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Bis(imino)-6,7-dihydro-5*H*-quinoline-cobalt complexes as highly active catalysts for the formation of vinyl-terminated PE waxes; steps towards inhibiting deactivation pathways through targeted ligand design†

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A set of five related bis(imino)-6,7-dihydro-5*H*-quinoline-cobalt(II) complexes, [2-(ArN = CPh)-8-(NAr)-C₉H₈N]CoCl₂ (Ar = 2,6-Me₂C₆H₃ **Co1**, 2,6-Et₂C₆H₃ **Co2**, 2,6-*i*-Pr₂C₆H₃ **Co3**, 2,4,6-Me₃C₆H₂ **Co4**, 2,6-Et₂-4-MeC₆H₂ **Co5**), have been synthesized in reasonable yield by the template reaction of cobalt(II) chloride hexahydrate, 2-benzoyl-6,7-dihydro-5*H*-quinolin-8-one and the corresponding aniline. The molecular structures of **Co1** and **Co4** highlight both the differences in the two imino-carbon environments (phenyl-capped chain vs. cyclic) and also the steric properties exerted by the bulky *N*_{imine}-aryl groups. On pre-treatment with either modified methylaluminoxane (MMAO) or methylaluminoxane (MAO), all complexes proved productive catalysts for the polymerization of ethylene. In particular, **Co1**/MAO was the most active reaching a very high level of 1.62 × 10⁷ g PE per mol (Co) per h over a 30 minute run time. Owing to the presence of the imino-phenyl substituent, **Co1**–**Co5** were able to exhibit good thermal stability by displaying appreciable catalytic activity at temperatures between 50 and 80 °C, generating polyethylenes with narrow dispersities (*M_w*/*M_n* range: 1.66–3.28). In particular, the least sterically bulky precatalysts, **Co1** and **Co4** formed polyethylene waxes (*M_w* range: 1.94–5.69 kg per mol) with high levels of vinyl unsaturation as confirmed by high temperature ¹H/¹³C NMR spectroscopy and by IR spectroscopy.

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Introduction

Late transition metal catalyzed polymerization of ethylene has been the subject of widespread interest^{1–5} ever since the groundbreaking discoveries that α-diimino-Ni(II)/Pd(II)^{1a} and bis(imino)pyridyl-Fe(II)/Co(II) complexes, under suitable activation, can exhibit very high catalytic activities. With particular regard to bis(imino)pyridyl-cobalt(II) halides (**A**, Chart 1),^{1b,2a} it

is often considered that the highly electrophilic nature of the active catalyst facilitates a low barrier towards ethylene insertion while bulky *ortho*-substituted *N*-aryl groups inhibit chain

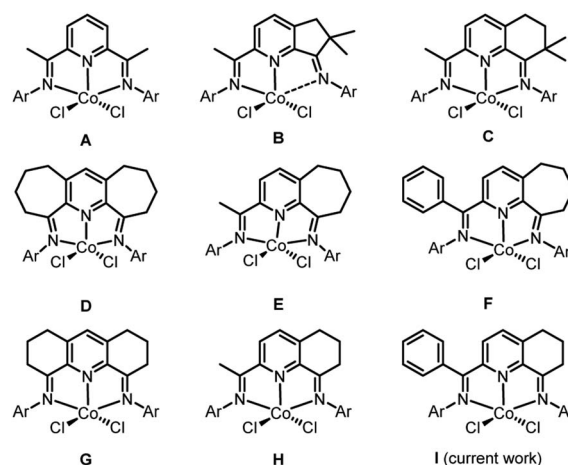


Chart 1 2,6-Bis(arylimino)pyridine-cobalt(II) halide **A** and its fused counterparts **B**–**I**.

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† Electronic supplementary information (ESI) available: X-ray crystallographic data. CCDC 2105336 (**Co1**) and 2105337 (**Co4**). For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d1ra07279j



transfer resulting in the formation of high molecular weight polyethylene. As a result, a wide array of substituent variations has over the years been made to the bis(imino)pyridine ligand skeleton.^{6,7}

Besides high molecular weight polyethylene products, linear polyethylene waxes of narrow dispersity and incorporating vinyl chain ends have gained increasing attention owing to their various applications such as lubricants and colorants for plastics processes.⁸ In view of this, alternative iron and cobalt ethylene polymerization catalysts have been investigated as a means to target such low molecular weight polymers (typical M_w range: 1–5 kg per mol). In particular, cobalt catalysts bearing bis(imino)pyridines fused with carbocyclic rings have shown a propensity towards the formation of such materials; broader polymer dispersities can be a drawback with their iron counterparts.⁶ However, the specific ring size of the fused carbocycle and its substitution pattern can be influential on not only the polymer specification but also the performance of the catalyst. For example, precatalysts **B**^{9a} and **C**^{9b} (Chart 1), that contain dimethyl-substituted five and six-membered fused rings respectively, exhibit quite different productivities while the molecular weight of their polyethylenes fall outside the desirable range for a polyethylene wax. On the other hand, precatalysts containing doubly or singly fused seven-membered rings (**D–F**,^{9c,9d,9e} Chart 1) or six-membered rings devoid of methyl substituents (**G**,^{9f} **H**,^{9g,9h} Chart 1) are able to form the target waxes whilst maintaining, in general, good catalytic performance. As an additional benefit of the fused ring is its ability to influence the thermal stability of the catalyst allowing the polymerization process to operate at a higher temperature. In a similar way, modifications to the imino-*C* substituent in the *N,N,N*-ligand can be influential on thermostability. For instance, **E** exhibits its optimal activity for ethylene polymerization at 50 °C,^{9d} while its phenyl-substituted counterpart **F** operates most effectively at 70 °C.^{9e}

In this article, we return to the cyclohexyl-fused **H**-type family of precatalysts with a view to replacing the imino *C*-methyl substituent with a phenyl group. It was envisaged that this change could impede deactivation pathways and in-turn increase the thermostability of the catalysts.^{9e,9g} Furthermore, it was of interest to examine how this modification would impact on catalytic activity and its ability to form polyethylene waxes. In particular, we target five examples of 2-(phenyl(arylimino)methyl)-8-arylimino-6,7-dihydro-5*H*-quinoline-cobalt(II) chloride (**I**, Chart 1) in which the steric and electronic properties of the *N*-aryl groups have been varied. In addition, a full catalytic evaluation for the polymerization of ethylene is reported that explores not only the effect of precatalyst structure but also the co-catalyst type, precatalyst/co-catalyst ratio, run time, temperature and pressure. The intrinsic properties of the polyethylenes will also be discussed (e.g., molecular weight, dispersity, melting temperatures, vinyl content) as will the synthetic and characterization data for the new cobalt complexes.

Results and discussion

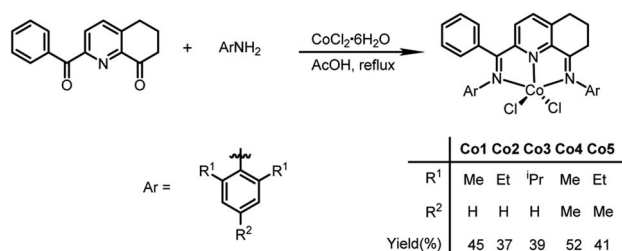
Synthesis and characterization of Co1–Co5

The 2-(phenyl(arylimino)methyl)-8-arylimino-6,7-dihydro-5*H*-quinoline-cobalt(II) chloride complexes, [2-(ArN = CPh)-8-(NAr)-C₉H₈N]CoCl₂ (Ar = 2,6-Me₂C₆H₃ **Co1**, 2,6-Et₂C₆H₃ **Co2**, 2,6-*i*-Pr₂C₆H₃ **Co3**, 2,4,6-Me₃C₆H₂ **Co4**, 2,6-Et₂-4-MeC₆H₂ **Co5**), were prepared in reasonable yield (37–52%) by the template reaction of 2-benzoyl-6,7-dihydro-5*H*-quinolin-8-one, cobalt dichloride hexahydrate and the corresponding aniline in acetic acid at reflux for 6 hours (Scheme 1). Attempts at trying to isolate the free bis(imino)-6,7-dihydro-5*H*-quinoline compounds by condensation of 2-benzoyl-6,7-dihydro-5*H*-quinolin-8-one with the aniline, proved unsuccessful.¹⁰ Complexes **Co1–Co5** have been characterized by a combination of IR spectroscopy, microanalysis and single crystal X-ray diffraction.

For the purposes of the X-ray determination, single crystals of **Co1** and **Co4** were grown by diffusing diethyl ether into dichloromethane solutions of the corresponding complex. Views of **Co1** and **Co4** are shown in Fig. 1 and 2; selected bond lengths and angles are gathered in Table 1.

Owing to the structural similarity, both **Co1** and **Co4** will be discussed together. In common to both structures, a cobalt center is surrounded by three nitrogen donors belonging to the tridentate 2-(phenyl(arylimino)methyl)-8-arylimino-6,7-dihydro-5*H*-quinoline chelating ligand (aryl = 2,6-dimethylphenyl **Co1**, 2,4,6-trimethylphenyl **Co4**) and two chlorides to complete a geometry best described as *pseudo*-square pyramidal. The three nitrogen atoms N1, N2, N3 and Cl1 fill the basal plane while Cl2 occupies the apical position. The cobalt atom sits above the square base by 0.600 Å for **Co4** and 0.496 Å for **Co1** in a manner akin to that seen with related (*N,N,N*)CoCl₂ complexes.^{9g} The exterior Co–N_{imine} bond lengths [2.213, 2.263 Å (**Co4**), 2.238, 2.209 Å (**Co1**)] are longer than the central Co–N_{pyridine} bond length [2.024 Å (**Co4**), 2.040 Å (**Co1**)] in line with the stronger donor ability of the pyridine. As with its analogues,^{11,12} the *N*-aryl groups are inclined towards perpendicular with respect to the *N,N,N*-Co coordination plane with dihedral angles of 71.06°, 76.73° **Co4** and 72.59°, 65.67° **Co1**. On account of the sp³-hybridization of carbon atoms, C5, C6 and C7, there is some puckering of this section of the fused 6-membered ring that folds away from the apical Cl2 ligand in both structures. There are no intermolecular contacts of significance.

The FT-IR spectra of **Co1–Co5** exhibited absorption bands for the two C=N_{imine} groups around 1610 cm^{−1}, values that



Scheme 1 Synthesis of complexes **Co1–Co5** via a one-pot route.



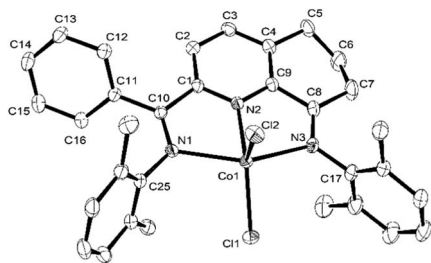


Fig. 1 OLEX2 representation of **Co1**. The thermal ellipsoids are shown at the 30% probability level, while all hydrogen atoms and a molecule of CH_2Cl_2 have been removed for clarity.

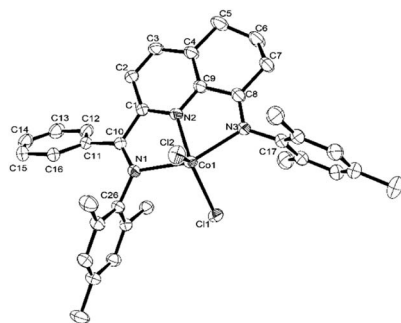


Fig. 2 OLEX2 representation of **Co4**. The thermal ellipsoids are shown at the 30% probability level, while all hydrogen atoms and two molecules of CH_2Cl_2 have been removed for clarity.

Table 1 Selected bond lengths and angles for **Co1** and **Co4**

	Co1	Co4
Bond lengths (Å)		
Co(1)–N(1)	2.238(2)	2.213(3)
Co(1)–N(2)	2.040(2)	2.024(3)
Co(1)–N(3)	2.209(2)	2.263(3)
Co(1)–Cl(1)	2.2579(8)	2.2475(10)
Co(1)–Cl(2)	2.2775(8)	2.2808(11)
Bond angles (deg)		
N(1)–Co(1)–N(2)	74.06(9)	75.31(11)
N(1)–Co(1)–N(3)	148.11(9)	148.73(10)
N(2)–Co(1)–N(3)	75.80(9)	75.15(11)
N(1)–Co(1)–Cl(2)	95.66(6)	99.48(8)
N(2)–Co(1)–Cl(2)	100.14(7)	105.69(9)
N(3)–Co(1)–Cl(2)	99.65(7)	98.14(9)
N(1)–Co(1)–Cl(1)	101.54(6)	98.24(8)
N(2)–Co(1)–Cl(1)	146.28(7)	136.48(10)
N(3)–Co(1)–Cl(1)	97.69(7)	96.08(9)
Cl(2)–Co(1)–Cl(1)	113.58(4)	117.79(4)

are characteristic of cobalt(II) bound imines.^{9b,9c} Further examination of the spectra revealed the absence of any absorptions corresponding to complexed $\text{C}=\text{O}$ groups or free diketone. In addition, the elemental analysis data were in good agreement with the proposed formulae for all five cobalt complexes.

Exploring the scope of **Co1–Co5** as precatalysts in ethylene polymerization

To facilitate the ethylene polymerization study, a direct *in situ* activation process was followed whereby the cobalt precatalyst (**Co1–Co5**) was pre-activated by the addition of co-catalyst and the resulting active species used to mediate the polymerization. In view of previous studies, most bis(imino)pyridine-cobalt(II) halide complexes have been shown to exhibit their optimal performance for ethylene polymerization when activated with either modified methyl aluminoxane (MMAO) or methylaluminoxane (MAO).^{13,14} With this in mind, the current work focuses on using these two types of co-catalyst with **Co4** selected as the test precatalyst to develop a suitable set of conditions that could then be used to screen the remaining cobalt complexes. For each **Co4**/co-catalyst combination, variations in reaction temperature, Al : Co molar ratio, run time and pressure will be independently conducted; the results are compiled in Tables 2 and 3. As a matter of routine, the polyethylene generated was characterized by gel permeation chromatography (GPC) and differential scanning calorimetry (DSC). In addition, high temperature NMR spectroscopy as well as IR spectroscopy was used to explore the microstructural properties of selected samples of polyethylene.

Catalytic evaluation of **Co1–Co5 using MAO.** With MAO as the co-catalyst, we first set about establishing conditions that would deliver the highest polymerization activity for **Co4**; these reaction conditions would then be extended to the screening of **Co1–Co3** and **Co5** (Table 2).

As the temperature of the polymerization run can play a key role in catalytic performance,^{11b,15} we first explored its effect on the catalytic activity of **Co4**/MAO. Specifically, five runs were conducted at fixed temperatures between 30 °C and 70 °C with the Al : Co molar ratio at 2000 : 1 (entries 1–5, Table 2). A peak of 5.80×10^6 g PE per mol (Co) per h was observed with the run temperature set at 50 °C. Even at 70 °C, the catalyst still performed well with only a 37% loss of activity. With respect to the molecular weights of the resulting polyethylenes, these gradually declined from 5.69 kg per mol to 2.93 kg per mol as the temperature was increased from 30 °C to 70 °C in keeping with an increased rate of chain transfer as the temperature was raised (Fig. 3).^{15b,16} Nevertheless, the distributions of the polyethylenes obtained across the temperature range remained reasonably narrow ($M_w/M_n = 2.09$ –2.93) highlighting the good control and single site-like behavior of the cobalt active species.

With the polymerization temperature retained at 50 °C, the impact of the Al : Co molar ratio was investigated by performing the runs at ratios of between 1750 : 1 and 2750 : 1 (entries 3 and 6–9, Table 2). Inspection of the data for **Co4**/MAO revealed the highest activity of 8.65×10^6 g PE per mol (Co) per h was achieved at a molar ratio of 2250 : 1. In terms of molecular weight of the polymer, it was noted that this steadily dropped as the Al : Co ratio was increased, which may suggest that chain transfer from the active cobalt species to aluminum co-catalyst is operative.^{9d,9e}

To explore the profile of **Co4**/MAO with respect to time, the polymerizations were undertaken at five distinct run times



Table 2 Catalytic evaluation of Co1–Co5 using MAO as co-catalyst^a

Entry	Precat.	Al : Co	<i>T</i> (°C)	<i>t</i> (min)	Activity ^b	<i>M_w</i> ^c	<i>M_w</i> / <i>M_n</i> ^c	<i>T_m</i> ^d (°C)
1	Co4	2000	30	30	2.20	5.69	2.27	126.6
2	Co4	2000	40	30	3.25	4.39	2.28	125.9
3	Co4	2000	50	30	5.80	3.64	2.93	125.0
4	Co4	2000	60	30	4.97	3.06	2.03	124.5
5	Co4	2000	70	30	4.23	2.93	2.34	124.1
6	Co4	1750	50	30	3.84	3.77	2.05	125.1
7	Co4	2250	50	30	8.65	3.17	2.05	124.2
8	Co4	2500	50	30	6.79	2.92	1.66	125.0
9	Co4	2750	50	30	5.16	2.77	1.95	123.3
10	Co4	2250	50	5	28.2	3.47	2.00	124.8
11	Co4	2250	50	15	10.4	3.48	2.19	124.9
12	Co4	2250	50	45	4.71	3.30	1.95	124.7
13	Co4	2250	50	60	3.78	3.20	1.95	124.7
14 ^e	Co4	2250	50	30	4.52	3.68	3.28	125.0
15	Co1	2250	50	30	16.18	2.67	2.43	123.3
16	Co2	2250	50	30	2.42	9.91	2.30	129.6
17	Co3	2250	50	30	0.16	146.43	1.67	131.2
18	Co5	2250	50	30	1.48	11.38	2.08	129.3

^a Conditions: 2.0 μmol of cobalt precatalyst, 10 atm of ethylene, 100 mL of toluene. ^b Activity: 10⁶ g PE per mol (Co) per h. ^c *M_w* in kg per mol. *M_w* and *M_w*/*M_n* measured by GPC. ^d Measured by DSC. ^e 5 atm.

namely, 5, 15, 30, 45 and 60 minutes (entries 7 and 10–13, Table 2 and Fig. S1†). After 5 minutes, the catalyst reached peak activity of 28.2×10^6 g PE per mol (Co) per h before steadily losing performance and reaching its lowest level of 3.78×10^6 g PE per mol (Co) per h after 60 minutes. Such a downward trend in activity over time would suggest some gradual deactivation of the catalyst. With the ethylene pressure lowered from 10 to 5 atm, the catalytic activity dropped by nearly a half (entries 14 vs. 7, Table 2), while the molecular weight of the polymer increased from 3.17 kg per mol to 3.68 kg per mol. Moreover, the

dispersity of the polyethylene generated was broader at the lower ethylene pressure, a finding that has been reported elsewhere for a cobalt analogue.¹⁷

By using the favored set of conditions established with Co4/MAO, the remaining cobalt precatalysts, Co1–Co3 and Co5, were additionally evaluated for ethylene polymerization (entries 15–18, Table 2). In general, all the cobalt complexes showed good catalytic activities across a wide range ($0.16\text{--}16.18 \times 10^6$ g PE per mol (Co) per h) with the relative levels following the order: Co1[2,6-di(Me)] > Co4[2,4,6-tri(Me)] > Co2[2,6-di(Et)] >

Table 3 Catalytic evaluation of Co1–Co5 using MMAO as co-catalyst^a

Entry	Precat.	Al : Co	<i>T</i> (°C)	<i>t</i> (min)	Activity ^b	<i>M_w</i> ^c	<i>M_w</i> / <i>M_n</i> ^c	<i>T_m</i> ^d (°C)
1	Co4	2000	40	30	2.27	4.11	2.28	126.3
2	Co4	2000	50	30	3.50	3.44	2.96	124.5
3	Co4	2000	60	30	3.98	2.87	1.67	124.7
4	Co4	2000	70	30	3.13	2.67	1.89	124.1
5	Co4	2000	80	30	2.45	2.59	2.05	124.3
6	Co4	1500	60	30	3.00	2.81	1.77	124.4
7	Co4	1750	60	30	3.25	2.89	1.98	124.8
8	Co4	2250	60	30	3.35	2.90	1.87	124.5
9	Co4	2500	60	30	2.34	2.43	1.92	124.6
10	Co4	2000	60	5	14.52	2.23	1.95	124.6
11	Co4	2000	60	15	6.30	1.94	1.96	124.6
12	Co4	2000	60	45	2.13	2.52	2.28	124.6
13	Co4	2000	60	60	2.05	2.89	2.35	124.7
14 ^e	Co4	2000	60	30	2.44	3.58	1.82	124.8
15	Co1	2000	60	30	7.14	2.43	1.96	123.2
16	Co2	2000	60	30	1.72	10.01	2.24	129.4
17	Co3	2000	60	30	0.18	166.57	2.67	132.7
18	Co5	2000	60	30	1.20	9.87	2.31	129.5

^a Conditions: 2.0 μmol of cobalt precatalyst, 10 atm of ethylene, 100 mL of toluene. ^b Activity: 10⁶ PE per mol (Co) per h. ^c *M_w* in kg per mol. *M_w* and *M_w*/*M_n* measured by GPC. ^d Measured by DSC. ^e 5 atm.



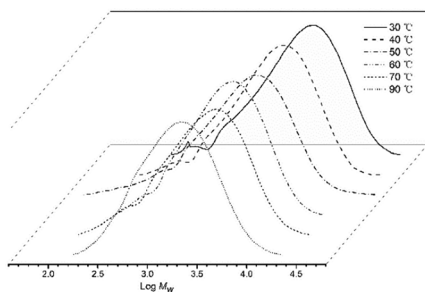


Fig. 3 GPC traces of the polyethylene produced using Co4/MAO at different run temperatures (entries 1–5, Table 2).

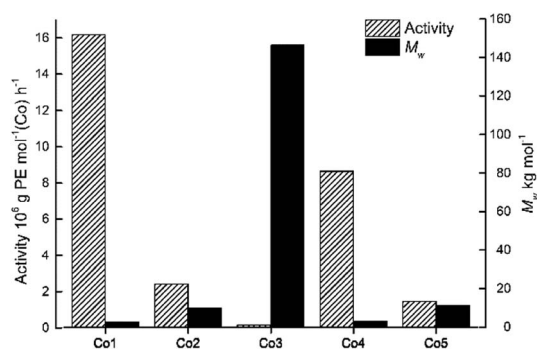


Fig. 4 Bar chart showing catalytic activity and molecular weight of the polyethylene as a function of the precatalyst employed (entries 7 and 15–18, Table 2); MAO used as co-catalyst in each case.

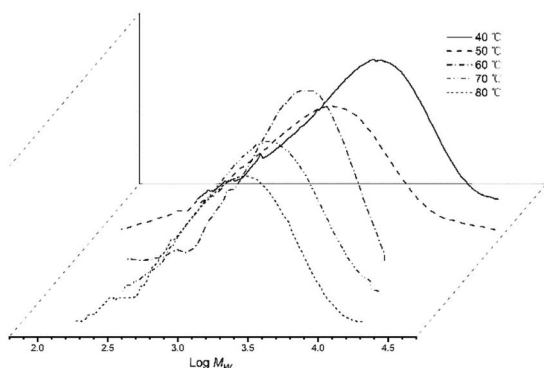


Fig. 5 GPC traces of the polyethylene produced using Co4/MMAO at different run temperatures (entries 1–5, Table 3).

Co5[2,6-di(Et)-4-Me] > Co3[2,6-di(i-Pr)]. It is noteworthy that the activity of Co1[2,6-di(Me)] (16.18×10^6 g PE per mol (Co) per h) exceeded that for Co3[2,6-di(i-Pr)] (0.16×10^6 g PE per mol (Co) per h) by an order of magnitude (Fig. 4).

It would seem that the steric properties exerted by the *ortho*-substituents are influential with increased hindrance generally leading to lower activity, a finding that can be credited to inefficient coordination of the monomer. On the other hand, the molecular weight of the polymers (M_w range: 2.67–146.43 kg per mol) followed the reverse order with respect to the precatalyst: Co3[2,6-di(i-Pr)] > Co5[2,6-di(Et)-4-Me] > Co2[2,6-di(Et)] > Co4

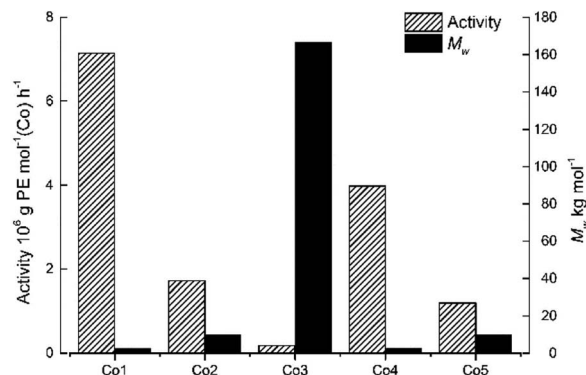


Fig. 6 Bar chart showing catalytic activity and molecular weight of the polyethylene as a function of the precatalyst employed (entries 3 and 15–18, Table 3); MMAO used as co-catalyst in each case.

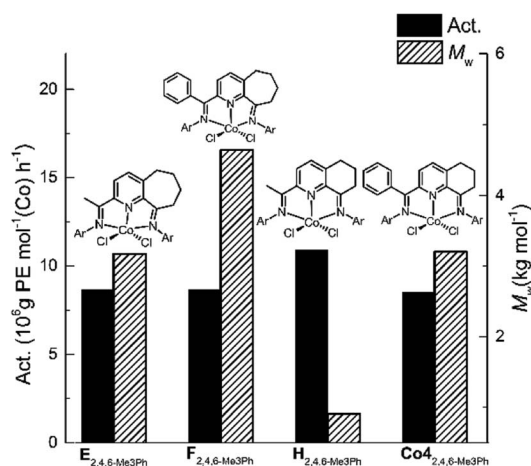


Fig. 7 Comparison of catalytic activity and polymer molecular weight for Co4 (entry 7, Table 2) with that produced by E, F and H (Ar = 2,4,6-Me₃Ph); all precatalysts screened with MAO as the co-catalyst at 10 atm C₂H₄.

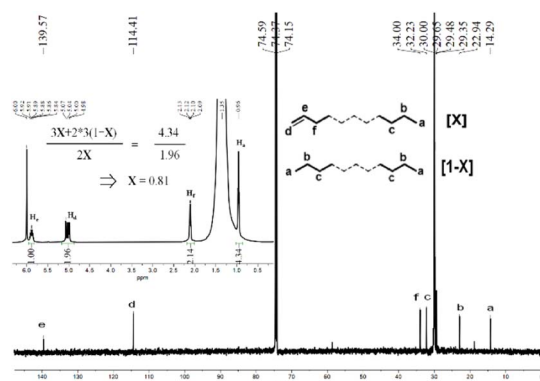


Fig. 8 ¹³C NMR spectrum of the polyethylene wax generated using Co4/MMAO along with an inset showing its ¹H NMR spectrum (entry 3, Table 3); recorded in tetrachloroethane-*d*₂ (δ_C 74.37, δ_H 6.0).

[2,4,6-tri(Me)] > Co1[2,6-di(Me)]. Indeed, the molecular weight of the polymer generated by Co1 (M_w = 2.67 kg per mol) was substantially less than that seen with Co3 (M_w = 146.43 kg per



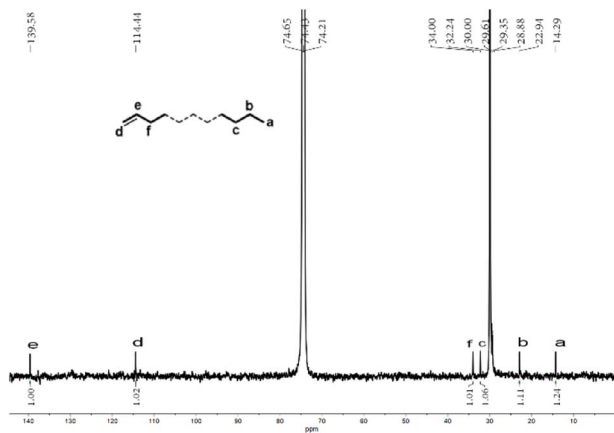


Fig. 9 Inverse-gated decoupled ^{13}C NMR spectrum of the polyethylene wax generated using **Co4**/MAO (entry 3, Table 3); recorded in tetrachloroethane- d_2 (δ_{C} 74.43).

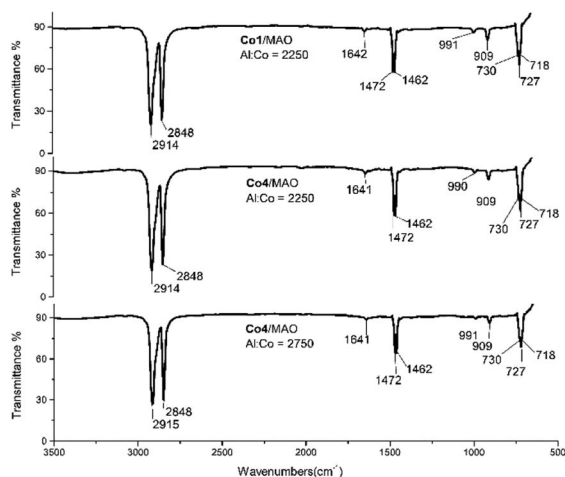


Fig. 10 FT-IR spectra of the polyethylene samples generated using **Co1**/MAO (top, entry 15, Table 2) and **Co4**/MAO (middle, entry 7, Table 2 and bottom, entry 9, Table 2).

mol) which highlights how steric effects can be beneficial to chain propagation. Furthermore, the dispersity (M_w/M_n) of the polyethylene generated using **Co1**–**Co5** fell in the range 1.67–2.96, which again underlines the good control of the polymerizations in accord with single-site active species.

Catalytic evaluation of **Co1–**Co5** using MMAO.** To investigate the effect of co-catalyst on catalyst performance and polymer properties, we also explored the use of MMAO as an activator for **Co1**–**Co5**. As before, **Co4** was employed as the test precatalyst to allow the establishment of a general set of conditions for the polymerization runs; the results are collected in Table 3.

On increasing the run temperature from 40 °C to 80 °C, the topmost level of activity for **Co4**/MMAO (3.98×10^6 g PE per mol (Co) per h) was seen at 60 °C (entries 1–5, Table 3). By comparison with **Co4**/MAO, this optimal temperature was 10 degrees lower albeit with a higher catalytic activity (5.80×10^6 g PE per mol (Co) per h, entry 3, Table 2). By raising the

temperature to 70 °C, a 27% loss of activity was noted with **Co4**/MMAO, while at 80 °C a further 28% loss was seen. With regard to the molecular weight of the polymer, this decreased steadily from 4.11 kg per mol at 40 °C to 2.59 kg per mol at 80 °C. Indeed, this downward trend in molecular weight is similar to that seen for **Co4**/MAO, with temperature-induced chain-transfer likely responsible (Fig. 5). Likewise, a relatively narrow dispersity of the polymers was again observed (M_w/M_n range: 1.67–2.96) which highlights the effective control of the polymerizations as well as single-site nature of the active species.

With the polymerization temperature kept at 60 °C, a series of runs were conducted by varying the amount of MAO with respect to **Co4**. In particular, with the Al : Co molar ratio adjusted between 1500 : 1 and 2500 : 1 (entries 3, 6–9, Table 3), a maximum level of catalytic activity was reached with 2000 molar equivalents. Notably, when the Al : Co molar ratio was increased to 2500 : 1, a noticeable drop in both catalytic activity and the molecular weight was observed which would suggest the onset of chain transfer to aluminum (Fig. S2†).^{9d,9e} In all cases the dispersity of polyethylene remained fairly narrow between 2.43 and 2.90.

To explore the time-activity profile for **Co4**/MMAO, the polymerizations runs were conducted at intervals between 5 and 60 minutes with the temperature and Al : Co molar ratio maintained at 60 °C and 2500 : 1, respectively (entries 3, 10–13, Table 3). The highest activity of 14.52×10^6 g PE per mol (Co) per h was seen in the 5 first minutes which compares with 28.2×10^6 g PE per mol (Co) per h with **Co4**/MAO (entry 10, Table 2). Upon prolonging the reaction time, the polymerization activity smoothly decreased, underlining the gradual deactivation of the cobalt active species; similar observations were seen for **Co4**/MAO and for their imino-*C* methyl counterparts (**H**, Chart 1). Furthermore, the activity of **Co4**/MMAO was observed to dramatically decrease when the ethylene pressure was reduced from 10 atm to 5 atm. Conversely, the molecular weight of the polyethylene generated at 5 atm was higher than that at 10 atm. Similar observations were noted with **Co4**/MAO and indeed the molecular weight of the polymer at 5 atm was comparable (3.68 kg per mol (MAO) vs. 3.58 kg per mol (MMAO)).

To probe the effect of precatalyst structure, the remaining cobalt catalysts, **Co1**–**Co3** and **Co5** were also screened for ethylene polymerization under the optimal conditions ascertained for **Co4**/MMAO. Collectively, the activities fell in the range 0.18 – 7.14×10^6 g PE per mol (Co) per h which is narrower when compared to 0.18 – 16.18×10^6 g PE per mol (Co) per h with MAO. Specifically, the activities decreased in the order, **Co1**[2,6-di(Me)] > **Co4**[2,4,6-tri(Me)] > **Co2**[2,6-di(Et)] > **Co5**[2,6-di(Et)-4-Me] > **Co3**[2,6-di(*i*-Pr)], which is identical to that seen with MAO and indeed similar to that reported elsewhere.¹⁸ Once again steric properties played a key role with greater hindrance resulting in lower activity. As for the molecular weight, a broad range in values from 2.43 kg per mol to 166.57 kg per mol were seen with bulky diisopropyl-containing **Co3** falling at the very top end, while dimethyl-containing **Co1** occupying the bottom; trends that are the inverse of that found for activity (Fig. 6).



In general, the catalytic activities to **Co1**–**Co5** with MAO as co-catalyst were higher than that with MMAO as has been noted with related cobalt catalysts.^{9d,9e,9g} Likewise, the molecular weights of the polyethylenes produced using MAO exceeded that found using MMAO with the exception of **Co2** and **Co3**. To allow a comparison to be made with previously reported bis(imino)pyridine-Co(II) precatalysts that contain six- and seven-membered fused rings,^{9d,9e,9g} the molecular weight and activity data for the 2,4,6-trimethylphenyl-containing precatalysts **H**, **E** and **F** (Chart 1) are displayed along with that for **Co4** in Fig. 7; all precatalysts have been screened with MAO as the co-catalyst at 10 atm C₂H₄.

Analysis of the data reveals that the introduction of a phenyl group at the imino-*C* had a beneficial effect on the molecular weight of the resulting polyethylene. For instance, the molecular weight obtained using **F** ($M_w = 4.6$ kg per mol)^{9e} was clearly higher than for **E** ($M_w = 3.2$ kg per mol),^{9d} while their activity was comparable. In a similar manner, the polyethylene produced using **Co4** ($M_w = 3.2$ kg per mol) was notably higher than that formed by **H** ($M_w = 0.91$ kg per mol).^{9g} This observation may be due to the protection imparted by the phenyl group on the active center.

Microstructural properties of the polyethylene waxes

On inspection of Tables 2 and 3, it is evident that the T_m values for all the polyethylenes produced with either MAO or MMAO exceeded 123 °C. Such a temperature range would suggest high linearity for these polymers.¹⁹ To gather more information about their microstructural properties, representative samples of polyethylene wax obtained from the separate MAO and MMAO investigations were selected, namely that formed using **Co4**/MAO ($M_w = 3.17$ kg per mol, entry 7, Table 2) and **Co4**/MMAO ($M_w = 2.87$ kg per mol, entry 3, Table 3), were characterized by ¹H and ¹³C NMR spectroscopy as well as by inverse-gated decoupled ¹³C NMR spectroscopy. Typically, these spectra were obtained at 100 °C in 1,1,2,2-tetrachloroethane-*d*₂ so as to engender suitable solubility of the polymers.^{9,16}

In the case of the sample obtained using **Co4**/MAO, the ¹H NMR spectrum displayed downfield multiplets at around δ 5.87 ppm (H_c ; integration = 1) and δ 5.04 ppm (H_d ; integration = 2) that are characteristic of a $-\text{CH}=\text{CH}_2$ end group (Fig. S3†). Corroboration of this assignment was provided by the ¹³C NMR spectrum which showed the corresponding unsaturated carbon signals at δ 114.4 ppm and 139.6 ppm (C_e and C_d in Fig. S4†).^{7c,19d} The methyl chain end protons (H_a) were visible at δ 0.96 ppm in the ¹H NMR spectrum with a relative integration of 4.58 which, on account of being >3H, suggested the presence of both vinyl-terminated polyethylene and some fully saturated polymer.^{7c} In an attempt to quantify the relative proportion of vinyl-terminated polyethylene to fully saturated polyethylene, we determined the molar fraction X (where $X = 1$ refers to 100% vinyl-terminated polyethylene) based on the integrations of the signals for H_a and H_d in the ¹H NMR spectrum (Fig. S5†).

A value of 0.79 was determined which implied a clear bias towards a vinyl-terminated polymer in line with β -H elimination constituting the key termination pathway.^{16,19d} As supporting

evidence, the inverse-gated decoupled ¹³C NMR spectrum of the sample displayed an integral ratio for C_d , C_e and C_a as 0.98 : 1.00 : 1.24 (Fig. S5†). Likewise, the sample of polyethylene wax obtained using **Co4**/MMAO revealed similar features with a sharp singlet at δ 30.0 ppm in its ¹³C NMR spectrum and an intense signal at δ 1.35 ppm in its ¹H NMR spectrum in support of linear polymer (Fig. 8). Once again, the value of X (0.81) highlighted the predominance of β -H elimination in chain termination with vinyl-terminated polyethylenes being the major type of chain end. Similarly, the inverse-gated decoupled ¹³C NMR spectrum for the polymer sample supported this finding by revealing a $C_d : C_e : C_a$ integral ratio of 1.00 : 1.02 : 1.24 (Fig. 9).

The two polymer wax samples formed using **Co4**/MAO ($M_w = 3.17$ kg per mol, entry 7, Table 2) and **Co4**/MMAO ($M_w = 2.87$ kg per mol, entry 3, Table 3), were also studied by FT-IR spectroscopy. As shown in Fig. 10 (middle) and Fig. S6 (middle),† weak bands at 1642 cm⁻¹, 990 cm⁻¹ and 909 cm⁻¹ could be assigned to the vinyl group in the form of a vibration for the C=C stretching mode, the in-phase plane vibration and the out of phase plane vibration, respectively. On the other hand, the stronger peaks at 700–750 cm⁻¹ (CH₂ rocking vibration) and at 1440–1490 cm⁻¹ (CH₂ wagging vibration) were characteristic of linear polyethylenes.²⁰ In the same way, polyethylene samples prepared using higher Al : Co molar ratios of **Co4**/MAO ($M_w = 2.77$ kg per mol, entry 9, Table 2) and **Co4**/MMAO ($M_w = 2.43$ kg per mol, entry 9, Table 3) showed related features through the intensity of the vinyl vibrations was visibly less.

In addition, samples of the polyethylene waxes generated using the more active catalysts, **Co1**/MAO ($M_w = 2.67$ kg per mol, entry 15, Table 2) and **Co1**/MMAO ($M_w = 2.43$ kg per mol, entry 15, Table 3), were also studied using IR spectroscopy. Once more, the spectra revealed the presence of vinyl-terminated polyethylenes (Fig. 10 (top) and Fig. S6 (top)†). Clearly, chain termination *via* β -H elimination represents a key pathway for all these polymerizations.

Conclusions

Five examples of **I**-type (Chart 1) cobalt precatalysts namely, **Co1**–**Co5**, each incorporating an imino *C*-phenyl group but differing in the imino *N*-aryl substitution pattern, have been successfully synthesized and characterized. Confirmation of the structural type by single crystal X-ray diffraction was also undertaken in the cases of **Co1** and **Co4**. On activation with MMAO or MAO, all cobalt complexes exhibited high activity with **Co1**/MAO reaching the maximum level of 1.62×10^7 g PE per mol (Co) per h generating strictly linear polyethylenes with narrow dispersity. In addition, all five cobalt(II) precatalysts displayed good thermal stability with **Co4**/MMAO achieving an appreciable activity of 2.45×10^6 g PE per mol (Co) per h when the polymerization run was conducted at a more industrially relevant temperature of 80 °C. It is believed that the introduction of the imino *C*-phenyl substituent improved this thermal stability by inhibiting deactivation pathways. In addition, the molecular weights of the polyethylene were found to fall in an exceptionally broad range (2.43–166.57 kg per mol) which is



controlled by the steric hindrance imparted by the *ortho*-substituents on the *N*-aryl groups. Notably the least sterically hindered **Co1** and **Co4** were able to form narrowly dispersed polyethylene waxes (M_w range: 1.94–5.69 kg per mol) with high levels of vinyl unsaturation. Subsequent investigations will focus on using these ligand frameworks and their derivatives as supports for other polymerization-active late transition metals.

Experimental

General procedures

All operations of air- and/or moisture-sensitive compounds were carried out using standard Schlenk techniques under an atmosphere of nitrogen. Toluene was purified by initially heating it to reflux over sodium (containing a small amount of benzophenone) and then distilled under a nitrogen atmosphere prior to use. Methylaluminoxane (MAO) (1.46 M solution in toluene) and modified methylaluminoxane (MMAO) (1.93 M in *n*-heptane) were purchased from Albemarle Corporation, while high purity ethylene was purchased from Beijing Yansan Petrochemical Company. Other reagents were obtained from Aldrich or local suppliers. ^1H and ^{13}C NMR spectroscopic measurements on the organic compounds were performed on Bruker DMX 400 MHz instruments at room temperature using TMS as an internal standard; δ values are given in ppm. A PerkinElmer System 2000 FT-IR spectrometer was used to collect spectra from 4000 cm^{-1} to 400 cm^{-1} at room temperature with a data interval of 1 cm^{-1} when resolution was set at 4 cm^{-1} . Elemental analysis was carried out with a Flash EA 1112 microanalyzer. Gel permeation chromatography (GPC) was undertaken with an Agilent PLGPC 220 GPC system using triple detection and with a polystyrene standard to determine molecular weights (M_w) and molecular weight distributions (M_w/M_n). Typically, polyethylene samples were dissolved in 1,2,4-trichlorobenzene at $150\text{ }^\circ\text{C}$ and heated for at least two hours to dissolve. The melting temperatures (T_m) of the polyethylene samples were measured by differential scanning calorimeter (DSC) under a nitrogen atmosphere on a PerkinElmer DSC-7 instrument. As a matter of course, polyethylene samples were performed in crimped aluminum pans and data were collected with heat/cool cycles (-20 to $160\text{ }^\circ\text{C}$ and 160 to $-20\text{ }^\circ\text{C}$) at a heating rate of $10\text{ }^\circ\text{C min}^{-1}$. All reported melting temperatures (T_m) were recorded from the second heating cycle. ^1H and ^{13}C NMR spectra of the polyethylene samples was recorded on a Bruker AVIII 500WB instrument. Typically, a weighed amount of polyethylene (20–40 mg) was first dissolved in 1,1,2,2-tetrachloroethane- d_2 (2 mL) with TMS as an internal standard. Spectra were then recorded at $100\text{ }^\circ\text{C}$ on a Bruker AVANCE III 500 MHz instrument in 5 mm standard glass tubes. Operating conditions used for the ^{13}C NMR spectra: number of scans between 2000 and 3000, spectral width 31 250.0 Hz, acquisition time 0.5243 s, relaxation delay 1.5000 s. Inverse-gated decoupled ^{13}C NMR spectra were recorded with 6 mg of the relaxation reagent chromium(III) acetylacetonate. The operating conditions were set as follows: spectral frequency 125.70 MHz, pulse width 10.0 μs , spectral width 21.3675 kHz, acquisition time 0.7668 s, relaxation delay 5.0000 s.

Synthesis of 2-benzoyl-6,7-dihydro-5H-quinolin-8-one

The compound, 2-benzoyl-6,7-dihydro-5H-quinolin-8-one, was prepared in an acceptable yield (*ca.* 30%) from 2-chloro-6,7-dihydro-5H-quinolin-8-one according to a literature method.^{9e} ^1H NMR (400 MHz, CDCl_3 , TMS): δ 8.27 (m, 2H, Py), 8.12 (d, $J = 8.0\text{ Hz}$, 1H, Ph), 7.86 (d, $J = 8.0\text{ Hz}$, 1H, Ph), 7.60 (m, 1H, Ph), 7.57 (m, 2H, Ph), 3.13 (d, $J = 6.0\text{ Hz}$, 2H, Ph), 2.86 (d, $J = 6.6\text{ Hz}$, 2H, Ph), 2.27 (m, 2H, Ph). ^{13}C NMR (100 MHz, CDCl_3 , TMS): δ 195.79, 192.35, 154.58, 147.00, 143.26, 138.95, 135.88, 133.39, 131.69, 128.39, 127.51, 39.98, 29.57, 22.61.

Synthesis of [2-(ArN = CPh)-8-(NAr)-C₉H₈N]CoCl₂ (Co1–Co5)

(a) **Ar = 2,6-Me₂C₆H₃ Co1.** A suspension of 2-benzoyl-6,7-dihydro-5H-quinolin-8-one (76 mg, 0.30 mmol), 2,6-dimethylaniline (81 mg, 0.66 mmol) and $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (64 mg, 0.27 mmol, 0.9 eq.) in glacial acetic acid (20 mL) was stirred and heated to reflux for 6 h. Once cooled to room temperature, an excess of diethyl ether was added to precipitate the product. The precipitate was then filtered, washed with diethyl ether ($4 \times 15\text{ mL}$) and dried under reduced pressure to give **Co1** as a yellow powder (97 mg, 61%). FT-IR (cm^{-1}): 2954(w), 2917(m), 2869(w), 1622(s), 1566(s), 1469(s), 1446(s), 1381(m), 1322(m), 1267(vs), 1201(m), 1144(w), 1106(w), 981(m), 870(m), 779(s), 728(m), 700(vs). Anal. calculated for $\text{C}_{32}\text{H}_{31}\text{Cl}_2\text{CoN}_3$ (587.45): C, 65.43; H, 5.32; N, 7.15. Found: C, 65.67; H, 5.02; N, 6.99%.

(b) **Ar = 2,6-Et₂C₆H₃ Co2.** Using the same procedure and molar ratios as that described for **Co1** but with 2,6-diethylaniline as the arylamine, **Co2** was obtained as a yellow powder (84 mg, 46%). FT-IR (cm^{-1}): 2966(m), 1622(s), 1570(s), 1447(s), 1322(m), 1266(vs), 1187(m), 1109(m), 1062(m), 981(m), 780(s), 730(s), 699(vs). Anal. calculated for $\text{C}_{36}\text{H}_{39}\text{Cl}_2\text{CoN}_3$ (673.63): C, 67.75; H, 6.73; N, 6.24. Found: C, 67.45; H, 6.23; N, 6.33%.

(c) **Ar = 2,6-i-Pr₂C₆H₃ Co3.** Using the same procedure and molar ratios as that described for **Co1** but with 2,6-diisopropylaniline as the arylamine, **Co3** was obtained as a yellow powder (103 mg, 55%). FT-IR (cm^{-1}): 3067(w), 2965(m), 2920(w), 2873(m), 1619(m), 1563(s), 1449(s), 1324(s), 1267(s), 1185(s), 1055(s), 982(m), 802(s), 778(m), 730(m), 699(m). Anal. calculated for $\text{C}_{40}\text{H}_{47}\text{Cl}_2\text{CoN}_3$ (699.67): C, 68.67; H, 6.77; N, 6.01. Found: C, 68.47; H, 6.99; N, 5.77%.

(d) **Ar = 2,4,6-Me₃C₆H₂ Co4.** Using the same procedure and molar ratios as that described for **Co1** but with 2,4,6-trimethylaniline as the arylamine, **Co4** was obtained as a yellow powder (87 mg, 52%). FT-IR (cm^{-1}): 2972(w), 2908(m), 1626(s), 1592(s), 1477(s), 1447(s), 1382(m), 1321(m), 1267(vs), 1140(w), 1066(w), 982(m), 854(m), 775(s), 733(m), 699(vs). Anal. calculated for $\text{C}_{34}\text{H}_{35}\text{Cl}_2\text{CoN}_3$ (615.51): C, 66.35; H, 5.73; N, 6.83. Found: C, 66.45; H, 5.67; N, 6.81%.

(e) **Ar = 2,6-Et₂-4-MeC₆H₂ Co5.** Using the same procedure and molar ratios as that described for **Co1** but with 2,6-diethyl-4-methylaniline as the arylamine, **Co5** was obtained as a yellow powder (75 mg, 41%). FT-IR (cm^{-1}): 2967(m), 2928(m), 1622(s), 1576(s), 1451(s), 1321(m), 1266(vs), 1183(m), 1103(m), 1062(m), 981(m), 780(s), 698(vs). Anal. calculated for $\text{C}_{38}\text{H}_{43}\text{Cl}_2\text{CoN}_3$ (671.62): C, 67.96; H, 6.45; N, 6.26. Found: C, 68.04; H, 6.15; N, 6.09%.



General procedure for ethylene polymerization

A stainless autoclave (250 mL), equipped with an ethylene pressure control system, a mechanical stirrer and a temperature controller, was firstly evacuated and backfilled with ethylene three times. Once the desired temperature was reached the precatalyst (2.0 μmol), pre-dissolved in freshly distilled toluene (25 mL), was injected into the autoclave under an atmosphere of ethylene (*ca.* 1 atm). More toluene (25 mL) was injected followed by the required amount of co-catalyst (MAO, MMAO) and then a final portion of toluene (50 mL) to complete the addition. The autoclave was immediately pressurized to the desired pressure (5 or 10 atm) and the stirring commenced. Once the designated time (5, 10, 30, 45 and 60 min) was reached, the reactor was cooled to room temperature with an external ice/water bath. The autoclave was then vented and the reaction quenched by the addition of 10% HCl in ethanol (200 mL). The polyethylene was filtered, washed with ethanol, and then dried overnight under reduced pressure in an oven at 60 °C.

Crystal data and structure refinement

Single crystal X-ray diffraction studies were conducted on a XtaLAB Synergy-R single-crystal diffractometer with mirror-monochromatic Cu-K α radiation ($\lambda = 1.54184$ Å) for **Co1** at 170(10) K and **Co4** at 220(13) K; the cell parameters were obtained by global refinement of the positions of all collected reflections. Structure solution was performed by using the SHELXT program (Sheldrick, 2015)^{21a} and structure refinement was performed by using the SHELXL program (Sheldrick, 2015).^{21b} All hydrogen atoms were placed in calculated positions. Crystallographic parameters for **Co1** and **Co4** are summarized in Table S1.†

Conflicts of interest

There are no conflicts to declare.

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