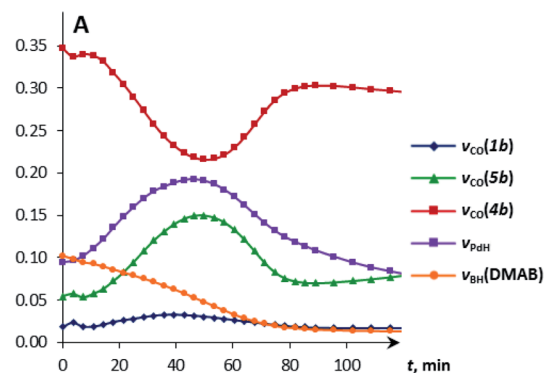


Fig. 4 Kinetic curves obtained from the IR spectra of DMAB (1 equiv. (top) and 5 equiv. (bottom)) dehydrogenation catalyzed by **4b** ($c = 0.003\text{ M}$, 298 K, THF).



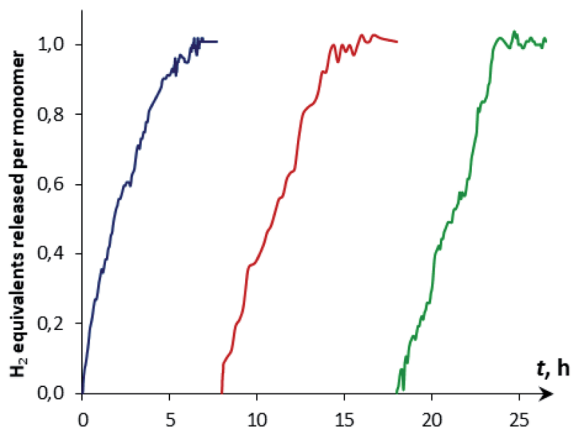


Fig. 5 Kinetic curves of DMAB ($c = 0.06$ M) dehydrogenation by reused $[\text{CpW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ (**4a**, 5 mol%, $c = 0.003$ M) in THF at 313 K.

signal ($\delta_{\text{p}} 92.0$), whereas the bimetallic complex ($\delta_{\text{p}} 75.6$) is first consumed and then restored (Fig. S9[†]). Changes in the IR spectra show the same trend: the ionic complex **5** is formed at the expense of **4**, whereas the ν_{PdH} band (1718 cm^{-1}) increases in intensity (Fig. S10 and S11[†]). The bands (ν_{CO} , ν_{PdH}) of **4**, **5** and (PCP)PdH restore their initial intensity after the complete conversion of DMAB (Fig. 4, top). An increase in DMAB loading increases the length of the quasi-stationary stage of the reaction when the concentration of both the catalyst and reaction intermediates is nearly constant (Fig. 4, bottom).

The reaction kinetics was measured following the decrease of ν_{BH} bond intensity of DMAB. The observed initial rate corresponds to the substrate disappearance (Fig. S6[†]). In the presence of excess DMAB, the kinetic curves ($c(\text{DMAB})$ vs. t) are linear (Fig. S12[†]). DMAB dehydrogenation occurs faster in the presence of the bimetallic complex **4a** with the Cp-ligand: the initial reaction rate ν_0 is $5.8 \times 10^{-6}\text{ M s}^{-1}$ for **4a** and $1.7 \times 10^{-6}\text{ M s}^{-1}$ for **4b** at 20 mol% catalyst loading. Keeping the substrate concentration constant, the increase of the catalyst **4a** concentration from 0.003 M to 0.006 M leads to the doubling of

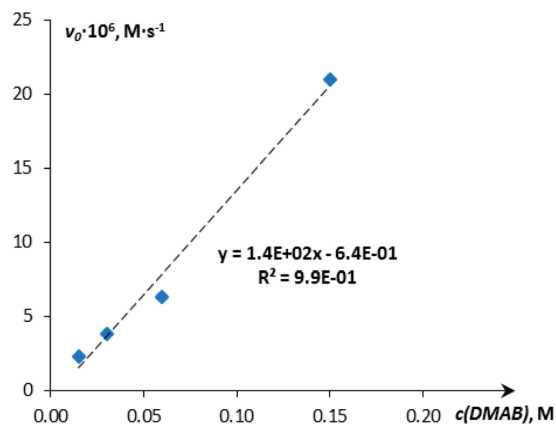


Fig. 6 Dependence of the initial reaction rate on the substrate (DMAB) concentration ($c = 0.015$ M, 0.03 M, 0.06 M, 0.15 M; THF; 313 K).

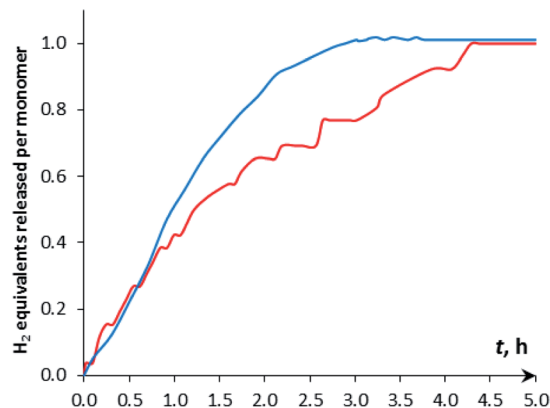


Fig. 7 Kinetic curves of DMAB dehydrogenation by $[\text{CpW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ (**4a**, $c = 0.003$ M) in THF at 313 K. (Red) $c(\text{DMAB}) = 0.015$ M, (5 eq.); (blue) $c(\text{DMAB}) = 0.15$ M, (50 eq.).

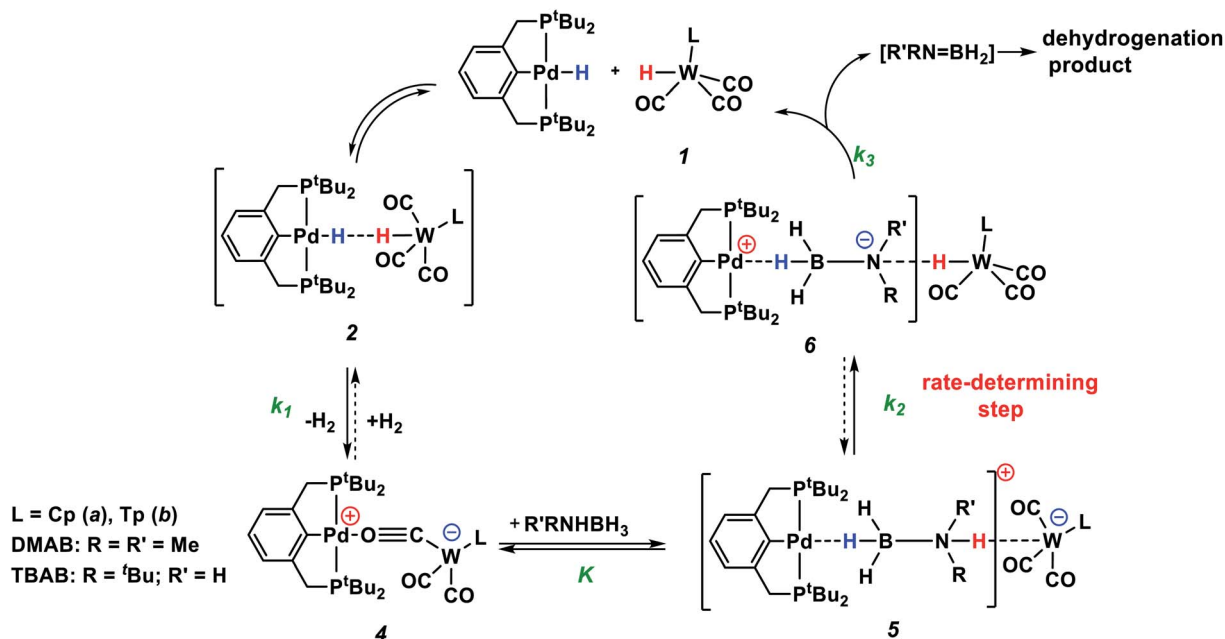
the initial reaction rate. It should be noted that the kinetic curves obtained by the integration of proton NMR spectra have very similar behavior to the IR curves (Fig. S13[†]). According to the spectral data obtained, the bimetallic species $[\text{LW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ (**4**) are regenerated at the end of the catalytic reaction and can be reused for the next substrate loading (Fig. 5). The initial reaction rates ν_0 are similar when DMAB is added to the mixture of two neutral hydrides or to the pre-generated bimetallic complex **4**, revealing that the reagents' mixing sequence does not affect the reaction rate (Fig. S11 and S12[†]).

The H_2 production in the reaction of DMAB with $[\text{CpW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ (**4a**) in THF was monitored using the Man on the Moon X103 kit at ambient temperature and 313 K (Fig. 5 and S14[†]). At 313 K and 1 : 5 catalyst : DMAB ratio, complete DMAB conversion is achieved in 4 h (Fig. 7 and Table S3[†]) with an initial reaction rate of $2.5 \times 10^{-6}\text{ M s}^{-1}$. The increase of substrate loading to 50 equivalents at the same concentration of **4a** ($c = 0.003$ M, 2 mol%; $T = 313$ K) leads to complete DMAB conversion in less than 3 hours (Fig. 7), and the initial reaction rate is $2.1 \times 10^{-5}\text{ M s}^{-1}$. Under these conditions, the TOF value reaches 26 h^{-1} at a half-conversion time (Table S3[†]). The initial reaction rate at a catalyst : DMAB ratio of 1 : 50 is higher than that at 1 : 5, since the increase of DMAB amount at a constant catalyst concentration shifts the equilibrium of complex **5** formation (Scheme 2 and Fig. 6).

Mechanism of $\text{R}'\text{RNH}(\text{BH}_3)$ dehydrogenation by bimetallic ion pairs

According to the experimental data, we suggest the reaction mechanism shown in Scheme 3 involving three important steps: (1) amine-borane molecule insertion between the two units of the bimetallic complex, (2) proton and hydride transfer, yielding aminoborane ($\text{Me}_2\text{N}=\text{BH}_2$) and neutral metal hydrides, (3) dihydrogen release as a result of the reaction between palladium and tungsten hydrides *via* the DHB complex. The insertion of the amine-borane molecule into the bimetallic complex **4** yields the η^1 -borane complex **5** (Scheme





Scheme 3 Mechanism of R'RNHBH₃ dehydrogenation by bimetallic ion pairs 4.

3). The latter is also an ionic pair with the tungsten anion [LW(CO)₃]⁻, which could explain the accumulation of LW(CO)₃⁻ in the reaction mixture (Fig. 4). In the next stage, the proton transfer from the NH group of coordinated amine-borane to tungsten atom gives the neutral hydride **1** and the zwitterionic complex **6**. The hydride transfer from the BH group to the palladium atom inside complex **6** leads to Me₂N=BH₂ elimination and the formation of (PCP)PdH. To close the catalytic cycle, the two neutral metal hydrides react with each other, regenerating the catalytically active complex **4**.

DFT calculations

This mechanism was supported by DFT calculations at the ωB97XD/def2-TZVP theory level taking the real system for the structure optimization in toluene, which was introduced within the SMD model. Complex **5** is 7.7 kcal mol⁻¹ above **4** on the Gibbs free energy scale (Fig. 8); however, its formation is possible (electronic energy difference ΔE(4-5) = -7.5 kcal mol⁻¹). The formation of **5** from **1** + Pd + DMAB is nearly ergoneutral (ΔG(1-5) = +1.9 kcal mol⁻¹) and thus this equilibrium can be effectively shifted toward **5** in the presence of excess DMAB. The key bonds in complex **5** (Fig. 9) are elongated (*r*(Pd-H_B) = 1.89 Å and *r*(B-H_{Pd}) = 1.28 Å) relative to the B-H bond in free DMAB (1.21 Å) and Pd-H in (PCP)PdH (1.64 Å). The amine group of the dimethylamine fragment is hydrogen-bonded to the tungsten anion (H_N-W 3.09 Å, ∠N-H-W 162.6°). Despite the B-H_{Pd} bond elongation, the relaxed potential energy surface (PES) scan for BH bond distance in the complex **5** did not lead to any TS, continuously going up to 50 kcal mol⁻¹. Hydride transfer does not occur at this step, due to the instability of NR₃BH₂⁺, in agreement with the rather low hydride-donating ability of amine-boranes.³¹ This is consistent with the suggested hydride transfer from DMAB to

the cationic Pt-catalyst³² where additional stabilization of [Me₂NHBH₂]⁺ by a strong nucleophile was required. For the DMAB molecule, the proton transfer from nitrogen to tungsten occurs with a barrier ΔG[‡] = 21.8 kcal mol⁻¹ and the formation of the zwitterionic complex **6** and LW(CO)₃H (13.0 kcal mol⁻¹ above complex **5**). B-H bond cleavage in complex **6** is almost barrierless (<2 kcal mol⁻¹) (Fig. S15†). As a result, the hydride transfer goes faster than the proton transfer (*k*₃ ≫ *k*₂); therefore, complex **6** could not be observed experimentally. The ensemble of the reaction products ((PCP)PdH, LW(CO)₃H and H₂B=NR₂) is only slightly higher in free energy than the starting adduct **5** (+1.0 kcal mol⁻¹), and the reaction is still reversible. However, since H₂ evolution from the two hydrides (*k*₁, Scheme 3)^{8,9,33} and the off-metal B=N dimerization^{34,35} are featured with a comparable barrier, the overall reaction of dehydrogenative DMAB coupling becomes irreversible.

Catalytic TBAB dehydrogenation

Dehydrogenation of mono-substituted *tert*-butylamine-borane (^tBuNH₂BH₃, TBAB) was also studied using complex **4a** as a catalyst. However, in this case, the IR spectroscopic picture and kinetic curves are not similar to those with DMAB (Fig. S16†). On the quasi-stationary stage, the concentration of neutral tungsten hydride **1a** is higher, and that of the η¹-borane complex **5a** is lower than in DMAB dehydrogenation. Measured 60 min after TBAB addition (δ_B -21.8), the ¹¹B NMR spectrum shows the formation of the intermediate ^tBuNH=BH₂ (δ_B 36.1, t) and cyclotriborazane [^tBuNH-BH₂]₃ (δ_B -3.7, t) – the product of the evolution of the first H₂ equivalent (Fig. S17†). At the end of the reaction, there is no initial TBAB, and the single product of the evolution of the second H₂ equivalent – borazine [^tBuN-BH]₃ (δ_B 28.1, d) – is observed. The formation of two equiv. H₂



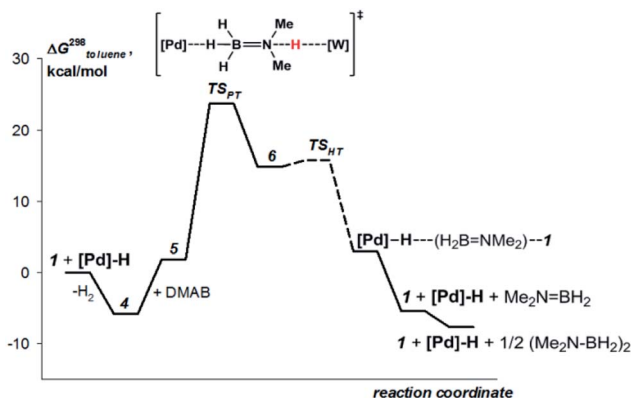


Fig. 8 Computed (DFT/ωB97XD) energy profile for the DMAB dehydrogenation by [TpW(CO)₂(μ-CO)···Pd(PCP)] (4b) in toluene (SMD model).

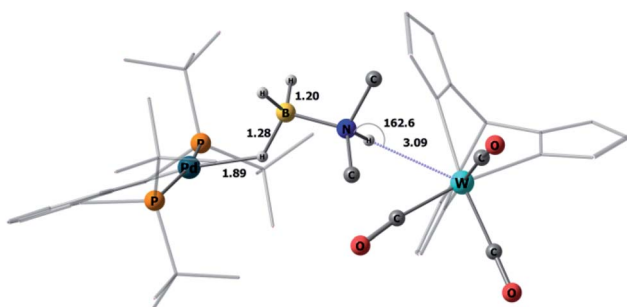
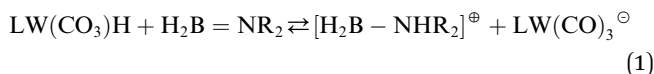


Fig. 9 DFT/ωB97XD-optimized structure of the complex 5b. Key bond lengths are reported in Å.

per ^tBuNH₂BH₃ molecule was also confirmed by volumetric measurements (Fig. S18†).

To explain the different ratios of ionic and molecular forms of catalyst units in the reaction with DMAB and TBAB, we considered the behavior of the first dehydrogenation product that accumulates in an unexpectedly high amount in the case of TBAB. Monomeric aminoborane H₂B=NR₂ is potentially basic³⁶ and appears to be able to deprotonate the tungsten hydride (eqn (1)).



The isolated [H₂B-NHR₂]⁺ cation is not stable, but it can be stabilized by interacting with nucleophilic atoms like the carbonyl or the THF's oxygen atoms. Indeed, the neutral CpWH(CO)₃/H₂B=NMe₂ complex appears to be almost ergoneutral relative to the ion pair CpW(CO)₃⁻/H₂B-NMe₂H⁺ (ΔE = -1.3 kcal mol⁻¹ in favor of the neutral form, Fig. S19†) stabilized by B···OC_W interactions. Introducing a THF molecule into the system also allows stabilizing the ion pair *via* B···O_{THF} interactions (ΔE = +0.2 kcal mol⁻¹; Fig. S19†), whereas CO groups of the CpW(CO)₃⁻ anion appear unbound, having a geometry similar to that of the CpW(CO)₃⁻/R₃NH⁺ ion pair.³⁷ It should be noted that H₂B=NH^tBu appears to be a weaker base

than H₂B=NMe₂; its neutral form is preferred by 3.8 kcal mol⁻¹. That suggests a lower deprotonation extent of LWH(CO)₃ when H₂B=NH^tBu is formed (Fig. S19 and S20†).

Based on these computational data, we could expect the equilibrium side process of LW(CO)₃H deprotonation, yielding the ion pair CpW(CO)₃⁻/H₂B-NMe₂H⁺ with CO···B interaction for DMAB dehydrogenation in toluene. When THF is used as a solvent, this ion pair is likely converted into the THF stabilized one. Consequently, the IR spectra in the ν_{CO} range in THF resemble those of the tungsten anion, while the IR spectra in toluene would rather resemble those of the ion pair 4 (Fig. S21†). Thus, this side process (eqn (1)) should affect the solution composition at the quasi-stationary stage of DMAB dehydrogenation, increasing the relative amount of LW(CO)₃⁻. Switching DMAB to TBAB diminishes the impact of LW(CO)₃H deprotonation leading to the presence of both metals in hydridic forms in the reaction mixture (Fig. S16†).

Reaction kinetics

Both DFT calculations and experimental data suggest that the N-H bond cleavage is the rate-determining step of the catalytic reaction. Under these conditions, the overall reaction rate is determined by the rate of proton transfer in complex 5 (5 → 6) $r_2 = k_2 \cdot [5]$. As shown by DFT calculations, the activation energy values for stepwise proton and hydride transfer (ΔG_{PT}[‡] = 21 kcal mol⁻¹ and ΔG_{HT}[‡] < 2 kcal mol⁻¹) indicate that $k_3 \gg k_2$ (Scheme 3). Thus, the hydride transfer goes much faster than proton transfer and conversion of 6 to products has no influence on the reaction rate. When DMAB is in excess, the reaction is pseudo-zero order in DMAB. Taking into account the pre-kinetic step of DMAB coordination to bimetallic complex 4 in the presence of excess DMAB, the reaction rate is $r = k_2 K_{C_0}(\text{DMAB}) \cdot [4] = k_{\text{eff}} \cdot [4]$, $K = [5]/([4][\text{DMAB}]])$. Thus, analysis of the experimental data gives the k_2 values of the rate-limiting step equal to 0.17 s⁻¹ and 0.03 s⁻¹ for 4a and 4b, respectively (Table S4†). These values correspond to ΔG_{298 K}[‡] 18.5 ± 0.1 and 19.5 ± 0.1 kcal mol⁻¹, in reasonable agreement with DFT calculations. For TBAB dehydrogenation by complex 4a, the k_2 value is 0.10 s⁻¹ (ΔG_{298 K}[‡] 18.8 ± 0.1 kcal mol⁻¹), indicating faster proton transfer. The overall activation free energy ΔG_{298 K}[‡] for the conversion of 4 to 6 is *ca.* 25 kcal mol⁻¹ that is in good agreement with DFT calculations (Fig. 8).

The use of deuterated amine-boranes NDMe₂BH₃ and NDMe₂BD₃ proves that the rate-limiting stage is the N-H bond cleavage, as is often for AB dehydrogenation.³⁸ The rate constant k_{eff} for NDMe₂BH₃ dehydrogenation catalyzed by 4a (Fig. S22 and S23†) is substantially lower ($k_{\text{eff}}^{\text{ND}} = 2.8 \times 10^{-7}$ M s⁻¹) than that for DMAB ($k_{\text{eff}}^{\text{NH}} = 5.8 \times 10^{-6}$ M s⁻¹) giving the kinetic isotope effect (KIE = $k^{\text{NH}}/k^{\text{ND}}$) of 20.6 ± 0.3. The use of the fully deuterated analogue NDMe₂BD₃ does not lead to a further decrease in the reaction rate, giving KIE = $k^{\text{H}}/k^{\text{D}}$ of 20.6 ± 0.7. The KIE obtained is substantially higher than typical KIEs in metal-catalyzed amine-borane dehydrogenation.³⁹⁻⁴³ This value also exceeds the KIEs for self-exchange in CpW(CO)₃H/[CpW(CO)₃]⁻ or for proton transfer of CpW(CO)₃H to aniline.^{44,45} Such high KIE values suggest a proton tunneling that is likely to occur when there is



a minimal geometry distortion along the reaction coordinate.⁴⁶ Our computations partially reproduce the observed magnitude of the isotope effect predicting increasing of the barrier upon NH to ND substitution by $\Delta\Delta G_{298}^\ddagger = 1.5 \text{ kcal mol}^{-1}$ ($\text{KIE} = k^{\text{NH}}/k^{\text{ND}} = 13$). One of the reasons for large KIE magnitudes (>10) is long (large acid–base separations) and strong H-bonded complexes.⁴⁷

Conclusions

In summary, bimetallic complexes $[\text{LW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ (**4**) act as “metallic-analogs” of typical main group FLPs. The presence of acidic and basic metal centers in these ionic pairs triggers the cooperative BH/NH bond activation in amine-boranes. In the first reaction stage, the η^1 -borane complex $[(\text{PCP})\text{Pd}(\sigma^1\text{-HBH}_2\text{-NR}_2\text{H})\cdots\text{W}(\text{CO})_3\text{L}]$ (**5**) is formed, in which BH is strongly bound to the palladium atom and the amine group is hydrogen-bonded to the tungsten atom. The step-wise proton transfer to W and hydride transfer to Pd yield the unsaturated B=N fragment and neutral metal hydrides. Molecular hydrogen evolution is the result of two metal hydrides interacting, regenerating bimetallic species **4**. One equivalent of H_2 can be produced from Me_2NHBH_3 , whereas dehydrogenation of ${}^t\text{BuNH}_2\text{BH}_3$ gives two equivalents of hydrogen per monomer at room temperature. The catalytic system can be easily generated through direct mixing of $\text{LW}(\text{CO})_3$, $(\text{PCP})\text{PdH}$ and amine-borane, without preliminary synthesis of the ionic catalyst **4**. The presence of ionic intermediates during the dehydrogenation cycle requires the use of polar solvents to achieve effective catalysis. The mechanistic study showed that proton transfer is the rate-determining step; therefore, the reaction is accelerated by a more basic anion $[(\text{CpW}(\text{CO})_3]^- > [\text{TpW}(\text{CO})_3]^-)$ while the hydride transfer is almost barrierless. The dehydrogenation process starts only in the presence of excess amine-borane due to the shift of the pre-kinetic equilibrium ($\text{H}_3\text{BNHR}_2 + \mathbf{4} = \mathbf{5}$) to the right, which in turn causes the increase of the initial reaction rate. Yet another notable observation is that the dimethylaminoborane monomer $\text{H}_2\text{B}=\text{NR}_2$ is able to deprotonate tungsten hydride in competition with H_2 evolution from two hydrides and $\text{H}_2\text{B}=\text{NR}_2$ oligomerization. Different basicities of dehydrogenated DMAB and TBAB monomers lead to a different impact of this side process to the overall catalytic reaction.

Prior studies have shown that bimetallic systems featuring a metal–metal bond⁴⁸ can indeed operate in a concerted way activating H_2 , C–H, B–H, and other bonds.^{49–52} These complexes activate the ONE bond, splitting it between two metals, often in an oxidative addition fashion. In our case, two transition metal-based building units of a bimetallic complex do not interact directly but act cooperatively as a Lewis acid and a Lewis base, splitting the N–H and B–H bond without changing the metals' oxidation state. So far, such reactivity has been reported only for $[\text{Cp}_2\text{ZrOC}_6\text{H}_4\text{P}^t\text{Bu}_2]^+$ which can be described as an early transition-metal-containing linked FLP.²¹ This behavior is similar to that of Stephan's main group FLPs¹⁸ and can be exploited for other catalytic conversions; these studies are underway in our laboratories.

Experimental section

All reactions were performed using standard Schlenk procedures under a dry argon atmosphere. Commercial reagents (dimethylamine-borane, *tert*-butylamine-borane) were purchased from Aldrich and used after preliminary sublimation. Tetrahydrofuran (THF) and toluene were dried over Na/benzophenone and distilled under an argon atmosphere. THF- d_8 (Aldrich) was stored over 4 Å molecular sieves and degassed before use by three freeze–pump–thaw cycles. NDMe_2BH_3 and NDMe_2BD_3 were prepared as described in the literature.^{40,53} Variable-temperature (VT) NMR spectra were recorded on Bruker AVANCE II and Varian Inova FT-NMR spectrometers operating at 300 and 400 MHz in the 200–320 K temperature range. ${}^1\text{H}$ chemical shifts are reported in parts per million (ppm) downfield of tetramethylsilane (TMS) and were calibrated against the residual resonance of the deuterated solvent, while ${}^{31}\text{P}\{{}^1\text{H}\}$ chemical shifts were referenced to 85% H_3PO_4 with downfield shift taken as positive. ${}^{11}\text{B}$ and ${}^{13}\text{C}\{{}^1\text{H}\}$ were referenced to $\text{BF}_3 \cdot \text{OEt}_2$. IR spectra were recorded at different temperatures (190–293 K) using a home-modified cryostat (Carl Zeiss Jena) with a Nicolet 6700 spectrometer using 0.05–0.2 cm CaF_2 cells. The accuracy of the experimental temperature was ± 0.5 °C. The cryostat modification allows transferring the reagents (premixed at either low or room temperature) under an inert atmosphere directly into the cells.

Elemental analyses were carried out in the Laboratory of Microanalysis of INEOS RAS. The classic manual technique was used. The sample was burned in a platinum crucible in a stream of oxygen at 950 °C followed by trapping CO_2 and water with Ascaris (asbestos impregnated with NaOH) and Anhydron (anhydrous magnesium perchlorate), respectively, and the analysis of the mass changes.

Synthesis of $[\text{LW}(\text{CO})_2(\mu\text{-CO})\cdots\text{Pd}(\text{PCP})]$ L = Cp (**4a**), Tp (**4b**)

Solid $\text{LW}(\text{CO})_3\text{H}$ (0.03 mmol) was placed in a Schlenk flask filled with argon (10 ml) together with 1 ml of tetrahydrofuran. Then, 1 ml of a $(\text{PCP})\text{PdH}$ THF solution (0.03 mmol of hydride in 2 ml of the solvent) was added to this solution. The resulting mixture of two colorless hydrides instantly became yellow colored. After an hour of stirring, the solvent was concentrated to *ca.* 0.1 ml under vacuum. Then, the Schlenk flask was filled with argon and left for several days standing at ambient temperature until a crystalline precipitate was obtained. The resulting solid (yellow needle-shaped crystals) was washed with a small amount of cold THF (2×0.2 ml) and dried in a vacuum. Yield: 90%.

4a. IR (THF, ν_{CO} , cm^{-1}): 1907, 1815, 1656. ${}^1\text{H}$ NMR (400 MHz, THF- d_8 , ppm): δ 6.93–6.85 (m, 3H, ArH); 5.10 (s, 5H, Cp-H), 3.28 (t, ${}^2J_{\text{P-H}} = 4.12$ Hz, 4H, $-\text{CH}_2-$); 1.43 (t, ${}^3J_{\text{P-H}} = 7.0$ Hz; 36H, $\text{C}(\text{CH}_3)_3$). ${}^{31}\text{P}\{{}^1\text{H}\}$ (161.9 MHz, THF- d_8 , ppm): δ 75.3. Anal. calc. (%) for $\text{C}_{32}\text{H}_{47}\text{O}_3\text{P}_2\text{PdW}$ (%): C 46.20, H 5.69. Found: C, 46.00; H, 5.78.

4b. IR (THF, ν_{CO} , cm^{-1}): 1901, 1796, 1650. ${}^1\text{H}$ NMR (600 MHz, THF- d_8 , ppm): δ 7.66 (s, 3H, Tp CH), 7.59 (s, 3H, Tp CH), 5.98 (s,



3H, Tp CH), 6.82–6.76 (m, 3H, ArH), 3.19 (pt, $^2J_{\text{P-H}} = 4.12$ Hz, 4H, $-\text{CH}_2-$), 1.24 (pt, $^3J_{\text{P-H}} = 7.0$ Hz, 36H, $\text{C}(\text{CH}_3)_3$). $^{31}\text{P}\{^1\text{H}\}$ (161.9 MHz, THF- d_8 , ppm): δ 76.8. Anal. calc. (%) for $\text{C}_{36}\text{H}_{53}\text{-BN}_6\text{O}_3\text{P}_2\text{PdW}$ (%): C 44.08, H 5.45. Found: C, 43.95; H, 5.53.

General procedure for DMAB dehydrogenation by **4**

For NMR kinetic studies. (PCP)PdH ($c = 0.01$ M) and **1a** or **1b** complexes ($c = 0.01$ M) and DMAB (5, 10 eq.) were mixed simultaneously at 290 K in THF- d_8 .

For variable temperature IR studies. Complex **4a** or **4b** was generated *in situ* by the reaction of (PCP)PdH ($c = 0.003$ – 0.0039 M) with **1a** or **1b** ($c = 0.003$ M), respectively. The reagents were dissolved in THF or toluene at 270 K and allowed to react for 30 min. Then, a chosen amount of DMAB was added, and the obtained mixture was monitored in the temperature range 190–290 K.

For kinetic IR studies. Reagents were prepared *via* three methods.

Method I. A portion of isolated complex **4a** or **4b** was dissolved in THF ($c = 0.003$ M). Then, a chosen amount of DMAB (5 equiv.) was added and the IR spectra were monitored until full catalyst regeneration.

Method II. Complex **4a** or **4b** was generated *in situ* by the reaction of (PCP)PdH ($c = 0.003$ – 0.0036 M) with **1a** or **1b** ($c = 0.003$ M), respectively. The reagents were dissolved in THF at 290 K and allowed to react for 20 min. Then, a chosen amount of DMAB (1–5 eq.) was added, and the mixture obtained was monitored until full catalyst regeneration.

Method III. (PCP)PdH ($c = 0.003$ – 0.0036 M), **1a** or **1b** complexes ($c = 0.003$ M) and DMAB (5 eq.) were mixed simultaneously at 290 K in THF. The reaction mixture was monitored until full catalyst regeneration.

Volumetric studies of amine-borane dehydrogenation.

Hydrogen evolution during dehydrogenation of amine-boranes was monitored using the Man on the Moon X103 kit. The volume of the system is 32 ml (two-necked round-bottom flask – 30 ml, three-way valve – 2 ml). The monitored solutions were prepared *via* three methods.

Method I. A portion of the isolated complex **4a** or **4b** (0.006 mmol) was dissolved in THF (1 ml) in an argon-filled flask of the device connected to a three-way valve. Then the flask was tightly closed with a septum cap, and the valve was opened to the pressure sensor. The chosen amount of DMAB (5–50 equiv.) in 1 ml of THF was injected with a syringe through a septum cap.

Method II. Complex **4a** or **4b** ($c = 0.003$ – 0.01 M) was generated *in situ* by mixing the solutions of (PCP)PdH (1–1.2 eq.) and **1** ($c = 0.003$ – 0.01 M, 1 eq.) in THF in an argon-filled flask of the device connected to a three-way valve. Then, the flask was tightly closed with a septum cap, and the valve was opened to the pressure sensor. A chosen amount of DMAB (5–50 eq.) in THF was injected with a syringe through a septum cap.

Method III. The solution of (PCP)PdH in THF (1–1.2 eq.) was prepared in an argon-filled flask of the device connected to a three-way valve. Then the flask was tightly closed with a septum cap, and the valve was opened to the pressure sensor.

The mixture of **1** ($c = 0.003$ – 0.01 M, 1 eq.) and amine-borane (5–50 equiv.) in THF was injected with a syringe.

The resulting mixture was stirred at 295 K. Data from a pressure sensor connected *via* a wireless network to a computer were recorded as a function of pressure *versus* time for 3–20 hours. The values accumulated were referenced by the pressure of THF in a blank experiment at 295 K and used for calculations of the H_2 equivalents evolved. The calculations were performed in the ideal gas approximation ($pV = nRT$).

Computational details

Calculations were performed with the Gaussian 09 (ref. 54) package at the DFT/ ω B97XD⁵⁵ level without any ligand simplification. For all atoms, the Def2-TZVP⁵⁶ basis set was applied, supplemented with an effective core potential⁵⁷ in the case of Pd and W. The structures of all complexes and transition states were fully optimized in toluene ($\epsilon_r = 2.3741$) described by the SMD model,⁵⁸ without any symmetry restrictions. The nature of all the stationary points on the potential energy surfaces was confirmed by vibrational analysis. Transition state (TS) structures showed only one negative eigenvalue in their diagonalized force constant matrices, and their associated eigenvectors were confirmed to correspond to the motion along the reaction coordinate under consideration using the Intrinsic Reaction Coordinate (IRC) method.⁵⁹

X-ray crystallography

For crystals of **4a** and **4b**, X-ray diffraction data were collected at 120 K with a Bruker ApexII DUO diffractometer using graphite monochromatic Cu- $K\alpha$ and Mo- $K\alpha$ radiation, respectively. Using Olex2,⁶⁰ the structures were solved with the ShelXT⁶¹ structure solution program using intrinsic phasing and refined with the XL⁶² refinement package using least-squares minimization. Hydrogen atom of the BH group in **4b** was found in the difference Fourier synthesis while positions of other hydrogen atoms were calculated, and they all were refined in the isotropic approximation within the riding model. Crystallographic data and structure refinement parameters are given in Table S1.† CCDC 2020560 and 2020559 contain the supplementary crystallographic data for **4a** and **4b**, respectively.†

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- R. M. Bullock and G. M. Chambers, *Philos. Trans. R. Soc., A*, 2017, **375**, 20170002.
- M. Bakos, A. Gyomai, A. Domjan and T. Soos, *Angew. Chem., Int. Ed.*, 2017, **56**, 5217–5221.
- L. Liu, B. Lukose and B. Ensing, *J. Phys. Chem. C*, 2017, **121**, 2046–2051.
- H. Grützmacher, *Angew. Chem., Int. Ed.*, 2008, **47**, 1814–1818.
- S. R. Flynn and D. F. Wass, *ACS Catal.*, 2013, **3**, 2574–2581.
- N. P. Mankad, *Chem. Commun.*, 2018, **54**, 1291–1302.
- N. V. Belkova, O. A. Filippov and E. S. Shubina, *Chem.–Eur. J.*, 2018, **24**, 1464–1470.
- V. A. Levina, A. Rossin, N. V. Belkova, M. R. Chierotti, L. M. Epstein, O. A. Filippov, R. Gobetto, L. Gonsalvi, A. Lledós, E. S. Shubina, F. Zanobini and M. Peruzzini, *Angew. Chem., Int. Ed.*, 2011, **50**, 1367–1370.
- E. S. Osipova, N. V. Belkova, L. M. Epstein, O. A. Filippov, V. A. Kirkina, E. M. Titova, A. Rossin, M. Peruzzini and E. S. Shubina, *Eur. J. Inorg. Chem.*, 2016, 1415–1424.
- C. Li, S. Cheng, M. Tjahjono, M. Schreyer and M. Garland, *J. Am. Chem. Soc.*, 2010, **132**, 4589–4599.
- T. J. Mazzacano, N. J. Leon, G. W. Waldhart and N. P. Mankad, *Dalton Trans.*, 2017, **46**, 5518–5521.
- A. Staubitz, A. P. Robertson and I. Manners, *Chem. Rev.*, 2010, **110**, 4079–4124.
- L. H. Jepsen, M. B. Ley, Y.-S. Lee, Y. W. Cho, M. Dornheim, J. O. Jensen, Y. Filinchuk, J. E. Jørgensen, F. Besenbacher and T. R. Jensen, *Mater. Today*, 2014, **17**, 129–135.
- R. Moury and U. B. Demirci, *Energies*, 2015, **8**, 3118–3141.
- A. Rossin and M. Peruzzini, *Chem. Rev.*, 2016, **116**, 8848–8872.
- A. L. Colebatch and A. S. Weller, *Chem.–Eur. J.*, 2019, **25**, 1379–1390.
- D. Han, F. Anke, M. Trose and T. Beweries, *Coord. Chem. Rev.*, 2019, **380**, 260–286.
- D. H. A. Boom, A. R. Jupp and J. C. Slootweg, *Chem.–Eur. J.*, 2019, **25**, 9133–9152.
- A. N. Marziale, A. Friedrich, I. Klopsch, M. Drees, V. R. Celinski, J. Schmedt auf der Günne and S. Schneider, *J. Am. Chem. Soc.*, 2013, **135**, 13342–13355.
- G. M. Adams, A. L. Colebatch, J. T. Skornia, A. I. McKay, H. C. Johnson, G. C. Lloyd-Jones, S. A. Macgregor, N. A. Beattie and A. S. Weller, *J. Am. Chem. Soc.*, 2018, **140**, 1481–1495.
- A. M. Chapman, M. F. Haddow and D. F. Wass, *J. Am. Chem. Soc.*, 2011, **133**, 8826–8829.
- S. Lau, D. Gasperini and R. L. Webster, *Angew. Chem., Int. Ed.*, DOI: 10.1002/anie.202010835.
- X. Chen, J.-C. Zhao and S. G. Shore, *Acc. Chem. Res.*, 2013, **46**, 2666–2675.
- A. Rossin, G. Tuci, L. Luconi and G. Giambastiani, *ACS Catal.*, 2017, **7**, 5035–5045.
- See ESI† for more details.
- T. Miyazaki, Y. Tanabe, M. Yuki, Y. Miyake and Y. Nishibayashi, *Organometallics*, 2011, **30**, 2394–2404.
- A. Rossin, G. Bottari, A. M. Lozano-Vila, M. Paneque, M. Peruzzini, A. Rossi and F. Zanobini, *Dalton Trans.*, 2013, **42**, 3533–3541.
- V. Skagestad and M. Tilset, *J. Am. Chem. Soc.*, 1993, **115**, 5077–5083.
- R. F. Jordan and J. R. Norton, *J. Am. Chem. Soc.*, 1982, **104**, 1255–1263.
- R. T. Baker, J. C. Gordon, C. W. Hamilton, N. J. Henson, P. H. Lin, S. Maguire, M. Murugesu, B. L. Scott and N. C. Smythe, *J. Am. Chem. Soc.*, 2012, **134**, 5598–5609.
- I. E. Golub, O. A. Filippov, E. S. Gulyaeva, E. I. Gutsul and N. V. Belkova, *Inorg. Chim. Acta*, 2017, **456**, 113–119.
- M. Roselló-Merino, J. López-Serrano and S. Conejero, *J. Am. Chem. Soc.*, 2013, **135**, 10910–10913.
- N. V. Belkova, L. M. Epstein and E. S. Shubina, *Eur. J. Inorg. Chem.*, 2010, **2010**, 3555–3565.
- E. M. Leitao, N. E. Stubbs, A. P. Robertson, H. Helten, R. J. Cox, G. C. Lloyd-Jones and I. Manners, *J. Am. Chem. Soc.*, 2012, **134**, 16805–16816.
- T. M. Gilbert, *Organometallics*, 1998, **17**, 5513–5520.
- M. H. Matus, D. J. Grant, M. T. Nguyen and D. A. Dixon, *J. Phys. Chem. C*, 2009, **113**, 16553–16560.
- V. A. Levina, O. A. Filippov, E. I. Gutsul, N. V. Belkova, L. M. Epstein, A. Lledós and E. S. Shubina, *J. Am. Chem. Soc.*, 2010, **132**, 11234–11246.
- A. L. Colebatch and A. S. Weller, *Chem.–Eur. J.*, 2019, **25**, 1379–1390.
- M. E. Sloan, A. Staubitz, T. J. Clark, C. A. Russell, G. C. Lloyd-Jones and I. Manners, *J. Am. Chem. Soc.*, 2010, **132**, 3831–3841.
- N. T. Coles, M. F. Mahon and R. L. Webster, *Organometallics*, 2017, **36**, 2262–2268.
- A. C. Ried, L. J. Taylor, A. M. Geer, H. E. Williams, W. Lewis, A. J. Blake and D. L. Kays, *Chem.–Eur. J.*, 2019, **25**, 6840–6846.
- R. J. Keaton, J. M. Blacquiere and R. T. Baker, *J. Am. Chem. Soc.*, 2007, **129**, 1844–1845.
- P. Bhattacharya, J. A. Krause and H. Guan, *J. Am. Chem. Soc.*, 2014, **136**, 11153–11161.
- R. T. Edidin, J. M. Sullivan and J. R. Norton, *J. Am. Chem. Soc.*, 1987, **109**, 3945–3953.
- K.-T. Smith, M. Tilset, S. S. Kristjansdottir and J. R. Norton, *Inorg. Chem.*, 1995, **34**, 6497–6504.
- R. M. Bullock and B. R. Bender, in *Encyclopedia of Catalysis*, ed. I. Horváth, John Wiley & Sons, Inc., 2002.
- P. M. Kiefer and J. T. Hynes, *J. Phys. Chem. A*, 2004, **108**, 11793–11808.
- J. Bauer, H. Braunschweig and R. D. Dewhurst, *Chem. Rev.*, 2012, **112**, 4329–4346.
- M. Devillard, R. Declercq, E. Nicolas, A. W. Ehlers, J. Backs, N. Saffon-Merceron, G. Bouhadir, J. C. Slootweg, W. Uhl and D. Bourissou, *J. Am. Chem. Soc.*, 2016, **138**, 4917–4926.
- J. Campos, *J. Am. Chem. Soc.*, 2017, **139**, 2944–2947.
- N. Hidalgo, J. J. Moreno, M. Pérez-Jiménez, C. Maya, J. López-Serrano and J. Campos, *Chem.–Eur. J.*, 2020, **26**, 5982–5993.
- H.-C. Yu, S. M. Islam and N. P. Mankad, *ACS Catal.*, 2020, **10**, 3670–3675.



- 53 L. Wirtz, W. Haider, V. Huch, M. Zimmer and A. Schäfer, *Chem.–Eur. J.*, 2020, **26**, 6176–6184.
- 54 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, T. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford CT, 2016.
- 55 J.-D. Chai and M. Head-Gordon, *Phys. Chem. Chem. Phys.*, 2008, **10**, 6615–6620.
- 56 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 57 D. Andrae, U. Haeussermann, M. Dolg, H. Stoll and H. Preuss, *Theor. Chim. Acta*, 1990, **77**, 123–141.
- 58 A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2009, **113**, 6378–6396.
- 59 K. Fukui, *Acc. Chem. Res.*, 1981, **14**, 363–368.
- 60 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.
- 61 N. Sebbar, M. Ellouz, E. Essassi, Y. Ouzidan and J. Mague, *Acta Crystallogr., Sect. E: Crystallogr. Commun.*, 2015, **71**, o999.
- 62 G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 2008, **64**, 112–122.

