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Active and passive adsorption of 47 trace atmospheric volatile organic compounds onto carbon nanotubes

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Carbon nanotubes (CNTs) of varying sizes and CNTs functionalised with carboxylic acids were examined by thermal desorption gas chromatography mass spectrometry (TD-GCMS) to determine the degree of surface contamination of atmospheric volatile organic compounds (VOCs). The CNTs could be purged of physisorbed VOCs by heating to 380 °C under a stream of purified nitrogen gas. As soon as the cleaned CNTs were exposed to atmospheric air they spontaneously adsorbed trace VOCs. As well as passive adsorption of VOCs, active sampling was carried out by pumping atmospheric air through the CNTs and comparing the results with the standard multisorbent materials Carboxipack X and Tenax that are used widely for VOC trapping and analysis. The CNTs were found to trap many VOCs at a comparable level to the standard sorbent materials. Therefore, to maintain the CNTs in a pristine condition, it is recommended that they are first heated under vacuum to remove residual physisorbed VOCs, and then stored under vacuum or in a purified inert gas atmosphere.

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Introduction

Carbon nanotubes (CNTs) have been proposed for use in many applications including wastewater treatment,¹ hydrogen adsorption,² and antimicrobial and anti-adhesive functions in medicine.³ These applications all rely on the specific surface properties of the CNTs promoting physisorption processes. Several studies have also shown that CNTs are excellent sorbents for organic species, and have been used in solid phase extraction, chromatographic analysis, and sensor development,^{4–23} although to the best of our knowledge there are few reports that examine adsorption of volatile organic compounds (VOCs) at their trace atmospheric level (ppb (v/v)).^{24–26} This leads to concerns whether gaseous VOCs that are naturally present in the atmosphere could spontaneously be introduced into the CNTs during their storage and transportation, thereby affecting their desirable surface features as well as influencing their toxicological properties. Adsorbed VOCs can potentially become involved in chemical reactions, generating unwanted side products, affecting reaction rates, or acquiring false positives or inaccurate data for CNTs when used in other applications. It has already been extensively proven that many of the proposed catalytic effects of CNTs are instead caused by trace metallic

impurities,^{27–29} therefore, trace organic adsorbents could also have deleterious effects on the surface properties of the CNTs.

Sensitive microscopy techniques such as SEM and TEM imaging of CNTs will not detect additional adsorbed organic molecules due to their small size and likely monolayer coverages. However, it is important to be able to identify any trapped organic impurities and interferences on CNTs after being subjected to exposure for a known duration of time under ambient conditions. There have been no studies aimed at determining the adsorption effects of naturally present atmospheric organic impurities at trace levels and comparing the results with standard multisorbents used for air quality testing, which have the highest levels of trapping performance and therefore, provide a good means of assessing the likely adulteration by unwanted VOCs. Therefore, in this study, gas chromatography mass spectrometry (GC-MS) coupled with thermal desorption (TD) trapping experiments were performed to determine the extent of contamination of CNTs with 47 trace atmospheric VOCs under passive and active adsorption conditions to determine which atmospheric VOCs are more readily adsorbed, with the results compared with the standard Carboxipack X/Tenax multisorbent. The study also investigated how the VOCs can be removed from the CNTs, and the required storage procedures for maintaining the CNTs in pristine condition.

Methods

Materials and chemicals

CNTs were purchased from Nanostructured & Amorphous Materials Inc. (Houston, Texas) and were purified using high

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temperature conditioning to remove adsorbed VOCs. The dimensions and purities of the CNTs as supplied by the manufacturer are given in Table 1. Thermogravimetric analysis (TGA), Raman spectroscopy and inductively coupled plasma-mass spectrometry (ICP-MS) characterisation of the CNTs have been reported previously.²⁴ VOC standards used for the analysis were obtained from Sigma-Aldrich (St Louis, USA), Merck (Hohenbrunn, Germany), Alfa Aesar (Heysham, Lancaster, UK) and Fluka (Buchs, Switzerland) with purities not less than 97%, except for the following compounds: 1,2,3-trimethylbenzene (93.9%) from Fluka and methacrolein (95%) from Sigma-Aldrich. The neat chemicals were diluted with the appropriate amount of methanol (Schedelco, Malaysia) to prepare (by serial dilutions) VOC standards solution for analysis.

Procedures and instrumentation for TD-GCMS experiments

Stainless steel sorbent tubes (89 mm length \times 6.4 mm outer diameter) that were pre-packed with 100 mg Carboxen 100 and 200 mg Tenax TA were obtained from Markes International Limited, Llantrisant, UK. The CNTs for analysis were packed into empty stainless sorbent tubes using a Markes Gauze Loading Rig and Plunger. Due to the higher volumes occupied by the CNTs, the mass of each CNT that could be packed into in the tube was less than for the standard Carboxen 100/Tenax multisorbent and was between 75–100 mg (Table 1). The CNTs inside the sorbent tubes were preconditioned to remove existing adsorbed VOCs by heating to 380 °C for up to 20 hours using a Markes tube conditioning dry purge unit (TC-20) under a 70 mL min⁻¹ flow of nitrogen gas. Thermal gravimetric analysis indicated that the CNTs did not undergo any decomposition at this temperature. Atmospheric air samples of the sorbent packed tubes were obtained using a calibrated SKC pocket air pump 210–1002, USA.

Thermal desorption experiments were conducted using a Markes UNITY Series 2 heating/trapping unit attached to a Markes Ultra TD-100 autosampler. The autosampler transported the sorbent containing tubes into the UNITY primary

desorption compartment that was joined in series to a Peltier cooled Tenax TA trap set at –10 °C, which was used for secondary desorption. Initial desorption involved heating the sorbent containing tubes to 375 °C for CNT sorbents or 280 °C for the Carboxen 100/Tenax multisorbent for 10 minutes under a 45 mL min⁻¹ flow of helium (99.999%), causing the desorbed VOCs to be transferred onto the Tenax TA cold trap. Secondary desorption involved heating the pre-concentrated VOCs on the Tenax TA trap to 300 °C for 7 minutes under a reversed flow of 6 mL min⁻¹ helium *via* 1/6 split mode to transfer the analytes into an Agilent 7890A gas chromatograph.

For separation of the analytes, an Agilent J & W DB-VRX (122–1564, 260 °C, 60 m \times 250 μ m \times 1.4 μ m, 1219.45766) column was utilised with helium carrier gas at a flow rate of 1.5 mL min⁻¹. The oven temperature for the GC was maintained at 30 °C for 12 minutes, increased to 60 °C over one minute, then increased at a rate of 40 °C per minute until 124 °C. The oven was isothermal at 124 °C for 2 minutes then raised at 9 °C per minute up to 200 °C and then isothermal at 200 °C for a further 3 minutes. Compounds exiting the column were detected using a 70 eV electron impact ionization Agilent Inert 5975C mass spectrometer with the ion source isothermal at 230 °C and the quadrupole mass analyzer maintained at 150 °C. Scans were performed over a mass range of 35–350 amu utilising scan mode. Origin Pro 8.1 and Microsoft Excel for Microsoft 365 MSO were used for statistical calculations.

Results and discussion

Passive sampling of VOCs

Experiments were first conducted to cleanse the CNTs of pre-adsorbed VOCs using 5 sorbent tubes each containing a nanomaterial: a multi-walled CNT (MWCNT), a carboxylated derivative of the MWCNT (COOH-MWCNT), a single-walled CNT (SWCNT), a functionalized SWCNT (COOH-SWCNT) and a SWCNT with shorter length (sSWCNT). The physical dimensions and residual organic artefacts of the commercially available nanomaterials are summarized in Table 1. During the

Table 1 Properties, masses, and organic artefacts present for CNTs packed into thermal desorption tubes^a

Sorbent	Reported CNT physical dimensions and surface areas	Reported CNT purity/%	Reported wt% COOH	TD tube packed mass/mg	Residual artefacts present after thermal conditioning/ng		
					Benzene	Toluene	Hexane
sSWCNT	Diameter: 1–2 nm Length: 1–3 μ m Surface areas: 300–380 m ² g ⁻¹	>95	N.A.	75	1.07	0	0.17
SWCNT	Diameter: 1–2 nm	>90	N.A.	75	1.41	0.11	0.81
COOH-SWCNT	Length: 5–30 μ m Surface areas: 300–380 m ² g ⁻¹	>95	2.59–2.87	75	0.67	0	0.13
MWCNT	Outer diameters: 50–80 nm Inner diameters: 5–15 nm	95	N.A.	100	1.71	0	2.20
COOH-MWCNT	Length: 10–20 mm Surface areas: 60–80 m ² g ⁻¹	95	0.47–0.51	100	1.78	0	0.52

^a N.A. = not applicable.



optimization of the CNT conditioning procedures to achieve acceptable CNT blanks prior to exposure, the total ion current (TIC) chromatograms revealed that multiple VOCs were adsorbed during prolonged storage and required numerous hours of thermal conditioning to be desorbed from the CNT materials.

It was found that a temperature of 380 °C under a constant flow of purified nitrogen gas for 20 hours was required to optimally clean the CNTs of residual VOCs. However, benzene

and hexane were found in all CNT blank chromatograms and toluene was also detected in the SWCNT sorbents even after optimum conditioning of the sorbent tubes (Table 1). This indicates that these molecules are likely to be always present as artefacts in the CNTs and cannot be removed through the high temperature conditioning processes. The carbon-containing compounds benzene and toluene are employed throughout catalytic CVD processes to produce CNTs consisting of

Table 2 GC-MS data on the absence (X) and presence (✓) of atmospheric VOCs adsorbed on different CNT sorbents that had been left inside thermal desorption tubes in ambient air for 72 hours

Target analytes	Qualifier ions ^a		<i>t_R</i> ^b (min)	VOCs detected in CNT sorbents				
	Q1	Q2		MWCNT	COOH-MWCNT	SWCNT	COOH-SWCNT	sSWCNT
Isopropyl alcohol	43 (17)	59 (5)	8.21	X	X	✓	X	X
Ethyl ether	45 (65)	73 (12)	8.8	✓	✓	✓	✓	✓
Isoprene	68 (69)	53 (54)	9.11	✓	✓	✓	✓	✓
Dichloromethane	49 (90)	86 (65)	10.27	✓	✓	✓	✓	✓
2-Methylpentane	43 (100)	42 (53)	13.01	✓	✓	✓	✓	✓
Methacrolein	41 (84)	39 (73)	13.25	✓	✓	✓	✓	✓
3-Methylpentane	56 (87)	41 (52)	13.63	✓	✓	✓	✓	✓
Hexane	41 (60)	43 (51)	14.21	✓	✓	✓	✓	✓
2-Butanone	43 (100)	57 (8)	14.26	X	X	✓	X	✓
Trichloromethane	85 (67)	47 (17)	14.7	✓	✓	✓	✓	✓
Ethyl acetate	61 (19)	70 (15)	14.79	X	X	✓	✓	✓
Methylcyclopentane	69 (48)	41 (42)	15.05	✓	✓	✓	✓	✓
Cyclohexane	56 (95)	41 (43)	15.98	✓	✓	✓	✓	✓
Benzene	77 (22)	51 (12)	16.16	✓	✓	✓	✓	✓
Heptane	43 (100)	57 (64)	16.89	✓	✓	✓	✓	✓
Trichloroethylene	132 (97)	134 (31)	16.97	✓	✓	✓	✓	✓
Methyl methacrylate	41 (85)	39 (46)	17.27	X	X	X	X	X
Methyl cyclohexane	55 (61)	98 (46)	17.58	✓	✓	✓	X	✓
Methyl isobutyl ketone	58 (48)	85 (25)	17.95	X	X	X	X	X
Pyridine	52 (47)	51 (21)	18.1	X	X	X	X	X
2-Methylheptane	43 (78)	70 (26)	18.37	✓	✓	✓	✓	✓
Toluene	92 (64)	65 (10)	18.7	✓	✓	✓	✓	✓
1-Octene	41 (77)	70 (90)	18.88	X	X	X	✓	✓
Octane	85 (71)	57 (49)	19.04	✓	✓	✓	✓	✓
Hexanal	57 (71)	72 (33)	19.14	✓	✓	✓	✓	✓
Tetrachloroethylene	164 (77)	129 (65)	19.58	✓	✓	✓	✓	✓
Furfural	95 (91)	39 (33)	19.95	X	X	✓	✓	✓
Ethylbenzene	106 (38)	77 (8)	20.63	✓	✓	✓	✓	✓
<i>m,p</i> -Xylene	106 (56)	77 (12)	20.86	✓	✓	✓	✓	✓
Nonane	43 (91)	85 (48)	21	X	X	✓	✓	✓
Heptanal	55 (66)	57 (55)	21.16	X	X	✓	✓	✓
Styrene	103 (46)	78 (37)	21.29	✓	✓	✓	✓	✓
<i>o</i> -Xylene	106 (54)	105 (21)	21.39	✓	✓	✓	✓	✓
Phenol	66 (24)	65 (20)	22.38	✓	✓	✓	✓	✓
3-Ethyltoluene	120 (42)	91 (14)	22.6	✓	✓	✓	✓	✓
4-Ethyltoluene	120 (39)	91 (12)	22.68	✓	✓	✓	✓	✓
Benzaldehyde	106 (97)	77 (87)	22.74	X	X	✓	✓	✓
1,3,5-Trimethylbenzene	120 (62)	91 (11)	22.85	✓	✓	✓	✓	✓
Decane	43 (74)	71 (45)	22.9	✓	✓	✓	✓	✓
2-Ethyltoluene	120 (42)	91 (13)	23.05	✓	✓	✓	✓	✓
Octanal	43 (94)	57 (94)	23.13	X	X	X	✓	X
Benzonitrile	76 (32)	50 (10)	23.18	✓	✓	✓	✓	✓
1,2,4-Trimethylbenzene	120 (59)	91 (11)	23.41	✓	✓	✓	✓	✓
1,2,3-Trimethylbenzene	120 (51)	91 (10)	24.03	✓	✓	✓	✓	✓
Acetophenone	77 (66)	120 (27)	24.82	✓	✓	✓	X	✓
Nonanal	41 (70)	70 (40)	25.03	X	X	X	X	X
Decanal	41 (81)	70 (58)	27.03	X	X	X	X	X

^a Identity and () relative abundance of qualifier ions. ^b Retention times of the analytes.



graphene sheets.³⁰ These reagents may remain as residual impurities within the synthesised CNTs or may have been produced during the elevated temperature conditioning process. Carboxylated CNTs were found to contain lesser quantities of hexane in comparison to the non-carboxylated analogues, with the most hexane detected in the MWCNT. This is possibly associated with hexane having a higher isosteric heat of adsorption on non-polar (*i.e.*, non-carboxylated) CNTs, that results in stronger adsorbent-adsorbate interactions,^{31,32} making it more difficult to desorb hexane from the non-functionalized of CNTs.

The thermally conditioned CNTs inside the sorbent tubes were then placed uncapped on the bench in an analytical laboratory and left exposed to atmospheric air for 72 hours. TD-GCMS analysis was performed for all CNT tubes to qualitatively identify the VOCs that were adsorbed. Mass spectrums of 47 VOC standards were used to determine the relative abundance of qualifier and quantifier ions with respect to the base ion for each compound. The 47 standards were selected based on pre-existing knowledge of background VOCs present in the atmosphere in Singapore.³³ Qualitative identification was performed by matching the relative abundance of qualifier ions and retention times (t_R) of unknowns to the standards. Table 2 summarizes the presence or absence of VOCs that were adsorbed onto the different CNTs, their t_R 's and the relative abundance ratios of various ions with respect to the base ion. Representative chromatograms of the desorbed contents of the sorbent tubes after 72 hours of ambient air exposure are shown in Fig. 1.

Many VOCs in the laboratory air were found to be adsorbed on the CNT materials during the 72 hours of exposure. A total of 33 VOCs were detected in the MWCNTs and between 37 to 40 compounds detected in the SWCNTs. The most visible signals present in all chromatograms belonged to 2-methylpentane, 3-methylpentane, hexane, benzene and toluene. VOCs that were adsorbed on some but not all CNTs were generally alkenes, carbonyl compounds and alcohols except for methyl cyclohexane and nonane.

The preliminary data showed that many organic compounds were retained on the CNTs during exposure to air. As a result, they could potentially participate in chemical reactions. With the garnering interest in utilizing CNTs as a reaction vessel or as a catalyst-support in chemical reactions, it is important to investigate methods of proper containment, transport and purification to simultaneously reduce inorganic and organic contaminants prior to their actual application.^{34,35}

Active sampling of VOCs

The ability of CNTs to adsorb and desorb gaseous organic compounds during the exposure to ambient air demonstrates the potential of these nanomaterials for trapping these molecules *via* passive diffusion. To evaluate the possibility of active sampling, air samples were collected using the tube containing the SWCNTs because it showed the highest amounts of VOCs trapped amongst all the nano-sorbents (Fig. 1 and Table 2). A conventional tube containing the widely used commercial Carpack X/Tenax multisorbent was utilized with which to

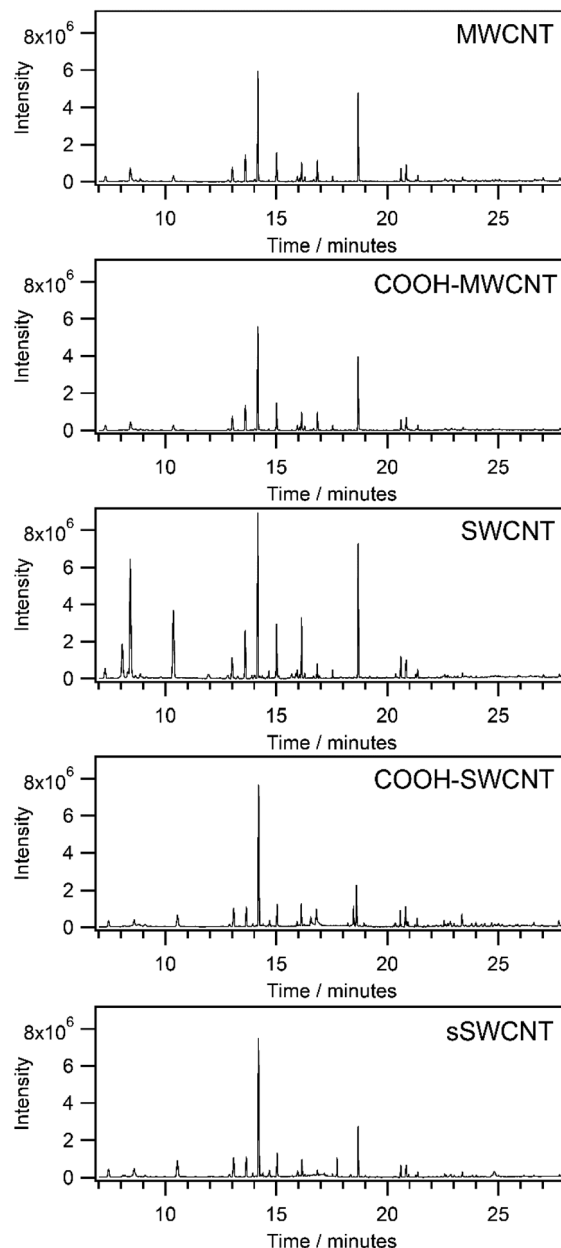


Fig. 1 Total ion current chromatograms of atmospheric VOCs adsorbed on different CNT sorbents that had been left inside thermal desorption tubes in ambient air for 72 hours.

compare the performance of the SWCNTs for trapping VOCs. Each sorbent tube was connected to a portable air pump and placed on the rooftop of the School of Physical and Mathematical Sciences (SPMS) building in the host University. The air flow was calibrated to 20 mL min⁻¹, with 2.4 L of air acquired after 2 hours with the sorbent tubes and then analysed using TD-GCMS. The overlaid total ion current chromatograms of the 2 sorbent tubes after sampling are shown in Fig. 2. Peak area ratios for each VOC detected were calculated using eqn (1) and are summarized in Table 3 to enable a comparison of the trapping of VOCs by the SWCNTs with respect to the standard commercially available Carpack X/Tenax. Ratios > 1 for eqn



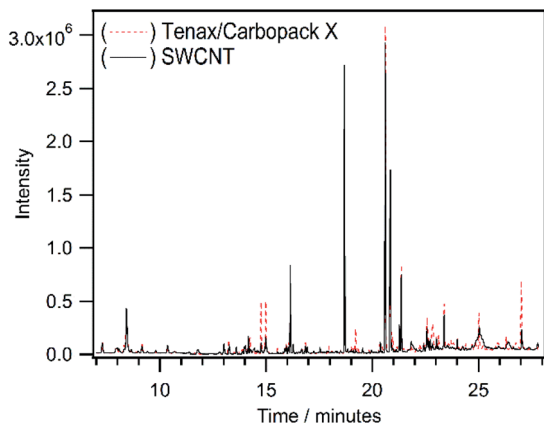


Fig. 2 Total ion current chromatograms of atmospheric VOCs adsorbed separately on SWCNTs and Carbopack X/Tenax multi-sorbents that were packed in thermal desorption tubes, with 2.4 L of atmosphere air passed through the sorbents over 2 hours.

(1) indicate that the SWCNTs are more efficient than the Carbopack X/Tenax multisorbent in trapping the particular VOC.

$$(\text{Peak area ratio})_{\text{VOC}} = \frac{(\text{VOC peak area})_{\text{CNT}}}{(\text{VOC peak area})_{\text{Tenax/Carbopack X}}} \quad (1)$$

10 of the targeted VOCs were absent from both sorbent tubes during sampling, due to their not being present in the atmosphere at that particular time (or below the detection limits). 37 VOC target analytes were identified in the multisorbent Carbopack X/Tenax tube sample whereas 29 VOC analytes were present in the SWCNT sample tube. Out of the 29 analytes identified in the SWCNT tube, only 7 VOCs (butanone, decanal, ethyl acetate, heptane, octane, nonane and 1,2,4-trimethylbenzene) had peak area ratios <0.7, indicating that most VOCs were favourably retained in comparison to the standard multisorbent material. Fig. 3 shows the quantifier ion peak area of selected target compounds in both sorbent tubes plotted corresponding to their primary functional groups.

VOCs that were found in the Carbopack X/Tenax multi-sorbent tube but not in the SWCNT sorbent tube are isopropyl alcohol, isoprene, benzene, nonanal, octanal, hexanal, methyl isobutyl ketone and acetophenone. The results from sampling agree with the functional group trends observed previously during the loading of the VOC standards onto the CNT materials,²⁴ except for benzene and 1,2,4-trimethylbenzene. The discrepancies between the benzene ratios in sample tubes and in VOC standards tubes were deemed to be an artifact interference error. Benzene is inherently generated from both sorbent materials during heating and the amount of benzene detected from atmospheric sampling was very low. After background benzene subtraction, zero was obtained on the CNT tube while very low signal intensity was acquired from the Carbopack X/Tenax tube.

Variations in the atmospheric temperature could possibly be the explanation for the lower peak area ratio of 1,2,4-trimethylbenzene during sampling. Although there is no breakthrough of 1,2,4-trimethylbenzene in the SWCNT during the

Table 3 Normalized peak area ratio of target analytes detected in SWCNT sorbent tube after collecting 2.4 L of air sample at the roof of the SPMS building^a

Functional group	VOC	Normalised peak area ratio
Alcohol	Isopropyl alcohol	0
Ether	Ethyl ether	N.A.
Alkene	Isoprene	0
Alkane	1-Octene	N.A.
	2-Methylpentane	1.22
	3-Methylpentane	1.11
	Hexane	1.32
	Methylcyclopentane	0.96
	Cyclohexane	1.08
	Heptane	0.66
	Methyl cyclohexane	1.07
	2-Methylheptane	N.A.
	Octane	0.36
	Nonane	0.22
	Decane	N.A.
	Halogenated alkanes	Dichloromethane
	Trichloromethane	N.A.
Halogenated alkenes	Trichloroethylene	1.18
	Tetrachloroethylene	1.13
Carbonyl compounds	2-Butanone	0.2
	Methyl isobutyl ketone	0
	Hexanal	0
	Heptanal	N.A.
	Octanal	0
	Nonanal	0
	Decanal	0.33
	Ethyl acetate	0.22
Aromatic compounds	Benzene	0
	Toluene	1.1
	Ethyl benzene	0.96
	<i>p,m</i> -Xylene	0.96
	<i>o</i> -Xylene	0.92
	2-Ethyltoluene	0.83
	3-Ethyltoluene	0.8
	4-Ethyltoluene	0.77
	1,3,5-Trimethylbenzene	1.07
	1,2,4-Trimethylbenzene	0.64
	1,2,3-Trimethylbenzene	0.73
Vinyl carbonyls	Methacrolein	3.17
	Methyl methacrylate	N.A.
Aromatic ketones	Acetophenone	0
Aromatic aldehydes	Benzaldehyde	1.1
Vinylbenzenes	Styrene	2.98
Hydroxybenzenes	Phenol	1.36
Heterocyclic compounds	Pyridine	N.A.
	Furfural	N.A.
Cyanobenzenes	Benzonitrile	N.A.

^a 0 represents not detected in the SWCNT while N.A. represents the absence in both SWCNT and Carbopack X/Tenax.

loading of standards,²⁴ it may occur during sampling when temperature in the atmosphere is sufficiently high.

Better recoveries were observed for one third of the VOCs detected in the SWCNTs when the peak areas were compared to the conventional sorbent during sampling (Table 3 and Fig. 3). Methacrolein had the highest ratio, and the signal response was approximately 3.2 times greater than the conventional multi-

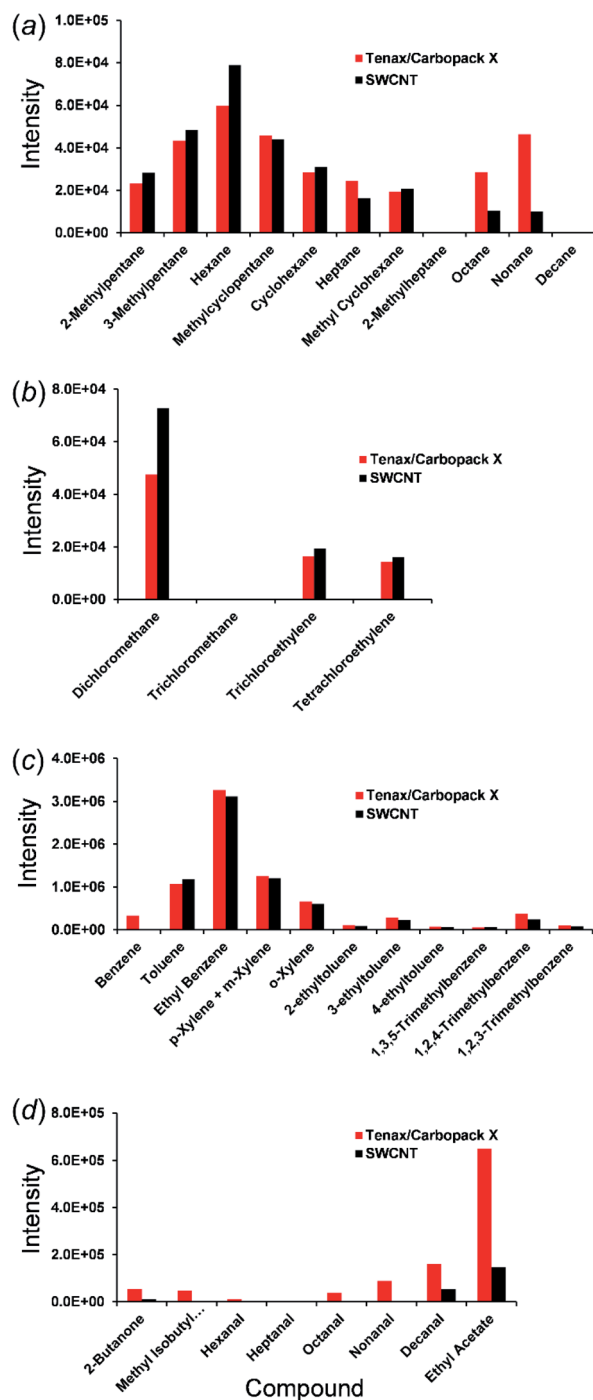


Fig. 3 Comparisons of the quantifier ion peak area of selected VOC signals in SWCNT and Carpack X/Tenax multisorbent, arranged according to their functional groups. (a) Saturated hydrocarbons, (b) saturated and unsaturated halides, (c) aromatic hydrocarbons, and (d) carbonyl compounds.

sorbent material. Styrene was next and had a peak area 3.0 times higher than found from desorption of VOCs from the Carpack X/Tenax. Dichloromethane had a peak abundance that was 1.5 times higher when using the SWCNT sorbent tube for active sampling. Other compounds having peak area ratios >1.1 include phenol, hexane and 2-methylpentane.

Statistical tests using the Carpack X/Tenax sorbent demonstrated excellent reproducibility with relative standard deviation (% RSD) values <10%, recoveries between 60% and 120%, tube desorption efficiencies close to 100% and excellent breakthrough values 5% or lower.³³ The CNT sorbents have been shown to have % RSD below 25% for most VOCs, which are acceptable values required by the USEPA for environmental sampling.²⁴ Similarly, breakthroughs and recoveries for most of the VOCs were acceptable for the CNTs.

For the 29 VOCs that were detected in both sorbent materials, the paired *t*-test for comparing individual differences to determine if the data were significantly different (eqn (2)). In eqn (2), *d* is the absolute value of the mean of the difference, *s_d* is the standard deviation of the difference, *n* is the number of VOCs, and *t_{cal}* is the calculated comparison value. The *t_{cal}* value was found to be 1.549, which indicates that the results obtained for the CNTs and Carpack X/Tenax sorbents are not significantly different at confidence levels of 90% and greater.

$$t_{cal} = \frac{|d|}{s_d} \sqrt{n} \quad (2)$$

The multi-sorbent Carpack X/Tenax still remains a superior general sorbent for atmospheric sampling when compared to the CNTs examined in this study since it can retain and desorb a higher number of VOCs. However, it is difficult to exactly quantify the sampling abilities of the CNTs with respect to the Carpack X/Tenax multisorbent because different amounts of the sorbent materials were used in the tests, due to the CNTs being considerably less dense than the Carpack X/Tenax. The humidity can also influence the adsorption abilities of the sorbent, but the effects are likely to be relatively small since CNTs are hydrophobic. Nevertheless, it is clear from the peak area ratios obtained from the chromatograms that many of the VOCs are retained and desorbed on the CNTs as efficiently as the multisorbent standard material.

Conclusions

Many studies have focused on the role that metallic impurities have on the catalytic properties of carbonaceous nanomaterials, often concluding that the enhanced properties of the nanomaterials are entirely due to the impurities. This study has shown that CNT sorbents passively exposed to a chemistry laboratory environment for 72 hours spontaneously absorbed a large number of trace atmospheric VOCs. Similarly, the active sampling of outdoor air using SWCNTs in stainless steel sorbent tubes indicated that they retained and desorbed many VOCs at a favorable level to the standard Carpack X/Tenax multisorbent that is used for sampling VOCs in air quality monitoring studies. Nevertheless, the Carpack X/Tenax multisorbents can still be considered superior to the CNTs since they are able to adsorb, and then release after heating a larger number of VOCs. Under atmospheric active trace sampling conditions, the CNTs are considerably less efficient than the Carpack X/Tenax at trapping a range of oxygenated compounds including isopropyl alcohol, isoprene, nonanal,



octanal, hexanal, methyl isobutyl ketone and acetophenone, in addition to benzene and isoprene.

Since the surface properties of the nanomaterials are considered essential to their applied functions, it is crucial to review the appropriate methods for eliminating organic impurities in purified CNTs prior to their applications and for transporting between apparatus. Although this study has been restricted to CNTs, it is highly likely that other nanoscale carbon-based materials such as graphene and its reduced/oxidized forms also suffer from extensive surface contamination from atmospheric VOCs. CNTs can be cleansed of VOCs by heating to 380 °C under a flow of an inert gas or under vacuum conditions. The cleaned CNTs need to be stored in an inert atmosphere otherwise they will immediately adsorb many atmospheric VOCs at a rate similar to materials utilized specifically as sorbents for atmospheric analysis.

Author contributions

GKSW and RDW conceptualised the project. GKSW conducted the investigation and developed the methodology. Both authors conducted formal analysis of the data. RDW supervised and managed the project.

Conflicts of interest

There are no conflicts to declare.

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