


Cite this: *RSC Adv.*, 2021, 11, 32942

An aerobic oxidation of alcohols into carbonyl synthons using bipyridyl-cinchona based palladium catalyst†

Ravi Kumar Cheedarala,^a Ramasamy R. Chidambaram,^b Ayyanar Siva^{*b} and Jung Il Song^{*a}

We have reported an aerobic oxidation of primary and secondary alcohols to respective aldehydes and ketones using a bipyridyl-cinchona alkaloid based palladium catalytic system (**PdAc-5**) using oxygen at moderate pressure. The **PdAc-5** catalyst was analysed using SEM, EDAX, and XPS analysis. The above catalytic system is used in experiments for different oxidation systems which include different solvents, additives, and bases which are cheap, robust, non-toxic, and commercially available on the industrial bench. The obtained products are quite appreciable in both yield and selectivity (70–85%). In addition, numerous important studies, such as comparisons with various commercial catalysts, solvent systems, mixture of solvents, and catalyst mole%, were conducted using **PdAc-5**. The synthetic strategy of oxidation of alcohol into carbonyl compounds was well established and all the products were analysed using ¹H NMR, ¹³CNMR and GC-mass analyses.

Received 2nd August 2021
Accepted 13th September 2021

DOI: 10.1039/d1ra05855j

rsc.li/rsc-advances

1. Introduction

Palladium-catalysed aerobic transformation has been one of the most challenging topics within organometallic chemistry. This system is widely used in oxidation, involving alcohols, alkenes, oxidative coupling and so on, and has triggered commercial interest in the above-mentioned chemistry. Generally, an alcohol oxidation reaction plays an important role in functional group manipulation and heteroatom incorporation in the synthesis of chemically and biologically active compounds. Thus, developing a new process in both academic and industrial sectors to beat conventional systems¹ becomes a bit challenging for chemists. These challenges force researchers to come up with novel catalysis design systems for the above-mentioned transformation.² For key factors like stoichiometric oxidants, their versatility, expense, and environmental impact need to be considered. A viable and attractive approach is the use of metal-catalysed oxidation in the presence of molecular oxygen as an oxidant.³

In the past 4–5 decades, there have been significant advances in the development of new transition metal catalysed oxidations

using improved catalytic systems⁴ with increased turnover numbers. The use of highly stable ligands has brought about these advances, thus initiating the search for new catalytic systems. The development of ligand-modulated catalytic systems, which do not require the use of a co-catalyst, greatly simplified the catalysis and has allowed for an elucidation of the mechanistic features of these reactions. A significant number of mechanistic studies^{5,6} have been performed by various research groups over the years and more are on their way. Recently, our group have reported a ligand-free recyclable palladium catalyst for (1) the oxidation of alcohols into carbonyl compounds, (2) the highly selective α -alkylation of ketones with alcohols and imines, and (3) the production of secondary amines by coupling benzyl alcohols and primary amines.⁷ However, these studies exhibited both important improvements and the current limitations of palladium oxidase reactions, which illustrate the challenges of developing new more efficient direct molecular oxygen-coupled processes.⁸ Based on these previous points, we concentrate on developing a Pd-based chiral organometallic catalyst for the aerobic oxidation reaction under mild reaction conditions.

2. Experimental section

2.1 Materials and methods

5,5-Dimethyl-2,2-bipyridine, AIBN, (+)-cinchonine, and *N*-bromosuccinimide were purchased from Sigma-Aldrich. 2-Methyl benzyl alcohol, 4-methyl benzyl alcohol, 2-methoxy benzyl alcohol, 4-methoxy benzyl alcohol, 3-methoxy 4-hydroxy benzyl alcohol, 2-fluoro benzyl alcohol, 3-fluoro benzyl alcohol, 4-fluoro

^aResearch Institute of Mechatronics, Department of Mechanical Engineering, Changwon National University, Changwon City, Republic of Korea. E-mail: ravi@changwon.ac.kr; jisong@changwon.ac.kr

^bSupramolecular and Organometallic Chemistry Lab, Department of Inorganic Chemistry, School of Chemistry, Madurai Kamaraj University, Madurai 625021, Tamil Nadu, India. E-mail: drasiva@gmail.com

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ra05855j



benzyl alcohol, 2-chloro benzyl alcohol, 4-chloro benzyl alcohol, 2,3-dichloro benzyl alcohol, 2-nitro benzyl alcohol, 3-nitro benzyl alcohol, 4-*tert*-butyl benzyl alcohol, mesitaldehyde, 2,6-dimethyl benzyl alcohol, 3,5-dibromo-2,4,6-trimethyl benzyl alcohol, 1-phenyl-1 ethanol, 1-phenyl-1 propanol, diphenyl methanol, cyclohexyl methanol, cyclohexanol, 1,4-dihydroxy cyclohexane, cyclododecanol, piperonyl alcohol, 4-acetamido benzyl alcohol, 4-trifluoromethyl benzyl alcohol, 2,4-trifluoromethyl benzyl alcohol, 4-phenyl benzyl alcohol, pyridine-2-yl methanol, pyridine-3-yl methanol, pyridine-4-yl methanol, naphthyl-1-methanol, and naphthyl-2-methanol were purchased from various sources as technical grade. All the solvents for the reactions were obtained as laboratory grade. All the solvents for analysis were purchased as analytical grade.

The melting points were measured in open capillary tubes and are uncorrected. The ^1H , and ^{13}C NMR spectra were recorded on a Bruker (Avance) 400 MHz NMR instrument using TMS as an internal standard, and CDCl_3 and DMSO as solvents. Standard Bruker software was used throughout. Chemical shifts are given in parts per million (δ -scale) and the coupling constants are given in Hertz. Silica gel-G plates (Merck) were used for TLC analysis with a mixture of *n*-hexane and ethyl acetate as an eluent. Column chromatography was carried out in silica gel (60–120 mesh) using *n*-hexane, dichloro methane (DCM), methanol, and ethyl acetate as eluents. Pressure reactions were carried out using a PREMEX A-Line Avalon Instrument. Distillation of the products was carried out using a glass oven – B585 Kugelrohr (Buchi made) instrument. Electrospray Ionization Mass Spectrometry (ESI-MS) analyses were recorded in an LCQ Fleet, Thermo Fisher Instruments Limited, US. ESI-MS was performed in positive ion mode. The collision voltage and ionization voltage were -70 V and -4.5 kV , respectively, using nitrogen as atomization and desolvation gas. The desolvation temperature was set at $300\text{ }^\circ\text{C}$. The relative amount of each component was determined from the LC-MS chromatogram, using the area normalization method. GC was recorded using an Agilent GC 6890 N, with split ratio: $1 : 100$, $250\text{ }^\circ\text{C}$, FID: $300\text{ }^\circ\text{C}$, flow rate: nitrogen 30 mL min^{-1} . Column: HP-5; $30\text{ m} \times d: 0.32\text{ mm} \times f: 0.25\text{ mm}$, $-60\text{ to }325\text{ }^\circ\text{C}$. GCMS was recorded using a 5977B GC/MSD single quadrupole analyser using helium as carrier gas.

3. Experimental section

3.1 Preparation of cinchona based palladium catalyst (PdAc-5)

3.1.1 Preparation of 5,5-bis(bromomethyl)-2,2-bipyridine (2). NBS (1.93 g, 10.85 mmol) was added in solid lots slowly to a mixture of **1** (1 g, 5.42 mmol) and AIBN (17 mg, 0.10 mmol) in CCl_4 (75 mL) and it was refluxed for 2 h under a nitrogen atmosphere. The precipitated succinimide was hot filtered and the solid was washed with CCl_4 (100 mL). The combined CCl_4 phases were evaporated. The remaining solid was dissolved in CH_2Cl_2 (100 mL) and extracted with 0.5 M $\text{Na}_2\text{S}_2\text{O}_3$ solution ($2 \times 150\text{ mL}$). The combined $\text{Na}_2\text{S}_2\text{O}_3$ fractions were extracted with CH_2Cl_2 (50 mL) and the combined CH_2Cl_2 layers were dried (Na_2SO_4) and concentrated. The crude product was purified by

column chromatography (silica gel, EtOAc/hexane, $1 : 4$) giving a white powder. Yield 63%. Mp: $193\text{--}194\text{ }^\circ\text{C}$. ^1H NMR (300 MHz, CDCl_3) δ 8.67 (d, $J = 3\text{ Hz}$, 2H), 8.39 (d, $J = 9\text{ Hz}$, 2H), 7.87–7.84 (m, 2H), 4.53 (s, 4H). ^{13}C NMR (75 MHz, CDCl_3) δ_{C} 149.78, 143.67, 131.82, 128.19, 115.44, 23.74. ESI-MS (M^+); 342.88.

3.1.2 Synthesis of cinchonine (contains free $-\text{OH}$) based CPTC (4).⁸ A mixture of 5,5-bis(bromomethyl)-2,2-bipyridine **2** (1 g, 2.9 mmol), cinchona (free OH) and **3** (6.38 mmol) was dissolved in THF (50 mL) and heated to reflux overnight. The resulting brown solid was filtered, washed with diethyl ether, and dried to yield **4**. Yield: 90%. ^1H NMR (400 MHz, DMSO) δ 9.12 (s, 2H), 9.00 (d, $J = 4\text{ Hz}$, 2H), 8.64 (d, $J = 8\text{ Hz}$, 2H), 8.41 (t, $J = 8\text{ Hz}$, 4H), 8.12 (d, $J = 8\text{ Hz}$, 2H), 7.88–7.71 (m, 6H), 6.86 (t, $J = 6\text{ Hz}$, 2H), 6.55 (s, 2H), 6.09–5.98 (m, 2H), 5.33–5.22 (m, 6H), 5.11 (d, $J = 16\text{ Hz}$, 2H), 4.00 (t, $J = 16\text{ Hz}$, 4H), 3.58–3.50 (m, 2H), 3.11 (d, $J = 8\text{ Hz}$, 2H), 2.67 (d, $J = 8\text{ Hz}$, 2H), 2.33 (t, $J = 12\text{ Hz}$, 2H), 1.91 (s, 2H), 1.79 (d, $J = 8\text{ Hz}$, 2H), 1.08 (t, $J = 8\text{ Hz}$, 2H), 0.87 (t, $J = 8\text{ Hz}$, 2H); ^{13}C NMR (100 MHz, DMSO) δ 155.60, 153.60, 150.14, 147.99, 144.89, 142.76, 137.05, 129.42, 128.79, 127.29, 125.02, 124.33, 120.75, 120.08, 118.94, 117.03, 67.38, 64.68, 59.37, 55.92, 54.08, 36.69, 27.51, 26.34, 22.98. ESI-MS (M^+): 931.28.

3.1.3 Synthesis of cinchona (contains free $-\text{OH}$) based CPTC–palladium catalyst (5).⁹ A mixture of **4** (2.0 g, 1.73 mmol), palladium acetate (0.4 g), and DMF (10 mL) was stirred at $100\text{ }^\circ\text{C}$ for about 16 h. The solvent was then removed under reduced pressure and the residual mass was diluted with diethyl ether (50 mL), filtered and washed thoroughly with diethyl ether (25 mL). It was then dried under vacuum for 12 h until it reached a constant weight. The yield was 2.1 g, 88%.

3.1.4 General procedure for the oxidation of alcohols (7b to 35b). A mixture of alcohol (6.0 mmol), toluene (5.0 mL), K_2CO_3 (5 mol%), and PdAc-5 (1 mol%) was pressurised in a mini-clave apparatus using oxygen gas (5 bar) at $80\text{ }^\circ\text{C}$ for about 8 h. The reaction was monitored by GC for reaction completion. The mass was then removed from the reactor, filtered over a hyflo bed and diluted with water (20 mL). It was then extracted using diethyl ether ($10\text{ mL} \times 2$). The combined organic layers were washed with water ($5\text{ mL} \times 5$), dried over sodium sulfate and concentrated to remove solvent. Then it was purified by column chromatography using silica gel 60–120 mesh size (hexane : ethyl acetate) as an eluent or by distillation using a Kugelrohr instrument. The purified product was then analysed by NMR and mass spectroscopy to confirm the product.

4. Characterization of products

4.1 2-Methylbenzaldehyde (7b)

Colourless liquid. Yield: 84.0%. GC purity: 94.6%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.26 (s, 1H), 7.79 (dd, $J = 7.6, 1.5\text{ Hz}$, 1H), 7.47 (td, $J = 7.5, 1.6\text{ Hz}$, 1H), 7.36 (td, $J = 7.5, 1.2\text{ Hz}$, 1H), 7.28–7.22 (m, 1H), 2.67 (s, 3H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 192.73, 140.53, 134.07, 133.58, 131.98, 131.70, 126.25, 19.52.

4.2 4-Methylbenzaldehyde (8b)

Colourless liquid. Yield: 89.5%. GC purity: 93.7%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.96 (s, 1H), 7.77 (d, $J = 8.1\text{ Hz}$, 2H), 7.33 (d, $J =$

7.8 Hz, 2H), 2.44 (s, 3H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 191.94, 145.49, 134.13, 129.71, 21.83.

4.3 2-Methoxybenzaldehyde (9b)

Pale yellow liquid. Yield: 80.5%. GC purity: 94.2%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.45 (s, 1H), 7.80 (dd, $J = 7.7, 1.8$ Hz, 1H), 7.52 (ddd, $J = 8.2, 7.4, 1.8$ Hz, 1H), 7.08–6.90 (m, 2H), 3.90 (s, 3H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 189.62, 161.75, 135.83, 128.38, 124.80, 120.55, 111.58, 55.53.

4.4 4-Methoxybenzaldehyde (10b)

Pale brown oil. Yield: 86.6%. GC purity: 99.10%, ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.88 (s, 1H), 7.88–7.86 (d, $J = 8$ Hz, 2H), 7.18–7.16 (d, $J = 8$ Hz, 2H), 3.83 (s, 3H), ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 55.8, 114.8, 129.2, 132.3, 166.4, 191.0.

4.5 3-Methoxy-4-hydroxybenzaldehyde (11b)

Off-white solid. Yield: 73.5%. GC purity: 99.62%, ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.81 (s, 1H), 7.43–7.41 (d, $J = 8$ Hz, 1H), 7.05–7.03 (d, $J = 8$ Hz, 1H), 6.75 (s, 1H), 5.65 (s, 1H), 3.93 (s, 3H), 4.22 (s, 1H), 3.83 (s, 3H), ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 191.16, 151.93, 147.30, 129.73, 127.58, 114.55, 108.96, 56.08.

4.6 2-Fluorobenzaldehyde (12b)

Yellow liquid. Yield: 61.99%. GC purity: 98.25%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.37 (d, $J = 0.8$ Hz, 1H), 7.88 (td, $J = 7.4, 1.9$ Hz, 1H), 7.61 (dddd, $J = 8.4, 7.3, 5.4, 1.9$ Hz, 1H), 7.33–7.09 (m, 2H); ^{13}C NMR (101 MHz, CDCl_3) δ_{C} 187.12, 166.34, 162.92, 136.29, 128.64, 124.58, 116.45.

4.7 3-Fluorobenzaldehyde (13b)

Yellow liquid. Yield: 66.58%. GC purity: 97.10%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.00 (d, $J = 1.9$ Hz, 1H), 7.69 (dt, $J = 7.5, 1.3$ Hz, 1H), 7.61–7.48 (m, 2H), 7.39–7.26 (m, 1H); ^{13}C NMR (101 MHz, CDCl_3) δ_{C} 190.82, 164.69, 161.38, 138.36, 130.74, 126.00, 121.52, 115.27.

4.8 4-Fluorobenzaldehyde (14b)

Yellow liquid. Yield: 76.58%. GC purity: 90.56%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.98 (s, 1H), 7.92 (dd, $J = 8.7, 5.4$ Hz, 2H), 7.22 (t, $J = 8.5$ Hz, 2H); ^{13}C NMR (101 MHz, CDCl_3) δ_{C} 190.45, 168.16, 164.76, 132.91, 132.18, 116.30.

4.9 2-Chlorobenzaldehyde (15b)

Light brown liquid. Yield: 69.6%. GLC purity: 92.11%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.47 (d, $J = 0.8$ Hz, 1H), 7.91 (dd, $J = 7.7, 1.8$ Hz, 1H), 7.53 (ddd, $J = 8.0, 7.1, 1.8$ Hz, 1H), 7.48–7.32 (m, 2H). ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 189.67, 137.81, 135.03, 132.32, 130.49, 129.24, 127.18.

4.10 4-Chlorobenzaldehyde (16b)

Light brown solid. Yield: 78.5%. GLC purity: 97.7%. Mp: 46–50 °C (lit. 47–51 °C). ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.98 (s, 1H),

7.82 (d, $J = 8.5$ Hz, 2H), 7.51 (d, $J = 8.4$ Hz, 2H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 190.75, 140.81, 134.61, 130.80, 129.34.

4.11 2,3-Dichlorobenzaldehyde (17b)

Pale yellow solid. Yield: 82.6%. GC purity: 90.15% (decomposes upon storage), Mp: 62–65 °C (64–67 °C, lit.). ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.47 (s, 1H), 7.84–7.82 (d, $J = 8$ Hz, 1H), 7.71–7.69 (d, $J = 8$ Hz, 1H), 7.33–7.37 (t, $J = 6$ Hz, 1H). ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 189.34, 136.02, 135.59, 134.33, 134.20, 127.42, 127.60.

4.12 2-Nitrobenzaldehyde (18b)

Brown solid. Yield: 88%. GLC purity: 95.89%. Mp: 42–44 °C (reported: 42–44 °C); ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.41 (s, 1H), 8.12–8.10 (m, 1H), 7.95–7.93 (m, 1H), 7.81–7.73 (m, 2H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 188.31, 149.69, 134.22, 133.84, 131.45, 129.75, 124.62.

4.13 3-Nitrobenzaldehyde (19b)

Pale yellow solid. Yield: 58.6%. GC purity: 92.64%, Mp 57.5–58.5 °C (lit. 57–58.5 °C). ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.14 (s, 1H), 8.72 (s, 1H), 8.51–8.49 (d, $J = 8$ Hz, 1H), 8.27–8.25 (d, $J = 8$ Hz, 1H), 7.82–7.78 (t, $J = 6$ Hz, 1H). ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 189.83, 148.76, 137.38, 134.74, 130.44, 128.59, 124.40.

4.14 4-*t*-Butylbenzaldehyde (20b)

Colourless liquid. Yield: 85.5%. GC purity: 90.41%, ^1H NMR (400 MHz, CDCl_3) δ_{H} 9.88 (s, 1H), 7.83–7.81 (d, $J = 8$ Hz, 2H), 7.48–7.46 (d, $J = 8$ Hz, 2H), 1.35 (s, 9H), ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 31.3, 34.2, 125.5, 129.5, 133.8, 157.1, 193.0.

4.15 2,6-Dimethylbenzaldehyde (21b)

Pale brown liquid. Yield: 78.66%. GC purity: 96.36%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.62 (s, 1H), 7.32 (t, $J = 7.6$ Hz, 1H), 7.09 (d, $J = 7.6$ Hz, 2H), 2.61 (s, 6H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 193.39, 140.99, 132.85, 132.30, 129.58, 20.34.

4.16 Mesitaldehyde (22b)

Colourless liquid. Yield: 80.14%. GC purity: 95.87%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.53 (s, 1H), 6.87 (s, 2H), 2.56 (s, 6H), 2.30 (s, 3H); ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 192.76, 143.66, 141.31, 130.38, 129.82, 21.31, 20.33.

4.17 3,5-Dibromo-2,4,6-trimethylbenzaldehyde (23b)

Colourless solid. Yield: 68.8%. GC purity: 95.45%. Mp: 162–164 °C (lit. 160 °C, decomposes); ^1H NMR (CDCl_3 , 400 MHz) δ_{H} 10.47 (s, 1H), 2.74 (s, 3H), 2.57 (s, 6H); ^{13}C NMR (CDCl_3 , 100 MHz) δ_{C} 194.0, 142.5, 137.6, 135.0, 127.7, 26.9, 20.4.

4.18 1-Phenylethan-1-one (24b)/(acetophenone)

Colourless oil. Yield: 87.55%. GC purity: 99.73%. ^1H NMR (400 MHz, CDCl_3) δ_{H} 7.94–7.57 (m, 5H), 2.52 (s, 3H). ^{13}C NMR (100 MHz, CDCl_3) δ_{C} 197.11, 133.14, 128.80, 128.64, 26.45.



4.19 Propiophenone (25b)

Colourless liquid. Yield: 82.33%. GLC purity: 98.8%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 7.94–7.91 (m, 2H), 7.53–7.48 (m, 1H), 7.43–7.39 (m, 2H), 2.95 (q, $J_1 = J_2 = 7.2$ Hz, 2H), 1.19 (t, $J_1 = J_2 = 2.4$ Hz). ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 200.79, 136.90, 132.86, 128.54, 127.95, 31.75, 8.22.

4.20 Benzophenone (26b)

Colourless solid. Yield: 88.12%. GLC purity: 97.32%. Mp: 48–50 °C (reported: 47–51 °C); ^1H NMR (400 MHz, CDCl_3): δ_{H} 7.81–7.79 (m, 2H), 7.60–7.56 (m, 1H), 7.50–7.46 (m, 2H); ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 196.89, 137.72, 132.53, 130.17, 128.39.

4.21 Cyclohexanecarbaldehyde (27b)

Pale yellow liquid. Yield: 83%. GLC purity: 95.23%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 9.59 (s, 1H), 2.24–2.19 (m, 1H), 1.90–1.20 (m, 10H); ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 205.17, 50.07, 26.08, 26.02, 25.12.

4.22 Cyclohexanone (28b)

Colourless liquid. Yield: 81%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 4.49–4.39 (m, 1H), 3.33–2.90 (m, 1H), 2.41–2.12 (m, 3H), 2.11–1.61 (m, 4H); ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 211.52, 42.00, 27.11, 25.07.

4.23 Cyclohexane-1,4-dione (29b)

Colourless solid. Yield: 77%. Mp: 75–77 °C (reported: 77–78.5 °C); ^1H NMR (400 MHz, CDCl_3): δ 2.67 (s, 4H); ^{13}C NMR (100 MHz, CDCl_3): δ 208.51, 36.73.

4.24 Cyclododecanone (30b)

Colourless solid. Yield: 70.6%. GC purity: 99.76%. Mp 57–59 °C (reported 58–60 °C). ^1H NMR (400 MHz, CDCl_3): δ_{H} 2.48–2.45 (m, 4H), 1.74–1.68 (m, 4H), 1.28–1.30 (m, 14H). ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 212.88, 40.25, 24.73, 24.58, 24.20, 24.20, 22.54, 22.31.

4.25 Piperonyl aldehyde (31b)

Light yellow waxy solid. Yield: 86.25%. GC purity: 96.78%. Mp: 35–39 °C (reported: 37–39 °C). ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.30 (s, 1H), 7.16 (dd, $J = 8.1, 1.7$ Hz, 1H), 7.14 (d, $J = 1.7$ Hz, 1H), 6.62 (d, $J = 8.1$ Hz, 1H), 5.99 (s, 2H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 186.81, 152.20, 151.48, 146.02, 128.14, 107.47, 105.06, 103.89.

4.26 4-Acetamidobenzaldehyde (32b)

Brown solid (slight decomposition observed in storage). Yield: 78.99%. GC purity: 98.25%. Mp: 146–152 °C (reported: 148–152 °C). ^1H NMR (400 MHz, CDCl_3): δ_{H} 9.91 (s, 1H), 8.00 (s, 1H), 7.8–7.79 (m, 2H), 7.71 (d, $J = 8.4$ Hz, 2H), 2.23 (s, 3H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 191.14, 168.86, 143.62, 132.14, 131.14, 119.22, 24.76.

4.27 4-(Trifluoromethyl)benzaldehyde (33b)

Light brown liquid. Yield: 74.52%. GC purity: 92.9%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.09 (s, 1H), 8.00 (dt, $J = 7.9, 0.9$ Hz, 2H), 7.80 (d, $J = 8.1$ Hz, 2H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 191.07, 138.63, 135.79, 135.36, 129.89, 126.11, 121.60, 117.99.

4.28 2,4-Bis(trifluoromethyl)benzaldehyde (34b)

Yellow liquid. Yield: 58.23%. GC purity: 87.25%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.04 (d, $J = 0.5$ Hz, 1H), 8.06 (d, $J = 1.5$ Hz, 1H), 7.90 (dd, $J = 7.5, 0.5$ Hz, 1H), 7.72 (dd, $J = 7.5, 1.5$ Hz, 1H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 189.42, 134.81, 134.80, 134.74, 133.42, 133.17, 129.59, 129.57, 129.56, 129.54, 126.95, 126.92, 126.88, 126.85, 125.22, 125.19, 125.16, 125.13, 124.91, 123.66, 122.76.

4.29 4-Biphenylcarboxaldehyde (35b)

Yellow solid. Yield: 86.54%. GC purity: 98.66%. Mp: 58–62 °C (reported: 57–59 °C). ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.06 (s, 1H), 7.96 (d, $J = 8.3$ Hz, 2H), 7.76 (d, $J = 8.3$ Hz, 2H), 7.68–7.60 (m, 2H), 7.53–7.36 (m, 3H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 191.84, 147.13, 139.65, 135.15, 130.21, 128.96, 128.42, 127.63, 127.3.

4.30 2-Pyridinecarboxaldehyde (36b)

Brown liquid. Yield: 82.77%. GC purity: 96.35%. ^1H NMR (400 MHz, CDCl_3): δ 10.09 (s, 1H), 8.81 (dt, $J = 4.8, 1.3$ Hz, 1H), 7.98 (dt, $J = 7.8, 1.2$ Hz, 1H), 7.91 (td, $J = 7.6, 1.6$ Hz, 1H), 7.57 (ddd, $J = 7.4, 4.8, 1.4$ Hz, 1H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 193.20, 152.60, 150.02, 136.89, 127.70, 121.51.

4.31 4-Pyridinecarboxaldehyde (37b)

Yellow liquid. Yield: 76.25%. GC purity: 91.25%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 9.98 (s, 1H), 8.77 (d, $J = 6.0$ Hz, 2H), 7.60 (d, $J = 6.0$ Hz, 2H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 191.29, 150.97, 141.15, 121.85.

4.32 3-Pyridinecarboxaldehyde (38b)

Light brown liquid (darkens upon storage). Yield: 68.5%. GC purity: 89.23%. ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.04 (s, 1H), 8.99 (d, $J = 2.1$ Hz, 1H), 8.75 (dd, $J = 4.9, 1.8$ Hz, 1H), 8.08 (dt, $J = 7.9, 2.0$ Hz, 1H), 7.41 (dd, $J = 7.9, 4.8$ Hz, 1H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 190.60, 154.55, 151.85, 135.58, 131.19, 123.87.

4.33 5-Nitrofuran-2-carbaldehyde (39b)

Colourless solid. Yield: 68%. Mp: 38–40 °C (reported: 37–39 °C). ^1H NMR (400 MHz, CDCl_3): δ_{H} 9.82 (s, 1H), 7.41 (d, $J = 4$ Hz, 1H), 7.34 (d, $J = 3.6$ Hz, 1H); ^{13}C NMR (100 MHz, CDCl_3): δ_{C} 178.45, 151.09, 118.92, 111.86.

4.34 1-Napthaldehyde (40b)

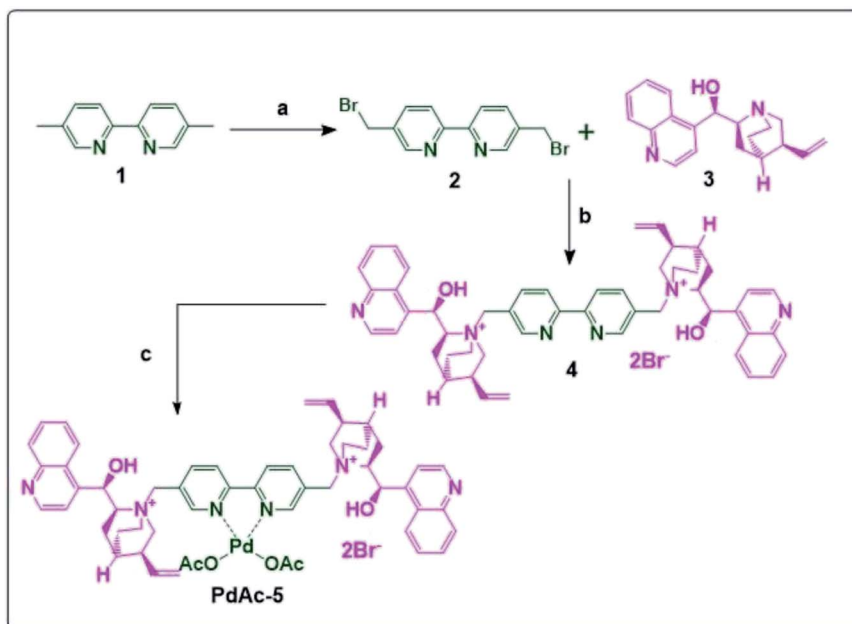
Light brown liquid. Yield: 86.23%. GC purity: 96.18%, ^1H NMR (400 MHz, CDCl_3): δ_{H} 10.38 (s, 1H), 9.26 (dd, $J = 8.6, 1.1$ Hz, 1H), 8.08 (dt, $J = 8.2, 1.0$ Hz, 1H), 8.01–7.83 (m, 2H), 7.79–7.44 (m, 3H); ^{13}C NMR (101 MHz, CDCl_3): δ_{C} 193.45, 136.60, 135.19, 133.60, 131.27, 130.41, 128.97, 128.38, 126.86, 124.77.

4.35 2-Naphthaldehyde (41b)

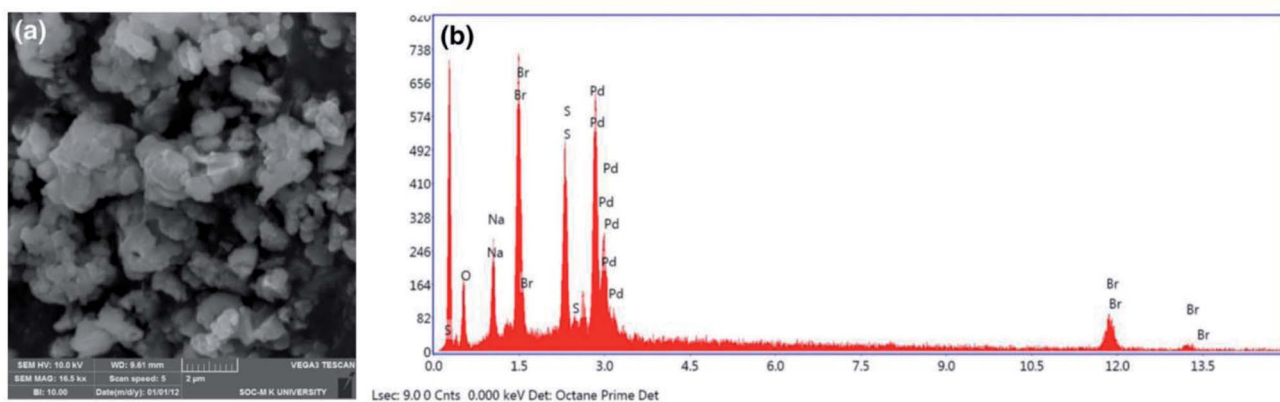
Pale yellow solid. Yield: 83.5%. GC purity: 94.68%. Mp: 59–63 °C (reported: 59–61 °C). ^1H NMR (400 MHz, CDCl_3) δ_{H} 10.16 (s, 1H), 8.34 (s, 1H), 8.05–7.84 (m, 4H), 7.62 (dddd, $J = 17.6, 8.2, 6.9, 1.4$ Hz, 2H); ^{13}C NMR (101 MHz, CDCl_3) δ_{C} 192.23, 136.42, 134.5, 134.08, 132.61, 129.50, 129.08, 128.05, 127.07, 122.74.

5. Results and discussion

Compound 2 was synthesised by the reaction between 1 and NBS in CCl_4 using AIBN as a catalyst in 80% yield. The resulting 2 was reacted with cinchona alkaloid (free base) (3) to obtain the bipyridyl-cinchona ligand (4) in 96% yield.⁸ It was then



Scheme 1 Synthesis of PdAc-5; reagents and conditions: (a): NBS, AIBN, CCl_4 , reflux 4 h; (b): THF, reflux, 16 h; and (c): palladium acetate, DMF, 16 h, 100 °C.



Element	Weight (%)	Atomic weight(%)	Net. int	Error (%)	K ratio	Z	R	A	F
O K	22.13	49.10	104.1	15.47	0.0361	1.2029	0.8692	0.1358	1.000
Na K	10.71	16.53	142.8	13.20	0.0293	1.1044	0.9000	0.2467	1.003
Br L	22.78	10.12	390.8	7.14	0.1318	0.8840	1.0785	0.6501	1.007
S k	12.21	13.52	451.0	7.71	0.0796	1.0972	0.9433	0.5861	1.014
Pd L	32.18	10.74	610.6	5.09	0.2390	0.8540	1.1435	0.8595	1.012

Fig. 1 SEM (a) and EDAX (b) spectrum of PdAc-5.



converted into its palladium acetate derivative (**PdAc-5**), as shown in Scheme 1.

The morphological evidence for **PdAc-5** was confirmed by SEM and EDAX analysis, as presented in Fig. 1.⁹ The SEM image of **PdAc-5** showed the presence of highly crystalline structures (Fig. 1a), which confirmed the incorporation of palladium metal in the catalytic system of **PdAc-5**. Further, the atomic weight (%) of **PdAc-5** is 29.4%, measured by atomic absorption spectroscopy.¹⁰

To investigate the oxidation state of palladium, the binding affinity between the metal and cinchonine-based bipyridine ligand was confirmed by X-ray photoelectron spectroscopy (XPS), as shown in Fig. 2. **PdAc-5** showed two binding energy values that appeared at 337 and 342.42 eV. These were

attributed to Pd 3d_{5/2} and Pd 3d_{3/2}, which were assigned to palladium being present in +2 oxidation states. In addition to that, the free palladium acetate binding energy value that appeared at 338.4 eV for our catalyst 5 showed a 1.24 eV negative shift, which is consistent with previously reported literature. This negative shift indicates that palladium acetate coordinated with the 2,2'-bipyridine nitrogen atom. Furthermore, the N 1s spectrum (Fig. 2b) shows an asymmetric and wide peak, which was fitted with three components at 397.8, 399.3, and 401.02 eV. The components were assigned to the quinoline nitrogen, pyridinic nitrogen, and quaternary nitrogen, respectively. Generally, the reported pyridine nitrogen (–N=) binding energy values appear at 398.2 eV, but we observed an N 1s core level binding energy of 399.3 eV, so the pyridine unit N 1s core level

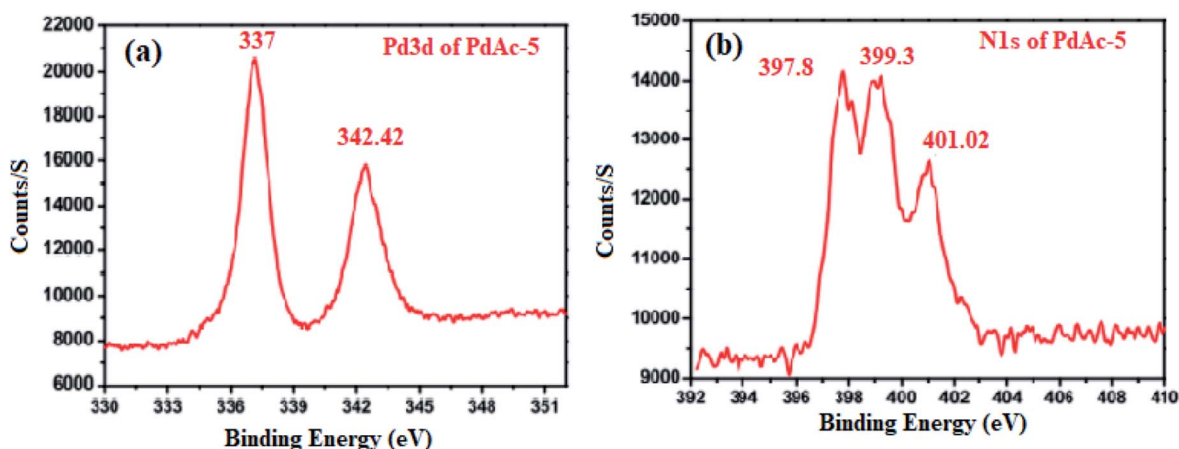


Fig. 2 (a) X-ray photo-electron spectrum of P 3d; and (b) N 1s core level spectrum of **PdAc-5**.

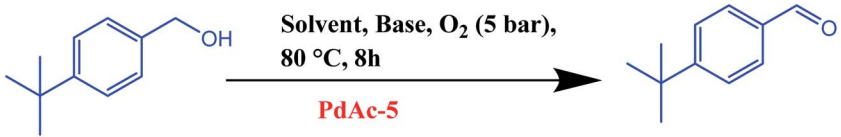
Table 1 Screening of oxidation system^a

				Conversion, %	
Entry	Solvent (mL)	Additive		(7) ^b	(8)
1	Toluene	None		NA	NA
2	Toluene	DBU (10 mg)		NA	NA
3	Toluene	K ₂ CO ₃ (10 mg)		NA	NA
4	Toluene	CH ₃ COONa (10 mg)		NA	NA
5	Acetonitrile	K ₂ CO ₃ (10 mg)		12.85%	NA
6	Acetonitrile	CH ₃ COONa (10 mg)		27.83%	NA
7	Water : acetic acid (1 : 1)	CH ₃ COONa (10 mg)		79.28%	6.5%
8	Water : acetic acid (1 : 1)	CH ₃ COONa (10 mg)		35.70% ^d	NA
9	Water : acetic acid (1 : 1)	CH ₃ COONa (10 mg)/DTBP ^c (50 mg)		38.87%	12.35
10	Water : acetic acid (1 : 1)	CH ₃ COONa (10 mg)/DTBP ^c (200 mg)		35.33%	36.22%

^a Reaction conditions: 4-butyl benzyl alcohol (**6**) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, 8 h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, –60 to 325 °C. ^c Di-tertiary butyl peroxide (aqueous solution: 70%). ^d Reduced catalyst (5 mg per gram of substrate).



Table 2 Screening of solvent system^a

				
Entry	Solvent (mL)	Catalyst	Additive	Conversion, % (7) ^b
1	Water : acetic acid (1 : 1)	1 mol%	CH ₃ COONa (5 mol%)	79.28%
2	Water : acetonitrile (1 : 1)	1 mol%	CH ₃ COONa (5 mol%)	74.8%
3	Water	1 mol%	CH ₃ COONa (5 mol%)	15.0%
4	Toluene	1 mol%	CH ₃ COONa (5 mol%)	10.46%
5	Acetic acid : toluene (1 : 1)	1 mol%	CH ₃ COONa (5 mol%)	27.93%
6	Water : acetic acid (1 : 1)	Nil	CH ₃ COONa (5 mol%)	NA

^a Reaction conditions: 4-butyl benzyl alcohol (6) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, 8 h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C.

of **PdAc-5** is slightly positively shifted by 1.1 eV. Finally, these results clearly showed that Pd(OAc)₂ was strongly coordinated with the bipyridine nitrogen atom as well as the palladium being present in +2 oxidation states. In addition to that, Pd was incorporated with the bipyridine ligand, which was easily identified by the ¹H NMR and ¹³C NMR spectra of ligand **4** and **PdAc-5**. Here, ligand **4** and **PdAc-5** bipyridine ligand 'N' adjacent to the aromatic proton appeared at 9.12–9.18 ppm, and also the acetate carbonyl and methyl protons appeared at 182.01 and 2.73 ppm in ¹H NMR and ¹³C NMR spectra, respectively (Table S1†).¹¹

To start with, we scrutinised various oxidation condition reactions with **PdAc-5** and oxygen gas using different solvents and additives.

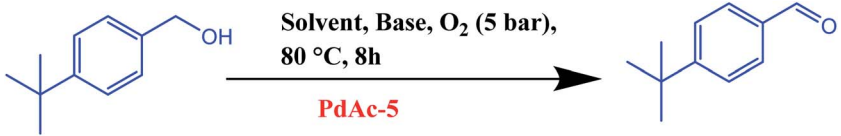
The initial entries in Table 1 (entries 1–4) reflect the fact that there is no product formation using toluene as a solvent in the presence of various additives. The solvent was shifted from toluene to polar entities like acetonitrile, a water–acetic acid

mixture and water. The conversion are very poor using acetonitrile (entries 5 and 6) as a solvent in the presence of an additive (K₂CO₃ or CH₃COONa), but appreciable conversion was observed using a water–acetic acid mixture (1 : 1) and sodium acetate as an additive (entry 7). The same system was explored using a reduced catalyst which resulted in poor conversion (entry 8). The above experiment was further explored using DTBP as a co-oxidant in the system, which resulted in the formation of acid **8** in substantial quantities apart from aldehyde **7** (entries 9 and 10).^{12–14}

In Table 2 we have evaluated the solvent system to carry out the oxidation process.

In this case, we found that the aqueous mixtures of acetic acid and acetonitrile (Table 2, entries 1 and 2) both gave better results than other systems like water, toluene, or an acetic acid : toluene mixture (entries 3–5), which yielded very poor conversions. The system was conducted in a water–acetic acid mixture without a catalyst, which yielded no conversion of the

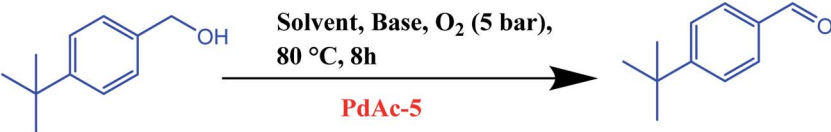
Table 3 Base screening for the oxidation reaction^a

			
Entry	Solvent (mL)	Additive (5 mol%)	Conversion, (7) ^b
1	Water : acetonitrile (1 : 1)	NaOAc	74.8%
2	Water : acetonitrile (1 : 1)	NaOH	40.0%
3	Water : acetonitrile (1 : 1)	Na ₂ CO ₃	67.1%
4	Water : acetonitrile (1 : 1)	K ₂ CO ₃	83.09%
5	Water : acetonitrile (1 : 1)	DBU	26.76%
6	Water : acetonitrile (1 : 1)	Triethanolamine	44.32%

^a Reaction conditions: 4-butyl benzyl alcohol (6) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, 8 h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100, 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C.

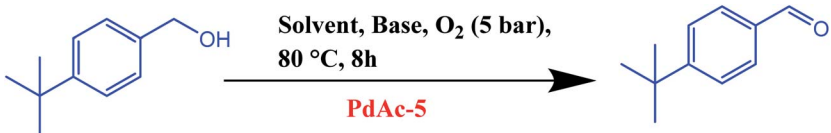


Table 4 Effect of catalyst and base quantity for the oxidation reaction^a

				
Entry	Solvent system (mL)	Catalyst	Additive	Conversion, (7) ^b
1	Water : acetonitrile (1 : 1)	2 mol%	K ₂ CO ₃ (5 mol%)	86.1%
2	Water : acetonitrile (1 : 1)	1 mol%	K ₂ CO ₃ (5 mol%)	83.09%
3	Water : acetonitrile (1 : 1)	0.5 mol%	K ₂ CO ₃ (5 mol%)	65.00%
4	Water : acetonitrile (1 : 1)	1 mol%	K ₂ CO ₃ (10 mol%)	84.00%
5	Water : acetonitrile (1 : 1)	—	K ₂ CO ₃ (10 mol%)	No reaction

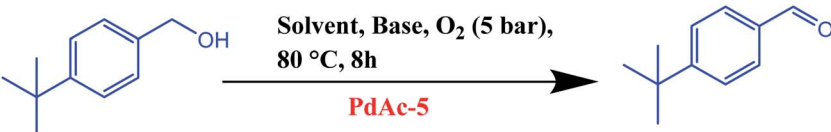
^a Reaction conditions: 4-butyl benzyl alcohol (**6**) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100, 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C.

Table 5 Solvent mixture screening for the oxidation reaction^a

		
Entry	Solvent mixture	Conversion, (7) ^b
1	Water : acetonitrile (1 : 1)	83.09%
2	Water	12.08%
3	Water : THF (1 : 1)	24.98%
4	Water : DMF (1 : 1)	17.82%
5	Water : acetone (1 : 1)	47.10%

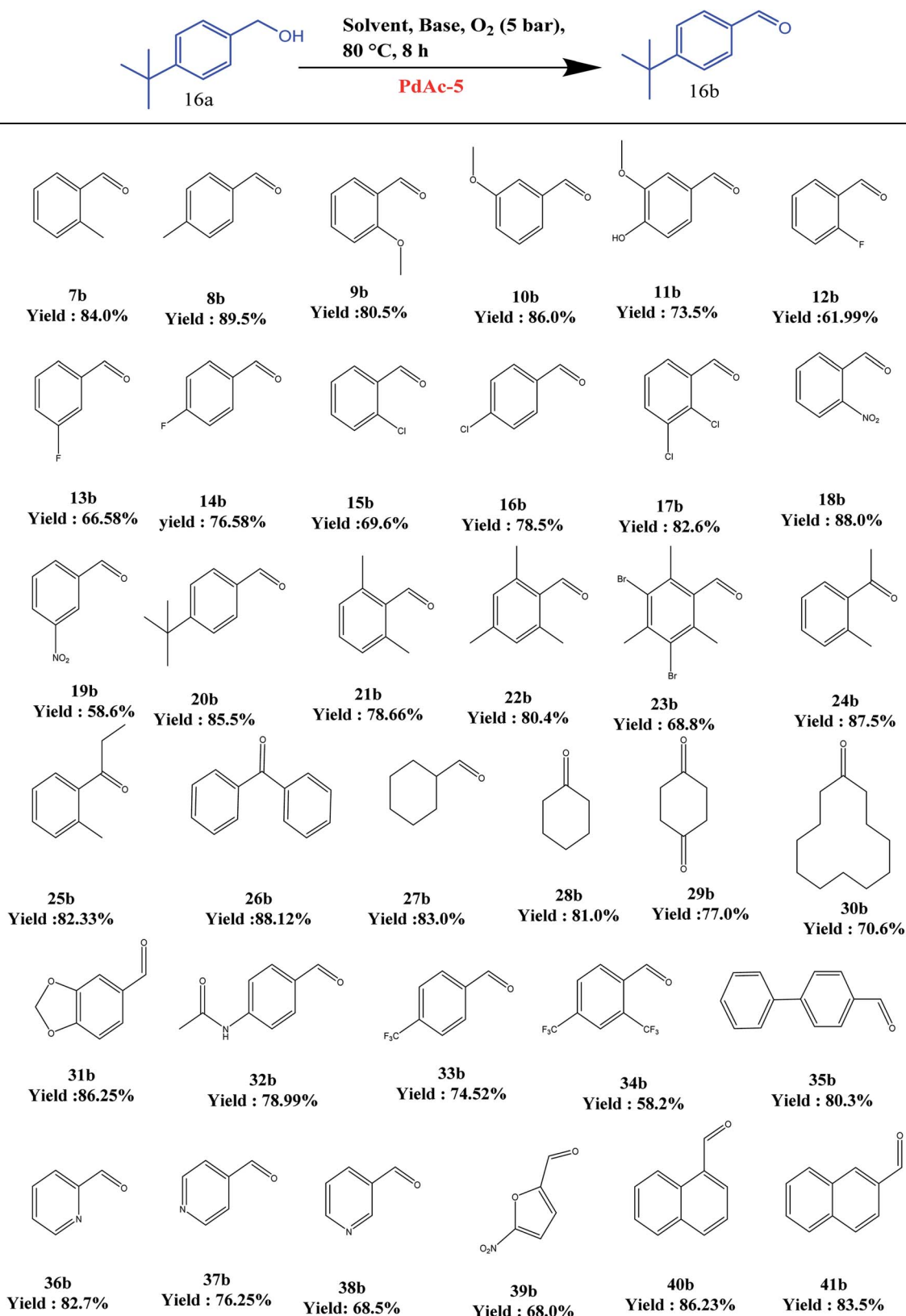
^a Reaction conditions: 4-butyl benzyl alcohol (**6**) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100, 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C.

Table 6 Effect of temperature of oxidation reaction^a

		
Entry	Temperature (°C)	Conversion, (7) ^b
1	25	4.08%
2	40	6.60%
3	60	31.30%
4	80	83.09%
5	125	75.90%

^a Reaction conditions: 4-butyl benzyl alcohol (**6**) (6.0 mmol), solvent (10 mL), **PdAc-5** (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, 8 h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100, 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C.



Table 7 Substrate screening for oxidation (1° and 2° alcohols) using PdAc-5^{a,b}

^a Reaction conditions: 4-tert-butyl benzyl alcohol (6) (6.0 mmol), solvent (10 mL), PdAc-5 (1 mol%), additive (5 mol%), oxygen (5 bar), 80 °C, 8 h. ^b Agilent GLC-GCMS: 6890 N, split ratio: 1 : 100, 250 °C, FID: 300 °C, flow rate: nitrogen 30 mL min⁻¹. Column: HP-5; 30 m × d: 0.32 mm × f: 0.25 mm, -60 to 325 °C. Yield of the product after column purification/distillation. All the products are matched against their reference standards.



substrate (entry 6). On the account of reactor stability, we opted to use a water–acetonitrile mixture as the solvent for further screening experiments rather than a water–acetic acid mixture (acetic acid corrodes the rubber gasket of the mini-clave at higher temperatures ($\text{pH} < 3$)).^{15–18}

The screening process studied the influence of various additives in our oxidation system. From Table 3, it is evident that additives like sodium acetate (entry 1) or potassium carbonate (entry 4) show better conversion than the other bases. Potassium carbonate is preferred over the former because of its alkaline pH (9–10), which ensures the stability of the reaction mass during the process. From the above and Table 3 we concluded that potassium carbonate in a water–acetonitrile mixture gave better conversion than other systems. On this accord, we also screened the above system with different base quantities (Table 4) and aqueous solvent mixtures (Table 5) which are shown below.^{17,18}

From Table 4 it is evident that a higher quantity of base (Table 4, entry 4) does not affect the conversion to a great extent. The same holds for an increase in the quantity of catalyst for the above reaction (entry 1). But conversion does drop with a lower catalyst loading (entry 3) which may be completed in a longer time period than usual.

From Table 5, we concluded that water–acetonitrile (Table 5, entry 1) works better than the other solvent mixtures (entries 2–4). The conversion is moderate in an acetone–water system (entry 5), but the rate of conversion is quite slow compared with a water–acetonitrile system. Last but not least, the reaction was also screened for its conversion rates at different temperatures as tabulated in Table 6.

As shown in Table 6, conversion is poor at lower temperatures (entries 1–3), and better conversion is observed in the temperature range of 80–90 °C (entry 4). Also, experiments were performed at much higher temperatures at which product decomposition was observed, giving rise to impurities and thus lowering the product ratio in the mass (entry 5). This may be due to the decomposition of **PdAc-5** along with the product.¹⁹

The above said catalyst was tried for different substrates for the oxidation reaction. In this regard, electron-donor substituents offered better yields but electron-withdrawing substituents ended up in poor conversions. As far as the secondary alcohols are concerned, the yields are better than for their primary counterparts (Table 7).

6. Catalyst recycling

To check the catalyst performance, the catalyst was recovered from the reaction mass, filtered and dried at 70 °C in a glass oven under vacuum. The dried catalyst was reintroduced into the new reaction system and then processed for the catalyst performance. The yields are quite similar for initial the 5 runs with a catalyst recovery of 90% per run (Fig. 3).^{20,21}

7. Scale-up study

The performance of the above-prepared catalyst on a commercial scale was checked using the oxidation of carveol (5-

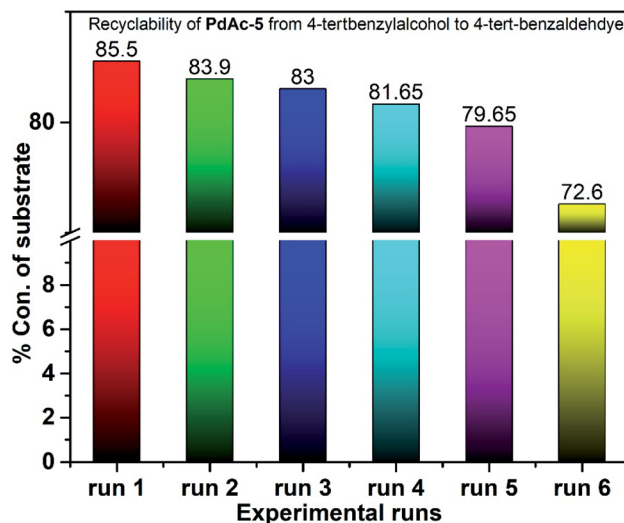


Fig. 3 Yields for catalyst recycling.

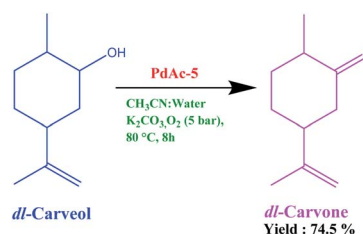


Fig. 4 Oxidation of carveol to carvone using **PdAc-5**.

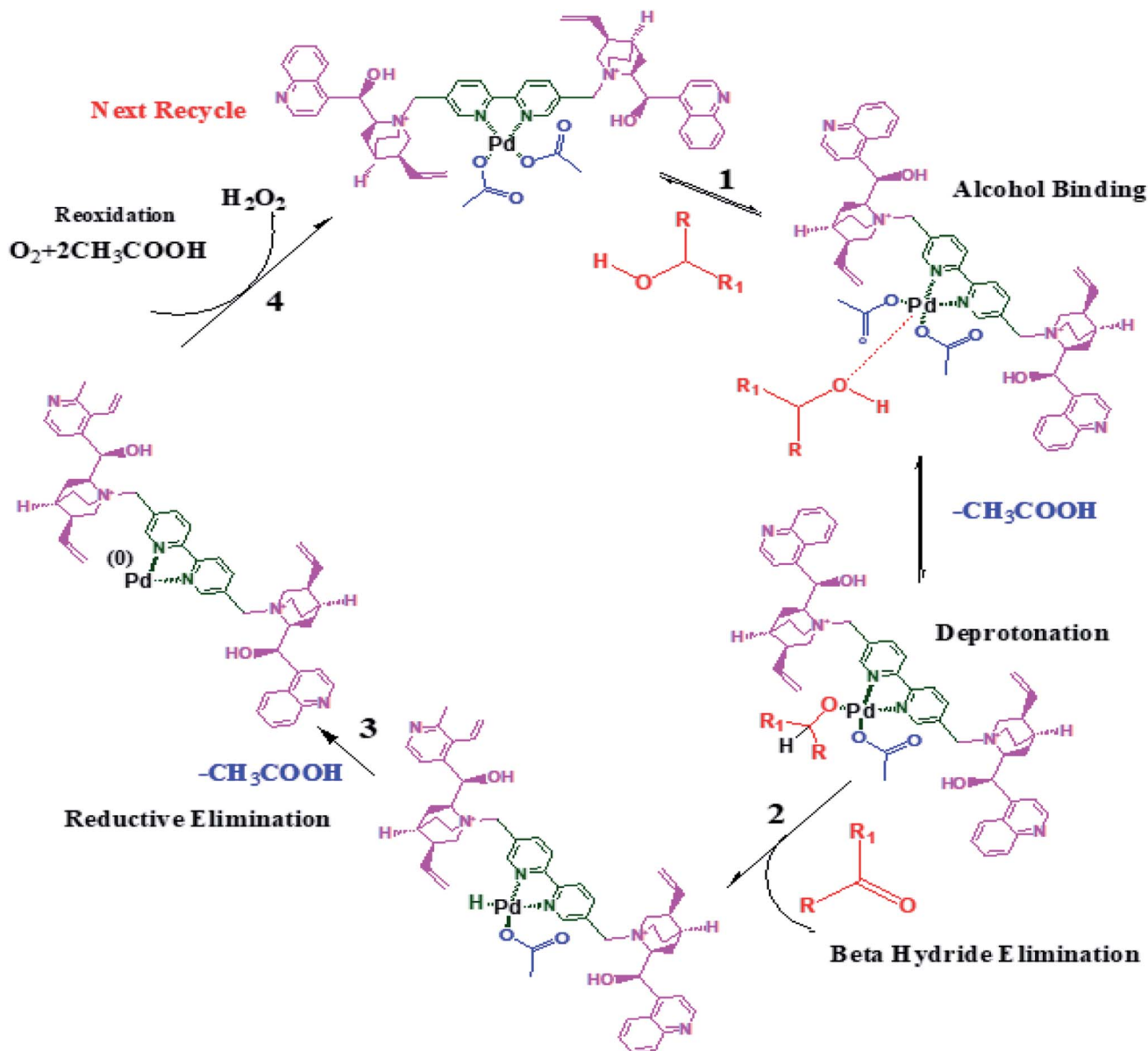
isopropenyl-2-methyl-cyclohexanol) to carvone (5-isopropenyl-2-methyl-cyclohexanone). As we know, there is a huge demand for synthetic carvone in the commercial market. The oxidation of carveol to carvone is currently performed using Oppenauer oxidation employing an aluminium isopropoxide and 2-butanone system on an industrial scale. We performed oxidation using **PdAc-5** under pressurised conditions using an acetonitrile–water system with potassium carbonate as base (Fig. 4). The isolated yield was 74.5% with 98.25% purity. This can be further enhanced using process engineering and process chemistry techniques.²²

8. Reaction mechanism

Based on the observed results, we proposed a suitable mechanism. The alcohol binds to the catalyst through oxidative addition, where it undergoes deprotonation followed by beta hydride elimination to yield the corresponding aldehyde/ketone. Thereafter, the catalyst undergoes a reductive elimination followed by redox transfer where it is regenerated for the next cycle (Scheme 2).²³

9. Conclusions

In summary, the organo-palladium catalytic system of **PdAc-5** has transformed primary and secondary alcohols into valuable



Scheme 2 Plausible mechanism for the oxidation of alcohols using PdAc-5.

aldehydes and ketones of industrial importance under an oxygen atmosphere with considerable yields. Despite our effort to achieve better results, some of the substrates were not good enough, and there is still much potential for the continued development of these reactions. The obtained products are quite appreciable in both yield and selectivity (70–85%). Concerning environmental and economic issues, the introduction of non-toxic, cheap, and at the same time efficient and universal catalyst systems, able to operate under mild reaction conditions in a highly selective manner and for a broad range of substrates, remains a challenge for future research.

Conflicts of interest

The authors declare no competing interest.

Acknowledgements

This work was supported by the National Research Foundation of Korea (NRF) grant funded by the Korea government (MSIP) No. 2017R1A2B2011730, No. 2018R1A6A1A03024509 and 2019R1A2C1011113.



References

- 1 (a) R. A. Sheldon and J. K. Kochi, *Metal Catalysed Oxidations of Organic Compound*, Academic Press, New York, 1981; (b) J. Mancuso and D. Swern, *Synthesis*, 1981, **1981**(3), 165–185; (c) K. Bowden, I. M. Heilbron, E. R. H. Jones and B. C. L. Weedon, *J. Chem. Soc.*, 1946, **3**, 39–45; (d) *Oxidation: Selectivity, Strategy & Efficiency in Modern Organic Chemistry*, ed. B. M. Trost, Elsevier, 1992; (e) *Oxidation of Alcohols to Aldehydes and Ketones, A Guide to Current Common Practice*, ed. G. Tojo and M. Fernández, Springer Science & Business Media, Inc., 2006; (f) J. E. Backvall, *Modern oxidation methods*, Wiley-VCH, Weinheim, Germany, 2nd edn, 2010.
- 2 (a) D. Daniel and C. Fabrizio, *Handbook of Advanced Methods and Processes in Oxidation Catalysis: From Laboratory to Industry*, 2014, vol. 24; (b) K. M. Gligorich and M. S. Sigman, *Chem. Commun.*, 2009, **14**(26), 3854–3867; (c) J. Liqun and L. Aiwen, *Sci. China: Chem.*, 2012, **55**, 2027.
- 3 (a) S. D. Pike and A. S. Weller, *Philos. Trans. R. Soc., A*, 2018, **373**, 2037; (b) C. Copéret, M. Chabanas, R. Petroff Saint-Arroman and J. M. Basset, *Angew. Chem., Int. Ed.*, 2003, **13**, 42; (c) N. R. Conley, L. A. Labios, D. M. Pearson, C. C. L. McCrory and R. M. Waymouth, *Organometallics*, 2007, **26**, 5447; (d) V. R. Gangwal, B. G. M. Van, W. Kuster and J. C. Schouten, *Chem. Eng. Sci.*, 2002, **57**, 5051; (e) D. C. Ebner, R. M. Trend, C. Genet, M. J. McGrath, P. O'Brien and B. M. Stoltz, *Angew. Chem., Int. Ed.*, 2008, **47**, 6367.
- 4 (a) I. A. Ansari and R. Gree, *Org. Lett.*, 2002, **4**, 1507; (b) T. Mallat and A. Baiker, *Chem. Rev.*, 2004, **104**, 3037; (c) P. Lahtinen, J. U. Ahmad, E. Lankinen, P. Pihko, M. Leskelä and T. Repo, *J. Mol. Catal. A: Chem.*, 2007, **275**, 228; (d) C. Keresszegi, T. Mallat and A. Baiker, *New J. Chem.*, 2001, **25**, 1163; (e) C. Keresszegi, T. Bürgi, T. Mallat and A. Baiker, *J. Catal.*, 2002, **211**, 244; (f) C. Parmeggiani and F. Cardona, *Green Chem.*, 2012, **14**, 547; (g) N. Gorgas, B. Stöger, L. F. Veiros and K. Kirchner, *ACS Catal.*, 2016, **6**, 2664; (h) J. F. Greene, J. M. Hoover, D. S. Mannel, T. W. Root and S. S. Stahl, *Org. Process Res. Dev.*, 2013, **17**, 1247.
- 5 (a) D. Könnig, F. D. Sypaseuth, C. C. Tzschucke and M. Christmann, *Chem. Commun.*, 2014, **50**, 5014; (b) G. Saravanan, N. Helfried, G. Dirk, T. Kerstin, J. Haijun and B. Matthias, *Chem.–Eur. J.*, 2013, **19**, 15979; (c) S. Abedi and A. Morsali, *New J. Chem.*, 2017, **41**, 5846; (d) Y. Li, J. Huang, X. Hu, F. Leung-YukLam, W. Wang and R. Luque, *J. Mol. Catal. A: Chem.*, 2016, **425**, 61–67; (e) J. Muzart, *Tetrahedron*, 2003, **59**, 5789; (f) S. E. Davis, M. S. Ide and R. J. Davis, *Green Chem.*, 2013, **15**, 17; (g) G. Mayan and R. Neumann, *Catal. Lett.*, 2008, **123**, 41; (h) K. Yamaguchi and N. Mizuno, *Chem.–Eur. J.*, 2003, **9**, 4353; (i) R. DiCosimo and G. M. Whitesides, *J. Phys. Chem.*, 1989, **93**, 768; (j) T. Matsumoto, M. Ueno, N. Wang and S. Kobayashi, *Chem.–Asian J.*, 2008, **3**, 196.
- 6 (a) B. Z. Zhan and A. Thompson, *Tetrahedron*, 2004, **60**, 2917; (b) B. A. Steinhoff and S. S. Stahl, *Org. Lett.*, 2002, **4**, 4179; (c) B. A. Steinhoff, I. A. Guzei and S. S. Stahl, *J. Am. Chem. Soc.*, 2004, **126**, 11268; (d) R. J. Nielsen and W. A. Goddard III, *J. Am. Chem. Soc.*, 2006, **128**, 9651; (e) H. Wu, Q. Zhang and Y. Wang, *Adv. Synth. Catal.*, 2005, **347**, 1356.
- 7 (a) C. Xu, C. Zhang, H. Li, X. Zhao, L. Song and X. Li, *Catal. Surv. Asia*, 2016, **20**, 13; (b) S. Zavahir, Q. Xiao, S. Sarina, J. Zhao, S. Bottle, M. Wellard, J. Jia, L. Jing, Y. Huang, J. P. Blinco, H. Wu and H. Y. Zhu, *ACS Catal.*, 2016, **6**, 3580; (c) P. A. Shapley, N. Zhang, J. L. Allen, D. H. Pool and H. C. Liang, *J. Am. Chem. Soc.*, 2000, **122**, 1079; (d) T. Nishimura, T. Onoue, K. Ohe and S. Uemura, *J. Org. Chem.*, 1999, **64**, 6750; (e) T. S. T. Komanoya, *Catal. Commun.*, 2009, **10**, 1095; (f) N. Kakiuchi, Y. Maeda, T. Nishimura and S. Uemura, *J. Org. Chem.*, 2001, **66**, 6620.
- 8 M. N. Kopylovich, A. P. C. Ribeiro, E. C. B. A. Alegria, N. M. R. Martins, L. M. D. R. S. Martins and A. J. L. Pombeiro, *Adv. Organomet. Chem.*, 2015, **63**, 91–174.
- 9 R. A. Fernandes, A. K. Jha and P. Kumar, *Catal. Sci. Technol.*, 2020, **10**, 7448–7470.
- 10 (a) B. Z. Zhan and A. Thompson, *Arabian J. Chem.*, 2004, **60**, 2917; (b) B. A. Steinhoff and S. S. Stahl, *Org. Lett.*, 2002, **4**, 4179; (c) B. A. Steinhoff, I. A. Guzei and S. S. Stahl, *J. Am. Chem. Soc.*, 2004, **126**, 11268; (d) R. J. Nielsen and W. A. Goddard III, *J. Am. Chem. Soc.*, 2006, **128**, 9651; (e) H. Wu, Q. Zhang and Y. Wang, *Adv. Synth. Catal.*, 2005, **347**, 1356.
- 11 (a) G. J. Brink, I. W. C. E. Arends and R. A. Sheldon, *Science*, 2000, **287**, 1636; (b) G. J. Brink, I. W. C. E. Arends and R. A. Sheldon, *Adv. Synth. Catal.*, 2002, **344**, 355; (c) T. Diao and S. S. Stahl, *J. Am. Chem. Soc.*, 2011, **133**, 14566.
- 12 J. F. Greene, J. M. Hoover, D. S. Mannel, T. W. Root and S. S. Stahl, *Org. Process Res. Dev.*, 2013, **17**, 1247.
- 13 G. Saravanan, N. Helfried, G. Dirk, T. Kerstin, J. Haijun and B. Matthias, *Chem.–Eur. J.*, 2013, **19**, 15979.
- 14 N. Decharin, V. P. Brian and S. S. Stahl, *J. Am. Chem. Soc.*, 2011, **133**, 13268.
- 15 (a) M. S. Kwon, N. Kim, C. M. Park, J. S. Lee, K. Y. Kang and J. Park, *Org. Lett.*, 2005, **7**, 1077–1079; (b) M. S. Kwon, N. Kim, S. H. Seo, I. S. Park, R. K. Cheedra and J. Park, *Angew. Chem., Int. Ed.*, 2005, **44**, 6913–6915; (c) M. S. Kwon, S. Kim, S. Park, W. J. Bosco, R. K. Cheedra and J. Park, *J. Org. Chem.*, 2009, **74**, 2877–2879.
- 16 (a) D. Daniel and C. Fabrizio, *Handbook of Advanced Methods and Processes in Oxidation Catalysis: From Laboratory to Industry*, 2014, **24**, p. 2005; (b) K. M. Gligorich and M. S. Sigman, *Chem. Commun.*, 2009, **14**, 3854; (c) J. Liqun and L. Aiwen, *Sci. China: Chem.*, 2012, **55**, 2027.
- 17 (a) P. K. Vijaya, S. Murugesan and A. Siva, *Tetrahedron Lett.*, 2015, **56**, 5209; (b) A. Siva, P. K. Vijaya, M. Mariyappan, V. Ashokkumar, V. Sadhasivam, S. Balakrishnan, C. Chinnadurai and S. Murugesan, *New J. Chem.*, 2017, **41**, 7980–7986; (c) C. Chinnadurai, P. K. Vijaya, M. Harikrishnan, N. Malini, M. Sepperumal and A. Siva, *Res. Chem. Intermed.*, 2021, **47**, 895–909.



- 18 R. R. Chidambaram, V. Sadhasivam, M. Mariyappan and A. Siva, *J. Iran. Chem. Soc.*, 2019, **16**, 373–384.
- 19 D.-W. Tan, J.-B. Xie, Q. Li, H.-X. Li, J.-C. Li, H.-Y. Li and J.-P. Lang, *Dalton Trans.*, 2014, **43**, 14061–14071.
- 20 Y. Yang, Z. Du, Y. Huang, F. Lu, F. Wang, J. Gao and J. Xu, *Green Chem.*, 2013, **15**, 1932–1940.
- 21 G. Zhang, X. Xie, Y. Wang, X. Wen, Y. Zhao and C. Ding, *Org. Biomol. Chem.*, 2013, **11**, 2947–2950.
- 22 A. M. Thomas, J. Peter, S. Nagappan, A. Mohan and C.-S. Ha, *Nanomaterials*, 2020, **10**, 2051.
- 23 G. Zhang, X. Wen, Y. Wang, X. Han, Y. Luan, L. Zheng, C. Ding and X. Cao, *RSC Adv.*, 2013, **3**, 22918–22921.

