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# Highly porous copper-supported magnetic nanocatalysts: made of volcanic pumice textured by cellulose and applied for the reduction of nitrobenzene derivatives†

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Herein, a novel designed heterogeneous catalytic system constructed of volcanic pumice magnetic particles (VPMPs), cellulose (CLS) as a natural polymeric matrix, and copper nanoparticles (Cu NPs) is presented. Also, to enhance the inherent magnetic property of VPMP, iron oxide (Fe<sub>3</sub>O<sub>4</sub>) nanoparticles have been prepared and incorporated in the structure *via* an *in situ* process. As its first and foremost excellent property, the designed composite is in great accordance with green chemistry principles because it contains natural ingredients. Another brilliant point in the architecture of the designed composite is the noticeable porosity of VPMP as the core of the composite structure (surface area: 84.473 m<sup>2</sup> g<sup>-1</sup>). This great porosity leads to the use of a small amount (0.05 g) of the particles for catalytic purposes. However, the main characterization methods, such as Fourier-transform infrared and energy-dispersive X-ray spectroscopy, thermogravimetric analysis, and electron microscopy, revealed that the spherical metallic particles (Fe and Cu oxides) were successfully distributed onto the surface of the VPMP and CLS matrices. Further, vibrating-sample magnetometer analysis confirmed the enhancement of the magnetic property (1.5 emu g<sup>-1</sup>) of the composite through the addition of Fe<sub>3</sub>O<sub>4</sub> nanoparticles. Further, the prepared (Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu) nanocomposite has been applied to facilitate the reduction reaction of hazardous nitrobenzene derivatives (NBDs) to their aniline analogs, with 98% conversion efficiency in eight minutes under mild conditions. Moreover, the good reusability of the catalytic system has been verified after recycling it ten times without any significant decrease in the performance.

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## 1. Introduction

As time goes on, microscale and nanoscale heterogeneous catalytic systems are gaining increasing attention for a variety of reasons, of which one of the most important is that they are able to create a high surface area for chemical reactions between the involved reactants and the catalytic substrate.<sup>1–3</sup> In this regard, heterogeneous catalytic systems based on iron oxide nanoparticles have become more important due to the possibility of

their easy separation by only holding an external magnet at the bottom of the reaction flask and decanting the content.<sup>4–11</sup> Heterogeneous catalytic composites are modified by various components, such as organic structures, polymers, inorganic particles, biological structures, copper nanoparticles, and palladium nanoparticles, so that the majority of the catalytic sites are available.<sup>12–20</sup> In recent years, more attention has been paid to green chemistry due to environmental concerns.<sup>20,21</sup> That is why researchers are constantly trying to design novel natural heterogeneous catalytic systems with a high degree of biocompatibility. Undoubtedly, volcanic pumice benefits from a very porous structure that can lead to a high surface area, and it possesses properties such as high magnetic property, biocompatibility, and great surface functionalization capability; thus, it can inherently act as an eco-friendly substrate to create a heterogeneous catalytic composite.<sup>21–25</sup>

The use of cellulose (CLS) in this work, because it is natural, has been a strong indication of how important green chemistry was to us in this work.<sup>26,27</sup> There are many reasons why cellulose, as a substantial polymeric substrate with a biocompatible

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origin, is used in this work; one of the most important of these is that cellulose, due to its large number of hydroxyl (–OH) functional groups, can form physicochemical hydrogen bonds with the –OH groups of pumice, which can lead to integration in the composite. On the other hand, the rest of the hydroxyl functional groups of cellulose, which are free of hydrogen bonds with the hydroxyl functional groups of pumice, are suitable for chelation by cationic metals such as copper and palladium and can form heterogeneous catalytic composites from natural components; this is one of the most important properties of this composite.<sup>28</sup> The synthesized composite is notable even in terms of mechanical properties because it is a stable hybrid structure.<sup>29,30</sup> As mentioned earlier, the synthetic heterogeneous catalytic composite in this work is suitable for scaling up and industrial applications because of its high biocompatibility and because it is made from components found in abundance in nature. In this work, after pumice, as an inorganic base with –OH groups functions as a suitable host for cellulose, as an organic base,  $\text{Cu}^{2+}$  ions are added to the system to act as catalytic sites to reduce the nitro functional groups of different compounds to amines after their conversion to Cu.<sup>31,32</sup>

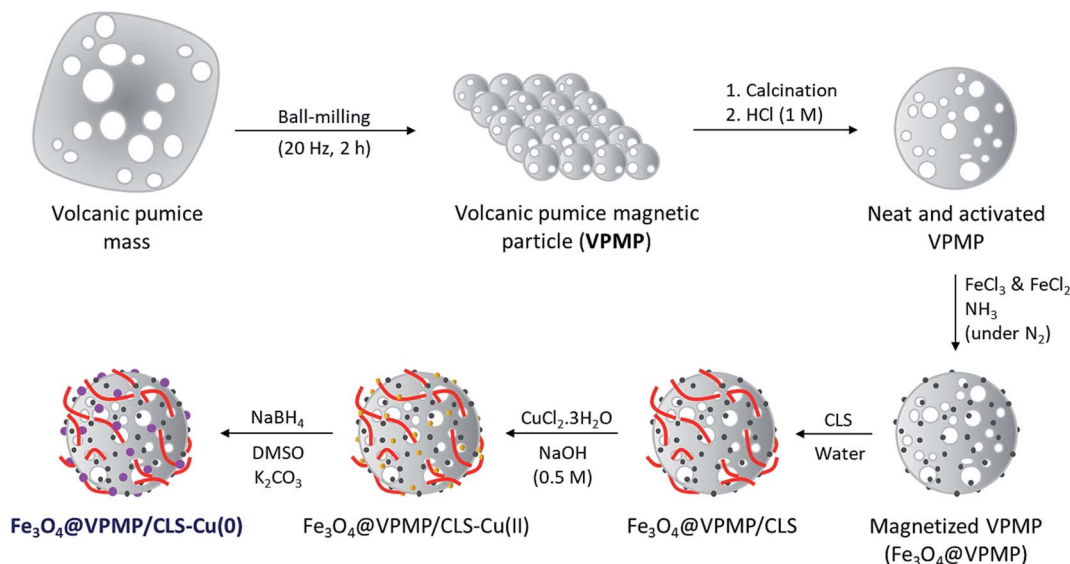
Herein, an attempt was made to introduce a novel and convenient method for the synthesis of amine derivatives from the corresponding nitro compounds using  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalyst, which was synthesized with natural VPMP, CLS, and also Cu NPs. Then, the magnetic behavior and other critical characteristic properties of the prepared catalytic system, such as its average size, porosity, present chemical state, metallic elements, and thermal stability, were carefully investigated using different analytical methods. Then, the ability of this catalyst to reduce nitrobenzene derivatives (NBDs) was carefully examined. In summary, it was found that using the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system at 70 °C led to a 98% reaction yield in just 8 minutes. Also, the catalytic system under study can be reused for 10 successive runs due to its magnetic property, which leads to easy magnetic separation of this catalyst

from the reaction medium. Also, in this study, water solvent was used as an eco-friendly solvent to evaluate the performance of the synthesized catalytic composite; this easily confirms the importance of green chemistry in this study.

## 2. Results and discussion

### 2.1. Preparation of the $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ nanocatalyst

As shown in Scheme 1, several steps are taken to prepare the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  composite. Initially, the pumice is ground by a ball-mill to form uniform particles.<sup>33</sup> After that, because the pumice is very porous and there may be useless fillers in its pores, it should be calcinated at a high temperature in a furnace.<sup>34</sup> Then, if electron microscopy (EM) confirms that the calcinated pumice has uniformity and the useless fillers have been successfully removed from its pores, the magnetic property of the pumice is increased by  $\text{Fe}_3\text{O}_4$  in order to separate it from the solvent on a laboratory scale, although the pumice itself has an inherent magnetic property. Next, it is dispersed and mixed in a concentrated solution of CLS under mild conditions. At the end of this step, the formed composite is magnetically separated and washed to remove the unbound CLS. Herein, a composite called  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS}$  was prepared that contains a large number of hydroxyl functional groups in its VPMP and CLS structures and can be used several times. In another flask, an aqueous solution of  $\text{CuCl}_2 \cdot \text{H}_2\text{O}$  was prepared, which was added to a solution of dispersed  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS}$  in which the hydroxyl functional groups were activated by an alkaline aqueous solution. After completion of the reaction, the resulting composite,  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$ , which was now brown and no longer black like magnetic pumice, was magnetically separated and washed to remove excess salts and ions trapped inside the pores and silicate network of the pumice. Ultimately,  $\text{Cu(II)}$  in the composite of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$  was reduced to  $\text{Cu(0)}$  by sodium borohydride under alkaline conditions.<sup>35</sup> Herein, we intend to



Scheme 1 Preparation route of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(0)}$  nanocomposite.

**Table 1** Optimization of the reduction reaction of nitrobenzene under various catalytic conditions

Entry	Cat.	Cat. (mg)	Cat. (mmol%)	NaBH <sub>4</sub> (mmol)	Temp. (°C)	Time (min)	Yield (%)
1	—	—	—	—	80	180	N.R.
2	—	—	—	2	80	180	Trace
3	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS	30	12	2	80	90	Trace
4	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	10	4	2	70	8	85
5	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	30	12	2	70	8	90
6	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	50	20	2	70	8	98 <sup>a</sup>
7	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	60	24	2	70	8	98
8	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	50	20	3	70	8	94
9	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	50	20	1	70	8	90
10	CLS-Cu	50	20	2	70	8	92
11	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	50	20	2	R.t.	8	81
12	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu	50	20	2	R.t.	30	97

<sup>a</sup> Optimum condition using the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocomposite (50.0 mg), nitrobenzene (1.0 mmol), NaBH<sub>4</sub> (2.0 mol), and at 70 °C. Calculations related to mmol% of catalyst has been given in the ESI file. R.t. stands for room temperature.

monitor the catalytic ability of the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocatalyst in the reduction reactions of NBDs, and the results obtained from the mentioned experiments are reported in Table 1, in the optimization section.

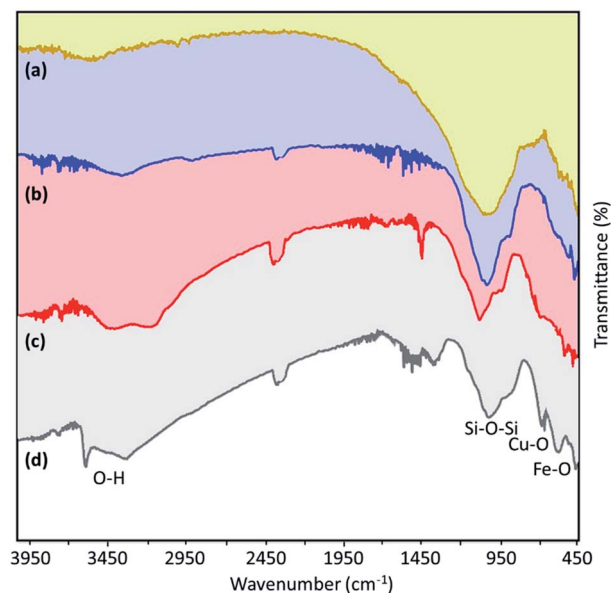
## 2.2. Characterization of Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocatalyst

**2.2.1. FTIR spectroscopy.** To examine the functional groups of the produced materials, FTIR spectra were obtained for the neat VPMP, Fe<sub>3</sub>O<sub>4</sub>@VPMP, Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS and Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu, as shown in Fig. 1. The results indicated that the existence of the peaks in all the materials spectra appearing at *ca.* 575–590 cm<sup>-1</sup>, 1000 cm<sup>-1</sup> and 3450 cm<sup>-1</sup> are related to the stretching vibrations of the metal oxide bonds (Fe–O), bending vibrations of the Si–O–Si bands of VPMP and stretching vibrations of the O–H bonds, respectively.<sup>36</sup> Furthermore, the presence of Si–OH and SiO–H bonds was

confirmed by the peaks that appeared at 950 cm<sup>-1</sup> and 870 cm<sup>-1</sup>, respectively; these peaks are present in the spectra of all three samples.<sup>37</sup> Finally, the peak that appeared at *ca.* 635 cm<sup>-1</sup> well confirms the formation of Cu–O bonds in the structure of the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocomposite.<sup>38</sup> The peak at 2300 cm<sup>-1</sup> can also be related to CO<sub>2</sub> in the atmosphere, and it is also visible in the spectra of all three samples.<sup>39</sup>

**2.2.2. EDX analysis.** Energy-dispersive X-ray (EDX) spectroscopy should be used to further confirm the presence of elements that are predicted to be present at various stages of composite preparation. As shown in Fig. 2, all the composition elements for the three materials in the current study, Fe<sub>3</sub>O<sub>4</sub>@VPMP, Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS and Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu, were detected and affirmed through the EDX peaks with different intensities. Furthermore, the peak intensity of carbon element in the EDS of the fabricated Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu (Fig. 2a) increased significantly compared to that of the neat VPMP (Fig. 2b) and VPMP@CLS (Fig. 2c) samples, which is due to the good composition of CLS with VPMP in the presence of Cu nanoparticles in the magnetic nanocomposite structure. In addition, it has been clearly demonstrated that 10.1 wt% of the total weight of the fabricated Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu belongs to the Cu NPs.

**2.2.3. Electron microscopy.** In order to investigate the size, morphology, and dispersion state of the prepared Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocomposite, scanning-electron microscopy (SEM) and transmission-electron microscopy (TEM) were used. As can be observed in Fig. 3a, the ground VPMP was obtained from the pumice mass, but there was no uniformity in size or shape. Also, the size of the ground particles is in the range of 500–1100 nm, which is not appropriate. Fig. 3b illustrates the powdered VPMP (obtained by ball-milling), which has high uniformity of its size and morphology. As can be seen, good dispersion of the spherical-shaped VPMP particles with the average sizes of *ca.* 64 nm have been obtained. Also, with a more accurate look at the image (b), it is recognized that the rough structure of the VPMP nanoparticles is quite porous. Fig. 3c exhibits the composition state of the iron oxide (Fe<sub>3</sub>O<sub>4</sub>) nanoparticles with the VPMP nanoparticles. It can be clearly



**Fig. 1** The FTIR spectra of (a) neat VPMP, (b) Fe<sub>3</sub>O<sub>4</sub>@VPMP nanoparticles, (c) Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS nanocomposite, and (d) Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu nanocomposite.





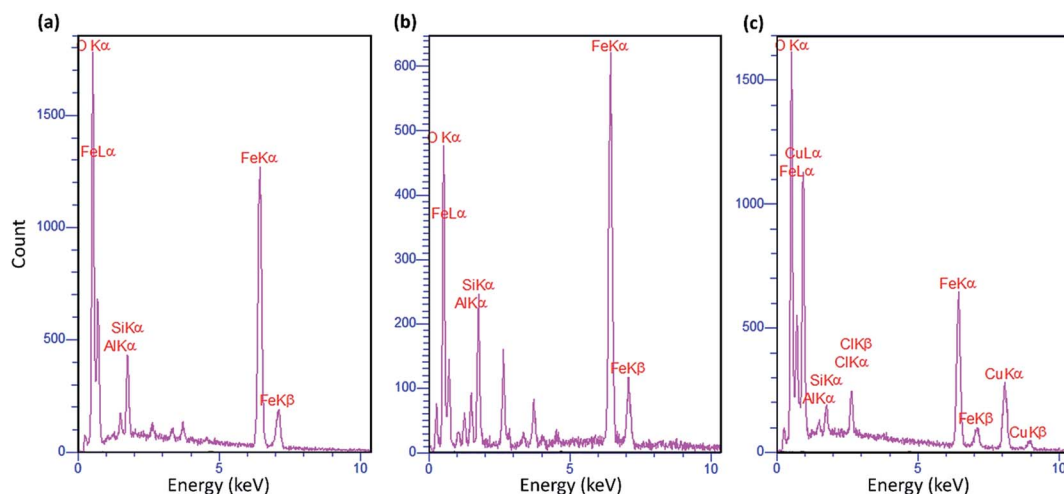


Fig. 2 EDX spectra of (a)  $\text{Fe}_3\text{O}_4@VPMP$ , (b)  $\text{Fe}_3\text{O}_4@VPMP/CLS$ , and (c)  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$ .

observed that the *in situ* preparation and incorporation of  $\text{Fe}_3\text{O}_4$  were successfully performed, as the  $\text{Fe}_3\text{O}_4$  seeds with a mean size of *ca.* 27 nm are finely distributed on the VPMP surfaces.

Moreover, the Fig. 3g-series demonstrates the SEM energy-mapping images of the  $\text{Fe}_3\text{O}_4@VPMP$  nanoparticles that are related to image (e). Fig. 3d shows that the  $\text{Fe}_3\text{O}_4@VPMP/CLS$

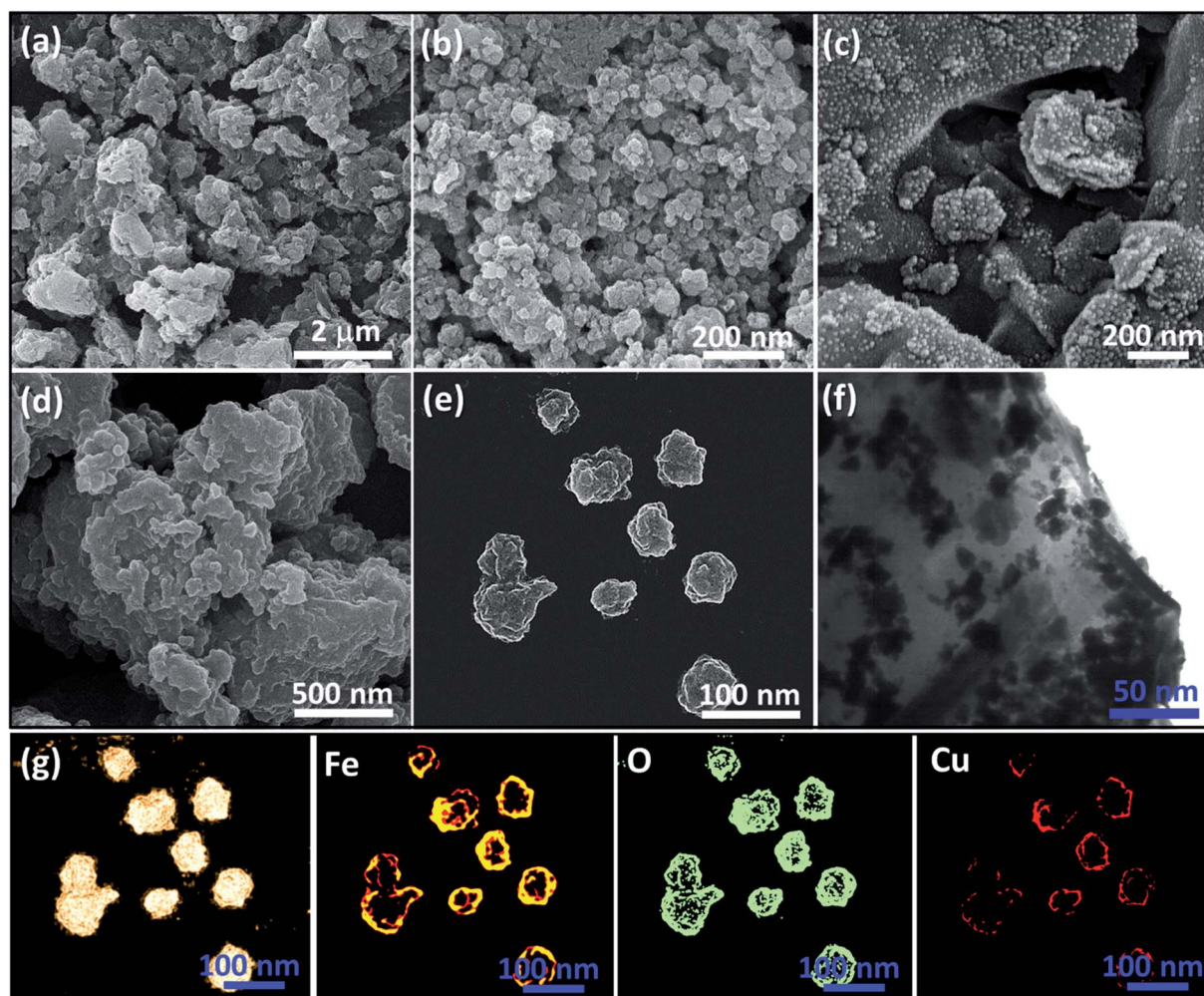


Fig. 3 The SEM images of (a) ground pumice mass, (b) powdered VPMP, (c)  $\text{Fe}_3\text{O}_4@VPMP$  nanoparticles, (d) the integrated  $\text{Fe}_3\text{O}_4@VPMP/CLS$  composite, and (e) the dispersed  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  nanocomposite; (f) the TEM image of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  nanocomposite; and (g-series) SEM energy-mapping images of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  nanocomposite, related to image (e).

particles are slightly agglomerated due to the integration of the CLS polymeric matrix. In this stage, good coating of  $\text{Fe}_3\text{O}_4$ @VPMP by the CLS occurred, and the particle agglomeration is temporary because the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS particles were further redispersed in an alkaline medium (NaOH, 0.5 M) and also washed several times. As can be observed in Fig. 3e, good dispersion of the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu nanocomposite was obtained through the washing of the particles and functionalization with the copper ions. Actually, the excess CLS matrix was removed through re-dispersion *via* the ultrasonication and additional washing. As can be seen in image (e), the uniformity of the particles (in size and shape) was preserved during the copper-functionalization process. Furthermore, the mean size of the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu particles was estimated to be *ca.* 81 nm, which is good for catalytic applications. Ultimately, Fig. 3f illustrates the TEM image of the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu composite, in which the dark spots are attributed to the  $\text{Fe}_3\text{O}_4$  and Cu nanoparticles on the surface of the VPMP/CLS substrate. The energy-mapping images of the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu composite related to image (e) have been illustrated as the “g series”, demonstrating the presence of Cu element on the surfaces. Generally, there are two justifications for incorporation of the Cu nanoparticles into the VPMP/CLS matrix; first, there are strong metal–oxide (M–O) electronic interactions between the Cu nanoparticles and abundant hydroxyl groups in the structure of CLS.<sup>40</sup> Second, the Cu nanoparticles are incorporated into the voids of the VPMP, which is a highly porous substrate.<sup>41</sup>

**2.2.4. VSM analysis.** The vibrating-sample magnetometer (VSM) technique was applied to detect the intrinsic magnetic performance of the arranged  $\text{Fe}_3\text{O}_4$ @VPMP and  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu nanocomposite materials. The hysteresis curves of the  $\text{Fe}_3\text{O}_4$ @VPMP nanocatalyst exhibit about a *ca.*  $10.0 \text{ emu g}^{-1}$  decrease in magnetic saturation, which is related to the phenomenon in which not all of the nanocatalysts were found to present magnetic features (Fig. 4). As is perceived from the magnetic-hysteresis curves, the magnetic behavior of the  $\text{Fe}_3\text{O}_4$ @VPMP nanocomposite has not been considerably reduced after addition of the CLS–Cu particles to form  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu nanocomposite materials. Additionally, the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu nanocomposite materials could be collected, through this excellent magnetic property, without difficulty *via* using a magnet, and they could be reused several times.

**2.2.5. TGA/DTA analysis.** In this section, the thermal decomposition of the neat VPMP and the fabricated  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu system is investigated by thermogravimetric analysis (TGA) and differential thermal analysis (DTA). As shown in Fig. 5, in the temperature range of 150–370 °C, about 10% of the mass of  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu (black curve) was lost, while this reduction in mass for the neat VPMP from 300 °C to 500 °C is about 5%. This mass reduction in the TGA curve of  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu is related to the removal of –OH groups in the CLS structure. According to the literature, the –OH groups of polymers leave the structure during the dehydration process at this thermal range. Also, this change of mass due to the removal of the –OH groups of CLS in the DTA curve shows itself as an endotherm. Above 400 °C, both the Cu nanoparticles and

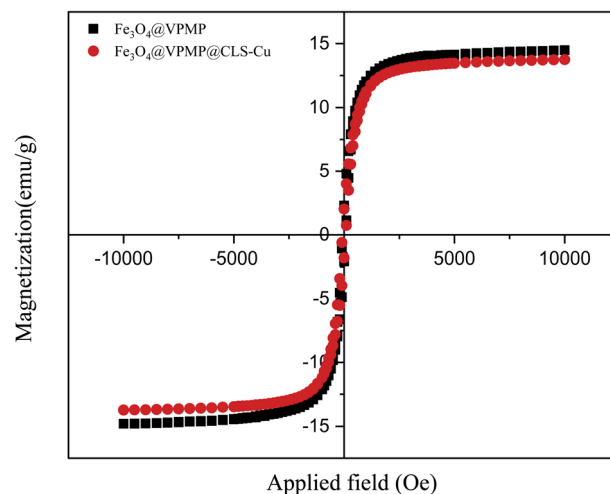


Fig. 4 The VSM curves of  $\text{Fe}_3\text{O}_4$ @VPMP and  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu.

VPMP NPs begin to decompose, where the reduction in mass of the prepared composite is very slow. This mass reduction in the DTA curve is shown as an endotherm. The silicate network has a high ability and capacity to trap various molecules, including water molecules, in its structure; therefore, it seems that the 5% reduction in the neat VPMP mass is probably related to the removal of water molecules trapped in the silicate structure of the VPMP structure.<sup>42</sup> In addition, the first shoulder in the TGA curve of neat VPMP is about 1 wt%, and it is related to the separation of physically adsorbed moisture onto the surfaces. However, this shoulder is not in the TGA curve of  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu, which seems to be due to the good drying of the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu sample under  $\text{N}_2$  before TGA. According to the DTA curve, the studied nanocomposite maintains its structure up to 600 °C, and no melting or crystallization action occurs up to the mentioned temperature in its structure (Fig. 5).

**2.2.6. XRD analysis.** The X-ray diffraction pattern (XRD) for the  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu compound is shown in Fig. 6. In this regard, the pattern of this nanocomposite was compared with the separate patterns of  $\text{Fe}_3\text{O}_4$  and CuO based on the JCPDS database (PDF # 892531 and PDF # 190629). According to the

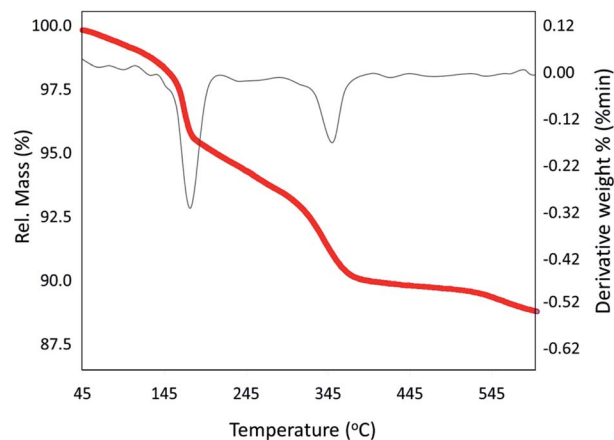


Fig. 5 The TGA (red) and DTA (gray) curves of the fabricated  $\text{Fe}_3\text{O}_4$ @VPMP/CLS–Cu nanocomposite.



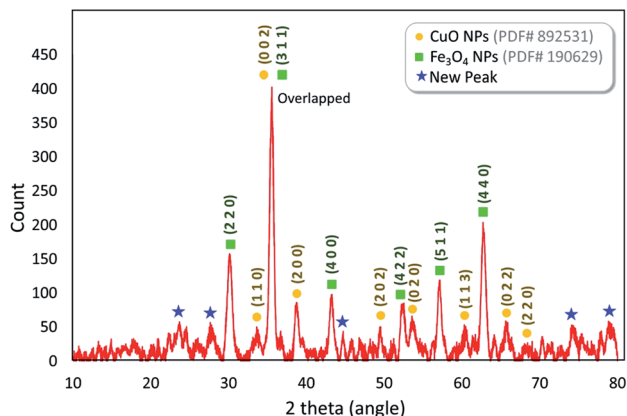


Fig. 6 The XRD pattern of the fabricated  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite in comparison with the reference patterns of CuO and  $\text{Fe}_3\text{O}_4$  NPs.

results, the peaks found at  $2\theta = \sim 30.7^\circ$ ,  $36.2^\circ$ ,  $43.8^\circ$ ,  $52.2^\circ$ ,  $57.7^\circ$ , and  $63.4^\circ$ , which are respectively assigned to the green Miller indices (2 2 0), (3 1 1), (4 0 0), (4 2 2), (5 1 1), and (4 4 0), are related to the XRD patterns of  $\text{Fe}_3\text{O}_4$  magnetic nanoparticles (MNPs). Also, the peaks that appeared at  $2\theta = \sim 33.1^\circ$ ,  $38.8^\circ$ ,  $49.3^\circ$ ,  $53.8^\circ$ ,  $61.7^\circ$ ,  $68.7^\circ$ , and  $69^\circ$ , which were respectively assigned to the yellow Miller indices (1 1 0), (0 0 2), (2 0 0), (2 0 2), (1 1 3), (0 2 2) and (2 2 0), in Fig. 6 are related to the diffraction pattern of the CuO structure in the composite of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ .<sup>43</sup> Also, the new peaks appearing at  $2\theta = \sim 22.3^\circ$ ,  $23.6^\circ$ ,  $44.6^\circ$ ,  $74.2^\circ$ , and  $79.6^\circ$  are related to the amorphous structure of the VPMP structure.

**2.2.7. BET analysis.** Due to the highly porous structure of the pumice used in the synthesis of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite, adsorption/desorption of  $\text{N}_2$  gas Brunauer–Emmett–Teller (BET) surface area analysis was used to confirm the porosity of the synthesized composite, as shown in Fig. 7. According to the obtained results, the pumice structure and  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite indicate the type IV isotherm of mesoporous materials. According to the obtained curves, all the pores of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  were filled by liquid  $\text{N}_2$  at  $P/P_0=0.99$ ; therefore, for our sample, according to the amount adsorbed at  $P/P_0 = 0.99$  ( $72.2 \text{ cm}^3 \text{ g}^{-1}$  STP) and its conversion to liquid volume (gas per liquid volume ratio is 647

for  $\text{N}_2$  at 77 K), we can estimate that the total volume of the composite pores is  $0.5164 \text{ cm}^3 \text{ g}^{-1}$ . Also, according to the results, the reduction of the average pore size from 40 nm for pumice powder to 24.452 nm in the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite along with the total surface area of the nanocomposite, which is  $84.473 \text{ m}^2 \text{ g}^{-1}$ , confirms that although a large number of pumice pores are filled with  $\text{Fe}_3\text{O}_4$  and CLS–Cu magnetic nanoparticles, large internal pores are still available for the catalytic performance of the nanocomposite.

**2.2.8. XPS analysis.** In order to evaluate the chemical state of the copper nanoparticles ( $\text{Cu}^0$ ,  $\text{Cu}^+$  or  $\text{Cu}^{2+}$ ) in the structure of the composite, X-ray photoelectron spectroscopy (XPS) was used. As demonstrated in Fig. 8a, the presence of Al/Si, Fe, and Cu elements in the structure of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite was confirmed; these elements originate from the VPMP,  $\text{Fe}_3\text{O}_4$  and Cu nanoparticles, respectively. In the expanded spectrum of the copper element related to Cu 2p (Fig. 8b), there are two peaks at *ca.* 933 and 936 eV marked by  $\text{Cu}^+$  and  $\text{Cu}^{2+}$ , respectively, confirming the formation of  $\text{Cu}_2\text{O}$  and CuO nanoparticles in the structure of the catalyst.<sup>44</sup> Therefore, it is concluded that the stage of reduction with sodium borohydride is required before beginning the catalytic process for the nitrobenzene derivatives. Also, the chemical state of the copper particles after the reduction process by sodium borohydride in alkaline conditions was investigated by this method. As can be observed in Fig. 8c and d, the peak of the copper element (Cu 2p) appeared at *ca.* 953 eV, confirming the chemical state  $\text{Cu}^0$  after the reduction process.<sup>45</sup>

### 2.3. Application of $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ nanocatalyst in the reduction reactions of nitrobenzene derivatives (NBDs)

**2.3.1. Optimization of the catalytic system in the reduction of nitrobenzene.** In order to determine the optimized conditions for the use of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system, the catalytic amount of the nanocomposite, amount of  $\text{NaBH}_4$ , medium, temperature and reaction time used in the reduction reaction of nitrobenzene (NB) must be carefully monitored. Details of this investigation are reported in the table below. As shown in the table, in order to prove the important role of pumice porosity in the catalytic activity of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ , CLS–Cu alone was applied under the same conditions

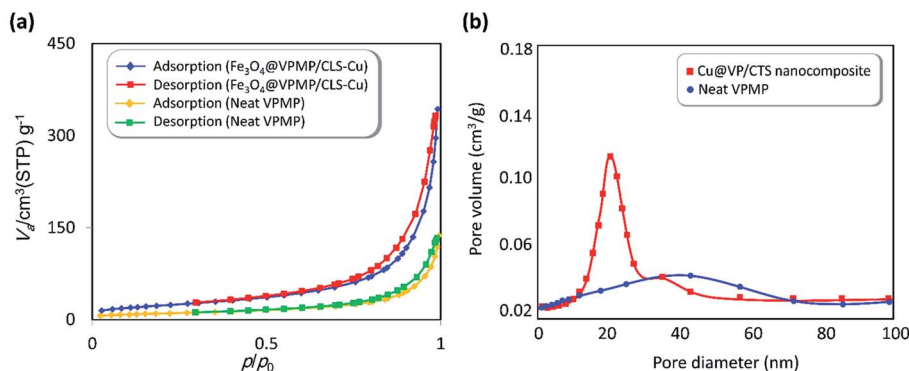


Fig. 7 (a) The BET curves of the neat VPMP and the fabricated  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite and (b) the mean pore volumes of the neat VPMP and the fabricated  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite, via adsorption/desorption of  $\text{N}_2$  under ambient conditions.





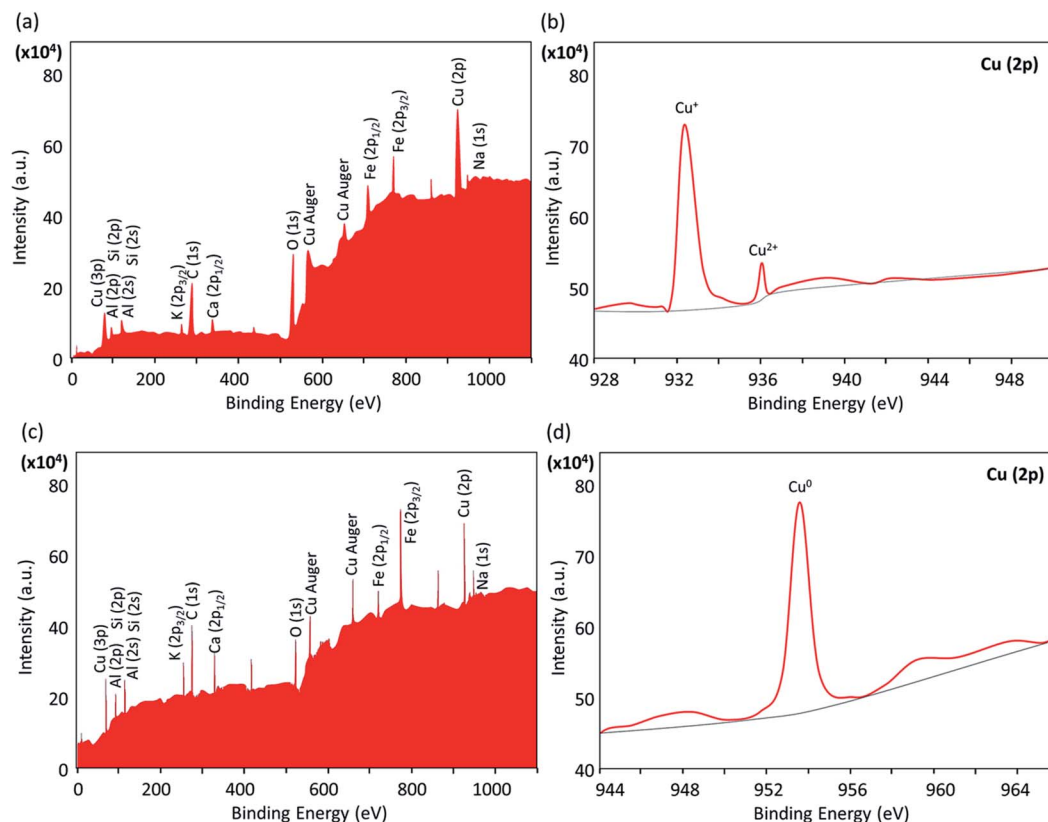


Fig. 8 The XPS spectra of the prepared  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite: (a and b) before and (c and d) after reduction by  $\text{NaBH}_4$ ; (a and c) standard form and (b and d) expanded form of a specific area.

during the reduction reaction, which showed a reduction in the reaction yield. This confirms the role of the pumice porosity in the catalytic activity of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ . Also, the role of copper as the main catalytic site was investigated, which, as predicted, showed that the non-use of copper in the catalytic system leads to a sharp decrease in the yield of the reaction. From the table, it has been revealed that the optimum conditions were provided by using 0.05 g of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalyst during 8 min of stirring at  $70^\circ\text{C}$ .

**2.3.2. Catalyzed synthesis of the aniline derivatives.** In order to study the catalytic efficiency of the prepared  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  composite, NBDs were utilized, and the results are shown in Table 2. Interestingly, by using the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system, in a short time, a high yield reaction was obtained from the reduction reactions of NBDs, which proves that the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system is a promising candidate for reduction of NBDs compared to other catalysts. In addition, the  $^1\text{H-NMR}$  and  $^{13}\text{C-NMR}$  spectra and spectral data of aniline as well as its derivatives are founded in the ESI file (Fig. S1–S20†).

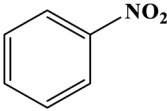
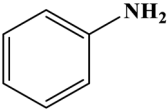
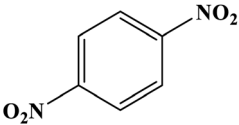
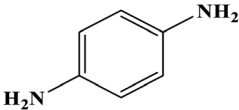
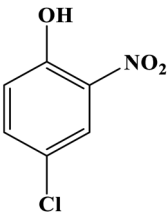
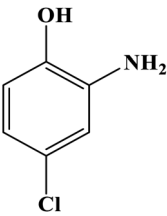
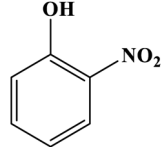
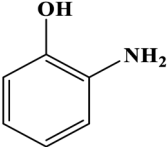
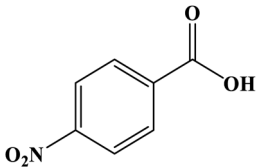
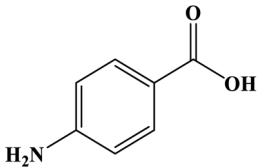
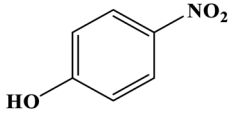
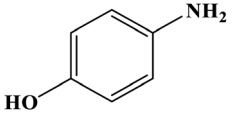
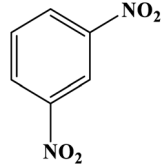
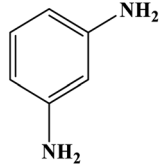
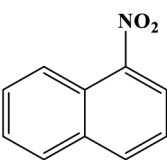
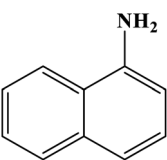
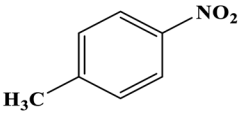
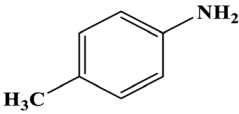
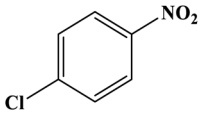
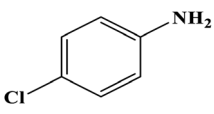
**2.3.3. Suggested mechanism.** As shown in Scheme 2, a plausible mechanism can be suggested for the reduction process of the NBDs under catalytic conditions provided by the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  nanocomposite. At the first stage, the  $\text{Cu}^{2+}$  ions as the main catalytic sites are reduced to  $\text{Cu}^0$  by sodium borohydride ( $\text{NaBH}_4$ ) under the alkaline conditions provided by potassium carbonate ( $\text{K}_2\text{CO}_3$ ).<sup>54</sup> In stage 2, both the NB and borohydride ( $\text{BH}_4^-$ ) structures are attracted onto the

catalyst surface through electronic and hydrogen bond interactions.<sup>55</sup> In third stage, hydridation of the NB structure is performed by the borohydride in the vicinity of the NB on the surface of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  particle. In the alkaline conditions, sodium metaborate ( $\text{NaB(OH)}_4$ ) is produced after completion of the reduction reaction of the NB. Finally, in stage 4, the produced aniline structure leaves the surface of the catalyst, and the particles are magnetically separated from the mixture, washed, and reused several times.

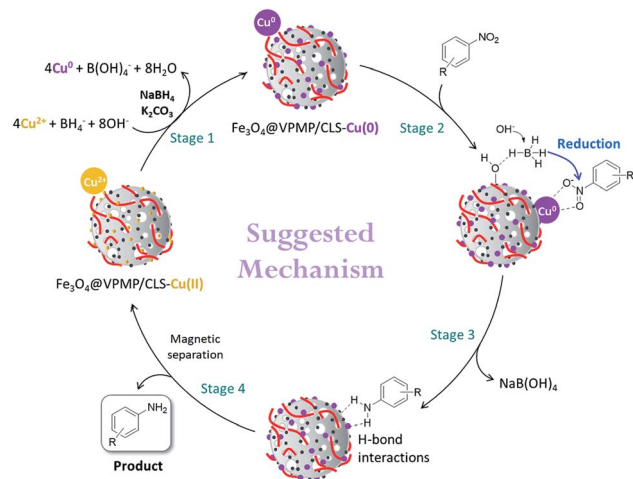
**2.3.4. Recyclability effect of the nanocomposite catalytic system.** In order to evaluate the reusability of the prepared  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system, a reasonable investigation was performed on the model reduction reaction of nitrobenzene (NB). Firstly, the half-life of the reaction was estimated in five successive runs of the reaction. The details of the experiment and the obtained results are reported in Table S1 and Fig. S21, ESI file.† As reported in Table S2 (in the ESI file†), the time of the half-life of the reduction reaction of the NB changed to higher values after each cycle of separation and re-preparation of the catalyst particles. Based on the obtained results, it can be concluded that the catalytic efficiency of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  system is gradually lost during the recycling and reusing processes. To investigate the possible reasons, probable leaching of the copper particles from the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  system as the main catalytic site was checked by inductively coupled plasma mass spectrometry (ICP-MS) analysis of the supernatant after the separation of the



Table 2 Gained yields after the reduction of NB derivatives to aniline analogues under the optimized conditions

Entry	NB structure	Product	Time (min)	Yield (%)	Mp (°C)	
					Found	Reported
1			8	98	Liquid sample	(Ref. 46)
2			13	97	145–146	144–146 (ref. 47)
3			7	96	135–138	136–138 (ref. 48)
4			3	98	172–175	171–173 (ref. 49)
5			4	96	187–189	186–188 (ref. 46)
6			8	95	184–186	185–187 (ref. 49)
7			13	89	65–66	65–66 (ref. 50)
8			12	86	51–53	52–53 (ref. 51)
9			7	97	42–44	42–43 (ref. 52)
10			9	97	70–71	70–72 (ref. 53)

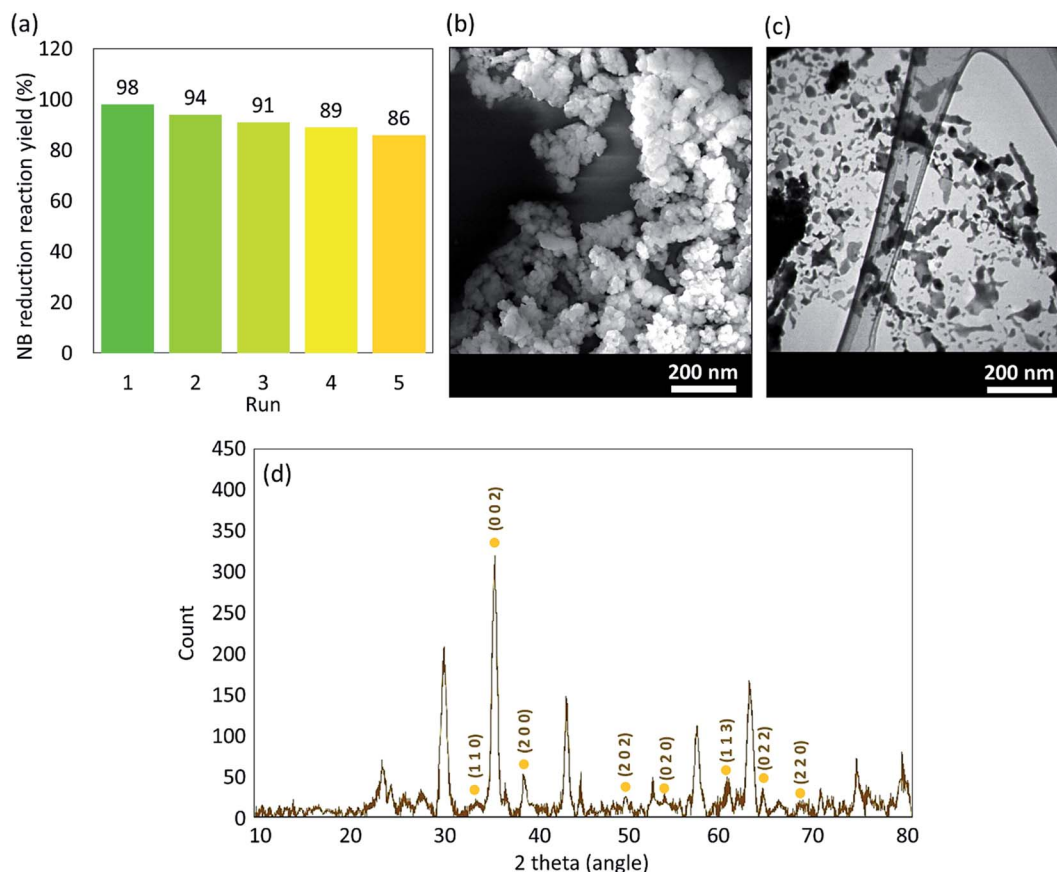




**Scheme 2** A plausible multi-stage mechanism for the reduction of the NBDs catalyzed by the fabricated  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  catalytic system.

magnetic particles after running each recycling process. As can be seen in Table S3 (in the ESI†), a small amount of leaching of the copper particles occurred during the separation, rinsing, drying and reusing of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  particles in five successive runs. Most likely, the copper leaching occurred during the ultrasonication process that is performed for dispersion of the particles before each run. The resulting

reaction yields during five-time reuse of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  particles are given in the bar chart in Fig. 9a. As can be seen, it can be concluded that the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  particles can be recovered and reused at least five successive times after magnetic separation, rinsing, and drying of the particles. Obviously, this has economic benefits, particularly in large scale applications. In order to investigate any possible change in the structure of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  nanocatalyst after five-time recycling, SEM and TEM imaging of the recovered particles were carried out. As illustrated in Fig. 9b and c, the uniformity and integrity of the VPMP,  $\text{Fe}_3\text{O}_4$  and Cu nanoparticles in the CLS matrix have been suitably maintained, which confirms the good physicochemical stability of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  composite. Also, the crystalline structure of the recovered  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  composite was studied by XRD spectroscopy. As can be observed in Fig. 9d, the same pattern with the fresh catalyst (Fig. 6) was obtained after recycling five times; only the intensity of the peaks related to the CuO nanoparticles (marked with Miller indices) decreased slightly due to the copper leaching. The content of Cu in the recovered structure was evaluated by inductively coupled plasma optical emission spectrometry (ICP-OES) analysis, and the obtained results, which have been given in Table S4 (in the ESI†), confirm the Cu-leaching from the catalytic system during the recycling. Besides, the graph of the optimization stages of reduction reaction of nitrobenzene under



**Fig. 9** (a) Recyclability diagram of the  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  catalytic system in the reduction reaction of NB to aniline; (b) SEM image, (c) TEM image, and (d) XRD pattern of the recovered  $\text{Fe}_3\text{O}_4@VPMP/CLS\text{-Cu}$  particles after recycling five times.



Table 3 Catalytic systems previously reported for reduction of NB derivatives to aniline analogues

Entry	Catalyst	Conditions	Cat. (mg)	Time (min)	Conversion (%)	Ref.
1	Ag-MPTA-1 <sup>a</sup>	80 °C/NaOH/N <sub>2</sub>	250	600	97	56
2	Ag-SBA-15 <sup>b</sup>	R.t./NaBH <sub>4</sub>	10	6	95	57
3	Nano Co <sup>c</sup>	R.t./N <sub>2</sub> H <sub>4</sub> ·H <sub>2</sub> O	5.6	30	50	58
4	PhNO <sub>2</sub> , Pd NPs/RGO <sup>d</sup>	50 °C/NaBH <sub>4</sub>	6.0	90	98	59
5	Cu NPs <sup>e</sup>	50 °C	6.0	120	95	60
6	Fe <sub>3</sub> O <sub>4</sub> @VPMP/CLS-Cu <sup>f</sup>	70 °C	5.0	8	98	—

<sup>a</sup> Poly-triallylamine (MPTA-1). <sup>b</sup> SBA-15 is a mesoporous silica. <sup>c</sup> Cobalt nanoparticles. <sup>d</sup> Reduced graphene oxide. <sup>e</sup> Copper nanoparticles. <sup>f</sup> The present work.

various catalytic conditions (related to Table 1) is illustrated in Fig. S22, ESI file.†

**2.3.5. Comparison of the prepared Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu with other catalytic systems.** In this section, we attempt to make a quick comparison between our presented nanocatalyst and a number of other previously reported systems that, like the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu catalyst, have the ability to convert NBDs to their aniline analogues. Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu has advantages over the other catalysts discussed in the table below: (1) in general, heterogeneous catalytic systems are preferred to homogeneous catalytic systems due to their convenient separation process. The VPMP, which is used in our presented nanocatalyst, has a magnetic property itself due to the presence of iron in its structure; however, to increase this property, Fe<sub>3</sub>O<sub>4</sub> was added to the synthetic composite, which leads to easier separation of this catalytic composite from the reaction medium. Significant conversion (98%) of different NB derivatives occurred in a short time (often less than 10 minutes), while in most cases (entries 1, 3, 4 and 5), the reaction time exceeded 90 minutes. In the design and preparation phase, it is very important to use materials that, in addition to benefiting from high degrees of biocompatibility, are inexpensive; therefore, the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS-Cu composite, due to the use of abundant natural materials (such as VP rock), which has many economic advantages, can be considered superior to the systems reported in the table below (entry 2), although the results obtained by the reported catalytic systems are excellent. In addition to the mentioned advantages, other advantages of this catalytic system are easy preparation of the composite in question and use of lightweight substrates with a porous structure, which leads to good dispersion of the catalyst particles in the reaction medium and an overall increase in performance (Table 3).

### 3. Experimental

#### 3.1. Chemicals and instruments

All the chemicals, reagents, and equipment used in this study are listed in Table 4.

#### 3.2. Preparation methods

**3.2.1. Preparation of magnetic Fe<sub>3</sub>O<sub>4</sub>@VPMP micro-plates.** First, the purchased pumice was placed in a milling bowl and ground by ball-milling (20 Hz) for two hours. The ground pumice was then poured into a crucible and placed in a furnace until its temperature reached 400 °C within an hour, and then

the temperature slowly reached the ambient temperature as it was inside the furnace and without any temperature shock. The pumice was then removed from the furnace and washed several times with HCl (1 M). Then, 2.2 g of the prepared pumice was dispersed with 5.0 mmol of FeCl<sub>3</sub>, 5.0 mmol of FeCl<sub>2</sub> and 50.0 mL of deionized water in a 100.0 mL balloon for 30 minutes at 45 °C *via* ultrasonication using a cleaner bath. It was then placed on a stirrer at 85 °C and stirred vigorously under N<sub>2</sub> atmosphere for two hours. After this time, 7.5 mL of 25% ammonia was added dropwise in one hour. After this process, ultimately, the product was magnetically separated from the water, washed with deionized water several times, and then dried in an oven.

**3.2.2. Preparation of the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS composite.** At this stage, 0.11 g of cellulose powder was initially dispersed in 20.0 mL of deionized water at room temperature. Then, 1.0 g of Fe<sub>3</sub>O<sub>4</sub>@VPMP particles were added with vigorous stirring to the prepared cellulose solution. After 6 hours, the Fe<sub>3</sub>O<sub>4</sub>@VPMP/CLS composite was magnetically separated and washed several times with deionized water.

Table 4 Chemicals and equipment used in this study

Materials and equipment	Purity and brand
Pumice stone	Granulated – Sigma Aldrich
Cellulose	Microcrystalline – Sigma Aldrich
FeCl <sub>2</sub> ·4H <sub>2</sub> O	Sigma Aldrich (98%)
FeCl <sub>3</sub> ·6H <sub>2</sub> O	Sigma Aldrich (≥98%)
Ammonia	Merck (25%)
Sodium hydroxide	Merck, 97%
CuCl <sub>2</sub> ·2H <sub>2</sub> O	Sigma Aldrich (≥99%)
Ethanol	Sigma Aldrich – 97%
Sodium borohydride	Sigma Aldrich (≥96%)
FT-IR analysis	Shimadzu IR-470 spectrometer
EDX analysis	Numerix DXP-X10P
SEM analysis	Sigma-Zeiss microscope
TEM analysis	Philips Cm 12 Instrument
VSM analysis	Lakeshore 7407
TGA/DTA analysis	STA504 device
XRD analysis	JEOL JDX-8030 (30 kV, 20 mA)
BET analysis	Micromeritics ASAP 2010
XPS analysis	K-Alpha+
ICP-OES analysis	DV 5300
ICP-MS analysis	ELAN 6100 DRC-e
NMR analysis	Varian Unity Inova 500 MHz
Ultrasound cleaning bath	KQ-250 DE (40 kHz, 250 W)
Melting point measurement	Electrothermal 9100, made in UK



**3.2.3. Preparation of the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$  composite.** At this stage, in order to prepare the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$  composite, initially, 75.0 mL of 0.04 M solution of  $\text{CuCl}_2\cdot 3\text{H}_2\text{O}$  was prepared. Then, in another flask,  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS}$  particles (1.0 g) were dispersed in deionized water *via* ultrasonication, and the as-prepared solution of NaOH (0.5 M, 50.0 mL) was added dropwise during the ultrasonication at room temperature. Afterward, the pre-prepared solution of  $\text{CuCl}_2\cdot 3\text{H}_2\text{O}$  was slowly added to the flask of solution containing  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS}$  particles and stirred vigorously. After 6 hours, the  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$  composite was magnetically separated and washed several times with deionized water.

### 3.3. A general procedure for reduction of nitroarenes with the $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(0)}$ system

To a round-bottom flask (25 mL) containing deionized water (5 mL), nitrobenzene (1 mmol, 0.123 g) and  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu(II)}$  (0.05 g) were added. Then, the pH of the mixture was tuned to *ca.* 8.0 by addition of  $\text{K}_2\text{CO}_3$  (0.04 g). Next,  $\text{NaBH}_4$  (2 mmol, 0.075 g) was added to the mixture, and the resulting mixture was stirred at 70 °C. After completion of the reaction, the catalyst was separated by an external magnet and the product was extracted with EtOAc. The organic layer was dried over anhydrous  $\text{Na}_2\text{SO}_4$ . Evaporation of the solvent under reduced pressure afforded the pure aniline in 98% yield.

### 3.4. Large scale experiment on the reduction of nitrobenzene by the $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$ catalytic system

To evaluate the efficiency of the prepared  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  catalytic system in industrial applications, a gram scale model experiment was carried out. For this purpose, in a round bottom flask (500 mL), 40 mmol (4.9 g) of nitrobenzene was dissolved in 160 mL of water. Then, the pH of the mixture was tuned to *ca.* 8.0 by addition of  $\text{K}_2\text{CO}_3$  (1.6 g), and the temperature of the flask was controlled with an ice bath. In a separate flask (100 mL), 2.0 g of  $\text{Fe}_3\text{O}_4\text{@VPMP/CLS-Cu}$  powder was well dispersed in 40 mL of water *via* ultrasonication by a cleaner bath (50 KHz, 100 W  $\text{L}^{-1}$ ) and finally added to the main flask (500 mL) containing the nitrobenzene solution. Then,  $\text{NaBH}_4$  (80 mmol, 3.0 g) was added to the mixture, and the resulting mixture was stirred in an ice bath. At this stage, the reaction temperature was maintained at around 20 °C. After completion of the reaction, the magnetic particles were separated with an external magnet several times. Next, the residue was divided into five 40 mL portions, and liquid-liquid extraction was performed using EtOAc (20 mL) (three times for each portion). Finally, the organic phase was dried over anhydrous  $\text{Na}_2\text{SO}_4$  (5.0 g), and the solvent was evaporated under reduced pressure (pure aniline in 98% yield).

## 4. Conclusion

In this study, a magnetic catalytic system using natural-based materials was presented that has the ability to reduce various derivatives of nitrobenzene (NB) in a very short time of 8 minutes. This efficient catalytic system was characterized by FT-IR spectroscopy, SEM, EDX spectroscopy, XRD spectroscopy, TGA, and VSM. This catalytic system has many advantages, such

as (1) convenient separation, which is due to the super-paramagnetic behavior of the nanoscale catalytic system, which was also proven by VSM analysis; (2) due to the high porosity of the pumice according to SEM, this catalytic system is able to provide a large active surface area; and (3) the use of low amounts of the heterogeneous catalyst (0.05 g) and many other advantages can introduce this composite as a suitable candidate for scaling up and industrial applications.

## Author contributions

R. T. L. designed the project and led the hypothesis. M. S. and F. S. Q. carried out the experimental sections and provided the required analyses. R. T. L. and M. S. prepared the first draft of manuscript. All the authors helped in interpreting the obtained results, writing, editing and revising the manuscript. A. M. and A. E. S. managed and supervised all the work sections.

## Conflicts of interest

The authors listed in this article have no conflict of interests.

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