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Ultrafast synthesis of near-zero-cost S-doped Ni(OH)₂ on C₃N₅ under ambient conditions with enhanced photocatalytic activity†

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Planting highly efficient and low-cost Ni-based noble-metal-free active sites on semiconductors is of great significance in the field of photocatalysis. Herein, taking wide visible-light-responsive 2D C₃N₅ as a model semiconductor, an impressive near-zero-cost 2D S-doped nickel hydroxide (S–Ni(OH)₂) is grown on C₃N₅ ultrafast within 30 min under ambient conditions by facile reaction between extremely low-cost Ni(NO₃)₂ and Na₂S in aqueous solution. The fabricated 2D S–Ni(OH)₂–C₃N₅ hybrid exhibits enhanced photocatalytic performance for both H₂ production from water and NO removal for air purification. The H₂ production rate on S–Ni(OH)₂–C₃N₅ is ~7 times higher than that of Ni(OH)₂–C₃N₅ and even slightly higher than that of Pt–C₃N₅, demonstrating its potential as a candidate for noble metal catalysts like Pt. In particular, an apparent quantum yield (AQY) value of 30.9% at 420 nm for H₂ production is reached on 1.0 wt% S–Ni(OH)₂–C₃N₅ due to quick internal charge transfer efficiency. In addition, ~42% of NO can be purified in a continuous flow reaction system. This work affords a cost-efficient strategy to steer the photocatalytic property of Ni-based catalysts.

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Introduction

In the past decades, the ever-increasing energy crisis and environmental pollution have drawn wide concern considering the sustainable development of human society and public health.^{1–3} Among the various clean technologies, the environmentally friendly semiconductor-based photocatalytic oxidation and reduction is deemed as the most potential one.^{4–7} In this field, photocatalytic solar H₂ production and NO removal lie in the core attributed to the huge energy consumption and the high risk respiratory diseases. Up to now, different semiconductors including metal oxides/sulfides (TiO₂, CdS, ZnIn₂S₄, *etc.*),^{8–12} metal-free materials (carbon nitride, COFs, conjugated polymers, *etc.*)^{13–17} and heterojunctions are all widely investigated to solve these issues.^{12,18–20} Of these strategies, constructing a heterojunction between different materials to combine the superiority of each one can significantly improve the whole photocatalytic efficiency. For example, K. Domen *et al.* selectively photodeposited Rh/Cr₂O₃ and CoOOH on different facets of the SrTiO₃:Al photocatalyst as cocatalysts for hydrogen and oxygen evolution respectively and an external quantum efficiency up to 96% at 350–360 nm was obtained.²¹

Recently, carbon nitride has drawn enormous attention due to its low-cost, high stability and suitable band position for water splitting;^{22–25} however, the band-gap of conventional carbon nitride (~2.7 eV for C₃N₄) is still too wide to obtain ideal photocatalytic activity. To alter the photocatalytic efficiency of carbon nitride, structural and band position regulation by doping (S, O, P, N- or C-rich modification) and fabricating heterojunctions (TiO₂/C₃N₄, CdS/C₃N₄, CNN/BDCNN, *etc.*) with type II or Z/S-scheme mechanisms are widely adopted.^{15,26–31} Recently, N-rich carbon nitride (*i.e.* C₃N₅) was reported;^{31–36} it possessed a narrower band-gap and higher catalytic activity than common C₃N₄, which represents a new direction for exploring carbon nitride based materials. For example, our previous work found that photocatalytic H₂ production activity over Pt–C₃N₅ was ~2.2 times higher than that of Pt–C₃N₄.³⁶ Nickel based materials exhibit excellent catalytic reduction and oxidation activities in the field of both photocatalysis and electrocatalysis,^{37–43} and thus can be used as co-catalysts on C₃N₅. Moreover, it has been reported that S doping can significantly improve the electrocatalytic performance of Ni/Fe (oxy)hydroxide materials;⁴⁴ thus doping of S in nickel-based co-catalysts may further improve their performance. However, these Ni-based species were commonly synthesized by hydrothermal or heat-treatment at high temperature, which was time and cost exhaustive. Therefore, construction of highly effective Ni-based catalysts in short time with low cost and energy consumption is of great importance for potential industrialization of photocatalytic technology.

Herein, taking 2D C₃N₅ as a model semiconductor, 2D S-doped Ni(OH)₂ (denoted as S–Ni(OH)₂) was facilely planted on its surface by a one-step precipitation method between Ni(NO₃)₂ and

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Na_2S in water (co-existence of S^{2-} and OH^-) under stirring within 30 minutes at room temperature, as revealed in Scheme 1. The hybrid photocatalyst ($\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$) was evaluated by photocatalytic activity for H_2 production and NO oxidation removal. Experimental results depicted that S-Ni(OH)_2 can promote the photogenerated e^-/ h^+ separation of C_3N_5 after light excitation and favor the generation of active oxygen species (ROS) participating in the subsequent photocatalytic procedure. In particular, the visible-light-induced photocatalytic H_2 production rate over $\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$ is even higher than that of $\text{Pt-C}_3\text{N}_5$, while the cost of S-Ni(OH)_2 (taking the H_2 production test as an example: in 45 mg 1.0 wt% $\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$) is calculated to be just ~ 0.0032 ¥ or ~ 0.00049 \$ based on the price of $\text{Ni(NO}_3)_2$ and Na_2S . The present work provides new insights into the design of low-cost and noble-metal-free catalysts for energy and environmental applications.

Experimental section

Material preparation

All chemicals used were purchased from Aladdin and used without purification. C_3N_5 was prepared by heating 3-amino-1,2,4-triazole at 500°C in air according to our previous work.³⁶ Generally, $\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$ was synthesized by mixing a certain amount of C_3N_5 , $\text{Ni(NO}_3)_2$ and Na_2S in water and stirring for 30 min. And then, the precipitate was washed with distilled water and absolute ethanol several times and dried in a vacuum at 80°C for 12 h. The loading amount of S-Ni(OH)_2 on C_3N_5 was calculated based on the initial input nickel mass fraction in the hybrid, and a series of $\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$ hybrids with S-Ni(OH)_2 loading amount varying between 0.5 and 3.0 wt% were prepared by changing the amount of $\text{Ni(NO}_3)_2$ and Na_2S accordingly. Pristine S-Ni(OH)_2 was also prepared by the same method in the absence of C_3N_5 . By the way, $\text{Ni(OH)}_2\text{-C}_3\text{N}_5$ was prepared by the same method using NaOH instead of Na_2S .

Catalyst characterization

XRD patterns were recorded on a powder X-ray diffractometer with $\text{Cu K}\alpha$ radiation (D8 Advance Bruker Inc., Germany). TEM and HRTEM images were acquired on an FEI TALOS F200. SEM and EDS mapping images were acquired on a Gemini Sigma 300. The Brunauer–Emmett–Teller (BET) specific surface area was evaluated using nitrogen adsorption–desorption apparatus (ASAP 2040,

Micrometrics Inc., USA) with all samples degassed at 120°C for 12 h prior to measurements. Valence state of each element in the catalyst was analyzed by X-ray photoelectron spectroscopy (XPS) on a Thermo Fisher ESCALAB 250Xi X-ray photoelectron spectroscope. A Shimadzu UV-3100 recording spectrophotometer and fluorescence spectrometer (F-4600, Hitachi Inc., Japan) were used to record UV-vis diffuse reflectance (DRS) and PL spectra measurements respectively. Time-resolved fluorescence spectra (TRFS) were obtained on an Edinburgh FLS 1000. Electron paramagnetic resonance (EPR) spectra were recorded on a JES-FA200 EPR spectrometer. 5,5-Dimethyl-1-pyrroline-*N*-oxide (DMPO) was used as a radical spin-trapping reagent for $\cdot\text{OH}$ and $\cdot\text{O}_2^-$. 2,2,6,6-Tetramethylpiperidine (TEMP) was used as the trapping agent for $^1\text{O}_2$.³ The photocurrent and electrochemical impedance measurements were recorded on a CHI 660D electrochemical workstation (Chenhua Instrument, Shanghai, China) with a conventional three electrode configuration. Pt foil and Ag/AgCl (saturated KCl) were used as the counter electrode and reference electrode respectively. The working electrode was made by spreading a catalyst/Nafion slurry on FTO glass and 0.1 M Na_2SO_4 aqueous solution was used as the electrolyte.

Photocatalytic activity tests and photoelectrochemical measurements

The photocatalytic H_2 production rate and AQY measurements were carried out on a top-irradiated reaction vessel containing the catalyst and sacrificial reagent (TEOA) connected to a closed gas system with a 300 W Xe-lamp (PLS SXE300, Beijing Perfectlight Inc., China) equipped with a cutoff filter ($\lambda > 420$ nm) or band-pass filters ($\lambda = 420, 500$ nm, etc.); the photocatalytic H_2 production system is shown in Fig. S1.† The H_2 production rate was detected using a GC (SP7820, TCD detector, 5 \AA molecular sieve columns, and Ar carrier), and AQY values were calculated according to our previous work:¹⁷

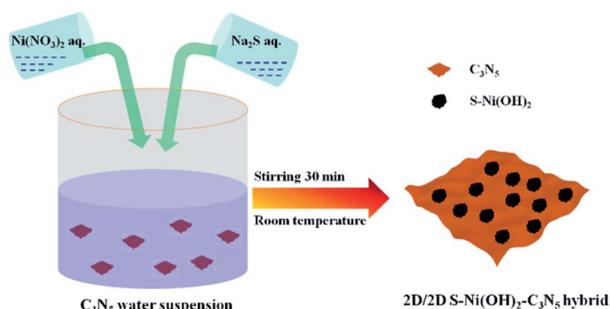
$$\text{AQY} = \frac{2 \times (\text{number of evolved } \text{H}_2 \text{ molecules})}{\text{number of incident photons}} \quad (1)$$

For photocatalytic NO removal, a continuous flow reactor under ambient conditions was adopted according to our previous work.⁷ The volume of the rectangular reactor was 4.5 L ($30\text{ cm} \times 15\text{ cm} \times 10\text{ cm}$). 25 mg catalyst was dispersed into the mixture of ethanol and water and then transferred into a culture dish with a diameter of 12 cm, and then the dish was placed in the reactor after the solvent was dried at 45°C . A 30 W visible LED (General Electric) was used as a light source. The gas (containing NO and air) flow rate through the reactor was controlled at 1000 mL min^{-1} using a mass flow controller with an initial NO concentration of ~ 600 ppb. The NO and NO_2 concentrations were recorded on a NO_x analyzer model T200 (Teledyne API). The generation of NO_3^- was detected by ion chromatography on a Thermo DIONEX ICS-900.

Results and discussion

Crystal phase and micro morphology analyses

The pristine S-Ni(OH)_2 prepared from the ultrafast reaction between Ni^{2+} and Na_2S in water under ambient conditions



Scheme 1 Illustration of ultrafast preparation of the $\text{S-Ni(OH)}_2\text{-C}_3\text{N}_5$ hybrid.



within 30 min is firstly detected by XRD, as depicted in Fig. S2† which demonstrates the formation of S-Ni(OH)₂ with low diffraction peaks corresponding to the PDF card of 73-1520 for Ni(OH)₂, indicating the low crystallinity of the present S-Ni(OH)₂ material. Herein, the confirmation of S doping will be illustrated in the following section. It should be noted that pristine Ni(OH)₂ without S doping is prepared by reaction between Ni²⁺ and NaOH for comparison, and the diffraction peaks of this Ni(OH)₂ are also in agreement with the same PDF card (Fig. S2†), indicating that the comparison is reasonable. For the XRD patterns of C₃N₅ and S-Ni(OH)₂-C₃N₅ hybrid (Fig. 1a), only two distinct peaks corresponding to (100) and (002) planes of C₃N₅,^{32,35,36,45} are observed even with a high loading amount of S-Ni(OH)₂ up to 3.0 wt%, which should relate to the low crystallinity of present S-Ni(OH)₂. In addition, no obvious shift happens in the main peaks of C₃N₅ after the loading of S-Ni(OH)₂, revealing that the procedure for hybrid preparation does not destroy the basic structure of C₃N₅.

The element component of 1.0 wt% S-Ni(OH)₂-C₃N₅ is analyzed by elemental analysis and ICP-MS (Table S1†). From the elemental analysis result, the C/N ratio of the catalysts was determined to be 3 : 4.7, which is much lower than that of C₃N₄ and very close to that of C₃N₅, and the result is in good agreement with previous literature.³⁵ As revealed in Fig. 1b, S-Ni(OH)₂-C₃N₅ is composed of C, N, Ni, O and S elements, demonstrating the existence of S-doping in Ni(OH)₂ during the reaction between Ni²⁺ and Na₂S preliminarily. In addition, the result depicts the successful combination of S-Ni(OH)₂ and C₃N₅. The high resolution spectra of C 1s (Fig. S3a†) and N 1s (Fig. S3b†) in S-Ni(OH)₂-C₃N₅ are similar to that of pristine C₃N₅ shown in our previous work,³⁶ and no obvious peak shifts are found, revealing that the structure of C₃N₅ is well kept during the loading of S-Ni(OH)₂. The existence of Ni ions is confirmed by the high resolution spectrum of Ni 2p, in which two distinct spin-orbit peaks at ~855.8 (Ni 2p_{3/2}) and ~873.5 eV (Ni 2p_{1/2}) along with two satellite peaks (identified as "Sat.") are clearly observed (Fig. 1c).^{40,42} The O 1s spectra mainly consist of

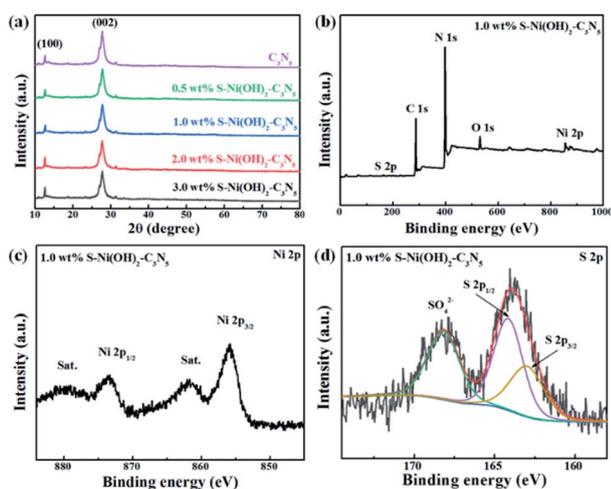


Fig. 1 (a) XRD patterns of C₃N₅ and S-Ni(OH)₂-C₃N₅ with different loading amounts of S-Ni(OH)₂. (b) XPS survey, (c) Ni 2p and (d) S 2p spectra of 1.0 wt% S-Ni(OH)₂-C₃N₅ respectively.

a peak at 531.79 eV, (see Fig. S3c†) which can be assigned to the O²⁻ species in Ni(OH)₂.⁴⁶ Most importantly, two peaks are observed in the high resolution spectrum of S 2p, in which the peak centered at ~164 eV represents the metal-sulfur bond and S²⁻, while the other peak in the high binding energy region can be considered as residual SO₄²⁻ adsorbed on the material (Fig. 1d).^{40,42,47} That is, the existence of the Ni-S bond can support the formation of S-doping in Ni(OH)₂.

To further confirm the successful fabrication of the S-Ni(OH)₂-C₃N₅ hybrid, TEM and SEM images are acquired and shown in Fig. 2. It has been reported in our previous work that the micro morphology of pristine C₃N₅ is nanosheet.³⁶ Fig. 2a and b display the TEM and HRTEM images of S-Ni(OH)₂. As shown in Fig. 2a, S-Ni(OH)₂ also exhibits a micro morphology of nanosheets and a clear lattice fringe of 0.23 nm corresponding to the (011) plane of S-Ni(OH)₂ is observed (Fig. 2b). In particular, no lattice fringe representing nickel sulfide is found, demonstrating the existence state of S-doping in S-Ni(OH)₂. Fig. 2c displays the TEM image of 1.0 wt% S-Ni(OH)₂-C₃N₅, in which 2d nanosheets containing S-Ni(OH)₂ and C₃N₅ are observed. In addition, S-Ni(OH)₂ and C₃N₅ come into tight contact with each other and it is difficult to distinguish each material due to their similar micro morphology. The high resolution region (Fig. 2d) of S-Ni(OH)₂-C₃N₅ can be clearly divided into an amorphous region corresponding to C₃N₅ and a crystalline region for S-Ni(OH)₂ with a clear lattice fringe of 0.23 nm corresponding to the (011) plane of S-Ni(OH)₂. By the way, the two different regions in Fig. 2d come into tight contact with each other, which will be favorable for the internal transfer of photogenerated carriers.

Fig. 2e depicts the SEM and element mapping images of 1.0 wt% S-Ni(OH)₂-C₃N₅. As revealed, the hybrid is basically composed of nanosheets containing S-Ni(OH)₂ and C₃N₅, which is in good agreement with the above results of TEM. The element mapping images convey that the hybrid contains C, N, Ni, O and S, and all of the elements disperse uniformly in the observed region. Moreover, the existence of S-doping in S-Ni(OH)₂ is confirmed from the mapping of S in the material. Based on the above-mentioned XRD, XPS, TEM and SEM analyses, the product of reaction between Ni²⁺ and Na₂S is named S-Ni(OH)₂ considering that no any information corresponding to nickel sulfide is detected

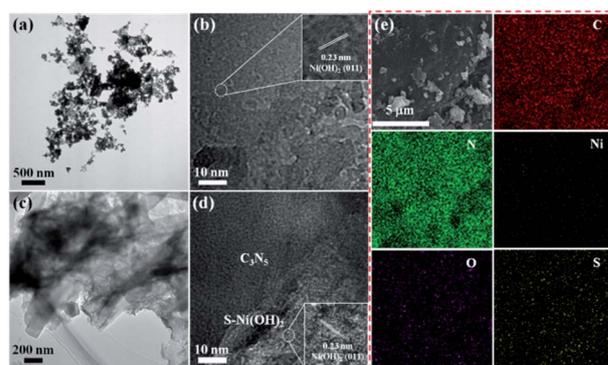


Fig. 2 TEM and HRTEM of S-Ni(OH)₂ (a and b) and 1.0 wt% S-Ni(OH)₂-C₃N₅ (c and d). SEM and EDS mapping of 1.0 wt% S-Ni(OH)₂-C₃N₅ (e).



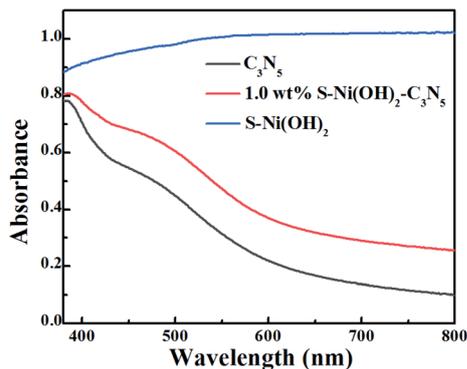


Fig. 3 DRS spectra of S-Ni(OH)₂, C₃N₅ and 1.0 wt% S-Ni(OH)₂-C₃N₅.

while the Ni-S bond exists in the hybrid, and the hybrid of 2D S-Ni(OH)₂-C₃N₅ is successfully constructed.

DRS and e⁻/h⁺ separation analyses

In the field of photocatalysis, the light responsive region plays a significant role in the performance of catalysts. The visible light absorption properties of S-Ni(OH)₂, C₃N₅ and S-Ni(OH)₂-C₃N₅ are compared in Fig. 3. In brief, C₃N₅ possesses a narrow band-gap of ~2.25 eV according to our previous work,³⁶ and can absorb the visible light region of 400–600 nm with high intensity, while pristine S-Ni(OH)₂ exhibits a strong absorption across the spectral range of 200–2400 nm (Fig. S6[†]). For the hybrid material, 1.0 wt% S-Ni(OH)₂-C₃N₅ exhibits a similar visible light absorption region to C₃N₅, but the absorption intensity is slightly increased by the hybridization of black S-Ni(OH)₂. In a word, this wide and strong visible light absorption property of the hybrid is favorable for achieving high photocatalytic performance.

Light absorption is the first step for a photocatalytic procedure, and the subsequent photogenerated e⁻/h⁺ separation is of great importance to the whole performance. Monitoring of the PL behavior of photocatalysts can find out some information on

e⁻/h⁺ separation after light excitation.^{26,48,49} As revealed in the steady-state PL spectrum in Fig. 4a, C₃N₅ exhibits strong PL intensity, demonstrating its high photogenerated e⁻/h⁺ recombination rate after light excitation. However, the emission intensity of C₃N₅ can be efficiently quenched by S-Ni(OH)₂. The lower PL intensity of 1.0 wt% S-Ni(OH)₂-C₃N₅ than C₃N₅ demonstrates that the recombination of e⁻/h⁺ pairs on C₃N₅ after excitation is restrained, which should be caused by the electron transfer from C₃N₅ to S-Ni(OH)₂ since Ni-based species such as nickel sulfide and Ni(OH)₂ are often used as co-catalysts for photocatalytic H₂ production.^{43,50–52} Mott-Schottky analysis was further conducted on S-Ni(OH)₂ to verify whether the CB position of S-Ni(OH)₂ is sufficient to capture the photo-generated electrons from C₃N₅ (Fig. S7[†]). The CB position of S-Ni(OH)₂ was determined to be -0.58 V vs. NHE, which is less negative than that of C₃N₅ (-0.98 V vs. NHE).³⁶ So, under irradiation, the photogenerated electrons of C₃N₅ could migrate to S-Ni(OH)₂ and be captured. The time-resolved fluorescence spectra (TRFS) of C₃N₅ and 1.0 wt% S-Ni(OH)₂-C₃N₅ are compared in Fig. 4b to further illustrate this viewpoint. As depicted, the PL intensity of C₃N₅ decays quickly after excitation, indicating its low fluorescence lifetime which is not favorable for photogenerated e⁻/h⁺ separation. However, the decay tendency of 1.0 wt% S-Ni(OH)₂-C₃N₅ exhibits a clear slowdown under the same conditions, demonstrating that the fluorescence lifetime is prolonged, which should be related to the internal electron transfer from C₃N₅ to S-Ni(OH)₂.^{48,53} For comparison, the fluorescence lifetime of C₃N₅ and 1.0 wt% S-Ni(OH)₂-C₃N₅ is calculated to be ~8.50 and 14.02 ns respectively via a biexponential fitting. The enhanced e⁻/h⁺ separation efficiency of 1.0 wt% S-Ni(OH)₂-C₃N₅ compared to pristine C₃N₅ is further confirmed by photocurrent-time and EIS behaviors. As revealed in Fig. 4c, sharp photocurrent is generated under light irradiation and then decreases suddenly when light is turned off, demonstrating that the current is light-switched and produced by the photogenerated e⁻/h⁺ separation after the semiconductor is excited.^{20,54,55} In comparison, 1.0 wt% S-Ni(OH)₂-C₃N₅ exhibits a much higher photocurrent value than pristine C₃N₅, indicating that the separation of photogenerated e⁻/h⁺ on C₃N₅ is facilitated by loading of the S-Ni(OH)₂ co-catalyst. In addition, a smaller arc radius is recorded on the S-Ni(OH)₂-C₃N₅ electrode, revealing the smaller resistance for charge transfer which is favorable for e⁻/h⁺ separation.^{53,56} All these results convey that the S-Ni(OH)₂-C₃N₅ hybrid possesses quick internal charge transfer efficiency, which will facilitate the subsequent photocatalytic performance.

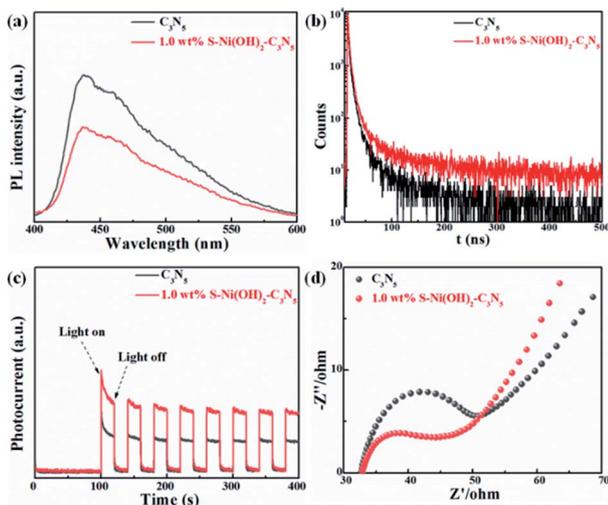


Fig. 4 Comparison of steady-state PL (a), time-resolved PL (b), photocurrent-time (c) and EIS (d) curves of C₃N₅ and 1.0 wt% S-Ni(OH)₂-C₃N₅.

Photocatalytic H₂ production performance

In consideration of the wide visible light absorption region and quick photogenerated e⁻/h⁺ separation efficiency of the S-Ni(OH)₂-C₃N₅ hybrid, its photocatalytic performance for H₂ production from water with TEOA as a sacrificial reagent under visible light irradiation ($\lambda > 420$ nm) is evaluated firstly. A blank experiment shows that almost no H₂ production is observed on pristine C₃N₅ due to the lack of a co-catalyst. And pristine S-Ni(OH)₂ also cannot produce H₂ under these conditions.



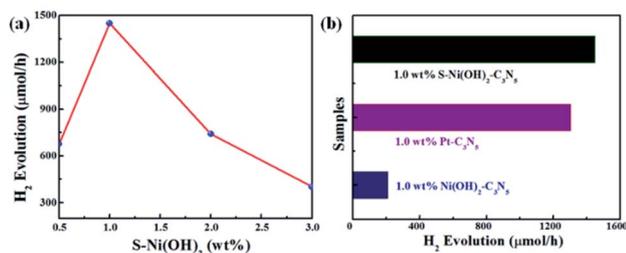


Fig. 5 (a) Influence of S-Ni(OH)₂ loading amount on the H₂ production activity of the S-Ni(OH)₂-C₃N₅ hybrid. (b) Comparison of H₂ production activity over 1.0 wt% S-Ni(OH)₂-C₃N₅, 1.0 wt% Ni(OH)₂-C₃N₅ and 1.0 wt% Pt-C₃N₅ catalysts. Conditions: 45 mg catalyst, 10 vol% TEOA/water aqueous, $\lambda > 420$ nm.

However, as displayed in Fig. 5a, the H₂ production rate is dramatically enhanced even with a low loading amount of 0.5 wt% S-Ni(OH)₂ on C₃N₅. Generally, a volcano-type relationship between the loading amount of a cocatalyst and the whole photocatalytic activity will be obtained,⁵⁷ suggesting that the fraction of S-Ni(OH)₂ in the hybrid must be screened to explore the highest H₂ production rate. As shown in Fig. 5b, the maximum H₂ production rate is 1450 $\mu\text{mol h}^{-1}$ on the catalyst of 1.0 wt% S-Ni(OH)₂-C₃N₅. Notably, excessive loading of S-Ni(OH)₂ will influence the visible light absorption of C₃N₅, and then the whole photoactivity is decreased. In addition, the initial concentration of Na₂S in the procedure of catalyst preparation will influence the concentration of OH⁻ and S²⁻, which participate in the formation of S-Ni(OH)₂. The influence of initial mol ratio between Ni²⁺ and Na₂S (Ni/S) in the preparation procedure on the photoactivity is shown in Fig. S4,[†] and the selected condition for S-Ni(OH)₂ preparation is Ni/S = 1/3.

To explore the influence of S doping on the photocatalytic H₂ production performance, the comparison of 1.0 wt% S-Ni(OH)₂-C₃N₅ and 1.0 wt% Ni(OH)₂-C₃N₅ is shown in Fig. 5b. Unexpectedly, 1.0 wt% Ni(OH)₂-C₃N₅ exhibits much lower H₂ production activity (209 $\mu\text{mol h}^{-1}$), which is only one seventh of the value over the 1.0 wt% S-Ni(OH)₂-C₃N₅ hybrid. That is to say, a novel method is explored to prepare a highly efficient Ni(OH)₂-based co-catalyst by simply changing the initial reactant from NaOH to Na₂S that reacts with Ni²⁺ under ambient conditions ultra-fast. The photocatalytic performance is also evaluated using Pt nanoparticles as a reference since Pt is generally considered as the most efficient co-catalyst for photocatalytic H₂ production.^{15,23} As depicted in Fig. 5b, the H₂ production rate on 1.0 wt% S-Ni(OH)₂-C₃N₅ is slightly higher than that of 1.0 wt% Pt-C₃N₅, demonstrating that the present S-Ni(OH)₂ is a highly efficient co-catalyst for H₂ production and is promising in consideration of the facile preparation procedure and low cost of Ni-based materials compared with noble metals like Pt.

The AQY values of 1.0 wt% S-Ni(OH)₂-C₃N₅ under a series of monochromatic light irradiation are shown in Fig. 6a, in which the AQY values decrease gradually with the increase of light wavelength and the highest AQY value (30.9%) is obtained at 420 nm. The change tendency of AQY is in agreement with the DRS spectrum of S-Ni(OH)₂-C₃N₅, demonstrating that the

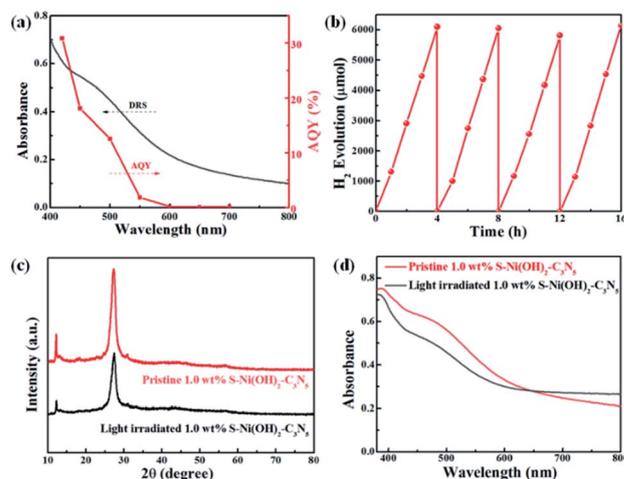


Fig. 6 (a) Changing tendency of AQY values of 1.0 wt% S-Ni(OH)₂-C₃N₅ with monochromatic light irradiation. (b) Cycle performance of 1.0 wt% S-Ni(OH)₂-C₃N₅ for H₂ production under visible light irradiation. Comparison of XRD (c) and DRS (d) spectra of 1.0 wt% S-Ni(OH)₂-C₃N₅ before and after photocatalytic performance.

photocatalytic performance of S-Ni(OH)₂-C₃N₅ is determined by its light absorption behavior. Long-term stability of a photocatalyst is an important norm in the field of H₂ production. Fig. 6b reveals the cycling performance of 1.0 wt% S-Ni(OH)₂-C₃N₅ and no obvious attenuation of H₂ production rate happens after 16 h and 4 runs of photocatalytic test, demonstrating the excellent photocatalytic stability of the present S-Ni(OH)₂-C₃N₅ hybrid. To further evaluate the stability of S-Ni(OH)₂-C₃N₅ in photocatalysis, the samples are collected after long-term light irradiation and monitored by XRD and DRS. As depicted in Fig. 6c and d, no obvious change happens in the XRD peaks of S-Ni(OH)₂-C₃N₅ experiencing the photocatalytic test, indicating that the basic structure of the hybrid is maintained. Meanwhile, the collected sample exhibits a similar visible light absorption region compared with fresh S-Ni(OH)₂-C₃N₅, demonstrating the stability of the hybrid again.

Photocatalytic NO oxidation performance

Except for H₂ production, the photocatalytic performance of S-Ni(OH)₂-C₃N₅ is also tested for NO oxidation to demonstrate the potential of the catalyst in purification of atmospheric pollution. A control experiment shows that pristine S-Ni(OH)₂ has no activity for NO removal. As revealed in Fig. 7a, the oxidation of NO can reach equilibrium in 25 min in the present continuous flow reaction system. Briefly, NO removal efficiencies on C₃N₅ and S-Ni(OH)₂-C₃N₅ with 0.5, 1.0, 2.0, and 3.0 wt% S-Ni(OH)₂ fraction after 25 min of light irradiation are 35.0, 37.0, 42.0, 35.0 and 35.0% respectively and the most efficient catalyst is 1.0 wt% S-Ni(OH)₂-C₃N₅ hybrid, indicating that the photocatalytic performance of C₃N₅ is enhanced by loading of S-Ni(OH)₂. And the above-mentioned quick internal charge transfer from C₃N₅ to S-Ni(OH)₂ under light excitation should play a significant role in this process. Fig. 7b evaluates the durability and stability of the 1.0 wt% S-Ni(OH)₂-C₃N₅ hybrid



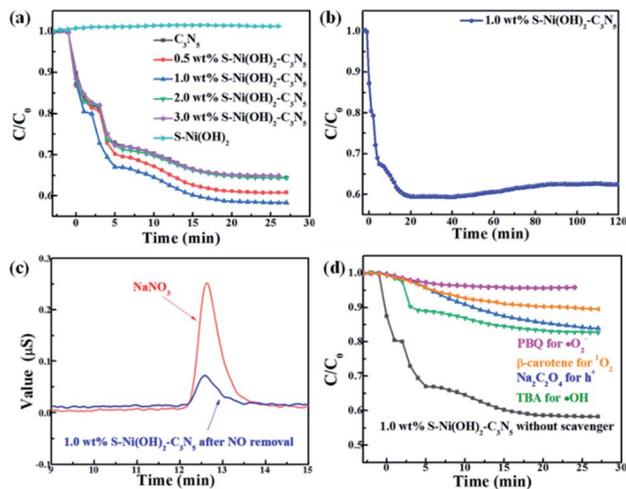


Fig. 7 (a) Influence of S-Ni(OH)₂ loading amount on NO removal ratio over the S-Ni(OH)₂-C₃N₅ hybrid. (b) Durability of 1.0 wt% S-Ni(OH)₂-C₃N₅ for NO removal. (c) Detection of NO₃⁻ by ion chromatography. (d) Photocatalytic NO removal performance of 1.0 wt% S-Ni(OH)₂-C₃N₅ in the presence of a series of trapping agents.

for NO removal, and only a slight decrease in NO removal ratio is observed during 120 nm light irradiation, demonstrating the good stability of the hybrid. Generally, the main products of photocatalytic NO oxidation are NO₂ and NO₃⁻ according to literature reports.^{58–60} The detection of NO₂ is carried out on the present NO_x analyzer, and the concentration of NO₂ along with NO oxidation is shown in Fig. S5.† It is found that most NO is converted to NO₃⁻ based on the amount of NO removal and NO₂ production. Therefore, the confirmation of produced NO₃⁻ in the reaction is shown in Fig. 7c with NaNO₃ as a reference by ion chromatography, in which a clear signal corresponding to NO₃⁻ is observed for the sample prepared with 1.0 wt% S-Ni(OH)₂-C₃N₅ experiencing photocatalytic NO oxidation reaction.

It has been reported that activation of molecular O₂ and generation of a series of ROS (*i.e.* [•]OH, [•]O₂⁻ and ¹O₂) play the main role in the procedure of photocatalytic NO oxidation.^{61–63} In addition, photogenerated h⁺ residual in the VB of the semiconductor is also of great importance.^{64,65} Thus, the photocatalytic NO removal performance of 1.0 wt% S-Ni(OH)₂-C₃N₅ in the presence of a series of trapping agents for different active species is compared in Fig. 7d. As displayed, when Na₂C₂O₄, *tert*-butanol (TBA), *p*-benzoquinone (PBQ) and β-carotene are used as scavengers for h⁺, [•]OH, [•]O₂⁻ and ¹O₂ respectively, the photocatalytic performance of 1.0 wt% S-Ni(OH)₂-C₃N₅ is seriously restrained, demonstrating that NO is removed by these ROS and residual h⁺ in the VB of C₃N₅. The NO₂ concentration during the photocatalytic NO removal on 1.0 wt% S-Ni(OH)₂-C₃N₅ was also monitored with the NO_x analyzer (see Fig. S8†). From these results, it can be seen that the [•]OH and ¹O₂ play important roles in the conversion of NO₂ to NO₃⁻. Electron spin-resonance spectroscopy (ESR) tests are conducted to monitor the change of signals corresponding to [•]OH, [•]O₂⁻ and ¹O₂. As revealed in Fig. 8, almost no signals are detected in

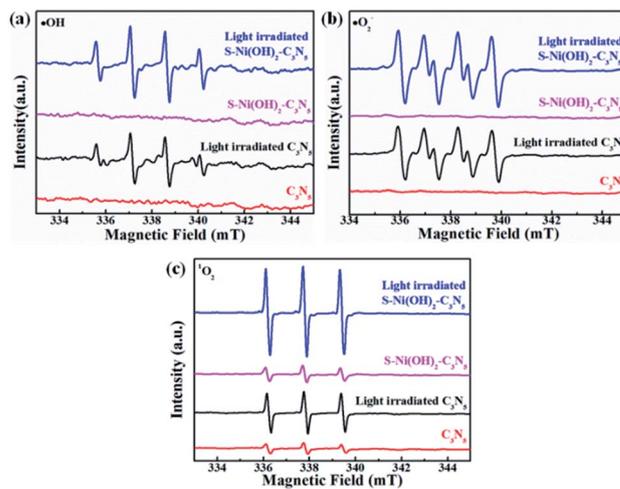
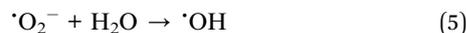


Fig. 8 ESR spectra of DMPO-[•]OH (a), DMPO-[•]O₂⁻ (b) and TEMPO (c) for ROS detection.

the dark on both 1.0 wt% S-Ni(OH)₂-C₃N₅ and C₃N₅, but strong peaks corresponding to distinct DMPO-[•]OH, DMPO-[•]O₂⁻ and TEMPO (product of TEMP oxidized by ¹O₂) signals are detected,^{3,7,64} demonstrating that the generation of these ROS is light controllable. Moreover, the light induced ESR signal intensity on 1.0 wt% S-Ni(OH)₂-C₃N₅ is higher than that of C₃N₅, revealing that more amount of ROS is generated on S-Ni(OH)₂-C₃N₅ due to the quick internal charge transfer. So, the process of photocatalytic NO removal on 1.0 wt% S-Ni(OH)₂-C₃N₅ should be as follows:



Based on all of the above results, the mechanism for photocatalytic H₂ production or NO oxidation over the S-Ni(OH)₂-C₃N₅ hybrid is proposed. Under visible light irradiation, C₃N₅ is excited and generates e⁻/h⁺ pairs. S-Ni(OH)₂ can promote the internal charge separation and harvest the e⁻ on the CB of C₃N₅, and then the trapped electrons participate in subsequent surface reactions. For photocatalytic H₂ production (Fig. 9a), the collected electrons on S-Ni(OH)₂ react with adsorbed H₂O molecules and H₂ is produced. The residual h⁺ on the VB of C₃N₅ is consumed by TEOA, and then the whole photocatalytic H₂ production procedure is accomplished. However, the situation for photocatalytic NO oxidation is different (Fig. 9b). O₂ is firstly activated and a series of ROS ([•]O₂⁻, ¹O₂ and [•]OH) can be generated. Generally, [•]O₂⁻ is produced by reaction between e⁻



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