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## Synthesis and photophysical properties of benzoxazolyl-imidazole and benzothiazolyl-imidazole conjugates†

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Materials that have higher fluorescence emission in the solid state than molecules in solution have recently been paid more attention by the scientific community due to their potential applications in various fields. In this work, we newly synthesized benzoxazolyl-imidazole and benzothiazolyl-imidazole conjugates, which show aggregation-induced emission (AIE) features in their solid and aggregate states. It was found that oxygen and sulfur substitutions can dramatically influence the molecular structures and polarities of the dyes, leading to different degrees of the AIE phenomenon. The benzothiazolyl-imidazole molecule has lower polarity compared to that of benzoxazolyl-imidazole; therefore, the dye bearing a benzothiazolyl group shows higher emission intensity and dual emission in aqueous solution. Theoretical calculation results suggest that the benzothiazolyl-imidazole molecules might have electrostatic interactions between sulfur and nitrogen atoms, explaining the experimental observations of lower critical aggregation concentration and photophysical properties both in solution and in the solid state. The theoretical calculations agree with the experimental data, thus demonstrating a potent strategy to gain a deep understanding of the structure–property relationships to design solid-state fluorescent materials.

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## 1. Introduction

Materials that have higher fluorescence emission in the solid and aggregate state than that of molecules in solution have recently been paid more attention by the scientific community due to their potential applications in various fields, including explosive detection, optoelectronics, photonics, sensing, biomedical applications and so on.<sup>1–11</sup> The concept of aggregation-induced emission (AIE) was discovered in 2001 and since then numerous AIEgens have been developed by researchers, academics and others.<sup>12</sup> Until now, several interesting types of AIEgen based on tetraphenylethene,<sup>13–15</sup> hexaphenylsilole,<sup>16,17</sup> tetraphenylpyrazine,<sup>18,19</sup> distyrylanthracene,<sup>20,21</sup> boron diimides<sup>22</sup> and imidazole<sup>23</sup> have been explored.

Among these chromophores, imidazole-based molecules were relatively rarely investigated as AIEgens, instead they were extensively utilized for anticancer agents, antibacterial agents, catalysts, dyes and ionic liquids.<sup>24–31</sup> On the other hand, cyanines are a popular class of  $\pi$ -conjugated donor–acceptor (D–A) chromophores with a characteristic of charged structures. The scientific interest in cyanine dyes is mainly related to their excellent spectral properties, for instance good fluorescence

quantum yields, high molar extinction coefficients, broad wavelength tunabilities, stability, and increased sensitivity.<sup>32,33</sup> Besides, benzoxazolyl and benzothiazolyl derivatives, such as oxazole yellow and thiazole orange, have been proven to achieve selective labeling of cancer cells due to their high affinity.<sup>34–36</sup> However, they have weak fluorescence in their aggregated state, limiting their applications.<sup>37–39</sup> Thus, this work aims to develop new types of donor–acceptor-conjugated cyanine-like molecules, which show AIE features in their solid and aggregate states. Herein, the newly discovered D– $\pi$ –A molecules, comprised of imidazole and benzoxazolyl as well as imidazole and benzothiazolyl units were designed and synthesized. The effect of heteroatoms (oxygen and sulfur) on the photophysical properties and molecular geometries were investigated. It was found that the oxygen and sulfur substitutions can dramatically influence molecular structures and polarities of the dyes, leading to the different degree of AIE phenomenon. Since the benzothiazolyl-imidazole molecule has lower polarity than benzoxazolyl-imidazole, the dye bearing benzothiazolyl group shows higher emission intensity and dual emission in aqueous solution, which accompanying the more minor AIE enhancement and red solid is observed. Furthermore, the theoretical calculations agree with the experimental observations, thus demonstrating a potent strategy to gain a deep understanding of the structure–property relationship to design solid-state fluorescent materials.

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## 2. Methods and materials

### 2.1. Synthesis of dyes

**2.1.1. Synthesis of **6a**.** Compound **3** (ref. 23) (0.14 mmol, 36.7 mg) and **5a**<sup>40</sup> (0.14 mmol, 40.1 mg) were dissolved in dry ethanol with trace amount of pyridine, and the reaction mixture was heated to reflux under nitrogen for 12 hours. After cooled to room temperature, the reaction mixture was concentrated by rotary evaporator and the residue was precipitated with hexane, then washed with acetonitrile to obtain compound **6a** (23.7 mg, 32%) as an orange-red solid. <sup>1</sup>H NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta$  = 3.02 (t, 2H, CH<sub>2</sub>, *J* = 6.8 Hz), 3.74 (s, 3H, CH), 5.11 (t, 2H, CH<sub>2</sub>, *J* = 6.8 Hz), 7.23–7.27 (m, 3H, 3CH), 7.45–7.79 (m, 4H, 4CH), 7.58–7.61 (m, 3H, 3CH), 7.81 (t, 1H, CH, *J* = 8.2 Hz), 7.90 (t, 1H, CH, *J* = 8.2 Hz), 8.12 (d, 1H, CH, *J* = 15.2 Hz), 8.18 (d, 1H, CH, *J* = 15.2 Hz), 8.35 (d, 1H, CH, *J* = 8.2 Hz), 8.47 (d, 1H, CH, *J* = 8.2 Hz). <sup>13</sup>C NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta$  = 32.12, 33.06, 45.40, 113.89, 117.38, 124.85, 127.23, 127.87, 128.57, 128.78, 128.76, 130.05, 130.98, 133.26, 133.90, 135.11, 141.43, 141.65, 142.85, 171.82, 171.96. HRMS (ESI<sup>−</sup>) *m/z* for C<sub>28</sub>H<sub>23</sub>BrN<sub>3</sub>O<sub>3</sub>, calcd 528.09228, found 528.27682.

**2.1.2. Synthesis of **6b**.** In a manner similar to that described above, a solution of **3** (ref. 23) (0.14 mmol, 36.7 mg), **5b**<sup>41</sup> (0.14 mmol, 42.3 mg) and the catalytic amount of pyridine in dry ethanol to yield compound **6b** (19.0 mg, 25%) as red solid. <sup>1</sup>H NMR (400 MHz, DMSO-*d*<sub>6</sub>):  $\delta$  = 3.00 (t, 2H, CH<sub>2</sub>, *J* = 6.3 Hz), 3.06 (s, 3H, CH), 5.08 (t, 2H, CH<sub>2</sub>, *J* = 6.3 Hz), 7.21–7.29 (m, 3H, 3CH), 7.43–7.46 (m, 4H, 4CH), 7.56–7.57 (m, 3H, 3CH), 7.79 (t, 1H, CH, *J* = 7.5 Hz), 7.87 (t, 1H, CH, *J* = 7.5 Hz), 8.10 (d, 1H, CH, *J* = 14.8 Hz), 8.15 (d, 1H, CH, *J* = 14.8 Hz), 8.31 (d, 1H, CH, *J* = 8.1 Hz), 8.43 (d, 1H, CH, *J* = 8.1 Hz). <sup>13</sup>C NMR (100 MHz, DMSO-*d*<sub>6</sub>):  $\delta$  = 32.11, 33.04, 45.38, 113.86, 117.38, 124.85, 127.21, 127.85, 128.56, 128.77, 128.85, 129.70, 129.75, 130.04, 130.97, 133.29, 133.93, 135.10, 141.43, 141.65, 142.85, 171.83, 171.98. HRMS (ESI<sup>−</sup>) *m/z* for C<sub>28</sub>H<sub>23</sub>BrN<sub>3</sub>O<sub>2</sub>S<sub>1</sub>, calcd 544.06943, found 544.07623.

### 2.2. Characterizations

The <sup>1</sup>H and <sup>13</sup>C NMR spectra were conducted with Bruker Avance NEO 400 MHz NMR spectrometer using DMSO-*d*<sub>6</sub> as the solvent. The UV-vis absorption and fluorescence emission spectra were recorded on Shimadzu UV-2550 spectrometer and Horiba FluoroMax®-4 spectrometer, respectively. Emission spectra of **6a** and **6b** were observed for the excitation wavelengths of 470 nm and 475 nm, respectively. Micro fluorometer cuvettes (1 mm light path, 0.35 mL volume) were used to avoid reabsorption of emission from the sample.

### 2.3. Computational methods

All the calculations were performed with the Gaussian98 program package.<sup>42</sup> The geometric structures of molecules **6a** and **6b** were optimized with density functional theory (DFT) using the B3LYP functional (Becke's three-parameter hybrid functional using the correlation functional of Lee, Yang, and Parr, which includes both local and non-local terms correlation functionals) and the 6-31G(d,p) basis set.<sup>43</sup> The Hartree–Fock

(HF)/6-31G(d,p) methods were also carried out to compare the DFT results and the experimental measurements. The transition energies were calculated at the TD-DFT/B3LYP/6-31G(d,p) and HF/6-31G(d,p) geometries, respectively.

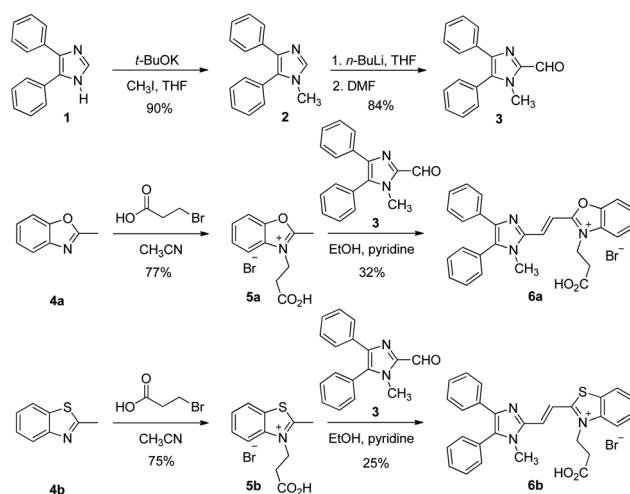
## 3. Results and discussion

### 3.1. Synthesis

The synthetic routes of dyes **6a** and **6b** are displayed in Scheme 1. Compounds **3**, **5a** and **5b** were synthesized according to literature procedures.<sup>23,40,41</sup> 1-Methyl-4,5-diphenylimidazole (**2**) was produced by methylation of 4,5-diphenylimidazole (**1**). Subsequently, compound **3** was furnished by formylation of **2** with dimethylformamide (DMF).<sup>23</sup> Meanwhile, the heterocyclic quaternary salts of **5a** and **5b** were carried out in the reaction of the 3-bromopropionic acid with 2-methylbenzoxazole (**4a**) and 2-methylbenzothiazole (**4b**), respectively.<sup>40,41</sup> The newly discovered dyes of **6a** and **6b** were prepared through condensation reactions of **3** and **4a** as well as **3** and **4b** in the presence of pyridine, respectively. The identification of the new compounds was determined by <sup>1</sup>H NMR, <sup>13</sup>C NMR and high resolution mass spectrometry (see Methods and materials section for a full description and data shown in the ESI†).

### 3.2. Spectral properties

With compounds **6a** and **6b** in hand, we investigated the effect of atom substitutions (oxygen and sulfur) on the photophysical properties. Since dyes **6a** and **6b** are ionic molecules, they are soluble in polar solvent like dimethyl sulfoxide (DMSO). Fig. 1 revealed the concentration-dependent UV-vis absorption spectra of **6a** and **6b** in DMSO. It was found that the absorption peak of **6a** was 470 nm at 5  $\mu$ M. In comparison, the dye **6b** exhibited a peak at 475 nm, with a shoulder at approximately 580 nm, which is similar to that of unsymmetrical trimethine cyanine dyes reported by Ge and Lu *et al.*<sup>44</sup> The absorption bands were slight bathochromic shift to longer wavelengths by



Scheme 1 Synthetic routes of dyes **6a** and **6b**.



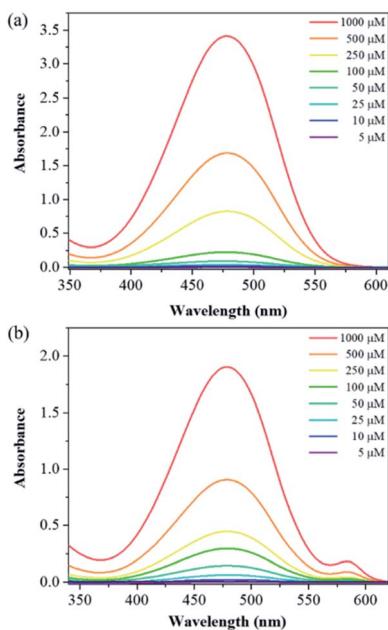


Fig. 1 Concentration-dependent UV-vis absorption spectra of (a) **6a** and (b) **6b** in DMSO.

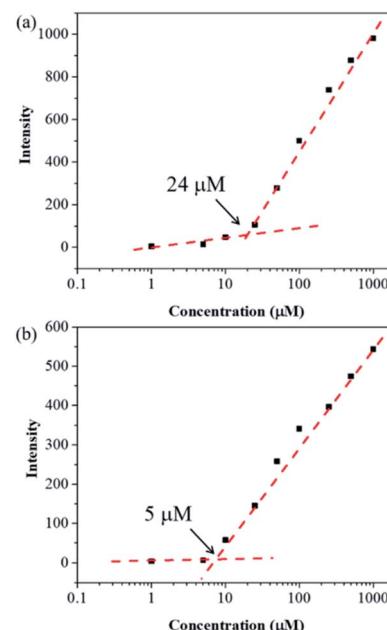


Fig. 3 CAC of (a) **6a** and (b) **6b** in DMSO.

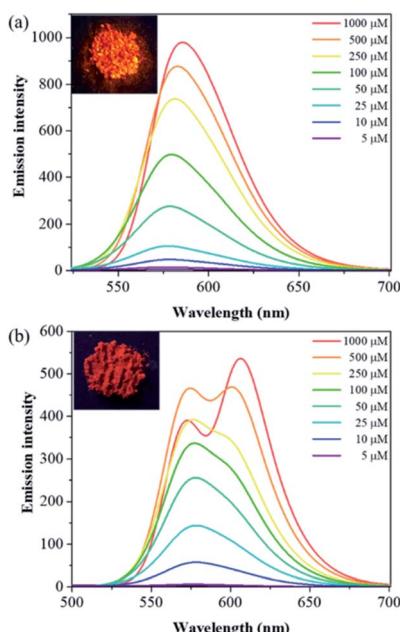


Fig. 2 Concentration-dependent fluorescence emission spectra of (a) **6a** and (b) **6b** in DMSO with excitation at wavelengths of 470 and 475 nm, respectively. The insets are the photographs of solid powders of **6a** and **6b** taken under UV illumination.

increasing the concentration of **6a** and **6b**, which is attributed to the ground state aggregate species were formed in both **6a** and **6b**. The corresponding fluorescence emission spectra of **6a** were obtained with excitation wavelength at 470 nm and displayed in Fig. 2a. As the concentration increased from 5 to 1000 μM, the emission intensities were enhanced *ca.* 1000 fold, indicating

that the dye of **6a** was AIE material. In contrast, the dual-wavelength emission was found to enhance emission intensities by increasing the concentration of **6b** (Fig. 2b). The insets of Fig. 2 showed the photographs of solid powders of **6a** and **6b** taken under UV irradiation at 365 nm, with bright orange-red and red emissions, respectively. Additionally, the excitation spectra were recorded for the emission wavelength of 620 nm, and the higher the concentration, the higher the excitation spectrum intensity at the 580 nm wavelength (Fig. S1†).

In order to get more deep understanding of the aggregate abilities and properties of dyes **6a** and **6b**, we studied their critical aggregation concentration (CAC) and the volume fraction of DMSO ( $x_{\text{DMSO}}$ ) in the binary water/DMSO mixtures. Fig. 3 presented that the CAC of **6a** and **6b** were 24 μM and 8 μM, respectively, suggesting that compound **6b** has more tendency to aggregate than **6a** in pure DMSO. The solubility of the molecule in a polar solvent can be confirmed from the dipole moment values. Therefore, the dipole moments of **6a** and **6b** cations were determined by two different calculation methods. The calculated dipole moments of **6a** and **6b** cations were 12.5897 D and 12.1788 D, respectively, by using DFT/B3LYP/6-31G(d,p) basis set. While the dipole moments calculated through HF/6-31G(d,p) method were found to be 8.8510 D for **6a** cation and 8.2871 D for **6b** cation. Moreover, the octanol–water partition coefficient ( $\text{clog } P$ ) was also used to evaluate the polarity of compounds, and it is known that the more polar, hydrophilic compound will have a lower  $\text{clog } P$ .<sup>45,46</sup> The  $\text{clog } P$  values of **6a** and **6b** cations were 1.2339 and 1.8299, respectively. These results suggested that compound **6a** would be more polar than **6b**, consequently leading to higher solubility in DMSO. To further examine the relationship between the ionic molecules and the solvent polarity, the various water/DMSO mixtures were conducted at 50 μM, based on the CAC evaluation of **6a** and **6b**.

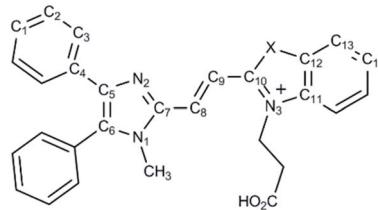


in DMSO. As can be seen from Fig. 4, the emission intensities were increased in lower polarity systems, showing that the AIE characteristics in the aggregate states of **6a** and **6b**. It is worth noting that **6a** with weak fluorescence intensity in pure water, owing to its polar molecule structure. By increasing the DMSO fraction, we can observe the considerable enhancement of the emission intensity (Fig. 4a). On the other hand, **6b** exhibited less polarity with respect to **6a**, which is expected to produce an aggregate state in water and result in bright emission (Fig. 4b). The calculation results of the dipole moments are comparable with the experimental observations.

### 3.3. Theoretical calculations

To achieve more insight into the intrinsic properties of the molecules **6a** and **6b**, molecular calculations were performed using the Gaussian program. The geometries were optimized at DFT/B3LYP/6-31G(d,p) level of theory. The optimized structures of **6a** and **6b** cations were displayed in Fig. S2,† the corresponding bond lengths, bond angles as well as dihedral angles were revealed in Table 1. The calculated bond lengths of C<sub>10</sub>–O and C<sub>10</sub>–S were 1.35 and 1.75 Å, respectively. The bond angles of C<sub>10</sub>–O–C<sub>12</sub> and C<sub>10</sub>–S–C<sub>12</sub> were 108.25° and 91.61°, respectively. These results suggest that the heteroatoms of oxygen and sulfur show a significant influence on the molecular structures of **6a** and **6b**. Furthermore, the dihedral angles of C<sub>4</sub>–C<sub>5</sub>–N<sub>2</sub>–C<sub>7</sub>, C<sub>7</sub>–C<sub>8</sub>–C<sub>9</sub>–C<sub>10</sub> and C<sub>10</sub>–O–C<sub>12</sub>–C<sub>13</sub> were in the range of 178.90° to 179.85° for **6a**; C<sub>4</sub>–C<sub>5</sub>–N<sub>2</sub>–C<sub>7</sub>, C<sub>7</sub>–C<sub>8</sub>–C<sub>9</sub>–C<sub>10</sub> and C<sub>10</sub>–S–C<sub>12</sub>–C<sub>13</sub> were in the range of 179.21° to 179.92° for **6b**. These values of dihedral angles for **6a** and **6b** indicated that they are π-

Table 1 Selected bond lengths (R, Å), bond angles (D, °) and dihedral angles (Di, °) of **6a**<sup>+</sup> and **6b**<sup>+</sup><sup>a</sup>



	<b>6a</b> <sup>+</sup> (DFT) <sup>b</sup>	<b>6b</b> <sup>+</sup> (DFT) <sup>b</sup>	<b>6a</b> <sup>+</sup> (HF) <sup>c</sup>	<b>6b</b> <sup>+</sup> (HF) <sup>c</sup>
R (C <sub>1</sub> –C <sub>2</sub> )	1.397	1.397	1.386	1.386
R (C <sub>2</sub> –C <sub>3</sub> )	1.392	1.392	1.383	1.383
R (C <sub>3</sub> –C <sub>4</sub> )	1.406	1.406	1.392	1.392
R (C <sub>4</sub> –C <sub>5</sub> )	1.472	1.471	1.480	1.480
R (C <sub>5</sub> –C <sub>6</sub> )	1.428	1.427	1.391	1.390
R (C <sub>6</sub> –N <sub>1</sub> )	1.364	1.365	1.355	1.357
R (N <sub>1</sub> –C <sub>7</sub> )	1.392	1.391	1.362	1.361
R (C <sub>5</sub> –N <sub>2</sub> )	1.344	1.345	1.339	1.341
R (N <sub>2</sub> –C <sub>7</sub> )	1.342	1.342	1.308	1.308
R (C <sub>7</sub> –C <sub>8</sub> )	1.414	1.416	1.433	1.435
R (C <sub>8</sub> –C <sub>9</sub> )	1.381	1.383	1.351	1.351
R (C <sub>9</sub> –C <sub>10</sub> )	1.404	1.412	1.426	1.434
R (C <sub>10</sub> –N <sub>3</sub> )	1.359	1.362	1.324	1.324
R (N <sub>3</sub> –C <sub>11</sub> )	1.405	1.407	1.407	1.410
R (C <sub>11</sub> –C <sub>12</sub> )	1.392	1.404	1.371	1.383
R (C <sub>10</sub> –X)	1.352	1.753	1.312	1.726
R (X–C <sub>12</sub> )	1.380	1.753	1.368	1.744
R (C <sub>12</sub> –C <sub>13</sub> )	1.383	1.395	1.374	1.387
R (C <sub>13</sub> –C <sub>14</sub> )	1.398	1.392	1.386	1.378
D (C <sub>5</sub> –N <sub>2</sub> –C <sub>7</sub> )	107.11	107.14	106.98	106.95
D (C <sub>6</sub> –N <sub>1</sub> –N <sub>7</sub> )	107.25	107.28	106.65	106.66
D (C <sub>10</sub> –N <sub>3</sub> –C <sub>11</sub> )	108.49	114.77	107.91	114.33
D (C <sub>10</sub> –X–C <sub>12</sub> )	108.25	91.61	108.77	91.31
Di (C <sub>4</sub> –C <sub>5</sub> –N <sub>2</sub> –C <sub>7</sub> )	179.22	179.21	179.99	179.97
Di (C <sub>7</sub> –C <sub>8</sub> –C <sub>9</sub> –C <sub>10</sub> )	178.90	179.52	179.90	178.53
Di (C <sub>10</sub> –X–C <sub>12</sub> –C <sub>13</sub> )	179.85	179.92	179.91	179.89

<sup>a</sup> For **6a**<sup>+</sup>, X is oxygen atom; for **6b**<sup>+</sup>, X is sulfur atom. <sup>b</sup> DFT/B3LYP/6-31G(d,p). <sup>c</sup> HF/6-31G(d,p).

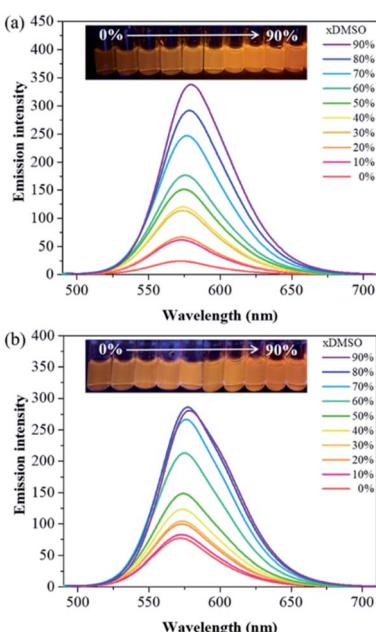


Fig. 4 Fluorescence emission spectra of (a) **6a** and (b) **6b** in water/DMSO mixtures (concentration = 50 μM;  $\lambda_{\text{ex}} = 470$  nm for **6a** and  $\lambda_{\text{ex}} = 475$  nm for **6b**). The insets are the photographs of **6a** and **6b** in solvent mixtures of various fractions (xDMSO), taken under UV irradiation at 365 nm.

conjugated compounds. Notably, the dihedral angles of the phenyl rings at 4, 5-position of imidazole and imidazole core were in the range of 28.85° to 55.69° for **6a** and **6b** (Fig. S2†), implying that the twisted conformations would be responsible for the AIE characteristic of dyes **6a** and **6b** (Fig. 2 and 4).<sup>23</sup> Meanwhile, HF/6-31G(d,p) method was achieved for comparison with DFT results, the calculated bond lengths, bond angles and dihedral angles values obtained with DFT and HF are comparable (Fig. S2† and Table 1).

Since the oxygen and sulfur substitutions can dramatically influence molecular structures of **6a** and **6b**, the natural bond orbital (NBO) analysis was used to further investigate the charge delocalization and noncovalent interactions.<sup>47</sup> NBO study offers an efficient method for investigating intra- and intermolecular bonding and interaction among bonds. It provides a convenient basis for understanding the arrangement of molecules in the aggregate states.<sup>48,49</sup> As can be seen from Table 2, the NBO charge of the oxygen atom was –0.43, while the charge of the

Table 2 Selected NBO charge distribution of **6a<sup>+</sup>** and **6b<sup>+</sup>**<sup>a</sup>

	<b>6a<sup>+</sup></b> (DFT) <sup>b</sup>	<b>6b<sup>+</sup></b> (DFT) <sup>b</sup>	<b>6a<sup>+</sup></b> (HF) <sup>c</sup>	<b>6b<sup>+</sup></b> (HF) <sup>c</sup>
C <sub>1</sub>	-0.21663	-0.21670	-0.21268	-0.21305
C <sub>2</sub>	-0.23122	-0.23140	-0.23099	-0.23111
C <sub>3</sub>	-0.19996	-0.20031	-0.19494	-0.19557
C <sub>4</sub>	-0.09745	-0.09674	-0.10388	-0.10296
C <sub>5</sub>	0.17663	0.17671	0.16744	0.16804
C <sub>6</sub>	0.22322	0.22190	0.26219	0.25901
C <sub>7</sub>	0.32061	0.32588	0.30455	0.31226
C <sub>8</sub>	-0.17008	-0.17616	-0.04400	-0.05185
C <sub>9</sub>	-0.35870	-0.33164	-0.45544	-0.42530
C <sub>10</sub>	0.61046	0.08347	0.79425	0.21800
C <sub>11</sub>	0.11319	0.14996	0.12736	0.17738
C <sub>12</sub>	0.29262	-0.22928	0.31879	-0.24872
C <sub>13</sub>	-0.25362	-0.22158	-0.24457	-0.20471
C <sub>14</sub>	-0.21563	-0.21512	-0.21352	-0.21879
N <sub>1</sub>	-0.33428	-0.33496	-0.40310	-0.40407
N <sub>2</sub>	-0.44915	-0.45462	-0.46350	-0.47120
N <sub>3</sub>	-0.37692	-0.35749	-0.47700	-0.45253
X	-0.42640	0.53972	-0.52358	0.52132

<sup>a</sup> For **6a<sup>+</sup>**, X is oxygen atom; for **6b<sup>+</sup>**, X is sulfur atom. <sup>b</sup> DFT/B3LYP/6-31G(d,p). <sup>c</sup> HF/6-31G(d,p).

sulfur atom was 0.54 using DFT calculation (the HF method had the same trend as the DFT data). Moreover, the NBO charges of the three nitrogen atoms of **6a** and **6b** were in the range from -0.33 to -0.45, suggesting that it might have electrostatic interaction between sulfur and nitrogen atoms. It is probably due to the opposite charges of the heteroatoms in **6b**, leading to an increase the tendency to aggregate and the formation of the dual emission. The visualization of charge distributions of **6a** and **6b** were depicted in Fig. 5 and S3,† it was clear to perceive

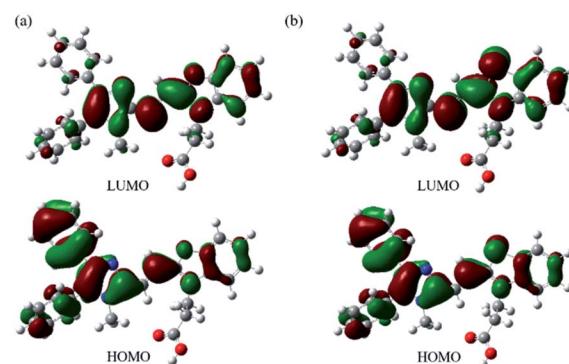


Fig. 6 Graphical demonstration of HOMO and LUMO energy levels of (a) **6a** and (b) **6b** calculated by using DFT at the B3LYP/6-31G(d,p) level.

that the positive and negative charged electrostatic potential in the molecules. The NBO charge analysis presented here explains the experimental observations of CAC and photophysical properties both in solution and in the solid state (Fig. 2–4).

Furthermore, from the calculated results of the dihedral angles in Table 1, we note that the compounds of **6a** and **6b** could be  $\pi$ -conjugated molecules. Therefore, we observed the electron densities of their highest occupied molecular orbitals (HOMO) are distributed in the whole molecule backbone, and that of their lowest unoccupied molecular orbitals (LUMO) are located on the core of imidazole and benzoxazolyl unit (as well as the core of imidazole and benzothiazolyl unit). Compared with an oxygen atom, the sulfur atom has a noticeable contribution to its HOMO and LUMO components (Fig. 6). Table S1† summarizes the TD-DFT results for the lowest energy vertical excitation. The transition energies of **6a** and **6b** cations obtained with DFT/B3LYP/6-31G(d,p) were 2.3319 eV and 2.2435 eV, respectively, indicating that compound **6b** exhibits longer wavelength absorption than that of **6a**. Additionally, HF/6-31G(d,p) was also tested for **6a** and **6b**, which provides qualitatively similar results. The calculation results obtained by DFT/B3LYP/6-31G(d,p) and HF/6-31G(d,p) agree with the experimental data of their optical properties.

#### 4. Conclusions

In summary, we have newly synthesized the benzoxazolyl-imidazole and benzothiazolyl-imidazole conjugates, which show aggregation-induced emission (AIE) features in their solid and aggregate states. The effect of heteroatoms (oxygen and sulfur) on the photophysical properties as well as molecular geometries were investigated. It was found that the oxygen and sulfur substitutions can dramatically influence molecular structures and polarities of the dyes, leading to the different degree of AIE phenomenon. Since the benzothiazolyl-imidazole molecule has lower polarity than benzoxazolyl-imidazole, the dye bearing benzothiazolyl group shows higher emission intensity and dual emission in an aqueous solution, which accompanying the more minor AIE enhancement and red solid is observed. Theoretical calculation results suggest that the

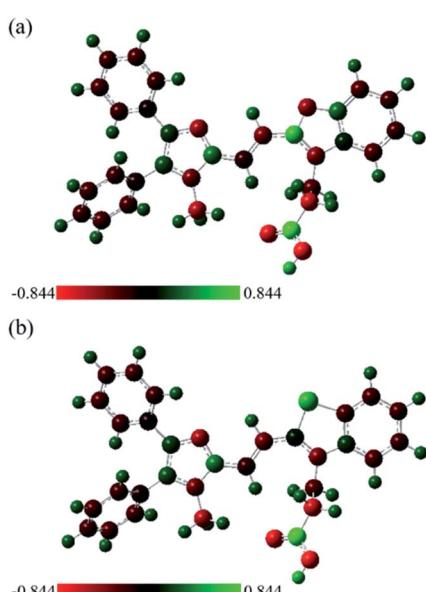


Fig. 5 NBO charge distribution on the optimized (a) **6a<sup>+</sup>** and (b) **6b<sup>+</sup>** designated by color change on the atoms with color scheme and scale (red for negative charge and green for positive charge), determined using DFT at the B3LYP/6-31G(d,p) level.



benzothiazolyl-imidazole molecule might have electrostatic interaction between sulfur and nitrogen atoms, explaining the experimental observations of lower critical aggregation concentration and photophysical properties both in solution and solid state. The theoretical calculations are in agreement with the experimental data, thus demonstrating a potent strategy to gain a deep understanding of the structure–property relationship to design solid-state fluorescent materials.

## Conflicts of interest

There are no conflicts to declare.

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