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A dual polymer composite of poly(3-hexylthiophene) and poly(3,4-ethylenedioxythiophene) hybrid surface heterojunction with g-C₃N₄ for enhanced photocatalytic hydrogen evolution†

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A surface heterojunction catalyst of g-C₃N₄-PEDOT/P3HT with P3HT and PEDOT as the polymer sensitizer and hole transport pathway is successfully prepared. The as constructed g-C₃N₄-PEDOT/P3HT composite exhibits a photocatalyst H₂ evolution rate up to 427703.3 μmol h⁻¹ g⁻¹ which is 1059 times higher than that of g-C₃N₄, 118 times higher than that of g-C₃N₄-PEDOT with ascorbic acid as sacrificial reagents. What's more, the g-C₃N₄-PEDOT/P3HT can even show an obviously enhanced photocatalytic H₂ evolution rate which is 6.1 times higher than that of pure g-C₃N₄ in pure water without any sacrificial reagent. Combining the experimental results and molecular dynamic (MD) simulation results, a possible mechanism can be drawn that the existed PEDOT possesses relatively higher hole mobility and can be used as a hole conductor between g-C₃N₄ and P3HT. Then, the photogenerated holes migration can be accelerated by PEDOT from the VB of g-C₃N₄ to the VB of P3HT. All those factors may benefit the synergy among g-C₃N₄, PEDOT and P3HT, which finally facilitates the rapid migration of photoinduced electron-hole pairs and eventually improves the photocatalytic H₂ activity process of g-C₃N₄-PEDOT/P3HT with visible light. The present work may provide useful insights for designing a surface heterojunction composite photocatalyst with high photocatalytic activity for H₂ production.

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Introduction

As a metal-free semiconductor photocatalyst, graphite carbon nitride (g-C₃N₄) has attracted abundant attention as a visible light induced water splitting catalyst for hydrogen production.¹ However, the visible light photocatalytic activities of pure g-C₃N₄ are still restricted by the relatively low conductivity of electrons and strong recombination of photogenerated electrons-holes.^{2,3} Much effort has been devoted to addressing this issue, like element doping,⁴⁻⁶ constructing heterojunctions with another narrowed band-gap semiconductor,⁷⁻⁹ and sensitization by organic dyes,¹⁰⁻¹³ to enhance the visible light driven photoactivity of g-C₃N₄.

Within these mentioned strategies, the heterojunction design was a common and effective way, due to the easy

integration effect of the different semiconductor components.¹⁴⁻¹⁸ However, previous research mainly focused on the inorganic-inorganic heterojunction or inorganic-polymer heterojunction.¹⁹⁻²¹ Recently, the fully polymer-polymer composite based heterojunction systems, applied in the photocatalytic H₂ production, have received much attention due to the encouraging progress results.²²

Among which, some typical conducting polymers like poly(3-hexylthiophene) (P3HT) and polypyrrole (PPy), were used as polymer composite photocatalyst to improve the H₂ production activity of g-C₃N₄. For example, Yan *et al.* developed a P3HT-Pt/g-C₃N₄ composite heterojunction for H₂ production with Na₂S-Na₂SO₃ solution as sacrificial reagent.²¹ Then, they found that the H₂ production activity increased to 300 times when 3 wt% P3HT was added and the quantum yield about 2.9% at 420 nm. After that, the same group prepared a new heterostructure of g-C₃N₄/Au/P3HT/Pt by a self-assembling method.²³ The as-prepared heterojunction g-C₃N₄/Au/P3HT/Pt promoted visible light hydrogen production activity with optimal amount P3HT at 0.5 wt% and obtained the highest H₂ evolution rate 320 μmol h⁻¹. Recently, Peng and coworkers introduced a surface heterojunction (SHJ) photocatalyst of P3HT-Pt/g-C₃N₄. By majorization the composition of P3HT and photoreaction factors, an impressive apparent quantum yield was achieved.²⁴ Besides,

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Chen's and coworkers prepared the g-C₃N₄/PPy heterojunction nanoparticles and found that the H₂ evolution from pure water was achieved by loading very low weight ratio dispersed PPy nanoparticles.²⁵

However, as the rapid charge recombination always happens during the diffusion of charges on the photocatalyst particle surface after photo-excitation, which is severely limiting the overall photocatalytic efficiency. To further suppress the charge recombination, one of the effective way is to design photocatalysts with spatially separated oxidative and reductive sites.²⁶

Usually, co-catalysts on semiconductor can collect photo-generated carriers, the maintain active sites for catalytic H₂ or O₂ evolution, and thus improving the activity and stability of photocatalysts by suppressing the photogenerated carriers recombination.²² Currently, plenty of the developed photocatalytic systems use noble metal co-catalysts, such as Pt, Pd, Ag, as the sink of photoinduced electrons to achieve high H₂ production activity.²² However, the behavior of photogenerated hole was nearly ignored. As the photogenerated hole may result in the byproduct like O₂, and the sacrificial electron donor was commonly used to eliminate these unfavourable photo-generated hole. In fact, these photogenerated hole may recombine with the photoinduced electrons again before it was consumed by the sacrificial reagent.

Poly(3,4-ethylenedioxythiophene) (PEDOT) has been considered to be a perfect conducting polymer in optoelectronic devices, with its perfect conductivities, thermal stability and transparency in the visible range.²⁶ Specially, because of its high hole mobility, PEDOT is often used as a hole conductor in dye sensitized solar cells (DSSCs).^{27,28} Moreover, Chen *et al.* found that the PEDOT also could catalyse water oxidation reaction in the typical photoelectrochemical cells.²⁹ What's more, Zong and Wang *et al.* constructed a g-C₃N₄-PEDOT composite and PEDOT acted as a hole transport channel.²⁶ Then, the g-C₃N₄-PEDOT-Pt system was created, in which PEDOT and Pt acted as oxidation and reduction reaction sites, respectively. In this way, the as-prepared g-C₃N₄-PEDOT-Pt photocatalyst exhibited 4 times higher hydrogen activity than the conventional photocatalyst without PEDOT. However, the photocatalytic activities of pure g-C₃N₄-PEDOT was still limited due to its relatively large bandgap (2.7 eV), as the PEDOT it self could not used as the light absorber to excite photocatalytic H₂ production, resulting in an insufficient sunlight harvesting.

Based on above analysis, we wonder whether it is possible to construct a g-C₃N₄-PEDOT/P3HT surface heterojunction (SHJ) catalyst with P3HT and PEDOT as the polymer sensitizer and

hole transport pathway, respectively. In this paper, we report a new type of g-C₃N₄ SHJ photocatalyst by co-loading PEDOT and P3HT onto g-C₃N₄. Moreover, as Pt is often deposited to g-C₃N₄ as electron sink and proton reduction sites, we successfully construct a g-C₃N₄-PEDOT/P3HT/Pt composite, in which PEDOT act as hole transport pathway, P3HT as the polymer sensitizer and Pt as the reduction reaction sites. In this way, the g-C₃N₄-PEDOT/P3HT composite loaded with Pt exhibits the better visible light harvest character, synergy with spatially separated oxidative and reductive sites, and finally remarkably enhanced the photocatalytic H₂ evolution activity.

Results and discussion

Structure and morphology of photocatalyst

As shown in Scheme 1, we present the typical synthetic route of g-C₃N₄-PEDOT/P3HT composite. This process involves two

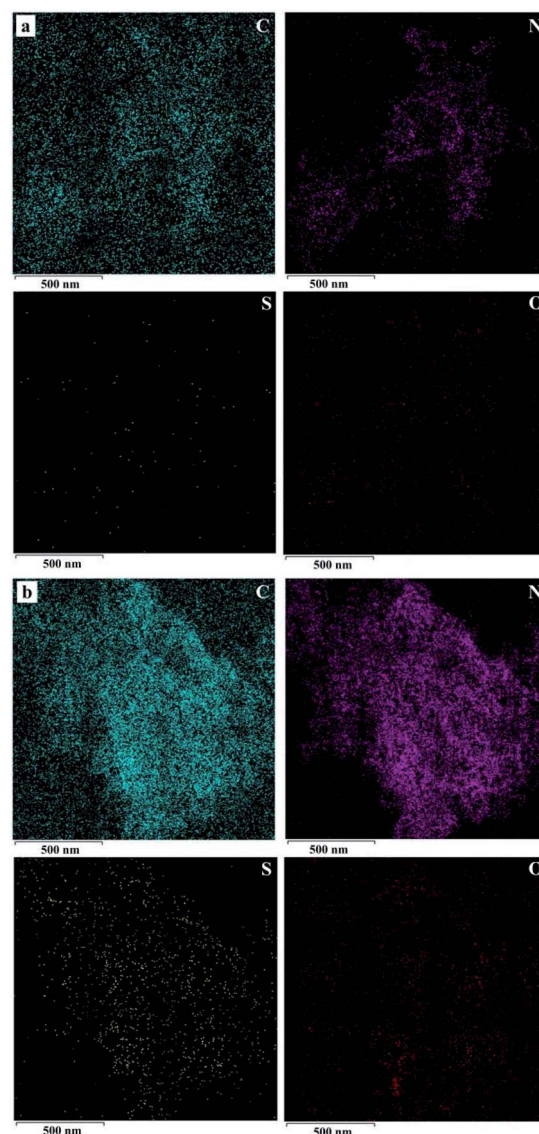
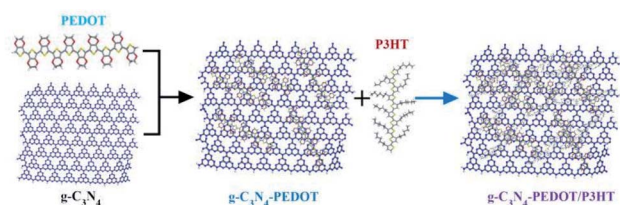


Fig. 1 Typical TEM images of the elemental mapping C (green), N (purple), S (yellow) and O (red) (a) g-C₃N₄, (b) g-C₃N₄-PEDOT/P3HT.



Scheme 1 Synthetic route of g-C₃N₄-PEDOT/P3HT composite.



main steps and the detailed process can be found in the experimental section.

The morphology of pure $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ was investigated with transmission electron microscopy (TEM) and high resolution TEM (HRTEM). From Fig. S1,[†] it can be observed that $g\text{-C}_3\text{N}_4$ possessed the typical crinkly structure as that reported in our former reports.³⁰ After the PEDOT and P3HT adsorbed to $g\text{-C}_3\text{N}_4$, the main morphologies of $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ almost maintain the similar crinkly structure as the pure $g\text{-C}_3\text{N}_4$ (shown in Fig. S2[†]). Besides, the HRTEM image of the pure $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ were also exhibited in Fig. S3(a and b).[†] It can be seen some small nanoparticles with sizes of 5–10 nm are located on the nanoplates of $g\text{-C}_3\text{N}_4$, which could be ascribed to the existence of PEDOT/P3HT. As the unobvious image contrast and amorphous structure between P3HT and PEDOT may lead us hard to distinguish the two polymersemiconductors under the TEM or even under the HRTEM. This kind of similar phenomenon or problem also reported in other literature.^{23,26} In order to further determine the presence of the PEDOT and P3HT in $g\text{-C}_3\text{N}_4$ and to confirm the distribution of the polymer, element mapping analysis was implemented by the energy dispersive X-ray spectroscopy (EDS). As shown in Fig. 1a and b, all C, N, S and O elements are evenly dispersed at the surface of $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$. The clearly appearing hyperintense S and O elements signals only come from the P3HT and PEDOT fragment, while the O elements signals only come from PEDOT fragment. As the contrast sample, the S and O elements disappeared in pure $g\text{-C}_3\text{N}_4$ which further indicated that the $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ composite was formed successfully.

The XRD pattern of the prepared samples are shown in Fig. 2. All the samples have the similar diffraction peaks at about 12.9° and 27.7° , corresponding to the typical characteristic diffraction peak of the pure $g\text{-C}_3\text{N}_4$.^{23,24} The diffraction peaks of the different $g\text{-C}_3\text{N}_4$ -polymer samples are very similar to those of $g\text{-C}_3\text{N}_4$, indicating that the addition of PEDOT and P3HT would not change the crystal structure of $g\text{-C}_3\text{N}_4$.

Fig. 3 shows the FTIR spectra of $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4\text{-P3HT}$, $g\text{-C}_3\text{N}_4\text{-PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$. As shown in Fig. 2, the obtained composite samples show the typical absorption peak

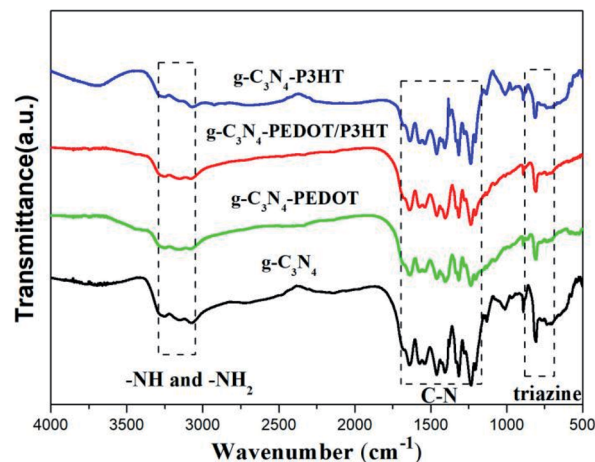


Fig. 3 The FTIR spectral of $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4\text{-P3HT}$, $g\text{-C}_3\text{N}_4\text{-PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$.

of pure $g\text{-C}_3\text{N}_4$. The typical absorption peak at 807 cm^{-1} is belonged to the breathing mode of the triazine units.³¹ In the scope of $1240\text{--}1650\text{ cm}^{-1}$, these peaks are corresponded to the stretching vibration of C-N.³² Additionally, the peaks at around $3150\text{--}3450\text{ cm}^{-1}$ are allocated to stretching vibration modes of the -NH_2 and -NH group of $g\text{-C}_3\text{N}_4$. Therefore, the composite photocatalysts $g\text{-C}_3\text{N}_4\text{-P3HT}$, $g\text{-C}_3\text{N}_4\text{-PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ show the similar characteristic peaks of pure $g\text{-C}_3\text{N}_4$ and no new peak appear in the spectrum, which means that the relatively low quantity of PEDOT or P3HT in the $g\text{-C}_3\text{N}_4$ -polymer composites or without chemical bond formed between $g\text{-C}_3\text{N}_4$ and PEDOT (or P3HT).

Optical and photoelectrochemical properties analysis

The UV-Vis diffuse reflectance spectra (DRS) of different polymer sensitized $g\text{-C}_3\text{N}_4$ samples are analyzed to show the optical absorption properties. From the Fig. 4, pure $g\text{-C}_3\text{N}_4$ only can absorb visible light with the absorption edge that is shorter than 460 nm, corresponding to a bandgap of $\sim 2.7\text{ eV}$. When P3HT is

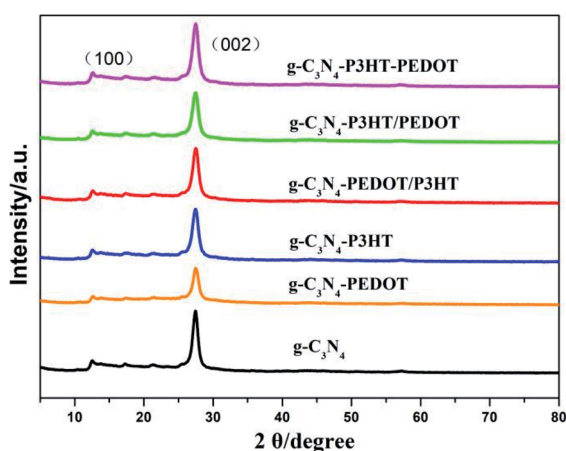


Fig. 2 XRD patterns of the as-prepared $g\text{-C}_3\text{N}_4$ -polymer samples.

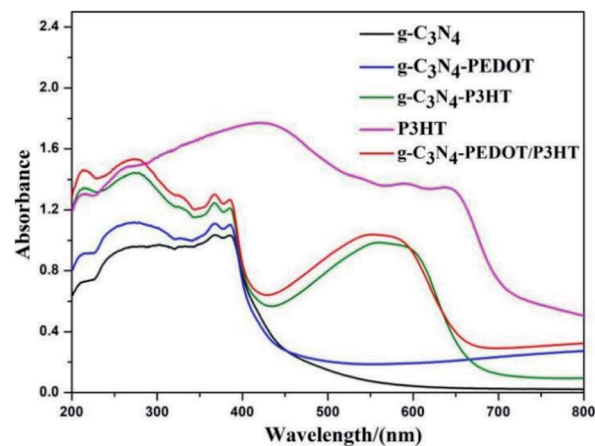


Fig. 4 UV-Vis diffuse reflectance spectra of the $g\text{-C}_3\text{N}_4$ -polymer samples.



loaded to $g\text{-C}_3\text{N}_4$, the $g\text{-C}_3\text{N}_4\text{-P3HT}$ shows a distinct spectral absorption band in the scope of 400–700 nm which is belong to the typical absorption area of the pure P3HT. However, the absorption edge of $g\text{-C}_3\text{N}_4$ or $g\text{-C}_3\text{N}_4\text{-P3HT}$ remained almost unchanged after the impregnation of PEDOT. The similar phenomenon can also be seen in $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$ (Fig. S4†). These results reveal that PEDOT did not change the electronic structure of $g\text{-C}_3\text{N}_4$ or $g\text{-C}_3\text{N}_4\text{-P3HT}$. The little increased baseline of the absorption by PEDOT could be due to the light blue color of PEDOT it self. Besides, the obvious colour change can be found in the ordinary optical photos of the as-prepared $g\text{-C}_3\text{N}_4$ -polymer samples also shown in Fig. S5.†

In order to investigate the role of PEDOT, P3HT and photo induced electron-hole separation abilities of obtained samples, the photoluminescence (PL) spectra analysis was performed and illustrated in Fig. 5. As can be seen, pure $g\text{-C}_3\text{N}_4$ shown a broad PL emission peak, corresponding to the fast carrier recombination of pure $g\text{-C}_3\text{N}_4$. When the PEDOT or P3HT was loaded to $g\text{-C}_3\text{N}_4$, the PL peak appeared gradually quenched for $g\text{-C}_3\text{N}_4\text{-P3HT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT}$. Significant PL quenching was observed in the polymer composite surface heterojunction (SHJ) catalysts. For the PL spectra of $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$, the intensity was much lower than that of pure $g\text{-C}_3\text{N}_4$ or $g\text{-C}_3\text{N}_4\text{-PEDOT}$. These results indicated that the efficient charge transfer or separation from both PEDOT and P3HT to $g\text{-C}_3\text{N}_4$ could be achieved successfully by the proposed strategy, which was benefit for the enhanced photocatalytic performance.

Fig. 6 shows the transient photo-current behavior of the obtained $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4\text{-P3HT}$, $g\text{-C}_3\text{N}_4\text{-PEDOT}$, $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$. For the pure $g\text{-C}_3\text{N}_4$, it appears the lower transient photocurrent response, while the photocurrent of $g\text{-C}_3\text{N}_4\text{-P3HT}$, $g\text{-C}_3\text{N}_4\text{-PEDOT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ or $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$ is relatively higher than that of the pure $g\text{-C}_3\text{N}_4$. It is obviously observe that the photocurrent of $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ is much higher than that of $g\text{-C}_3\text{N}_4\text{-P3HT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT}$. The much

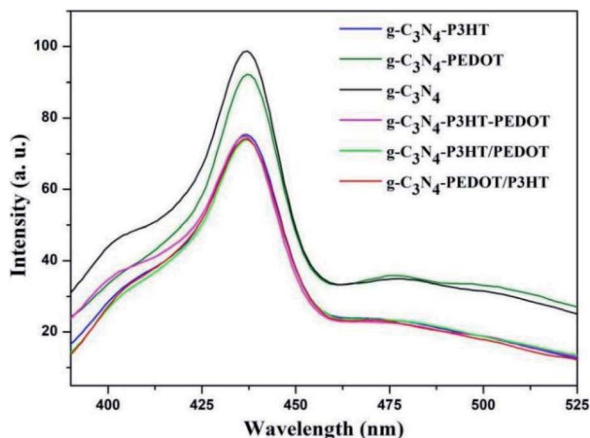


Fig. 5 Photoluminescence emission spectra of pure $g\text{-C}_3\text{N}_4$ and all the $g\text{-C}_3\text{N}_4$ -polymer composite samples.

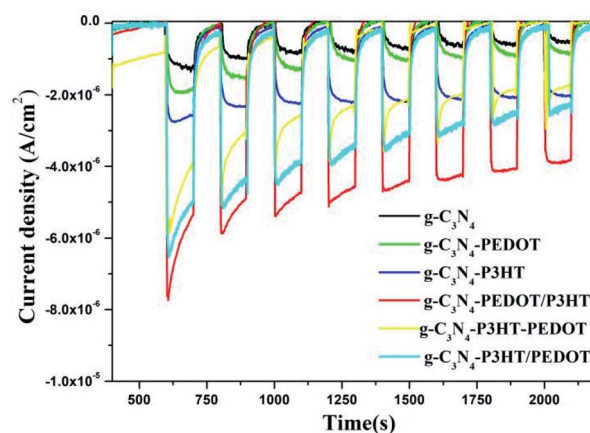


Fig. 6 Transient photocurrent performance of the obtained $g\text{-C}_3\text{N}_4$ and the $g\text{-C}_3\text{N}_4$ -polymer composite samples in Na_2SO_4 solution (0.1 M).

improved photocurrent over $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ further intimates the synergy among $g\text{-C}_3\text{N}_4$, PEDOT and P3HT. The highest photocurrent density also demonstrates that $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ is superior to $g\text{-C}_3\text{N}_4\text{-P3HT}$ and $g\text{-C}_3\text{N}_4\text{-PEDOT}$ in separating photo induced electron-hole pairs.

Photocatalytic activity of the $g\text{-C}_3\text{N}_4$ composite samples

The photocatalytic H_2 evolution performance of the obtained photocatalysts were assessed under visible light using ascorbic acid (AA) as sacrificial reagent. As shown in the Fig. 7a, when loaded with Pt as a cocatalyst, $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4\text{-PEDOT}$ shows a photocatalyst H_2 evolution rate of $403.7 \mu\text{mol h}^{-1} \text{g}^{-1}$ and $3633.9 \mu\text{mol h}^{-1} \text{g}^{-1}$, respectively. However, after P3HT added to $g\text{-C}_3\text{N}_4$, the H_2 evolution rate of $g\text{-C}_3\text{N}_4\text{-P3HT}$ reaches at $353233.3 \mu\text{mol h}^{-1} \text{g}^{-1}$, which is over 875 times higher than that of $g\text{-C}_3\text{N}_4$ or 97 times higher than that of $g\text{-C}_3\text{N}_4\text{-PEDOT}$. As reported, the wide light absorption and efficient charge migration between P3HT and $g\text{-C}_3\text{N}_4$ was responsible for the magnificent photoactivity of the polymer/polymer surface heterojunction catalyst. Based on the aforementioned assumptions, we want to construct a $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$ surface heterojunction catalyst with P3HT and PEDOT as a polymer sensitizer and hole transport pathway to improve the photocatalytic activity, respectively. In order to verify the above hypothesis, the photocatalyst H_2 evolution rate of $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$ are also tested. When the PEDOT is added to $g\text{-C}_3\text{N}_4\text{-P3HT}$, the H_2 evolution rate is enhanced obviously, which is of $427703.3 \mu\text{mol h}^{-1} \text{g}^{-1}$, $421030.0 \mu\text{mol h}^{-1} \text{g}^{-1}$ and $407256.7 \mu\text{mol h}^{-1} \text{g}^{-1}$ for $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$, $g\text{-C}_3\text{N}_4\text{-P3HT/PEDOT}$ and $g\text{-C}_3\text{N}_4\text{-P3HT-PEDOT}$, respectively. The results also confirm that the order of adding P3HT and PEDOT to $g\text{-C}_3\text{N}_4$ has little effect on the properties of the composite sample. As the best sample $g\text{-C}_3\text{N}_4\text{-PEDOT/P3HT}$, the photocatalyst H_2 evolution rate is up to $427703.3 \mu\text{mol h}^{-1} \text{g}^{-1}$ which is 1059 times higher than that of $g\text{-C}_3\text{N}_4$, 118 times higher than that of $g\text{-C}_3\text{N}_4\text{-PEDOT}$ or even 1.2 times higher than that of $g\text{-C}_3\text{N}_4\text{-P3HT}$. Besides, the apparent

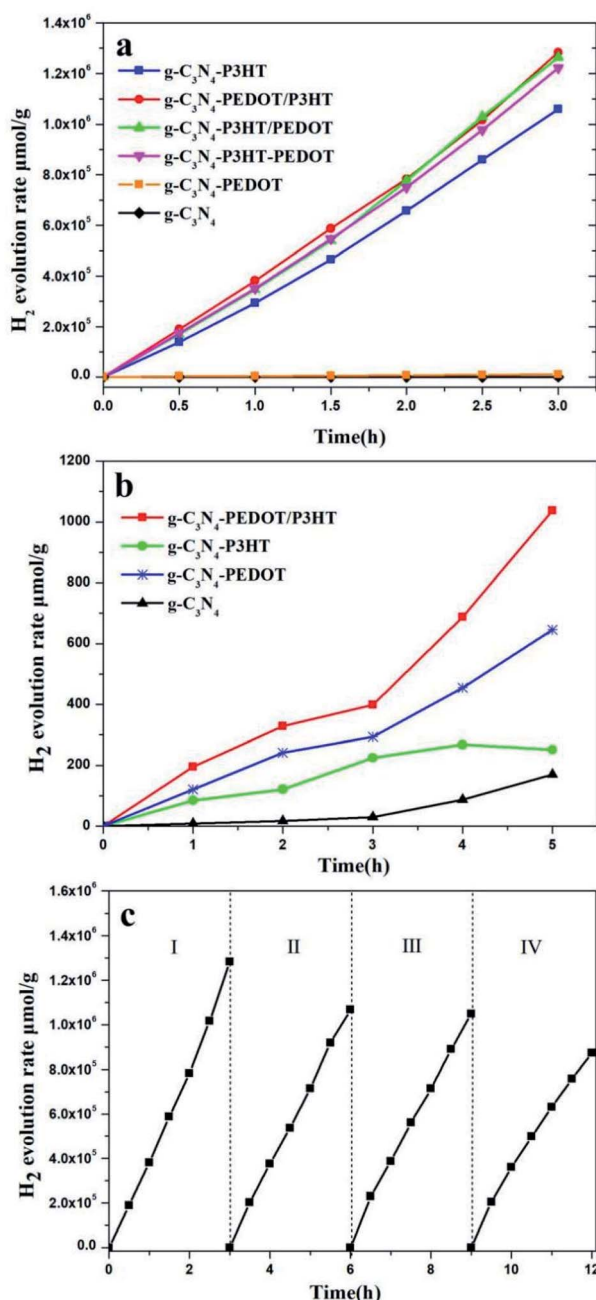


Fig. 7 The photocatalytic H₂ evolution as a function of reaction time. (a) The comprehensive comparison with AA as sacrificial reagent, (b) the comprehensive comparison from pure water, (c) the recyclable stability of H₂ evolution for g-C₃N₄-PEDOT/P3HT under visible light ($\lambda \geq 400$ nm).

quantum yields (AQY) of g-C₃N₄-PEDOT/P3HT were also tested, the AQY values were 4.4%, 10.6% and 14.9% under 420 nm, 520 nm and 600 nm monochromatic light, respectively. In addition, the comparison of the H₂ evolution activity and AQY between this work and other similar polymer composite photocatalysts reported in the literature was added which was shown in Table S1 (ESI†).

In order to further confirm the separation situation of the photo generated electron-hole, the photocatalytic H₂ evolution from pure water is investigated too. The results in Fig. 7b show that the amount of H₂ evolution rate of g-C₃N₄, g-C₃N₄-P3HT, g-C₃N₄-PEDOT and g-C₃N₄-PEDOT/P3HT is 56.9 $\mu\text{mol h}^{-1} \text{g}^{-1}$, 83.8 $\mu\text{mol h}^{-1} \text{g}^{-1}$, 215.4 $\mu\text{mol h}^{-1} \text{g}^{-1}$ and 345.7 $\mu\text{mol h}^{-1} \text{g}^{-1}$. The g-C₃N₄ and g-C₃N₄-P3HT exhibit the lower photocatalytic H₂ evolution rate, while the g-C₃N₄-PEDOT and g-C₃N₄-PEDOT/P3HT show the obviously enhanced photocatalytic H₂ evolution rate which is 3.7 and 6.1 times higher than that of g-C₃N₄. Besides, the photocatalytic evolved gas is only hydrogen, no oxygen is detected. And these kind of similar phenomenon were also observed in other work, like g-C₃N₄-Pt or PPy/g-C₃N₄ in pure water, and the possible product of water oxidation is hydrogen peroxide.²⁵

The encouraging result indeed confirms the synergy between the P3HT and PEDOT when both of them are binding with g-C₃N₄, which is beneficial for the separation of the photo induced electron-hole. The recyclable stability of g-C₃N₄-PEDOT/P3HT is also evaluated by the cycling experiments under $\lambda \geq 400$ nm light irradiation conditions (shown Fig. 7c). After four cycles, H₂ evolution activity still maintains about 75.0% of

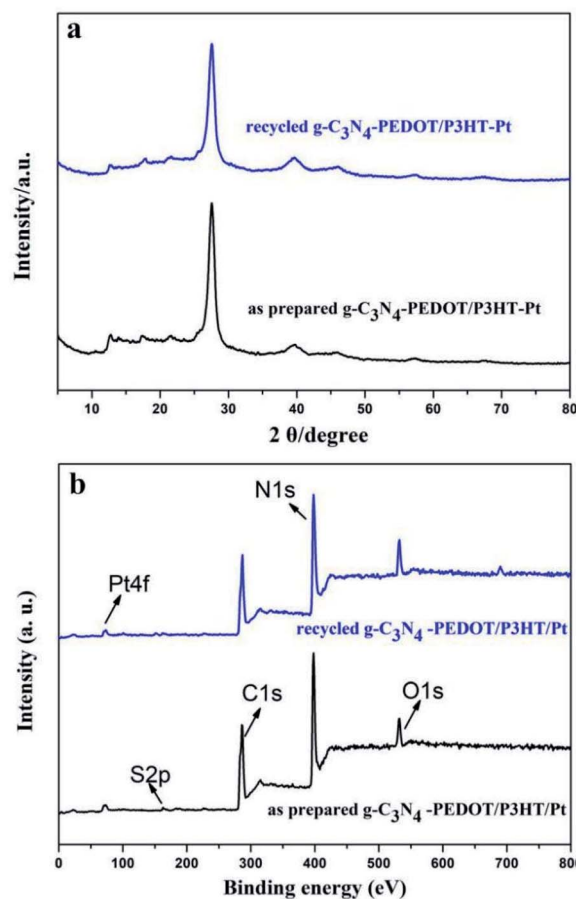


Fig. 8 (a) The XRD patterns of the g-C₃N₄-PEDOT/P3HT-Pt composite before and after stability test in AA solution, (b) the total XPS element spectrum of g-C₃N₄-PEDOT/P3HT-Pt composite before and after stability test in AA solution.

the initial one. Besides, the additional characterization of the recycled g-C₃N₄-PEDOT/P3HT/Pt composite after stability test were performed to further demonstrate the stability of the target photocatalyst. From Fig. 8a and b, it can be seen that the typical crystal structure and main elements characteristic peaks (the specific element of C, N, O and Pt shown in Fig. S6†) appeared no obvious change after four cycles, which further confirmed the photocatalyst was quite stable under the reaction condition.

Photocatalytic mechanism of the g-C₃N₄-polymer catalyst

In this part, MD simulations were conducted to investigate the possible interaction between g-C₃N₄ and PEDOT or P3HT polymer. In order to better clarify the microstructure of g-C₃N₄ hybriide with PEDOT or P3HT, the radial distribution functions (RDF) and mean square displacement (MSD) were calculated. Fig. 9(I) and (III) showed the equilibrated configurations and partial enlarged view of g-C₃N₄-polymer. As can be seen, the g-C₃N₄ structure was nearly unchanged, and only weak van der Waals forces presented between g-C₃N₄ and PEDOT or P3HT. From the RDF results in Fig. 9(II), the distance between C and N from g-C₃N₄ with S from PEDOT was mainly centered at the range about 3.9 Å ~ 4.1 Å, while the distance between C and N

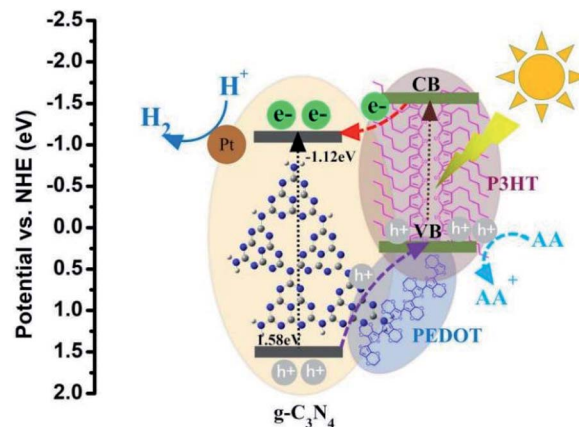


Fig. 10 A proposed photocatalytic H₂ production mechanism of g-C₃N₄-PEDOT/P3HT/Pt SHJ catalyst structure.

from g-C₃N₄ with S from P3HT was disperse with a relatively large range about 3.8 Å ~ 5.0 Å which meant a relatively stronger interaction between g-C₃N₄ with PEDOT. The similar results were also confirmed by MSD and shown in Fig. 9(IV). For g-C₃N₄-PEDOT, the diffusion coefficient of g-C₃N₄ and PEDOT were kept the same level about $5.18 \times 10^{-14} \text{ m}^2 \text{ s}^{-1}$ (g-C₃N₄) and $7.28 \times 10^{-14} \text{ m}^2 \text{ s}^{-1}$ (PEDOT). However, the diffusion coefficient of g-C₃N₄ and P3HT appeared a larger difference, about $5.47 \times 10^{-14} \text{ m}^2 \text{ s}^{-1}$ (g-C₃N₄) and $1.04 \times 10^{-12} \text{ m}^2 \text{ s}^{-1}$ (P3HT). The smaller difference of the diffusion coefficient between g-C₃N₄ and PEDOT illustrated a relatively stronger interaction between the two compared with that of g-C₃N₄ and P3HT, which may attributed to the relatively small molecular volume of PEDOT and without any side chain resistance. Then we may imagine that the g-C₃N₄-PEDOT/P3HT SHJ catalyst could be formed with P3HT as a polymer sensitizer is the upper layer, while the PEDOT as hole transport pathway is interlayer between P3HT and g-C₃N₄, and the three different part can form an effective mutual synergy system eventually.

The possible mechanism of photocatalysts for the superior photocatalytic performance of H₂ evolution is shown in the Fig. 10. As formerly reported, the LUMO level of P3HT is more negative than the conduction band (CB) of g-C₃N₄ which is favorable for the photo induced electrons transferring from P3HT to g-C₃N₄, and then trapped by the Pt for H₂ production. Besides, as the valence band (VB) of g-C₃N₄ is more positive than that of P3HT, the holes may have the trend of spontaneous migration from the VB of g-C₃N₄ to the VB of P3HT. More importantly, the existing PEDOT possesses the relatively higher hole mobility and can be used as a hole conductor between g-C₃N₄ and P3HT. Then, the photogenerated holes migration can be accelerated by PEDOT from the VB of g-C₃N₄ to the VB of P3HT. Finally, the holes on the VB of P3HT are depleted by AA. All those factors may benefit the rapid migration of photo generated electron-hole pairs and eventually facilitate the photocatalytic H₂ activity process of g-C₃N₄-PEDOT/P3HT under visible light.

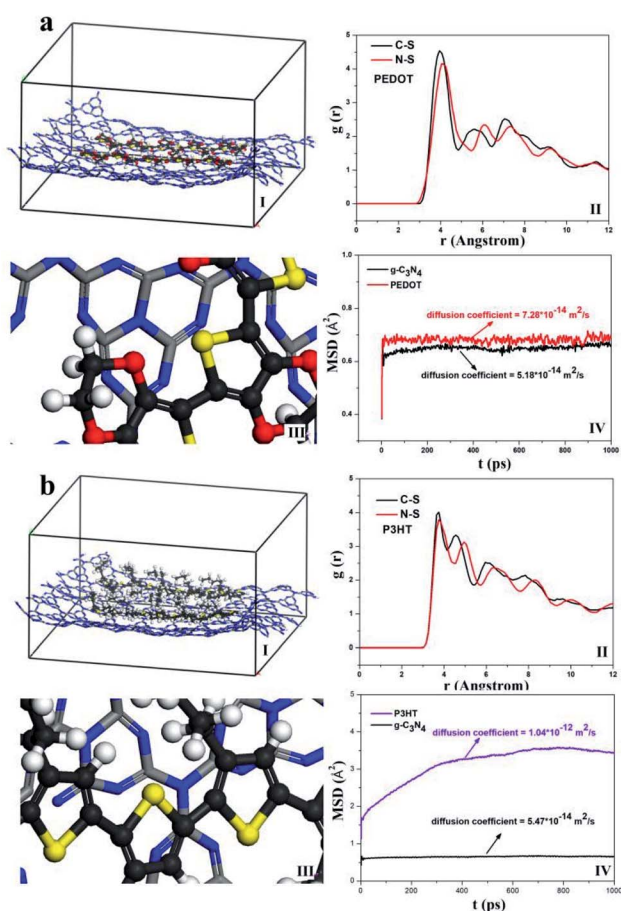


Fig. 9 MD simulated equilibrated configurations (I) g-C₃N₄-polymer (II) RDF (III) partial enlarged view of g-C₃N₄-polymer and (IV) MSD (a) g-C₃N₄-PEDOT, (b) g-C₃N₄-P3HT.



Conclusions

In conclusion, the g-C₃N₄-PEDOT/P3HT surface heterojunction (SHJ) catalyst with P3HT and PEDOT as a polymer sensitizer and hole transport pathway was successfully constructed. Moreover, as Pt particles are often deposited to g-C₃N₄ as electron sink and proton reduction sites, we successfully construct a g-C₃N₄-PEDOT/P3HT/Pt composite, in which PEDOT act as hole transport pathway, P3HT as the polymer sensitizer and Pt as the reduction reaction sites. In this way, the g-C₃N₄-PEDOT/P3HT/Pt composite exhibits the photocatalyst H₂ evolution rate is up to 427703.3 $\mu\text{mol h}^{-1} \text{g}^{-1}$ which is 1059 times higher than that of g-C₃N₄, 118 times higher than that of g-C₃N₄-PEDOT. The dramatically enhanced photocatalytic performance can be assigned to the synergy among g-C₃N₄, PEDOT and P3HT, which finally facilitate the separated rapid migration of photoinduced electron-hole pairs. This work may provide a useful insights for designing the surface heterojunction (SHJ) composites photocatalyst with high photocatalytic activity for H₂ production.

Experimental

Materials

The poly(3-hexylthiophene-2,5-diyl) (regioregular) was acquired from Sigma-Aldrich, poly(3,4-ethylenedioxythiophene)-poly(styrenesulfonate) (1.5% in water) was purchased from Macklin, and both used without any further treatment. The rest reagents and solvents were used as received from commercial suppliers without additional treatment.

Synthesis of g-C₃N₄-polymer composite photocatalysts

g-C₃N₄ was synthesized according to our previous reported work.³² Briefly, melamine (3 g), urea (7.5 g) and ammonium chloride (2 g) were mixed in an alumina crucible with a cover and the precursor was heated at 600 °C for 4 hours in air with the rate of 5 °C min⁻¹. After cooling to the room temperature, the product was gathered for later use.

g-C₃N₄-P3HT. 18 mg P3HT was solved in 20 mL chloroform with 100 mL flask and 600 mg g-C₃N₄ was added, then the suspension was sonicated for 20 minutes with stirring for 12 h at room temperature. The obtained suspension was then evaporated by rotary evaporator and dried in a 80 °C vacuum oven for 2 h, the g-C₃N₄-P3HT sample was obtained.

g-C₃N₄-PEDOT. 600 mg g-C₃N₄ was dispersed in 20 mL THF with 100 mL flask and 800 μL PEDOT was added, then the suspension was sonicated for 20 minutes with stirring for 12 h at room temperature. The obtained suspension was then evaporated by rotary evaporator. In order to make the PEDOT insoluble, the obtained powder was subsequently immersed in 20 mL ethylene glycol (EG) and stirred for 1 hour. After treated with EG, the g-C₃N₄-PEDOT composite was centrifuged, washed with deionized water and anhydrous ethanol for several times and dried in a 100 °C vacuum oven for 2 h, the g-C₃N₄-PEDOT sample was obtained.

g-C₃N₄-P3HT/PEDOT. The initial component amount remains the same as above g-C₃N₄-P3HT and 800 μL PEDOT

was further added, the specific synthetic procedure was same as g-C₃N₄-PEDOT to obtain the g-C₃N₄-P3HT/PEDOT sample.

g-C₃N₄-PEDOT/P3HT. The initial component amount remains the same as above g-C₃N₄-PEDOT and 18 mg P3HT was further added, the specific synthetic procedure was same as g-C₃N₄-P3HT to obtain the g-C₃N₄-PEDOT/P3HT sample.

g-C₃N₄-P3HT-PEDOT. 600 mg g-C₃N₄ was dispersed in 20 mL THF with 100 mL flask, then 18 mg P3HT and 800 μL PEDOT were added in the same time, then the suspension was sonicated for 20 minutes with stirring for 12 h at room temperature. The obtained suspension was then evaporated by rotary evaporator. In order to make the PEDOT insoluble, the obtained powder was subsequently immersed in 20 mL EG and stirred for 1 hour. After treated with EG, the g-C₃N₄-P3HT-PEDOT composite was centrifuged, washed with deionized water and anhydrous ethanol for several times and dried in a 100 °C vacuum oven for 2 h, the g-C₃N₄-P3HT-PEDOT sample was obtained.

Photocatalysts characterization

TEM and HRTEM characterization were characterized on a JEOL-2010 electron microscope at an acceleration voltage of 200 kV. The crystal phases of the samples were analyzed by the Rigaku D/MAX 2500X X-ray powder diffractometer with Cu K α radiation. XPS measurements were performed on a Thermo Scientific ESCALAB Xi+ (monochromatic Mg K α , 1253.6 eV). The UV-vis spectrometer was carried out on an Agilent Cary 5000 UV-vis-NIR spectrophotometer with BaSO₄ as a reference. Photoluminescent (PL) spectra were recorded by a spectrofluorophotometer (Cary Eclipse) with excitation slit widths of 5 nm and emission slit widths of 5 nm. The infrared pattern was analysed by FTIR spectrometer (PE Spectrum Two).

Photocatalytic activity analyse

Photocatalytic activity tests were carried out on a CEL-SPH2N-S9 photocatalytic water splitting device (Ceaulight Beijing). The Pt content with 3 wt% was deposited to the catalyst according to the method in our previous work.³² Consequently, the various g-C₃N₄-polymer composites photocatalysts were prepared first and then the Pt was loaded to composite g-C₃N₄-polymer catalyst. For specific photocatalytic H₂ experiment, 10 mg catalyst loaded with 3 wt% Pt sample was added into 10 mL deionized water containing 1.8 M ascorbic acid as the sacrificial reagent aqueous solution by ultrasonication for 10 min. Prior to the light irradiation, the system was eliminated for 30 min to remove air. Next, the solution temperature was maintained at 3 °C. The measurements were performed under illumination of a Xenon lamp (300 W, Ceaulight Beijing) with a 400 nm cutoff filter and the corresponding incident light intensity was 100 mW cm⁻². The photocatalytic H₂ evolution rate was detected by a gas chromatography (GC-7920, Ceaulight Beijing) with high purity N₂ as carrier gas. For apparent quantum yields (AQY) test, 20 mg as prepared catalyst sample was added into 45 mL 0.6 M H₂SO₄ solution containing 13.3 g ascorbic acid as sacrificial reagent by ultrasonication for 10 min. Before irradiation, the system was evacuated for 30 min to remove air and the



temperature of the solution was maintained at 3 °C. A 300 W Xenon lamp (Ceaulight Beijing) with a 400 nm cutoff filter was used for irradiation the mixture for 60 min and evacuated for another 30 min, then various band-pass filters (420 nm, 520 nm and 600 nm) was used as the irradiation source to measure the monochromatic AQY for the H₂. Besides, the specific AQY processing may reference our former work.³²

Photo electrochemical measurements

The carbon cloth (CeTech Co. Ltd, WO1002) was cleaned by sonication sequentially in acetone, ethanol and deionized water for 30 min each, followed by drying in vacuum at 100 °C. Then, chemical activation was operated by submerging carbon cloth into a mixed solution with concentrated sulfuric acid (H₂SO₄, 98%) and concentrated nitric acid (HNO₃, 68%) in a volume ratio of 1 : 3 (H₂SO₄/HNO₃) for 12 hours. Afterwards, the activated carbon cloth was washed with deionized water until neutral, and then placed for standby. Subsequently, 5 mg obtained photocatalysts sample were mixed with 990 μL ethanol, 10 μL Nafion solution (5 wt%, Aladdin). Then, the sample mixture solution was ultrasonic treatment for 30 min. The obtained suspension (about 100 μL) was dropped onto the activated carbon cloth electrode and dried naturally in the air. A three-electrode system was employed where the as prepared carbon cloth electrode, a platinum piece, and a saturated calomel electrode (SCE) were utilized as a working electrode, a counter electrode, and a reference electrode, respectively. Besides, the Na₂SO₄ solution (0.1 M) was used as electrolyte for the transient photocurrent *I*-*T* curve testing with above three-electrode system. The measurements were performed under illumination of a Xenon lamp (300 W, Ceaulight Beijing) with a 400 nm cutoff filter and the corresponding incident light intensity was 100 mW cm⁻².

Computational methods

In this work, MD simulation with a full atomistic force field was conducted to investigate the possible interaction between g-C₃N₄ and PEDOT or P3HT polymer.³³ The COMPASS module from Materials Studio developed by Accelrys Inc was used to conduct the computations. For the target system, the initial simulation system was optimized with 50 000 steps. Then, the NVT dynamic equilibration was running for 3 ns with time step 1 fs. Besides, the mean square displacement (MSD) and radial distribution functions (RDF) was obtained by analyzing the data collected at 1000 steps intervals during the last 1 ns. The molecular structure of monolayer g-C₃N₄, the chain of PEDOT, the chain of P3HT and g-C₃N₄-polymer composites were depicted in Fig. S7.†

Conflicts of interest

There are no conflicts to declare.

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