Electrochemical properties of \( \text{La}_{0.5}\text{Sr}_{0.5}\text{Fe}_{0.95}\text{Mo}_{0.05}\text{O}_{3-\delta} \) as cathode materials for IT-SOEC

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Solid oxide electrolysis cells (SOECs) are a new type of high-efficiency energy conversion device that can electrolyze CO\(_2\) efficiently and convert electricity into chemical energy. However, the lack of efficient and stable cathodes hinders the practical application of CO\(_2\) electrolysis in SOECs. Herein, a novel perovskite oxide \( \text{La}_{0.5}\text{Sr}_{0.5}\text{Fe}_{0.95}\text{Mo}_{0.05}\text{O}_{3-\delta} \) (LSFMo) is synthesized and used as a cathode for SOECs. The introduction of Mo significantly improves the CO\(_2\) tolerance of the cathode material and solves the problem of SrCO\(_3\) generation in the \( \text{La}_{0.5}\text{Sr}_{0.5}\text{Fe}_{0.95}\text{O}_{3-\delta} \) material. Mo ion doping promotes the conductivity in a reducing atmosphere and increases the oxygen deficiencies of the material, which lowers the ohmic resistance (\( R_\text{p} \)) of the material and significantly improves the CO\(_2\) adsorption and dissociation in the middle-frequency of polarization resistance (\( R_\text{p} \)). For example, \( R_\text{p} \) decreases from 0.49 to 0.24 \( \Omega \) cm\(^2\) at 800 °C under 1.2 V. Further, the reduction of \( R_\text{f} \) and \( R_\text{p} \) increases the performance improvement, and the current density is increased from 1.56 to 2.13 A cm\(^{-2}\) at 800 °C under 2 V. Furthermore, LSFMo shows reasonable short-term stability during the 60 h stability test.

Moreover, the problems of particle agglomeration and carbon deposition also lead to the decline of electrolytic performance.\(^{7-12}\) Recently, perovskite materials have attracted wide attention due to excellent redox stability, flexible composition, sulfur resistance and carbon resistance. The perovskite materials could meet the needs of complex working conditions by selecting different elements at the A and B sites, respectively. The SOECs cathode are exposed in reducing environment, so the materials are basically selected from anode materials of solid oxide fuel cells, including Ti-, Fe- and Cr-based perovskite oxides,\(^{13,14}\) such as \( \text{La}_{0.75}\text{Sr}_{0.25}\text{Cr}_{0.5}\text{Mn}_{0.5}\text{O}_{3-\delta}, \text{La}_{0.75}\text{Sr}_{0.25}\text{Sr}_{0.5}\text{Fe}_{1.5}\text{O}_{6-\delta}, \text{La}_{0.8}\text{Sr}_{0.2}\text{Fe}_{1.5}\text{O}_{6-\delta} \), and \( \text{Sr}_{1-x}\text{Fe}_{x}\text{Mn}_{1-x}\text{O}_{3-\delta} \).\(^{15-17}\)

Compared with Cr-based and Mn-based materials, Fe-based materials exhibit stronger catalytic activity due to their high oxygen vacancies and ionic conductivity, suggesting they are potential substitutes for Ni-YSZ.\(^{14,16}\) However, the poor stability in reducing atmosphere leads to the phase decomposition and SrCO\(_3\) generation when Fe-based materials are used as fuel electrode of solid oxide cells, which affect the cell performance and operation stability. The structure of perovskite could be stabilized by doping high-valent elements at the B site, which makes the Fe-based perovskite keep good reduction stability and CO\(_2\) tolerance so that it meets the requirements of SOECs cathode.\(^{17-20}\)

In this work, a small quantity of Mo is introduced into the B site of \( \text{La}_{0.5}\text{Sr}_{0.5}\text{Fe}_{0.95}\text{O}_{3-\delta} \) (LSF) to partly substitute Fe, and the \( \text{La}_{0.5}\text{Sr}_{0.5}\text{Fe}_{0.95}\text{Mo}_{0.05}\text{O}_{3-\delta} \) (LSFMo) material is synthesized by a solid-phase method. The crystal structure and electrical conductivity of LSFMo material in oxidizing and reducing
atmosphere are tested and discussed by X-ray photoelectron spectroscopy (XPS) analysis. The electrochemical performance and stability of cells supported by La$_{0.8}$Sr$_{0.2}$Ga$_{0.83}$Mg$_{0.17}$O$_3$–δ (LSGM) electrolyte are studied.

**Experimental**

**Power synthesis**

The LSF and LSFMo materials were synthesized by the solid-state reaction method with La$_2$O$_3$, SrCO$_3$, Fe$_2$O$_3$, and MoO$_3$ as the raw materials. Each of powders was weighed with stoichiometric amounts and mixed by ball-milling with ethanol for 10 h. After drying at 80 °C for 2 h, the mixture was pressed into pellets and roasted at 1250 °C for 10 h. Finally, the samples were crushed and ground into powder.

The LSGM powder was synthesized by the same technological process. The dense electrolyte pellets were obtained by pressing powder and sintering at 1450 °C in air for 10 h.

**Sample preparation**

In the case of the conductivity test, the LSF and LSFMo powders were mixed with polyvinyl alcohol (PVA) binder and pressed into rectangular samples. The dense bars were obtained by sintered at 1400 °C in air for 6 h. During the preparation of the electrolysis cells, the cathode slurries were made by mixing the powders with isopropyl alcohol, ethylene glycol and glycerin (10 : 2 : 1) for 1 h. Similarly, the La$_{0.6}$Sr$_{0.4}$CoO$_{2.57}$Fe$_{0.5}$Ga$_{0.5}$O$_{3}$–δ (LSCF) and Gd$_{0.2}$Ce$_{0.8}$O$_2$ (GDC) were mixed with 1 : 1 ratio to form anode slurry. The LSGM pellets with 300 μm thickness were used as the supporter, and the LSF(LSFMo) and LSCF-GDC slurries were sprayed on the two sides of electrolyte. The template with 0.26 cm$^2$ area was used to keep the effective area of the electrode. The sprayed cells (LSF(LSFMo)|LSGM|LSCF-GDC) were calcined at 1000 °C for 2 h to make the electrodes appear loose and porous structure. The Ag paste was drawn on both sides of the electrode as collector to test performance.

**Characterization and test**

The phase structure of the materials was characterized by X-ray diffraction (XRD) in the range of 20–80°. The LSF and LSFMo samples were placed in CO$_2$ : CO = 1 : 1 at 800 °C for 20 h to study CO$_2$ tolerance, and the treated powders was analyzed by XRD. In conductivity test, the Ag wire was fixed in the end of the sample as a current electrode, and the Ag wire in the middle of the sample is used as voltage electrode. The conductivity of the electrode bars was tested via a four-terminal method from 850 °C to 500 °C using Keithley 2460 Source Meter. The cooling rate was 5 °C min$^{-1}$ and the temperature was stable for 30 min before test. In the reducing atmosphere, the flow rate of CO$_2$ and CO were both 50 mL min$^{-1}$. The valence state change of the elements in the electrode surface at oxidation state and reduction state were tested by XPS meth. The data of each element were fitted using Xpspeak41 software.

The electrolytic cells were sealed on Al$_2$O$_3$ tubes and connected to PARSTAT2273 advanced electrochemical workstation for electrochemical test. The cathode was sealed inside the tube and fed with pure CO$_2$. The anode is directly placed in air. The electrochemical performances were performed when open circuit voltage of cells was steady. The $I$–$V$ curves from 0 V to 2 V were recorded via the linear voltage sweep. The electrochemical impedance spectra (EIS) at different voltages were measured with a frequency range of 100 kHz–10 mHz, and the data were treated by Zview to explain each step of the electrochemistry process under different voltages. The electrolytic stability under 1.2 V was measured by Keithley 2460 Source Meter at 800 °C. The micromorphology of the electrode and electrolyte were viewed by scanning electron microscope.

**Results and discussion**

**Crystal structure and chemical compatibility**

Fig. 1(a) and (b) display the XRD patterns of LSF and LSFMo powders after calcinating at 1250 °C for 10 h in air. A pure rhombohedral perovskite structure is formed after Mo element doping without any impurities, suggesting that 5% Mo could completely enter the LSF lattice. The main peaks of 30°–35° are magnified to further analyze the change in the two perovskite materials. After Mo is doped into the LSF structure, the peak is moved toward a small angle (from 32.6° to 32.5°), indicating an expansion of the lattice caused by charge compensation. Due to the addition of high-valence Mo$^{5+/6+}$, the valence of Fe is...
reduced to compensate for charge deficiency. Although the Mo5+/6+ radius is smaller than that of Fe3+ and Fe2+, the Fe ionic radius gradually increases with the reduction of Fe ionic to the lower valence state, and the radius of Mo5+/6+ is also larger than that of Fe4+, which lead to an increase in the unit volume. This phenomenon could be verified by XPS results of the Fe element in the following section. This phenomenon could be verified by XPS of Fe element results in the following section. (ionic radius: Fe3+ = 58.5 pm, Fe2+ = 64.5 pm, Fe4+ = 78 pm, Mo5+ = 59 pm and Mo6+ = 61 pm).

The stability of SOEC cathode materials is further investigated, and LSF and LSFMo materials were explored in CO2 : CO = 1 : 1 atmosphere at 800 °C for 20 h treatment. The XRD results of the treated powders are shown in Fig. 1(c) and (d). The SrCO3 phase and little Fe metal are generated after LSFMo is treated in CO2 : CO = 1 : 1, whereas LSFMo maintains a pure phase of perovskite after treatment under the same condition, implying that Mo doping improves the CO2 tolerance and reduction stability of the electrode. Although a small amount of Fe metal has a catalytic effect on CO2, a large amount of SrCO3 affects the conductivity and covers active sites of LSF, which is harmful to the electrochemical performance.

Conductivity

To study the influence of conductivity after Mo doping in LSF material, the conductivity of the two materials in oxidizing and reducing atmosphere was tested, respectively. Fig. 2(a) demonstrates the electrical conductivity curves of the two materials in the air with the change in temperature from 500 °C to 850 °C. Conductivity first increases and subsequently decreases with an increase in temperature, which is similar to that of La0.8Sr0.2Co1–xFe3xO3, Sr2Mg1–xCo1+xO3, and Nd0.7Sr0.3Fe1–xCo3O7 materials. The variation trend of LSF and LSFMo conductivity presents the small polar hopping conductivity mechanism.23 At 650 °C, the conductivity of LSF reaches the maximum value (301.5 S cm−1), whereas the value of LSFMo is only 109.2 S cm−1. The electrical conductivity of materials changes to metallic conductance with a further increase in temperature, thus, they possess the properties of a metallic conductor, i.e., the resistance increases and the electrical conductivity decreases as the temperature rises.22–24 For Fe-based perovskite, the conductivity of the materials is mainly determined by electrical conductivity,25–28 which is performed by the B–O–B bond form of the Zerner exchange mechanism, expressed as Fe4+–O22––Fe3+ → Fe3+–O2––Fe4+ → Fe3+–O2––Fe4+, so the higher concentration of Fe3+/Fe4+ redox couples make the higher conductivity for LSF. After doping of Mo in the LSF, the high-valence Fe is partially reduced, leading to the reduction of Fe3+/Fe4+ couples and decrease in conductivity.

The LSF and LSFMo samples were placed in a CO2 : CO = 1 : 1 atmosphere to record the conductivity in the reducing atmosphere, as shown in Fig. 2(b). The conductivity of LSF and LSFMo increases with the rise of temperature, which is similar to the trend of Sr2Fe1.6Mo0.4O3, La0.3Sr0.7FeO3, TiO2, and materials.27–28 Compared with LSF material, the conductivity of LSFMo increased at the same temperatures. For example, it increased from 0.24 S cm−1 to 0.41 S cm−1 at 850 °C. This is because the electrode presents n-type conductance at low oxygen partial pressure, in which the conductivity is dominated by the free electrons.27–29 Fe4+ and Mo6+ are reduced to generate the oxygen vacancies during this condition, as shown in formulas (1)–(3),

\[
2\text{Fe}^{4+} + \text{O}_2^\bullet + \frac{1}{2}\text{O}_2 \rightarrow 2\text{Fe}^{2+} + \text{O}_2^\bullet
\]

\[
2\text{Mo}^{6+} + \text{O}_2^\bullet + \frac{1}{2}\text{O}_2 \rightarrow 2\text{Mo}^{4+} + \text{O}_2^\bullet
\]

\[
\text{O}_2^\bullet + \frac{1}{2}\text{O}_2^\bullet + 2e' \rightarrow \text{O}^\bullet + 2e'
\]

where FeFe' and MoMo' represent the ions of Fe4+ and Mo6+, respectively. The reduction of Fe4+ and Mo6+ is accompanied by the consumption of lattice oxygen, resulting in the formation of oxygen vacancies. The formation of oxygen vacancies increases free electrons, which leads to the increase of electrical conductivity. Meanwhile, the increase in oxygen vacancies also improves the rate of the oxygen reduction reaction.

As shown in Fig. 2(b), the corresponding Arrhenius curves and the activation energy (Ea) of LSF and LSFMo are calculated by the following equation:

![Fig. 2](image-url) The conductivity of LSF and LSFMo sample (a) in air, (b) in CO2 : CO = 1 : 1. Arrhenius plot of the conductivity values with various temperatures shown as an inset.
\[
\sigma = \frac{A}{T} \exp \left( \frac{-E_a}{KT} \right)
\]

\(E_a\) is related to the electron transfer of the electrode for the oxygen reduction reaction. The value of LSF is 0.823 eV, while that of LSFMo is reduced to 0.746 eV, suggesting that the introduction of Mo ions improves the electron transfer of material. From the stability and conductivity of two materials in the reducing atmosphere, it can be concluded that LSFMo is more suitable for SOEC cathode than LSF.

**XPS**

The element valence of the LSF and LSFMo powders was tested by XPS. Fig. 3 shows the XPS spectra of each sample in the oxidation and reduction states, respectively. The data were fitted by Xpspeak41, as shown in Table 1.

There are three peaks in the Fe 2p spectrum, which correspond to Fe 2p3/2, Fe 2p1/2, and Fe satellite, respectively. For the Fe 2p3/2 spectra of LSF oxidation (denoted as LSF(Ox)) and LSFMo oxidation (denoted as LSFMo(Ox)), three interval fixed peaks of Fe element are further analyzed. The binding energy of approximately 709.2, 710.3 and 711.9 eV correspond to Fe2+, Fe3+ and Fe4+ peaks, respectively.30,31 In the oxidized LSF sample (Fig. 3(a)), Fe2+, Fe3+ and Fe4+ account for 13%, 45.1% and 41.9%, respectively, while these values change to 17%, 48.2%, and 34.8% in LSFMo(Ox) (Fig. 3(b)). The content of Fe4+ in LSFMo(Ox) is significantly decreased, verifying the above XRD and conductivity results. The fitting results show that the average valence state of Fe in oxidized LSF is 3.29, whereas the

<table>
<thead>
<tr>
<th>Cond.</th>
<th>Samples</th>
<th>Fe 2p3/2 (%)</th>
<th>Mo 3d5/2 (%)</th>
<th>O 1s</th>
<th>Average</th>
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<td></td>
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<td>Fe4+</td>
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<tr>
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<td>35.66</td>
<td>30.66</td>
<td>20.03</td>
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<td>33.83</td>
<td>39.87</td>
<td>26.30</td>
<td>20.03</td>
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**Fig. 3** XPS spectra of Fe 2p, Mo 3d and O 1s in LSF(Ox) (a) and (g); LSFMo(Ox) (b), (e) and (h); LSF(Re) (c) and (i); LSFMo(Re) (d), (f) and (j).
value of LSFMo decreases to 3.18. It is evident that in the charge compensation mechanism, partial replacement of high-valence Mo could reduce the Fe valence state. Fe⁴⁺/Fe³⁺ electron pairs play a crucial role in the process of electron transport. The decrease in Fe⁴⁺/Fe³⁺ electron pairs blocks the electron transition channel in LSFMo, reducing the conductivity of the electrons.

According to the binding energy intensity, the spectrum can be divided into four peaks. The first peak at 528.5 eV is attributed to the lattice oxygen (Olat). The second peak is O⁻/O₂⁻² species in the form of highly oxidized oxygen at 529.7 eV, and the third peak at O represents adsorbed oxygen or hydroxyl groups at 531.2 eV. The fourth peak at approximately 532.2 eV is related to the molecular water adsorbed on the surface (O₃OH).²⁷⁻³⁹ In the O 1s spectrum, both O⁻/O₂⁻² and O constitute adsorbed oxygen (Oads), representing the oxygen defect concentration. As shown in Table 1, the relative content of Oads increases after Mo doping, indicating that O⁻/O₂⁻² will be generated as a charge compensation when Fe⁴⁺ is replaced by higher valence Mo⁶⁺, which increases the concentration of oxygen defects.

The ratio of Oads/Olat can be reflected as the content of oxygen vacancy in materials, and the ratio values of LSF(Ox) and LSFMo(Ox) are 2.10 and 2.17, respectively, indicating that Mo doping can effectively increase the oxygen vacancy concentration of electrode materials. In the reduction state, due to the reduction of Fe and Mo, the values are changed to 2.22 and 2.42, respectively. Compared with LSF, LSFMo shows more oxygen vacancies in both oxidation and reduction states, which would provide more active sites and higher oxygen ion conduction in the reaction process.

**Electrochemical performance**

Fig. 4 exhibits the variation of the current density of LSF(LSFMo)/LSGM/LSCF-GDC electrolysis cells with voltage ranging from 0 to 2 V for pure CO₂ electrolysis. The measured OCV of the cells is approximately 0.1 V, suggesting good airtightness. The current density increases as the voltage increases during the entire voltage range. In addition, the electrolytic current density of the LSFMo cell is higher than that of the LSF cell under the same applied voltage, showing that the catalytic performance of Mo-doped material is improved. The formation of oxygen vacancies and increases of conductivity in a reducing atmosphere.

The spectra of Mo 3d in LSFMo(Ox) and LSFMo(Re) are shown in Fig. 3(e) and (f). The orbital spin splitting peak of Mo 3d shows two wide peaks, one is 3d₅/₂ with low energy and the other is 3d₃/₂ with high energy. Both of them consist of Mo⁵⁺ and Mo⁶⁺ peaks. The fitting results manifest that many Mo⁶⁺ ions (89.42%) existed in LSFMo(Ox), whereas some Mo⁶⁺ ions are converted into Mo⁵⁺ after reduction, which further promotes the formation of oxygen vacancies and increases of conductivity in a reducing atmosphere.
current density of the LSF cell is 1.56 A cm\(^{-2}\) when voltage reaches 2 V, while the value of LSFMo cell is increased to 2.13 A cm\(^{-2}\) with 36.5% promotion. The performance of LSFMo cell at 1.5 V (1.07 A cm\(^{-2}\)) is also higher than other Fe-based materials, such as La\(_{0.5}\)Sr\(_{0.5}\)Fe\(_{0.95}\)V\(_{0.05}\)O\(_{3-\delta}\) (0.53 A cm\(^{-2}\) at 1.5 V), La\(_{0.6}\)Sr\(_{0.4}\)Fe\(_{0.8}\)Ni\(_{0.2}\)O\(_{3-\delta}\) (0.57 A cm\(^{-2}\) at 1.5 V) and La\(_{0.6}\)Ca\(_{0.4}\)Fe\(_{0.8}\)Ni\(_{0.2}\)O\(_{3-\delta}\) (0.75 A cm\(^{-2}\) at 1.5 V),\(^{42}\) suggesting the excellent electrolysis performance.

To analyze the difference in the electrolysis performance between LSF and LSFMo electrolytic cells, the EIS for CO\(_2\) electrolysis were tested under different voltages at 800 °C, as shown in Fig. 5. The intercept between the impedance spectrum and the X-axis represents \(R_s\), and the voltage gradually decreases as the voltage increases, which is caused by the thermal effect of electric current under high voltage. Meanwhile, \(R_p\) also diminishes due to the reduction activation of the electrode. Under the same voltage, LSFMo shows lower \(R_s\) and \(R_p\). For example, the \(R_s\) and \(R_p\) of LSF are 0.33 \(\Omega\) cm\(^{2}\) and 0.49 \(\Omega\) cm\(^{2}\) at 1.2 V, respectively, whereas those of LSFMo are decreased to 0.22 \(\Omega\) cm\(^{2}\) and 0.24 \(\Omega\) cm\(^{2}\), respectively. The diminution of \(R_s\) and \(R_p\) together promotes the electrochemical performance of LSFMo. Since CO is produced in the fuel electrode during the electrolysis process and the conductivity of the LSF electrode is low in the reducing atmosphere, the decrease in \(R_p\) is mainly caused by the increase of conductivity in the LSFMo electrode, while the change in \(R_p\) may be due to the promotion of catalytic activity caused by oxygen defects. The \(R_p\) is further decreased to 0.13 \(\Omega\) cm\(^{2}\) when the voltage is increased to 1.5 V, which is significantly lower than La\(_{0.6}\)Sr\(_{0.4}\)Fe\(_{0.8}\)Ni\(_{0.2}\)O\(_{3-\delta}\) (0.211 \(\Omega\) cm\(^{2}\) at 2 V),\(^{41}\) La\(_{0.6}\)Sr\(_{0.4}\)Fe\(_{0.8}\)Ni\(_{0.2}\)O\(_{3-\delta}\) (0.2 \(\Omega\) cm\(^{2}\) at 1.6 V),\(^{40}\) and Sr\(_2\)Fe\(_{1.5}\)Mo\(_{0.5}\)O\(_{4}\)–SDC (0.19 \(\Omega\) cm\(^{2}\)) at 1.5 V\(^{28}\) reported in the literature, demonstrating excellent catalytic activity. To further explain the \(R_p\) change, EIS are fitted by Zview with the model \(L_0 - R_s - (R_{th}/CPE1) - (R_{th}/CPE2) - (R_{th}/CPE3)\), where \(L_0\) represents the inductance coefficient, which is related to the equipment and connecting cables; \(R_s\) is the total ohmic resistance generated by the electrolyte, electrode and lead wires; and \(R_{th}\), \(R_m\) and \(R_L\) represents polarization resistances in high-frequency, middle-frequency and low-frequency, respectively. For CO\(_2\) electrolysis, \(R_s\) is related to the charge transfer process, and the middle-frequency \(R_{th}\) is attributed to the adsorption and dissociation of CO\(_2\) at the porous electrode, while \(R_m\) is ascribed to the process of CO\(_2\) diffusion. The fitting results at 1.2 V are listed in Table 2. Compared with LSF, \(R_{th}\) and \(R_m\) of LSFMo are significantly decreased, indicating that LSFMo has better CO\(_2\) adsorption and dissociation, which is attributed to more oxygen vacancy concentrations. Moreover, SrCO\(_3\) is produced in reduction atmosphere for LSF, which will cover the active sites on the surface of material and affects the electrolysis performance.\(^{42}\) The Mo doping suppresses the formation of SrCO\(_3\), which may be another reason of performance improvement.

The electrolysis stability of the LSFMo cell was measured at 800 °C and 1.2 V constant voltage, as shown in Fig. 6(a). The attenuation rate is 0.044 mA cm\(^{-2}\) min\(^{-1}\) (0.008% A cm\(^{-2}\) min\(^{-1}\)), which is lower than that of LSF (0.022% A cm\(^{-2}\) min\(^{-1}\))\(^{28}\) and La\(_{0.75}\)Sr\(_{0.25}\)Cr\(_{0.5}\)Mn\(_{0.5}\)O\(_3\) (4 mA cm\(^{-2}\) min\(^{-1}\))\(^{43}\) implying that LSFMo cell has reasonable stability. Fig. 6(b) exhibits the microstructures of LSFMo/LSGM/LSCF-GDC cells after the stability test. The thickness of the LSGM electrolyte and LSFMo electrode is approximately 300 μm and 15 μm, respectively. The LSFMo electrode has a porous structure, whereas the LSGM electrolyte shows a dense structure. The electrode and electrolyte maintain good contact without obvious cracks, indicating good thermal compatibility during operation. In addition, the

### Table 2

The fitting results of EIS about LSF and LSFMo cells at 1.2 V

<table>
<thead>
<tr>
<th>Samples</th>
<th>(R_s) ((\Omega) cm(^{2}))</th>
<th>(R_m) ((\Omega) cm(^{2}))</th>
<th>(R_L) ((\Omega) cm(^{2}))</th>
<th>(R_p) ((\Omega) cm(^{2}))</th>
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<tbody>
<tr>
<td>LSF</td>
<td>0.049</td>
<td>0.35</td>
<td>0.091</td>
<td>0.49</td>
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<tr>
<td>LSFMo</td>
<td>0.047</td>
<td>0.177</td>
<td>0.018</td>
<td>0.242</td>
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Fig. 6  (a) The stability of LSFMo cell at 1.2 V, (b) SEM of cross-section of cathode after stability test, (c) XRD of cathode after stability test.
cathode after the test is detected by XRD as shown in Fig. 6(c). The result that no SrCO₃ and Fe are generated verifies the excellent CO₂ tolerance and reduction stability of the LSFMo electrode.

Conclusion

In summary, a new Mo-doped LSF perovskite material was synthesized using a solid-phase method, and the electrochemical properties of SOECs as cathode materials were studied. Mo doping significantly improves the CO₂ tolerance of LSF materials and solves the problem of SrCO₃ generation. 5% Mo addition increases the oxygen defects of the material surface, which improves the conductivity in the reducing atmosphere and the catalytic activity of cathode material. Compared with LSF, LSFMo exhibits smaller Rₐ and Rₚ as the SOEC cathode, which makes the LSFMo electrolytic cell possess higher electrolytic performance. For example, the current densities increase from 1.56 A cm⁻² to 2.13 A cm⁻² under 2 V. The LSFMo cell shows excellent stability in the 60 h electrolysis test. Therefore, LSFMo can be considered as a promising cathode material for SOECs at intermediate temperatures.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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