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# Role of oxide support in Ni based catalysts for CO<sub>2</sub> methanation

 Ye Hwan Lee,<sup>a</sup> Jeong Yoon Ahn,<sup>a</sup> Dinh Duc Nguyen,<sup>b</sup>  Soon Woong Chang,<sup>b</sup> Sung Su Kim <sup>\*b</sup> and Sang Moon Lee <sup>\*b</sup>

The CO<sub>2</sub> methanation reaction of reduced and unreduced Ni based CeO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, TiO<sub>2</sub> and Y<sub>2</sub>O<sub>3</sub> supported catalysts was investigated. The Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalysts exhibited similar CO<sub>2</sub> conversions at all reaction temperatures. The catalysts were studied by X-ray diffraction (XRD), H<sub>2</sub> chemisorption, H<sub>2</sub> temperature-programmed reduction (TPR), and *in situ* diffuse reflection infrared Fourier transform spectroscopy (DRIFTS); the results suggested that the reducibility of both metal and support at low temperature, strong metal support interaction and small Ni particle size are important factors for low-temperature CO<sub>2</sub> methanation. Based on the DRIFT studies, the difference in the CO<sub>2</sub> adsorption properties and reaction pathway depending on the reduced and unreduced Ni based supported catalysts was discussed.

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## 1 Introduction

The CO<sub>2</sub> methanation reaction is important in industry because it produces methane and removes CO<sub>2</sub>. Among the CO<sub>2</sub> conversion reactions using catalysts, the CO<sub>2</sub> methanation reaction or so-called Sabatier reaction, is the most advantageous reaction.<sup>1</sup> CO<sub>2</sub> methanation is a simple reaction which makes it possible to integrate the transformation of methane into biogas through the process of anaerobic digestion to generate power, or into other industrial plants with CO<sub>2</sub>-rich exhaust gases.<sup>2</sup> CO<sub>2</sub> methanation is a highly exothermic reaction, and it can produce the heat. CO<sub>2</sub> methanation also has a large equilibrium constant at a lower temperature, so that methane can be generated with a high conversion at lower temperatures; however, the conversion was still low due to kinetic limitation and catalyst performance. Catalytic performance is dependent on various parameters, such as the kinds of support, preparation methods, and the addition of promoters.<sup>3</sup> Especially, many researchers have studied the effect of support to enhance the performance of the CO<sub>2</sub> methanation reaction. The support has a significant effect on the redox properties, metal-support interaction, metal dispersion and adsorption properties.<sup>4–6</sup> Many efforts have been made to develop Ni-based supported catalysts for low-temperature CO<sub>2</sub> methanation.<sup>7–11</sup> The CO<sub>2</sub> methanation activity has been investigated using Ni-based catalysts deposited on various supports such as SiO<sub>2</sub>,<sup>12</sup> α-Al<sub>2</sub>O<sub>3</sub>,<sup>13</sup> MgO,<sup>14</sup> ZrO<sub>2</sub>,<sup>15</sup> Y<sub>2</sub>O<sub>3</sub>,<sup>16</sup> CeO<sub>2</sub>,<sup>17</sup> TiO<sub>2</sub>,<sup>18</sup> β-zeolite,<sup>19</sup> and

their composite supports such as Ni/CeO<sub>2</sub>-ZrO<sub>2</sub>,<sup>20</sup> Ni/γ-Al<sub>2</sub>O<sub>3</sub>-ZrO<sub>2</sub>-TiO<sub>2</sub>-CeO<sub>2</sub>,<sup>21</sup> Ni/CeO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub>,<sup>22</sup> Frontera *et al.*<sup>23</sup> reported that the catalytic activity strongly effects on the characteristics of the support for Ni based catalysts. Le *et al.*<sup>24</sup> founded that the high Ni dispersion and strong CO<sub>2</sub> adsorption plays an important role in the high CO and CO<sub>2</sub> methanation activities for Ni/CeO<sub>2</sub> catalyst. Tada *et al.*<sup>13</sup> reported that the Ni/CeO<sub>2</sub> catalyst showed a high conversion compare to Ni/α-Al<sub>2</sub>O<sub>3</sub>. A large CO<sub>2</sub> adsorption amount and high CO<sub>2</sub> reduction ability could result in high CO<sub>2</sub> conversion at a low temperature using a Ni/CeO<sub>2</sub> catalyst. Muroyama *et al.*<sup>16</sup> found that the Ni/Y<sub>2</sub>O<sub>3</sub> catalyst exhibited a high CO<sub>2</sub> conversion and CH<sub>4</sub> yield compare to Ni/CeO<sub>2</sub>, Ni/Al<sub>2</sub>O<sub>3</sub>, Ni/ZrO<sub>2</sub>, Ni/La<sub>2</sub>O<sub>3</sub> and Ni/Sm<sub>2</sub>O<sub>3</sub> catalysts. They expected that the promotion of the decomposition of formate species would lead to high catalytic activity. Abello *et al.*<sup>25</sup> have reported that the Ni-Al-activated catalyst prepared by a co-precipitation method exhibited a high CO<sub>2</sub> conversion and stability at a high space velocity and highly loaded and dispersed small Ni nanoparticles (*ca.* 6 nm) dispersed over NiO-alumina by partial reduction of the mixed oxide. Vogt *et al.*<sup>26</sup> reported that CO<sub>2</sub> hydrogenation over Ni is considered to follow two steps; direct dissociation and H-mediated and they proved how structure sensitivity effects the mechanism of CO<sub>2</sub> hydrogenation over Ni/SiO<sub>2</sub>. Zeolite and TiO<sub>2</sub> supports are also used as Ni catalysts during CO<sub>2</sub> methanation. Liu *et al.*<sup>18</sup> reported that the 15 wt% Ni/TiO<sub>2</sub> catalyst prepared by deposition-precipitation method showed excellent CO<sub>2</sub> methanation activity (conversion: 96%; selectivity: 99%) at 260 °C. The good dispersion of Ni particles with high exposure of active sites, which may lead to enhanced exposure of active sites that facilitate the generation of surface-dissociated hydrogens. Recent reports have shown substantial improvement in the CO<sub>2</sub>

<sup>a</sup>Department of Environmental Energy Engineering, Graduate School of Kyonggi University, 94-6 San, Iui-dong, Youngtong-ku, Suwon-si, Gyeonggi-do, 442-760, Korea

<sup>b</sup>Department of Environmental Energy Engineering, Kyonggi University, 94-6 San, Iui-dong, Youngtong-ku, Suwon-si, Gyeonggi-do, 442-760, Korea. E-mail: sskim@kyonggi.ac.kr; leesangm@kyonggi.ac.kr


methanation rate by changing the support's properties and/or by the addition of promoters such as Ni–W–Mg,<sup>27</sup> Ni–La,<sup>28</sup> Ni–Cu or Fe/Al<sub>2</sub>O<sub>3</sub> (ref. 29) catalysts. Many studies have examined the performance of highly active Ni-based catalysts supported on various metal oxides for CO<sub>2</sub> methanation at low temperature. However, the effect of metal–support interaction and role of support on catalytic performances during Ni based CO<sub>2</sub> methanation reaction is yet to be unraveled. In this work, CeO<sub>2</sub>, Y<sub>2</sub>O<sub>3</sub>, TiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub> supports were selected as representative Ni catalysts and prepared by wet impregnation method. We investigated the catalytic activities of all catalysts for CO<sub>2</sub> methanation and examined their physicochemical properties. Moreover, the adsorbed species on the reduced and unreduced catalyst surface were compared.

## 2 Experimental

### 2.1 Preparation of catalysts

Ni-Based catalysts supported on CeO<sub>2</sub> (Sigma Aldrich Co.),  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> (Alfa Aesar Co.), Y<sub>2</sub>O<sub>3</sub> (Sigma Aldrich Co., St. Louis, MO, USA) and TiO<sub>2</sub> (G-5, Cristal Global Co.) were prepared by a wet impregnation method. Nickel nitrate hexahydrate (Ni(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O; Sigma Aldrich Co.) was dissolved in distilled water at 80 °C. After impregnation, the moisture was removed at 80 °C using a rotary vacuum evaporator and then dried at 103 °C oven for 12 h. The samples were calcined at 400 °C for 2 h and then samples were reduced at 420 °C for 2 h with a 30% H<sub>2</sub>/N<sub>2</sub>. The obtained samples were grounded and sieved using a 40–50 mesh. Ni loading was fixed at 10 wt%.

### 2.2 Catalytic activity

The CO<sub>2</sub> methanation experimental apparatus consisted of a continuous flow-type fixed-bed reactor comprising a quartz tube (inner diameter: 8 mm; height: 600 mm) and a catalytic bed (Fig. 1). To measure the gas temperature, another K-type thermocouple was installed at the top of the catalytic bed. Prior to the experiment, the catalysts were pretreated at 420 °C for 1 h with a 30% H<sub>2</sub>/N<sub>2</sub> at a flow rate of 100 cm<sup>3</sup> min<sup>-1</sup>. The feed gases comprised 16.67% CO<sub>2</sub>, 66.66% H<sub>2</sub> and 16.67% N<sub>2</sub>. The total flow through the reactor was 120 cm<sup>3</sup> min<sup>-1</sup> and

a space velocity of 14 400 l kg<sup>-1</sup> h<sup>-1</sup> was obtained. The outlet gas-supply pipe was made of stainless steel and wrapped with a heating band set at 180 °C to prevent water condensation. The concentrations of the reactants and products were measured as follows: the inlet and outlet gas concentration were analyzed using a gas chromatograph with a thermal conductivity detector (GOW-MAC, series 580). The CO<sub>2</sub> conversion and yield can be calculated as the following:

$$\text{CO}_2 \text{ conversion (\%)} = \frac{(\text{CO}_{2\text{in}} - \text{CO}_{2\text{out}})}{\text{CO}_{2\text{in}}} \times 100$$

$$\text{CH}_4 \text{ yield (\%)} = \frac{\text{CH}_{4\text{out}}}{\text{CO}_{2\text{in}}} \times 100$$

$$\text{CO yield (\%)} = \frac{\text{CO}_{\text{out}}}{\text{CO}_{2\text{in}}} \times 100$$

### 2.3 Characterization

The surface areas of Ni based catalysts were obtained by the Brunauer–Emmett–Teller (BET) equation using an ASAP 2010 instrument (Micrometrics). The Ni dispersion was evaluated by H<sub>2</sub> chemisorption at 35 °C using a Micrometrics ASAP 2020 instrument (Micrometrics). All catalysts were reduced under a H<sub>2</sub> airflow at 300 °C for 1 h and then cooled into 35 °C. X-Ray diffraction (XRD) analysis was measured on an X'Pert PRO MRD instrument (PANalytical) with a Cu K $\alpha$  ( $\lambda = 1.5056 \text{ \AA}$ ) radiation. Field emission-transmission electron microscope (TEM) analysis was carried out on a JEM-2100 F (JEOL) microscope (200 keV voltage). All samples were prepared by suspending an ultrasonicated catalyst powder in ethanol and placing the suspension on a Cu grid. For H<sub>2</sub> chemisorption analysis, the catalysts were activated with 10% H<sub>2</sub> at 300 °C for 0.5 h and then cooled to 50 °C and saturated with H<sub>2</sub> pulses. The temperature-programmed reduction (TPR) of H<sub>2</sub> was measured by 10% H<sub>2</sub>/Ar using 0.3 g of the catalyst at a total flow rate of 50 cm<sup>3</sup> min<sup>-1</sup>. Before the H<sub>2</sub> TPR measurement, the catalyst was pretreated in a flow of air at 400 °C for 0.5 h, followed by cooling to 50 °C. The catalyst was placed in dilute hydrogen, and the consumption of hydrogen was monitored using Autochem 2920 (Micrometrics) by increasing the temperature to 900 °C at a rate of 10 °C min<sup>-1</sup>. X-ray photoelectron spectroscopy (XPS) analysis was performed using an ESCALAB 210 (VG Scientific), and Al K $\alpha$  monochromate (1486.6 eV) was used as an excitation source. Fourier-transform infrared (FT-IR) spectroscopy experiments were conducted in a diffuse reflection cell equipped with a CaF<sub>2</sub> window using an FT-IR spectrometer (Nicolet IS 10, Thermo-Fisher), and diffuse reflectance (DR) 400 accessory was used. The spectra included 30 accumulated scans at resolutions of 4 cm<sup>-1</sup>, which were obtained using a mercury–cadmium–telluride (MCT) detector. To investigate the characteristics of CO<sub>2</sub> adsorption and CO<sub>2</sub> methanation reaction, the gas flowing over the samples pretreated by H<sub>2</sub> and air were switched to CO<sub>2</sub> or CO<sub>2</sub> + H<sub>2</sub> for 20 min at 200 and 300 °C.

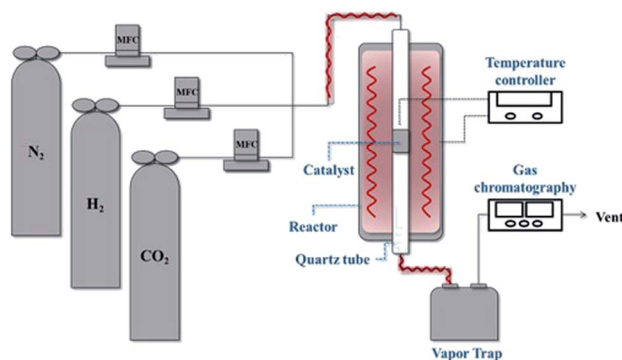


Fig. 1 Schematic of the experimental fixed bed catalytic reactor for CO<sub>2</sub> methanation.



### 3 Results and discussion

#### 3.1 Catalytic activities

Fig. 2 shows the CO<sub>2</sub> conversions of Ni-based catalysts supported on CeO<sub>2</sub>, TiO<sub>2</sub>, Y<sub>2</sub>O<sub>3</sub> and Al<sub>2</sub>O<sub>3</sub> at different reaction temperatures for the CO<sub>2</sub> methanation reaction. Among the catalysts, the reduced Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalysts exhibited the similar CO<sub>2</sub> conversion over a range from 250 to 400 °C. For the reduced Ni/Al<sub>2</sub>O<sub>3</sub> catalyst, the CO<sub>2</sub> conversion decreased at 250–280 °C and then reached the maximal value at 350–380 °C. Especially, the Ni/TiO<sub>2</sub> catalyst exhibited a very poor CO<sub>2</sub> conversion at all temperature range. The CO<sub>2</sub> to CH<sub>4</sub> and CO conversions are also shown in Fig. 2(b) and (c). CO<sub>2</sub> was fully converted to the CO (0.1–2.1%) in all temperature range for Ni/TiO<sub>2</sub> catalyst. From the properties of the equilibrium conversion, the reverse water gas shift reaction is not favored at a low temperature. Thus, it is indicated that the CO<sub>2</sub> should be able to convert CO on the reduction sites on a TiO<sub>2</sub> support. Except for the Ni/TiO<sub>2</sub> catalyst, CO yields were nearly zero at a reaction temperature under 300 °C, indicating that all CO<sub>2</sub> conversion is mostly concordant with CO<sub>2</sub> conversion to CH<sub>4</sub>. The order of catalytic activity regardless of the reaction temperature was as follows: Ni/CeO<sub>2</sub> ≈ Ni/Y<sub>2</sub>O<sub>3</sub> > Ni/Al<sub>2</sub>O<sub>3</sub> > Ni/TiO<sub>2</sub>. It could be concluded that the reduced Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalysts showed superior activity at a low temperature. Although catalytic activity of the Ni-based catalyst differs with operation

conditions such as temperature, pressure, catalyst loading, and gas component, the general results showed similarities to previously reported studies by other researchers.<sup>11,30,31</sup> Cai *et al.*<sup>30</sup> reported that CO<sub>2</sub> conversion of Ni/Ce<sub>x</sub>Zr<sub>1-x</sub>O<sub>2</sub> catalyst showed the 72.21% at 390 °C. The higher reducibility of the Ce-rich supported highly-dispersed Ni catalyst was considered to be an important factor for long-term stability.<sup>30,32,33</sup> However, the role of Ni dispersion in the stability of the Ni/CeO<sub>2</sub> catalyst is difficult to elucidate in this work, and further studies will be necessary to fully address this point.

#### 3.2 Physicochemical properties

The BET surface areas and metal dispersions of the reduced and unreduced samples are summarized in Table 1. The order of support's surface area was as follows: TiO<sub>2</sub> > CeO<sub>2</sub> > Al<sub>2</sub>O<sub>3</sub> > Y<sub>2</sub>O<sub>3</sub> (TiO<sub>2</sub>: 316.226, CeO<sub>2</sub>: 29.557, Al<sub>2</sub>O<sub>3</sub>: 7.375, Y<sub>2</sub>O<sub>3</sub>: 2.615 m<sup>2</sup> g<sup>-1</sup>; not shown in the figure). It can be observed that the Ni/TiO<sub>2</sub> catalyst had a much larger specific surface area than other catalysts. Zhang *et al.*<sup>34</sup> showed that the catalyst's specific surface area is not directly related to the catalytic activity in CO<sub>2</sub> methanation for Ni-based catalysts. Our results were in good agreement with previously mentioned evidence from the literature. The reduced Ni/TiO<sub>2</sub> catalyst had the lowest value of Ni dispersion (0.0116%). The reduced Ni/Y<sub>2</sub>O<sub>3</sub> catalyst showed the highest metal dispersion (2.7883%) of all the unreduced Ni-

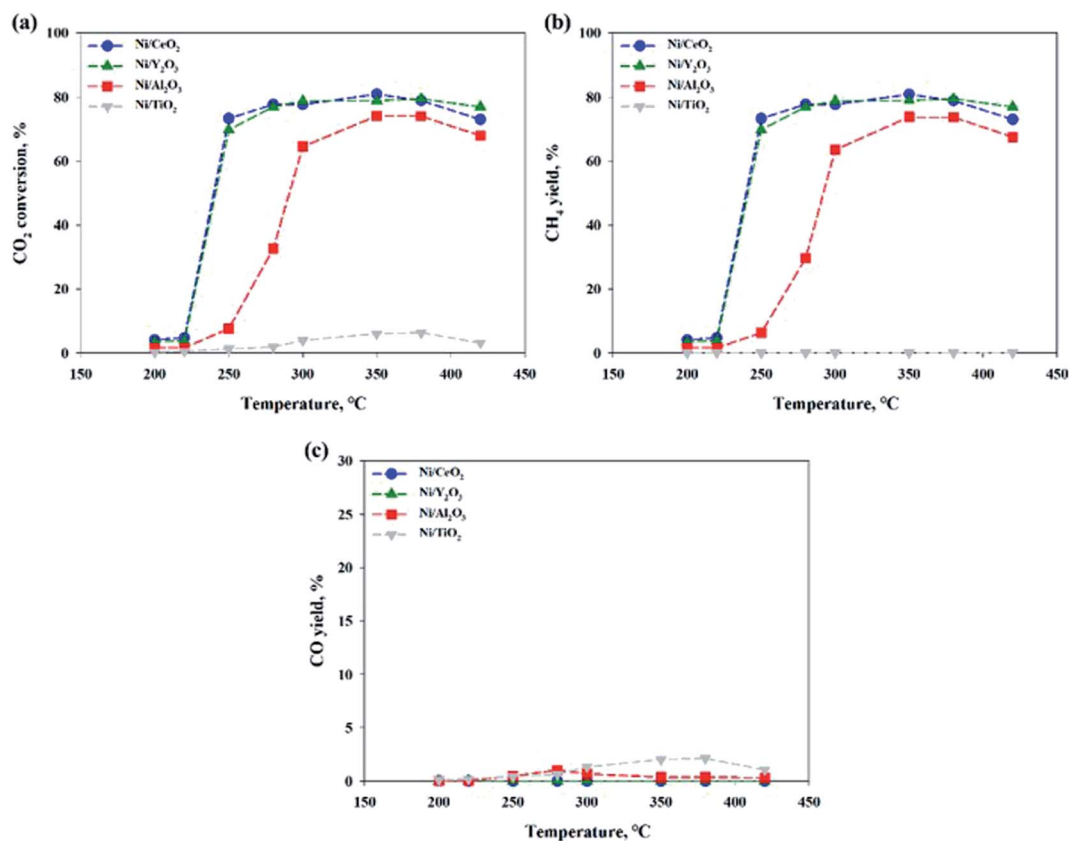


Fig. 2 (a) CO<sub>2</sub> conversions, (b) CH<sub>4</sub> and (c) CO yield in CO<sub>2</sub> methanation over 10 wt% Ni/support catalysts. CO<sub>2</sub> : H<sub>2</sub> : N<sub>2</sub> = 1 : 4 : 1, GHSV: 14 400 l kg<sup>-1</sup> h<sup>-1</sup>.



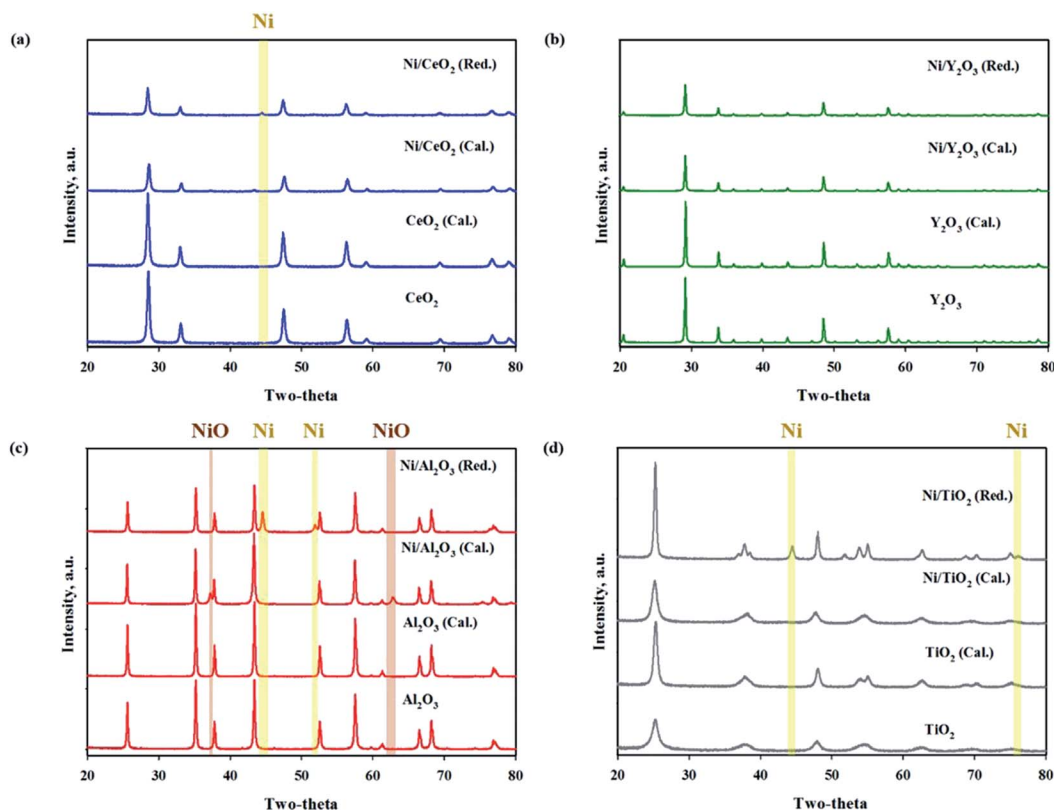
**Table 1** BET surface area and metal dispersion of reduced and unreduced Ni/metal oxide catalysts

	Surface area ( $\text{m}^2 \text{g}^{-1}$ )		Metal dispersion <sup>a</sup> (%)	
	Unreduced	Reduced	Unreduced	Reduced
Ni/CeO <sub>2</sub>	29.907	22.845	4.932	1.823
Ni/Y <sub>2</sub> O <sub>3</sub>	16.794	21.233	3.438	2.788
Ni/Al <sub>2</sub> O <sub>3</sub>	9.867	8.698	2.323	0.027
Ni/TiO <sub>2</sub>	121.74	43.606	0.067	0.012

<sup>a</sup> Calculated by H<sub>2</sub> pulse chemisorption.

based catalysts. Reduction degree of surface area and metal dispersion for Ni based catalysts with heat treatment by air or H<sub>2</sub> can be attributed to complex interdependency on metal-support interactions. This would suggest that a decrease of metal dispersion by reduction treatment could be explained by a strong interaction between Ni and oxide support. The order of strength of the metal and support interaction was as follows: Ni/Al<sub>2</sub>O<sub>3</sub> < Ni/TiO<sub>2</sub> < Ni/CeO<sub>2</sub> < Ni/Y<sub>2</sub>O<sub>3</sub>. Fig. 3(a)–(d) shows the XRD patterns of raw supports, calcined supports, calcined Ni based catalysts and calcined Ni based catalysts. The XRD pattern of Ni consists of three main peaks at 45°, 53°, and 76° corresponding to (111), (200), and (220) planes, respectively. The XRD pattern of NiO consists of two main peaks at 37° and 62° corresponding to (111) and (220) planes, respectively.<sup>35</sup> For the TiO<sub>2</sub> powder,

the main peaks were observed to be at  $2\theta = 25.3, 37.1$  and  $47.5^\circ$ , corresponding to typical anatase TiO<sub>2</sub> peaks. The sharp peaks were observed for the crystallite TiO<sub>2</sub> structure when the amorphous TiO<sub>2</sub> powder was calcined at a temperature of 400 °C. NiO peaks were not observed in calcined Ni/TiO<sub>2</sub> catalyst. It is interesting to note that Ni crystallite main peaks were detected at  $2\theta =$  at 44.5, 51.8 when using the reduced Ni/TiO<sub>2</sub> catalyst. The release of bonding oxygen atoms within Ni–O–Ti by H<sub>2</sub> reduction will help to move the Ni particles, then the Ni particles are easily agglomerated. Ni and NiO peaks were observed in both reduced and unreduced Ni/Al<sub>2</sub>O<sub>3</sub> catalysts. Unlike the Ni/TiO<sub>2</sub> catalyst, it confirmed that Ni and NiO particles were agglomerated by heat treatment regardless of the migration of oxygen atom for Ni supported on irreducible Al<sub>2</sub>O<sub>3</sub>. CeO<sub>2</sub> and Y<sub>2</sub>O<sub>3</sub> support had quite low surface area. Nevertheless, the Ni and NiO peaks were not observed in reduced and unreduced Ni/Y<sub>2</sub>O<sub>3</sub> catalyst and very low Ni and NiO peaks were observed in reduced and unreduced Ni/CeO<sub>2</sub> catalyst, which is related to strong metal-support interactions (SMSI effect). This phenomenon is in agreement with results from previous metal dispersion. TEM analysis was conducted to estimate the Ni particle sizes of Ni based catalysts. The TEM and mapping images of the reduced Ni catalysts are presented in Fig. 4(a)–(l). It could be seen that the Ni/Y<sub>2</sub>O<sub>3</sub> mean particle size of Ni was 12.8 nm, while that of Ni/CeO<sub>2</sub>, Ni/Al<sub>2</sub>O<sub>3</sub> and Ni/TiO<sub>2</sub> were 16.5 nm, 19.9 nm and 21 nm, respectively. Vogt *et al.*<sup>25</sup> investigated the particle size effect of Ni/SiO<sub>2</sub> catalysts prepared by



**Fig. 3** XRD patterns of metal oxide supports and reduced (red.) and unreduced (cal.) Ni supported catalysts ((a) Ni/CeO<sub>2</sub>, (b) Ni/Y<sub>2</sub>O<sub>3</sub>, (c) Ni/Al<sub>2</sub>O<sub>3</sub>, (d) Ni/TiO<sub>2</sub>).



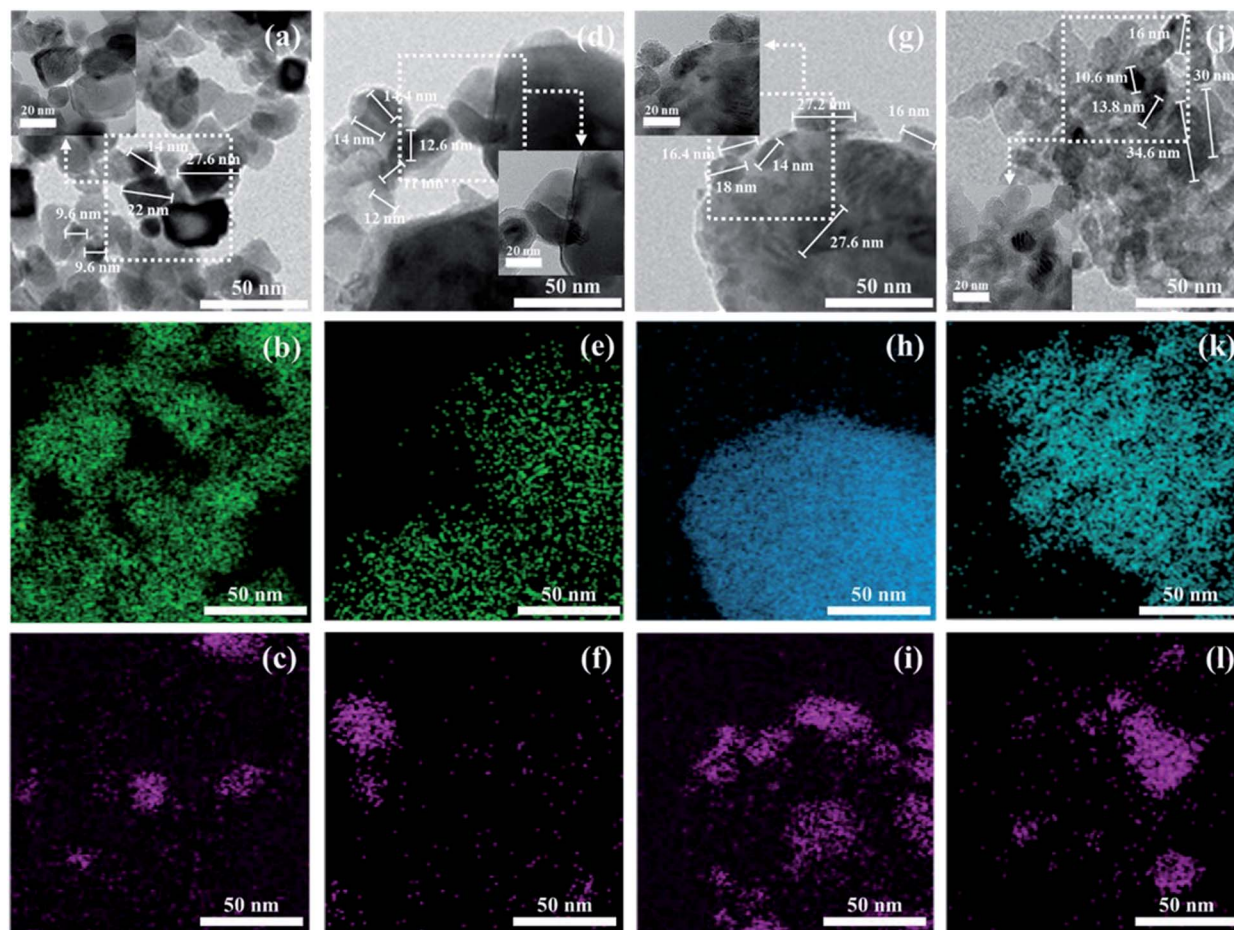


Fig. 4 TEM and mapping images of Ni/support catalysts ((a) Ni/CeO<sub>2</sub>, (b) Ce, (c) Ni, (d) Ni/Y<sub>2</sub>O<sub>3</sub>, (e) Y<sub>2</sub>O<sub>3</sub>, (f) Ni, (g) Ni/Al<sub>2</sub>O<sub>3</sub>, (h) Al<sub>2</sub>O<sub>3</sub>, (i) Ni, (j) Ni/TiO<sub>2</sub>, (k) TiO<sub>2</sub>, (l) Ni).

homogeneous deposition precipitation and co-precipitation with different Ni loadings (1–60 wt%). It was concluded that the Ni based CO<sub>2</sub> methanation is structure sensitive from 1–7 nm for Ni/SiO<sub>2</sub> catalyst. Many studies have attempted to enhance the Ni dispersion by increasing H<sub>2</sub> adsorption as active sites.<sup>11,36–39</sup> In this study, the 10 wt% Ni-based catalysts prepared by the impregnation method have relatively large Ni particle size of 12.8–21 nm, but it has low temperature CO<sub>2</sub> methanation activity. In the case of Ni/Al<sub>2</sub>O<sub>3</sub> and Ni/TiO<sub>2</sub> catalyst, it was confirmed that the difference in activity at high temperature (280–420 °C) was clearly displayed despite the similar particle size and metal dispersion. The fact that the Ni particles play an important role in the adsorption of hydrogen as active sites, but other factors such as CO<sub>2</sub> adsorption characteristics and oxygen transfer by hydrogen can influence the catalytic performance.

The reducibility of Ni-based catalysts was investigated by a H<sub>2</sub>-temperature programmed reduction, and the profiles are shown in Fig. 5(a). A similar H<sub>2</sub>-TPR result for the Ni/CeO<sub>2</sub> catalyst was previously reported.<sup>13,17,40</sup> Zhou *et al.*<sup>17</sup> reported that three reduction peaks can be seen at around 220 °C, 280–330 °C and 380 °C. The first low temperature peak attributed to the reduction of highly dispersed NiO species. The second peak attributed to the reduction of NiO species on the subsurface of

the Ni/CeO<sub>2</sub> catalyst and highly dispersive NiO species. The last peak can be assigned to the reduction of bulk NiO species. Tada *et al.*<sup>13</sup> reported that the two reduction peaks can be observed at 340 and 420 °C, which were attributed to the reduction of NiO on CeO<sub>2</sub> at about 400 °C. Ding *et al.*<sup>40</sup> also reported that the main reduction peaks appeared at 350–450 °C; these peaks were attributed to the reduction of NiO. In this study, the Ni/CeO<sub>2</sub> catalyst exhibited four-hydrogen-consumption maximum peaks at 165 °C, 225 °C, 290 °C and 785 °C, which were attributed to highly dispersive NiO species or Ni(OH)<sub>2</sub>, NiO species on the subsurfaces, Ni–O–Ce species, and bulk CeO<sub>2</sub> reduction peaks, respectively. It should be noted that all the reduction peaks are shifted to a lower temperature compared to previously reported studies. It can be suggested that, although the catalytic composition was the same, the reduction trend can differ, depending on the preparation conditions *e.g.*, the kinds of precursor, calcination temperature/time, and metal loadings. For the Ni/Al<sub>2</sub>O<sub>3</sub> catalyst, two reduction peaks were observed at 250–400 °C and 400–500 °C. The first peak is assigned to the free Ni species and second peak is attributed to the Ni species combined with the Al<sub>2</sub>O<sub>3</sub> support (Ni–O–Al).<sup>40</sup> Evidence from the literature showed that the last reduction peak can be observed for the Ni/Al<sub>2</sub>O<sub>3</sub> catalyst in the high-temperature



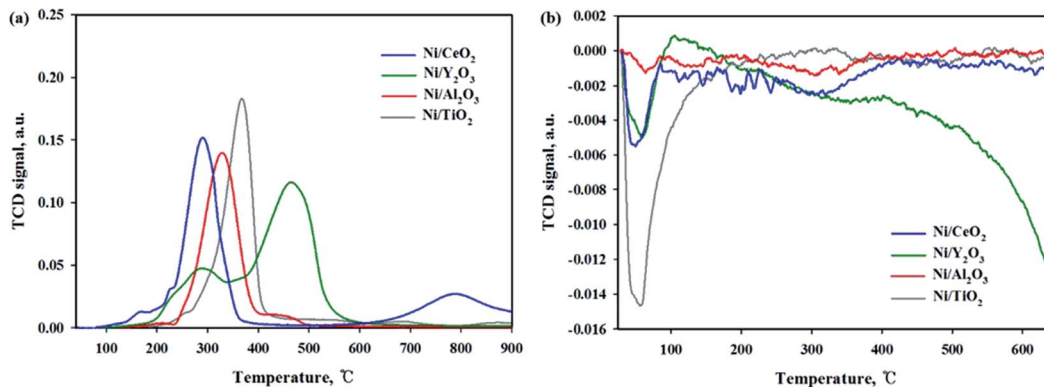


Fig. 5 (a)  $\text{H}_2$ -TPR and (b)  $\text{CO}_2$  TPO profiles of Ni/metal oxide catalysts.

region (750–850 °C), suggesting a stronger interaction between NiO and the  $\text{Al}_2\text{O}_3$  support.<sup>21</sup> But this peak was not observed in this sample; this might be due to a low calcination temperature. For the  $\text{Ni}/\text{Y}_2\text{O}_3$  catalyst, the two broad reduction peaks were observed in the 200–350 °C and 370–600 °C ranges, respectively. The first peak indicate the existence of a interaction between NiO and the  $\text{Y}_2\text{O}_3$  support and the second reduction peak is attributed to the bulk  $\text{Y}_2\text{O}_3$  support. The  $\text{Ni}/\text{TiO}_2$  catalyst exhibited only one broad hydrogen consumption peak at 250–420 °C, which was attributed to the Ni species combined with the  $\text{TiO}_2$  support (Ni–O–Ti). It is expected that the 10 wt% Ni based catalysts may have a metallic Ni form at 420 °C of reduction temperature. According to the  $\text{H}_2$  TPR and activity test results of all catalysts, enhancement of the  $\text{CO}_2$  conversion could be mainly due to the higher amount of NiO species able to be reduced at low temperature.  $\text{CO}_2$  temperature-programmed oxidation (TPO) was performed to observe the  $\text{CO}_2$  conversion for reduced catalysts at 420 °C as shown in Fig. 5(b). For the reduced  $\text{Ni}/\text{TiO}_2$  catalyst, the one huge  $\text{CO}_2$  consumption peak was observed in the 50–150 °C. This result indicated that the reduced  $\text{Ni}/\text{TiO}_2$  catalyst was able to accept oxygen by introduction of  $\text{CO}_2$  on reducible sites. However, on the Ni surface H<sub>2</sub> adsorption is difficult due to low Ni dispersion, which makes it difficult to react the  $\text{CO}_2$  methanation. The reduced  $\text{Ni}/\text{CeO}_2$  and  $\text{Ni}/\text{Y}_2\text{O}_3$  catalysts exhibited the similar  $\text{CO}_2$  consumption peak at 50–100 °C. For the reduced  $\text{Ni}/\text{Al}_2\text{O}_3$  catalyst, the  $\text{CO}_2$  consumption peak was not observed at all temperature range, it may be concluded that the  $\text{CO}_2$  molecules cannot be dissociated on reduced  $\text{Ni}/\text{Al}_2\text{O}_3$  catalyst. Among the catalysts used in this study, the  $\text{Ni}/\text{CeO}_2$  and  $\text{Ni}/\text{Y}_2\text{O}_3$  catalyst had the higher dispersion and enhanced reducibility of Ni particles, as well as strong metal and support interactions, which could become an important factor for the catalytic activity during low-temperature  $\text{CO}_2$  methanation.

### 3.3 In situ DRIFTS

To investigate the interaction of  $\text{CO}_2$  with catalysts, *in situ* DRIFT studies were performed on the reduced and unreduced catalysts at 200 and 300 °C as shown in Fig. 6. Fig. 6(a) shows the  $\text{CO}_2$  adsorption profile on the  $\text{Ni}/\text{CeO}_2$  catalyst surface. The

assignment of  $\text{CO}_2$  adsorption bands for the  $\text{Ni}/\text{Ce}_{0.5}\text{Zr}_{0.5}\text{O}_2$  catalyst was performed by reported in the previous literature.<sup>41,42</sup> The spectra of  $\text{CO}_2$  adsorption on  $\text{Ni}/\text{CeO}_2$  are similar to those of the  $\text{Ni}/\text{Ce}_{0.5}\text{Zr}_{0.5}\text{O}_2$  catalyst. The band observed at 1595  $\text{cm}^{-1}$  can be assigned to formate species and the bands centered at 1367  $\text{cm}^{-1}$  was assigned to monodentate carbonate. These two main bands were still maintained at 300 °C, implying that the formation of formate and monodentate carbonate species are favored at high temperature, as the main intermediate during direct hydrogenation of  $\text{CO}_2$ .<sup>41,43–46</sup> The reduced and unreduced  $\text{Ni}/\text{CeO}_2$  catalysts exhibited similar  $\text{CO}_2$  adsorption bands, regardless of calcination or reduction treatments, indicating that the  $\text{CO}_2$  adsorption properties on  $\text{Ni}/\text{CeO}_2$  catalyst may depend on the  $\text{CeO}_2$  support, not NiO and metallic Ni species.  $\text{CO}_2$  adsorption bands were compared for  $\text{Ni}/\text{Y}_2\text{O}_3$  and  $\text{Ni}/\text{Al}_2\text{O}_3$  in Fig. 6(b) and (c). The main adsorption bands were detected at 1541, 1290, 1215 and 1046  $\text{cm}^{-1}$  (formate), and 1571  $\text{cm}^{-1}$  (bidentate carbonate) for  $\text{Ni}/\text{Y}_2\text{O}_3$  catalyst and at 1606, 1406 and 1364  $\text{cm}^{-1}$  (formate), and 1364  $\text{cm}^{-1}$  (bidentate carbonate) for  $\text{Ni}/\text{Al}_2\text{O}_3$ . The formate bands were maintained at 300 °C for  $\text{Ni}/\text{Al}_2\text{O}_3$  and  $\text{Ni}/\text{Y}_2\text{O}_3$  catalysts. The intensity of main peaks increased with higher temperature (300 °C) for  $\text{Ni}/\text{Al}_2\text{O}_3$  catalyst. In Fig. 6(d), the  $\text{CO}_2$  adsorption bands were observed for the  $\text{Ni}/\text{TiO}_2$  catalyst. The bands centered at 1219 and 1067  $\text{cm}^{-1}$  were attributed to monodentate species and the bands centered at 1620 and 1219  $\text{cm}^{-1}$  were assigned to hydrogen carbonate peaks. The intensity of  $\text{CO}_2$  absorbed peaks decreased at calcined catalyst while these peaks increased with  $\text{H}_2$  reduction treatment, indicating that a small part of catalytic sites become active by surface OH groups, and  $\text{CO}_2$  adsorption peaks are increased. Fig. 7 shows the  $\text{CO}_2$  adsorption profiles of all reduced and unreduced catalysts in the  $\text{CO}_2$ - $4\text{H}_2$  atmosphere as a function of different temperatures. The two main bands attributed to formate and monodentate carbonate were still maintained at 200 and 300 °C for the calcined  $\text{Ni}/\text{CeO}_2$  catalyst when introducing  $\text{CO}_2$  and  $\text{H}_2$  (Fig. 7(a)). These peaks were maintained at the reaction temperature of 200 °C for reduced  $\text{Ni}/\text{CeO}_2$  catalyst but the intensity of peaks decreased obviously at reaction temperature of 300 °C and the two bands were detected at 1900 and 2038  $\text{cm}^{-1}$  which was assigned bridged CO and linear CO



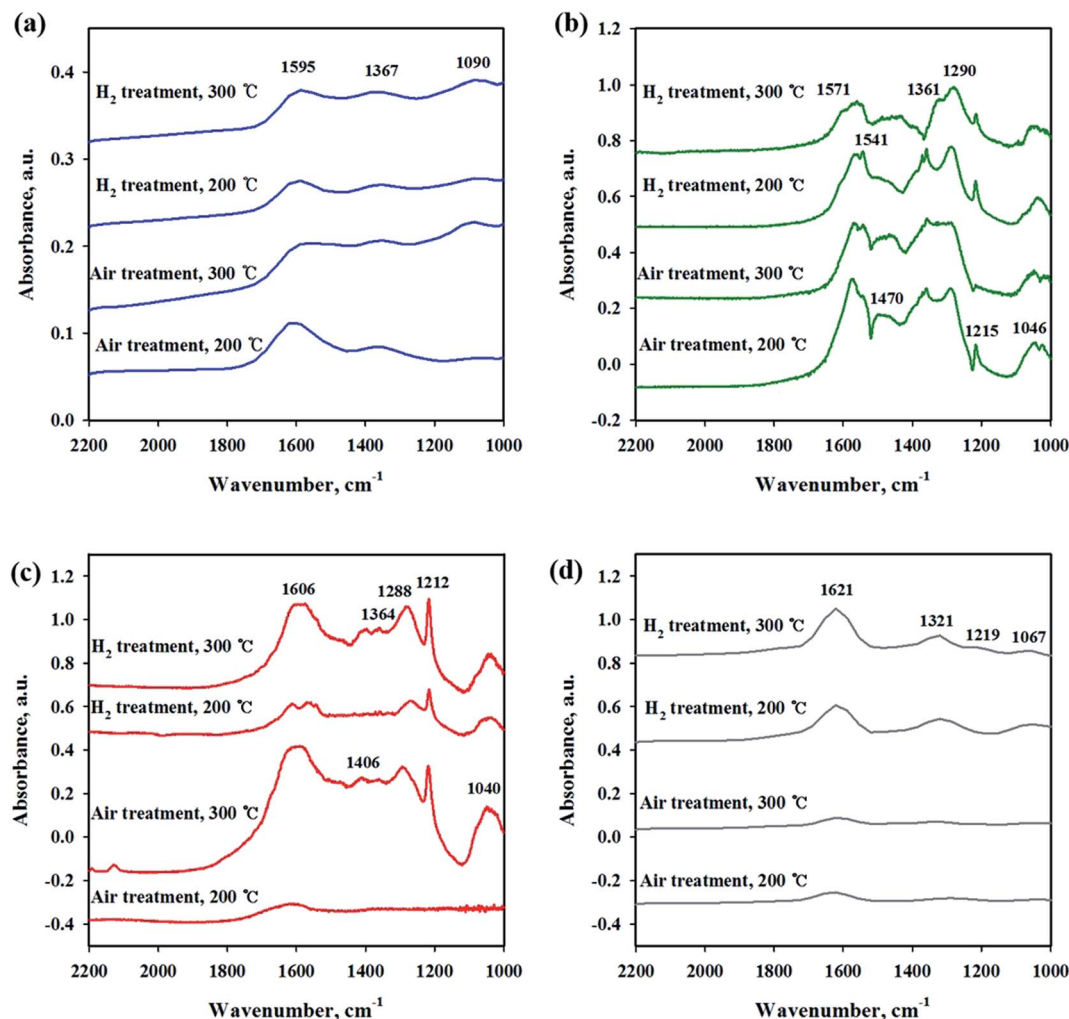


Fig. 6 *In situ* DRIFTS studies of CO<sub>2</sub> adsorption on Ni/metal oxide catalysts with different pretreatment and temperature conditions. (a) Ni/CeO<sub>2</sub>, (b) Ni/Y<sub>2</sub>O<sub>3</sub>, (c) Ni/Al<sub>2</sub>O<sub>3</sub> and (d) Ni/TiO<sub>2</sub>.

bands adsorbed on Ni, respectively.<sup>47</sup> This result indicated that CO<sub>2</sub> was readily dissociated to CO and it does participate in the CO<sub>2</sub> methanation reaction at high temperature for Ni/CeO<sub>2</sub> catalyst. Aldana *et al.*<sup>44</sup> reported that methanation proceeds through formate species originated from the hydrogenation of carbonates and CO was formed by a redox cycle on reduced ceria support. At 300 °C, only monodentate carbonate bands well remained. This result indicated that the monodentate carbonates are not easily hydrogenated; this result was in good agreement with Pan *et al.*<sup>41</sup> As shown in Fig. 7(b), although hydrogen is injected, the main formate and bidentate carbonate bands were still maintained at 200 and 300 °C for the calcined Ni/Y<sub>2</sub>O<sub>3</sub> catalyst. The small and broad formate bands observed at 300 °C for the reduced Ni/Y<sub>2</sub>O<sub>3</sub> catalyst and the two small bands were detected at 1900 and 2038 cm<sup>-1</sup> which was assigned bridged CO and linear CO bands adsorbed on Ni, respectively, it form a similar pattern with reduced Ni/CeO<sub>2</sub> catalyst. This formation of active species as formate leads to higher catalytic activity at high temperature. It is expected that sufficient reduction pretreatment of reducible supports such as CeO<sub>2</sub> and

Y<sub>2</sub>O<sub>3</sub> supported catalysts efforts on CO<sub>2</sub> decomposition into CO which is an crucial step in CO<sub>2</sub> conversion into CH<sub>4</sub> at high temperature and thus catalytic activity using Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalysts is enhanced.<sup>13</sup> In Fig. 7(c), the CO<sub>2</sub> absorption bands were still maintained at 200 °C for the calcined Ni/Al<sub>2</sub>O<sub>3</sub> catalyst, it forms a similar pattern with CO<sub>2</sub> adsorption pattern as shown in Fig. 6(c). CO<sub>2</sub> methanation is not progressed in this low temperature for Ni/Al<sub>2</sub>O<sub>3</sub> catalyst. The similar bands were detected at 300 °C for the calcined Ni/Al<sub>2</sub>O<sub>3</sub> catalyst, but the small band observed at 2027 cm<sup>-1</sup> which was assigned linear CO bands adsorbed on Ni. It is known from the H<sub>2</sub> TPR results that the release of bonding oxygen atoms within Ni–O–Al by H<sub>2</sub> reduction at 300 °C, CO<sub>2</sub> was readily dissociated to CO on reduction active site for the calcined Ni/Al<sub>2</sub>O<sub>3</sub> catalyst. The spectra for the reduced Ni/Al<sub>2</sub>O<sub>3</sub> catalyst exhibited small overlapped bands of formate species and the intensity of bridged CO and linear CO peaks centered at 1870 and 2027 cm<sup>-1</sup> increased. This result indicated that CO<sub>2</sub> was readily dissociated to CO and it does participate in the CO<sub>2</sub> methanation reaction at high temperature for reduced Ni/Al<sub>2</sub>O<sub>3</sub> catalyst. The Ni/TiO<sub>2</sub> catalyst



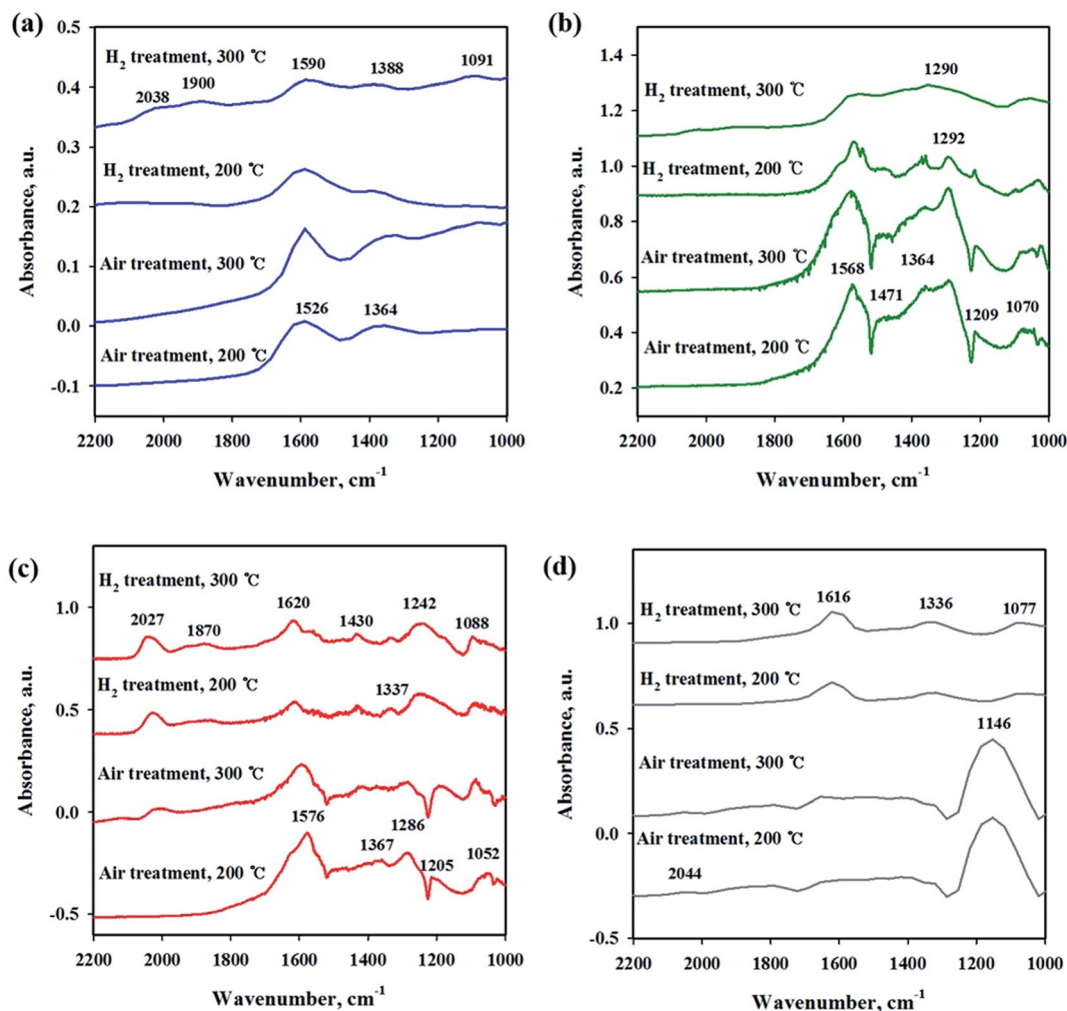


Fig. 7 *In situ* DRIFTS studies of CO<sub>2</sub> methanation (CO<sub>2</sub> + H<sub>2</sub>) on Ni/metal oxide catalysts with different pretreatment and temperature conditions. (a) Ni/CeO<sub>2</sub>, (b) Ni/Y<sub>2</sub>O<sub>3</sub>, (c) Ni/Al<sub>2</sub>O<sub>3</sub> and (d) Ni/TiO<sub>2</sub>.

showed a different adsorption bands than other catalyst as shown in Fig. 7(d). The intensity of broad hydrogen carbonate peak centered at 1000–1300 cm<sup>-1</sup> increased for calcination Ni/TiO<sub>2</sub> catalyst and the two small bands were observed at 1336 and 1077 cm<sup>-1</sup> which was assigned monodentate peaks regardless of reaction temperature. The amount of CO<sub>2</sub> adsorbed onto Ni/TiO<sub>2</sub> was much larger than that onto other catalyst, but CO<sub>2</sub> methanation reaction cannot occur regardless of reaction temperature. Due to the low Ni dispersion, it can be inferred that adsorbed monodentate and hydrogen carbonate species are difficult to react with adsorbed hydrogen atoms on active metallic Ni site to form methane so that CO<sub>2</sub> methanation reaction does not take place at all temperatures. Numerous studies about reaction mechanism of CO<sub>2</sub> methanation have been investigated. The reaction mechanisms are normally classified into two reaction pathways. One involves CO<sub>2</sub> conversion to CO as intermediate, which then follows the same mechanism as CO methanation.<sup>48</sup> The other one involves formate and carbonate as main intermediate, which directly hydrogenate without forming CO.<sup>44</sup> From the results in this

study, it was assumed that the methanation reaction of CO<sub>2</sub> will follow the first mechanism for Ni/Al<sub>2</sub>O<sub>3</sub> catalyst and formate and carbonate as main intermediate mechanism at low temperature (200 °C) for Ni/Y<sub>2</sub>O<sub>3</sub> and Ni/CeO<sub>2</sub> catalysts. However, the methanation reaction of CO<sub>2</sub> will follow the first mechanism at high temperature (300 °C) for Ni/Y<sub>2</sub>O<sub>3</sub> and Ni/CeO<sub>2</sub> catalysts. These different mechanisms might be one of the reasons why the role of supports. The Ni supported irreducible Al<sub>2</sub>O<sub>3</sub> catalyst favors the CO<sub>2</sub> conversion to CO and then follows the same mechanism as CO methanation. While carbonate and formate species were found to be the main intermediate on surface oxygen vacancy site such as Ce<sup>3+</sup> sites, which could enhance the catalytic activity at low temperature compare to Ni/Al<sub>2</sub>O<sub>3</sub> catalyst.<sup>3</sup> The carbonate and formate species were mainly present as an adsorption intermediate on surface oxygen vacancy site (Ce<sup>3+</sup> sites), while the CO species were found to be the main adsorption intermediate for the irreducible Al<sub>2</sub>O<sub>3</sub> supported Ni catalyst. According to TPR and DRIFT results, the interaction between Ni and Ce or Y may facilitate the formation of the OH groups on the Ni–OH or Y–OH, Ce–OH bonds. The





dissociated hydrogen atoms on the Ni metal spillover onto the ceria or yttria support surface, and undergo surface diffusion at low temperature by introduction of hydrogen. Superior reducibility of both Ni-rich surface species and CeO<sub>2</sub> support by strong interaction between Ni and CeO<sub>2</sub> or Y<sub>2</sub>O<sub>3</sub> supports are important factor and then weakly adsorbed CO<sub>2</sub> species such as formate and carbonate on the surface oxygen vacancy site easily reacted with hydrogen dissociation on the metal followed by spillover at low temperature.

## 4 Conclusions

The Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalysts show highly enhanced catalytic activity for low-temperature CO<sub>2</sub> methanation, as compared with the Ni/Al<sub>2</sub>O<sub>3</sub> and Ni/TiO<sub>2</sub> catalysts. The increase in the low-temperature CO<sub>2</sub> methanation activity can be directly correlated with an enhancement in the reducibility and small Ni particle size of the Ni/CeO<sub>2</sub> and Ni/Y<sub>2</sub>O<sub>3</sub> catalyst, which is caused by a strong interaction between reduced CeO<sub>2</sub> or Y<sub>2</sub>O<sub>3</sub> oxide support and Ni. When Ni species are dispersed on “reducible” oxides, such as CeO<sub>2</sub> and Y<sub>2</sub>O<sub>3</sub>, carbonate and formate species were mainly present as an adsorption intermediate on the surface oxygen vacancy site such as Y<sup>(3-x)+</sup> or Ce<sup>3+</sup> sites at low temperature, which plays a key role in enhancing the low catalytic activity. Due to the low Ni dispersion of Ni/TiO<sub>2</sub> catalyst, it can be concluded that adsorbed CO<sub>2</sub> species are difficult to react with hydrogen atoms on active metallic Ni site to form methane. When the Ni species were dispersed on “irreducible” oxides, species such as Al<sub>2</sub>O<sub>3</sub> and CO were found to be the main adsorption intermediates. Overall, CeO<sub>2</sub> and Y<sub>2</sub>O<sub>3</sub> supports are promising support for the Ni-based catalyst and further improvement in low-temperature CO<sub>2</sub> methanation activity can be made by modification of the mixed CeO<sub>2</sub> and Y<sub>2</sub>O<sub>3</sub> support with small Ni particle size and oxygen vacancies.

## Author contributions

Ye Hwan Lee: conceptualization, writing – original draft. Jeong Yoon Ahn: experiment and evaluation. Dinh Duc Nguyen: data curation, validation, formal analysis. Soon Woong Chang: writing – review & editing. Sung Su Kim: writing – review & editing. Sang Moon Lee: supervision, project administration.

## Conflicts of interest

There are no conflicts to declare.

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