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A soft on/off switch based on the electrochemically reversible H–J interconversion of a floating porphyrin membrane†

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Soft molecular assemblies that respond reversibly to external stimuli are attractive materials as on/off switches, in optoelectronic, memory and sensor technologies. In this Edge Article, we present the reversible structural rearrangement of a soft porphyrin membrane under an electrical potential stimulus in the absence of solid-state architectures. The free-floating porphyrin membrane lies at the interface between immiscible aqueous and organic electrolyte solutions and is formed through interfacial self-assembly of zinc(II) *meso*-tetrakis(4-carboxyphenyl)porphyrins (ZnPor). A potential difference between the two immiscible electrolyte solutions induces the intercalation of bis(triphenylphosphoranylidene) ammonium cations from the organic electrolyte that exchange with protons in the porphyrin membrane. *In situ* UV/vis absorbance spectroscopy shows that this ionic intercalation and exchange induces a structural interconversion of the individual porphyrin molecules in the membrane from an H- to a J-type molecular configuration. These structural rearrangements are reversible over 30 potential cycles. *In situ* polarisation-modulation fluorescence spectroscopy further provides clear evidence of structural interconversion of the porphyrin membrane, as intercalation of the organic electrolyte cations significantly affects the latter's emissive properties. By adjusting the pH of the aqueous phase, additional control of the electrochemically reversible structural interconversion can be achieved, with total suppression at pH 3.

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Introduction

Controlling the molecular arrangements of chromophores in supramolecular assemblies remains a grand challenge with functional dyes such as porphyrins.¹ In these assemblies, non-covalent intermolecular interactions precisely control the position and orientation of single molecules, thereby governing the physicochemical and photophysical properties of the ensemble.^{2–4} For example, in the solid-state, changes in molecular packing may turn non-emissive organic molecules such as terpyridine into good emitters,⁵ or impact significantly the emission of anthracene assemblies.⁶ Methodologies that reversibly modulate such arrangements using external stimuli

are long sought after given their potential importance in the development of on/off switching behaviour in optoelectronic, memory and sensor technologies.^{7–10} Among these stimuli, light,¹¹ pH,¹² and electric fields¹³ can provide flexible external control with a certain degree of reversibility.

Compared to the interface between solid electrodes and electrolytes, the interface between two immiscible electrolyte solutions (ITIES) is dynamic and lacks inherent defects such as step edges, domain boundaries, terraces and kinks.¹⁴ Thus, the uniform templating of adsorbed molecules at such “soft” interfaces can provide a pathway to molecular assemblies with unique structural arrangements and functionalities.¹⁵ Given the weak interaction with the interface, intermolecular interactions in molecular assemblies adsorbed at the ITIES can be externally manipulated by applying an electrical potential stimulus across the liquid|liquid boundary.¹⁶ Recently, we developed an experimental approach to induce the self-assembly of Zn(II) *meso*-tetrakis(4-carboxyphenyl)porphyrin (ZnPor) through a hydrogen bond-mediated mechanism at the ITIES.¹⁷ By optimising the pH of the aqueous phase to closely match the pK_a of the carboxyl groups ($pK_a = 5.8$),¹⁸ films of ZnPor nanostructures self-assembled at the liquid|liquid interface creating a porphyrin membrane intermediate between the two electrolyte solutions. Upon polarisation, carboxyl groups in the ZnPor membrane act

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as cationic exchange sites and give rise to purely ionic electrochemical signals resembling those of redox-type reactions in solid substrates.¹⁹

In this Edge Article, we use *in situ* spectroelectrochemical techniques to demonstrate that ion intercalation and exchange reactions in these soft ZnPor membranes are accompanied by a change in the stacking arrangement of the individual porphyrin molecules within the nanostructures. *In situ* UV/vis absorbance and polarisation-modulation fluorescence spectroscopies in total internal reflection mode (TIR-UV/vis and PM-TIRF, respectively) show that exchange of protons in the porphyrin membrane by bulky bis(triphenylphosphoranylidene)ammonium cations (R_2N^+) from the organic electrolyte induces a stacking interconversion of the porphyrins from a compact H-type nanostructure to an open J-type. This interconversion can be manipulated by an interfacial Galvani potential difference between the two electrolyte solutions ($\Delta\phi^w$) and is reversible over 30 potential cycles.

Results and discussion

Monitoring reversible H- to J-type interconversion of the interfacial porphyrin membrane by *in situ* UV/vis spectroelectrochemistry

The potential induced interconversion between H- and J-type nanostructures in the ZnPor membrane was analysed in a 4-electrode configuration by cyclic voltammetry (CV), and *in situ* TIR-UV/vis and PM-TIRF; the configuration of the electrochemical cell used is presented schematically in Fig. 1 and in detail in Fig. S1 of the ESI.[†] Schematics and experimental

optimisation procedures for the *in situ* TIR-UV/vis absorbance and PM-TIRF spectroelectrochemical setups, respectively, are detailed in Fig. S2–S6.[†] Soft ZnPor membranes were obtained from the self-assembly of ZnPor at the water/ α,α,α -trifluorotoluene (TFT) interface as described in detail in the ESI[†] and our previous work.¹⁷ For CV and TIR-UV/vis absorbance experiments, ZnPor molecules not adsorbed at the liquid/liquid interface were flushed out, leaving only the membrane between the two electrolyte solutions of approximately 200 nm in thickness (Fig. S7[†]). For PM-TIRF experiments, after the ZnPor membrane was self-assembled, non-adsorbed ZnPor molecules remained in the aqueous phase during the course of the experiment.

Fig. 2a shows CVs in the presence and absence of the ZnPor membrane, respectively. At potentials where the aqueous phase is polarised negatively with respect to the organic phase ($\Delta\phi^w < 0$), an increase in the interfacial concentration of R_2N^+ near the organic electrolyte/membrane interface leads to a reversible intercalation/exchange reaction associated with peaks A and B in the voltammogram.¹⁹ Scanning negatively, R_2N^+ ions penetrate the membrane binding to anionic carboxyl sites and displacing protons through ion exchange at neutral carboxyl sites

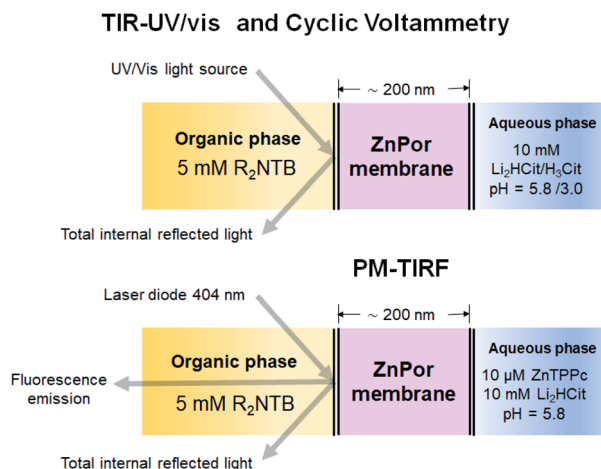


Fig. 1 Schematic configurations for spectroelectrochemical studies with the ZnPor membrane system. The techniques used were cyclic voltammetry (CV), *in situ* UV/vis absorbance and polarisation-modulation fluorescence spectroscopies in total internal reflection mode (TIR-UV/vis and PM-TIRF, respectively). The organic phase was α,α,α -trifluorotoluene (TFT) containing 5 mM bis(triphenylphosphoranylidene)ammonium tetrakis(pentafluorophenyl)-borate (R_2NTB). The aqueous phase was 10 mM lithium citrate (Li_2HCit) at pH 5.8, or 10 mM citric acid (H_3Cit) at pH 3.0. The ZnPor membrane was prepared from 10 μM solutions of ZnPor in contact with the TFT as described in the ESI.[†]

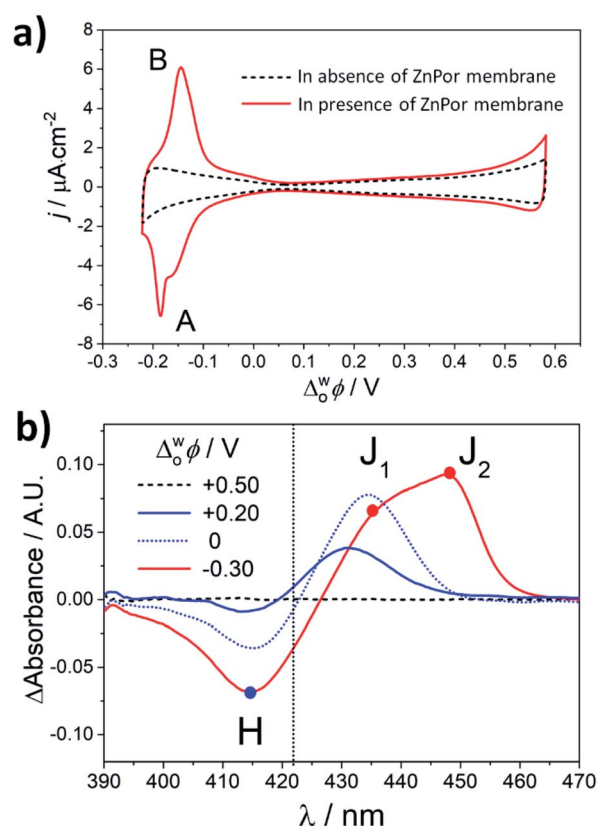


Fig. 2 (a) Cyclic voltammetry (CV) in the absence and presence of the ZnPor membrane in the configuration presented schematically in Fig. 1a. The scan rate was 25 mV s⁻¹. The 4-electrode electrochemical cell setup used for these measurements is presented in Fig. S1.[†] (b) Differential TIR-UV/vis absorbance spectra of the ZnPor membrane at different $\Delta\phi^w$. The spectrum of the membrane at +0.50 V was taken as a blank.



producing peak A. The latter process reverses as the potential direction is switched and scanned to more positive potentials, yielding peak B.¹⁹

In situ TIR-UV/vis absorbance measurements were carried out at different $\Delta\phi^w$, as presented in Fig. 2b. Differential spectra were obtained by subtracting the spectrum of the ZnPor membrane obtained at +0.5 V from spectra obtained at each $\Delta\phi^w$. A stepwise interconversion between three dominant supramolecular nanostructures is observed. Using the λ_{\max} of ZnPor in solution at pH 5.8 ($\lambda_{\max} = 422$ nm) as a reference (Fig. S8†), these supramolecular nanostructures were assigned as either H-type ($\lambda_{\max} = 414$ nm) or J-type (J_1 , $\lambda_{\max} = 435$ nm and J_2 , $\lambda_{\max} = 448$ nm). Control differential spectra in the absence of the ZnPor film were featureless (Fig. S4†).

The differential TIR-UV/vis absorbance spectra in Fig. 2b indicate that the structure of the ZnPor membrane is dynamic across the whole potential range studied. The potential of zero charge (PZC) at pH 5.8 in the presence of the ZnPor membrane was determined previously as +0.25 V.¹⁹ At potentials approaching or more positive than the PZC, *i.e.*, +0.2 and +0.5 V, ZnPor molecules primarily adopt a compact H-type configuration as observed for other porphyrin assemblies at the ITIES.²⁰ Note that at these positive potentials, no electrochemical signals were observed in the presence of the ZnPor membrane, and thereby a more compact structure could be expected. At potentials more negative than the PZC, *i.e.*, 0 and −0.30 V, R_2N^+ accumulates at the organic electrolyte|membrane interface. We propose that electrostatic interactions with R_2N^+ cause a partial opening of the nanostructures in the membrane and the appearance of an initial J-type structure (J_1). The subsequent intercalation of bulky R_2N^+ species and their exchange with substantially smaller protons in the membrane cause further structural rearrangements into a more open J-type configuration (J_2). A structural interconversion was also observed in the presence of other bulky and hydrophobic ions, such as tetraoctylammonium cations (TOA⁺), but absent with smaller

species, such as tetraethylammonium (TEA⁺) or tetrabutylammonium (TBA⁺) cations (data not shown). These observations suggest that features such as size and hydrophobicity are central to the ability of these cations to intercalate into the ZnPor membrane (Fig. S13 and S14†).

Of particular note is that differential TIR-UV/vis absorbance measurements clearly indicate that H–J interconversion begins to take place even at potentials considerably positive of peaks A and B. The latter is due to the continuous change in the double layer and the gradual build-up of R_2N^+ species at the organic electrolyte|membrane interface as $\Delta\phi^w$ is scanned negatively.

The reversibility of the structural H–J interconversion was demonstrated by comparing the magnitude of the absorbance of the H and J_2 nanostructures as a function of the number of potential cycles, as presented in Fig. 3a (one cycle = potential steps from +0.50 V to −0.25 V to +0.50 V). A near constant absorbance value for 30 cycles indicates an excellent, robust reversibility of the optical properties of the membrane. This observation is in line with the negligible charge differential ($\Delta Q \approx 0 \mu\text{C cm}^{-2}$) determined electrochemically for peaks A and B in the voltammogram, presented in Fig. S9.† The electrochemically reversible H–J interconversion of the ZnPor membrane at the water|TFT interface is shown schematically in Fig. 3b.

Suppression and reactivation of the interfacial porphyrin membrane H–J interconversion by pH

The pH of the aqueous phase significantly effects the structural interconversion and electrochemistry of the ZnPor membrane. As presented in Fig. 4, by decreasing the pH from 5.8 to 3.0, the electrochemical ion intercalation/exchange was suppressed with all electrochemical signals related to the presence of the ZnPor membrane disappearing. Furthermore, the H–J interconversion was suppressed at pH 3.0, with no changes in the differential absorbance spectra observed at different $\Delta\phi^w$. Thus, the spectroelectrochemical analysis clearly demonstrates that

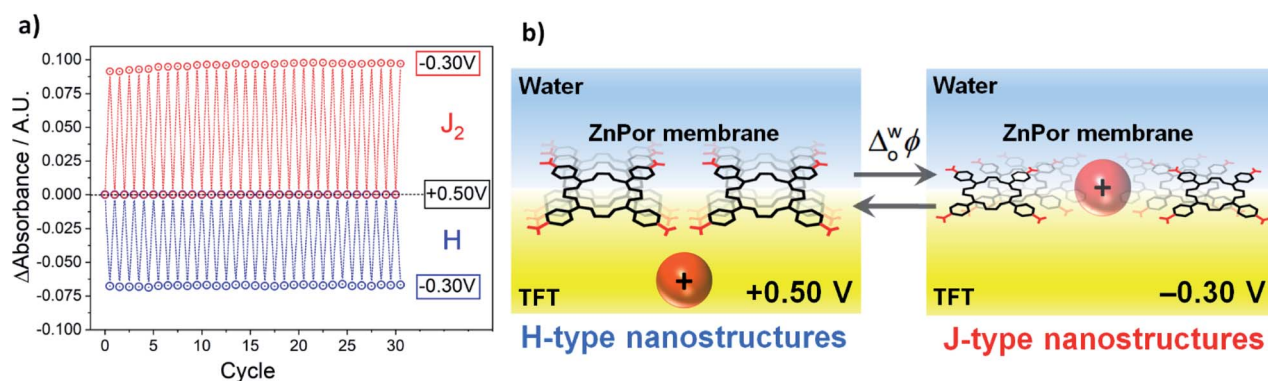


Fig. 3 (a) Reversible H–J interconversion of the ZnPor membrane as a function of the potential cycle. One cycle = potential steps from +0.50 V to −0.25 V to +0.50 V. The absorbance of H and J_2 was measured at 414 and 448 nm, respectively, from the differential TIR-UV/vis absorbance spectra presented in Fig. S10 and animated in a movie, see ESI.† (b) Schematic of the electrochemically reversible H–J interconversion of the ZnPor membrane at the water|TFT interface. For clarity, only the porphyrins near the interface are depicted, but R_2N^+ intercalation occurs throughout the *ca.* 200 nm thickness of the ZnPor membrane and not just at the monolayer of porphyrin near the water|TFT interface. Thus, upon variations of the applied $\Delta\phi^w$, the thickness of ZnPor membrane is not fixed and will expand/contract during the reversible intercalation/de-intercalation process.

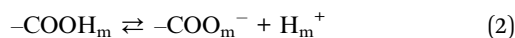


the potential induced intercalation/exchange reaction is responsible for the H-J interconversion observed. Consequently, when the former is suppressed, the porphyrin membrane retains a compact H-type configuration. Control TIR-UV/Vis absorbance experiments show that even though the ZnPor membrane initially self-assembles only at a pH value near the pK_a of the carboxyl groups ($pK_a = 5.8$), once assembled, the membrane is stable even at very acidic conditions (Fig. S11†). This excludes any effects of redissolution of the membrane on the analysis presented.

Both the applied $\Delta_o^w\phi$ and pH of the aqueous phase shift the chemical equilibrium involving carboxyl groups in the ZnPor membrane, aqueous protons, and bulky organic R_2N^+ cations. The overall equilibrium of the ion exchange in the membrane, m, can be represented as follows



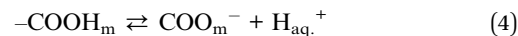
This equilibrium is similar to that present in membranes used in ion selective electrodes (ISE).²¹ In effect, within a certain potential range, the porphyrin membrane acts as an intermediate phase separating the aqueous and organic electrolyte solutions.¹⁹ The ion exchange reaction can be considered as the combination of acid dissociation



and binding of the organic cation



In turn, the equilibrium constant of the acid dissociation reaction is dictated by ionic equilibrium at the membrane|aqueous electrolyte interface



where aq. is the aqueous phase. Furthermore, the equilibrium constant of the binding reaction is dictated by ionic equilibrium at the organic electrolyte|membrane interface (*i.e.* intercalation)



where org. is the organic phase. The interfacial Galvani potential difference necessary for the ion exchange reaction ($\Delta_o^w\phi^{\text{IE}}$) is therefore a function of the concentration of the carboxyl groups in the ZnPor membrane, organic R_2N^+ cations, and aqueous protons. As the pH of the aqueous phase decreases, more energy is required to shift the chemical equilibria described in eqn (1)–(5) to favour acid dissociation and R_2N^+ binding in the membrane. Thus, $\Delta_o^w\phi^{\text{IE}}$ shifts to more negative potentials and eventually falls outside the polarisable potential window, as observed when the pH of the aqueous electrolyte is 3.0 in Fig. 4.

In situ polarisation-modulation fluorescence spectroscopy of the interfacial porphyrin membrane H-J interconversion

Further spectroelectrochemical evidence of the reversible structural interconversion of the ZnPor membrane was obtained using PM-TIRF to probe the effect of interfacial polarisation on the emitted fluorescence. PM-TIRF signals ($\Delta F^{\text{P-S}}$) as a function of emission wavelength (PM-TIRF spectra) were recorded at different $\Delta_o^w\phi$ (Fig. 5a.). As demonstrated in previous works, PM-TIRF spectra are interface-specific, with signals arising from randomly oriented molecules in the bulk phases effectively cancelled out.^{22,23} At $\Delta_o^w\phi = +0.59, +0.29, -0.06$ and -0.15 V, two emission peaks were observed ($\lambda_{\text{em.}} = 613$ and 663 nm), similar to the emission peaks for the monomeric porphyrin (see Fig. S15†). The negative values of $\Delta F^{\text{P-S}}$ at each $\Delta_o^w\phi$ confirm the relatively horizontal orientation of individual ZnPor molecules within the membrane with respect to the plane of the interface.²³ PM-TIRF spectra measured at -0.06 and -0.15 V show a major decrease in $-\Delta F^{\text{P-S}}$ compared to the spectrum at $+0.29$ V, with no spectral shift. The decrease in $-\Delta F^{\text{P-S}}$ may be associated with a change in molecular orientation of the ZnPor molecules within the membrane at the more negative potentials. The lack of a spectral shift indicates association of the ZnPor molecules in the membrane mainly through hydrogen bonds between their carboxyl groups.¹⁶ This finding is in line with the pH dependency of the ZnPor membrane formation, as detailed in our recent work,¹⁷ and the cooperative hydrogen bonding model proposed by Girault and co-workers in their photoelectrochemical studies of ZnPor at a polarised water|1,2-dichloroethane interface.²⁴ At $\Delta_o^w\phi < -0.30$ V, the PM-TIRF spectra red-shifted and a third emission peak ($\lambda_{\text{em.}} = 694$ nm) appeared, confirming close π - π stacking of the porphyrin ring at the interface and the formation of J-type nanostructures.

As presented in Fig. 5b, by switching the potential across the interface between the positive and negative potentials, the emission peak at 694 nm can be effectively controlled, further demonstrating the reversibility of the structural interconversion and the control over the emission properties of the porphyrin

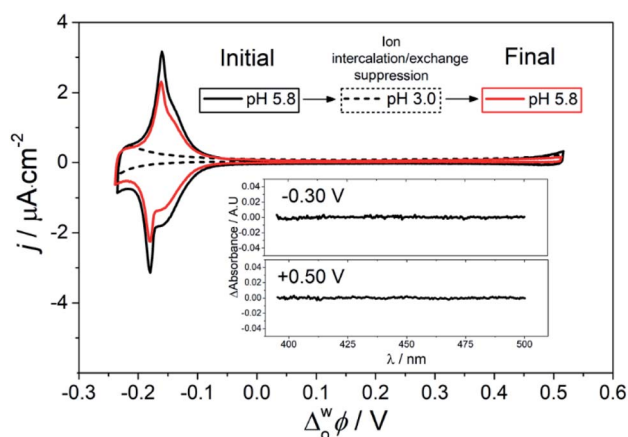


Fig. 4 Suppression and reactivation of the ZnPor membrane H-J interconversion monitored by CV and differential TIR-UV/vis absorbance. The scan rate for the CV was 5 mV s^{-1} . Inset: differential absorbance of the membrane at pH 3.0 with $\Delta_o^w\phi$ equal to -0.30 V and $+0.50$ V, respectively. The TIR-UV/vis absorbance spectrum at $+0.50$ V was taken as the blank, and spectra at other $\Delta_o^w\phi$ across the full potential range studied are shown in Fig. S12.†



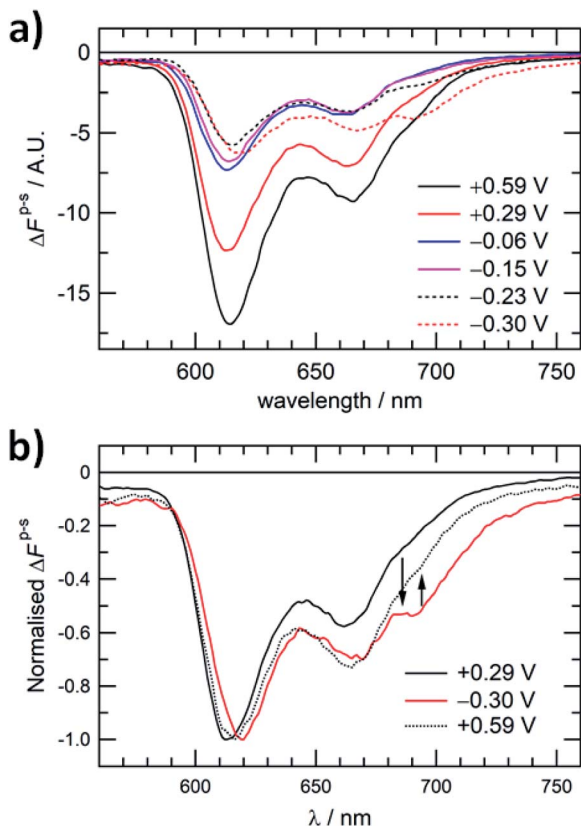


Fig. 5 (a) PM-TIRF spectra of the ZnPor membrane at different $\Delta\phi$. (b) PM-TIRF spectra of the ZnPor membrane after a potential cycle from +0.29 V to -0.30 V and back to +0.59 V. $\lambda_{\text{exc.}} = 404$ nm.

membrane. The rate of the dissociation of the J-aggregate was rather slow at +0.29 V, and therefore a higher positive potential of +0.59 V was applied on switching back from negative to positive potentials.

Conclusions

In this Edge Article, we demonstrate the reversible structural interconversion of a soft porphyrin membrane from an H- to a J-type molecular configuration in the absence of any solid-state architectures. The interconversion was controlled by applying an interfacial Galvani potential difference between two immiscible electrolyte solutions. The potential difference induced an ion intercalation/exchange reaction that affected reversibly the supramolecular structure of the membrane. The latter was demonstrated electrochemically by cyclic voltammetry and spectroscopically by *in situ* UV/vis absorbance and polarisation-modulated fluorescence spectroscopies in total internal reflection mode. The H- to J-interconversion was robust and repeatable over at least 30 potential cycles. The interconversion was suppressed within the polarisable potential window at pH 3.0, but recoverable by returning the pH of the aqueous phase to pH 5.8.

The work herein opens new opportunities to investigate the use of structurally dynamic soft supramolecular porphyrin

assemblies at the ITIES as switchable on/off responsive nanomaterials, with applications in optoelectronic, memory and sensor devices. The structural interconversion was not solely restricted to the intercalation of one specific organic cation, such as bis(triphenylphosphoranylidene)ammonium cations, but also observed with tetraoctylammonium cations. Thus, this type of electrochromic behavior can be manipulated by introducing different hydrophobic and bulky cations in the organic phase.

Data availability

All pertinent experimental data and detailed procedures for implementation and optimisation of the spectroelectrochemical experiments at the interface between two immiscible electrolyte solutions (ITIES) are provided in the ESI.

Author contributions

A. F. M.-O. and M. D. S. designed the study. A. F. M.-O., S. Y. and A. G.-Q. developed and optimised the TIR-UV/vis experimental setup. A. F. M.-O., S. Y. and I. R.-M. carried out the TIR-UV/vis experiments and analysed the data. S. Y. carried out the PM-TIRF experiments and analysed the data under the supervision of H. N. M. D. S. and H. N. were responsible for overall project administration and supervision as well as funding acquisition. All authors contributed to the discussion and writing the article.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- G. Magna, D. Monti, C. Di Natale, R. Paolesse and M. Stefanelli, *Molecules*, 2019, **24**, 4307.
- K. Cai, J. Xie, D. Zhang, W. Shi, Q. Yan and D. Zhao, *J. Am. Chem. Soc.*, 2018, **140**, 5764–5773.
- M. F. J. Mabesoone, A. J. Markvoort, M. Banno, T. Yamaguchi, F. Helmich, Y. Naito, E. Yashima, A. R. A. Palmans and E. W. Meijer, *J. Am. Chem. Soc.*, 2018, **140**, 7810–7819.
- R. van der Weegen, A. J. P. Teunissen and E. W. Meijer, *Chem.-Eur. J.*, 2017, **23**, 3773–3783.



- 5 T. Mutai, H. Satou and K. Araki, *Nat. Mater.*, 2005, **4**, 685–687.
- 6 Y. Dong, B. Xu, J. Zhang, X. Tan, L. Wang, J. Chen, H. Lv, S. Wen, B. Li, L. Ye, B. Zou and W. Tian, *Angew. Chem., Int. Ed.*, 2012, **51**, 10782–10785.
- 7 A. Mammanna, A. D'Urso, R. Lauceri and R. Purrello, *J. Am. Chem. Soc.*, 2007, **129**, 8062–8063.
- 8 S. Ishihara, J. Labuta, W. Van Rossom, D. Ishikawa, K. Minami, J. P. Hill and K. Ariga, *Phys. Chem. Chem. Phys.*, 2014, **16**, 9713–9746.
- 9 S. Yagai, in *Supramolecular Soft Matter: Applications in Materials and Organic Electronics*, ed. T. Nakanishi, John Wiley & Sons, Inc, 2011, pp. 77–95.
- 10 T. Eder, T. Stangl, M. Gmelch, K. Remmerssen, D. Laux, S. Höger, J. M. Lupton and J. Vogelsang, *Nat. Commun.*, 2017, **8**, 1641.
- 11 J. Wang, A. Kulago, W. R. Browne and B. L. Feilnga, *J. Am. Chem. Soc.*, 2010, **132**, 4191–4196.
- 12 K. Adachi, Y. Ura and N. Kanetada, *J. Porphyrins Phthalocyanines*, 2018, **22**, 658–669.
- 13 M. Mrinalini, S. S. Pathak, B. S. Achary, L. S. Panchakarla and S. Prasanthkumar, *Chem.-Asian J.*, 2019, **14**, 537–541.
- 14 H. H. Girault, in *Developments in Electrochemistry: Science Inspired by Martin Fleischmann*, ed. D. Pletcher, Z.-Q. Tian and D. E. Williams, John Wiley & Sons, Ltd, 2014, pp. 295–308.
- 15 J. F. Stoddart, *Nat. Chem.*, 2009, **1**, 14–15.
- 16 Z. Samec, *Electrochim. Acta*, 2012, **84**, 21–28.
- 17 A. F. Molina-Osorio, D. Cheung, C. O'Dwyer, A. A. Stewart, M. Dossot, G. Herzog and M. D. Scanlon, *J. Phys. Chem. C*, 2020, **124**, 6929–6937.
- 18 N. C. Maiti, S. Mazumdar and N. Periasamy, *J. Phys. Chem. B*, 1998, **102**, 1528–1538.
- 19 A. F. Molina-Osorio, J. A. Manzanares, A. Gamero-Quijano and M. D. Scanlon, *J. Phys. Chem. C*, 2020, **124**, 18346–18355.
- 20 A. J. Olaya, D. Schaming, P. F. Brevet, H. Nagatani, H. J. Xu, M. Meyer and H. H. Girault, *Angew. Chem., Int. Ed.*, 2012, **51**, 6447–6451.
- 21 E. Bakker, P. Bühlmann and E. Pretsch, *Chem. Rev.*, 1997, **97**, 3083–3132.
- 22 S. Yamamoto, H. Nagatani and H. Imura, *Langmuir*, 2017, **33**, 10134–10142.
- 23 S. Yamamoto, H. Nagatani, K. Morita and H. Imura, *J. Phys. Chem. C*, 2016, **120**, 7248–7255.
- 24 H. Jensen, J. J. Kakkassery, H. Nagatani, D. J. Fermín and H. H. Girault, *J. Am. Chem. Soc.*, 2000, **122**, 10943–10948.

