The photophysical properties of naphthalene bridged disilanes†

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Structural isomers of naphthalene-bridged disilanes were prepared via catalytic intramolecular dehydrocoupling of disilyl precursors using Wilkinson’s catalyst. Interestingly, it was observed that interchanging the side groups on the silicon atoms altered the photophysical properties of the bridged disilanes. Herein, we report the first example of naphthalene bridged disilanes forming excimers in non-polar solvents. Cyclic voltammetry experiments and DFT calculations were performed to analyse the band gaps of the compounds and $\sigma - \pi$ mixing in the bridged disilanes.

Introduction

The unique sigma electron delocalisation ($\sigma$-conjugation) present across catenated Si atoms grants polysilanes useful properties, such as absorption in the UV region and semi-conducting capability, when compared to the analogous carbon-based polymers.1–4 This feature is attributed to the fact that silicon has a lower ionization energy and a higher electron affinity compared to carbon, resulting in a lower singlet $\sigma_{\text{sil}} - \sigma_{\text{si}}$ excitation energy in Si-Si bonds compared with the $\sigma_{\text{CC}} - \sigma_{\text{CC}}^*$ excitation energy in C-C bonds.4 Polysilanes typically absorb in the range 250–350 nm and the absorption is greatly influenced by the side groups attached to the silicon atoms. Previous reports have shown that the presence of aromatic groups on the silicon atoms cause a red shift of 25–35 nm in the UV absorption spectra and enhance $\sigma - \pi$ mixing thereby increasing the conjugation across the system.5 Organo-silanes have a number of applications in both medicinal chemistry6–8 and as energy materials.9 The photoluminescent properties of silicon based materials have provided a new platform for polysilanes to act as chemo sensors for detecting explosive materials such as 2,4,6-trinitrotoluene (TNT) and picric acid.10–12 Being photoactive in nature, long chain organosilanes have also been tested for photocconductivity and photocurrent generation.13–15

Despite having the potential to be used in various applications, the Si-Si sigma bonds cleave upon excessive exposure to UV radiation, therefore limiting the applications of these materials.5 For example, successive blue shifts were observed in the UV spectra of oligosilanes over time indicating photodecomposition.16 Sakurai et al. proposed a radical mechanism for the photodecomposition of the Si-Si bonds as they successfully isolated the silyl radicals, as detected by electron paramagnetic resonance spectroscopy, for an arylsilane in cyclohexane.17

We propose that reinforcing the Si-Si bond in a disilane, using a tether in the form of a covalent bridge, is a potential way to support the weak Si-Si bonds. In the presence of a bridge, upon UV irradiation, the silyl radicals formed during photodecomposition will remain in close proximity to each other, therefore creating a greater chance for the two silyl radicals to recombine, reforming the covalent Si-Si bond. The bridge could be either alkyl or aromatic. Since aromatic groups attached to the silicon atoms leads to $\sigma - \pi$ mixing, attaching aromatic substituents will presumably enhance the semi-conducting properties of the material owing to increased conjugation. Evidence for this was observed in studies conducted by Klausen et al.19 wherein the conductivity of the disilanes containing sulfur tethers were measured by a scanning tunnelling microscope (STM) break junction technique.

In our quest to prepare oligosilanes that are robust, with increased conjugation that can ultimately be used for optoelectronics, substituted naphthalene bridged disilanes seemed to be the best candidates. We have previously reported20 that naphthalene disilanes increase the overall conjugation owing to the $\sigma - \pi$ mixing and also benefit the geometry of the resulting five membered ring system as the Si-Si bond length matches with that of the 1,8-substitution of naphthalene. In this study, we investigated the effects of modifying the substitution around the silicon atoms in the naphthalene bridged disilanes on their photophysical properties in organic solvents of varying polarity. The three target bridged disilanes (1b, 2b and 3b, Fig. 1), two of which are structural isomers of each other, were prepared.
Results and discussion

The target naphthalene bridged disilane precursors 1a, 2a and 3a (Scheme 1) were synthesised using previously reported methodologies (Fig. S1–S7†). A two-step lithiation process to synthesize 1a was used in order to introduce the two different silanes. 1,8-Dibromonaphthalene was monolithiated with n-butyllithium and further reacted with a chlorodimethylsilane (ClSiMe2H) to form 1. After purification, 1 was further lithiated and reacted with chlorodiphenylsilane (ClSiPh2H) to produce 1a as a white solid. The disilyl naphthalene derivatives (2a and 3a) were prepared from 1,8-dibromonaphthalene, which was reacted with n-butyllithium to form 1,8-dilithiated naphthalene in situ. Subsequently, followed by a slight excess of the respective chlorosilanes (e.g. PhMeSiCH or Me2SiCH) the formation of the required chiral and tetramethyl substituted disilyl naphthalene derivatives, 2a and 3a respectively, were obtained.

The disilyl precursors (1a, 2a and 3a) were subjected to intramolecular dehydrocoupling using 5 mol% of RhCl(PPh3)3 in toluene to produce the asymmetric naphthalene bridged disilane 1b (Fig. S8–S10†) and the symmetric naphthalene bridged disilane 2b and 3b (Scheme 2). Interestingly, it was observed that 1b and 3b existed as a crystalline solid, whereas, the chiral 2b was obtained as a waxy substance. Product 1b was dissolved in pentane and stored at −20 °C to grow single crystals (Fig. 2). However, repeated attempts to crystallise 2b, such as freezing at −20 °C and −40 °C in pentane for 24 h and 2 h, respectively, did not yield any crystalline solids, leaving the waxy substance after removal of pentane.

The five-membered ring in 1b was observed to be slightly distorted from planarity. The C13–Si1–Si2 bond angle was 93.02° compared with the C21–Si2–Si1 bond angle of 91.86° indicating the effect of steric bulk on the two silicon centres (Table S1†). The bond angles around the silicon centre bearing the bulkier phenyl rings were narrower compared with the silicon centre containing methyl groups. The Si1–Si2 bond distance was 2.33 Å which was similar to the 3b derivative.19 NMR spectroscopy and GC-MS analysis of 2b (Fig. S11–S15†) confirmed that the product obtained was a mixture of two isomers. The cis and trans isomers of the product 2b could not be separated by silica gel column chromatography or high-pressure flash chromatography as the mixture showed no separation on TLC despite varying the solvent polarities.

For the studies of the photophysical properties, naphthalene was selected as the parent compound. Absorption and emission spectra of naphthalene and the three-bridged disilanes (1b, 2b and 3b) were obtained in organic solvent of varying polarity (Table 1 and Fig. S17–S20†). In comparison to naphthalene, the absorption spectra of the bridged disilanes 1b, 2b and 3b displayed red-shifted absorption, thereby indicating the extended conjugation in the bridged systems. The absorption maxima of the naphthalene and compounds 1b–3b showed very little dependency on solvent polarity. The emission profiles were recorded in the same set of solvents (Fig. 3) at an excitation wavelength of 305 nm for 1b, 2b and 3b and 286 nm for naphthalene. In THF and acetonitrile, the emission spectra of the bridged silanes (1b, 2b and 3b) exhibited vibronic fine structure (maxima at ca. 335 and 345 nm), displaying a close resemblance to that of naphthalene. However, upon moving to non-polar solvents (cyclohexane and toluene), an additional broad emission band at longer wavelengths (>400 nm) is observed, suggesting the formation of excimers (excited state dimers, Fig. 3). The presence of this additional emission band in non-polar solvents was not observed for naphthalene.

Fig. 1 Substituted disilyl naphthalene precursors (1a, 2a and 3a) and naphthalene bridged disilanes (1b, 2b, and 3b).

Scheme 1 Synthesis of disilyl precursors 1a, 2a and 3a.

Scheme 2 Intramolecular dehydrocoupling of 1a, 2a and 3a to produce the bridged disilanes 1b, 2b and 3b.
Excimers form between identical molecules, when a monomer in the excited state interacts with another monomer in the ground state, which subsequently relaxes by dissociation and involves emission of photons. Concentration studies were conducted in cyclohexane (Fig. 4) and THF (Fig. S21) to provide further evidence for the presence of excimers in non-polar solutions. As excimers are favoured at higher concentrations, it was anticipated that the broad band observed at >400 nm in cyclohexane would decrease upon dilution. As illustrated in Fig. 4(ii), the emission spectrum of compound 1b in cyclohexane is highly dependent on concentration. Upon dilution of a 2.54 × 10−4 M solution of 1b, the broad emission band at 407 nm significantly decreases. The same trend was also observed for compounds 2b and 3b (Fig. 4(iii and iv)). In contrast, the emission spectrum of naphthalene in cyclohexane at varying concentrations remained unchanged (Fig. 4(i)). Measurements of the photoluminescence quantum yields of the naphthalene bridged disilanes (1b, 2b and 3b) corroborated these results (Table 1). In THF, the three derivatives exhibit weak to moderate fluorescence (compounds 1b and 3b \( \Phi_F < 0.01 \); compound 2b \( \Phi_F = 0.09 \)). However, in non-polar solvents (cyclohexane and toluene) all three compounds show an increase in fluorescence quantum yields (\( \Phi_F = 0.15–0.32 \)) due to the formation of excited state dimers. Interestingly, for 2b, a high fluorescence quantum yield in acetonitrile \( (\Phi_F = 0.49) \) is also observed.

Several literature examples of excimer formation in molecules containing a naphthalene moiety have been reported. Pandeeswer et al. and Boer et al. observed broad peaks in the visible regions for naphthalene diimides in polar as well as non-polar solvents. Similarly, L’Her et al. observed broad peaks in the visible regions for imidazolium-naphthalene salts in DCM. However, to the best of our knowledge, this is the first example where excimers are observed through intermolecular interactions between naphthalene in bridged disilane compounds. It is also important to note that this phenomenon was not observed in the disilyl naphthalene precursors, prior to the formation of the Si-Si bond. This was evident from the detailed photophysical study conducted by Maeda et al. on various silyl naphthalene compounds, similar to 1b–3b in non-polar solvents. As such, it appears that the Si-Si bond plays a key role in the formation of excimers in the compounds 1b–3b. The work of Karatsu et al. also highlighted the significance the presence of the Si-Si bond has on excimer formation. In their investigation, the photophysical properties of a series of permethyloligosilanes of varying length \((\text{Me}_2\text{Si})_n, n = 1–4, 6\), containing naphthalene or anthracene groups on the terminal silicon atoms, were explored. The two aromatic groups on the ends of the oligosilane chains have the ability interact via \( \pi \)-stacking, forming an intramolecular sandwich. In cases where the number of silicon atoms in the chain was greater than 1, intramolecular excimer formation

![Fig. 2 Molecular structures of 1a (left) and 1b (right, only one of the two crystallographically independent molecules is shown) with thermal ellipsoids shown at 50% probability level.](image)

Table 1 Spectroscopic data for naphthalene and compounds 1b, 2b, 3b

<table>
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<tr>
<th>Compound</th>
<th>Solvent</th>
<th>( \lambda_{\text{abs}} ) (nm)</th>
<th>( \epsilon_{\text{max}} ) (M(^{-1}) cm(^{-1}))</th>
<th>( \lambda_{\text{em}} ) (nm)</th>
<th>( \Phi_F )</th>
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<td>nd</td>
<td>nd</td>
<td>nd</td>
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<td>335, 427</td>
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<td></td>
<td>ACN</td>
<td>290</td>
<td>nd</td>
<td>335</td>
<td>0.13</td>
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\( \text{nd} \) = not determined. Wavelength of the absorption maximum. Wavelength of the emission maximum; for naphthalene, excitation wavelength was 286 nm. For compounds 1b–3b, excitation wavelength was 305 nm. Fluorescence quantum yield for compounds 1b–3b in different solvent, excitation wavelength range 285–295 nm. Emission band ascribed to the formation of excimers. Emission band ascribed to the formation of excimers. Naphthalene was not measured in toluene due to the strong absorption of toluene in the UV region in which naphthalene absorbs.
was observed in cyclohexane. Interestingly, when the oligo-
silane chain was tetramethyl disilane \((n = 2)\), the strongest
excimer fluorescence was observed,\(^{27}\) indicating the disilane
tether provides the optimal distance between the two end
groups (i.e. naphthalene or anthracene). Furthermore, unlike
in our study, excimer formation occurred for the shorter
chain oligomers \((n = 2, 3\) for naphthalene and \(n = 2–4\) for
anthracene) in polar solvent (e.g. acetonitrile) as well.

Fig. 3 Absorption (dashed) and emission (solid) profiles of (i) naphthalene, (ii) \(1b\), (iii) \(2b\), and (iv) \(3b\) in cyclohexane (black), toluene (red), THF
(blue), acetonitrile (green). Absorption spectrum is that recorded in THF.

Fig. 4 Concentration dependent emission of (i) naphthalene upon excitation at 286 nm \((1.76 \times 10^{-4} \text{ M (black)} \rightarrow 2.33 \times 10^{-5} \text{ M (blue)})\), (ii) \(1b\)
upon excitation at 305 nm \((2.54 \times 10^{-4} \text{ M (black)} \rightarrow 3.10 \times 10^{-6} \text{ M (blue)})\), (iii) \(2b\) upon excitation at 305 nm \((1.36 \times 10^{-4} \text{ M (black)} \rightarrow 1.61 \times
10^{-6} \text{ M (blue)})\) and (iv) \(3b\) upon excitation at 305 nm \((3.20 \times 10^{-5} \text{ M (black)} \rightarrow 1.20 \times 10^{-6} \text{ M (blue)})\) in cyclohexane.
The electrochemical behaviour of the bridged disilanes (1b, 2b, 3b) was determined using cyclic voltammetry (CV) in dry acetonitrile under inert atmosphere with ferrocene as an internal standard to calibrate the potentials (Fig. S25–S27†). The electrochemical oxidation (E_{ox}) and reduction (E_{red}) potentials were used to calculate the HOMO–LUMO band gaps of the three bridged disilanes relative to ferrocene/ferrocenium (4.8 eV below the vacuum level).\(^{31,32}\) The band gaps for 1b and 2b were found to be similar (3.62 eV and 3.67 eV, respectively) and lower than the calculated value for compound 3b (3.85 eV) (Table S2†). This was expected due to the extra \(\pi\)-conjugation afforded by the phenyl substituents in compounds 1b and 2b.

DFT calculations were performed using Gaussian 09 software package,\(^{33}\) for 1b, the two isomers of 2b (2b-cis and 2b-trans, Fig. 3), 3b and naphthalene to calculate their band gaps in the gas phase and in solution phases using B3LYP/6-31++G**. The band gaps of 1b and 2b isomers varied form 4.01 eV to 4.07 eV and were lower in value than 3b and naphthalene (Table 2) similar to the trend observed by the CV experiments. The lower relative band gaps for compounds 1b and 2b can be attributed to the presence of phenyl substituents on each silicon atom, increasing the overall conjugation. The diagrams depicting the HOMO of the same set of compounds (Fig. 6) indicated that the naphthalene also extended the overall conjugation of the system. This extended conjugation, a result of the \(\sigma-\pi\) mixing responsible for the reduction in the band gaps, is well established in the literature.\(^{34}\) Another way to lower the band gap is to increase the number of Si atoms in the chain.\(^{18}\) It was observed that the electronic contribution from the naphthalene rings was significantly more substantial than the phenyl rings attached to the Si centres, and that all the aromatic rings contributed to the \(\sigma-\pi\) mixing therefore explaining the difference in the band gap values between the various bridged silanes (1b and 2b) and 3b. Similar results were obtained for the other naphthalene based bridged systems recently reported by our group.\(^{19}\)

**Table 2** Calculated band gaps of 1b, 2b-cis, 2b-trans and 3b in both gas and solution phases by TD-SCF(DFT).\(^{35}\)

<table>
<thead>
<tr>
<th>Compound</th>
<th>Gas phase</th>
<th>THF</th>
<th>Cyclohexane</th>
</tr>
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<tr>
<td></td>
<td>Band gap (eV)</td>
<td>(\lambda_{\text{max}}) (nm)</td>
<td>Band gap (eV)</td>
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<tr>
<td>1b</td>
<td>4.06</td>
<td>305.0</td>
<td>4.02</td>
</tr>
<tr>
<td>2b-cis</td>
<td>4.07</td>
<td>304.2</td>
<td>4.05</td>
</tr>
<tr>
<td>2b-trans</td>
<td>4.04</td>
<td>306.4</td>
<td>4.01</td>
</tr>
<tr>
<td>3b</td>
<td>4.14</td>
<td>299.0</td>
<td>4.10</td>
</tr>
<tr>
<td>Naphthalene</td>
<td>4.32</td>
<td>286.5</td>
<td>4.29</td>
</tr>
</tbody>
</table>

Fig. 5 Optimized geometries of 1b, 2b(cis/trans) and 3b. (Hydrogen atoms are omitted for clarity.)

Fig. 6 HOMO diagrams from the optimised structures of 1b, 2b(cis/trans) and 3b.
Conclusions

A series of naphthalene bridged disilanes (1b, 2b, 3b) containing different arrangements of methyl and phenyl substituents at silicon were synthesised, characterised and the photophysical properties were investigated. Experimental results and DFT calculations suggested that the Si–Si bonds as well as the presence of aromatic groups, particularly naphthalene, attached to the Si atoms contributed to the lowering of the band gaps. For example, $\lambda_{\text{max}}$ was higher for 1b and 2b than 3b. Furthermore, interchanging the substitution on the silicon atoms in the disilane derivatives resulted in significant changes in the photophysical properties that were observed. The naphthalene disilane derivatives 1b (asymmetric with SiMe$_2$ and SiPh$_2$ substitution) and 3b (no Ph substituents) both absorbed in the UV region and showed weak emission in the near-UV region, similar to naphthalene. In contrast, the symmetric, naphthalene disilane mixture 2b(eis/trans) showed an intense emission peak in polar solvents. On switching to non-polar solvents such as cyclohexane and toluene, broadening in the emission peak in polar solvents. A decrease in the concentration led to a decrease in the broadening in the emission peak in non-polar solvents such as cyclohexane and toluene, broadening in the visible region, corresponding to the presence of excimers, was observed in the emission profile for all three disilane molecules. As expected, concentration studies performed showed that a decrease in the concentration led to a decrease in the broad band at ca. 430 nm. The photoluminescence observed for 2b presents opportunities to use this compound in applications such as light emitting devices or as a photocurrent generator in solar cells. Further derivatising these molecules as well as studying their potential in applications is ongoing in our laboratory.

Experimental

General information

All the experiments and manipulations were performed under nitrogen atmosphere in an MBraun Unilab 1200/780 glovebox or using conventional Schlenk techniques. All glassware was oven-dried overnight in a 135 °C oven. 1,8-Dibromonaphthalene and Wilkinson’s catalyst were purchased from AK Scientific, chlorodimethylsilane and chlorodiphenylsilane were purchased from Sigma-Aldrich, chloromethylphenylsilane was purchased from Gelest. Materials were used directly as received. Ether and toluene were obtained from solvent purification system and later dried over activated molecular sieves (3 Å).

Instrumentation

$^1$H, $^{13}$C and $^{29}$Si($^1$H) DEPT NMR, $^1$H–$^{29}$Si HSQC NMR spectra were recorded on Bruker DPX-400 (400 MHz) spectrometer. Chemical shifts were reported in parts per million (ppm) relative to tetramethylsilane and are referenced to residual protium in the NMR solvent (e.g., CHCl$_3$ = 7.26 ppm) as the internal standard. The silicon NMR resonances were determined with a DEPT pulse sequence. Data are represented as follows: chemical shift ($\delta$), multiplicity (s = singlet, d = doublet, t = triplet, q = quartet, p = quintet, sep = septet, m = multiplet), coupling constants in Hertz (Hz) and integration. High resolution mass spectrometry measurements were made on a Bruker microTOF-QII mass spectrometer, equipped with a KD Scientific syringe pump, in positive ion ESI mode. Hard ionization mass spectrometry analysis was performed on Agilent 7890A GC + 5975C EI-MS with Agilent auto-sampler. X-ray diffraction analysis of single crystals of 1a and 1b was performed on a Rigaku Oxford Diffraction XtaLAB-Synergy-S single crystal diffractometer with a PILATUS 200 K hybrid pixel array detector using Cu K$_\alpha$ radiation ($\lambda = 1.54184$ Å; ESI Table 1†). The data was processed with the SHELXS2018/3 and Olex2 1.3 software packages. All non-hydrogen atoms were refined anisotropically. Hydrogen atoms were inserted at calculated positions and refined with a riding model or located in the difference Fourier map and refined without restrictions (silane hydrogen atoms). Mercury 2020.3.1 was used to visualize the molecular structures. UV-Vis spectra were recorded on a Shimadzu UV-3600 Plus spectrophotometer and the fluorescence emission spectra were recorded on a Jasco FP-8600 spectrofluorometer. Samples were recorded in a macro quartz cuvette (light path = 10 × 10 mm). Fluorescence quantum yields were determined using an Edinburgh Photonics FLS-980 spectrofluorometer equipped with an integrating sphere, using a 450 W ozone free xenon arc lamp as the light source. CV experiments were performed using a BAS CGME controlled mercury stand cell. All the DFT calculations were performed using Gaussian 09. The geometries of all the subject molecules were optimised using the B3LYP functional and 6-311++G** basis set.

Preparation of 1a

1a was synthesized via a literature procedure and the NMR data was in excellent agreement with the published compound.

Preparation of 2a

To a 100 mL Schlenk flask equipped with a stir bar was added 1,8-dibromonaphthalene (2 mmol, 0.571 g) and 15 mL THF solvent. The solution was cooled to −78 °C using a dry ice-acetone bath, to which was added n-BuLi (2.0 M in cyclohexane, 3.0 equiv., 6 mmol, 3 mL) dropwise and the mixture was stirred at −78 °C for 1 hour. To the lithiated mixture was added chloromethylphenylsilane (3 equiv., 6 mmol, 0.9 mL) and was stirred at 22 °C overnight. The reaction mixture was quenched with saturated solution of NaHCO$_3$ (5 mL) and the aqueous layer was extracted with diethyl ether (3 × 10 mL). All the organic layers were combined and dried over Na$_2$SO$_4$. The crude product was obtained upon removal of solvents under vacuum as a yellow gummy substance. The crude was purified over silicagel column, using 100% hexanes as the eluent, to obtain non-separable diastereomers of 2a as a blueish-yellow gum in 64.3% yield.

$^1$H NMR (400 MHz, CDCl$_3$): $\delta$ 7.92–7.26 (m, Ph, 16H), 5.59 (p, SiH, 2H), 0.69 (dd, CH$_3$, 6H). $^{13}$C($^1$H) NMR (100.6 MHz, CDCl$_3$): $\delta$ 144.19, 139.80, 139.68, 139.59, 137.46, 137.42, 137.05, 137.00, 135.88, 133.66, 131.10, 129.85, 129.79, 126.53, −0.08. $^{29}$Si($^1$H) NMR (79.5 MHz, CDCl$_3$): $\delta$ −20.37. EI: Found for C$_{24}$H$_{24}$Si$_2$: 368.19 m/z, calc’d: 368.14 m/z.
Preparation of 3a

To a 50 mL Schlenk flask was added 1a (1 mmol, 0.37 g), Wilkinson’s catalyst (5 mol%, 0.06 mmol, 55 mg) and 10 mL toluene. The mixture was heated at 90 °C for 72 hours. To the mixture was added 15 mL hexane and the mixture was filtered through a Florisil column. The colorless solution was concentrated under high vacuum to remove the solvents and obtain the crude product. The crude was washed with cold pentane (2 × 5 mL) and was stored at −20 °C overnight to obtain colorless crystals of 3a in 30.0% yield.

1H NMR (400 MHz, CDCl3): δ 8.02–7.38 (m, Ph, 16H), 0.56 (s, CH3-6H). 13C{1H} NMR (100.6 MHz, CDCl3): δ 141.86, 134.8, 133.47, 133.11, 131.78, 130.89, 128.62, 126.16, 126.94, 124.90, −3.54. 29Si{1H} NMR (79.5 MHz, CDCl3): δ −20.62. El: found for C24H22Si2: 366.14 m/z, calcd: 366.12 m/z.

Preparation of 3b

To a 50 mL Schlenk flask was added 2a (1.2 mmol, 0.47 g), Wilkinson’s catalyst (5 mol%, 0.05 mmol, 46 mg) and 10 mL toluene. The mixture was heated at 90 °C for 72 hours. To the mixture was added 15 mL hexane and the mixture was filtered through a Florisil column. The colorless solution was dried under high vacuum to obtain the crude product. The crude was washed with cold pentane (2 × 5 mL) to obtain pure 3b as blueish-yellow gum in 24.5% yield.

From the proton NMR spectrum, it was evident that the product was a mixture of cis and trans isomers in 2:1 ratio respectively. Similar results were observed by Kim et al., and based on their findings, it is concluded that the cis isomer was present in higher amounts than the trans isomer.

1H NMR (400 MHz, CDCl3): δ 7.90–7.10 (m, Ph, 24H), 0.75 (s, cis-CH3-6H), 0.60 (s, trans-CH3-3H). 13C{1H} NMR (100.6 MHz, CDCl3): δ 145.70, 140.09, 140.03, 135.47, 133.87, 133.70, 133.02, 132.09, 131.71, 129.83, 128.70, 128.40, 127.87, 126.87, 126.87, 126.61, 125.00, 123.89, −5.71, −6.09. 29Si{1H} NMR (79.5 MHz, CDCl3): δ −22.90, −23.33. El: found for C24H22Si2: 366.12 m/z, calcd: 366.16 m/z.

Preparation of 3a and 3b

3a and 3b were synthesized via a literature procedure and the NMR data was in good agreement with what was reported.

Conflicts of interest

There are no conflicts to declare.

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Notes and references