


 Cite this: *RSC Adv.*, 2021, 11, 33734

Solution combustion synthesis of Ni/La₂O₃ for dry reforming of methane: tuning the basicity *via* alkali and alkaline earth metal oxide promoters†

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The production of syngas *via* dry reforming of methane (DRM) has drawn tremendous research interest, ascribed to its remarkable economic and environmental impacts. Herein, we report the synthesis of K, Na, Cs, Li, and Mg-promoted Ni/La₂O₃ using solution combustion synthesis (SCS). The properties of the catalysts were determined by N₂ physisorption experiments, scanning electron microscopy (SEM), transmission electron microscopy (TEM), X-ray diffraction (XRD), X-ray photoelectron spectrometry (XPS), and H₂-TPR (temperature programmed reduction). In addition, their catalytic performance towards DRM was evaluated at 700 °C. The results demonstrated that all catalysts exhibited porous structures with high specific surface area, in particular, Mg-promoted Ni/La₂O₃ (Mg–Ni–La₂O₃) which depicted the highest surface area and highest pore volume (54.2 m² g⁻¹, 0.36 cm³ g⁻¹). Furthermore, Mg–Ni–La₂O₃ exhibited outstanding catalytic performance in terms of activity and chemical stability compared to its counterparts. For instance, at a gas hourly space velocity (GHSV) of 30 000 mL g⁻¹ h⁻¹, it afforded 83.2% methane conversion and 90.8% CO₂ conversion at 700 °C with no detectable carbon deposition over an operating period of 100 h. The superb DRM catalytic performance of Mg–Ni–La₂O₃ was attributed to the high specific surface area/porosity, strong metal-support interaction (MSI), and enhanced basicity, in particular the strong basic sites compared to other promoted catalysts. These factors remarkably enhance the catalytic performance and foster resistance to coke deposition.

 Received 18th July 2021
 Accepted 11th October 2021

DOI: 10.1039/d1ra05511a

rsc.li/rsc-advances

1. Introduction

The dramatic increment in energy demand and rapid depletion of fossil fuels have spurred remarkable efforts towards green alternative energy sources. One of the dedicated approaches to address these issues is the dry reforming of methane (DRM) in which two greenhouse gases *i.e.* carbon dioxide and methane are converted to syngas (CO + H₂), the feedstock for synthesis of methanol and other value-added chemicals. DRM is a highly endothermic reaction that becomes spontaneous at 640 °C.¹ DRM can be catalysed by noble elements such as Rh, Ir, Pt, Pd, and Ru, which afford outstanding catalytic activity and stability.^{2–6} However, their high cost and scarcity prohibit wide-scale adoption.⁷ Several transition metals such as Co, Ni, and Fe have been suggested as substitutes for noble metals. Among them, Ni has drawn remarkable interest owing to its low cost and high catalytic activity. However, the deactivation of active

sites induced by high temperature sintering and extensive deposition of coke are considered as the main stumbling blocks towards commercialization.⁸ Different strategies were employed to improve the performance of Ni-based catalysts towards DRM, such as alloying of Ni with other metals,⁹ addition of promoters,¹⁰ and confinement of Ni into mesoporous support.¹¹

According to previous reports, there is a consensus view that metal-support interaction (MSI) remarkably influence the catalytic performance of the DRM catalyst. Based on this, different metal oxides have been hired as supports for Ni-based DRM catalysts such as SiO₂,¹² Al₂O₃,¹³ TiO₂,¹¹ CeO₂,¹⁴ ZrO₂,¹⁵ and La₂O₃.¹² Among them, La₂O₃ received great interest ascribed to its parallel function as support and promoter, where it is capable to afford basic sites that enhance the adsorption of CO₂ *via* formation of La₂O₂CO₃. This endows the removal of deposited coke through the following reaction: La₂O₂CO₃ + C → La₂O₃ + 2CO.^{16,17} Concurrently, La₂O₃ enables strong interaction with Ni which promotes high dispersion of Ni particles, reduces the Ni particle size, and prevents the high temperature sintering.¹⁸

The addition of basic promoters, in particular, the alkaline earth metals oxides remarkably enhance the catalytic performance of Ni-based catalysts.¹⁹ Among these promoters, MgO received great interest attributed to its remarkable catalytic

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ra05511a



performance and enhanced coking resistance.¹⁰ For example, the catalytic performance of Ni-supported Mg–La mixed oxides (with different La³⁺/Mg²⁺ ratios) towards DRM was studied.²⁰ Activity and stability measurements depicted that Ni/10MgO–La₂O₃ afforded the best performance and mainly forms monoclinic La₂O₂CO₃, which enhanced the coke removal contrarily to the hexagonal phase, which has no impact on the catalytic activity.²⁰ In another work, the DRM performance of MgO-promoted Ni/Al₂O₃ catalyst synthesized by loading of MgO on Ni/Al₂O₃ prepared by atomic layer deposition of Ni on Al₂O₃ nanoparticles was studied.²¹ It was found that the addition of MgO enhanced the amount and strength of the catalyst basic sites, and increase the intensity of surface oxygenated species that enhance the adsorption and activation of CO₂. In addition, MgO promoted the resistance against coke formation, in particular, graphitic carbon, which is responsible for catalyst deactivation.²¹ Although the promotion of Ni/La₂O₃ with alkaline earth metals was previously investigated in the literature, however, the impact of alkali metals was not emphasized enough.

In terms of the synthesis approach, solution combustion synthesis (SCS) received increased interest as it permits the synthesis of wide range of materials in nanoscale dimensions such as metal oxides, sulfides, phosphates, metals and alloys.²² SCS allows the controlling of size, composition, and nano-architecture of materials through self-sustained exothermic reactions. This synthesis approach endows several merits such as simplicity, low cost, high porosity of synthesized materials, and small particle size.²³ Nevertheless, several preparation routes were reported for the synthesis of Ni/La₂O₃ catalysts such as incipient wetness impregnation method²⁴ and sol–gel approach,²⁵ whereas its synthesis *via* solution combustion procedure was not previously investigated.

Triggered by the above discussions, herein, we introduce the synthesis of alkali metals oxides and MgO-promoted Ni/La₂O₃ using SCS for DRM. The freshly reduced samples as well as spent samples (after DRM operation) were investigated using different structural techniques. Besides, the catalytic activity and stability of studied catalysts was evaluated at 700 °C.

2. Experimental

2.1. Materials and reagents

Nickel nitrate hexahydrate (≥98.5%), sodium nitrate (≥98.5%), magnesium nitrate hexahydrate (≥99.0%), lithium nitrate, cesium nitrate (99.0%), potassium nitrate (99.0%), and glycine (≥98.5%) were purchased from Sigma–Aldrich. Lanthanum nitrate hexahydrate (99.99%) was purchased from Carl–Roth, Germany.

2.2. Catalyst synthesis

2.2.1. Synthesis of Ni–La₂O₃. To a beaker containing 50 mL DI H₂O, (0.291 g, 1 mmol) of nickel nitrate hexahydrate, (2.165 g, 5 mmol) of lanthanum nitrate hexahydrate, and (0.788 g, 10.5 mmol) of glycine were added. The mixture was stirred for 60 minutes until it became homogeneous.

Afterwards, the solution was heated at 250 °C until complete dryness, ignition, and combustion took place. The as-formed solid was collected and calcined in air at 550 °C for 2 h at a ramp rate of 1 °C min^{−1}.

2.2.2. Synthesis of X–Ni–La₂O₃ (X = Cs, Li, K, Na, or Mg). The promoted Ni–La₂O₃ was synthesized with the same procedure of Ni–La₂O₃ with addition of 0.5 mmol of the nitrate precursor of the promoting species.

2.3. Catalyst characterization

The catalysts morphology was investigated by field emission scanning electron microscopy (FESEM) using Philips XL-30 microscope. The morphology of reduced catalysts was investigated using bright field transmission electron microscopy (TEM) *via* FEI Tecnai G2 TF20 UT microscope at an accelerating voltage of 200 kV. The samples composition was investigated by inductively coupled plasma optical emission spectrometry (ICP-OES) *via* a spectrometer (PerkinElmer, Optima 5300 DV). The textural properties were examined *via* N₂ sorption experiments at liquid nitrogen temperature (77 K) using the Brunauer–Emmett–Teller (BET) method. X-ray diffraction (XRD) spectra was recorded using X’Pert-Pro MPD diffractometer (PANalytical Co., Netherlands) using of Cu–K α X-ray source ($\lambda = 1.54059 \text{ \AA}$) as a radiation source in the 2θ range (10–80). The chemical composition and the elemental oxidation states were investigated using XPS spectrophotometer Kratos Axis Ultra XPS equipped with a monochromatic Al–K α radiation source (1486.6 eV) under ultra-high vacuum (UHV) (*ca.* 5×10^{-9} torr).

The thermogravimetric analysis was executed using TGA 4000 analyzer (PerkinElmer, USA). Measurements were performed at the temperature range 25–850 °C at a heating rate of 10 °C min^{−1} under air flow. Samples reducibility was examined using temperature-programmed reduction of hydrogen (H₂-TPR). Measurements were carried out with Micromeritics Autochem 2920 chemisorption analyzer with H₂ uptake recorded *via* TCD detector. 50 mg of the supported catalyst was placed into a quartz U-shape tube and 10% H₂ balanced in Ar was passed over the test sample in a flow rate of 10 mL min^{−1} until a stable baseline is attained. The sample was then heated in a ramp rate of 5 °C min^{−1} in the temperature range 30–850 °C. The basicity was examined by CO₂-TPD. 50 mg of the sample was first reduced in a flow of 10% H₂ at 800 °C for 2 h, then the sample was cooled down to room temperature. Then, the reduced sample was degassed at 300 °C in a flow of Ar for 1 h followed by cooling to 50 °C. Afterwards, 10% CO₂ is injected at a rate of 30 mL min^{−1} for 1 h at 50 °C followed by purging of Ar to remove excess CO₂, then the sample was heated under Ar flow at a ramp rate 10 °C min^{−1} to 900 °C and the signal of desorbed CO₂ is detected by TCD.

2.4. DRM activity

DRM activity measurements were carried out in a fixed bed quartz tube reactor ($d = 6 \text{ mm}$) at atmospheric pressure. 100 mg of the catalyst was loaded and fixed by quartz wool. Reactor temperature was recorded *via* K-type thermocouple located at the fixed bed. The catalyst was Pre-reduced with 10% H₂ flow at



800 °C for 2 h. After which, N₂ was purged to discard excess H₂. The feed gas composition was (10% CH₄ and 10% CO₂ balanced with Ar) and was fed to the reactor in a rate of 50 mL min⁻¹ giving rise to a weight hourly space velocity (WHSV) of 30 000 mL g⁻¹ h⁻¹. The composition of the outlet gas was investigated *via* a gas chromatograph (Agilent 7890B, Agilent Technologies, USA) coupled to a thermal conductivity detector (TCD). The catalytic performance was evaluated based on the change in the concentrations of CH₄, CO₂ in the inlet and outlet flow mixtures according to:

$$\% \text{conversion of CH}_4 (X_{\text{CH}_4}) = \frac{F_{\text{CH}_4(\text{in})} - F_{\text{CH}_4(\text{out})}}{F_{\text{CH}_4(\text{in})}} \times 100$$

$$\% \text{conversion of CO}_2 (X_{\text{CO}_2}) = \frac{F_{\text{CO}_2(\text{in})} - F_{\text{CO}_2(\text{out})}}{F_{\text{CO}_2(\text{in})}} \times 100$$

$$\% \text{H}_2 \text{ yield } (Y_{\text{H}_2}) = \frac{F_{\text{H}_2(\text{out})}}{2[F_{\text{CH}_4(\text{in})} - F_{\text{CH}_4(\text{out})}]} \times 100$$

$$\% \text{CO yield } (Y_{\text{CO}}) = \frac{F_{\text{CO}(\text{out})}}{F_{\text{CH}_4(\text{in})} + F_{\text{CO}_2(\text{in})}} \times 100$$

$$R_{\text{H}_2/\text{CO}} = \frac{F_{\text{H}_2(\text{out})}}{F_{\text{CO}(\text{out})}} \times 100$$

where $F_{i(\text{in})}$ and $F_{i(\text{out})}$ are the molar flow rate of inlet and outlet of species i , respectively.

3. Results and discussions

Un-promoted and promoted Ni-La₂O₃ were prepared using one-pot solution combustion synthesis. This procedure has several structural merits such as simplicity, energy effectiveness, and uniform distribution of components into the material's matrix, facile control of composition *via* change of oxidizer-to-fuel ratio.²⁶ The Ni content of as-prepared catalysts was investigated by ICP-OES. The wt% of Ni showed comparable values ranging between 6.63–7.14% (Table 1). The morphology of as-synthesized catalysts was explored by SEM (Fig. 1).

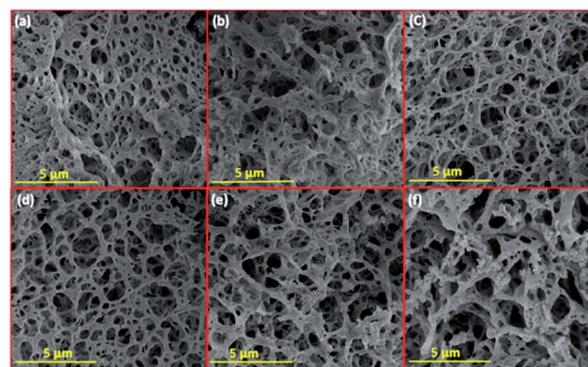


Fig. 1 SEM images of fresh reduced samples; (a) Li-Ni-La₂O₃, (b) Na-Ni-La₂O₃, (c) K-Ni-La₂O₃, (d) Mg-Ni-La₂O₃, (e) Cs-Ni-La₂O₃, and (f) Ni-La₂O₃.

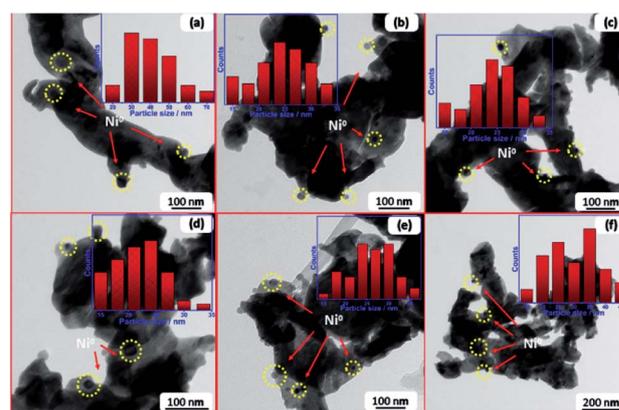


Fig. 2 TEM images and particle size distribution of reduced samples; (a) Li-Ni-La₂O₃, (b) Na-Ni-La₂O₃, (c) K-Ni-La₂O₃, (d) Mg-Ni-La₂O₃, (e) Cs-Ni-La₂O₃, and (f) Ni-La₂O₃.

All samples depict low-density spongy-like morphology with remarkable porous architecture containing large density of voids. This morphology can be attributed to the liberation of gases within short duration during the combustion process. Samples show similar morphology except Mg-Ni-La₂O₃, which exhibited higher density of voids with smaller size, which is an indication to higher porosity compared to other catalysts.

TEM of freshly reduced samples was investigated to get more deep insights on the structural properties of catalysts. The TEM

Table 1 Structural properties of the investigated catalysts

Catalyst	wt% of Ni	Specific surface area S_{BET} (m ² g ⁻¹)	Pore volume (cm ³ g ⁻¹)	Pore radius (nm)	Ni average particle size ^a (nm)	Crystallite size ^b (nm)
Ni-La ₂ O ₃	7.14	37.3	0.18	3.7	10.3	10.8
Cs-Ni-La ₂ O ₃	6.94	32.9	0.15	2.1	13.4	14.0
K-Ni-La ₂ O ₃	6.69	29.7	0.16	2.1	11.6	12.1
Li-Ni-La ₂ O ₃	6.63	42.9	0.24	2.5	14.3	14.9
Mg-Ni-La ₂ O ₃	6.78	54.5	0.36	2.8	14.8	15.6
Na-Ni-La ₂ O ₃	6.95	26.8	0.13	2.2	13.0	13.6

^a Calculated from TEM. ^b Calculated from XRD.



images of reduced samples are shown in Fig. 2. They demonstrate good dispersion of Ni nanoparticles in the matrix of support, whereas the promoter nanoparticles cannot be detected owing to high dispersion within the support. The particle size distribution of Ni (Fig. 2 inlets) was estimated in all cases and the average size is given in Table 1. Results depicted that the incorporation of promoter did not enhance the dispersion of Ni in the La₂O₃ matrix, so far, the particle size of Ni increases into the promoted samples compared to pristine Ni–La₂O₃. This may be attributed to partial weakening in metal-support interaction encountered by the insertion of promoter into the La₂O₃ matrix.²⁷

Fig. 3 displayed N₂ adsorption–desorption isotherms and pore size distributions of investigated Ni-based catalysts. According to the International Union of Pure and Applied Chemistry (IUPAC) classification, all catalysts exhibited type IV isotherms with H3 type hysteresis loop, which is characteristic feature for mesoporous materials.²⁸ Pristine Ni–La₂O₃ and promoted-catalysts exhibited similar shape of the isotherm with a hysteresis loop extending from over a wide range of relative pressure P/P^0 of 0.2 to 1, which implies a non-uniformity of pore structure and a wide pore size distribution. The similarity in textural properties between Ni–La₂O₃ and X–Ni–La₂O₃ reflects that the addition of promoter has no remarkable impact of the porous structure.²⁹ The specific surface area and pore dimensions (pore volume and pore diameter) were calculated using BJH (Barrett, Joyner, and Halenda) method and the data are given in Table 1. According to the obtained results, Mg–Ni–La₂O₃ exhibited the highest surface area and highest pore volume compared to other samples. This means that a greater number of accessible active sites available for the reactants, which can enhance the catalytic activity of Mg–Ni–La₂O₃ compared to other catalysts.

Fig. 4a represents wide-angle XRD diffraction patterns of promoted and un-promoted Ni–La₂O₃ after reduction in H₂ at 800 °C for 2 h. All samples revealed the presence of the diffractions at 2θ of 26.1, 29.1, 30.0, 39.5, 46.1, 52.2, 53.7, 55.4, 56.0, 60.4, 62.3, 66.8, 72.2, 73.5, 75.3, and 79.2, which are corresponding to (100), (002), (101), (102), (110), (103), (200), (112), (201), (004), (202), (104), (203), (210), (211), and (114) which are the characteristic peaks of La₂O₃ (JCPDS card no. 05-0602),

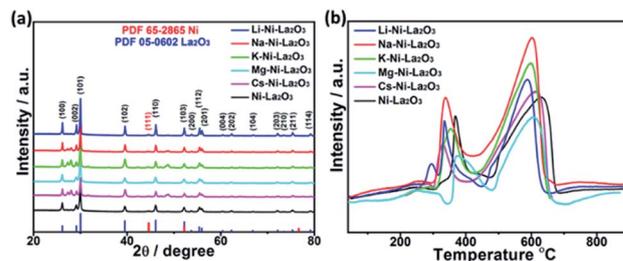


Fig. 4 (a) XRD diffraction patterns of reduced catalysts and (b) H₂-TPR profiles of as-prepared catalysts.

respectively. In addition, all samples exhibited a small diffraction peak at about 44.5°, corresponding to (111) diffraction of metallic Ni (JCPDS card no. 65-2865). No additional peaks were observed in the XRD patterns of promoted catalysts compared to un-promoted Ni–La₂O₃. This affirms that promoters have very small particle size and/or high dispersion in the La₂O₃ matrix, which is consistent with previous studies.³⁰ The average crystallite size of different catalysts was estimated by Scherrer's equation using Ni(111) peak. The calculated values are given in Table 1. The results show that incorporation of promoters slightly increases Ni crystallite size, which is consistent with the data obtained from TEM.

Samples reducibility and redox properties of freshly prepared materials was probed by H₂-TPR. Fig. 4b depicts the H₂-TPR profiles of as-prepared Ni–La₂O₃ and promoted Ni–La₂O₃ samples. Similar reduction profiles were obtained for un-promoted and promoted samples (except Li–Ni–La₂O₃). Excluding Li-promoted catalyst, all other samples depicted two reduction peaks; the first peak in the temperature range of 320–410 °C, which is corresponding to weakly interacting NiO.^{31,32} Whereas, the second peak observed above 600 °C can be assigned to strongly interacting NiO in the form of perovskite.^{33,34} Li–Ni/La₂O₃ demonstrated three reduction peaks in the H₂-TPR profile, the first peak can be attributed to the reduction of surface weakly interacting NiO species, the second peak can be assigned to bulk NiO, and the third peak corresponds to the reduction of strongly interacting NiO species in the form of perovskite. Furthermore, the lithium-promoted sample exhibits low reduction temperatures compared to other samples. The different behavior encountered by Li–Ni–La₂O₃ may arise from that Li₂O remarkably weakens the interaction between Ni and La₂O₃, which is consistent to previous studies.^{35,36} Compared to un-promoted Ni–La₂O₃, all promoted samples reveal a negative shift in the reduction temperatures. This is consistent to the previous studies, which implies the same trend.³⁷ This was attributed to competition between Ni and promoter species on the interaction with La₂O₃ support.³⁷ This means that the insertion of promoter slightly decreases the interaction of Ni with La₂O₃ and shift the reduction of Ni species towards lower temperatures.³⁸ Compared to other promoted catalysts, Mg–Ni–La₂O₃ demonstrated higher reduction temperatures, which affirms stronger MSI.

The surface composition and chemical entities of the freshly reduced catalysts were explored by XPS analysis. The XPS survey

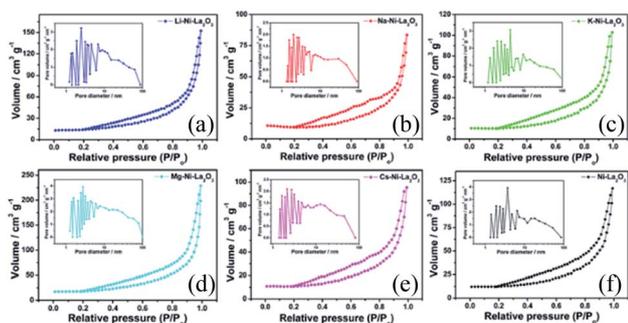


Fig. 3 N₂ adsorption/desorption isotherms and pore size distribution (inlet) of (a) Li–Ni–La₂O₃, (b) Na–Ni–La₂O₃, (c) K–Ni–La₂O₃, (d) Mg–Ni–La₂O₃, (e) Cs–Ni–La₂O₃, and (f) Ni–La₂O₃.



spectra of reduced catalysts depicted the presence of the promoter species in each case together with La, Ni, and O, which confirms their chemical composition (see Fig. 5 and S1†). The binding energies of Ni, La, and O were not significantly which confirms that their electronic environment was not affected by modification with the promoters (Fig. 5a–c).

Fig. 6a demonstrates the high resolution spectra of Ni 2p in the reduced Mg–Ni–La₂O₃ catalyst. Deconvolution of Ni 2p region depicted two components. The low binding energy components at about 852.4 eV is assigned to metallic Ni and the other at higher binding energy 856.3 eV corresponds to NiO.³⁹ The values of binding energies are relatively higher than normal values, which implies strong interaction between Ni and La₂O₃ support.⁴⁰ Deconvoluted high resolution spectra of O 1s region manifested two peaks. The first peak at 530.4–531 eV is attributed to lattice oxygen, whereas, the other peak at about 532.5 eV can be assigned to adsorbed oxygen species such as hydroxyls and carbonates.⁴¹

The basicity of samples was investigated by CO₂-TPD. The desorption profiles are shown in Fig. 7. Three sets of peaks were observed. The desorption peaks below 400 °C can be attributed to weakly basic sites, the peaks 400–600 °C are assigned to basic sites of moderate strength, and the at temperatures higher than 600 °C corresponds to strong basic sites.^{42,43} The total amount of desorbed CO₂ is larger in Mg–Ni–La₂O₃ compared to other catalysts, especially CO₂ desorbed from strong basic sites. This implies greater amount and enhanced strength of basic sites which can facilitate the adsorption and activation of CO₂ and enhance the DRM activity.

To investigate the effect of promoter type on the catalytic performance of Ni–La₂O₃, the DRM activity of investigated samples was examined at 700 °C for an operating time of 100 h using GHSV of 30 L g⁻¹ h⁻¹. Prior to the reaction, samples were reduced by 10% H₂/Ar at 800 °C for 2 h to convert Ni species to active metallic Ni. After that, the gas mixture was purged through the fixed bed reactor and the catalytic activity was tested by evaluating the percentage of conversions of CO₂ and CH₄. Fig. 8 depicted the catalytic activity of promoted and unpromoted catalysts displayed in terms of variation of CO₂ conversion, CH₄ conversion, and H₂/CO ratio with time-on-stream. The reaction was maintained for 100 h to explore long-term stability. Interestingly, all samples revealed slow increase in the conversion of CO₂ and CH₄ with time. According to previous reports, this increase in activity with time can be assigned to the incomplete reduction of Ni oxide species during the pre-treatment step.⁴⁴ Other reports ascribed this

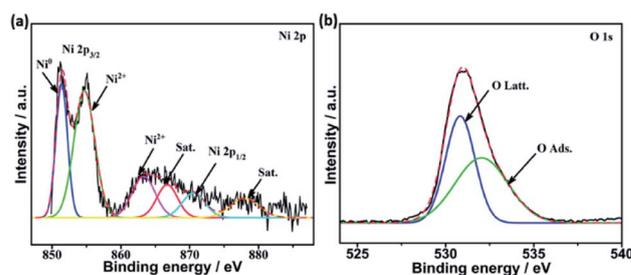


Fig. 6 Deconvoluted high resolution XPS spectra of (a) Ni 2p and (b) O 1s in reduced Mg–Ni–La₂O₃.

enhancement during the initial conduction period to the slow establishment of equilibrium concentration of La₂O₂CO₃, which is generated as an intermediate from the reaction of La₂O₃ with CO₂.^{20,45} Intriguingly, all catalysts manifested low initial conversions for CH₄ and CO₂ followed by subsequent increase with time-on-stream until a steady state was attained which is consistent with previous studies.⁴⁶ After this steady state CO₂ and CH₄ conversions remains constant until the end of operating time (100 h) with decay in the activity observed in CO₂ and CH₄ conversions in case of Ni–La₂O₃ and Li–Ni–La₂O₃. For all catalysts, CO₂ conversions exhibited higher values than CH₄ and the H₂/CO ratio is less than unity. This can be attributed to the concurrent reverse water gas shift (RWGS), CO₂ + H₂ → CO + H₂O.⁴⁷

It is well known that the ratio of H₂/CO is influenced by DRM as well as the co-occurring side reactions (Boudouard, RWGS, and methane decomposition).¹² The actual ratio of H₂/CO is affected by coexisting side reactions such as RWGS, methane decomposition, and Boudouard reaction. All catalysts displayed a H₂/CO ratio, which is fluctuating around almost constant value (Fig. 8c). This fluctuation is attributable to the concurrent carbon deposition and carbon gasification reactions.⁴⁸ The value of H₂/CO for all catalysts lies between 0.8 and 0.9, which affirms that RWGS is the dominating side reaction.

All promoted catalysts exhibited higher CO₂ and methane conversions compared to Ni/La₂O₃, which provides a clear evidence on the role of promoter in enhancing the catalytic activity. The maximum CO₂ conversion follows the order: Mg–Ni–La₂O₃ (90.8%) > Na–Ni/La₂O₃ (86.9%) > K–Ni–La₂O₃ (83.7%) > Cs–Ni–La₂O₃ (83.0%) > Li–Ni–La₂O₃ (80.6%) > Ni–La₂O₃ (74.4%). Similarly, the maximum CH₄ conversion follows the

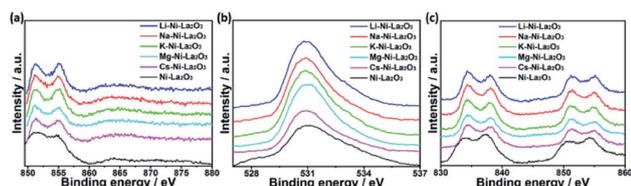


Fig. 5 High resolution XPS spectra of (a) Ni 2p, (b) O 1s, and (c) La 3d in reduced samples.

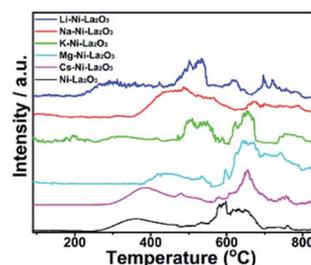


Fig. 7 CO₂-TPD profiles of investigated catalysts.



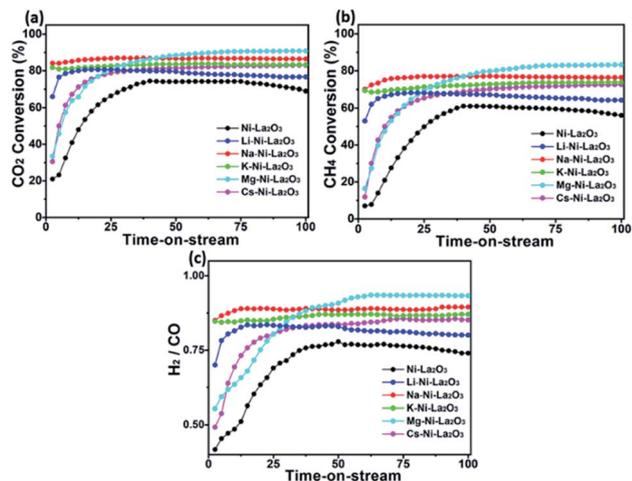


Fig. 8 DRM catalytic performance measurements; (a) carbon dioxide conversion, (b) methane conversion, and (c) H₂/CO ratio measured at 700 °C.

same order with values of 83.2%, 76.3%, 73.7%, 72.5%, 64.3%, and 56.0%, in case of Mg-Ni-La₂O₃, Na-Ni-La₂O₃, K-Ni-La₂O₃, Cs-Ni-La₂O₃, Li-Ni-La₂O₃, and Ni-La₂O₃, respectively. After reaching the maximum conversion, all catalysts demonstrated negligible decay in the activity until the end of DRM operating time (100 h) except Li-Ni-La₂O₃ and un-promoted catalyst, which revealed a significant decay in the activity with time. Li-Ni-La₂O₃ demonstrated 5.0 and 5.7% decay in the CO₂ and CH₄ conversion, respectively, while Ni-La₂O₃ exhibited 7.3% and 21.3% in the conversions of CO₂ and CH₄, respectively. This provides a clear evidence for the stabilizing impact of promoting species on the long-term stability of catalysts compared to Li-Ni-La₂O₃ and Ni-La₂O₃.

Fig. 9 displays the TEM images and Ni particle size distributions of spent catalysts after running for 100 h at 700 °C. Some Ni particles were agglomerated owing to high temperature sintering. The average particle size of Ni in the spent samples is 23.1, 30.1, 26.3, 34.2, 29.4, and 28.6 nm in case of

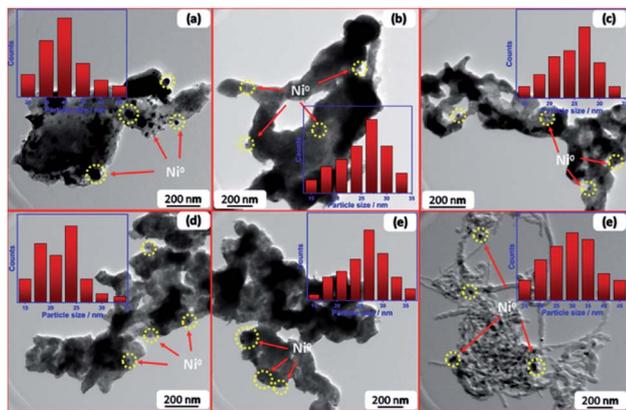


Fig. 9 TEM images of spent samples; (a) Li-Ni-La₂O₃, (b) Na-Ni-La₂O₃, (c) K-Ni-La₂O₃, (d) Mg-Ni-La₂O₃, (e) Cs-Ni-La₂O₃, and (f) Ni-La₂O₃.

Mg-Ni-La₂O₃, Na-Ni-La₂O₃, K-Ni-La₂O₃, Li-Ni-La₂O₃, Cs-Ni-La₂O₃, and Ni-La₂O₃, respectively. Based on the ratio of change in the Ni particle size during the reaction, Li-Ni-La₂O₃ and Ni-La₂O₃ demonstrated the greatest increase in the Ni particle size, whereas, Mg-Ni-La₂O₃ and Na-Ni-La₂O₃ revealed the lowest ratio of increase in the Ni particle size (Fig. 9). This can explain the rapid decay in the DRM activity with time in case of un-promoted and Li-promoted samples compared to other catalysts. Owing to the increase in the particle size by sintering, the number of Ni active centers decreased by agglomeration. This decreased the catalytic activity and diminished the long-term stability. This was confirmed by TEM of spent samples, which manifested the formation of filamentous type carbon in the un-promoted catalyst in contrast to all promoted samples that reveal no observable coke formation (Fig. 9).

Fig. 10a depicts XRD of spent samples after DRM operating time of 100 h at 700 °C. Similar diffraction patterns were obtained by promoted Ni-La₂O₃ and no observable peaks were obtained for the alkali metal oxides. New diffractions peaks were observed at 2θ of 15.6, 27.5, 28.0, 42.5, 47.4, 48.6, 50.1, 69.9, 70.9, and 75.7° corresponding to La(OH)₃ (JCPDS card no. 36-1481). La(OH)₃ was formed as a result of reaction between La₂O₃ and H₂O that produced by RWGS *i.e.* La₂O₃ + 3H₂O → 2La(OH)₃. In addition, new diffractions were observed at 14.2, 23.1, and 29.2° which can be indexed to La₂O₂CO₃ (JCPDS card no. 48-1113). During the DRM reaction, La₂O₃ reacts with CO₂ to form La₂O₂CO₃.⁴⁹ Unfortunately, we couldn't calculate the Ni crystallite size in the spent samples from XRD since the Ni(111) diffraction peak overlapped with the diffractions of La₂O₂CO₃ at about 44.5°. Furthermore, the graphite peak at 2θ of 26° could not be identified owing to the interference with (100) diffraction peak of La₂O₃ (JCPDS card no. 05-0602).

Fig. 10b represents the TGA profiles of spent catalysts after DRM 100 h duration time at 700 °C. The slight weight loss before 200 °C can be assigned to the release of physisorbed water in La(OH)₃.⁵⁰ The slight weight gain in the temperature range 300–400 °C can be ascribed to the oxidation of metallic Ni to NiO.⁵¹

The weight decline in the temperature range 400–600 °C owing to oxidation of coke. Only un-promoted sample revealed a noticeable weight loss in 400–600 °C, which is attributed to the oxidation of filamentous carbon, whereas, no weight loss was observed for all promoted samples in the same temperature range owing to the absence of coke deposition, which is

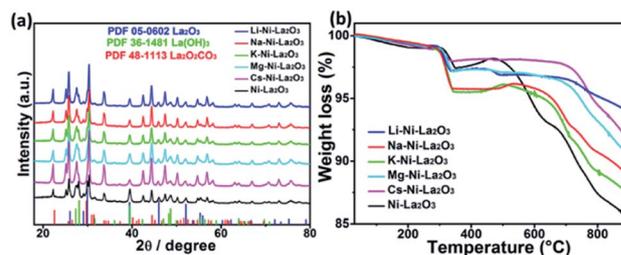


Fig. 10 (a) XRD patterns and (b) TGA profiles of spent samples after 100 h operation time at 700 °C.

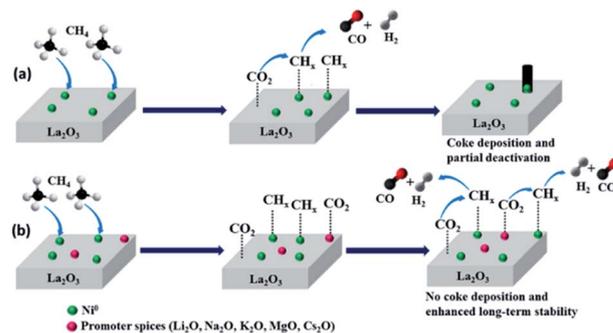


consistent to TEM results of the spent samples. On the other hand, the mass loss above 700 °C can be attributed to the decomposition of $\text{La}_2\text{O}_2\text{CO}_3$.⁵²

The XPS of spent samples was investigated to probe the change in the nature and the chemical states of different elements. High resolution spectra of C 1s can provide relevant information about the nature of carbon species, respectively at the utmost surface of the spent samples. Deconvolution of C 1s spectrum results in 3 peaks at 288.4 eV, 287.5 eV and 284.6 eV which can be assigned to CO_3^{2-} , C–O, and C–C species, respectively.⁴⁴ Whereas, the peak at lower binding energy (283.9 eV) can be assigned to Ni carbide species at the catalyst surface.⁵³ The Ni 2p peak in spent samples was shifted in all catalysts towards higher binding energies assigned to NiO, implying partial decrease in the metallic character of Ni (Fig. S2a†).⁵⁴ The differentiation between La^{3+} species in oxide and hydroxides is extremely difficult, thereon, this can be executed through the O 1s peaks.⁵⁴ It is reported that the O 1s in La_2O_3 appear at lower binding energies than O 1s in $\text{La}(\text{OH})_3$. Based on this, it could be understood that La in the oxide has a greater tendency to donate electrons to the adjacent oxygen atoms compared to the La in the hydroxide. This shifted the binding energy of oxygen to lower values.²⁰ This confirms that La_2O_3 was partially converted to $\text{La}(\text{OH})_3$ in spent samples which confirms the XRD analysis of spent catalysts (Fig. S2b†).

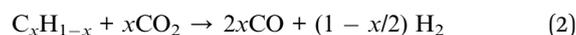
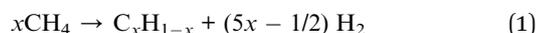
DRM requires bifunctional catalysts capable of promoting two simultaneous processes *i.e.* the cracking of methane and the oxidation of carbon species. It is well-known that deactivation of Ni-based catalysts during DRM is mainly assigned to the catalyst sintering as well as coke deposition.¹² It is also noteworthy that, metal-support interaction (MSI) remarkably influence the catalytic activity and durability of the catalyst, where it enhances the stability catalyst against sintering and also against deactivation through carbon deposition. La_2O_3 has strong MSI with Ni catalyst, so far, it enhances Ni dispersion and stability against sintering at high reaction temperatures.⁵⁵ It enhances the adsorption of CO_2 to form $\text{La}_2\text{O}_2\text{CO}_3$ which is capable of removal of deposited coke *via* the reaction: $\text{La}_2\text{O}_2\text{CO}_3 + 2\text{C} \rightarrow \text{La}_2\text{O}_3 + 2\text{CO}$.⁵⁶ However, the role of MSI is not the prevailing factor affecting the catalytic performance of the investigated catalyst. This can be evident from the inferior catalytic activity and durability of un-promoted catalyst although it has higher MSI compared to promoted catalysts. Intriguingly, the particle size of Ni in fresh reduced Ni- La_2O_3 was smaller than all promoted samples, so far, it was expected to afford comparable catalytic performance towards DRM.⁵⁷ However, Ni- La_2O_3 expressed the lowest activity and its spent catalyst showed the largest increase in the Ni particle size. Besides, only un-promoted sample manifested the formation of coke compared to all promoted samples. All these factors confirms imply that the catalyst basicity has the dominant role that affect the catalytic performance. Hence, Mg-Ni- La_2O_3 afforded the best catalytic activity and the highest stability.

According to the obtained results, the mechanism of DRM on un-promoted and promoted Ni- La_2O_3 can be represented by Scheme 1. Methane is firstly adsorbed at Ni sites and decomposed to CH_x intermediate species according to eqn (1).⁵⁸



Scheme 1 The mechanism of DRM over un-promoted and promoted catalysts.

Simultaneously, CO_2 is adsorbed especially at the more basic sites on La_2O_3 or at the promoter oxide. CH_x species are gasified according to eqn (2).^{58,59} However, the dissociation of methane on active Ni is usually a rapid process and the formed intermediate CH_x fragments may not be adjacent to activated CO_2 molecules, which results in further growth of carbon on the active Ni centers and deposition of coke.⁶⁰ The presence of promoter species not only enhance the basicity of La_2O_3 but the promoter itself can act as active sites for adsorption and activation of CO_2 . The chemisorption of CO_2 can generate CO and active oxygen species, which enhance the gasification of CH_x species and prevent the coke deposition.^{60,61} This was affirmed by the results, which revealed coke deposition in case of un-promoted sample, whereas promoted samples do not exhibit observable deposited carbon after DRM durability tests. The MSI play an important role regarding to long-term stability, so far, Li-Ni- La_2O_3 exhibited weakened MSI encountered by Li_2O , which results in sintering at high temperatures and diminishes the catalytic performance.



To summarize, the addition of promoters enhances both the catalytic activity and durability of Ni- La_2O_3 through three different roles. Firstly, the addition of MgO or other alkali metals oxides to Ni- La_2O_3 not only enhance the basicity of the catalyst but also it can create surface oxygen ionic species that enhance the stability of $\text{La}_2\text{O}_2\text{CO}_3$ phase which is liable for gasification of deposited coke, protecting Ni active centers from deactivation.^{33,61}

The greater number of oxygen vacancies and the higher oxygen mobility enhances the resistance to coke formation.⁶² Furthermore, the incorporation of alkali metals and alkaline earth metals increase the density of lattice oxygen surface species which enhances the activation of methane C–H bond and thus foster the catalytic activity.²¹ Last but not least, the promoters enhance the surface alkalinity through increasing the number of basic sites which in turn enhance the activation of CO_2 and facilitate the gasification of deposited carbon at the



active sites.⁶³ This was confirmed by the absence of carbon deposits in all spent catalysts except Ni–La₂O₃ owing to low basicity compared to Mg and other alkali metals.

4. Conclusions

Promoted and un-promoted Ni–La₂O₃ catalysts were successfully synthesized *via* one-step solution combustion route and their activity towards DRM was studied. All samples manifested high specific surface area and enhanced porosity owing to rapid ignition/combustion process and evolution of large amount of gases. In terms of DRM performance, all samples afforded high catalytic activity with CO₂ and CH₄ conversions higher than the equilibrium values at 700 °C. Promoted catalysts demonstrated enhanced catalytic activity and durability compared to un-promoted sample, which was ascribed to enhanced Lewis basicity. Particularly, Mg–Ni–La₂O₃ exhibited outstanding anti-coking capability compared to other samples. This can be assigned to two main factors. The first is the high surface area and high porosity compared to its counterparts, thereon, it can provide more active sites for DRM which enhances the catalytic activity. In addition, the incorporation of Mg afforded the increased number of basic sites while maintaining strong MSI compared to other promoters. This enhances the adsorption of CO₂ and the supply of surface oxygen species, which inhibit the coke deposition and retard the catalyst deactivation *via* high temperature sintering. This is evidenced by comparing the percentages of adsorbed oxygen in promoted and unpromoted catalysts. In the unpromoted sample, the adsorbed oxygen constitutes 12.0% of the total oxygen, while in the promoted catalysts, the ratio of adsorbed oxygen ranges between 44–54% (see Fig. S3 and Table S1†). On the other hand, among all promoted catalysts, only Li–Ni–La₂O₃ revealed decay in long-term stability. This was assigned to the high temperature sintering endowed by weakened MSI.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This publication was supported by Qatar University, internal grant number QUCG-CENG-19/20-7. The findings achieved herein are solely the responsibility of the authors. The authors also acknowledge the technical support of Central Laboratories Unit (CLU), Center for Advanced Materials (CAM), and Gas Processing Center (GPC), Qatar University, Doha, Qatar.

Notes and references

- 1 S. Wang, G. Lu and G. J. Millar, Carbon dioxide reforming of methane to produce synthesis gas over metal-supported catalysts: state of the art, *Energy Fuels*, 1996, **10**, 896–904.
- 2 M. M. Souza, D. A. Aranda and M. Schmal, Reforming of methane with carbon dioxide over Pt/ZrO₂/Al₂O₃ catalysts, *J. Catal.*, 2001, **204**, 498–511.
- 3 A. Erdoheily, J. Cserényi and F. Solymosi, Activation of CH₄ and its reaction with CO₂ over supported Rh catalysts, *J. Catal.*, 1993, **141**, 287–299.
- 4 F. Wang, L. Xu, J. Yang, J. Zhang, L. Zhang, H. Li, Y. Zhao, H. X. Li, K. Wu and G. Q. Xu, Enhanced catalytic performance of Ir catalysts supported on ceria-based solid solutions for methane dry reforming reaction, *Catal. Today*, 2017, **281**, 295–303.
- 5 M. Soria, C. Mateos-Pedrero, A. Guerrero-Ruiz and I. Rodríguez-Ramos, Thermodynamic and experimental study of combined dry and steam reforming of methane on Ru/ZrO₂–La₂O₃ catalyst at low temperature, *Int. J. Hydrogen Energy*, 2011, **36**, 15212–15220.
- 6 R. K. Singha, A. Yadav, A. Shukla, M. Kumar and R. Bal, Low temperature dry reforming of methane over Pd–CeO₂ nanocatalyst, *Catal. Commun.*, 2017, **92**, 19–22.
- 7 G. P. Figueredo, R. L. Medeiros, H. P. Macedo, Â. A. de Oliveira, R. M. Braga, J. M. Mercury, M. A. Melo and D. M. Melo, A comparative study of dry reforming of methane over nickel catalysts supported on perovskite-type LaAlO₃ and commercial α -Al₂O₃, *Int. J. Hydrogen Energy*, 2018, **43**, 11022–11037.
- 8 T. Xie, L. Shi, J. Zhang and D. Zhang, Immobilizing Ni nanoparticles to mesoporous silica with size and location control via a polyol-assisted route for coking-and sintering-resistant dry reforming of methane, *Chem. Commun.*, 2014, **50**, 7250–7253.
- 9 Z. Wu, B. Yang, S. Miao, W. Liu, J. Xie, S. Lee, M. J. Pellin, D. Xiao, D. Su and D. Ma, Lattice Strained Ni–Co alloy as a High-Performance Catalyst for Catalytic Dry Reforming of Methane, *ACS Catal.*, 2019, **9**, 2693–2700.
- 10 Q. Zhang, X. Feng, J. Liu, L. Zhao, X. Song, P. Zhang and L. Gao, Hollow hierarchical Ni/MgO–SiO₂ catalyst with high activity, thermal stability and coking resistance for catalytic dry reforming of methane, *Int. J. Hydrogen Energy*, 2018, **43**, 11056–11068.
- 11 Y. Cao, M. Lu, J. Fang, L. Shi and D. Zhang, Hexagonal boron nitride supported meso SiO₂-confined Ni catalysts for dry reforming of methane, *Chem. Commun.*, 2017, **53**, 7549–7552.
- 12 Q. Zhang, T. Zhang, Y. Shi, B. Zhao, M. Wang, Q. Liu, J. Wang, K. Long, Y. Duan and P. Ning, A sintering and carbon-resistant Ni-SBA-15 catalyst prepared by solid-state grinding method for dry reforming of methane, *J. CO₂ Util.*, 2017, **17**, 10–19.
- 13 C. A. Schwengber, F. A. da Silva, R. A. Schaffner, N. R. C. Fernandes-Machado, R. J. Ferracin, V. R. Bach and H. J. Alves, Methane dry reforming using Ni/Al₂O₃ catalysts: evaluation of the effects of temperature, space velocity and reaction time, *J. Environ. Chem. Eng.*, 2016, **4**, 3688–3695.
- 14 M. Khajenoori, M. Rezaei and F. Meshkani, Dry reforming over CeO₂-promoted Ni/MgO nano-catalyst: effect of Ni loading and CH₄/CO₂ molar ratio, *J. Ind. Eng. Chem.*, 2015, **21**, 717–722.
- 15 Y. Lou, M. Steib, Q. Zhang, K. Tiefenbacher, A. Horváth, A. Jentys, Y. Liu and J. A. Lercher, Design of stable Ni/ZrO₂



- catalysts for dry reforming of methane, *J. Catal.*, 2017, **356**, 147–156.
- 16 V. Tsipouriari and X. Verykios, Carbon and oxygen reaction pathways of CO₂ reforming of methane over Ni/La₂O₃ and Ni/Al₂O₃ catalysts studied by isotopic tracing techniques, *J. Catal.*, 1999, **187**, 85–94.
 - 17 X. E. Verykios, Catalytic dry reforming of natural gas for the production of chemicals and hydrogen, *Int. J. Hydrogen Energy*, 2003, **28**, 1045–1063.
 - 18 L. Mo, X. Zheng, Q. Jing, H. Lou and J. Fei, Combined Carbon Dioxide Reforming and Partial Oxidation of Methane to Syngas over Ni–La₂O₃/SiO₂ Catalysts in a Fluidized-Bed Reactor, *Energy Fuels*, 2005, **19**, 49–53.
 - 19 Z. Alipour, M. Rezaei and F. Meshkani, Effect of alkaline earth promoters (MgO, CaO, and BaO) on the activity and coke formation of Ni catalysts supported on nanocrystalline Al₂O₃ in dry reforming of methane, *J. Ind. Eng. Chem.*, 2014, **20**, 2858–2863.
 - 20 J. Ni, L. Chen, J. Lin, M. K. Schreyer, Z. Wang and S. Kawi, High performance of Mg–La mixed oxides supported Ni catalysts for dry reforming of methane: the effect of crystal structure, *Int. J. Hydrogen Energy*, 2013, **38**, 13631–13642.
 - 21 B. Jin, S. Li and X. Liang, Enhanced activity and stability of MgO-promoted Ni/Al₂O₃ catalyst for dry reforming of methane: role of MgO, *Fuel*, 2021, **284**, 119082.
 - 22 A. Varma, A. S. Mukasyan, A. S. Rogachev and K. V. Manukyan, Solution combustion synthesis of nanoscale materials, *Chem. Rev.*, 2016, **116**, 14493–14586.
 - 23 V. Danghyan, A. Kumar, A. Mukasyan and E. Wolf, An active and stable NiOMgO solid solution based catalysts prepared by paper assisted combustion synthesis for the dry reforming of methane, *Appl. Catal., B*, 2020, **273**, 119056.
 - 24 L. Mo, K. K. M. Leong and S. J. C. S. Kawi, A highly dispersed and anti-coking Ni–La₂O₃/SiO₂ catalyst for syngas production from dry carbon dioxide reforming of methane, *Catal. Sci. Technol.*, 2014, **4**, 2107–2114.
 - 25 C. Chen, X. Wang, L. Zhang, X. Zou, W. Ding and X. Lu, Synthesis of mesoporous Ni–La₂O₃/SiO₂ by poly(ethylene glycol)-assisted sol–gel route as highly efficient catalysts for dry reforming of methane with a H₂/CO ratio of unity, *Catal. Commun.*, 2017, **94**, 38–41.
 - 26 W. Wen and J.-M. Wu, Nanomaterials via solution combustion synthesis: a step nearer to controllability, *RSC Adv.*, 2014, **4**, 58090–58100.
 - 27 C. Montero, A. Ochoa, P. Castaño, J. Bilbao and A. G. Gayubo, Monitoring NiO and coke evolution during the deactivation of a Ni/La₂O₃– α -Al₂O₃ catalyst in ethanol steam reforming in a fluidized bed, *J. Catal.*, 2015, **331**, 181–192.
 - 28 L. Xu, H. Song and L. Chou, Carbon dioxide reforming of methane over ordered mesoporous NiO–Al₂O₃ composite oxides, *Catal. Sci. Technol.*, 2011, **1**, 1032–1042.
 - 29 N. Habibi, Y. Wang, H. Arandiyani and M. Rezaei, Biogas Reforming for Hydrogen Production: A New Path to High-Performance Nickel Catalysts Supported on Magnesium Aluminate Spinel, *ChemCatChem*, 2016, **8**, 3600–3610.
 - 30 Z. Taherian, M. Yousefpour, M. Tajally and B. Khoshandam, A comparative study of ZrO₂, Y₂O₃ and Sm₂O₃ promoted Ni/SBA-15 catalysts for evaluation of CO₂/methane reforming performance, *Int. J. Hydrogen Energy*, 2017, **42**, 16408–16420.
 - 31 I. Rivas, J. Alvarez, E. Pietri, M. J. Pérez-Zurita and M. R. Goldwasser, Perovskite-type oxides in methane dry reforming: effect of their incorporation into a mesoporous SBA-15 silica–host, *Catal. Today*, 2010, **149**, 388–393.
 - 32 L. Mo, K. K. M. Leong and S. Kawi, A highly dispersed and anti-coking Ni–La₂O₃/SiO₂ catalyst for syngas production from dry carbon dioxide reforming of methane, *Catal. Sci. Technol.*, 2014, **4**, 2107–2114.
 - 33 N. Wang, X. Yu, Y. Wang, W. Chu and M. Liu, A comparison study on methane dry reforming with carbon dioxide over LaNiO₃ perovskite catalysts supported on mesoporous SBA-15, MCM-41 and silica carrier, *Catal. Today*, 2013, **212**, 98–107.
 - 34 G. Valderrama, M. R. Goldwasser, C. U. de Navarro, J. M. Tatibouet, J. Barrault, C. Batiot-Dupeyrat and F. Martínez, Dry reforming of methane over Ni perovskite type oxides, *Catal. Today*, 2005, **107**, 785–791.
 - 35 S.-W. Park, D. Lee, S.-I. Kim, Y. J. Kim, J. H. Park, I. Heo, T. S. Chang and J. H. Lee, Effects of Alkali Metals on Nickel/Alumina Catalyzed Ethanol Dry Reforming, *Catalysts*, 2021, **11**, 260.
 - 36 F. Pompeo, N. N. Nichio, M. G. González and M. Montes, Characterization of Ni/SiO₂ and Ni/Li–SiO₂ catalysts for methane dry reforming, *Catal. Today*, 2005, **107**, 856–862.
 - 37 L. Xu, H. Song and L. Chou, Carbon dioxide reforming of methane over ordered mesoporous NiO–MgO–Al₂O₃ composite oxides, *Appl. Catal., B*, 2011, **108**, 177–190.
 - 38 J. Dias and J. Assaf, Influence of calcium content in Ni/CaO/ γ -Al₂O₃ catalysts for CO₂-reforming of methane, *Catal. Today*, 2003, **85**, 59–68.
 - 39 J. Requies, V. Barrio, J. Cambra, M. Güemez, P. Arias, V. La Parola, M. Pena and J. Fierro, Effect of redox additives over Ni/Al₂O₃ catalysts on syngas production via methane catalytic partial oxidation, *Fuel*, 2008, **87**, 3223–3231.
 - 40 N. Wang, X. Yu, K. Shen, W. Chu and W. Qian, Synthesis, characterization and catalytic performance of MgO-coated Ni/SBA-15 catalysts for methane dry reforming to syngas and hydrogen, *Int. J. Hydrogen Energy*, 2013, **38**, 9718–9731.
 - 41 J. Requies, M. Cabrero, V. Barrio, M. Güemez, J. Cambra, P. Arias, F. Pérez-Alonso, M. Ojeda, M. Peña and J. Fierro, Partial oxidation of methane to syngas over Ni/MgO and Ni/La₂O₃ catalysts, *Appl. Catal., A*, 2005, **289**, 214–223.
 - 42 A. Nagu, K. Vasikerappa, P. Gidyonu, C. Prathap, M. V. Rao, K. S. R. Rao and B. D. Raju, Additive-free vapour-phase hydrogenation of benzonitrile over MgO-supported Ni catalysts, *Res. Chem. Intermed.*, 2020, **46**, 2669–2681.
 - 43 L. Xu, W. Liu, X. Zhang, L. Tao, L. Xia, X. Xu, J. Song, W. Zhou, X. Fang and X. Wang, Ni/La₂O₃ Catalysts for Dry Reforming of Methane: Insights into the Factors Improving the Catalytic Performance, *ChemCatChem*, 2019, **11**, 2887–2899.
 - 44 M. Yu, K. Zhu, Z. Liu, H. Xiao, W. Deng and X. Zhou, Carbon dioxide reforming of methane over promoted Ni_xMg_{1-x}O



- (111) platelet catalyst derived from solvothermal synthesis, *Appl. Catal., B*, 2014, **148–149**, 177–190.
- 45 Z. Zhang and X. E. Verykios, Carbon dioxide reforming of methane to synthesis gas over Ni/La₂O₃ catalysts, *Appl. Catal., A*, 1996, **138**, 109–133.
- 46 H. Zhou, T. Zhang, Z. Sui, Y.-A. Zhu, C. Han, K. Zhu and X. Zhou, A single source method to generate Ru–Ni–MgO catalysts for methane dry reforming and the kinetic effect of Ru on carbon deposition and gasification, *Appl. Catal., B*, 2018, **233**, 143–159.
- 47 H. Wang, X. Dong, T. Zhao, H. Yu and M. Li, Dry reforming of methane over bimetallic Ni–Co catalyst prepared from La(Co_xNi_{1-x})_{0.5}Fe_{0.5}O₃ perovskite precursor: catalytic activity and coking resistance, *Appl. Catal., B*, 2019, **245**, 302–313.
- 48 L. Xu, Z. Miao, H. Song and L. Chou, CO₂ reforming of CH₄ over rare earth elements functionalized mesoporous Ni–Ln (Ln = Ce, La, Sm, Pr)–Al–O composite oxides, *Int. J. Hydrogen Energy*, 2014, **39**, 3253–3268.
- 49 E.-h. Yang and D. J. Moon, CO₂ Reforming of Methane over Ni⁰/La₂O₃ Catalyst Without Reduction Step: Effect of Calcination Atmosphere, *Top. Catal.*, 2017, **60**, 697–705.
- 50 Q. Mu and Y. Wang, Synthesis, characterization, shape-preserved transformation, and optical properties of La(OH)₃, La₂O₂CO₃, and La₂O₃ nanorods, *J. Alloys Compd.*, 2011, **509**, 396–401.
- 51 Y. Chai, Y. Fu, H. Feng, W. Kong, C. Yuan, B. Pan, J. Zhang and Y. Sun, A Nickel-Based Perovskite Catalyst with a Bimodal Size Distribution of Nickel Particles for Dry Reforming of Methane, *ChemCatChem*, 2018, **10**, 2078–2086.
- 52 X. Li, D. Li, H. Tian, L. Zeng, Z.-J. Zhao and J. Gong, Dry reforming of methane over Ni/La₂O₃ nanorod catalysts with stabilized Ni nanoparticles, *Appl. Catal., B*, 2017, **202**, 683–694.
- 53 J. Shabaker, D. Simonetti, R. Cortright and J. Dumesic, Sn-modified Ni catalysts for aqueous-phase reforming: characterization and deactivation studies, *J. Catal.*, 2005, **231**, 67–76.
- 54 M. Akri, S. Pronier, T. Chafik, O. Achak, P. Granger, P. Simon, M. Trentesaux and C. Batiot-Dupeyrat, Development of nickel supported La and Ce-natural illite clay for autothermal dry reforming of methane: toward a better resistance to deactivation, *Appl. Catal., B*, 2017, **205**, 519–531.
- 55 S. Li, H. Tang, D. Gong, Z. Ma and Y. Liu, Loading Ni/La₂O₃ on SiO₂ for CO methanation from syngas, *Catal. Today*, 2017, **297**, 298–307.
- 56 G. S. Gallego, F. Mondragón, J.-M. Tatibouët, J. Barrault and C. Batiot-Dupeyrat, Carbon dioxide reforming of methane over La₂NiO₄ as catalyst precursor—Characterization of carbon deposition, *Catal. Today*, 2008, **133**, 200–209.
- 57 J.-H. Kim, D. J. Suh, T.-J. Park and K.-L. Kim, Effect of metal particle size on coking during CO₂ reforming of CH₄ over Ni-alumina aerogel catalysts, *Appl. Catal., A*, 2000, **197**, 191–200.
- 58 S. Y. Foo, C. K. Cheng, T.-H. Nguyen and A. A. Adesina, Kinetic study of methane CO₂ reforming on Co–Ni/Al₂O₃ and Ce–Co–Ni/Al₂O₃ catalysts, *Catal. Today*, 2011, **164**, 221–226.
- 59 C. C. Chong, H. D. Setiabudi and A. A. Jalil, Dendritic fibrous SBA-15 supported nickel (Ni/DFSBA-15): a sustainable catalyst for hydrogen production, *Int. J. Hydrogen Energy*, 2020, **45**, 18533–18548.
- 60 X. Du, D. Zhang, R. Gao, L. Huang, L. Shi and J. Zhang, Design of modular catalysts derived from NiMgAl-LDH@m-SiO₂ with dual confinement effects for dry reforming of methane, *Chem. Commun.*, 2013, **49**, 6770–6772.
- 61 N. Wang, K. Shen, L. Huang, X. Yu, W. Qian and W. Chu, Facile Route for Synthesizing Ordered Mesoporous Ni–Ce–Al Oxide Materials and Their Catalytic Performance for Methane Dry Reforming to Hydrogen and Syngas, *ACS Catal.*, 2013, **3**, 1638–1651.
- 62 N. Wang, W. Qian, W. Chu and F. Wei, Crystal-plane effect of nanoscale CeO₂ on the catalytic performance of Ni/CeO₂ catalysts for methane dry reforming, *Catal. Sci. Technol.*, 2016, **6**, 3594–3605.
- 63 S. Podila, Y. A. Alhamed, A. A. AlZahrani and L. A. Petrov, Hydrogen production by ammonia decomposition using Co catalyst supported on Mg mixed oxide systems, *Int. J. Hydrogen Energy*, 2015, **40**, 15411–15422.

