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# Analysis of sodium generation by sodium oxide decomposition on corrosion resistance materials: a new approach towards sodium redox water-splitting cycle†

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In this study, the investigation of materials with corrosion resistance was carried out to prevent side reactions caused by sodium oxide ( $\text{Na}_2\text{O}$ ) in the Na-redox thermochemical water splitting cycle, and essential operational conditions for sodium (Na) generation from  $\text{Na}_2\text{O}$  were also investigated. Thermal desorption spectroscopy and X-ray diffraction techniques at altered conditions were mainly used for the experimental investigation. Numerous types of materials were tested to find materials with high resistance towards corrosion and to understand essential thermal decomposition processes of  $\text{Na}_2\text{O}$ . In addition, under different temperatures and pressure conditions, the thermodynamic calculation of Gibbs free energy was performed to obtain experimental results. As a result, a Ti alloy showed significant resistance towards the corrosive reaction by  $\text{Na}_2\text{O}$ . The obtained experimental and simulated results support the direct decomposition of  $\text{Na}_2\text{O}$  to form Na and  $\text{O}_2$  below 600 °C under low partial pressure conditions. The optimized conditions for Na generation with the Ti alloy sample can be used for low temperature water splitting.

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## Introduction

Excessive utilization of non-renewable sources and their hazardous by-products lead to the pollution and global warming problems. The search for alternative energy sources has become a necessity at present. The most important requirements for alternative energy sources include energy efficiency at the large scale and clean by-products.<sup>1–4</sup> For decades, researchers have devoted extensive efforts to find a suitable energy source. As of now, hydrogen ( $\text{H}_2$ ) is the best energy source of renewable energy such as natural energy in this regard. The advantages of  $\text{H}_2$  as an energy source are security, reduction in environmental problems, clean energy carrier, abundant element on the Earth, and the potential to fulfil the society's energy requirement.<sup>5–8</sup>

Currently,  $\text{H}_2$  is produced at the industrial level from non-renewable sources such as methane steam reforming,<sup>9,10</sup> gasification,<sup>11,12</sup> and partial oxidation.<sup>13</sup>  $\text{H}_2$  production from non-renewable sources generates some adverse by-products such as  $\text{CO}_2$  and other harmful gases. On the other hand,  $\text{H}_2$  is a potential green energy medium when  $\text{H}_2$  is produced from

renewable sources.<sup>14–17</sup> The hydrogen production *via* water electrolysis combined with power generation by natural energy such as solar and wind is an attractive method because the related technology has already been established. However, it is difficult to obtain benefits of scale-up because of the 2-dimensional electrode reactions. Thermochemical water splitting is one of the promising and encouraging methods for the large scale  $\text{H}_2$  production,<sup>18,19</sup> where larger merit can be expected by scale up as chemical plants. The main hurdle in direct water splitting includes the requirement of high temperatures around 4000 °C and efforts, which is not suitable from the economic point of view. Thermochemical water-splitting cycles are attractive methods. In this method, a cycle composed of chemical reactions takes place, and at the end, all chemicals regenerate again except  $\text{H}_2\text{O}$ . Various types of thermochemical cycles have been reported: tin oxide ( $\text{SnO}_2/\text{SnO}$ ),<sup>20,21</sup> cobalt oxide,<sup>22,23</sup> iron oxide ( $\text{Fe}_3\text{O}_4/\text{FeO}$ ),<sup>24,25</sup> zinc oxide ( $\text{ZnO}/\text{Zn}$ ),<sup>26,27</sup> germanium oxide ( $\text{GeO}_2/\text{GeO}$ ),<sup>28</sup> cerium-based oxides,<sup>29–31</sup> sulfur-iodine,<sup>32,33</sup> sodium manganese mixed ferrite,<sup>34</sup> and sodium-redox (Na-redox) cycles.<sup>35–37</sup> Among the above-mentioned works, specific temperature and pressure conditions are required due to kinetic and thermodynamics reasons.<sup>38–40</sup> For all the conventional cycles, high temperature heat energy (800–1500 °C) is required, and thereby utilizable heat sources are limited to solar heat plant with high concentration (tower-type) and nuclear high temperature gas-cooled reactors.

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The Na-redox cycle is composed of three types of reactions: hydrogen generation by reaction between NaOH and Na, Na metal separation by the thermal decomposition of sodium oxide (Na<sub>2</sub>O) to form sodium peroxide (Na<sub>2</sub>O<sub>2</sub>), and oxygen generation (O<sub>2</sub>) by the hydrolysis of Na<sub>2</sub>O<sub>2</sub> (Table 1). H<sub>2</sub> and O<sub>2</sub> generation reactions are not difficult thermodynamically and can be operated at moderate temperature and atmospheric pressure conditions.<sup>41</sup> The challenging reaction is the decomposition of sodium oxide (Na<sub>2</sub>O) due to its high thermodynamic stability. In general, the thermodynamics of chemical reactions are expressed using Gibbs free energy change  $\Delta G$ , as shown in eqn (1):

$$\Delta G = \Delta H - T\Delta S \quad (1)$$

where  $\Delta H$ ,  $T$ , and  $\Delta S$  are the enthalpy change, temperature, and entropy change of chemical reactions, respectively. The enhancement of  $\Delta S$  decreases the required temperature for the reaction progress. In previous studies, thermodynamic equilibrium was shifted by controlling the partial pressure of Na vapor. As a result, the Na generation was found below 500 °C. Thus, the feasibility of the Na-redox cycle as a potential hydrogen production method operated below 500 °C has been demonstrated. If the low temperature operation of thermochemical cycles could be realized, numerous types of heat sources such as solar heat and unused heat can be utilized for hydrogen production. In addition, effective heat storage materials such as molten salts (nitrates and carbonates) can also be used for stable and long running reaction. On the other hand, the details of the reaction process are not understood completely because of the strong corrosion property of Na<sub>2</sub>O and/or Na<sub>2</sub>O<sub>2</sub>, which is the most critical issue of the Na-redox cycle and limits the development for practical use.

In this study, the target reaction is the decomposition of Na<sub>2</sub>O to generate Na<sub>2</sub>O<sub>2</sub> and Na at a low temperature in the cycle, as shown in Table 1. The Na-redox water splitting cycle can be a potential hydrogen production technique after understanding the essential reaction processes and establishing control methods. The main hurdle is to handle the Na oxides at high temperatures because of their high corrosive properties as explained above. Na<sub>2</sub>O corrodes base materials as side reactions, and undesirable stable by-products are formed. Therefore, one of the objectives in this study is finding a suitable base material for the Na<sub>2</sub>O decomposition step. By using

the attractive base materials, the essential thermal decomposition properties of Na<sub>2</sub>O were investigated. Furthermore, the experimental results were discussed with thermodynamic calculations.

## Materials and methods

The starting material Na<sub>2</sub>O containing Na<sub>2</sub>O<sub>2</sub> was purchased from Sigma-Aldrich, and its composition was 80% Na<sub>2</sub>O + 20% Na<sub>2</sub>O<sub>2</sub>. Various types of base materials were used for the experimental testing of the Na<sub>2</sub>O decomposition and corrosion resistance properties: graphite powder, boron nitride (BN), single crystal Si, aluminium oxide (Al<sub>2</sub>O<sub>3</sub>), aluminium nitride (AlN), graphite polished, and Ti Alloy (0.8% Ni, 0.4% Mo, 98.8% Ti). All the sample materials were used as received without any pre-treatments. Na<sub>2</sub>O with each base material was heated and analyzed under high vacuum conditions (10<sup>-11</sup> MPa) by thermal desorption spectroscopy (TDS, ESCO EMD-WA 1000 s d<sup>-1</sup>), which is specially designed to detect metal vapor such as Na. The base materials are put on a quartz boat as a sample holder, and then the Na<sub>2</sub>O powder was put on the base materials in a glove box (Miwa MFG) filled with purified Ar (>99.9999%). The above samples were moved into the TDS chamber without exposing them to air by a transfer vessel. After the TDS experiment, all the samples were identified by the X-ray diffraction (XRD) technique (Rigaku-RINT 2500 V, Cu K $\alpha$  radiation:  $\lambda = 1.5408 \text{ \AA}$ ) to understand the decomposition properties of Na<sub>2</sub>O and obtain the information of side reactions, where the samples were covered by a polyimide sheet (Du Pont-Toray Co., Ltd., Kapton) to avoid the reaction with oxygen and moisture. All the above-mentioned processes were carried out without exposing the samples into air to avoid the influence of oxygen and moisture. Thermodynamic calculations were performed by the MALT software to expect possible reactions during the heat treatment of Na<sub>2</sub>O with the base materials. Here, the results for Ti alloy, AlN, graphite powder are shown and discussed as representative cases in the manuscript. The results of other samples' data are explained in ESI.<sup>†</sup>

## Results and discussion

### Base sample comparison for corrosion resistance

For the TDS experiments, to minimize the contribution of gases adsorption by the quartz boat and base material, baking and

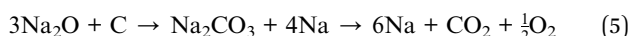
**Table 1** Summarised table for the Na-redox cycle and modified one. Where,  $\Delta H$ ,  $\Delta S$ , and  $\Delta G$  are representation of the enthalpy change, entropy change, and Gibbs free energy respectively at ambient conditions

|         | Equation   | $\Delta H$ (kJ) | $\Delta S$ (J mol <sup>-1</sup> ) | $\Delta G$ (kJ mol <sup>-1</sup> ) | Eq. no |
|---------|--|-----------------|-----------------------------------|------------------------------------|--------|
| Cycle 1 | 2NaOH(s) + 2Na(l) $\rightarrow$ 2Na <sub>2</sub> O(s) + H <sub>2</sub> (g)                                     | 11              | 36                                | 0.27                               | (2)    |
|         | 2Na <sub>2</sub> O(s) $\rightarrow$ Na <sub>2</sub> O <sub>2</sub> (s) + 2Na(g)                                | 540             | 250                               | 465                                | (3)    |
|         | Na <sub>2</sub> O <sub>2</sub> (s) + H <sub>2</sub> O(l) $\rightarrow$ 2NaOH(s) + $\frac{1}{2}$ O <sub>2</sub> | -55             | 66                                | -75                                | (4)    |
|         | H <sub>2</sub> O(l) $\rightarrow$ H <sub>2</sub> (g) + $\frac{1}{2}$ O <sub>2</sub>                            |                 |                                   |                                    |        |
| Cycle 2 | 2NaOH(s) + 2Na(l) $\rightarrow$ 2Na <sub>2</sub> O(s) + H <sub>2</sub> (g)                                     | 11              | 36                                | 0.27                               | (2)    |
|         | 2Na <sub>2</sub> O(s) $\rightarrow$ 4Na(s) + O <sub>2</sub> (g)  | 830             | 260                               | 752                                | (7)    |
|         | 2Na(s) + 2H <sub>2</sub> O(g) $\rightarrow$ 2NaOH(s) + H <sub>2</sub> (g)                                      | -368            | -221                              | -302                               | (4)    |
|         | 2H <sub>2</sub> O(g) $\rightarrow$ 2H <sub>2</sub> (g) + O <sub>2</sub> (g)                                    |                 |                                   |                                    |        |

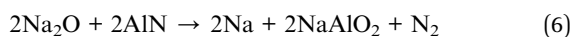


blank tests were performed before the sample measurements. Observed TDS data for the baking and blank test does not show any peak deflection from the regular pattern (see Fig. S1–S3†).

Temperature calibration was also performed in the blank test experiment to know the exact temperature for Na<sub>2</sub>O analyses (see Fig. S4†). The difference between the targeted temperature and real temperature is not small and mainly originated from the quartz boat and base material layer (see Fig. S5†). For precise comparison, the TDS data of graphite, AlN, and Ti alloy are shown in Fig. 1. Other tested materials and their corresponding data have been summarized in the ESI (see Fig. S6–S11†). Fig. 1(a) shows data for  $m/z = 44$  spectra, which correspond to carbon dioxide (CO<sub>2</sub>) gas. It is determined from the spectra that graphite has the highest intensity peak of CO<sub>2</sub> gas compared with the results for AlN and Ti alloy samples. The occurrence of the CO<sub>2</sub> peak is due to the formation of Na<sub>2</sub>CO<sub>3</sub> as an intermediate. At a high temperature, clear peaks corresponding to Na and CO<sub>2</sub> were found. Based on the CO<sub>2</sub> peak and MALT calculations (see Fig. S12†), the formation of Na<sub>2</sub>CO<sub>3</sub> is expected, and subsequently Na<sub>2</sub>CO<sub>3</sub> decomposes into Na and CO<sub>2</sub> at the targeted temperature of about 650 °C, as shown in eqn (5).



The base of AlN also shows a side reaction with Na<sub>2</sub>O, which is plotted in Fig. 1(b). The spectrum of  $m/z = 28$  corresponding to nitrogen (N<sub>2</sub>) gas shows a clear peak. The appearance of this peak is due to the desorption of N<sub>2</sub> molecules from AlN around 450 °C (targeted temperature), suggesting that Na<sub>2</sub>O reacts with AlN and produces NaAlO<sub>2</sub> and N<sub>2</sub> molecules at about 450 °C according to eqn (6).



The spectra in Fig. 1(c) show the peak of  $m/z = 23$ , which is Na. Except for the Ti alloy, other samples are showing a peak of Na due to the side reactions. All other samples (Si, polished

graphite, boron nitride, and so on) also showed corrosion reactions of Na<sub>2</sub>O at the high temperature. Among all the tested samples from TDS data, it was concluded that Ti alloy was the best material to exhibit corrosion resistance properties. The Ti alloy would also form a very thin, stable, and uniform oxide layer on the surface. It is expected that this oxide layer prevents the inside of Ti from any side reactions at any conditions. The corrosion resistance of a metal surface always depends on the stability of the oxide layer. The amount of Ni and Mo within the Ti alloy is supportive to make a stable oxide layer and enhance the corrosion resistance properties.<sup>42</sup> Therefore, the Ti alloy does not show strong corrosion towards the Na<sub>2</sub>O decomposition process. Fig. 2 shows the TDS spectra for Na<sub>2</sub>O on the Ti alloy base heated up to 800 °C (calibrated temperature: ~600 °C), and the major profiles in the observed data correspond to  $m/z = 32$  and 23, as plotted in Fig. 2, where all the possible release gases were observed in the experiments (see Fig. S7†). These gases correspond to O<sub>2</sub> and Na, respectively. Comparing the O<sub>2</sub> and Na peaks, the first peak is observed almost at 25 min in the O<sub>2</sub> spectra. Since all the TDS experiments performed at the heating rate of 5 °C min<sup>-1</sup>, the targeted temperature at this point is 125 °C (calibrated temperature: ~75 °C).

Because of the low temperature, the peak almost at 25 min is not related to the decomposition of Na<sub>2</sub>O thermodynamically. This peak would be due to the release of the O<sub>2</sub> molecule, which is absorbed on the Na<sub>2</sub>O sample. In fact, the TDS data of other samples also shows (see Fig. S6–S11†) the same O<sub>2</sub> peak at the same temperature region. The second peak is around after 90 min and the calibrated temperature here is 325 °C. In the comparison of both O<sub>2</sub> and Na spectra, only O<sub>2</sub> spectra show the high intensity peak. By observing the information obtained from software calculation (see Fig. 3) and XRD data (Fig. 5), this peak appears due to the decomposition of Na<sub>2</sub>O<sub>2</sub>. Na<sub>2</sub>O<sub>2</sub> is included as impurity in the as-purchased Na<sub>2</sub>O sample, and it decomposes into Na<sub>2</sub>O and O<sub>2</sub>. The last peak that is observed in both spectra after 160 min and the corresponding calibrated temperature at this point is 600 °C. At this point, both the spectra of Na and O<sub>2</sub> show a high intensity synchronized peak. By the observation, it can be concluded that the appearance of

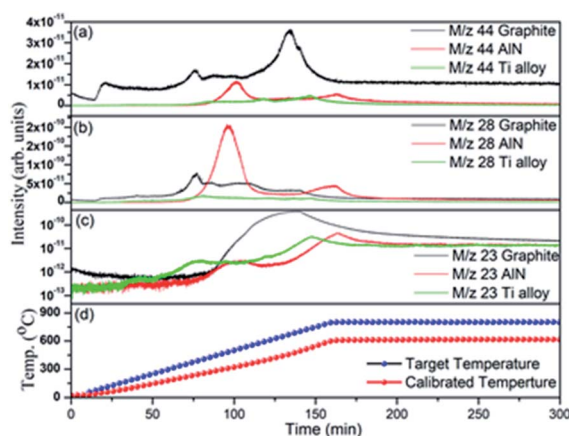


Fig. 1 Comparison TDS data for graphite, AlN, and Ti alloy of (a)  $m/z = 44$ , (b)  $m/z = 28$ , (c)  $m/z = 23$ , and (d) targeted and calibrated temperature profiles.

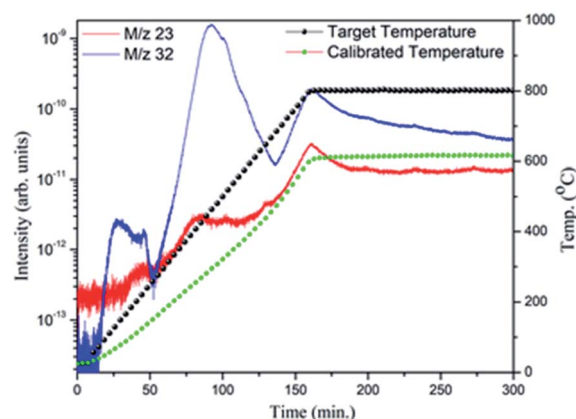


Fig. 2 TDS spectra of Na<sub>2</sub>O with the Ti alloy base performed at 800 °C.



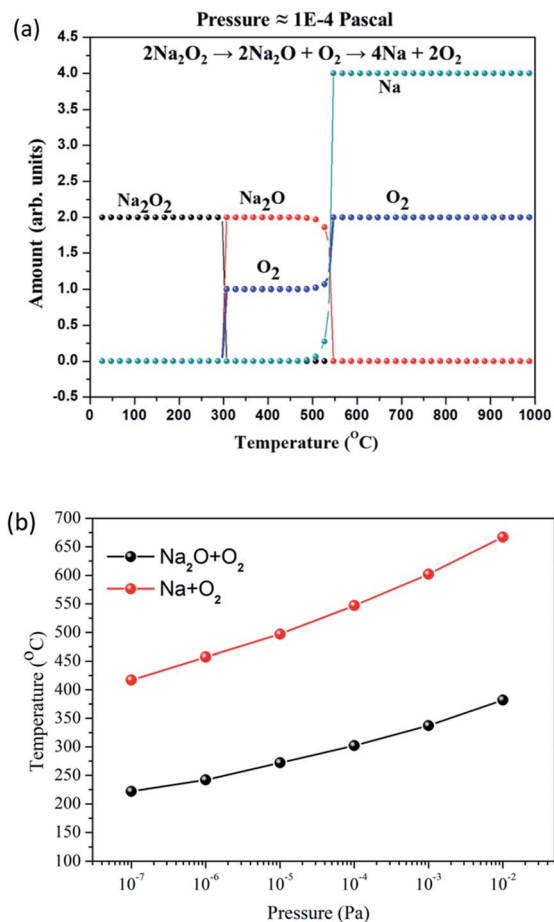
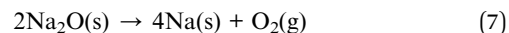


Fig. 3 Graphs based on the MALT software calculation: (a) temperature information for  $\text{Na}_2\text{O}$  and  $\text{Na}_2\text{O}_2$  decomposition under static pressure conditions and (b) temperature information for  $\text{Na}_2\text{O}$  and  $\text{Na}_2\text{O}_2$  decomposition for variable pressure conditions.

this is due to the decomposition of  $\text{Na}_2\text{O}$  into  $\text{Na}$  and  $\text{O}_2$ . The software calculation (see Fig. 3(a)) at this temperature also suggests the decomposition of  $\text{Na}_2\text{O}$  at this point.

Fig. 3(b) shows pressure dependence of the decomposition temperatures of  $\text{Na}_2\text{O}$  and  $\text{Na}_2\text{O}_2$ . From the results, the alteration in the partial pressure condition leads to the decrement in the decomposition temperature. The temperature difference for the  $\text{Na}_2\text{O}$  decomposition can be achieved nearby 250 °C just by reducing the pressure from the atmosphere to  $1 \times 10^{-2}$  Pa. Likewise, the  $\text{Na}_2\text{O}_2$  decomposition also proceeds at a low temperature in low partial pressure conditions. From the MALT software calculations, it is found that the used  $\text{Na}_2\text{O}$  sample decomposes in two stages.<sup>43,44</sup> First stage, the pre-existing  $\text{Na}_2\text{O}_2$  (in the  $\text{Na}_2\text{O}$  sample) decomposes into  $\text{Na}_2\text{O}$  and  $\text{O}_2$  at a low temperature. In the second stage,  $\text{Na}_2\text{O}$  decomposes into  $\text{Na}$  and  $\text{O}_2$ . In the case of  $1 \times 10^{-4}$  Pa pressure,  $\text{Na}_2\text{O}_2$  decomposes at about 300 °C and  $\text{Na}_2\text{O}$  decomposes at around 540 °C into  $\text{Na}$  and  $\text{O}_2$  according to TDS results. Owing to the above facts, it is difficult to control the decomposition of  $\text{Na}_2\text{O}$  into  $\text{Na}_2\text{O}_2$  and  $\text{Na}$  in our experimental conditions (Table 1; eqn (3)). Therefore, the reactions of the water-splitting cycle should be restructured,

as shown in Table 1. In this water-splitting cycle,  $\text{Na}_2\text{O}$  is directly decomposed into  $\text{Na}$  and  $\text{O}_2$  according to eqn (7). To complete the cycle,  $\text{Na}$  generated by reaction (7) is divided to half amount and used for reactions (2) and (4) to produce  $\text{H}_2$ .



### Temperature optimization for Na generation

For the temperature optimization of the Na generation, the TDS experiments were performed at different temperature conditions. The TDS measurements were performed for  $\text{Na}_2\text{O}$  on the Ti alloy base from 400 to 800 °C at the difference of 100 °C. After each TDS experiment, the same Ti alloy base sample heated for 900 °C without  $\text{Na}_2\text{O}$  in the TDS chamber (see detailed procedure in Fig. S13†). This heating process of the Ti alloy removes the background noise (impurities) of the previous experiment. All the TDS results for  $\text{Na}$  and  $\text{O}_2$  obtained at 400–800 °C are shown in Fig. 4. The spectra of  $\text{Na}$  and  $\text{O}_2$  for all the experiments are similar, suggesting that the Ti alloy possesses high stability for the corrosion of  $\text{Na}_2\text{O}$ . The  $\text{O}_2$  desorption due to the  $\text{Na}_2\text{O}_2$  decomposition was observed below 400 °C (see Fig. S14†) for all the spectra because the commercial  $\text{Na}_2\text{O}$  including  $\text{Na}_2\text{O}_2$  as an impurity was loaded for each experiment. The Na generation originated in the direct decomposition of  $\text{Na}_2\text{O}$  was found from 700 °C (calibrated temperature: ~540 °C) even though the reaction yield would be low compared with that at 800 °C (calibrated temperature: 600 °C).

The XRD experiments have been performed for the  $\text{Na}_2\text{O}$  sample before and after the above TDS experiments, and the results are shown in Fig. 5. From the observed XRD patterns, it can be clarified that the peaks corresponding to  $\text{Na}_2\text{O}_2$  in the initial sample (80%  $\text{Na}_2\text{O}$  + 20%  $\text{Na}_2\text{O}_2$ ) disappeared between 400 °C and 500 °C. That indicates that  $\text{Na}_2\text{O}_2$  is decomposed in this temperature region. This result is consistent with the results of TDS and thermodynamic calculation obtained by the MALT software. The intensity of peaks corresponding to the  $\text{Na}_2\text{O}$  phase was slightly lowered. When the direct  $\text{Na}_2\text{O}$  decomposition proceeded, the products of  $\text{Na}$  and  $\text{O}_2$  were

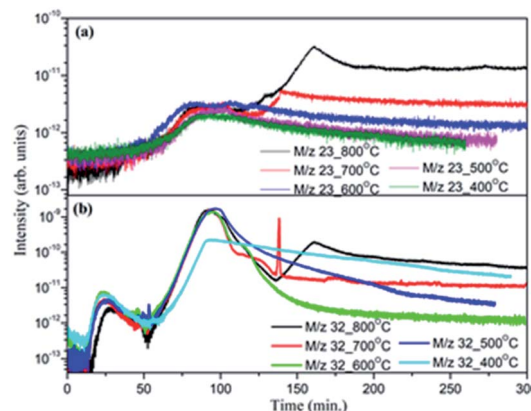


Fig. 4 Comparison of (a) Na peak ( $m/z = 23$ ) and (b)  $\text{O}_2$  peak ( $m/z = 32$ ) from TDS data obtained at the difference of 100 °C from 400–800 °C.





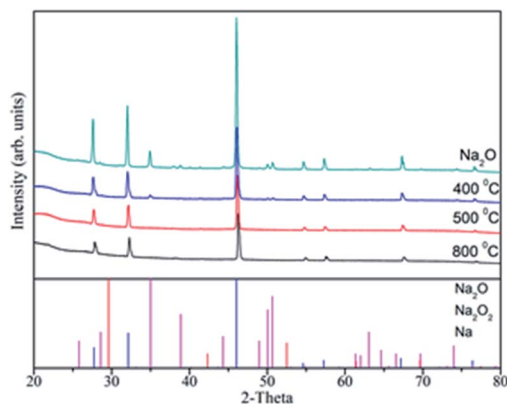


Fig. 5 XRD data for  $\text{Na}_2\text{O}$  at the various targeted temperature over the Ti alloy base.

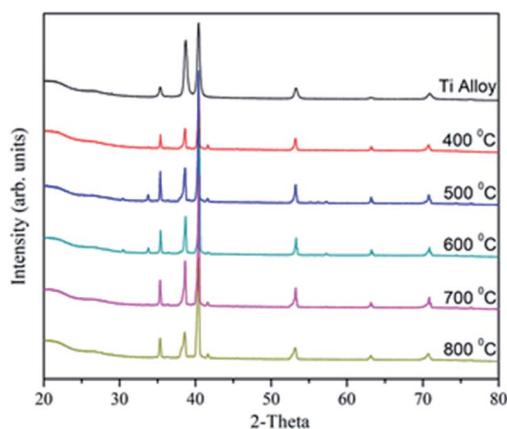


Fig. 6 XRD data for the Ti alloy after different targeted temperature TDS experiments.

released as gaseous phases in the TDS chamber. Moreover, the remaining material after the TDS experiments should be the undecomposed  $\text{Na}_2\text{O}$ . This is the reason that no peak was observed corresponding to Na even at 800 °C (calibrated temperature: 600 °C), which is enough temperature for the  $\text{Na}_2\text{O}$  decomposition as discussed above. Fig. 6 shows the XRD patterns of Ti alloy after the TDS experiments at 400–800 °C. The diffraction peaks assigned to the Ti alloy changed to be sharp and intense, causing annealing effects. On the other hand, the corrosion phases were not observed at all the temperature. Thus, these results also suggest that the Ti alloy is recognized as the promising corrosion-resistance material for the  $\text{Na}_2\text{O}$  decomposition process at high temperatures.

## Conclusion

In this study, finding corrosion resistance materials and investigation of the essential properties of the  $\text{Na}_2\text{O}$  decomposition reaction in the Na-redox cycle were carried out. For the  $\text{Na}_2\text{O}$  decomposition, different base materials such as graphite powder, BN, single-crystal Si,  $\text{Al}_2\text{O}_3$ , AlN, graphite polished, Ti Alloy have been tested for the corrosion resistance. From the

overall observation of all the materials, the Ti alloy shows outstanding corrosion resistance properties with  $\text{Na}_2\text{O}$ . Because a stable and uniform oxide layer on the Ti alloy appears to prevent the corrosion. Furthermore, the TDS experiments with different temperatures were performed for  $\text{Na}_2\text{O}$  on the Ti alloy base. The experiential results of TDS give details that  $\text{Na}_2\text{O}$  decomposes from around 600 °C in the high vacuum condition. TDS results also resemble the data obtained by the XRD and the theoretical calculation by the MALT software. This systematic study would be beneficial to establish the hydrogen production system based on the one for Na-redox thermochemical water splitting cycle.

## Conflicts of interest

There are no conflicts to declare.

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