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A novel and facile synthesis strategy for highly stable cesium lead halide nanowires†

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As promising low-dimensional semiconductor materials, cesium lead halide (CsPbX₃, X = Cl, Br, I) perovskite-like nanowires (NWs) can be widely applied to the field of semiconductor devices and integrated optoelectronics. Therefore, developing a facile and efficient synthesis method of cesium lead halide perovskite-like NWs can bring both fundamental and practical impacts to the field of optoelectronics. Here, we developed a synthesis strategy of all-inorganic cesium lead halide CsPbI₃ perovskite-like NWs under catalyst-free, solution-phase, and low-temperature conditions. The synthesis strategy was designed such that no inert gas is required and thus enables the synthesis to be carried out in air, which significantly reduces temperature, steps, time, and cost required for the reaction. The as-synthesized NWs were 7 μm in length and 80–100 nm in diameter with ideal morphology. Most of the CsPbI₃ NWs were crystallized in orthorhombic phases that were arranged orderly with a uniform growth direction. In addition, the CsPbI₃ NWs showed a photoluminescence peak near 610 nm and the fluorescence lifetime was 7.34 ns. The photoluminescence mechanism of CsPbI₃ NWs involves the self-trapping behaviour in the radiative recombination process. The composition of CsPbI₃ NWs is highly related to the synthesis temperature. The facile synthesis strategy has opened up a novel path for the synthesis of perovskite-like NWs, laying the foundation for the application of nano-optoelectronic devices, fluorescent anti-counterfeiting, and fluorescent composite materials.

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1. Introduction

Metal halide perovskite-like materials with ABX₃ structure (A = Rb, Cs; B = Ge, Sn, Pb; X = Cl, Br, I) have been widely used in solar cells,^{1–4} light-emitting diodes (LEDs),^{5–8} photodetectors,^{9–13} lasers^{14–17} and other fields because of their excellent optical and electrical properties, low cost preparation, and simple solution processing, and thus have attracted widespread attention in recent years.¹⁸ Compared to traditional semiconductor nanocrystals or nano-structured metal halide perovskite-like materials, all-inorganic lead halide perovskite-like materials (APbX₃, X = Cl, Br, I) have better optical properties, including photoluminescence effect and a narrow emission peak width. In addition, the emission spectra can be adjusted to the entire

visible region simply by modulating chemical components or changing the luminescence band size. Cesium lead iodide perovskite-like (CsPbI₃), specifically, has a suitable band gap with fluorescent properties that achieve high luminescence efficiency and performance.¹⁹ The excellent physical properties of CsPbI₃ have also prompted researchers to explore its use in high-performance optoelectronic applications.

Various methods to synthesize CsPbX₃ (X = Cl, Br, I) nanocrystals have been reported. Protesescu *et al.* first utilized a hot injection method to synthesize CsPbX₃.²⁰ Different nanostructures of CsPbX₃ nanocrystals, including nanocubes (NCs), nanosheets (NSs), nanoplates (NPLs), quantum dots (QDs),^{21,22} and nanowires (NWs),^{23–25} have been synthesized by varying ligands²⁶ and ratio, reaction time, or reaction temperature. The common synthesis method of CsPbX₃ nanocrystals involves using a standard anaerobic Schlenk line device at high temperatures (140–200 °C), with the use of inert gas and vacuum drying environment.²⁷ However, this common method does not apply well to all nanostructures of CsPbX₃ nanocrystals. For example, although the NWs synthesized by the common method have excellent luminescence properties, the synthesis process is complex and costly. Moreover, the synthesis of CsPbI₃ NWs requires higher temperature (most requires higher than 180 °C) than that required for CsPbBr₃ NWs and CsPbCl₃ NWs.²⁸ Therefore, it is important to develop a facile and low-temperature synthesis method for CsPbI₃ NWs.

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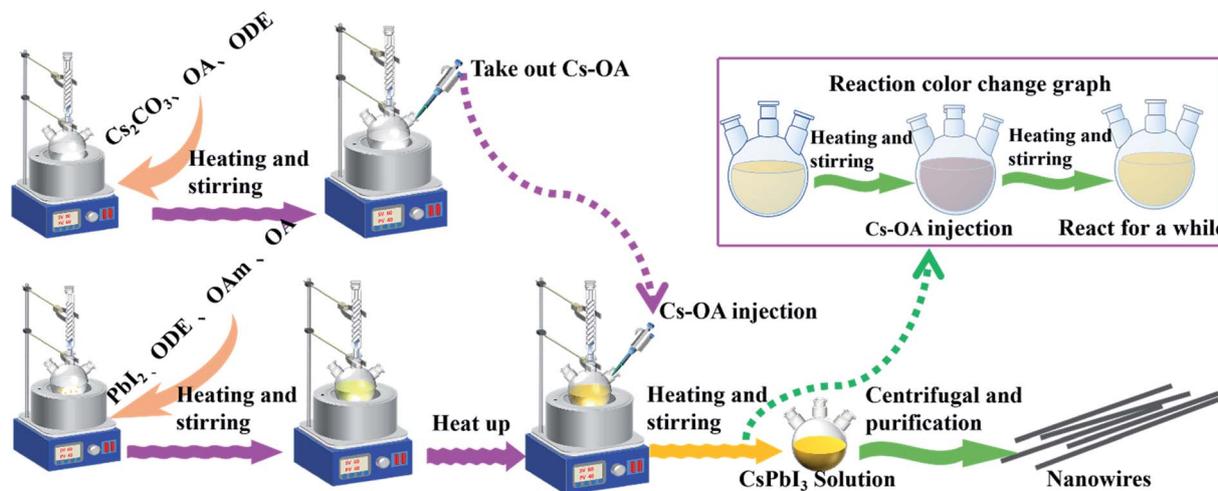


Fig. 1 Scheme of CsPbI₃ NWs preparation.

Here, a novel and facile synthesis strategy for highly stable cesium lead halide nanowires (CsPbX₃, X = Cl, Br, I) is reported. Compared with the conventional synthesis technique, our synthesis strategy uses a lower reaction temperature and does not require inert gas, vacuuming, and degassing processes. Our synthesis strategy with simplified reaction steps not only reduces the cost and time, but also improves the efficiency of the CsPbX₃ synthesis. The as-synthesized CsPbI₃ NWs using our strategy are formed as single crystals that crystallize in an orthorhombic phase and uniform growth directions. Also, the CsPbI₃ NWs were demonstrated to have superior optical properties. In sum, this novel and facile synthesis method is ideal to study the structure–function relationship, which is highly important for the applications and development of nano-optoelectronics and semiconductor devices.²⁹

2. Results and discussion

The preparation of CsPbI₃ NWs was performed under air-free conditions. The preparation process did not require the participation of an inert gas, and the vacuuming and drying processes were omitted in the experiment. The scheme of the CsPbI₃ NW preparation is shown in Fig. 1. Basically, the reaction was carried out by reacting Cs-oleic acid (Cs-OA) with lead halide (PbI₂) in the presence of OA and oleylamine (OAm) in octadecene (ODE). Different reaction temperatures were first tested and optimized at 155 °C based on the diameter of the CsPbI₃ NWs. Then, different reaction times at this temperature were tested and the reacted solution was cooled to room temperature on an ice-water bath. The as-synthesized nanocrystalline was obtained *via* centrifugation and purification and its morphology was examined *via* X-ray diffraction (XRD) and scanning electron microscopy (SEM).

The dynamics of CsPbI₃ NW formation was studied *via* SEM. In the initial reaction stage ($t < 10$ min), the major morphology of the nanocrystalline was nanocube with the size ranging from 3 nm to 9 nm. After 10 min, a few CsPbI₃ NWs with diameters around

100 nm and the length of about 3 μm were observed (Fig. 2(a)). After a longer reaction time (15–30 min), an increasing number of CsPbI₃ NWs were formed with a longer length of about 7 μm (Fig. 2(b) and (c)), and the formed CsPbI₃ NWs arranged more orderly (Fig. 2(d)). However, when the reaction time reached 30 min, the length of the CsPbI₃ NWs began to shorten gradually while maintaining the ordered arrangement (Fig. 2(e) and (f)). Accordingly, the optimized reaction temperature of 155–160 °C and the optimized reaction time of 15–30 min will be used for the synthesis of all-inorganic CsPbI₃ NWs, which leads to ideal morphology, ordered arrangement, and high crystallinity. Compared with CsPbI₃ NWs with diameters ranging from tens to hundreds of nanometers,²⁸ the CsPbI₃ NWs synthesized here were more stable with a controllable diameter to about 100 nm. The HR-TEM images (ESI, Fig. S5[†]) show that the CsPbI₃ NWs are single-crystalline.

The SEM image (ESI, Fig. S4[†]) and Fig. 3 show the evolution of CsPbI₃ NWs. In the initial stage of the reaction, nanowires mainly exist in the form of nanocubes. Simultaneously, some nanocubes began to evolve. Fig. 4 shows that intermediates precursors similar to nanowires are derived from the interior of the nanocube. In addition, from Fig. 2, it is suggested that the derivatives of nanocubes are embryonic nanowires. As the reaction progresses, the NWs fall off from the inside of the nanocubes, and as the reaction time increases, the morphology of NWs is complete. No discrete reaction intermediates of CsPbI₃ NWs were formed at a specific reaction time, as observed from the above experiments. Different intermediates coexisted with the product at a given reaction time. The growth and shape control of perovskite-like nanoparticles is a kinetically-controlled process, related to the binding of ligands and the temperature-dependent dynamic of surface ligands. The formation of perovskite-like nanoparticles is mainly divided into two steps. The first step involves a seed-mediated nucleation, and the second step involves further growth through self-assembly³⁰ and the directed attachment. The lead component helps stabilize the nucleation of the perovskite-like structure. In addition, the aging of the precursor solution could promote the generation of CsPbI₃ NWs.³¹



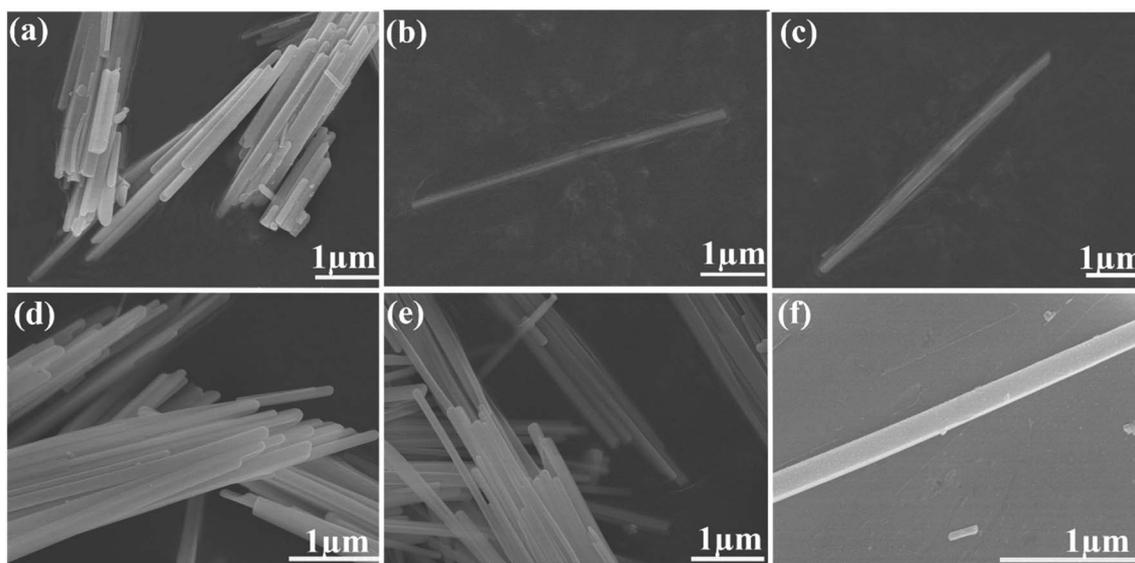


Fig. 2 Synthesis of CsPbI₃ NWs. (a) The reaction time is 10 min. (b–d) The reaction time is set as 15, 20 and 25 min, respectively. (e and f) The reaction time is 30 min. The images represent the morphological characterization of CsPbI₃ NWs. Scale bar, 1 μm.

Another characteristic of the synthesis of CsPbI₃ is the colour change during the synthesis. As the Cs-OA solution was added to the PbI₂ precursor solution, the colour of the solution instantly turned reddish black, and then gradually changed to yellow with the increase in the reaction time and stayed yellow (Fig. 4(a)). The synthesized CsPbI₃ NWs with simplified experimental steps show higher intensity of diffraction peaks in the XRD pattern, where the peaks correspond to the peaks in the spectrum of standard orthorhombic CsPbI₃ (PDF # 741970) materials. In addition, the intensity of the diffraction peaks of CsPbI₃ NWs gradually increased with time, which can be attributed to increase in time and coordination of anions and cations in the reaction solution,³² thereby promoting the growth and crystallinity of perovskite-like crystals. The results suggested that even if the experimental steps are simplified, the crystallinity of CsPbI₃ NWs is still excellent and belongs to the orthorhombic phase. In addition, with the colour changed from

reddish black to yellow in the reaction solution, CsPbI₃ NWs were shown to undergo a phase transition^{33,34} from cubic phase to orthorhombic phase during the formation process. The characterization further supports that the synthesized CsPbI₃ NWs have good morphology and completed crystal morphology.

Perovskite-like nanocrystals emit luminescence through the radiative recombination process of which the capture of a free electron by an ion occurs and an excess energy is radiated as photons.^{35,36} Defect states of traditional nanocrystals have energy levels in between the conduction band and the valence band and thus reduces the luminescence efficiency. In contrast,

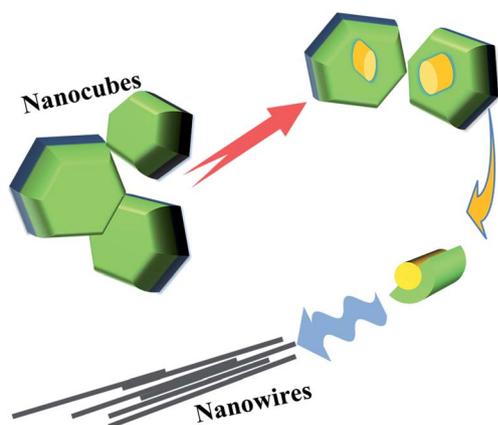


Fig. 3 The evolution of CsPbI₃ NWs.

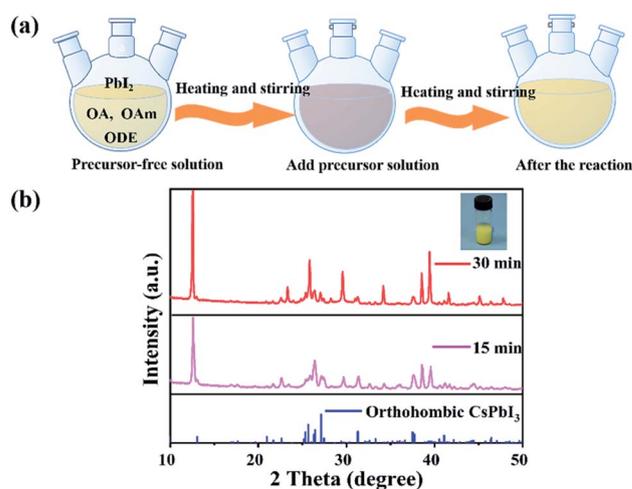


Fig. 4 (a) Colour change of solution after adding precursor. (b) Standard XRD patterns (bottom) of orthorhombic CsPbI₃, and experimental XRD patterns (top two) of CsPbI₃ NWs. The pink colour is the reaction time of 15 min, the red colour is the reaction time of 30 min, respectively. Inset: CsPbI₃ NWs distributed in hexane solution, showing yellow colour.



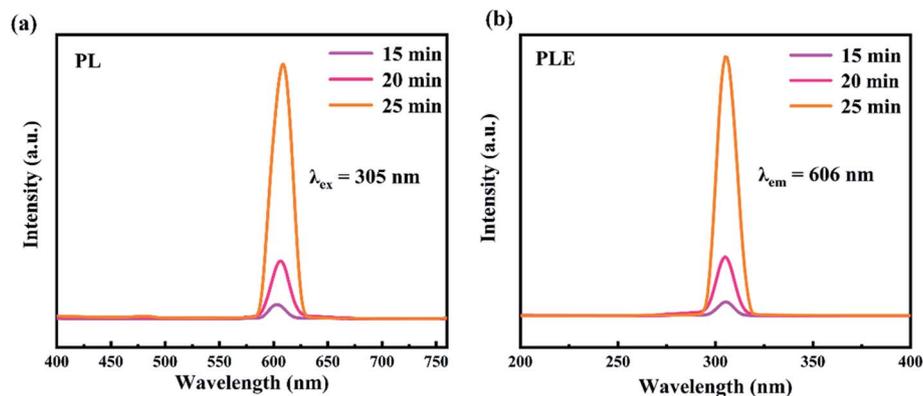


Fig. 5 Photoluminescence spectra of CsPbI₃ NWs (a) PL ($\lambda_{\text{ex}} = 305$ nm) and (b) PLE ($\lambda_{\text{em}} = 606$ nm) spectra. Both are researches on CsPbI₃ NWs products that have reacted for 15 min, 20 min and 25 min.

the energy levels of defect states of perovskite-like CsPbX₃ nanoparticles are only formed inside the valence band and conduction band, avoiding the generation of intermediate state defects, thereby improving the luminescence efficiency.^{37,38} Therefore, compared with traditional luminescent materials, perovskite-like materials possess the advantages of narrow and symmetrical emission spectrum, controllable emission wavelength, high luminescence efficiency, and stable luminescence performance, making it an ideal luminescent material.

We further examined the optical properties of the experimentally synthesized CsPbI₃ NWs. The photoluminescence spectra and fluorescence lifetime spectrum of the CsPbI₃ NWs were measured at the optimized temperature during synthesis. Fig. 5(a) shows how the photoluminescence emission (PL) spectra of CsPbI₃ NWs changed at 15 min, 20 min, and 25 min of reaction times. The formation of the NWs morphology was complete, as suggested by the single peak in the spectra. PL spectra ($\lambda_{\text{ex}} = 305$ nm) of CsPbI₃ NWs show broad bands from 400 to 750 nm. The fluorescence intensity increased with the increase in the reaction time. As time rises from 15 to 25 min,

the peak in the PL spectra red-shifted from 603 nm to 609 nm, suggesting that the nanoparticles became larger as the time increased. The red-shift phenomenon can be attributed to the effects of the local state energy level on the emission of the CsPbI₃ NWs, thereby affecting the quantum confinement.^{39,40} In addition, the emission peak is symmetrically distributed and narrow, which indicates that the size distribution of the CsPbI₃ NWs is uniform and thus the NWs are mono-dispersed. The photoluminescence excitation (PLE) spectra ($\lambda_{\text{em}} = 606$ nm) in Fig. 5(b) shows the excitation bands peaked around 305 nm. The peak of 25 min was stronger than the peak of 15 min, and the peak intensity also increased over time, further indicating that the size of nanoparticles increased. Furthermore, at the same temperature, the fluorescence intensity of reaction for 15 min, 30 min, and 45 min is also measured. From the PL spectra (Fig S1a†) and PLE spectra (Fig S1b†), it can be seen that the fluorescence peak of the NWs reacted for 45 min blue-shifted relative to the fluorescence peaks of the 20 and 30 min. It is most likely due to the reduction in the size of the nanocrystals. The results further prove that the CsPbI₃

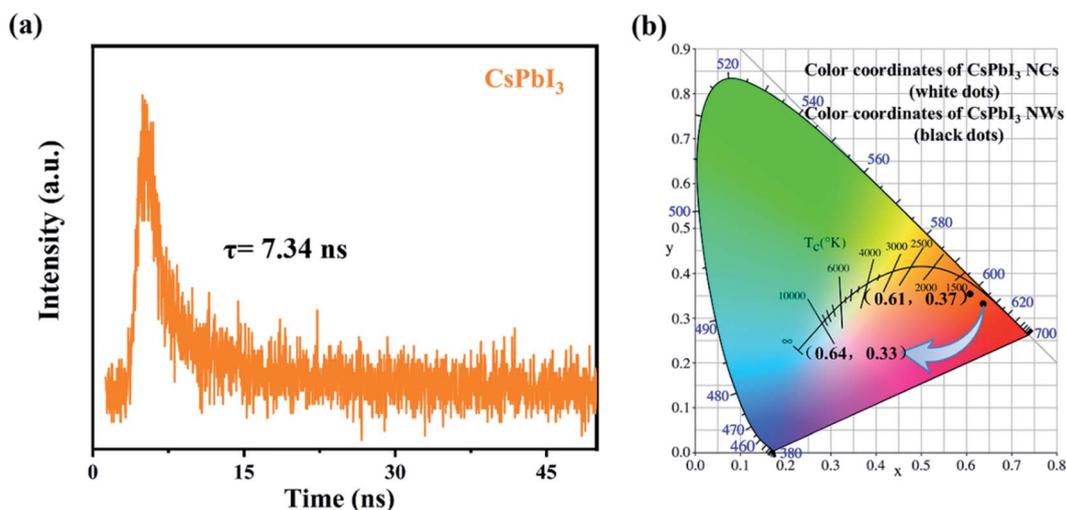


Fig. 6 (a) Fluorescence decay curve of CsPbI₃ NWs reacted at 155 °C for 15 min. (b) Colour coordinates of CsPbI₃ NCs and CsPbI₃ NWs.



nanowires synthesized within 15–30 min are more desirable. Fig. 6(a) shows the fluorescence lifetime spectrum of CsPbI₃ NWs synthesized at 155 °C for 15 min; the fluorescence lifetime was 7.34 ns. As time increased, the fluorescence intensity first increased and decreased afterwards. The result implies that the luminescence emission not only resulted from the radiative recombination of excitons, but is also affected by other factors such as luminescence from defects as well as electron and exciton recombination, and particle size and structure. The fluorescence decay curve can be fitted by a biexponential equation and the fluorescence lifetime can be calculated as follows:^{20,41,42}

$$I(t) = A_1 \exp(-t/\tau_1) + A_2 \exp(-t/\tau_2)$$

where τ is the fluorescence lifetime, of which the short component can be attributed to hole traps, while the long component is related to electron traps, t is the decay time, I is the relative fluorescence intensity, and A_1 and A_2 are constants. The fitting parameters of fluorescence decay are summarised in Table 1, and the lifetime is calculated to be 7.34 ns. The optical measurements show that the CsPbI₃ NWs synthesized by the simplified steps have a high radiation/non-radiation recombination ratio and maintain excellent fluorescence performance. In addition, Fig S2† shows that the fluorescence intensity was relatively low at excitation wavelengths around 200 nm. When excited at a single wavelength, as the excitation wavelength increases, the peak intensity of the emission spectrum gradually shows a red-shift and can cover around 480–760 nm. In addition, when the excitation wavelength is increased to longer than 350 nm, strong luminescence can be observed in the wavelength range of 400–480 nm. In consequence, the emission beam would contain the whole visible light range (400–760 nm).

Fig. 6(b) shows the chromaticity coordinates of CsPbI₃ nanocubes (NCs) and CsPbI₃ NWs under excitation at a wavelength of 305 nm. The colour coordinates of CsPbI₃ nanocubes (NCs) and NWs are (0.61, 0.37) and (0.64, 0.33), respectively. The coordinate values gradually red-shifted with the increase in time, which correspond to the PL spectrum at the same time. Moreover, we did a stability test on the CsPbI₃ NWs. It was found that the experimentally synthesized CsPbI₃ NWs may exhibit fluorescence decay phenomenon under UV irradiation (Fig. S3a†). Next, we performed modified experiments on the synthesized CsPbI₃ NWs. The same fluorescence test showed a negligible fluorescence intensity change of the NWs (Fig. S3b†), suggesting that the stability can still be improved through modified experiments. The results further proved that the experimental synthesis steps are simple and effective. The study will be continued to further improve the stability of CsPbI₃ NWs.

Table 1 Parameters used for fitting the PL decays of CsPbI₃ NWs, the CsPbI₃ NWs is reacted for 15 min at 155 °C

Sample	A ₁ (%)	τ ₁ (ns)	A ₂ (%)	τ ₂ (ns)	t _{avg} (ns)
CsPbI ₃ NWs	52.71	1.72	11.64	11.24	7.34

3. Conclusion

In summary, we have developed a novel and facile synthesis strategy for all-inorganic CsPbI₃ NWs, of which no inert gas protection during the entire process is required. Also, the synthesis strategy was optimized with reduced reaction steps, time, cost and temperature without the need of catalysts. From the material characterization, the as-synthesized CsPbI₃ NWs were crystallized as a single crystal in an orthogonal phase that was arranged orderly with a uniform growth direction. The nanocrystals underwent phase transition from a cubic phase to an orthorhombic phase in a short time during the reaction process and finally remained in the orthorhombic phase. In addition, the CsPbI₃ NWs showed a photoluminescence peak at 610 nm and the fluorescence lifetime was 7.34 ns. To summarize, the synthesis strategy reported here simplifies the traditional synthesis method and opens up a new synthesis route for applications in the fields of optoelectronics, semiconductor devices and fluorescent anticorrosion.

4. Experimental section

4.1 Materials

Cesium carbonate (Cs₂CO₃, 99.99%), 1-octadecene (ODE, >90%, GC), lead iodide (PbI₂, 99.9%), oleylamine (OAm, 80–90%), and oleic acid (OA, AR), *n*-hexane (>99%, GC). All chemicals were purchased from Aladdin and used as received without further purification.

4.1.1 Preparation of cesium oleate precursors. Cs₂CO₃ (0.4 g), OA (1.25 mL) and ODE (15 mL) were added to a 50 mL three-neck flask and stirred at 120 °C for 1 h to completely dissolve Cs₂CO₃ and form cesium oleate (Cs-OA). The Cs-OA solution was stored and preheated to 150 °C before the use.

4.2 Synthesis of CsPbI₃ NWs

PbI₂ (1.392 g), ODE (80 mL), OAm (8.8 mL), and OA (16 mL) were added into a 250 mL three-neck flask and stirred at 120 °C for 1 h to completely dissolve PbI₂. The temperature was then raised to 155–160 °C, and then 2.4 mL of the pre-prepared and preheated Cs-OA solution was quickly injected into the PbI₂ solution. After stirring for 15–25 min, the reaction mixture with the synthesized CsPbI₃ NWs was cooled in an ice-water bath for the purification process.

4.3 Isolation and purification of CsPbI₃ NWs

A centrifugation process was performed at 6000 rpm for 5 min to separate the CsPbI₃ NWs from the crude solution. After discarding the supernatant, hexane was added to dissolve the precipitate. The centrifugation process was repeated several times to obtain the purified final product of CsPbI₃ NWs.

4.4 Characterization

Photoluminescence and photoluminescence excitation spectra were recorded at room temperature by Edinburgh FLS1000 with a xenon lamp as the excitation source. Moreover, an Edinburgh steady/transient fluorescence spectrometer FLS1000 was used



for testing the decay of PL. Scanning electron microscopy (SEM) images were recorded using a SU8010 (Hitachi Company of Japan) operated at 10 kV. Transmission Electron Microscopy (TEM) images were obtained on a JEOL JEM-F200 electron microscope. Powder X-ray diffraction (XRD) patterns of the obtained products were measured on a PANalytical Empyrean power X-ray diffractometer using monochromatic Cu K α radiation. The sample powder to be measured was spread on the glass substrate.

Author contributions

Kun Nie, Ranran Zhou, and Hua Wang conceived the concept and idea of the present paper, worked on the study design strategy and selected the topics to be discussed. Ranran Zhou completed the main experiment and paper writing, Kun Nie, and Hua Wang were responsible for guidance and supervision work, and at the same time, guided and revised the final paper writing. Chi-An Cheng provided writing assistance and article revisions. Siying Qiu, and Jiayi Chen conducted literature searches, and abstracted the data from the eligible full text articles. Mengyun Wu, and Panlong Lin provided experimental guidance and experimental tests. Luoxin Wang, and Lefu Mei supervised the study. All of the authors discussed the results and commented on the manuscript.

Conflicts of interest

The authors declare no competing financial interest.

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Notes and references

- H. Wang, H. Li, W. Cai, P. Zhang, S. Cao, Z. Chen and Z. Zang, *Nanoscale*, 2020, **12**, 14369–14404.
- J. Burschka, N. Pellet, S. J. Moon, R. Humphry-Baker, P. Gao, M. K. Nazeeruddin and M. Gratzel, *Nature*, 2013, **499**, 316–319.
- M. Liu, M. B. Johnston and H. J. Snaith, *Nature*, 2013, **501**, 395–398.
- H. Tsai, W. Nie, J. C. Blancon, C. C. Stoumpos, R. Asadpour, B. Harutyunyan, A. J. Neukirch, R. Verduzco, J. J. Crochet, S. Tretiak, L. Pedesseau, J. Even, M. A. Alam, G. Gupta, J. Lou, P. M. Ajayan, M. J. Bedzyk and M. G. Kanatzidis, *Nature*, 2016, **536**, 312–316.
- H. Guan, S. Zhao, H. Wang, D. Yan, M. Wang and Z. Zang, *Nano Energy*, 2020, **67**, 104279.
- Z. K. Tan, R. S. Moghaddam, M. L. Lai, P. Docampo, R. Higler, F. Deschler, M. Price, A. Sadhanala, L. M. Pazos, D. Credgington, F. Hanusch, T. Bein, H. J. Snaith and R. H. Friend, *Nat. Nanotechnol.*, 2014, **9**, 687–692.
- M. Yuan, L. N. Quan, R. Comin, G. Walters, R. Sabatini, O. Voznyy, S. Hoogland, Y. Zhao, E. M. Bearegard, P. Kanjanaboos, Z. Lu, D. H. Kim and E. H. Sargent, *Nat. Nanotechnol.*, 2016, **11**, 872–877.
- X. Huang, Q. Sun and B. Devakumar, *J. Lumin.*, 2020, **225**, 117373.
- H. Wang, P. Zhang and Z. Zang, *Appl. Phys. Lett.*, 2020, **116**, 162103.
- L. Dou, Y. M. Yang, J. You, Z. Hong, W. H. Chang, G. Li and Y. Yang, *Nat. Commun.*, 2014, **5**, 5404.
- Z. L. Wang and J. Song, *Science*, 2006, **312**, 242–246.
- H. Li, W. Lin, L. Ma, Y. Liu, Y. Wang, A. Li, X. Jin and L. Xiong, *RSC Adv.*, 2021, **11**, 11663–11671.
- W. Zhai, J. Lin, C. Li, S. Hu, Y. Huang, C. Yu, Z. Wen, Z. Liu, Y. Fang and C. Tang, *Nanoscale*, 2018, **10**, 21451–21458.
- Z. Hu, Z. Liu, Z. Zhan, T. Shi, J. Du, X. Tang and Y. Leng, *Adv. Photonics*, 2021, **3**, 034002.
- G. Xing, N. Mathews, S. S. Lim, N. Yantara, X. Liu, D. Sabba, M. Gratzel, S. Mhaisalkar and T. C. Sum, *Nat. Mater.*, 2014, **13**, 476–480.
- H. Zhu, Y. Fu, F. Meng, X. Wu, Z. Gong, Q. Ding, M. V. Gustafsson, M. T. Trinh, S. Jin and X. Y. Zhu, *Nat. Mater.*, 2015, **14**, 636–642.
- X. Tang, Z. Hu, W. Yuan, W. Hu, H. Shao, D. Han, J. Zheng, J. Hao, Z. Zang, J. Du, Y. Leng, L. Fang and M. Zhou, *Adv. Opt. Mater.*, 2017, **5**, 1600788.
- X. Bao, M. Li, J. Zhao and Z. Xia, *J. Mater. Chem. C*, 2020, **8**, 12302–12307.
- Y. Tang, A. Lesage and P. Schall, *J. Mater. Chem. C*, 2020, **8**, 17139–17156.
- L. Protesescu, S. Yakunin, M. I. Bodnarchuk, F. Krieg, R. Caputo, C. H. Hendon, R. X. Yang, A. Walsh and M. V. Kovalenko, *Nano Lett.*, 2015, **15**, 3692–3696.
- S. Zhao, Y. Zhang and Z. Zang, *Chem. Commun.*, 2020, **56**, 5811–5814.
- D. Yan, T. Shi, Z. Zang, T. Zhou, Z. Liu, Z. Zhang, J. Du, Y. Leng and X. Tang, *Small*, 2019, 1901173.
- S. Sun, D. Yuan, Y. Xu, A. Wang and Z. Deng, *ACS Nano*, 2016, **10**, 3648–3657.
- A. Swarnkar, R. Chulliyil, V. K. Ravi, M. Irfanullah, A. Chowdhury and A. Nag, *Angew. Chem., Int. Ed.*, 2015, **54**, 15424–15428.
- S. W. Eaton, M. Lai, N. A. Gibson, A. B. Wong, L. Dou, J. Ma, L. W. Wang, S. R. Leone and P. Yang, *Proc. Natl. Acad. Sci. U. S. A.*, 2016, **113**, 1993–1998.
- Y. Luo, T. Tan, S. Wang, R. Pang, L. Jiang, D. Li, J. Feng, H. Zhang, S. Zhang and C. Li, *Nanoscale*, 2021, **13**, 4899–4910.
- Y. Li, H. Zhang, P. Liu, Y. Wang, H. Yang, Y. Li and H. Zhao, *Electrochem. Commun.*, 2015, **51**, 6–10.
- D. Zhang, S. W. Eaton, Y. Yu, L. Dou and P. Yang, *J. Am. Chem. Soc.*, 2015, **137**, 9230–9233.
- K. Nie, Q. An and Y. Zhang, *Nanoscale*, 2016, **8**, 8791–8797.
- Y. Bekenstein, B. A. Koscher, S. W. Eaton, P. Yang and A. P. Alivisatos, *J. Am. Chem. Soc.*, 2015, **137**, 16008–16011.
- T. Udayabhaskararao, M. Kazes, L. Houben, H. Lin and D. Oron, *Chem. Mater.*, 2017, **29**, 1302–1308.



- 32 B. A. Koscher, J. K. Swabeck, N. D. Bronstein and A. P. Alivisatos, *J. Am. Chem. Soc.*, 2017, **139**, 6566–6569.
- 33 J. A. Steele, H. Jin, I. Dovgaliuk, R. F. Berger, T. Braeckvelt, H. Yuan, C. Martin, E. Solano, K. Lejaeghere, S. M. J. Rogge, C. Notebaert, W. Vandezande, K. P. F. Janssen, B. Goderis, E. Debroye, Y.-K. Wang, Y. Dong, D. Ma, M. Saidaminov, H. Tan, Z. Lu, V. Dyadkin, D. Chernyshov, V. Van Speybroeck, E. H. Sargent, J. Hofkens and M. B. J. Roeffaers, *Science*, 2019, **365**, 679–684.
- 34 D. M. Trots and S. V. Myagkota, *J. Phys. Chem. Solids*, 2008, **69**, 2520–2526.
- 35 X. Tang, X. Hou, D. Shu and P. Zhai, *Sci. China: Technol. Sci.*, 2015, **58**, 1712–1716.
- 36 Y. L. Lai, C.-P. Liu and Z.-Q. Chen, *Appl. Phys. Lett.*, 2005, **86**, 121915.
- 37 H. Huang, M. I. Bodnarchuk, S. V. Kershaw, M. V. Kovalenko and A. L. Rogach, *ACS Energy Lett.*, 2017, **2**, 2071–2083.
- 38 J. M. Ball and A. Petrozza, *Nat. Energy*, 2016, **1**, 16149.
- 39 X. Shen, S. Wang, C. Geng, L. Li, E. Zhao, J. Sun, W. Wu, L. An and K. Pan, *J. Phys. Chem. C*, 2021, **125**, 5278–5287.
- 40 D. Chen, Y. Liu, C. Yang, J. Zhong, S. Zhou, J. Chen and H. Huang, *Nanoscale*, 2019, **11**, 17216–17221.
- 41 P. Li, W. Xie, W. Mao, Y. Tian, F. Huang, S. Xu and J. Zhang, *J. Alloys Compd.*, 2020, **817**, 153338.
- 42 X. Bao, M. Li, J. Zhao and Z. Xia, *J. Mater. Chem. C*, 2020, **8**, 12302–12307.

