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PEI functionalized NaCeF₄:Tb³⁺/Eu³⁺ for photoluminescence sensing of heavy metal ions and explosive aromatic nitro compounds†

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This work reports an eco-friendly hydrothermal approach for the synthesis of hexagonal NaCeF₄:Tb³⁺/Eu³⁺ nanophosphors. The phase, morphology and optical properties were characterized by Powder X-ray diffraction (PXRD), field emission scanning electron microscopy (FE-SEM), transmission electron microscopy (TEM), Fourier transform infrared (FT-IR) spectroscopy and photoluminescence (PL) spectroscopy respectively. Herein, the as-synthesized nanophosphor was functionalized with amine rich polyethylenimine (PEI) resulting in development of a luminescent nanoprobe bearing dual sensing functions for hazardous nitroaromatics and heavy metal ions. The strong photoluminescence emission of Eu³⁺ ions was selectively quenched upon addition of toxic analytes at concentrations from 10 to 100 ppm due to complex formation between the analytes and PEI functionalized nanostructure. The synthesized nanomaterial shows sharp emission peaks at 493, 594, 624, 657 and 700 nm. Significantly, the peak at 594 nm shows a noticeable quenching effect on addition of toxic analytes to the aqueous solution of the nanocrystals. The nanophosphors are sensitive and efficient for the PA and Fe³⁺ ion detection with an LOD of 1.32 ppm and 1.39 ppm. The Stern–Volmer (SV) quenching constant (*K*_{SV}) is found to be 2.25 × 10⁵ M⁻¹ for PA and 3.8 × 10⁴ M⁻¹ for Fe³⁺ ions. The high *K*_{SV} value and low LOD suggest high selectivity and sensitivity of the nanosensor towards PA and Fe³⁺ ions over other analytes. Additionally, a reduced graphene oxide and nanophosphor based nanocomposite was also synthesized to investigate the role of energy transfer involving delocalized energy levels of reduced graphene oxide in regulating the luminescence properties of the nanophosphor. It was observed that PEI plays central role in inhibiting the quenching effect of reduced graphene oxide on the nanophosphor.

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Introduction

Lanthanide doped luminescent nanomaterials have gained immense attention due to their potential applications in various fields such as drug delivery,¹ catalysis,² optoelectronics^{3,4} and chemical sensing.⁵ These nanocrystals are widely used for the identification of hazardous heavy metal ions (Fe, Cr, Co, Cd) and explosive aromatic nitro compounds (*p*-nitro aniline, picric acid, *p*-nitrophenol, 4-nitrotoulene).^{6–8} The toxic chemical compounds are heavily spread over surface of soil *via* various anthropogenic activities like machinery and electronic manufacturing, fossil fuel, oil refining and electroplating.^{9,10} The discharge of toxic chemical compounds into water disturb the whole ecosystem and has become a challenge worldwide. The pollution caused by chemical compounds originates from rapid industrial development and urbanization.¹¹ The release of

the hazardous compounds into the environment has increased drastically in underdeveloped countries and may cause massive affect to all living creatures.^{12,13} These chemicals can exert harmful effect on respiratory system, nervous system, connective – epithelial tissue, irritation to the eyes, inflammation in nose and skin.¹⁴ The non-biodegradable nature of many chemical compounds in the environment deteriorates the situation than ever before.¹⁵ A multipronged strategy is needed to tackle these deleterious compounds. Researchers all around the world have employed various analytical techniques for the detection of heavy metal ions and aromatic nitro compounds like inductively coupled plasma mass spectrometry (ICPMS),¹⁶ atomic absorption spectroscopy (AAS),¹⁷ gas chromatography coupled with electron capture detection,¹⁸ ion mobility spectroscopy (IMS),¹⁹ Raman spectroscopy²⁰ and high-pressure liquid chromatography (HPLC).²¹ However, these techniques are expensive and time-consuming since they depend on sophisticated equipment. These factors limit fast and convenient detection.

The metal ion (Fe³⁺) is one of the most significant metal ions that play an important role in biological systems.²² The Fe³⁺ metal ion participates in many biological processes such as RNA

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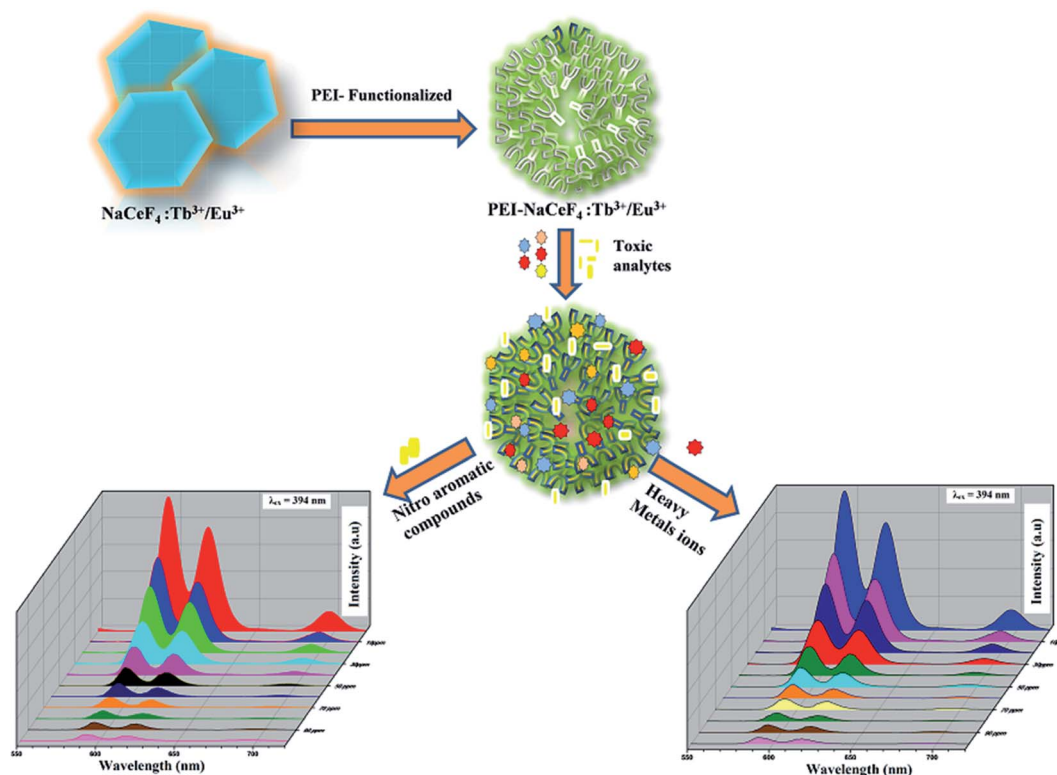
and DNA synthesis and helps in the oxygen transport capacity of haemoglobin. Deficiency and excess of this metal ion can cause serious disorders and diseases in human beings. The deficiency of Fe^{3+} ions can lead to anaemia in the body, whereas its excess can cause life-threatening conditions, such as heart problems, liver and pancreas disease and diabetes.²³ Picric acid (PA) is also an essential chemical source used as an antiseptic agent for the ministration of burns, small pox, and malaria.²⁴ The PA continues to accumulate in the atmosphere as a highly toxic and poorly degradable compound, which poses pernicious influence to both the human (central nervous system and cardiovascular systems) and the environment.²⁵ Therefore, the detection of these toxic analytes becomes crucial for monitoring human health and the environment.

Lanthanide based photoluminescent nanocrystals have been proposed as potential material as new class of luminescent probes.²⁶ They exhibit exceptional chemical and optical properties particularly sharp emission peaks with large Stokes shift and longer lifetimes due to forbidden nature of intra 4f–4f transitions.^{27–29} Moreover, the lanthanide-doped nanostructures shows a distinct set of sharp emission peaks independent of particle size compared to that of single and broad emission peaks observed for quantum dots or other nanoparticles.³⁰ Furthermore, the emission peak position is independent of the excitation wavelength.³¹ The lanthanide-doped host matrix such as fluorides,³² oxides,³³ phosphates³⁴ and oxysulfides³⁵ are thermally and photo-chemically stable.³⁶ The fluorides are appropriate among the hosts compared to other host matrices.

This is due to their low phonon energies that can quickly reduce non-radiative relaxations, resulting in higher photoluminescence quantum yields.^{37,38}

Malik and co-workers developed PEI functionalized $\text{NaYF}_4\text{:Gd}^{3+}/\text{Tb}^{3+}$ for selective detection of 2,4,6-trinitrotoluene *via* photoluminescent sensing technique. The PEI helps in detection of the hazardous nitroaromatic compound by the formation of the Meisenheimer complex and quenching the photoluminescence.³⁹ Tashi *et al.* fabricated down converting serine functionalized $\text{NaGdF}_4\text{:Eu}^{3+}/\text{Ce}^{3+}@/\text{NaGdF}_4\text{:Tb}^{3+}$ core-shell nanomaterial for the selective and sensitive detection of picric acid.⁴⁰ Abbas and co-workers developed C-encapsulated three-dimensional Cu nano-pyramids ($\text{C}@3\text{DCu-PY}$) sensors for the control and monitoring of phenolic compounds.⁴¹ Saleemi *et al.* reported a novel OV's rich spinel ZnFe_2O_4 micro-spheres (ZF-Ms) sensor for the effective detection of As^{3+} metal ion present in underground water.⁴² Abbas *et al.* fabricated a novel Cu nanoplates functionalized carbon nanoparticles ($\text{Cu}@\text{CNP}$) for the selective and sensitive detection of dopamine.⁴³

Herein, we have synthesized PEI functionalized $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ (see Scheme 1) for the selective detection of picric acid and Fe^{3+} metal ion in the aqueous solution. The presence of PEI residues on the surface of synthesized nanocrystals hinders aggregation and increases water solubility of probe nanostructure. Moreover, acid-base reactivity of this luminescent probe facilitates effective binding of analytes on the surface of nanomaterial due to the formation of Meisenheimer complexes between electron-rich amine groups and electron deficient analytes.⁴⁴



Scheme 1 Schematic procedure for the synthesis of $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) and PEI functionalized $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) nanophosphors.



Additionally, to unravel the effect of graphene derivatives on the photoluminescence properties of the nanophosphor, R-GO@PEI-NaCeF₄:Tb³⁺/Eu³⁺ nanocomposite was synthesized with and without the functionalizing material. The results showed a remarkable improvement in the PL intensity resulting in the red emission of nanocomposite with the addition of the amine rich PEI ligand rather than without the PEI attached nanocomposite.

Experimental section

Chemicals and materials

All the chemicals used in the work including terbium(III) nitrate hexahydrate Tb(NO₃)₃·6H₂O (99.9%), cerium(III) nitrate hexahydrate Ce(NO₃)₃·6H₂O (99.5%), europium(III) nitrate hexahydrate Eu(NO₃)₃·6H₂O (99.7%), ammonium tetrafluoroborate NH₄BF₄ and ethylenediaminetetraacetic acid (EDTA) were purchased from Alfa Aesar. Nitric acid (HNO₃), phosphoric acid (H₃PO₄), potassium permanganate KMnO₄ and ethanol all were of analytical grade. Graphite flakes and polyethylenimine (PEI) were procured from Sigma Aldrich. All the chemicals were used without further purification. Triple distilled water was used for experiments.

Synthesis of NaCeF₄:Tb³⁺ nanophosphors

In a usual procedure, a series of Tb doped NaCeF₄ phosphors were fabricated *via* modified hydrothermal method at varying temperature conditions. In a typical synthesis method Tb(NO₃)₃·6H₂O (0.05 mmol, 0.0217 g), Ce(NO₃)₃·6H₂O (0.95 mmol, 0.4125 g) were dissolved in 10 mL of distilled water and further stirred for 10 minutes for proper mixing of both the metal ions. Separately, 0.37 g (1 mmol) of Na₂EDTA was added to 5 mL of distilled water and stirred for 15 minutes. After continuous stirring, the two solutions were mixed together and agitated for half an hour at room temperature. To the resultant reaction mixture, 5 mL aqueous solution of NaOH was drop wise added under continuous stirring followed by addition of 10 mL of ethanol. Lastly, 2 mL aqueous solution of NH₄BF₄ (6 mmol, 0.629 g) was drop wise added into the reaction solution followed by vigorous stirring for 1 hour at room temperature. The resultant precursor solution was finally transferred into a 50 mL Teflon-lined autoclave and heated at 120 °C for 24 h. The reaction vessel was naturally cooled down at room temperature. The obtained product was separated by centrifugation and repeatedly washed with distilled water followed by 10% ethanol. The product obtained was finally dried at 70–80 °C for 12 h. Three more samples with the same chemical composition were synthesized under similar experimental conditions, at varying temperatures (140 °C, 160 °C, and 180 °C) for 24 h.

Synthesis of NaCeF₄:Tb³⁺ co-doped Eu³⁺

In a similar procedure, NaCeF₄:Tb³⁺ co-doped Eu³⁺ nanostructures were synthesized with 5% Tb³⁺ doping and with varying Eu³⁺ ion concentrations (1%, 3%, 5%, 7%). Na₂EDTA was chosen as surfactant or surface directing agent for better morphology of the nanoparticles. Tb(NO₃)₃·6H₂O (0.05 mmol, 0.0217 g), Ce(NO₃)₃·6H₂O (0.94 mmol, 0.408 g) and Eu(NO₃)₃·6H₂O (0.01 mmol, 0.004 g) were dispersed in 7 mL of distilled

water and stirred for about 10 minutes until complete dissolution. 0.37 g (1 mmol) of Na₂EDTA was added to 7 mL of distilled water and stirred for further 10 minutes. The two solutions were then mixed together and continuously stirred for extra 30 minutes. Additionally, 5 mL aqueous solution of NaOH was added drop wise followed by 10 mL of ethanol and 2 mL solution of NH₄BF₄ (6 mmol, 0.629 g) under vigorous stirring for 1.5 h at room temperature. The resultant reaction mixture was then transferred into a 50 mL Teflon-lined autoclave and the reaction vessel was heated at 160 °C for 24 h. The reaction vessel was cooled down naturally and the product so obtained was centrifuged and washed three times with distilled water followed by ethanol. The product was finally dried at 60 °C for 24 h.

Under similar experimental conditions, three more samples were synthesized at same temperature conditions like (160 °C at 24 h) but different Eu ion concentrations (3%, 5% and 7%).

PEI-surface functionalization of nanophosphor

In a typical hydrothermal method, 0.01 g of as synthesized NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) sample was dispersed in a 10 mL of water and sonicated for 30 minutes. Furthermore, 0.02 g of PEI was separately mixed in 10 mL of water and sonicated for about 15 minutes until proper mixing. Finally, dropwise PEI solution was added to the above sonicated solution with further addition of 5 mL ethanol followed by vigorous stirring for about 1 h. The surface-functionalized nanomaterial PEI @NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) so obtained was separated by centrifugation, washed twice with distilled water followed by ethanol washing and finally vacuum dried at room temperature for 24 h.

Synthesis of graphene oxide (GO)

Graphene oxide (GO) was prepared by one-pot synthesis gleaned from new improved Hummers method.⁴⁵ Natural graphite flakes were used as the precursor material and KMnO₄ as the oxidising agent for the preparation of GO. In a typical procedure, 1 g of pure graphite flakes were dissolved in a solution containing stoichiometric ratio 9 : 1 of sulphuric acid and phosphoric acid and vigorously stirred for about 6 h. The suspension mixture was then kept in an ice bath and the temperature was adjusted below 5 °C. Next, 8 g of KMnO₄ was added slowly into the mixture while stirring followed by stirring for 1.5 h. The resulting green color suspension was further stirred in a water bath (40 °C) for 2 h. The mixture was then adjusted to a linearly high temperature of 70 °C and stirred for about 45 minutes. Later, distilled water was added to the brown color suspension mixture and then transferred into beaker containing 400 mL of additional water. In order to terminate the oxidation process 15 mL of H₂O₂ was added and mixed vigorously for 10 minutes imparting yellow color. The product obtained was finally separated by centrifugation and then repeatedly decant off with distilled water followed by 10% HCl solution. Lastly, the solid product was dried under vacuum at room temperature and product was obtained as GO powder. Additionally, the reduced graphene oxide (R-GO) was synthesized by simple hydrothermal method using hydrazine hydrate and ammonia as the reducing agent as described in the literature.⁴⁶



Synthesis of R-GO@PEI@NaCeF₄:Tb³⁺/Eu³⁺

A facile hydrothermal route was used for the synthesis of nanocomposite material. 0.02 g of R-GO was dispersed in 15 mL distilled water and sonicated for 40 minutes. Separately, 0.08 g of as-synthesized PEI-functionalized NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) nanomaterial was dispersed in 30 mL of distilled water and sonicated for 20 minutes. The two solutions were mixed together followed by addition of 7 mL ethanol and stirred for 1 h. Finally, the suspension mixture was centrifuged and washed twice with distilled water followed by 10% HCl solution. The solid product was vacuum dried at room temperature for 24 h.

Spectroscopic and microscopic measurement

X-Ray diffraction (XRD) data was used to determine the phase purity and crystallite size of as-synthesized nanophosphors. The XRD data was collected using D8 X-ray diffractometer (Bruker) at scanning rate of 12 min⁻¹ in the 2θ range of 10 to 80° with Cu Kα radiation (λ = 0.15405 nm). Field emission scanning electron micrographs (FE-SEM) were recorded on FEI Nova Nano SEM-450 to examine the external morphology and particles size of as-synthesized nanophosphors. The internal composition, lattice parameters, grain size and morphology was recorded by using transmission electron microscopy (TEM) and high-resolution TEM (HR-TEM). The high-resolution TEM (HR-TEM) was recorded on Tecnai G2 S-Twin transmission electron microscope having field emission gun operating at 200 kV. The presence of surface directing agent as well as the functionalized materials which was assimilated on the surface of nanocrystals was explored by Shimadzu Fourier Transform Infrared Spectrophotometer (FT-IR) using standard KBr pellet technique (4000–400 cm⁻¹). For UV-visible down conversion photoluminescence emission spectra of as-synthesized nanophosphors Hitachi F-4700 Fluorescence Spectrophotometer having xenon lamp as excitation source was employed.

Sensing of hazardous nitroaromatics and heavy metal ions

The Photoluminescence detection experiment for the sensing of heavy metal ions and explosive aromatic nitro compounds were carried out on Hitachi F-4700 Fluorescence Spectrophotometer at room temperature. PEI-functionalized NaCeF₄:Tb³⁺/Eu³⁺ nanostructure (0.03 g) was dispersed in 30 mL of distilled water and sonicated for around 20 minutes. Six stock solutions of explosives were prepared by dissolving 0.1 g of each of nitroaromatic compounds as well as heavy metal ions in 100 mL of water (1000 ppm solution). From the stock solutions, ten different solutions of each compound were prepared by simple dilution having concentrations ranging from 10–100 ppm. To each dilute solution, 2 mL of as prepared sample suspension was added. The fluorescence spectrum of each solution was recorded immediately. Herein, for sensing of nitro-compounds as well as heavy metals ions the luminescence spectra was analyzed by exciting the PEI-functionalized NaCeF₄:Tb³⁺/Eu³⁺ luminescent material selectively at 394 nm for the solution phase at variant ppm concentrations.

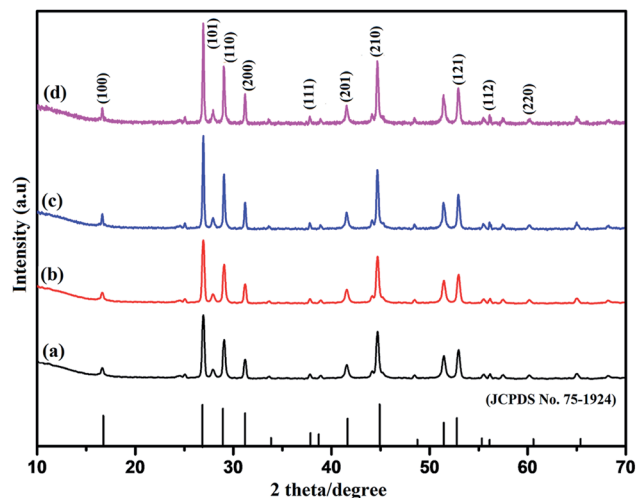


Fig. 1 PXR patterns of as-synthesized NaCeF₄:Tb³⁺/Eu³⁺ (x%) nanoparticles with different Eu³⁺ ion concentrations: (where x = (a) 1% (b) 3% (c) 5% (d) 7%).

Result and discussion

Powder X-ray diffraction

The composition, phase evolution, purity, and size of the as-prepared sodium cerium fluoride core doped with rare earth metals ions were confirmed by PXR technique. A series of NaCeF₄:Tb³⁺ nanocrystals samples were synthesized keeping the reaction time and concentration of the metals ions as well as the concentration of the surfactant (EDTA) constant while varying the reaction temperature. Fig S1: ESI† shows the diffraction pattern of synthesized NaCeF₄:Tb³⁺ nanophosphors at different heating temperatures. All the diffraction peaks are in good agreement with the JCPDS no. 75-1924.⁴⁷ Furthermore, to explore the effect of doping ions concentration on the structure of nanophosphors, another series of NaCeF₄:Tb³⁺/Eu³⁺ nanocrystals with different Eu³⁺ ion concentration were

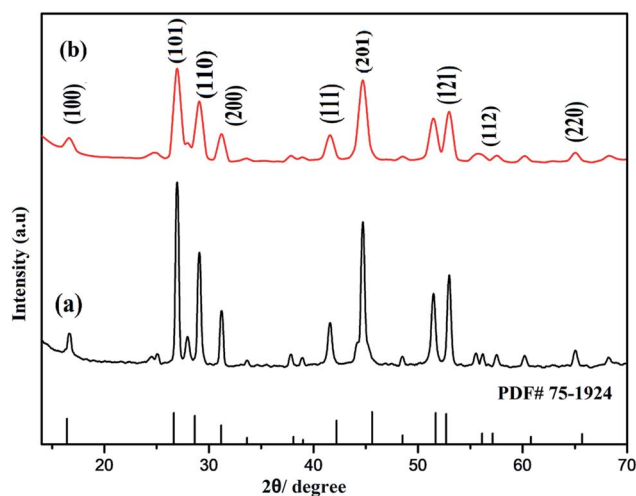


Fig. 2 Comparison of PXR patterns of prepared nanostructures: (a) NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) (b) PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%).



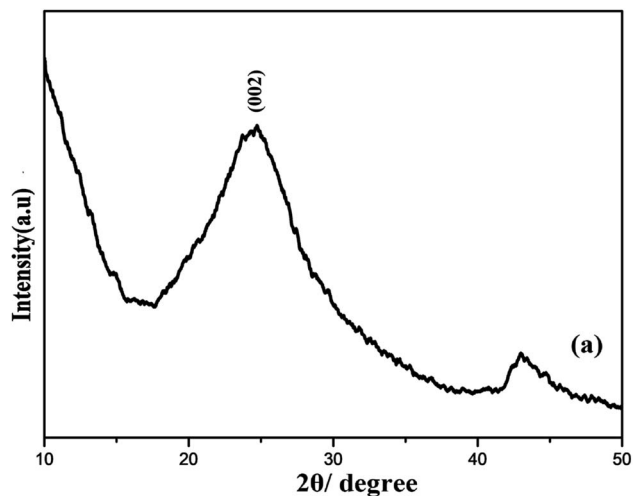


Fig. 3 PXRD pattern of (a) R-GO.

prepared keeping the reaction time, temperature and molar concentration of surfactant constant. Fig. 1 shows the diffraction pattern of the as-synthesized nanophosphors. It was observed that all the diffraction peaks of nanocrystals are in accordance with the JCPDS no. 75-1924 and are perfectly indexed to pure hexagonal phase. From the XRD pattern (Fig. 1), the Eu^{3+} concentrations do not cause any difference in the crystal phase and structure of nanophosphors. All the diffraction peaks can be readily indexed as the hexagonal NaCeF_4 , agreeing well with the data reported in the JCPDS no. 75-1924.⁴⁸ However, only small variation in particles size can be seen. The powder XRD patterns of as-synthesized nanocrystals were

systematically investigated and no extra impurity peak was observed clearly indicating that Tb^{3+} and Eu^{3+} ions were perfectly dissolved in the NaCeF_4 host lattice without further alteration in the crystal structure.

It was also observed that no additional peak was noticed in the spectra of nanophosphors with amine rich ligand (poly-ethylenimine (PEI)). However the diffraction pattern gets broadened which signify the nanoscale range of as-synthesized nanocrystals as shown in Fig. 2.

Scherrer's eqn (1) was applied to calculate the crystallite sizes of the as-synthesized nanomaterial:⁴⁹

$$L = \frac{K\lambda}{\beta \cos \theta} \quad (1)$$

Here, L (nm) is the crystallite size, K is the Scherrer constant (equals to 0.89), λ (nm) is the wavelength of the Cu $K\alpha$ radiant, $\lambda = 0.15405$ nm, 2θ is the diffraction angle, β is the full-width at half-maximum (FWHM) of the diffraction peak. The average crystallite size was calculated by using the major diffraction peaks of nanocrystals. The average crystallite size of $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ for nanostructures were observed in the range of 15 to 21 nm.

Furthermore, the XRD pattern of R-GO was systematically examined. A strong characteristics peak centered at $2\theta = 23^\circ$ corresponding to (002) plane and additional small peak around $2\theta = 45^\circ$ were observed as shown the Fig. 3. The presence of wide peak width clearly demonstrates the small crystallite size and short layered stacked structure of R-GO.⁵⁰

Energy dispersive X-ray spectroscopic analysis

The energy dispersive X-ray spectroscopy (EDS) analysis was used to determine the exact elemental content in the prepared

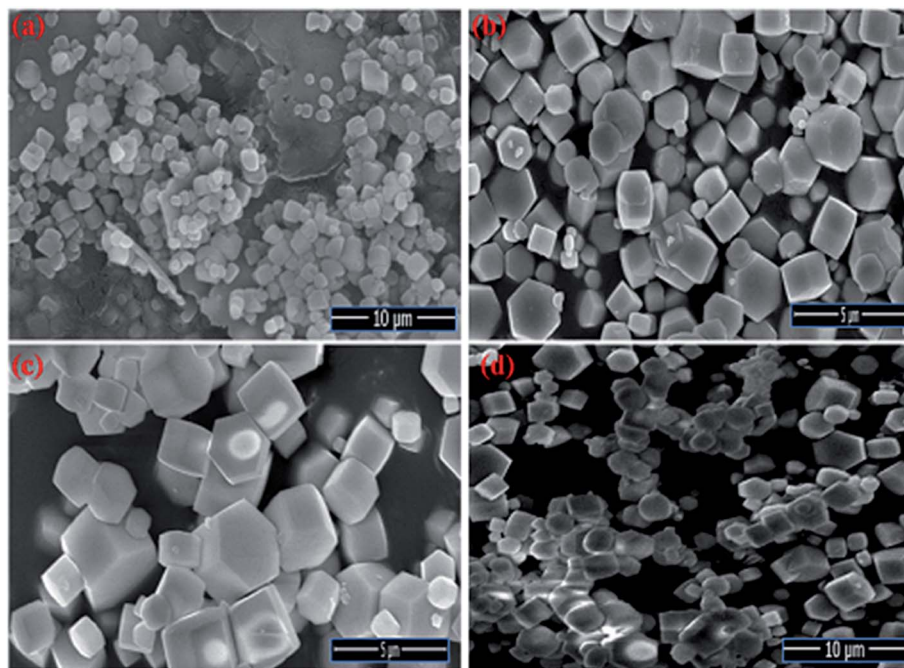


Fig. 4 FESEM micrographs depicting the variation in the morphology of nanophosphors from cubic to hexagonal $\text{NaCeF}_4:\text{Tb}^{3+}$ at different temperatures (a) 120 °C (b) 140 °C (c) 160 °C (d) 180 °C.



samples. The result obtained is tabulated in Tables S1 and S2: ESI.† The spatial distribution of elements involves Na, Ce, Tb, Eu, and F in the as-prepared $\text{NaCeF}_4\text{:Tb}^{3+}$ and $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ nanophosphors. Furthermore, the presence of additional elements like N, O and C in the EDS spectrum validates the formation of PEI-functionalized nanoparticles. The presence of crucial element like C in the R-GO and R-GO@PEI@ $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ nanocomposite samples confirms the formation of carbon based R-GO and R-GO doped nanocomposite. The spectral peaks of samples computing the presence of Y, Ce, Gd, Eu, Tb, Na, C, N, O and F elements are shown in Fig. S2–S4: ESI.†

Scanning electron microscopy and high-resolution transmission electron microscopy

The Field emission Scanning Electron Microscopy (FE-SEM) Transmission Electron Microscopy (TEM) and high-resolution Transmission Electron Microscopy (HR-TEM) analysis was utilized to monitor the shape, size and morphology of nanomaterials. Fig. 4 displays the FE-SEM micrographs of synthesized $\text{NaCeF}_4\text{:Tb}^{3+}$ at varying reaction parameters. The morphology of the nanophosphors was systematically tuned by heating temperature from 120 °C to 180 °C. During the growth mechanism, $\text{NaCeF}_4\text{:Tb}^{3+}$ exhibited two different crystal morphologies and phase transformation occurred from kinetically stable cubic shape (α -phase) at 120 °C to thermodynamically stable hexagonal shape (β -phase) at around 160 °C.⁵¹ It was observed from the micrographs that the 160 °C is a suitable temperature resulting in smooth-surfaced pure hexagonal-shaped nanoparticles. To further monitor the effect of Eu^{3+} ions, as-synthesized $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ aliquots were analyzed

(with varying Eu^{3+} contents from 1 to 7%) and results are shown in Fig. 5. The Eu^{3+} ion concentration caused tunable effect on the morphology and size of the resultant nanoparticles architecture, ranging from smooth-surfaced 2D hexagonal plate-like morphology in the early stage (1–3%) to aggregated broad sized hexagonal rods at the later stage of the reaction bearing a higher concentration of Eu^{3+} (7%). The overall result showed that Eu^{3+} ion exhibits dual-mode character, showing active luminescent property and also notable morphological agent for the synthesized nanostructure reaction. Hence, the reaction temperatures as well as metal ion concentration emerge as important factors in adapting an improved morphology and shape of nanomaterial. Fig. S5: ESI† displays the TEM micrographs of $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ showing hexagonal shape and PEI- $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ confirming the presence of a aggregated layer-like structures on to the surface of PEI capped hexagonal nanophosphor.⁵²

Furthermore, the FE-SEM and TEM micrographs of composites of R-GO with nanophosphor were also analyzed to explore the morphological effect of sheets like R-GO on the luminescent nanophosphor as shown in Fig. 6 and 7. The result noticed the transparent crumbled sheet like the layered structure of R-GO. Fig. 6(b) and 7(b) display the FE-SEM and TEM micrographs of the synthesized R-GO@PEI- $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ nanocomposite depicting the exquisite doping of hexagonal-shaped PEI- $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ nanoparticles onto the R-GO sheets.

The high-resolution transmission electron microscopy (HR-TEM) images confirm the high crystallinity of the prepared nanophosphors (see Fig. S6 and S7: ESI†). From the micrographs (shown in Fig. 8), it is indicated that the lattice fringe spacing between interlayer planes lies in the range of 0.27 nm to 0.3 nm, which is well indexed with the (101) lattice plane of the hexagonal

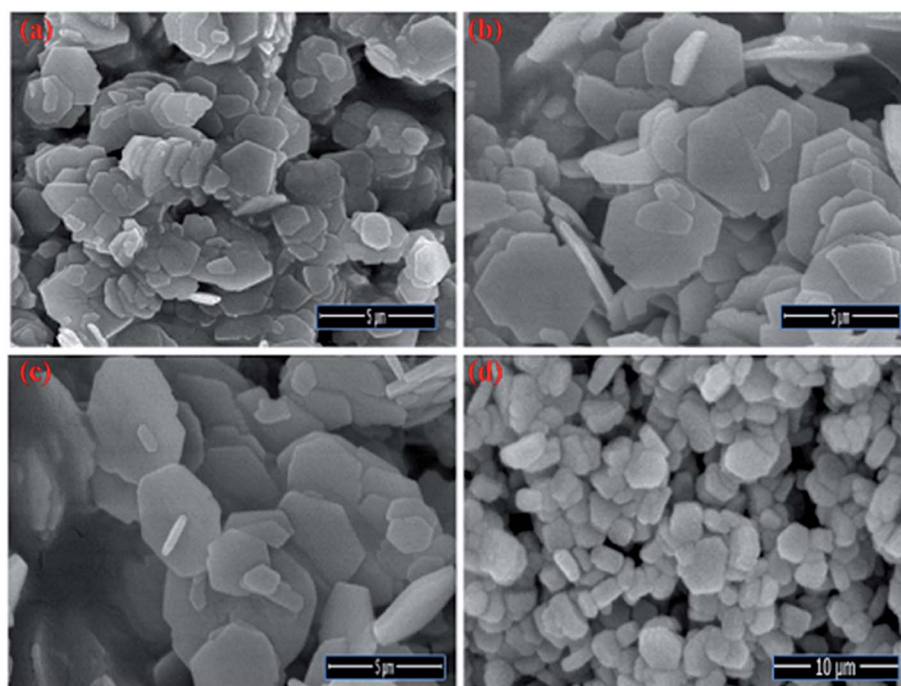


Fig. 5 FE-SEM images visualizing the morphological effect of Eu^{3+} ion contents on the hexagonal $\text{NaCeF}_4\text{:Tb}^{3+}/\text{Eu}^{3+}$ nanophosphors (a) 1% (b) 3% (c) 5% (d) 7%.



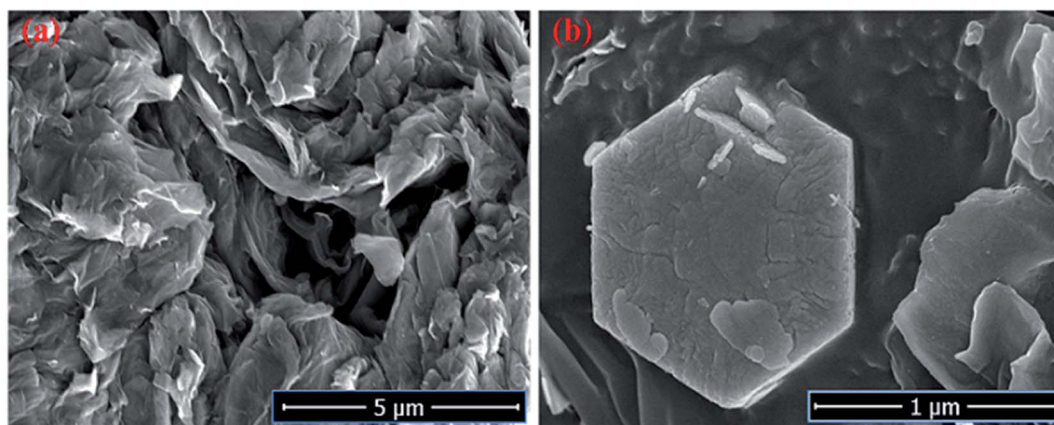


Fig. 6 FE-SEM micrographs of (a) R-GO (b) RGO@PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) nanocomposite.

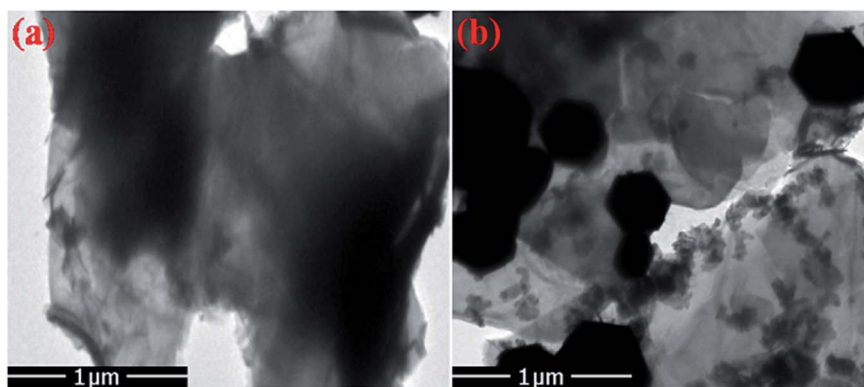


Fig. 7 TEM images of (a) R-GO (b) RGO@PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) nanocomposite showing the presence of hexagonal nanoparticles anchored on the R-GO sheets.

NaCeF₄ structure.⁵³ The existence of ordered lattice fringes corresponding to the hexagonal phase is in good agreement with the results of the X-ray diffraction (XRD) patterns. The nanophosphors result in preferential growth orientation along the 002 direction.

Infrared spectroscopy

FT-IR spectroscopy was used to explore the presence of functional groups in the synthesized NaCeF₄:Tb³⁺/Eu³⁺ nanocrystals

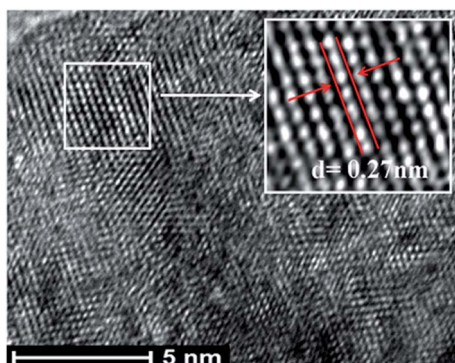


Fig. 8 HR-TEM micrograph showing PEI-functionalized NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) nanostructure.

as well as for the PEI functionalized NaCeF₄:Tb³⁺/Eu³⁺ nanophosphors as shown in Fig. S8; ESI.† The FT-IR spectra were selectively recorded within range of the wave number 4000–

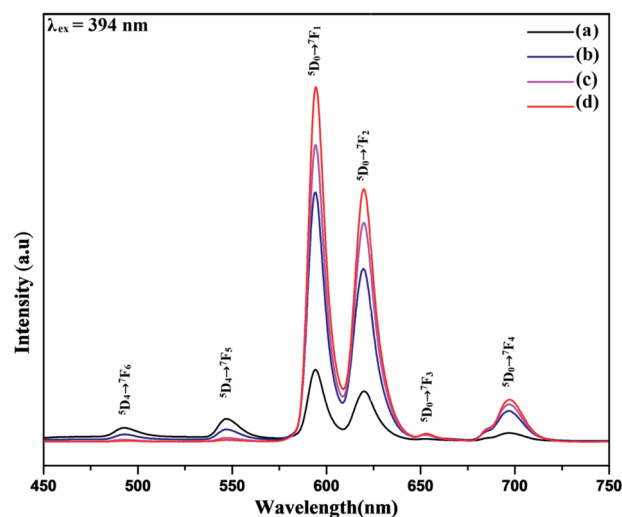


Fig. 9 Room-temperature photoluminescent emission spectra of NaCeF₄:Tb³⁺/Eu³⁺ displaying different emissions at varying Eu³⁺ concentrations (a) 1% (b) 3% (c) 5% (d) 7%.



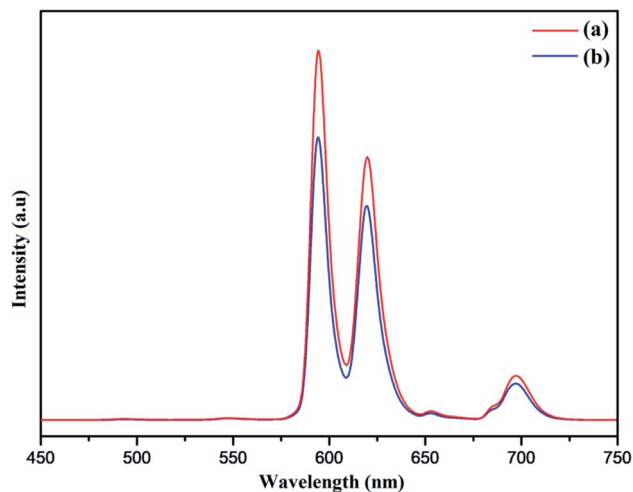


Fig. 10 Room-temperature photoluminescent emission spectra of (a) PEI-NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) (b) NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%).

400 cm⁻¹. The IR spectra revealed intense broad absorption band at 3400 cm⁻¹ assigning to stretching vibration of the adsorbed -OH molecules on the surface of NaCeF₄:Tb³⁺/Eu³⁺ nanophosphors. The bands at 2924 and 2850 cm⁻¹ correspond to asymmetrical and symmetrical stretching vibration

modes of the methylene group (-CH₂) respectively. The bands around 1640 and 1400 cm⁻¹ are attributed to asymmetrical and symmetrical stretching of the carboxylate (-COO) group.⁵⁴ The appearance of these IR bands confirms the presence of EDTA in the NaCeF₄:Tb³⁺ and NaCeF₄:Tb³⁺/Eu³⁺ nanocrystals.

Meanwhile, to monitor the proper functionalization of PEI onto the surface of nanoparticles the FT-IR spectrum of the prepared sample was recorded and compared with that of pure NaCeF₄:Tb³⁺/Eu³⁺ nanocrystals fabricated without PEI Fig. S9: ESI.† The existence of PEI is confirmed by the presence of the characteristic absorption peaks at 3389 cm⁻¹ corresponds to N-H stretching vibration of the NH₂ group. The band around 1380–1630 cm⁻¹ is due to vibration of the amide group.⁵⁵ For the additional bonding with any external ligand, the presence of free amine groups on the surface of the nanoparticles is of extreme importance as these can bind with the outer molecules. The band at 1527 cm⁻¹ (Fig. S9: ESI†) corresponding to amine groups confirms the presence of free amine groups still vacant on the surface nanoparticles for further attachment.⁵⁶ An absorption band around 1391 cm⁻¹ is attributed to stretching vibrations of the C-N bond of amino acid (see Fig. S9: ESI†). Meanwhile, it was interesting to note that all the above bands were not observed in the pure EDTA capped

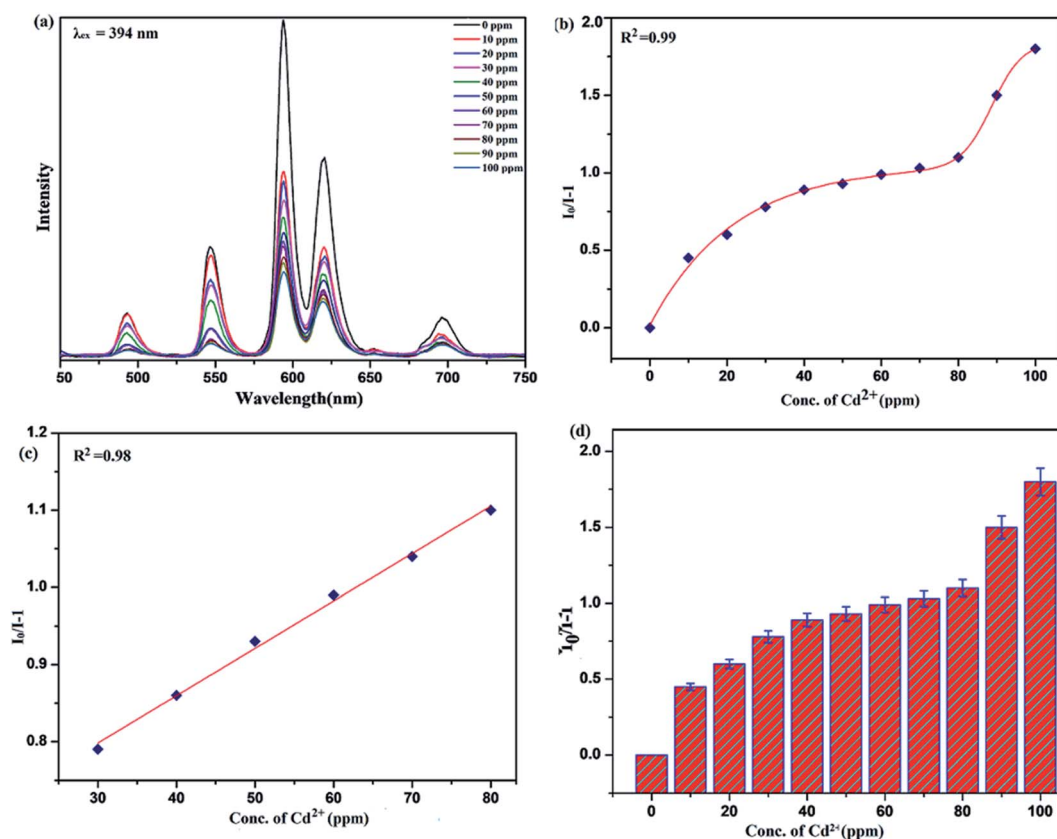


Fig. 11 Variation in photoluminescence intensity of nanosensor in presence of different Cd²⁺ concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394$ nm: (a) emission spectra of nanophosphors with addition of Cd²⁺ ion (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (Cd²⁺) concentrations (c) linear Stern–Volmer fitting and (d) error bar.



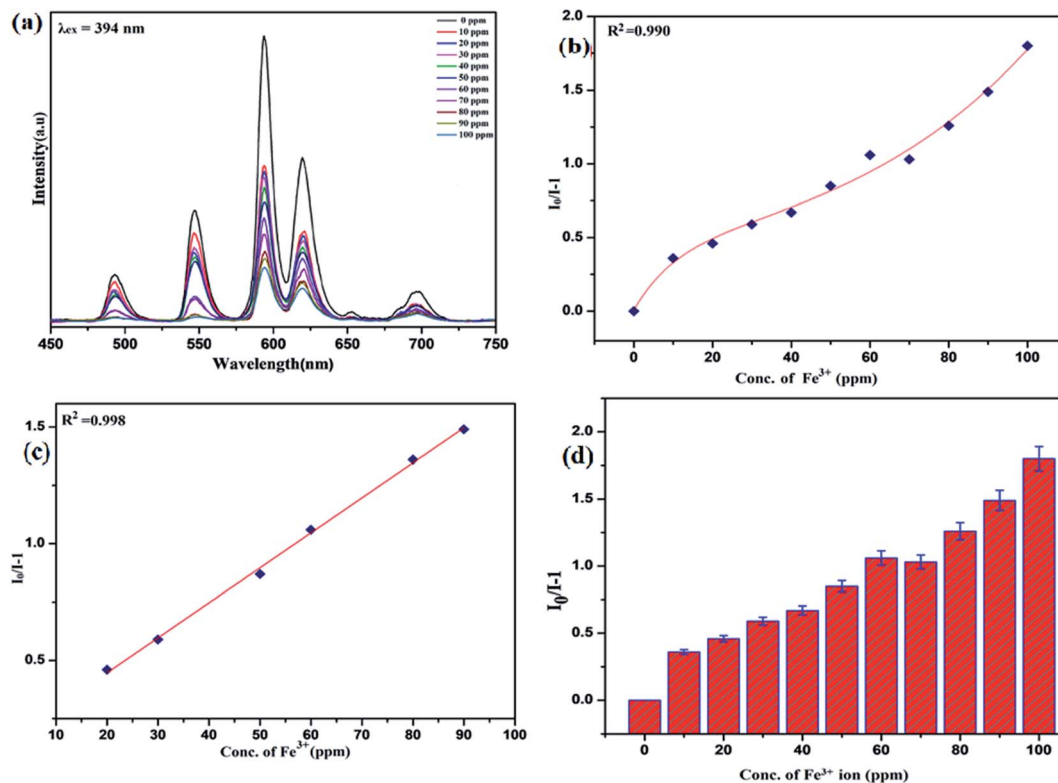


Fig. 12 Variation in photoluminescence intensity of nanosensor in presence of different Fe^{3+} concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394 \text{ nm}$: (a) emission spectra of nanophosphors with addition of Fe^{3+} ion (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (Fe^{3+}) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

nanocrystals confirming the successful functionalization of PEI on to the nanophosphor.

Additionally, FT-IR spectrum was further recorded to investigate the presence of functional groups and bonding interaction of the GO, R-GO and the $\text{PEI-NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}@\text{R-GO}$ nanocomposite as depicted in Fig. S10: ESI†. The occurrence of characteristic peak at 3430 cm^{-1} is attributed to the stretching vibration of hydroxyl groups ($-\text{OH}$) present on the GO sheets. The IR spectrum of GO demonstrates the presence of abundant oxygen containing functional moieties including $\text{C}=\text{O}$ functional groups (band around 1734 cm^{-1}) and epoxy groups (asymmetric vibration at 1232 cm^{-1}).⁵⁷ There occurs a drastic change in the IR spectrum of R-GO where shift in bands is observed from 1628 cm^{-1} to around 1586 cm^{-1} due to the reduction process.⁵⁸ Fig. S11: ESI† displays the FT-IR spectrum of the fabricated $\text{PEI-NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}@\text{R-GO}$ nanocomposite confirming the successfully binding interaction of various functional moieties remaining on the R-GO sheets with the nanophosphor. The shift in the band positions confirms the sufficient anchoring of metal nanophosphor onto the R-GO layered structure.

Photoluminescent properties

The photoluminescence spectra of the synthesized Tb^{3+} doped NaCeF_4 nanostructures was monitored under single excitation wavelength 247 nm as shown in Fig. S12: ESI†. The emission

peaks centred at 490, 545, 585 and 621 nm are corresponding to various transitions resulting from excited state $^5\text{D}_4$ to ground state $^7\text{F}_j$ ($J = 6, 5, 4, 3$) respectively. The presence of strong peak at 545 nm *i.e.*, $^5\text{D}_4 \rightarrow ^7\text{F}_5$ originates due to the magnetic-dipole and electric-dipole allowed transition.⁵⁹ The emission from Tb^{3+} is sensitized by Ce^{3+} .

The photoluminescence spectra of $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphors with varying Eu^{3+} concentrations (1%, 3%, 5% and 7%) were also recorded to investigate the luminescence behaviour and effect of different Eu^{3+} ion concentrations on the PL intensity as displayed in Fig. 9. The Ce^{3+} ion (having 4f–5d allowed transition) can also sensitize Eu^{3+} ion which can act as an activator and improve the luminescence competency of the nanophosphors.⁶⁰ However, direct sensitization of Eu^{3+} from Ce^{3+} is prohibited due to metal to metal electron charge transfer (MMCT) *i.e.*, $\text{Ce}^{3+} + \text{Eu}^{3+} \rightarrow \text{Ce}^{4+} + \text{Eu}^{2+}$ resulting in abrupt fall off photoluminescence intensity.⁶¹ Tb^{3+} ion (green-emitter) can be employed as bridging material for improved energy transfer from the Ce^{3+} ion to Eu^{3+} ion with minimum quenching effect. The results observed indicate effective energy transfer with superior luminescence emission.⁶² Thus, apart from being an efficient activator, Tb^{3+} can also be employed as a competent mediator between the sensitizer (Ce^{3+}) and activator (Eu^{3+}) resulting in minimizing the MMCT quenching.⁶³ Consequently, there occur well-built green orange emissions under ultraviolet excitation occurring due to energy transfer (ET) from $\text{Ce}^{3+} \rightarrow \text{Tb}^{3+}$ and $\text{Tb}^{3+} \rightarrow \text{Eu}^{3+}$.

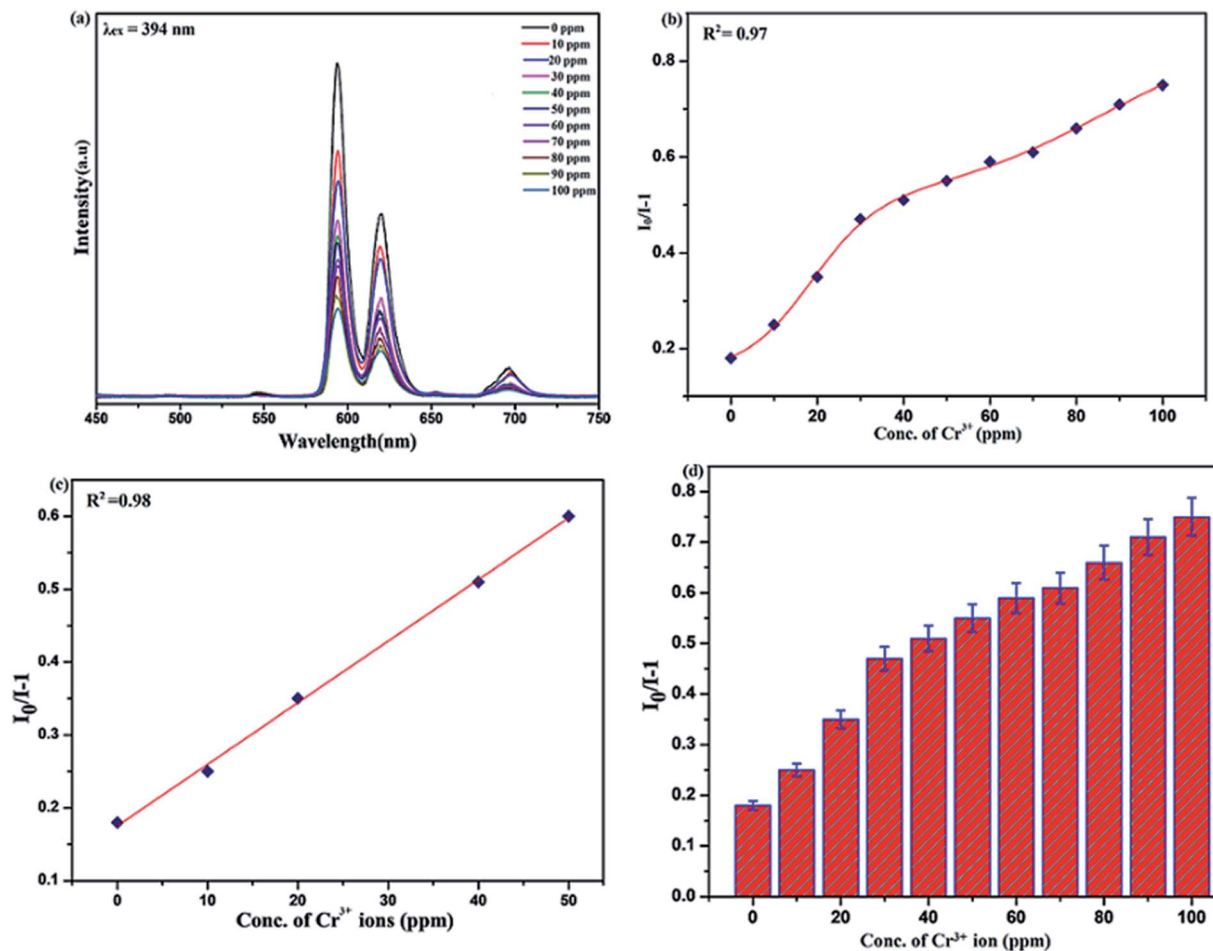


Fig. 13 Variation in photoluminescence intensity of nanosensor in presence of different Cr³⁺ concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{ex} = 394$ nm (a) emission spectra of nanophosphors with addition of Cr³⁺ ion (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (Cr³⁺) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

The Photoluminescence spectra of the NaCeF₄:Tb³⁺/Eu³⁺ with different Eu³⁺ concentrations displayed in Fig. 9 reveal the existence of four characteristics emission peaks (under 394 nm excitation) arising from different transitions centered at 594 (⁵D₀ → ⁷F₁), 624 (⁵D₀ → ⁷F₂), 657 (⁵D₀ → ⁷F₃), 700 (⁵D₀ → ⁷F₄). The peak centered at 657 nm (⁵D₀ → ⁷F₃) and 700 nm (⁵D₀ → ⁷F₄) may be due to magnetic and electric-dipole forbidden transitions of NaCeF₄:Tb³⁺/Eu³⁺ nanophosphors.⁶⁴ The PL intensity varies with the change in Eu³⁺ ion concentrations. The PL spectrum (Fig. 9) shows an enhancement in the luminescence intensity with the increase in the Eu³⁺ concentrations ranging from 1% to 7% as displayed. From the graph it is observed that sample having Eu³⁺ ion 7% content achieved maximum luminescence intensity showing efficient energy transfer to the Eu³⁺ ion. It is quite interesting to note that the peaks occurring at 493 and 548 nm correspond to the characteristic emission peaks of Tb³⁺. The decrease in the Tb³⁺ peaks with increase in the Eu³⁺ concentration manifest the effective energy transfer from Tb³⁺ to Eu³⁺ ions as shown in Fig. 9.

Furthermore, the PL spectrum of the PEI functionalized nanophosphor was also investigated showing improved photoluminescence intensity of the nanophosphor as depicted in Fig. 10. The PL emission spectrum shows an enhancement in the emission intensity as compared to the NaCeF₄:Tb³⁺/Eu³⁺ nanophosphor. In order to investigate the energy transfer process in the nanophosphor the life time decay curve was measured. The lifetime of NaCeF₄:Tb³⁺ and NaCeF₄:Tb³⁺/Eu³⁺ nanocrystals were recorded by taking into consideration the excitation and emission wavelength at 394 and 610 nm respectively (see Fig. S13 and S14: ESI†). The luminescence decay curves NaCeF₄:Tb³⁺/Eu³⁺ are in good agreement with the single exponential function $I = I_0 + A \exp(t/\tau)$ (where I and I_0 are luminescence intensities at the time ' t ' and ' 0 ', the term τ stands for photoluminescence decay time). The lifetimes of NaCeF₄:Tb³⁺ measured are in the range of 125 ms to 128 ms while the lifetimes of the Eu³⁺ doped NaCeF₄:Tb³⁺ (at different Eu³⁺ concentrations) and PEI functionalized NaCeF₄:Tb³⁺/Eu³⁺ varies from 131 ms to 138 ms. The NaCeF₄:Tb³⁺/Eu³⁺ exhibit much longer lifetime as compared to the Tb³⁺ doped nanophosphors suggesting an



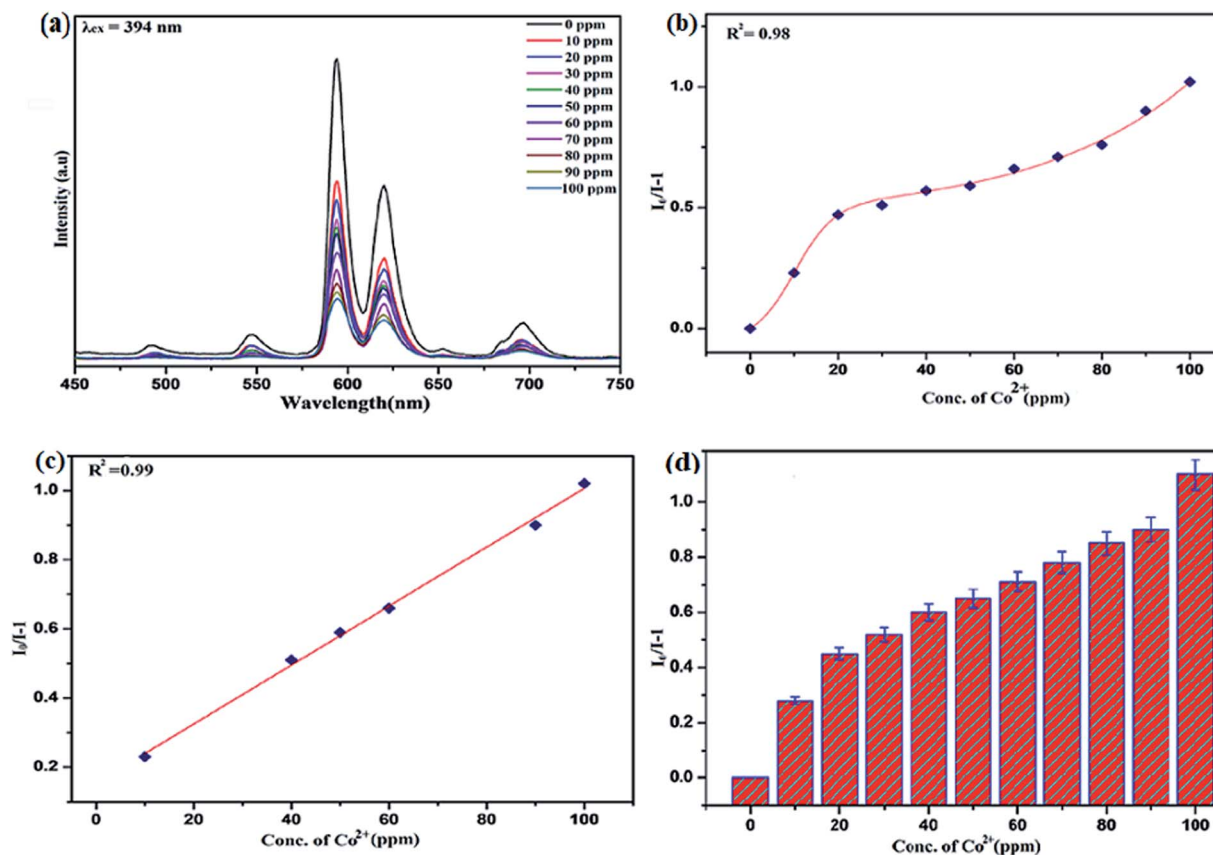


Fig. 14 Variation in photoluminescence intensity of nanosensor in presence of different Co^{2+} concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394 \text{ nm}$ (a) emission spectra of nanophosphors with addition of Co^{2+} ion (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (Co^{2+}) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

improved energy transfer process between the adjacent metals ions.

Luminescent behavior and sensing properties

Lanthanide-doped nanocrystals due to their unique luminescence properties (such as their excellent chemical stability, narrow bandwidth emission, photo stability, tunable emission color) turn out promising candidates for luminescent based materials.⁶⁵ The photoluminescent spectra of synthesized PEI functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphor was explored for the fluorescence detection of toxic metal ions and hazardous nitro-aromatics compounds. It has been systematically demonstrated that PEI-functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphor can easily and efficiently sense the aromatic nitro compounds as well as toxic heavy metal ions even at trace level. All the photo luminescent spectra were recorded at room temperature using water as the dispersion medium. As shown in Fig. 10, PEI capped $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphor displays a sharp emission peak centered at 594 and 624 nm under $\lambda_{\text{ex}} = 394 \text{ nm}$ when dispersed in water.

Sensing of toxic heavy metal ions

To unfold the fluorescence sensing ability of the synthesized nanophosphor utilized for the detection of toxic heavy metal

ions (Cd^{2+} , Fe^{3+} , Cr^{3+} , Co^{2+}) a series of luminescence-based experiments were performed. Equal amounts of the metals ions were added to the dispersed solution of nanophosphor. The suspension was taken in a 4 mm cuvette for luminescence measurement and compared with the luminescence intensity of the bare sample without the addition of analytes. Similar experiments were performed with other metal ions. The results demonstrated that the luminescence emission intensity of nanophosphor gets significantly influenced with the addition of different metal analytes bearing different concentrations ranging from 0 to 100 ppm. It was observed that the PL intensity is significantly quenched by the addition of metals ions and results obtained are displayed in Fig. 11 to 14 (for various metals analytes). The detection mechanism was strictly based on the interaction of the lone pair of NH_2 group present on the surface of nanophosphor with the metals ions leading to charge transfer from the electron rich nanophosphor material to the vacant metal ions resulting in luminescence quenching and ultimate metal sensing. Meanwhile, it was observed that 100 ppm solution of metal ion completely quenched the luminescence intensity of the nanophosphor.

The gradual increased luminescence quenching with increase in the concentration of analytes was analyzed by using the Stern–Volmer (SV) equation.⁶⁶



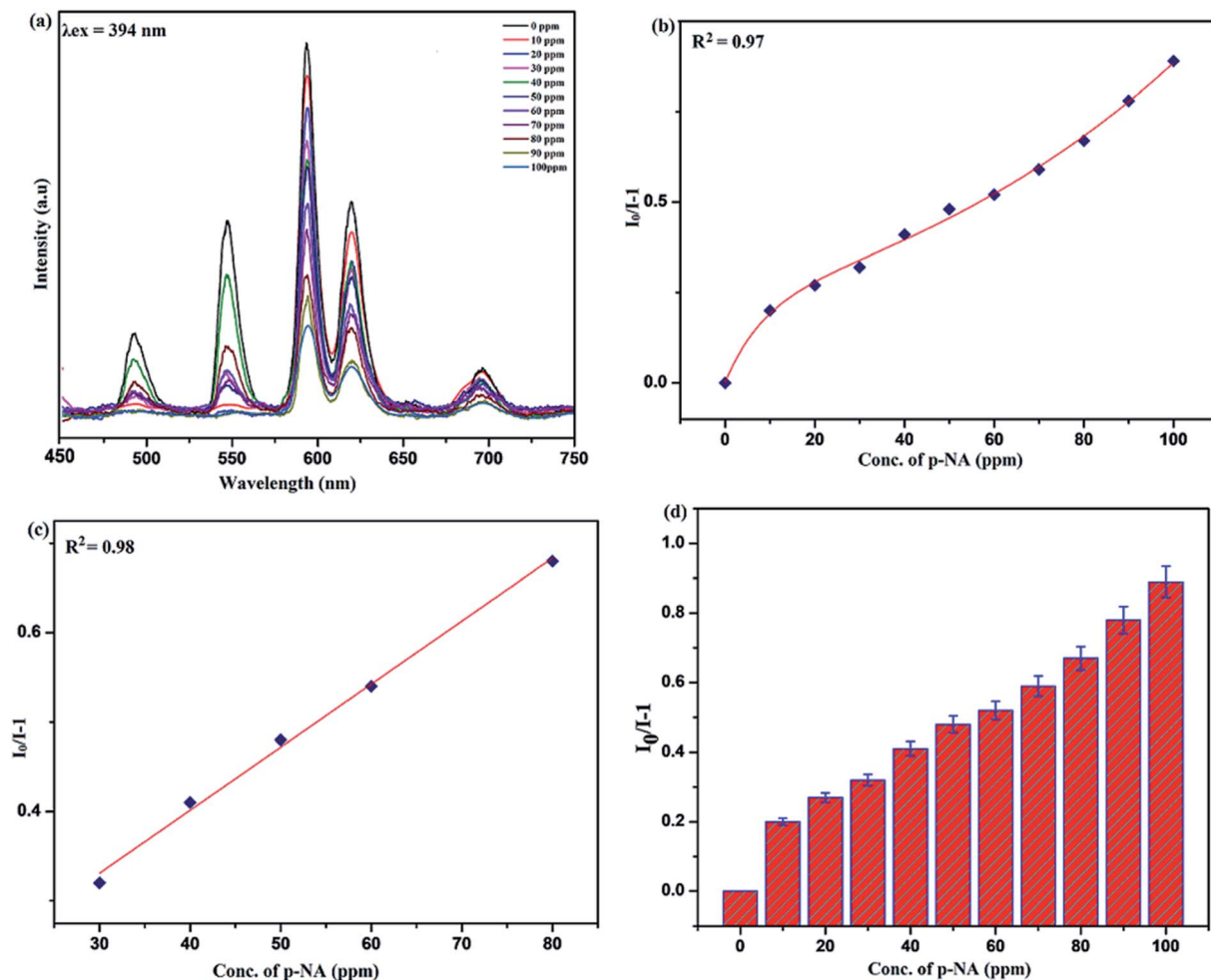


Fig. 15 Variation in photoluminescence intensity of nanosensor in presence of different *p*-NA concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394$ nm: (a) emission spectra of nanophosphors with addition of *p*-NA (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (*p*-NA) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

$$I_0/I = 1 + K_{\text{SV}}[Q]$$

$$K_{\text{SV}} = k_q\tau_f$$

Here, I_0 is the luminescence intensity in the absence of analytes, and I is the luminescence intensity in the presence of metals ions, Q is the concentration of analytes and K_{SV} is the Stern–Volmer (SV) quenching constant, k_q is the bimolecular quenching rate constant and τ_f is radiative life time. The affinity between PEI-functionalized nanophosphor (Eu-7%) and analytes (heavy metal ions and aromatic nitro-compounds) corresponding to K_{SV} follows the order $\text{Fe}^{3+} > \text{Cr}^{3+} > \text{Co}^{2+} > \text{Cd}^{2+}$. The calculated K_{SV} values for Fe^{3+} , Cr^{3+} , Co^{2+} , Cd^{2+} are $3.8 \times 10^4 \text{ M}^{-1}$, $2.6 \times 10^4 \text{ M}^{-1}$, $1.8 \times 10^4 \text{ M}^{-1}$, $0.85 \times 10^4 \text{ M}^{-1}$ respectively. The Fe^{3+} ions exhibited highest K_{SV} value as compared to other metals ions which suggest high selectivity of the PEI-functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) nanophosphors for Fe^{3+} ion. The quenching efficiency for investigated heavy metal ions is found in a range of 73 to 90.02% (see Fig. S13: ESI†).

Sensing of nitro-aromatic compounds

Benefiting from the above sensing properties of fabricated nanophosphor which could significantly quench the emission in aqueous media, the ability of this nanoprobe was again investigated to sense harmful nitro aromatics at even very low concentration. Similar detection experiment were performed considering selected toxic aromatic nitro compounds like picric acid (PA), *para*-nitrophenol (*p*-NP) *para*-nitroaniline (*p*-NA), and 4-nitrotoluene (4-NT). The photoluminescence results clearly indicate the selective and sensitive detection of nitro analytes by PEI-functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphor at trace level. The sensing principle is purely based on the charge transfer from the lone pairs of the amine group present on the nanophosphor with the electron deficient aromatic nitro compounds. This electronic interaction between the electron rich and electron deficient moieties leads to photoluminescence quenching. The effect of these different moieties on photo physical properties of nanophosphor was studied, by analyzing the emission spectra of PEI-functionalized compound in



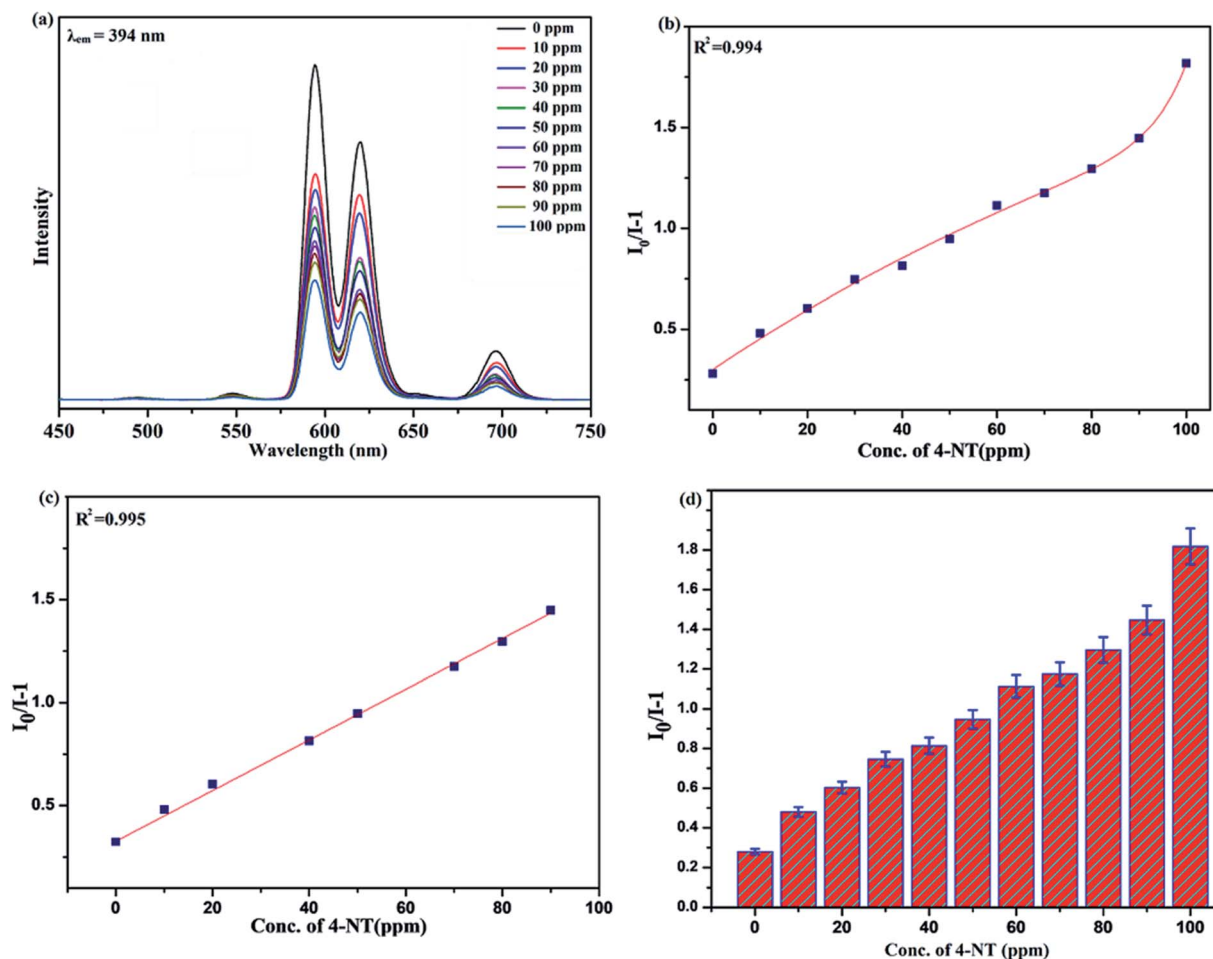


Fig. 16 Variation in photoluminescence intensity of nanosensor in presence of different 4-NT concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394$ nm (a) emission spectra of nanophosphors with addition of 4-NT (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (4-NT) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

distilled water excited at 394 nm in presence of varying concentrations of analytes. It was interestingly noticed that the luminescent intensity of the nanophosphor was significantly affected by addition of various concentration of aromatic nitro-compound ranging from 0 to 100 ppm. The experiments were also performed under similar conditions with $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) nanophosphor without PEI functionalization. The results showed subdued response of luminescence intensity of nanophosphor in presence of analytes. Furthermore, nitro compounds lead to significant quenching of the luminescence intensity of the PEI functionalized nanophosphor and the results obtained are displayed in Figs. 15 to 18 (for different nitro analytes). It is again observed that for 100 ppm solution of picric acid the quenching efficiency is about 97.05%. Thus, when compared with the other nitro analyte solutions, picric acid causes maximum quenching of the luminescence intensity as depicted in Fig. S14: ESI.† Consequently, it can be postulated that the PEI-functionalized $\text{Tb}^{3+}/\text{Eu}^{3+}$ co-doped NaCeF_4 nanophosphor can act as an effective chemo sensors for the sensing of ultra trace amount of harmful picric acid analyte

in aqueous solution. The calculated K_{SV} of nitrocompounds follow the order PA ($2.25 \times 10^5 \text{ M}^{-1}$) > $p\text{-NP}$ ($1.8 \times 10^5 \text{ M}^{-1}$) > $p\text{-NA}$ ($1.1 \times 10^5 \text{ M}^{-1}$) > 4-NT ($0.88 \times 10^5 \text{ M}^{-1}$). The high K_{SV} value for PA suggest the high selectivity of synthesized nanocrystals towards PA. Consequently, it was concluded that the K_{SV} of Fe^{3+} and PA is quite high as compared to other analyte compounds suggesting superior selectivity of the chemosensor for Fe^{3+} and PA.

By using linear regression method detection limit (LOD) for the sensing of different analytes were calculated using formula $3\sigma/S$ where 3 denotes the 95% confidence level factor; S is the slope of linear calibration curve and σ is the standard deviation of the observed intensity for the blank PEI-functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) nanostructure. The LOD for different aromatic nitrocompounds and heavy metal ions are as follow: 1.32 ppm (PA), 2.88 ppm ($p\text{-NP}$), 4.24 ppm ($p\text{-NA}$), 5.59 ppm (4-NT), 1.39 ppm (Fe^{3+}), 2.58 ppm (Cr^{3+}), 3.85 ppm (Co^{2+}) and 7.48 ppm (Cd^{2+}) respectively. The PA and Fe exhibit low detection limit as compare to that of other analytes and thereby suggested high sensitivity of the nanoprobe. As shown in Tables 1 and 2, the corresponding

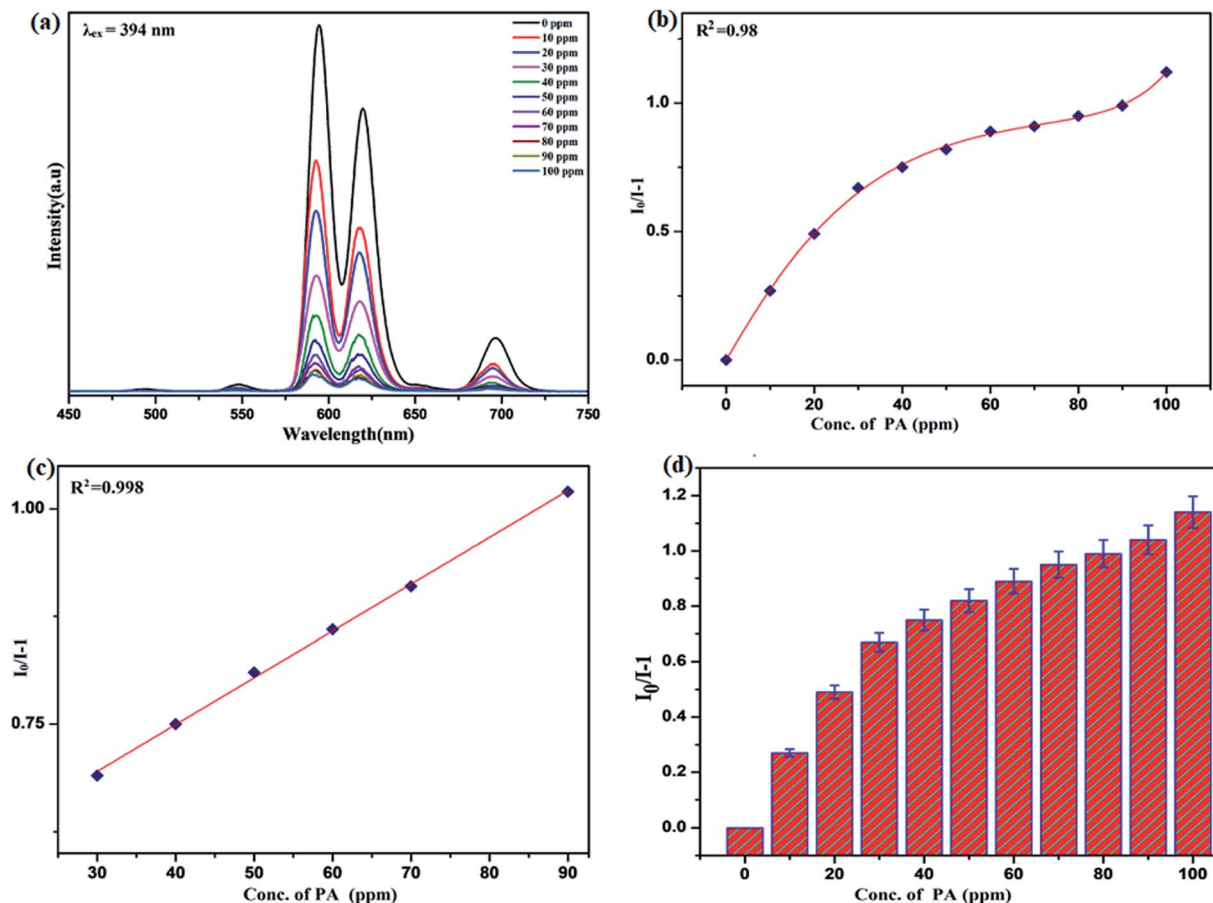


Fig. 17 Variation in photoluminescence intensity of nanosensor in presence of different PA concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{\text{ex}} = 394$ nm: (a) emission spectra of nanophosphors with addition of PA (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (PA) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

K_{SV} and LOD of PA and Fe is higher than those of some previously reported.^{67–76}

Quenching mechanism of metal ions and nitro-aromatic analytes

The possible quenching mechanism for the detection of analytes is based on the donor–acceptor interaction between PEI-functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ (Eu-7%) nanophosphor with different concentrations of analytes (nitroaromatics and heavy metal ions). The electron deficient nitro groups on the nitro aromatics interact with the lone pairs of the amine group present on the nanophosphor. Consequently, charge transfer takes place between the electron donor and electron acceptor group resulting in formation of stable acid–base Jackson Meisenheimer complex having negative charge delocalized over the ring of nitro analytes.⁷⁷ The stable σ -complex so formed consequently leads to luminescence quenching due to energy transfer from luminescence nanophosphor to analyte.⁷⁸ The incessant decrease in the luminescence intensity induces ultimate detection of the nitroaromatic analytes. For the detection of heavy metal ions similar donor–acceptor charge transfer mechanism is

considered to account for the luminescence quenching effect.⁷⁹ The charge transfer occurs from the electron rich PEI functionalized nanophosphor having lone pair of electrons to the electron deficient metals ions having vacant orbitals. The charge migration between the luminophore to the acceptor result in the quenching of luminescence intensity and ultimate detection of heavy metal ions by the PEI doped nanophosphor. The above results thus, specify PEI functionalized nanophosphor as a good candidate for the exquisite detection of nitro aromatics as well as for the heavy metals ions.

Photoluminescence spectra of R-GO@PEI@ $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanocomposite

Additionally, to unravel the impact of reduced graphene oxide on photoluminescence behavior and quenching phenomenon of the nanophosphor, the composites of R-GO and nanophosphor was investigated. Firstly, the photoluminescence spectrum of the RGO@ $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ was recorded at room temperature at excitation wavelength of 394 nm as shown in Fig. 19. Complete quenching of luminescence of the nanocomposite was observed due to the strong quenching effect of



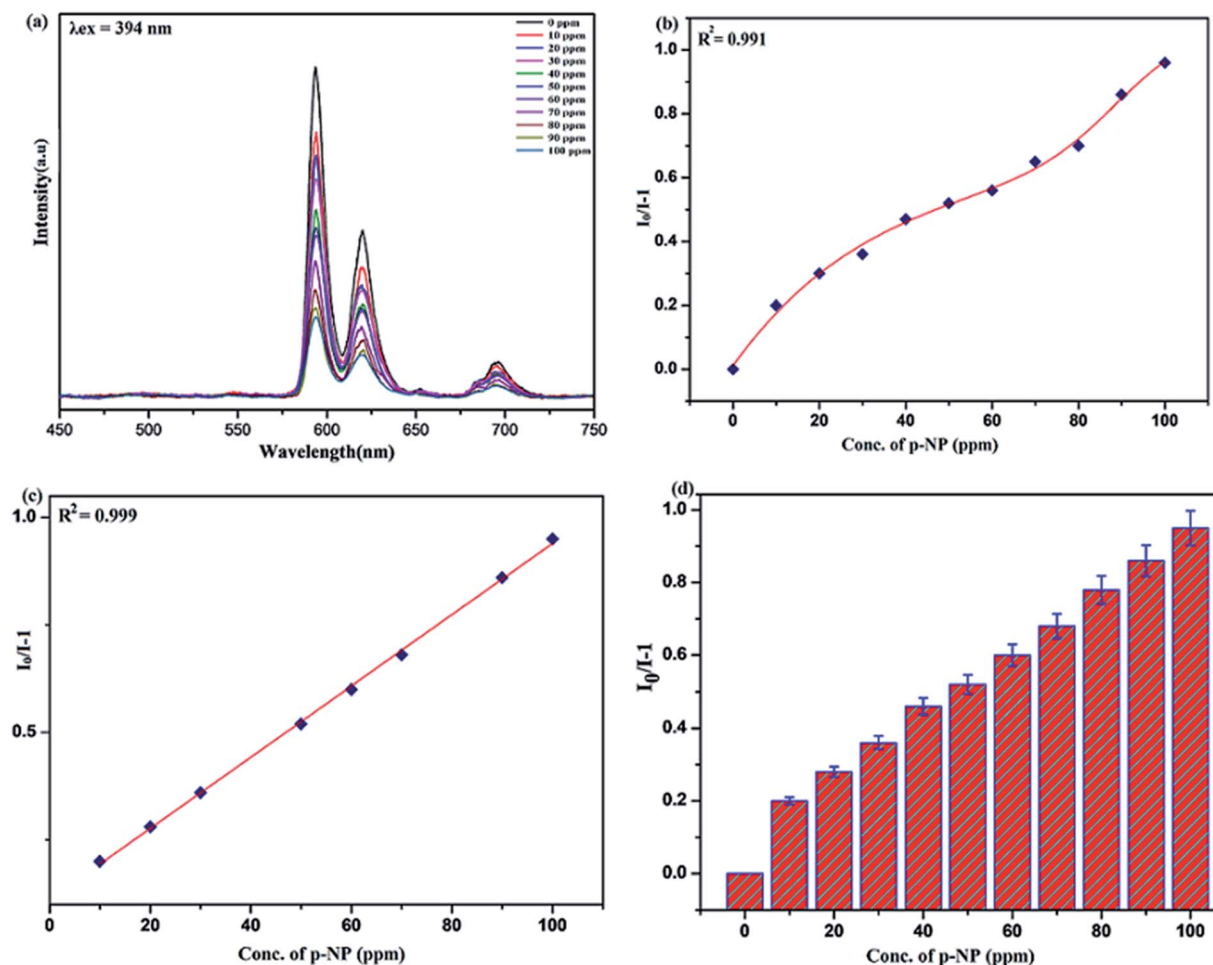


Fig. 18 Variation in photoluminescence intensity of nanosensor in presence of different *p*-NP concentrations (from 0 to 100 ppm) selective excitation at $\lambda_{ex} = 394$ nm: (a) emission spectra of nanophosphors with addition of *p*-NP (b) nonlinear Stern–Volmer plot of $I_0/I - 1$ versus (*p*-NP) concentrations (c) linear Stern–Volmer fitting and (d) error bar.

Table 1 Detection of nitroaromatic compounds by various methods from different nanomaterials

Sensing method	Materials	Detection of NACs	K_{SV} (M^{-1})	LOD	Reference
Electrochemical method	Cu ₂ O	PA	—	39 μ M	67
Electrochemical method	Nano-gold/glassy carbon	NP	—	8 μ M	68
Electrochemical method	AuBFE	PA	—	13.1 μ M	69
Fluorescence method	TPE-CMPs	2,4-DNT	1.84×10^3	—	70
		2,6-DNT	1.70×10^3		
		TNP	1.19×10^4		
Fluorescence method	Tb-MOFs	PA	7.47×10^4	—	71
Fluorescence method	FlrPicOMPbHCz	PA	1.9×10^4	10 μ M	72
		<i>p</i> -NP	5.6×10^3	44 μ M	
		DNP	9.1×10^3	13 μ M	
		DNT	1.2×10^3	44 μ M	
		NP	5.6×10^3	44 μ M	
		NBA	2.1×10^3	25 μ M	
Fluorescence method	[CH ₃ -dtpb]2[Mg ₃ (1,4-NDC)4(μ H ₂ O)2(CH ₃ OH)(H ₂ O)]·1.5H ₂ O	PA	2.8×10^4	—	73
		2,4-DNA	0.8×10^4	—	
Fluorescence method	Serine-functionalised NaYF ₄ :Ce ³⁺ /Gd ³⁺ /Eu ³⁺ @NaGdF ₄ :Tb ³⁺	PA	4.364×10^4	1.67 ppm	40
		<i>p</i> -NP	1.337×10^4	4.57 ppm	
		<i>p</i> -NT	0.824×10^4	6.36 ppm	
Fluorescence method	PEI-functionalised NaCeF ₄ :Tb ³⁺ //Eu ³⁺	PA	$2.25 \times 10^5 M^{-1}$	1.32 ppm	Present work
		<i>p</i> -NP	$1.8 \times 10^5 M^{-1}$	2.88 ppm	
		<i>p</i> -NA	$1.1 \times 10^5 M^{-1}$	4.24 ppm	
		<i>p</i> -NT	$0.88 \times 10^5 M^{-1}$	5.59 ppm	



Table 2 Detection of heavy metal ions by various methods from different nanomaterials

Sensing method	Materials	Detection of heavy metal ions			Reference
		Detection of heavy metal ions	K_{SV} (M^{-1})	LOD	
Electrochemical method	RGO/NiWO ₄ nanocomposite	Cd ²⁺	—	1.11×10^{-10} M	74
		Pb ²⁺	—	2.04×10^{-10} M	
		Cu ²⁺	—	2.80×10^{-10} M	
		Hg ²⁺	—	3.05×10^{-10} M	
Fluorescence method	[(CH ₃) ₂ NH ₂] ⁺ [Tb(bptc)] ⁻	Fe ³⁺	—	76.7 ppm	75
Fluorescence method	[CH ₃ -dtpb] ₂ [Mg ₃ (1,4-NDC)4(μ H ₂ O)2(CH ₃ OH)(H ₂ O)] · 1.5H ₂ O	Fe ³⁺	1.12×10^4	—	73
Fluorescence method	Water-soluble graphene oxide (GO)	Cr ³⁺	0.15×10^4	—	
		Fe ²⁺	5.6×10^3	—	76
Fluorescence method	PEI-functionalised NaCeF ₄ :Tb ³⁺ //Eu ³⁺	Hg ²⁺	2.1×10^3	—	
		Fe ³⁺	3.8×10^4	1.39 ppm	Present work
		Cr ³⁺	2.6×10^4	2.58 ppm	
		Cd ²⁺	1.8×10^4	3.85 ppm	
		Co ²⁺	0.88×10^4	7.48 ppm	

reduced graphene oxide.⁸⁰ Interestingly, the PL spectrum of RGO@PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) functionalized nanocomposite was recorded at excitation wavelength of 394 nm. It was interestingly observed that PL intensity of the nanophosphor recovers in the functionalized nanocomposite which was earlier quenched. The recovery of the PL intensity may be due to the presence of the PEI ligand which was capped around the nanophosphor surface. The PEI acts as bridging ligand which interconnects reduced graphene oxide and nanophosphor core. It was inferred that the PEI functionalized luminescent nanophosphor does not directly communicate with the graphene oxide. The quenching process taking place in the pure nanophosphor can be explained because of the direct interaction of core material with the RGO sheets leading to energy transfer and ultimate quenching of the PL in nanophosphor. On the other hand in the functionalized nanocomposite the PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) complex is anchored on the RGO *via* the electrostatic interaction as well as

hydrogen bonding between the NH₂ group of PEI and the functional moieties of RGO. This interaction keeps the core materials at far distance from the oxygen functionalizing group and thus reducing the quenching effect.⁸¹ It may be assumed that firstly, there occur strong direct interactions between oxygen functionality of RGO and core materials leading to quenching of the luminescence intensity. However, when the PEI was used as the additional ligand, the NH₂ group present on the PEI molecule strongly interacts with the different functional moieties present on the negatively charged RGO sheets thus weakening the interaction between core materials and ultimately inhibiting quenching leading to recovery of the PL in nanophosphor.

In conclusion, the self assembly of the nanophosphor with the RGO not only achieves recovery in the photoluminescence intensity but there occur red PL emission from the NaCeF₄:Tb³⁺/Eu³⁺ as shown in Fig. 19. The red emission of the nanocomposite was confirmed by the presence of highly intense peak at 624 nm occurring due to transition from ⁵D₀ → ⁷F₂ responsible for red emission in the PL spectrum.⁸² Thus, culminated that the binding of PEI as capping agent not only minimizes the quenching of luminescence intensity of nanophosphor but also results in the red emission of the luminescence nanocomposite providing it as a novel candidate for emerging applications as optical sensors.

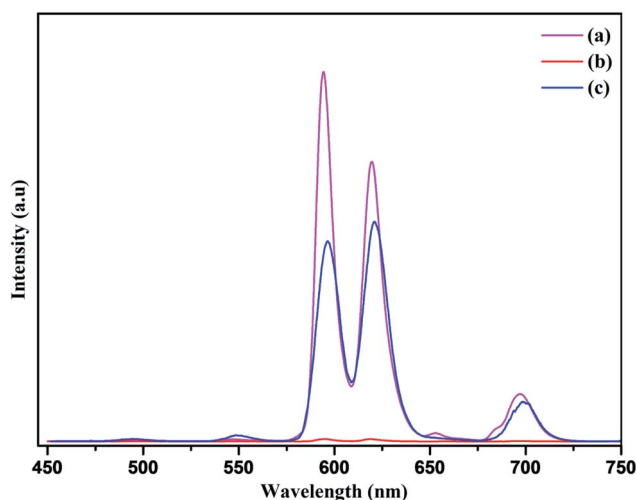


Fig. 19 Photoluminescence emission spectra of (a) NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) nanophosphor (b) R-GO@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%) (c) R-GO@PEI@NaCeF₄:Tb³⁺/Eu³⁺ (Eu-7%).

Conclusions

In summary, we have elaborately designed a highly selective and rapid luminescence nanosensor that can effectively detect the toxic analytes in the waste water. The presence of amine groups on the surface of PEI functionalized NaCeF₄:Tb³⁺/Eu³⁺ nanostructure enhance the efficiency of the detection of heavy metal ions and nitroaromatic compounds extensively. Additionally, with increased amounts of heavy metal ions and nitroaromatics, the luminescence emission of the fabricated nanoparticles is exponentially quenched. The quenching efficiency for the Fe³⁺ ion and PA are found to be 90.02% and 97.05% respectively. Furthermore, K_{SV} follows the order of (Fe³⁺ > Cr³⁺ >



$\text{Co}^{2+} > \text{Cd}^{2+}$ with highest K_{SV} value for Fe^{3+}) and for aromatic nitro compound as ($\text{PA} > p\text{-NP} > p\text{-NA} > 4\text{-NT}$). The data suggests high selectivity of novel PEI functionalized $\text{NaCeF}_4:\text{Tb}^{3+}/\text{Eu}^{3+}$ nanophosphor for the detection of Fe^{3+} and picric acid (PA).

Author contributions

Richa Singhaal: conceptualization, formal analysis, investigation & writing original draft. Lobzang Tashi: methodology and writing. Zaib ul Nisa: data curation & writing. Nargis Akhter Ashashi: software, review & editing. Charanjeet Sen and Swaita Devi: resources & validation. Haq Nawaz Sheikh: project administration & supervision.

Conflicts of interest

The authors declare no competing financial interest.

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