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Ultrafast conversion of carcinogenic 4-nitrophenol into 4-aminophenol in the dark catalyzed by surface interaction on BiPO₄/g-C₃N₄ nanostructures in the presence of NaBH₄†

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The heterogeneous catalytic conversion of pollutants into useful industrial compounds is a two-goals at once process, which is highly recommended from the environmental, economic, and industrial points of view. In this regard, design materials with high conversion ability for a specific application is required to achieve such a goal. Herein, the synthesis conditions for the fabrication of BiPO₄ nanorod bundles supported on g-C₃N₄ nanosheets as heterojunction composites was achieved using a facile *ex situ* chemical deposition for the reductive conversion of carcinogenic 4-nitrophenol (4-NP) into 4-aminophenol (4-AP). To better understand the mechanistic reduction pathways, BiPO₄/g-C₃N₄ composites with varying ratios were obtained. The morphology and structure of BiPO₄/g-C₃N₄ composites were checked using several methods: XRD, FE-SEM, HRTEM, XPS, and FT-IR, and it was found that hexagonal phase BiPO₄ nanorod bundles were randomly distributed on the g-C₃N₄ nanosheets. Overall, the reduction ability of BiPO₄/g-C₃N₄ composites was far better than bare BiPO₄ and g-C₃N₄. A total reductive conversion of 4-NP at a concentration of 10 mg L⁻¹ into 4-AP was found with 50% BiPO₄/g-C₃N₄ composite within only one minute of reaction. Moreover, the presence of reducing agent (NaBH₄) enhanced the kinetic rate constant up to 2.914 min⁻¹ using 50% BiPO₄/g-C₃N₄, which was much faster than bare BiPO₄ (0.052 min⁻¹) or g-C₃N₄ (0.004 min⁻¹). The effects of some operating parameters including the initial concentration of 4-NP and catalyst dosage were also evaluated during the experiments. BiPO₄/g-C₃N₄ showed great stability and recyclability, wherein, the catalytic reduction efficiency remains the same after five runs. A plausible 4-NP reduction mechanism was discussed. The high catalytic activity with the good stability of BiPO₄/g-C₃N₄ make it a potential candidate for the reduction of nitroaromatic compounds in real wastewaters.

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1. Introduction

Petrochemical and pesticide manufacturing industries discharge huge amounts of nitroaromatic compounds, *e.g.*, nitrophenols, which are known as potential environmental hazards to human health because of their high degree of carcinogenic risk and non-biodegradability in the aquatic or/and soil environments.^{1,2} Many efforts have been made to

remove nitrophenols from the environment using different physicochemical or photonic techniques.^{3–7} Of these, the reduction of nitrophenol is a great strategy as it achieves two goals at the same time, including the removal of toxic compound from the medium, plus the production of aminophenol as a valuable compound for industrial use such as in pharmaceutical and drug processing.⁸ Therefore, the intensification of green technologies for enhanced and selective reduction of nitrophenol into aminophenol has received much attention recently due to the important environmental and economic impact.^{9,10} In this regard, over the last decades, semiconductor catalysts have attracted wide attention for removing environmental pollutants or/and converting them into useful compounds for industrial use.^{11,12} Numerous semiconductor materials with appropriate physicochemical or/and photochemical properties have been widely developed and used for the catalytic and photocatalytic transformation of compounds.^{13,14}

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g-C₃N₄ has been considered as one of the most successful π -conjugated polymers over the last decades, due to its high stability, good absorption in the visible light and photo-efficiency.¹⁵ Owing to the existence of the lamellar structure, g-C₃N₄ is always well-crystallized, allowing the transfer of charges.¹⁶ In view of its unique structure, and photochemical characteristics, g-C₃N₄ has been widely employed in various fields, for instance, removal of pollutants from water or air,¹⁷ reduction of hazardous metals,¹⁸ water splitting,¹⁹ hydrogen evolution.²⁰ Therefore, g-C₃N₄ as a metal-free material is identified to be a “sustainable develop” catalyst. Regardless, the photocatalytic effectiveness of g-C₃N₄ is still limited due to several issues including the fast recombination of electron/hole charges and low adsorption capacity.^{21,22} To limit the recombination of charges and enhance the catalytic ability of g-C₃N₄, the surface modification through metal doping²³ or non-metal doping processes²⁴ were extensively reported. In terms of 4-nitrophenol reduction, Pd, Ag, Au doped g-C₃N₄ have been utilized to improve the catalytic activity of 4-nitrophenol reduction.^{25–28} On top of that, the combination of g-C₃N₄ with other semiconductors to form heterojunction systems which gives a range of synergistic benefits, wherein, the interfacial charge-transfer (IFCT) is the key advantage as it can enhance significantly the yield of long lifetime of separated redox charge carriers.^{29,30} For a successful construction of a heterojunction systems, the semiconductors should be properly chosen. In recent decades, the scientific community has addressed widely the application of bismuth based nanomaterials such as BiOBr, Bi₂S₃, BiPO₄, BiVO₄, Bi₂WO₆, Bi₂O₂CO₃ in different catalytic applications thanks to their appropriate physicochemical and photonic characteristics.^{31–36} In particular, BiPO₄ as a new bismuth salt catalyst, which has been developed for the first type by Zhu's group, has been revealed as promising candidate in semiconductor catalysis due to its highly photocatalytic efficiency under UV light which was found to be superior than TiO₂ (P25, Degussa) towards the oxidation of dyes.³⁷ However, the catalytic performance of BiPO₄ has been greatly restricted. Like other semiconductors, a lot research has been carried out on BiPO₄ to improve its (photo)-catalytic efficiency and visible light response. One of the used strategies is the combination of BiPO₄ with other semiconductors to get multifunctional heterojunctions with combined options, which could enhance the surface interactions, charges transfer and better visible light response.^{38–40} Other ways have been also reported including surface hybridization with π -conjugated materials,^{41–43} creation of surface oxygen vacancies,^{44,45} and phase junction system.^{46,47}

The combination of g-C₃N₄ with BiPO₄ has been proved to be very successful due to the excellent interfacial interaction between the two nanostructured materials in terms of high yield of separated charged and enhanced (photo)-catalytic reactions.^{48–50} However, the catalytic reduction application and mechanism of heterojunction BiPO₄ with g-C₃N₄ have not been thoroughly investigated.

In this work, we report the fabrication of BiPO₄ nanorod bundles/g-C₃N₄ nanosheets composite architectures with ultrafast catalytic efficacy towards the reduction of 4-nitrophenol. As far as we know, this is the first investigation on the use of BiPO₄/g-C₃N₄ to initiate the reduction of 4-NP in dark condition. The catalytic efficiencies bare BiPO₄ and g-C₃N₄ were studied as well for the purpose of comparison. Additionally, the effect of different BiPO₄:g-C₃N₄ ratios on the catalytic effectiveness was checked. The effect of some operating parameters such as the initial concentration of 4-NP and catalyst amount on the catalytic performance was studied. The stability of BiPO₄/g-C₃N₄ composite and the feasible mechanistic pathway for the catalytic reduction of 4-NP were addressed.

2. Experimental

All chemicals and characterization instruments were reported in the ESI.†

2.1 Synthesis of BiPO₄ nanorod bundles

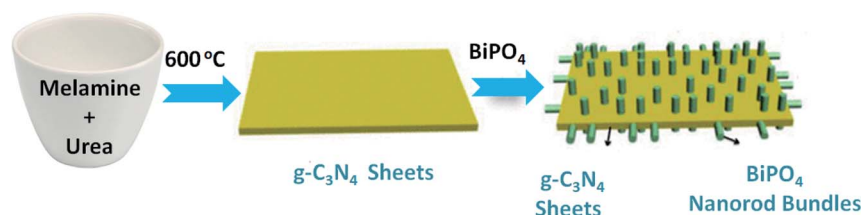
The details of the synthesis of BiPO₄ were reported in our previous study.⁵¹ Typically, 1.94 g of Bi(NO₃)₃·5H₂O and 0.5 g CTAB were stirred in 40 mL dimethyl sulfoxide. Then 0.524 g (NH₄)₂HPO₄ dissolved in 40 mL of water was dropped into the previous solution under vigorous stirring for 1 h. After the formation of precipitate, the solid was recovered with centrifugation, and then washed with water and ethanol, and finally it was dried for 6 h at 90 °C.

2.2 Synthesis of g-C₃N₄ nanosheets

g-C₃N₄ was synthesized through the direct heating of mixture of melamine and urea (5 : 5 g) at 600 °C in a muffle furnace for 2 h with 2 °C min^{−1} heating rate. The obtained yellow solid was ground.

2.3 Synthesis of BiPO₄ nanorod bundles/g-C₃N₄ nanosheets

To develop BiPO₄/g-C₃N₄ heterojunction photocatalyst, the *ex situ* route was adopted, as shown in Scheme 1, the growth of



Scheme 1 Diagram representation for the synthesis and structure of BiPO₄/g-C₃N₄ heterojunction.



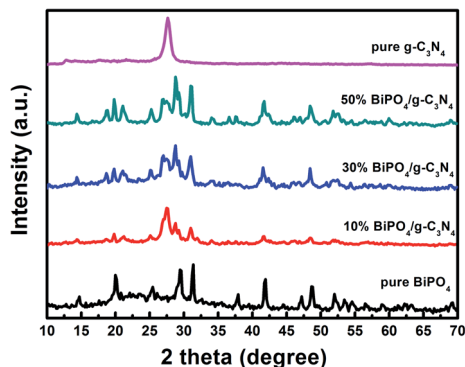


Fig. 1 Powder XRD patterns of pure BiPO_4 , $\text{g-C}_3\text{N}_4$, 10% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, and 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$.

BiPO_4 nanorod bundles within sheets of $\text{g-C}_3\text{N}_4$ was obtained. In details, 0.5 g of $\text{g-C}_3\text{N}_4$ was mixed with 50 mL methanol and stirred for 1 h. Different amounts of BiPO_4 was ultrasonically dispersed into $\text{g-C}_3\text{N}_4$ suspension for 10 min and complete stirring for 1 h to denote as 10% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, and 70% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ samples. Finally, the composites were washed with water and dried for 6 h at 90°C .

2.4 Catalytic reduction test of 4NP

The catalytic behavior was investigated towards the reductive conversion of 4-nitrophenol into 4-aminophenol in the presence of NaBH_4 . First, 50 mL of an aqueous solution of 4-nitrophenol (10 ppm) was mixed with NaBH_4 (16 mM), which leads to form a strong yellow solution. Then, 25 mg of the catalyst was added in dark to yellow solution and the reaction was

performed until the solution became colorless. A 3 mL aliquots were taken at different time intervals, and the solid suspension was removed *via* centrifugation. The concentration of compounds was followed by the use of UV-vis spectrophotometer. The recycling of $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ was carried out to check the stability. For this, the catalyst was recovered after the reaction, and the washed with double distilled water and dried at 90°C for subsequent tests.

3. Results and discussion

3.1 Characterization

Fig. 1 illustrates the XRD patterns of bare BiPO_4 , $\text{g-C}_3\text{N}_4$, and $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ hybrid composites. The diffraction peaks of BiPO_4 at 2-theta angles 14.7° , 20.1° , 25.5° , 29.5° , 31.3° , 41.9° , and 48.7° which corresponding to the crystal orientations of (100), (101), (110), (200), (102), (211), and (212) planes of BiPO_4 . These diffraction peaks correspond to JCPDS card no. 15-0766.⁵² It can be seen from the diffraction peaks that BiPO_4 is a hexagonal phase (HBIP, space group: $P3_121$). $\text{g-C}_3\text{N}_4$ pattern shows an intense peak at $2\theta = 27.7^\circ$ corresponding to its crystal plane (002) of stacking conjugated aromatic system, while the peak at 12.9° is due to the trietiazine units.⁵³ The diffraction peaks of $\text{g-C}_3\text{N}_4$ from mixture of urea and melamine is more intense than as prepared by urea as shown in Fig. S1.† No other peaks or impurities are observed, implying that the final products of BiPO_4 and $\text{g-C}_3\text{N}_4$ are of pure phases. The diffraction pattern of $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ shows that the composite is composed of the hexagonal phase BiPO_4 and $\text{g-C}_3\text{N}_4$. The intensity of diffraction peaks of BiPO_4 are gradually strengthen with increasing BiPO_4 content. Consequently, the above results confirm the successful formation $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ heterojunction.

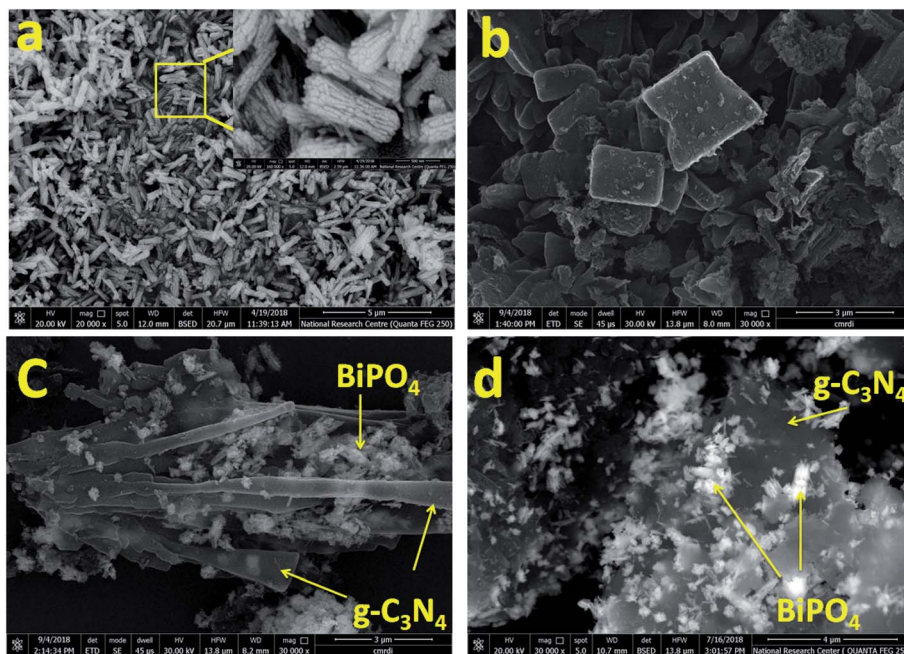


Fig. 2 FE-SEM images of BiPO_4 (a), $\text{g-C}_3\text{N}_4$ (b), and 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (c), and 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (d).

The surface morphologies of bare BiPO_4 , $\text{g-C}_3\text{N}_4$, 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, and 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ catalysts were characterized by FE-SEM. As depicted in Fig. 2, the diameter of nanorod bundles in BiPO_4 was found to be 300 nm, and lengths of 800 nm (Fig. 2a). Whereas the $\text{g-C}_3\text{N}_4$ displayed a typical sheet-like morphology with several crystallites (Fig. 2b). After the heterojunction was successfully formed in 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, and 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ samples, the agglomeration of BiPO_4 could be observed by the insertion of BiPO_4 nanorods bundles in the inter sheets of $\text{g-C}_3\text{N}_4$ (Fig. 2c and d), which implies that the $\text{g-C}_3\text{N}_4$ sheets are coated with BiPO_4 nanorods bundles.

The TEM image (Fig. 3a) revealed that BiPO_4 was composed of nanorods with diameter ranging from 30 to 40 nm. HRTEM (Fig. 3b) showed that BiPO_4 nanorod bundles have a lattice fringe spacing of 0.349 nm, which can be indexed to (110) crystal plane of BiPO_4 (JCPDS card no. 15-0766). The SAED pattern (Fig. 3c) confirmed the polycrystalline nature of the material due to the presence of concentric ring shape like. The TEM images shown in Fig. 3d indicated that BiPO_4 nanorod

bundles are bonded to $\text{g-C}_3\text{N}_4$ nanosheets. HRTEM (Fig. 3e) shows a fringe spacing of 0.350 nm due to (110) crystal planes of BiPO_4 . In addition, the polycrystalline phase was observed in SAED pattern (Fig. 3f). It can be seen from TEM images (Fig. 3g) that $\text{g-C}_3\text{N}_4$ has an unsmooth surface, which could boost the catalytic reaction by offering a higher surface area or/and favors the immobilization of other semiconductor nanoparticles. From the results of materials characterization, it can be concluded that the BiPO_4 nanorod bundles was successfully immobilized on the $\text{g-C}_3\text{N}_4$ nanosheets to form heterojunction system.

The chemical oxidation state of $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ hybrid structure was verified using XPS analysis. Fig. 4a illustrates the full scan survey which indicate the presence of Bi, P, O, C and N elements, implying the successful heterojunction of BiPO_4 and $\text{g-C}_3\text{N}_4$. The high resolution Bi 4f, P 2p, O 1s, C 1s, and N 1s are shown in Fig. 4b–f. The binding energies peaks at 159.81 eV and 165.11 eV (Fig. 4b) are due to the Bi 4f_{7/2} and Bi 4f_{5/2} levels in BiPO_4 , respectively.⁵⁴ The Bi 4f_{7/2} and Bi 4f_{5/2} peaks showed

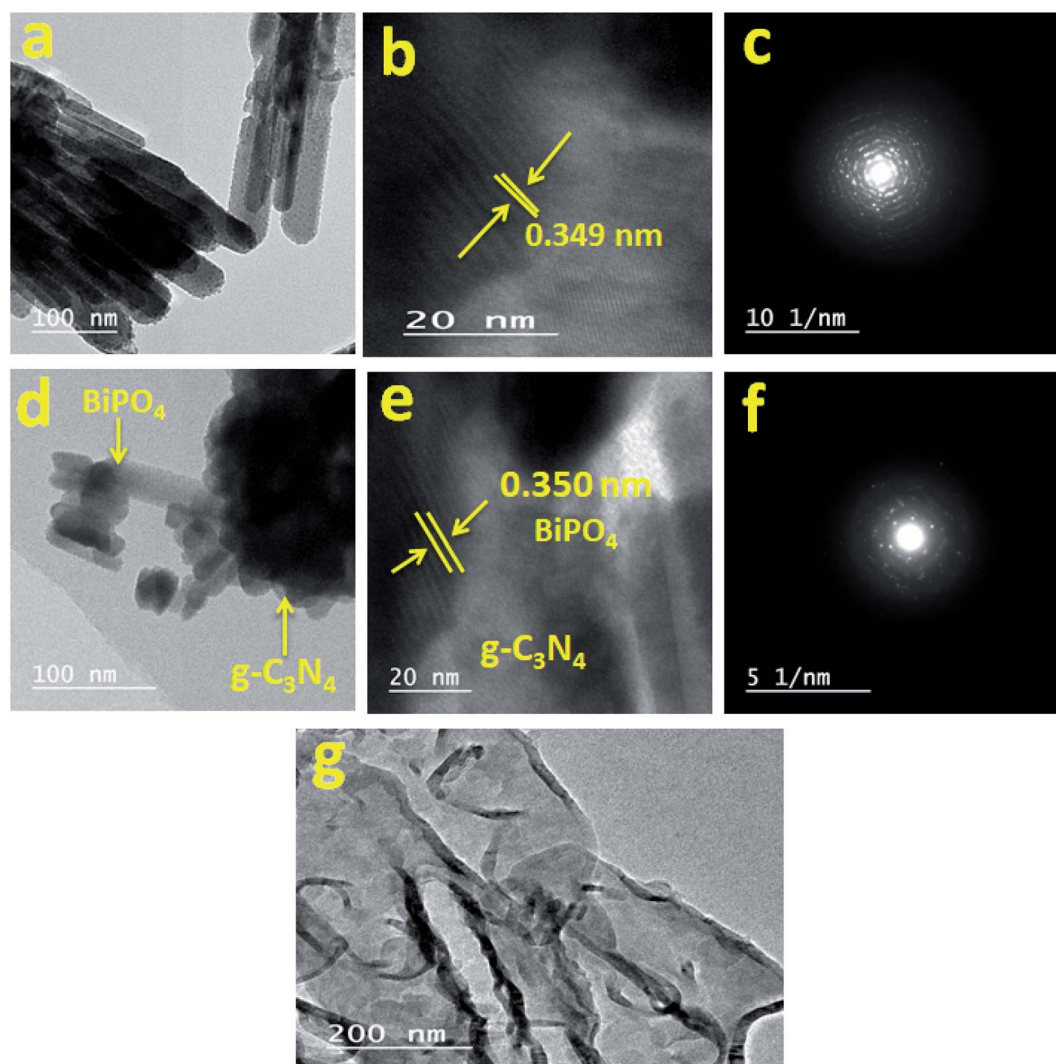


Fig. 3 TEM, HRTEM images and SAED pattern of BiPO_4 (a, b, and c, respectively), $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ hybrid structure (d, e, and f, respectively), and TEM image of $\text{g-C}_3\text{N}_4$ (g).



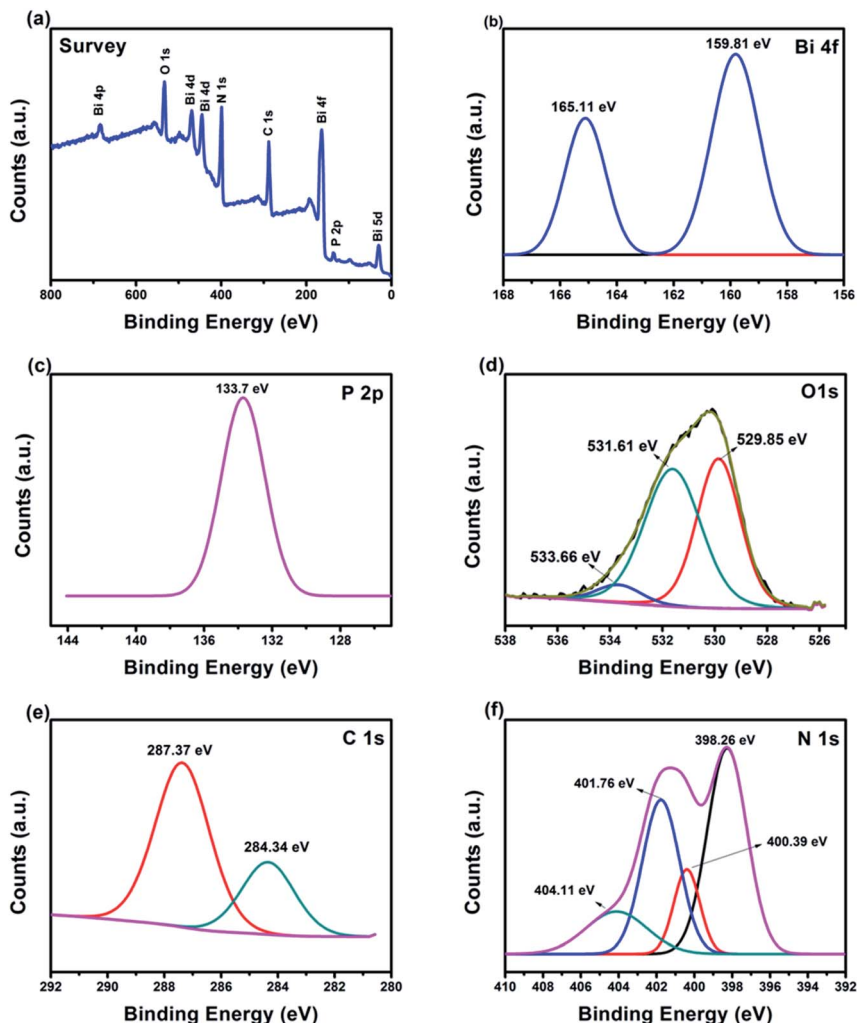


Fig. 4 XPS spectra of $\text{BiPO}_4/\text{g-C}_3\text{N}_4$. (a) survey spectra, (b) Bi 4f, (c) P 2p, (d) O 1s, (e) C 1s, (f) N 1s.

a splitting of 5.3 eV which confirms that Bi atoms have a valence state of +3.⁵⁵ The characteristic peak of P 2p can be found at 133.7 eV (Fig. 4c), which ascribed to the P^{5+} in BiPO_4 .⁵⁶ In Fig. 4d, the peaks at 529.85, 531.61 and 533.66 eV correspond to O^{2-} bonded to Bi, surface bound and adsorbed oxygen, respectively.⁵⁷ The two peaks in C 1s spectra appeared at 284.34 and 287.37 eV are due to the sp^2 bonded carbon atoms such as carbon-carbon bonding (C-C), and carbon-nitrogen double bond (C=N), respectively (Fig. 4e).^{58,59} The high resolution N 1s XPS spectra in Fig. 4f is divided into four bands with energies of 398.26, 400.39, 401.76 and 404.11 eV correspond to C-N=C, N-(C)₃ groups, and (C)₂-N-H, respectively.^{50,60} The XPS data is an accordance with XRD, FE-SEM, and TEM analysis.

FTIR spectra of bare semiconductors and $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ composites are shown in Fig. 5. In terms of BiPO_4 , bending vibration peaks due to PO_4 group are detected at 587 and 534 cm^{-1} . The band at 974 cm^{-1} is assigned to PO_4 stretching vibration.⁴¹ The band at 3495 cm^{-1} corresponds to hydroxyl stretching vibration groups of the surface adsorbed or/and structure water molecules in BiPO_4 . $\delta(\text{H-O-H})$ bending vibrations can be observed also at 1602 cm^{-1} . $\text{g-C}_3\text{N}_4$ spectrum shows a large band at 3155 cm^{-1} as a result of stretching

vibration in NH and NH_2 groups.⁶¹ The heptazine heterocyclic ring stretching vibrations in $\text{g-C}_3\text{N}_4$ units were appeared at 1200–1650 cm^{-1} . In addition, the peaks at 1310 and 1629 cm^{-1} correspond respectively to the stretching vibrations of C-N and C=N.⁶¹ The characteristic peak at about 800 cm^{-1} could be assigned to the triazine units.⁶² For $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ samples, the

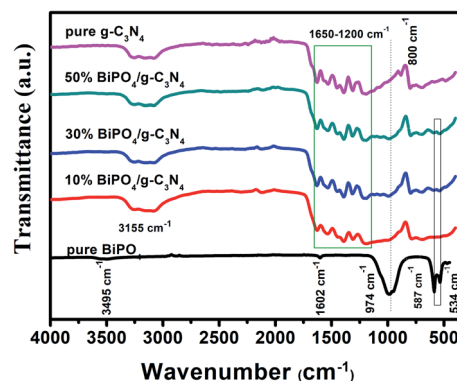


Fig. 5 FT-IR spectra of pure BiPO_4 , $\text{g-C}_3\text{N}_4$, 10% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$, and 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$.

entire characteristic peaks of (PO_4) group (bending, stretching vibration) in BiPO_4 are gradually strengthen with increasing BiPO_4 content, implying that the BiPO_4 and $\text{g-C}_3\text{N}_4$ are of harmonious coexistence.

3.2 4-NP catalytic reduction

The catalytic performances of as-prepared catalysts towards the reduction of 4-NP to 4-AP in the presence of NaBH_4 are shown in Fig. 6. The maximum absorption peak of aqueous solution of 4-

NP (Fig. 6a) is shown at 318 nm but after addition of NaBH_4 the peak shifted to 400 nm with a color conversion from light yellow to strong yellow due to NaBH_4 converted 4-nitrophenol to 4-nitrophenolate.^{63,64} The coexistence of these two peaks indicates the copresence of 4-nitrophenol and nitrophenolate during the equilibrium, which is due to 4-nitrophenol self-ionization.¹⁰ In terms of $\text{g-C}_3\text{N}_4$, no significant reduction was observed for pure $\text{g-C}_3\text{N}_4$ in presence of NaBH_4 (Fig. 6b). As seen in Fig. 6c–f, when 0.5 g L^{-1} of $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ composite was added to the reaction medium, a fast disappearance of yellow color was found, and

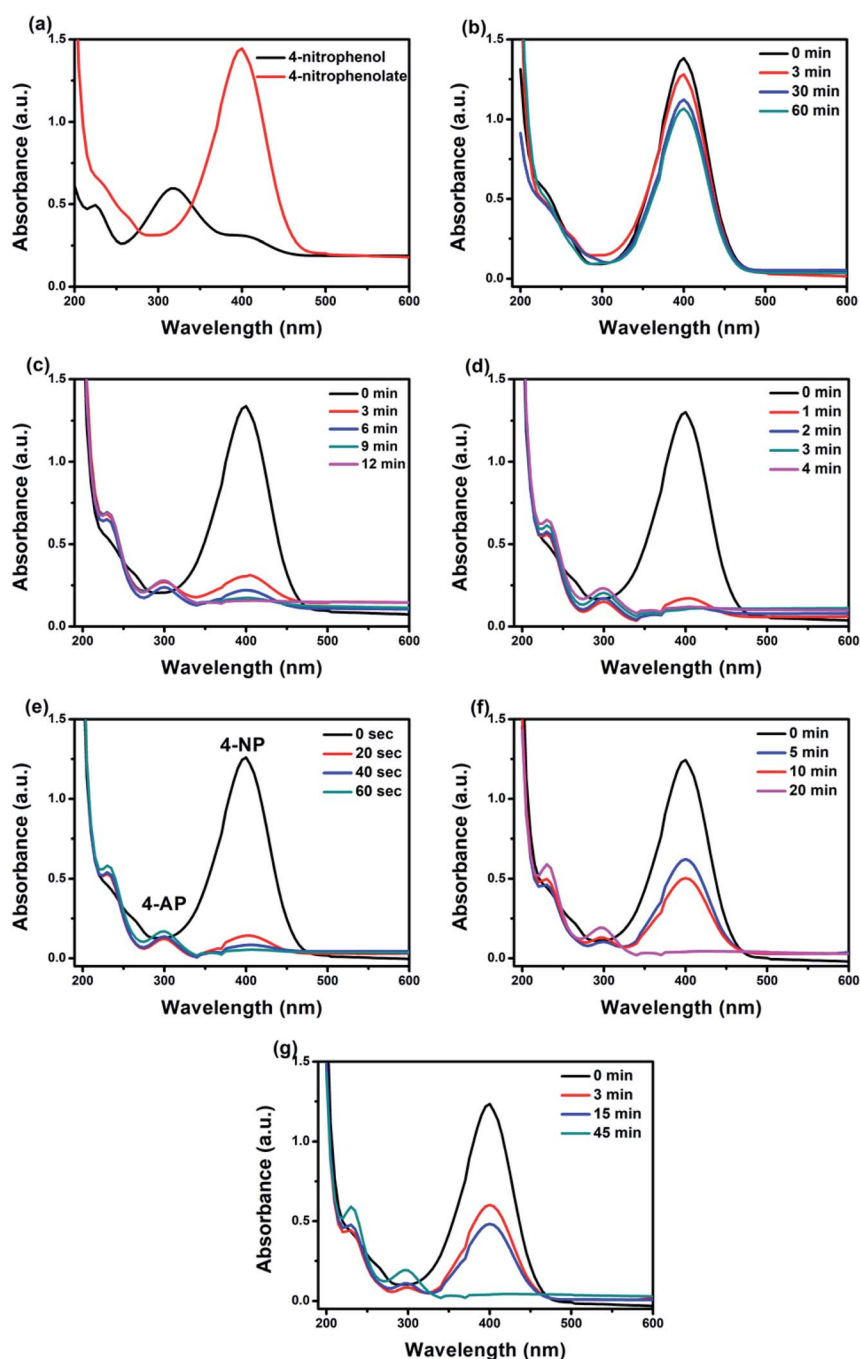


Fig. 6 UV-vis spectral changes of 4-NP before and after adding NaBH_4 solution (a); the successive reduction of 4-NP to 4-AP in presence of pure $\text{g-C}_3\text{N}_4$ (b), 10% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (c), 30% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (d), 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (e), 70% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ (f) and pure BiPO_4 (g). ($[\text{4-NP}] = 10 \text{ mg L}^{-1}$, $[\text{catalyst}] = 0.5 \text{ g L}^{-1}$).



a quick decrease in the peak intensity of 400 nm, and at the same time the characteristic peak of 4-AP at 300 nm was appeared, reflecting the reduction reaction.⁶⁵ Time consumed for reduction of 4-NP is 12, 4, 1, and 20 min for 10% BiPO₄/g-C₃N₄, 30% BiPO₄/g-C₃N₄, 50% BiPO₄/g-C₃N₄, and 70% BiPO₄/g-C₃N₄, respectively. In addition, bare BiPO₄ showed similar spectral tendency within 45 min (Fig. 6g). The order of the catalytic performances was as follows 50% BiPO₄/g-C₃N₄ > 30% BiPO₄/g-C₃N₄ > 10% BiPO₄/g-C₃N₄ > 70% BiPO₄/g-C₃N₄ > BiPO₄ > g-C₃N₄. Increasing percentage of BiPO₄ to 70% may cover surface of g-C₃N₄ and decreased the catalytic reduction of 4-NP. The enhancement of catalytic performance ability could be attributed to the synergistic heterojunction effect between BiPO₄ and g-C₃N₄ nanosheets, high electrical conductivity, and enlarged reaction active sites.

The reduction conversion kinetics of 4-NP were studied by the first-order simplification of Langmuir–Hinshelwood (L–H), using the following equation:⁶⁶

$$\ln(A_t/A_0) = -k_{\text{app}}t$$

Table 1 The k_{app} values for the reduction of 4-nitrophenol with and without catalysts in the presence of NaBH₄

S. No.	Catalyst	K_{app} (min ^{−1})
1	C ₃ N ₄	0.004
2	10% BiPO ₄ /g-C ₃ N ₄	0.186
3	30% BiPO ₄ /g-C ₃ N ₄	0.677
4	50% BiPO ₄ /g-C ₃ N ₄	2.914
5	70% BiPO ₄ /g-C ₃ N ₄	0.13
6	BiPO ₄	0.052
7	Without catalyst	0.001

where A_0 and A_t are absorbance of 4-NP in solution at times 0 and t , respectively, and k is the first-order rate constant (min^{−1}). The k_{app} value is obtained from a liner plot of $\ln(A_t/A_0)$ versus time (t), as shown in Fig. 7e. The rate constant values was 0.004, 0.052, 0.186, 0.677, and 2.914 min^{−1} for pure g-C₃N₄, BiPO₄, 10% BiPO₄/g-C₃N₄, 30% BiPO₄/g-C₃N₄, 50% BiPO₄/g-C₃N₄, and 70% BiPO₄/g-C₃N₄, respectively as shown in Fig. 7e

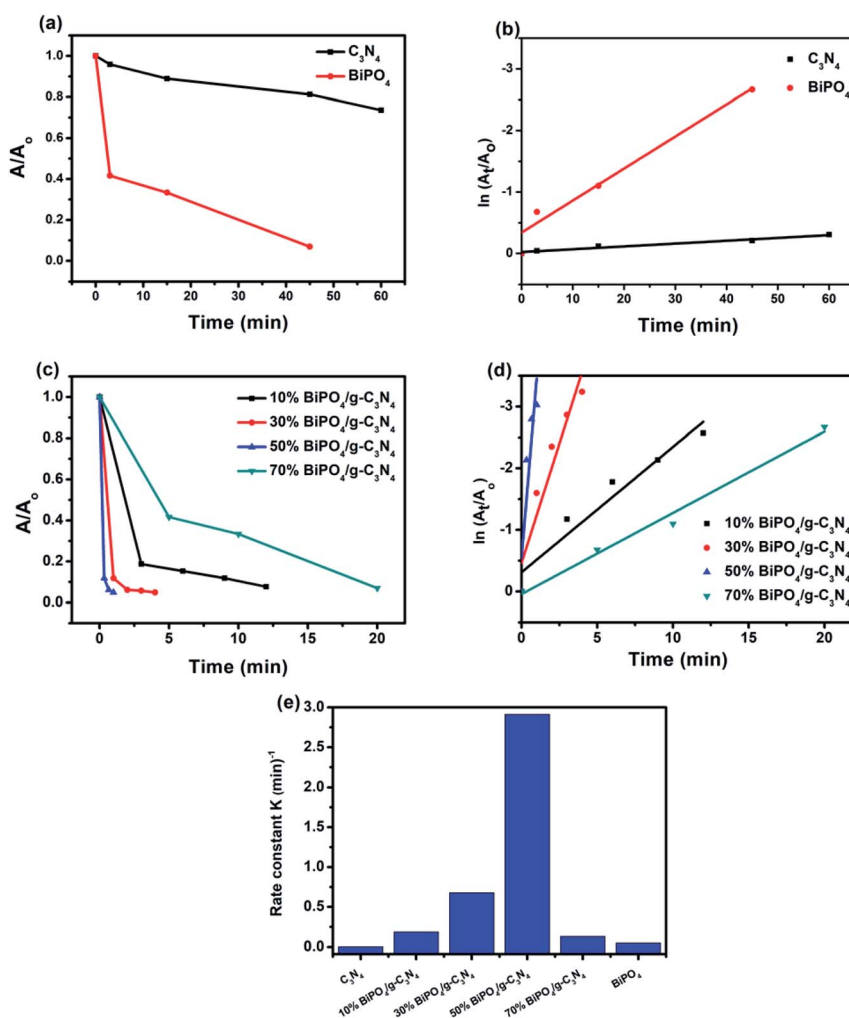


Fig. 7 Reduction rate of 4-NP over different catalysts (a and c); corresponding linear transform $\ln(A_t/A_0) = f(t)$ of the 4-NP reduction kinetics curves (b and d); comparison of the rate constant value for the reduction of the 4-NP over different catalysts (e). ([4-NP] = 10 mg L^{−1}, [catalyst] = 0.5 g L^{−1}).



Table 2 Impact of temperature on k_{app} values for the reduction of 4-nitrophenol ([4-NP] = 30 mg L⁻¹, [catalyst] = 0.5 g L⁻¹)

S. No.	Temperature (K)	K_{app} (min ⁻¹)
1	298	1.1
2	303	1.25
3	308	3.50
4	313	4.67
5	318	5.24

and Table 1. It is clear the increase of the amount of BiPO₄ in the composite leads to increase proportionally the rate constant of the reduction reaction. Therefore, 50% BiPO₄/g-C₃N₄ sample gives the highest k value which was 728.5 and 56 times superior than the values recorded for g-C₃N₄ and BiPO₄, respectively. The impact of temperature on reduction of 4-NP was further investigated using 50% BiPO₄/g-C₃N₄. The k_{app} values improved with the increase in temperature due to increasing diffusion rate of reactant molecules (Table 2).

The turnover number (TON) and the turn over frequency (TOF) of 50% BiPO₄/g-C₃N₄ heterojunction have been further investigated to show the catalytic efficiency of the catalyst. The TON of catalyst is the number of 4-NP molecules that can convert into products using 1 g of catalyst, while TOF is calculated as TON/time. The concentration of 4-NP and dosage of catalyst are 7.18×10^{-5} M, 0.025 g, respectively. The TOF was found to be 0.249 molecules per g.s for 50% BiPO₄/g-C₃N₄ heterojunction using the following equation:⁶⁴

$$TOF = \frac{m_i X x}{100 W t}$$

where m_i is the initial number of moles nitrophenol, X is the conversion of 4-NP, x is the molecular weight of 4-NP, W is the mass of catalyst used in the reaction (g), and t is the reaction time (h).

3.2.1 Effect of initial concentration of 4-NP. The effect of initial concentration of 4-NP on the catalytic efficiency rate using 50% BiPO₄/g-C₃N₄ catalyst was carried by varying the concentration from 10 to 70 mg L⁻¹, and the obtained results as shown in Fig. 8a. Interestingly, 50% BiPO₄/g-C₃N₄ was able to reduce all 4-NP solutions at concentrations from 10 to 70 mg L⁻¹ reflecting the highly efficiency of such a catalyst towards this 4-NP reduction. At lower concentration, superior constant rate was recorded due to the availability of large number of catalytic sites per a given amount of 4-NP moles. And *vice versa*, the higher the concentration, the lower the rate constant (Fig. 8b) fitting pseudo-first-order reaction, due to the high competition of 4-NP molecules on the limited sites. In addition, the number of molecules adsorbed at the surface of the BiPO₄/g-C₃N₄ heterojunction increases with the increase in the concentration of 4-nitrophenol and hence, the surface becomes saturated by 4-nitrophenol molecules. This leads to a decrease in concentration of BH₄⁻ ions approaching the surface of the BiPO₄/g-C₃N₄ heterojunction, hence lowering the rate of hydrogen transfer from BH₄⁻ ion to the 4-nitrophenol molecule. This confirms that the BiPO₄/g-C₃N₄ catalyzed reduction of nitrophenol occurs according to the L-H mechanism.

3.2.2 Effect of catalyst amount. We also investigated the effect of catalyst dosage (0.012–0.05 g) on the reduction 4-NP (70 mg L⁻¹). As seen in Fig. 9, with the increase of catalyst dosage, the removal efficiency of 4-NP improved markedly. With further

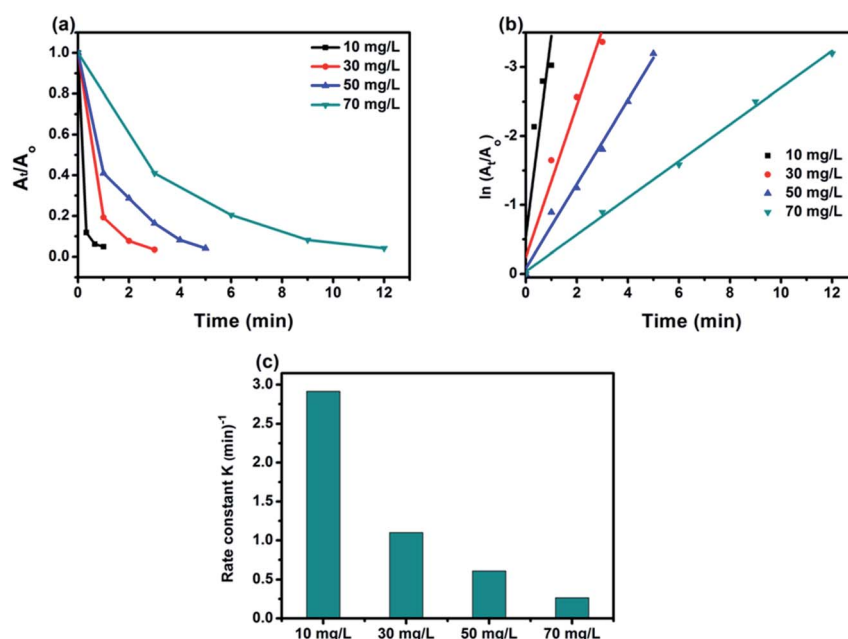


Fig. 8 Effect of initial concentration on reduction rate of 4-NP (a); corresponding linear transform $\ln(A_t/A_0) = f(t)$ of the 4-NP reduction kinetics curves (b); comparison of the rate constant value for the reduction of the 4-NP over initial different concentration (c). ([4-NP] = 10–70 mg L⁻¹, [catalyst] = 0.5 g L⁻¹).



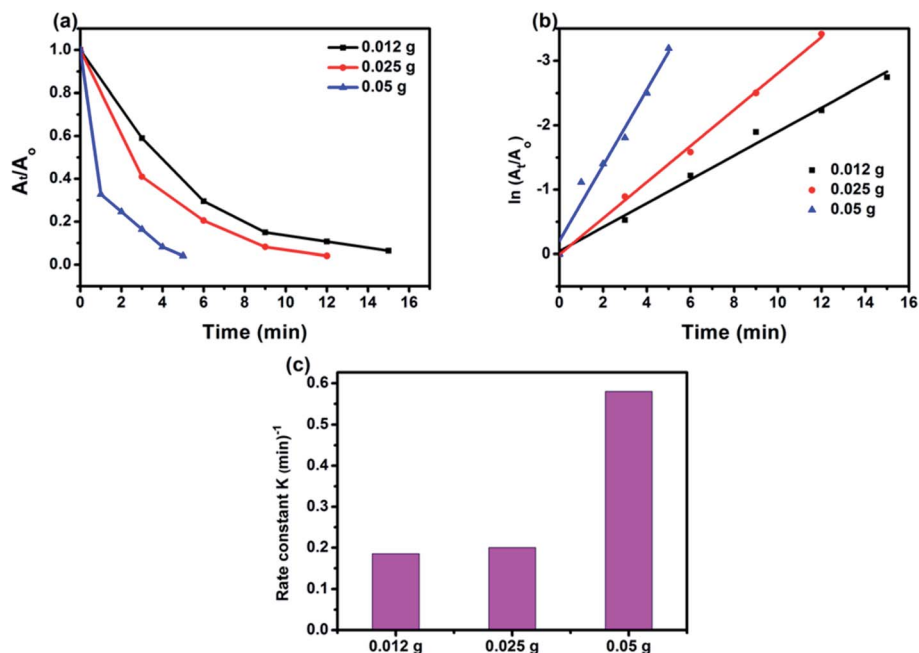
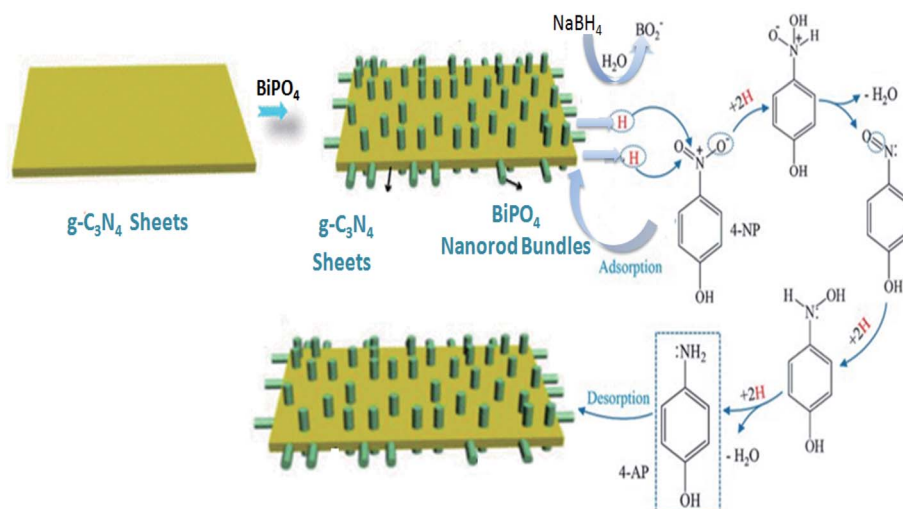


Fig. 9 Reduction rate of 4-NP over catalyst amount (a); corresponding linear transform $\ln(A_t/A_0) = f(t)$ of the 4-NP reduction kinetics curves (b); comparison of the rate constant value for the reduction of the 4-NP over catalyst amount (c). ($[4\text{-NP}] = 70 \text{ mg L}^{-1}$, $[\text{catalyst}] = 0.012\text{--}0.05 \text{ g}$ in 50 mL $[4\text{-NP}]$).

increase to 0.05 g, the reduction efficiency value is 99% within 5 min for 4-NP reduction. The 50% $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ composite (0.05 g) exhibited the highest rate constant of 0.58 min^{-1} compared with 0.025 g (0.26 min^{-1}), and 0.012 g (0.18 min^{-1}) catalyst dosage as shown in Fig. 9c. As above-discussed, the relationship between the active sites on the surface of the catalyst and 4-NP molecules is very significant in terms of reaction kinetics. The use of higher amount of catalyst is required to proceed a faster reduction reaction at high 4-NP concentration.

3.3 Feasible mechanism for 4-NP reduction

Overall, the obtained results showed that the bare $\text{g-C}_3\text{N}_4$ has very low catalytic efficiency towards the reduction of 4-NP from one side, and from another side, $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ catalysts exhibited ultrafast catalytic reductive conversion of 4-NP, compared to bare BiPO_4 . In addition, the higher amount BiPO_4 in the composite, the faster reduction kinetics. These results lead to deduce that BiPO_4 plays an important key especially when it is bonded to $\text{g-C}_3\text{N}_4$ nanosheets. The Langmuir-Hinshelwood (LH) model was proposed as mechanism for our



Scheme 2 The plausible mechanism for the catalytic reduction of 4-NP by $\text{BiPO}_4/\text{g-C}_3\text{N}_4$ composite using NaBH_4 as a reducing agent.

Table 3 The comparison of present work with the recent literature of catalytic reduction of 4-NP to 4-AP using various catalysts with its reaction conditions

Catalyst	Conc. of 4-NP (mol dm ⁻³)	Conc. of NaBH ₄ (mol dm ⁻³)	Amount of catalyst (g L ⁻¹)	Time (min)	Rate constant (min ⁻¹)	Number of cycles	Ref.
g-C ₃ N ₄ /CuS	1.26 × 10 ⁻²	0.5	3.3	50	0.0411	5	68
Au@g-C ₃ N ₄	0.01	0.1	1	10	0.109	10	69
Au/CeO ₂ @g-C ₃ N ₄	0.12 × 10 ⁻³	0.04	0.5	0.6	6.36	5	27
g-C ₃ N ₄ /Bi ₂ S ₃	0.03	0.2 × 10 ⁻³	0.01	60	0.016	5	70
Cu@g-C ₃ N ₄	0.13 × 10 ⁻³	1.64 × 10 ⁻⁵	5	25	1.11	4	71
N/graphene	0.1 × 10 ⁻³	0.04 × 10 ⁻³	1.5	18	0.029	3	72
Ag/g-C ₃ N ₄ /V ₂ O ₅	0.01	0.5	5	60	0.373	3	73
50% BiPO ₄ /g-C ₃ N ₄	7.18 × 10 ⁻⁵	16 × 10 ⁻³	0.5	1	2.914	5	This work

reaction. The plausible mechanism for the 4-NP in the presence of NaBH₄ and BiPO₄/g-C₃N₄ is shown in Scheme 2. NaBH₄ undergoes an ionization and produce BH₄⁻ species, which adsorb on the catalyst surface, followed by BO²⁻ species generation as a result of self-hydrolysis. Simultaneously, adsorbed BH₄⁻ species can induce the transfer of active hydrogen species to form hydride complex. Afterwards, such active hydrogen species react with nitro groups to reduce them into amino groups, as reported in previous studies.^{66,67} The produced aminophenol is released from the surface of the catalyst to the medium. The catalytic reduction activity of BiPO₄/g-C₃N₄ composites were significantly enhanced compared with bare BiPO₄ and g-C₃N₄ due to addition of BiPO₄ nanorod bundles immobilized on the g-C₃N₄ nanosheets to form heterojunction system may form more active surface-hydrogen species, suggesting higher yield of 4-NP reduction. In addition, the conductive g-C₃N₄ can transfer electrons and H⁺ radicals from the adsorbed BH₄⁻ to 4-NP through the catalytic BiPO₄ nanorod bundles more readily, resulting in the production of 4-AP. Furthermore, the heterojunction of BiPO₄/g-C₃N₄ renders all active sites readily accessible that can significantly facilitate the rapid transport and diffusion of hydrogen radicals and electrons, resulting in the enhanced catalytic reduction on 4-NP. The efficiency of BiPO₄/g-C₃N₄ towards the reduction of 4-NP was comparatively listed in Table 3 with the recent previous

reported catalysts for 4-NP reduction. The value of rate constant was calculated to be 2.914 min⁻¹ for 50% BiPO₄/g-C₃N₄ nanocomposite that is much better than other nanocomposites, as shown in Table 3.

3.4 Stability evaluation

Recycling of any heterogeneous catalyst has a significant role, especially for the economic and environmental points of view. The catalytic performance of BiPO₄/g-C₃N₄ catalyst was checked five times in a row towards the reduction of 4-NP under the same conditions (50 mg L⁻¹, and 25 mg 50% BiPO₄/g-C₃N₄). As shown in Fig. 10, the reduction rate was slightly decreased from 98% to 95%, after five times recycling. The results demonstrated that the as-prepared 50% BiPO₄/g-C₃N₄ heterojunction was comparably stable under the studied conditions, which is an advantage of using the catalysts for industrial applications.

4. Conclusions

Highly efficient and stable BiPO₄ nanorod bundles supported on g-C₃N₄ sheets composites with variable BiPO₄ loadings were developed using *ex situ* chemical deposition. The as synthesized composites showed superior catalytic activities towards reduction of 4-NP to 4-AP in wastewater, unlike bare g-C₃N₄ and BiPO₄. It was found that the increasing BiPO₄ loading enhanced catalytic reduction of 4-NP. The 50% BiPO₄/g-C₃N₄ sample performed the highest rate constant for 4-NP reduction which was 728.5 and 56 times superior than the values for g-C₃N₄ and BiPO₄, respectively. The mechanism for high-efficient catalytic activity have been discussed. In addition, the synthesized BiPO₄/g-C₃N₄ heterojunction does not show a significant reduce in its catalytic efficiency during the five times recycling tests. Therefore, the BiPO₄/g-C₃N₄ composite indicated the potential application for removing the hazardous wastes such as *p*-nitrophenol from the environment as a whole.

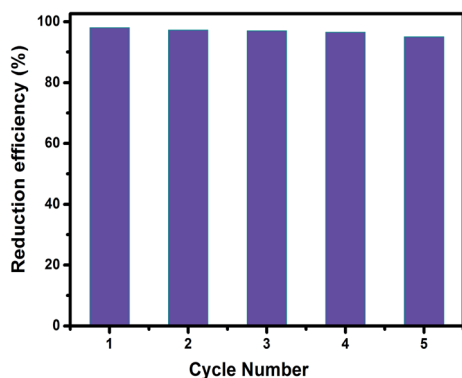


Fig. 10 Cyclic runs for the reduction of 4-NP on 50% BiPO₄/g-C₃N₄ heterojunction in the presence of NaBH₄. ([4-NP] = 50 mg L⁻¹, [catalyst] = 0.5 g L⁻¹).

Conflicts of interest

There are no conflicts to declare.



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