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# Single-atom niobium doped BCN nanotubes for highly sensitive electrochemical detection of nitrobenzene

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Herein, single-atom niobium-doped boron–carbon–nitrogen nanotubes (SANb-BCN) were synthesized and utilized to fabricate an electrochemical sensor for the detection of nitrobenzene (NB), an environmental pollutant. SANb-BCN were characterized through scanning transmission electron microscopy, scanning electron microscopy, transmission electron microscopy, X-ray diffraction analysis, and Raman spectroscopy. The Nb-BCN material modified on a glassy carbon electrode (GCE) showed an excellent electrochemical response behavior toward NB. The SANb-BCN-modified GCE (SANb-BCN/GCE) gave rise to a prominent NB reduction peak at  $-0.6$  V, which was positively shifted by 120 mV from the NB reduction peak of the bare GCE. Furthermore, the NB peak current ( $55.74 \mu\text{A}$ ) obtained using SANb-BCN/GCE was nearly 42-fold higher than that using the bare GCE ( $1.32 \mu\text{A}$ ), indicating that SANb-BCN/GCE is a highly sensitive electrochemical sensor for NB. An ultralow limit of detection ( $0.70 \mu\text{M}$ ,  $S/N = 3$ ) was also achieved. Furthermore, the SANb-BCN/GCE sensor was found to possess favorable anti-interference ability during NB detection; thus, the presence of various organic and inorganic coexisting species, including  $\text{Mg}^{2+}$ ,  $\text{Cr}^{6+}$ ,  $\text{Cu}^{2+}$ ,  $\text{K}^{+}$ ,  $\text{Ca}^{2+}$ ,  $\text{NH}_4^{+}$ ,  $\text{Cd}^{2+}$ , urea, 1-bromo-4-nitrobenzene, 3-hydroxybenzoic, terephthalic acid, 1-iodo-4-nitrobenzene, and toluene, minimally affected the NB detection signal. Notably, the SANb-BCN sensor material exhibited high sensitivity and specificity toward detection of NB in environmental samples. Thus, the use of the proposed sensor will serve as an effective alternative method for the identification and treatment of pollutants.

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## Introduction

Nitrobenzene (NB) is a raw material used in various organic syntheses, such as the production of aniline and pesticides.<sup>1,2</sup> However, it is a persistent organic pollutant, and its presence in the environment must be monitored. NB can enter the human body through the skin and respiratory tract, and its total retention rate can reach  $\sim 80\%$ .<sup>3</sup> In addition, intermediate substances produced by the biotransformation of NB,<sup>4</sup> such as *m*-nitrophenol and *p*-nitrophenol, are more toxic than NB itself. Due to the high toxicity of NB in humans and animals, the World Health Organization listed it as a 2B carcinogen in 2017.

NB is often unavoidably released into the environment, and regular monitoring of NB in water systems is the only preventive measure to avoid permanent damage to human and animal health.<sup>5–7</sup> To this end, the development of efficient, rapid, and convenient NB detection methods is of great significance.

Numerous methods have been developed for detecting NB in water systems, which employ techniques such as high-phase liquid chromatography, meteorological chromatography, and electrochemical analysis.<sup>8,9</sup> Although rapid and selective NB detection is achievable through chromatography, several disadvantages including low recognition ability, complex operation, and bulky equipment prevent its application as a portable technique for on-site detection.<sup>10</sup> In contrast, electrochemical methods allow for rapid, portable, convenient analysis as well as exhibit low limits of detection (LODs).<sup>11</sup> In addition, electrochemical methods allow for the integration of various analysis modes such as differential pulse voltammetry, and linear sweep voltammetry (LSV), cyclic voltammetry, square wave voltammetry.

Single-atom catalysis has emerged as an effective tool for numerous typical heterogeneous catalytic processes,<sup>12</sup> including CO oxidation,<sup>13,14</sup> the water-gas shift reaction,<sup>15</sup> hydrogen evolution reaction,<sup>16,17</sup> hydrogenation,<sup>18,19</sup> oxygen

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reduction reaction,<sup>20,21</sup> and degradation of persistent organic pollution.<sup>22</sup> Single-atom catalysts exhibit excellent performance toward various reactions owing to their unique coordination structure as well as high atom utilization rate, activity, and selectivity.<sup>23,24</sup> The use of such catalysts is expected to bridge the gap between homogeneous<sup>25</sup> and heterogeneous catalysis.<sup>26,27</sup> The synthetic technology for accessing single-atom catalysts has gradually matured.<sup>28</sup> Yin *et al.* synthesized single atom Fe catalysis to catalyze Fenton-type reactions,<sup>29</sup> and Xu *et al.* developed single cobalt atom catalyst improving peroxymonosulfate oxidation of several organic pollutants.<sup>30</sup> Wu *et al.* developed a monoatomic iron catalyst for nitrate reduction and electrochemical ammonia synthesis.<sup>31</sup>

Every structural unit of monoatomic materials cannot be ignored. Similar to metal enzymes in nature, in fact, in addition to the central metal atom, the structure of the single-atom sensor also has adjacent coordination atoms, as well as the supported substrate and the functional groups on its surface, all of which play an important role. The reason for the excellent performance of monoatomic materials in electrochemical sensing applications is the unique monodisperse metal sites and electronic structure. Although such representative studies have demonstrated the effectiveness and applicability of single-atom catalysis, only a few single-atom sensors have been reported for the detection of NB.<sup>32,33</sup> Therefore, it is important to develop a single-atom sensor for the detection of pollutants.

Herein, we reported the composite of single-atom niobium-doped boron–nitrogen–carbon nanotubes (SANb-BCN) for highly sensitive electrochemical detection of NB. Fig. 1 illustrates the preparation process and detection mechanism of the sensor. A single-atom Nb sensor supported on B, N, and C co-doped carbon nanotubes was prepared using the template method.<sup>34,35</sup> The following characterization methods were applied successively to determine the morphology and element distribution: X-ray diffraction (XRD), scanning electron microscopy (SEM), scanning transmission electron microscopy (STEM), transmission electron microscopy (TEM) analysis, and Raman spectroscopy. A modified electrode (SANb-BCN/GCE) was prepared as the sensor, and the electrochemical behavior of NB on the electrode was studied by controlling the electrochemical experimental conditions. Under optimized conditions, SANb-BCN/GCE exhibited high selectivity for NB and

a low LOD. Finally, SANb-BCN/GCE was used to analyze the target contaminant in actual samples and showed excellent results.

## Experimental

### Reagents and instruments

All reagents were of analytical grade. Polyethylene glycol ( $\text{HO}(\text{CH}_2\text{CH}_2\text{O})_n\text{H}$ ), urea ( $\text{CH}_4\text{N}_2\text{O}$ ), and boric acid ( $\text{H}_3\text{BO}_3$ ) were purchased from Aladdin (Shanghai, China). Niobium oxalate ( $\text{C}_{10}\text{H}_5\text{NbO}_{20}$ ) and NB ( $\text{C}_6\text{H}_5\text{NO}_2$ ) were obtained from Macklin (Shanghai, China) and were used directly without purification. The PBS buffer solution contained a certain concentration of sodium dihydrogen phosphate (Macklin) and disodium hydrogen phosphate (Macklin) and was freshly prepared. Ultrapure water ( $\geq 18.25 \text{ M}\Omega \text{ cm}^{-1}$ ) was used to configure all solutions used in the experiments.

LSV responses were recorded using an electrochemical workstation named CHI-660E (Shanghai Chenhua Instrument Co., Ltd., China) equipped with a three-electrode system at room temperature. In this system, the counter electrode is a platinum wire electrode, the reference electrode uses Ag/AgCl electrode, and the modified electrode and GCE were used as the working electrode.

### Synthesis of SANb-BCN

Briefly, niobium oxalate (50 mg), urea (10 g), polyethylene glycol (1 g), and boric acid (0.3 g) were dissolved in deionized water with stirring. Then, a melamine foam template (0.3 g) was soaked in the solution for 15 hours. After solvent removal, the sample was vacuum dried at  $80^\circ\text{C}$  for 24 hours to obtain a solid powder. The powder was calcined and carbonized in a tube furnace under a nitrogen atmosphere at  $900^\circ\text{C}$  ( $2^\circ\text{C min}^{-1}$ ) to obtain SANb-BCN.

### Preparation of SANb-BCN/GCE

The GCE was gently wiped with moist lens cleaning paper to ensure that the electrode surface was clean and smooth. The electrode surface was then successively polished with 300 nm and 50 nm polishing powder using a chamois. And in each polishing stage, use deionized water, ethanol and acetone in order to ultrasonic for 15 s, and then dried in a nitrogen atmosphere for later use. SANb-BCN (2 mg) was ultrasonically dispersed in deionized water (1 mL), and the dispersion (10  $\mu\text{L}$ ) was dropped onto the surface of the vertical GCE, followed by air drying for 48 h to obtain SANb-BCN/GCE, at room temperature.

### Electrochemical measurements

LSV was performed at a scan rate of  $50 \text{ mV s}^{-1}$  in the range from  $-0.1$  to  $-1 \text{ V}$  potential range. Prior to each measurement,  $\text{N}_2$  was bubbled through the electrolyte for 15 min to remove any dissolved oxygen. The amperometric  $i-t$  (IT) curve allows NB to be deposited on the electrode surface for LSV testing. The IT curve was obtained by stirring and enriching NB for 100 s (enrichment time) at a voltage of  $-0.3 \text{ V}$  (deposition potential).

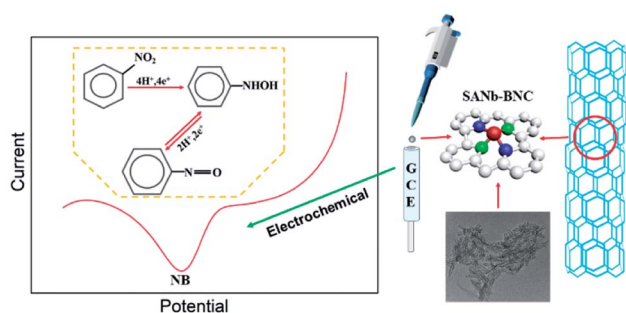


Fig. 1 Detection mechanism and preparation process of Nb-BCN/GCE.



The electrolyte used for the LSV test comprised PBS (10 mL, 0.1 M, pH 7.0) containing 0.1 mM NB.

## Results and discussion

### Characterization of SANb-BCN

As seen in the TEM and SEM images shown in Fig. 2a and b, respectively, the prepared SANb-BCN presented in the form of rod-shaped nanotubes morphology. The length of the nanotubes varied, and their diameter was approximately 50 nm. Single Nb sites, as opposed to Nb clusters, can be clearly observed in the STEM images. In addition, the EDS elemental mapping results confirmed that Nb, B, N, and C were uniformly distributed. These findings verified that the Nb atoms were monodispersed in the nanotubes. Such uniform dispersion can maximize atom utilization, activity, and contact with the test substance.

Two peaks were observed at 1330 and 1585  $\text{cm}^{-1}$  in the Raman spectrum of Nb-BNC (Fig. 3a), characteristic of carbon atom crystals. Peak D represents defects in the lattice of C atoms, and peak G is attributed to the in-plane stretching vibration of  $\text{sp}^2$  hybridized C atoms.<sup>33</sup> A simple calculation [ $I(\text{D})/I(\text{G}) = 1.16$ ] was employed to reveal the existence of defects in the C atom crystal. In the XRD pattern shown in Fig. 3b, the two diffraction peaks at  $27^\circ$  and  $42^\circ$  correspond to the two faces of graphitic carbon (002 and 101). Nb diffraction peaks were not observed in the entire scanned spectrum. Based on these results, it was concluded that SANb-BCN did not contain Nb-related crystals. X-ray absorption fine structure (EXAFS) of the Nb-BCN nanotubes confirmed that each Nb atom in Nb-BNC nanotubes was atomically coordinated with two B atoms and two N atoms fourfold.<sup>36</sup> The dispersion of single Nb atoms on

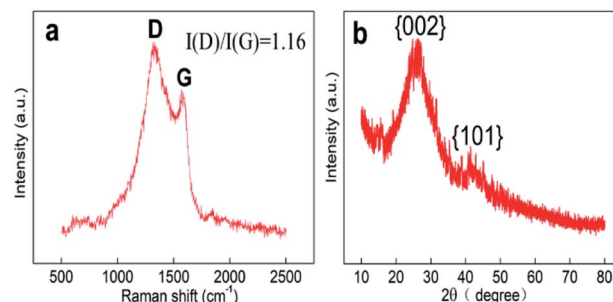


Fig. 3 (a) Raman spectrum and (b) XRD pattern of Nb-BNC.

the nanotube provides abundant active sites on SANb-BCN, thereby enhancing its electrochemical sensing performance.

### Electrode comparison

To evaluate the electrochemical sensing performance of SANb-BCN/GCE for NB, we compared the LSV curves of a bare GCE and those of SANb-BCN/GCE obtained under the same NB concentration (Fig. 4). In the case of the bare electrode, the curves obtained in the presence (red curve) and absence of NB (black curve) differed only slightly, with the former showing a weak reduction peak at  $-0.72$  V. In contrast, the two modified electrodes displayed a relatively strong current signal response in the presence of NB. SANb-BCN/GCE and BCN/GCE gave rise to a strong reduction peak at  $-0.6$  V, which was positively offset by 120 mV from that of the bare GCE. The NB peak current attained using BCN/GCE ( $11.83 \mu\text{A}$ ) was nearly 9-fold higher than that of the bare GCE ( $1.32 \mu\text{A}$ ). Moreover, the NB peak current attained using SANb-BCN/GCE ( $55.74 \mu\text{A}$ ) was nearly 42-

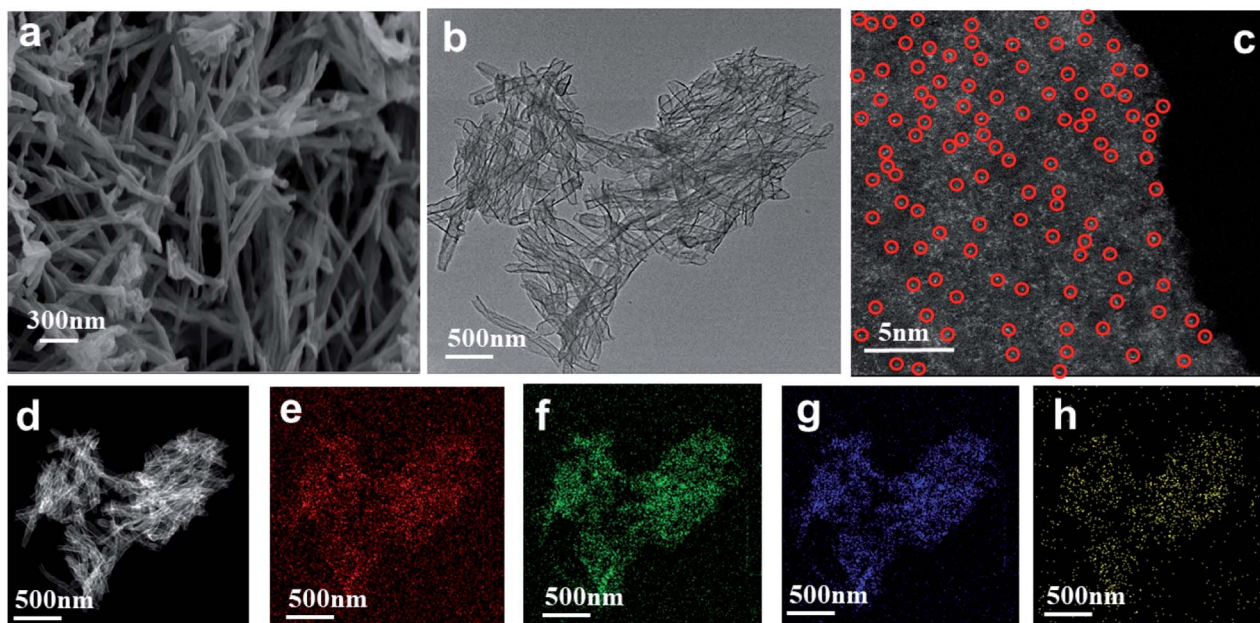


Fig. 2 (a) SEM, (b) TEM, (c) STEM images of SANb-BCN; some SANb sites are highlighted by red circles. (d–h) Elemental maps of as-synthesized SANb-BNC showing the distribution of B (e), C (f), N (g), Nb (h).





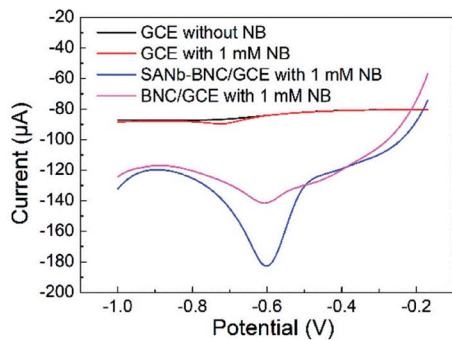


Fig. 4 LSV curves of bare GCE and SANb-BNC/GCE and BNC/GCE in 0.1 M PBS with 0.1 mM NB, and the LSV of bare GCE without NB, scan rate:  $50 \text{ mV s}^{-1}$ .

fold higher than that of the bare GCE. These electrochemical test results indicate the applicability of SANb-BCN/GCE as an electrochemical sensor for NB detection.

### Optimization of experimental conditions

To maximize the NB current signal of the modified electrode, a series of conditions were optimized. The amount of SANb-BCN on the surface of the modified electrode affects its response.<sup>37</sup> Hence, we first explored the influence of SANb-BCN concentration on the current signal. Within a certain range, the current signal intensity increased rapidly with increasing SANb-BCN concentration, reaching a maximum at a concentration of  $2 \text{ mg mL}^{-1}$  (Fig. 5a). Further increase in the concentration led to the reduction of the current signal. Interestingly, the amount of modification applied to the electrode surface had a significant impact on the current signal (Fig. 5b). As the amount of SANb-BCN dropped onto the surface of the electrode increased, the current signal continued to increase, stabilizing with the deposition of  $\sim 8 \text{ μL}$ .

The pH of a solution is indicative of the total hydrogen ion concentration and is an important parameter affecting redox activity.<sup>38</sup> Therefore, the NB reduction performance of SANb-BCN/GCE was evaluated at a pH range of 4–10; the corresponding results are shown in Fig. 5c. As the pH increased, the current signal first increased and then decreased, reaching a maximum at pH 7.0. This may be because the concentration of  $\text{H}^+$  ions affect the electrochemical performance of SANb-BCN/GCE when the pH is not neutral. Moreover, the NB reduction peak position shifted as the pH increased. Fig. 5d illustrates that the NB reduction peak potential and the pH are linearly dependent, with  $R^2 = 0.994$ . This result confirms that  $\text{H}^+$  concentration affects the electrochemical behavior of the electrode surface, and that the optimal pH for NB reduction is 7.0. In addition, by analyzing the relationship between pH and voltage, it can be concluded that NB is first reduced to phenylhydroxylamine by accepting four electrons, followed by conversion to nitrosobenzene through a two-electron oxidation process.<sup>39</sup>

With the optimized electrode and electrolyte in hand, we commenced the optimization of the experimental program,

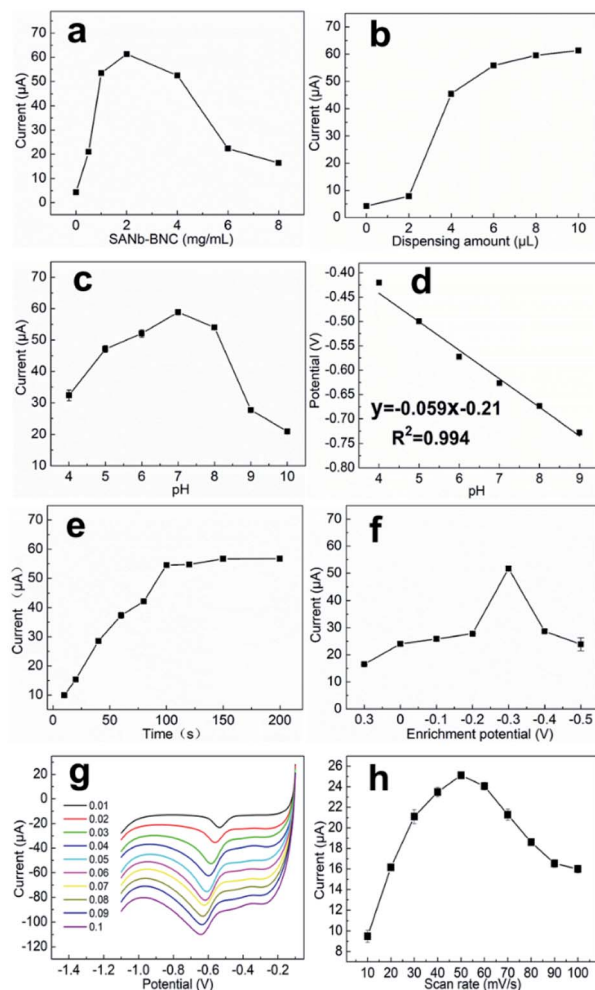


Fig. 5 Effect of SANb-BCN (a) concentration and (b) modification amount on the signal response. (c) signal response of SANb-BCN/GCE in 0.1 M PBS containing 0.1 mM NB at pH ranging from 4 to 10. (d) Relationship between pH and voltage. (e) The effect of enrichment time (10–200 s) and (f) enrichment potential (0.3, 0,  $-0.1$ ,  $-0.2$ ,  $-0.3$ ,  $-0.4$ ,  $-0.5 \text{ V}$ ) on current response. (g) LSV curves of different scan rates (10–100  $\text{mV s}^{-1}$ ), (h) the effect of different scan rates on current response.

including the enrichment time and enrichment potential. The default enrichment time of the instrument is 100 s, and the substance to be tested is generally enriched in equilibrium for this duration. Thus, we tested the current response at enrichment time in the range of 10–200 s (Fig. 5e). Evidently, when the enrichment time reached 100 s, the current response signal was balanced. As the enrichment time increased further, the current signal increased negligibly. With regard to the enrichment potential, it was found that the current signal reached a maximum at a potential of  $-0.3 \text{ V}$ , and as the voltage increased further, the current response shows a downward trend (Fig. 5f). Presumably, when the enrichment voltage exceeds a certain threshold, the oxidation of NB is no longer the dominant reaction and the incidence of side reactions increases.

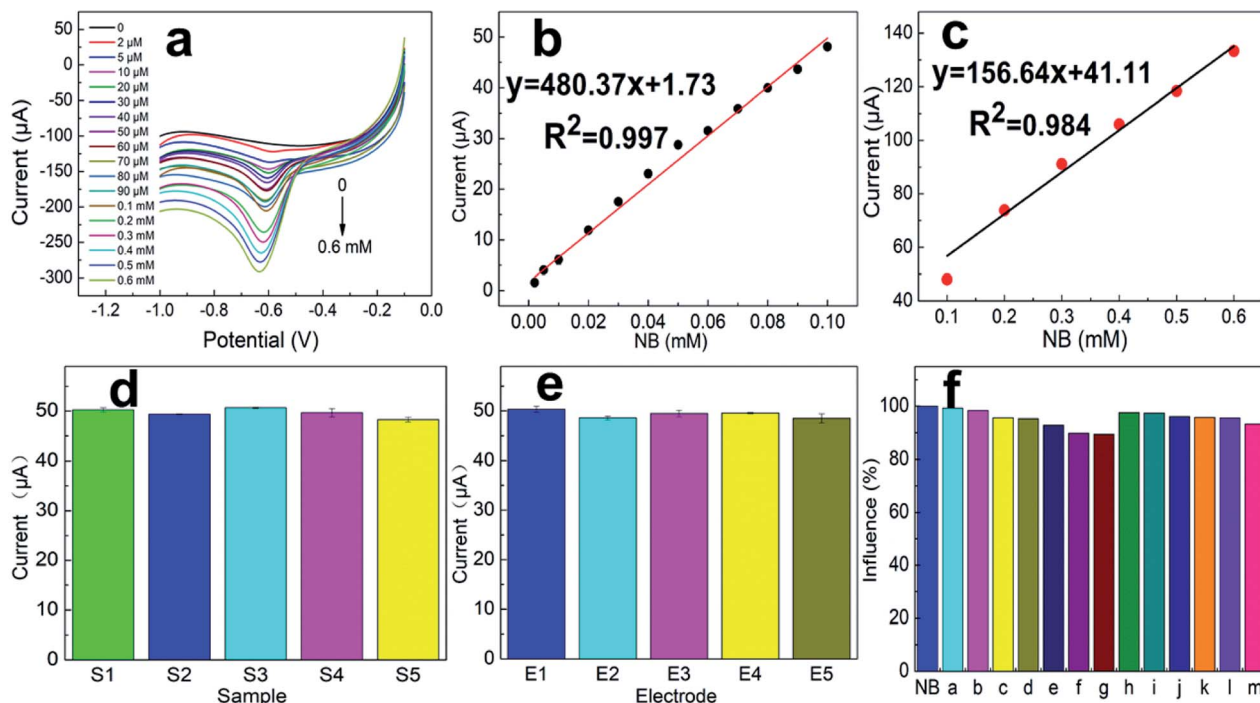


Fig. 6 (a) LSV curves obtained at NB concentrations of 0–600  $\mu\text{M}$  under optimized conditions (pH 7.0, 0.1 M PBS). Current response at NB concentrations of (b) 2–100  $\mu\text{M}$  and (c) 100–600  $\mu\text{M}$ . (d) Current signal responses of a single SANb-BCN/GCE acquired using five individual samples. (e) Current signal responses of five SANb-BCN/GCEs obtained using the same sample. (f) Current signal interference resulting from the presence of  $\text{Mg}^{2+}$ ,  $\text{Cr}^{6+}$ ,  $\text{Cu}^{2+}$ ,  $\text{K}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{NH}_4^+$ ,  $\text{Cd}^{2+}$ , urea, 1-bromo-4-nitrobenzene, 3-hydroxybenzoic, terephthalic acid, 1-iodo-4-nitrobenzene, and toluene (labeled a–m, all at a 100-times higher concentration than that of NB).

The LSV curves of different scanning rates is to analyze the specific reaction mechanism of the electrode surface. As shown in Fig. 5g and h, the LSV curves with the scan rates of 10–100  $\text{mV s}^{-1}$  presents a different current background. As the scan rate increases, the current response first increases and then decreases, and the current value reaches the maximum at 50  $\text{mV s}^{-1}$ . At the same time, as the scanning speed increases, the peak position slowly moves to a more negative potential. This is because a too fast scan rate will increase the charging current density of the electric double layer and the decrease in the solution ohmic point, which will reduce the utilization of the material. A scan rate that is too slow will cause the NB on the electrode surface to stay for too long and hinder the reduction of other NB molecules.

### Linear fit

To further understand the electrochemical NB sensing mechanism of SANb-BCN/GCE, LSV curves were acquired under optimized conditions at varying concentrations of NB. As shown in the LSV curve in Fig. 6a, the current response increased with increasing NB concentration. Based on the two graphs in Fig. 6b and c, it can be concluded that the current response is linearly correlated with NB concentration. It is worth noting that we have partitioned the linear range into two concentration ranges: 2–100  $\mu\text{M}$  and 100–600  $\mu\text{M}$ . Generally, pollutants must be adsorbed onto the sensor modified on the electrode surface to be reduced or oxidized, thereby generating a certain current response to be detected.

When the NB concentration is low, the active sites on the sensor surface are not all occupied. When the concentration of the reactant increases to a certain point, adsorption saturation is reached to a degree, and the adsorption enhancement effect gradually weakens. Hence, the current response growth rate decreases.

Therefore, the linear range was fitted sectionally, as is commonly practiced. The regression equations obtained by linear fitting are expressed as follows:

2–100  $\mu\text{M}$ :

$$I (\mu\text{A}) = 480.37C (\text{NB per mM}) + 1.73, R^2 = 0.997,$$

100–600  $\mu\text{M}$ :

$$I (\mu\text{A}) = 156.64C (\text{NB per mM}) + 41.11, R^2 = 0.984.$$

The formula  $\text{LOD} = 3S_b/S$  is used to calculate LOD, where  $S_b$  is the standard deviation of the blank experiment data and  $S$  is the slope of the calibration curve. The LOD of SANb-BCN/GCE for NB was computed to be 0.70  $\mu\text{M}$  ( $S/N = 3$ ). Several typical materials that have been developed for the detection of NB are listed in Table 1 along with their corresponding LODs, linear ranges, and linear correlation coefficients ( $R^2$ ). The LOD belonging to SANb-BCN/GCE is lower than the domestic minimum NB emission standard. Through comprehensive comparison, it can be concluded that SANb-BCN/GCE exhibits notable NB detection performance.



**Table 1** Comparison of analytical performance of several modified electrodes for electrochemical determination of NB<sup>a</sup>

Modified electrode	Detection method	Linear range ( $\mu\text{M}$ )	LOD ( $\mu\text{M}$ )	$R^2$	Reference
PAA-AgNPs/GC	Amperometry	10–600	1.68	0.9735	42
ZSO-gCN/GCE	LSV	30–100, 100–1000	2.2	0.971, 0.994	43
Ni/Fe-LDH	DPV	1–10, 10–100, 100–350	0.327	0.999, 0.984, 0.999	44
Ni/Fe(SDS)-LDH	DPV	1–10, 10–100, 100–350	0.093	0.983, 0.995, 0.999	44
UiO-66-NH <sub>2</sub>	Fluorescence quenching	0–30	0.91	0.995	6
TiO <sub>2</sub> GO-1	CV	2–8	2.64	0.968	45
TiO <sub>2</sub> GO-2	CV	0–4		0.969	45
ATP-Ag/GCE	LSV	3–30	1.1	0.997	46
BiF/CPE	SWV	1–100	0.83	0.998	47
SANb-BCN/GCE	LSV	2–100, 100–600	0.70	0.997, 0.984	This work

<sup>a</sup> PAA: poly amic acid, AgNPs: silver nanoparticles, GC: glassy carbon electrode, ZSO-gCN: stannate-graphitic carbon nitride, LDH: layered double hydroxides, SDS: sodium dodecyl sulfate, UiO-66-NH<sub>2</sub>: Zr-based metal organic frameworks, GO: oxide graphene sheets, ATP: attapulgite, BiF: bismuth-film, GPE: carbon paste electrode.

### Repeatability and reproducibility

The consistency between consecutive measurement results obtained using a single device under similar conditions represents repeatability. The consistency between the measurement results when the measurement conditions are changed indicates the reproducibility.<sup>40</sup> The repeatability and reproducibility of an electrode are indicative of its reliability and applicability. Thus, at a scan rate of 50 mV s<sup>-1</sup>, the response of one electrode was recorded in five different 0.1 M PBS solutions containing 0.1 mM NB. The results were shown in Fig. 6d; the current response error of the electrode for the five solutions was less than 5%. In addition, under the same conditions, the current signal responses of five individual electrodes to the same solution were comparable (Fig. 6e). These results indicate that SANb-BCN/GCE is capable of generating repeatable and reproducible data.

### Evaluation of anti-interference

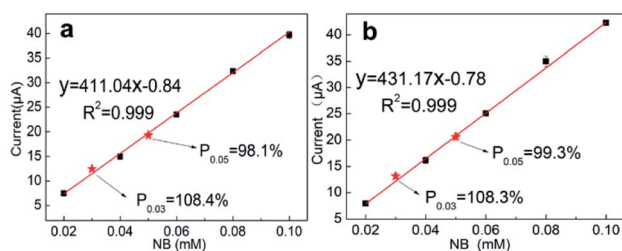
Natural water systems contain various ions and chemicals<sup>41</sup> that may affect pollutant detection results to varying degrees. The selectivity of SANb-BCN/GCE for NB was studied in the presence of various inorganic and organic substances. Thus, the same detection method and environment were used to detect the current signal response for NB in the presence of potential interference species, namely, Mg<sup>2+</sup>, Cr<sup>6+</sup>, Cu<sup>2+</sup>, K<sup>+</sup>, Ca<sup>2+</sup>, NH<sup>4+</sup>, Cd<sup>2+</sup>, urea, 1-bromo-4-nitrobenzene, 3-hydroxybenzoic,

terephthalic acid, 1-iodo-4-nitrobenzene, and toluene, all at a 100-times higher concentration than that of NB (Fig. 6f, (a–m)). Not surprisingly, the presence of the above-mentioned interfering species at such high concentrations resulted in a slight change in the NB current response, but the interference was less than 9%. This result demonstrates the excellent selectivity of SANb-BCN/GCE for NB detection.

### Real sample analysis

The recovery rate reflects the operational skill level of the analyst, but more importantly, it reflects whether the analytical method is suitable for testing a specific substrate.<sup>48</sup> To further verify the applicability of SANb-BCN/GCE, real water samples were analyzed, including tap water and lake water. NB (30  $\mu\text{M}$  and 50  $\mu\text{M}$ ) was added to the actual water samples to simulate water pollution, and the recovery rate was calculated. The tap water sample was obtained from a faucet in the laboratory of Yunnan University without dilution and filtration. Lake water was sampled from Qilu lake, one of the eight major plateau lakes in China. Because the lake water sample was relatively turbid, it was filtered through a membrane with a diameter of 0.45  $\mu\text{m}$ .

In the case of tap water (Fig. 7a),  $R^2 = 0.999$  and the recovery rates of 30 and 50  $\mu\text{M}$  NB were 108.4% and 98.1%, respectively. For lake water (Fig. 7b),  $R^2 = 0.999$ , and the recovery rates of 30 and 50  $\mu\text{M}$  NB were 108.3% and 99.3%, respectively. The value of the linear correlation coefficient and excellent recovery rates indicate that SANb-BCN/GCE is applicable for electrochemical sensing in actual water samples.



**Fig. 7** NB detection using SANb-BCN/GCE in (a) tap water and (b) Qilu lake water samples; the red dots represent the standard samples.

## Conclusions

Monoatomic Nb doped BNC nanotubes were prepared *via* the template method, and the prepared materials were modified on GCE. The resultant sensor was used for the detection of NB in aqueous media. SANb-BCN was characterized through STEM, TEM, EDS, SEM, and XRD analysis to determine its morphology and element distribution. In addition to revealing the morphology of the nanotubes, the characterization results



indicated that monoatomic Nb was well-dispersed. The mono-dispersing of Nb is advantageous for enhancing the electrochemical sensing performance of the modified electrode. Moreover, the optimal monitoring conditions were determined by optimizing the modification amount and concentration, enrichment time and potential, and pH. Furthermore, a good linear relationship, relatively low LOD, and a relatively wide linear range were achieved using the developed sensing method. Finally, the successful detection of NB in lake and tap water samples confirmed the practical utility of SANb-BCN/GCE. Owing to the high selectivity of SANb-BCN/GCE for NB and the low LOD, its application for the detection of NB in the environment is feasible. Therefore, the proposed detection system can be used for on-site environmental sample analysis.

## Author contributions

Meng Li: investigation, validation, original draft writing. Xianyun Peng and Xijun Liu: investigation and validation. Huaisheng Wang: investigation, methodology. Shusheng Zhang and Guangzhi Hu: supervision, resources, writing-review.

## Conflicts of interest

There are no conflicts to declare.

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