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Ratchet-like mechanism in a long-life photoproduct of salicylideneaniline enclathrated in a pillared-layer guanidinium disulfonate structure†

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Salicylideneaniline (SA) was found to exhibit extraordinary long-life photochromism upon being included in a cavity of guanidinium organosulfonate and exhibited a lifetime of ~ 30 times that of a pure SA crystal. Crystal structure analysis suggested that the sulfonate molecule in the apo-host provided a flexible cavity space that kinetically trapped SA in its photo-isomerized form, as if it was locked by a ratchet-like mechanism.

Crystal engineering is a promising strategy for the modification of the chemical and physical properties of molecular solids to obtain desirable functions.¹ For example, salicylideneanilines (SAs), which are known to exhibit either photochromism or thermochromism depending on their molecular environment,^{2–4} are often employed to exemplify the effects of polymorphs, cocrystallization, salification, or inclusion in a porous coordination network.^{5–8} The chromism of SA originates in the interconversion of three tautomers: the enol, *cis*-keto, and *trans*-keto forms (Chart 1).^{9,10} Therefore, the modification of the energy landscape by changing the surroundings of the SA chromophore facilitates chromism. Guanidinium organosulfonate (GS), which is known to spontaneously form a pillared-layer structure,^{11–14} is a promising candidate for providing a variety of molecular hosts to accommodate SA molecules as guests. In the present study, the use of biphenyl-4,4'-disulfonate as a pillar resulted in the clathrate compound, SA@1, which was a photochromic compound with an extraordinarily long life, whereas the use of naphthalene-2,6-

disulfonate afforded a photo-inactive clathrate compound, SA@2.

Molecular apo-hosts **1** and **2** were prepared in accordance with studies presented by Ward *et al.*^{14,15} After their preparation, **1** and **2**, were dissolved in methanol with 3 equivalents (to guanidinium) of SA, respectively, and were left to sit until the open evaporation initiated crystallization.† The products were examined by single-crystal X-ray structural analysis, which revealed that **1** and **2** afforded the molecular clathrates SA@1 and SA@2, respectively. Several crystal structures have been reported for several clathrates and solvates of **1** and **2**,¹⁴ and their connection topology that features hydrogen-bonding networks is known to vary with the type of guest and solvent molecules. Chart 2 illustrates the framework of the putative apo-hosts **1** and **2** in SA@1 and SA@2. The ratio of SA : disulfonate was 1 : 2, which implied that each unit cell (with volumes of

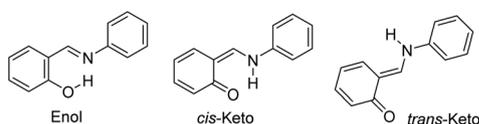


Chart 1 Three tautomers of salicylideneaniline relevant to its chromic behaviour.

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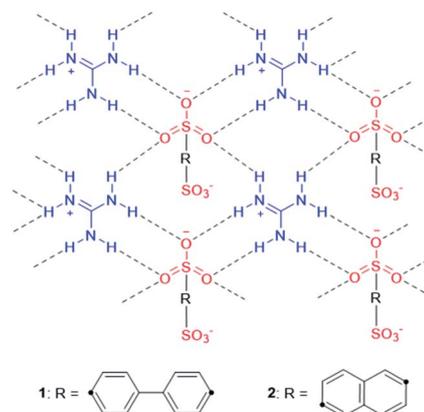


Chart 2 Schematic representation of apo-hosts **1** and **2**. The connection topology featuring hydrogen-bonding networks can be observed in their SA-clathrates, SA@1 and SA@2.



1227.95 Å³ for SA@1 and 1098.28 Å³ for SA@2) contained one molecule of SA (Fig. 1). However, the NMR profiles revealed that the ratio (0.44 and 0.41) of SA to sulfonate molecules was lower than that (0.50) derived from stoichiometry (Fig. S1†), implying that the bulk product contained an appreciable amount of the apo-hosts. FTIR spectra¹⁶ (Fig. S2†) and powder XRD patterns (Fig. S3†) also implies the contamination of the apo-host.

Upon irradiation with UV light, the color of the SA@1 crystal changes from pale yellow to orange, whereas that of the SA@2 crystal does not. Absorption spectra were subsequently obtained with a microscope-based UV-vis absorption spectrometer equipped with a temperature controller and a high-pressure Hg lamp (through a 330–380 nm band path filter) to investigate this.^{17–19} Upon irradiation with the Hg lamp at 15 °C, the SA@1 sample immediately shows an intense absorption band at ~500 nm with a blunt vibronic structure (Fig. 2a). An unmixed SA molecule is known to exhibit three crystalline polymorphs, *i.e.* α_1 -, α_2 -, and β -forms,²⁰ among which the α_1 -form is the most common photochromic species. The absorption spectrum of SA@1 in the photo-stationary state is quite similar to those of α_1 -SA,⁵ suggesting that the newly appeared band can be attributed to the photo-induced *trans*-keto form. When the sample is cooled to –60 °C, the spectrum in the photo-stationary state

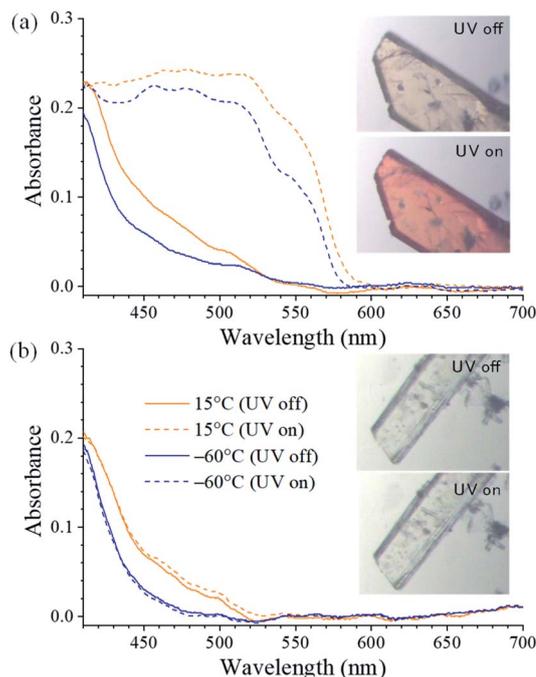


Fig. 2 Single-crystal UV-vis absorption spectra obtained with and without UV irradiation in (a) SA@1 and (b) SA@2. (inset) Microscope images at 25 °C.

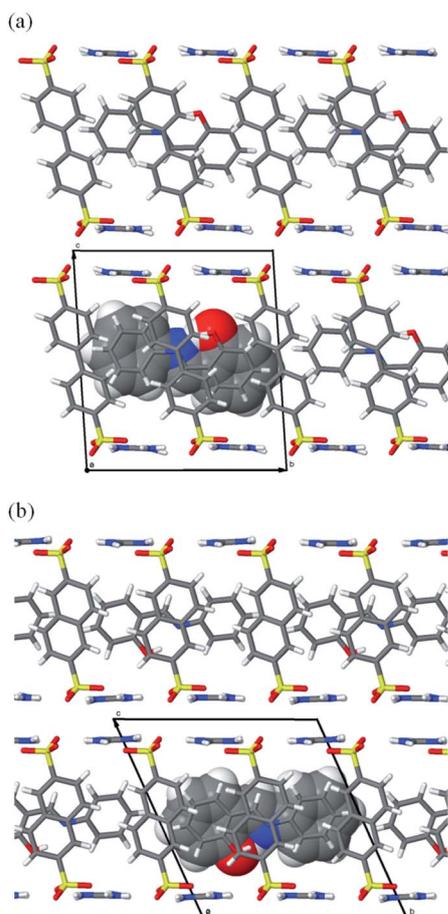


Fig. 1 Crystal structures of (a) SA@1 and (b) SA@2. Views along other directions are available in Fig. S4.†

reveal a clearer structure, whereas that in the dark-adapted state is almost unchanged. The spectra of SA@2 without the Hg lamp are noted to be similar to those of SA@1; however, they did not show significant changes upon UV irradiation (Fig. 2b). Both samples were photoluminescent but with different spectral profiles: SA@1 exhibited an orange emission peak at ~590 nm, and SA@2 exhibited a yellowish-green emission peak at ~520 nm (Fig. 3). This difference is similar to the behavior observed in the photochromic and thermochromic polymorphs of 3,5-di-*tert*-butylsilylideneaniline.¹⁷ There are two possible reasons for the spectral shift: (1) the alteration of electronic state in SA molecule due to conformational change and/or intermolecular interactions inside the cavity; (2) reabsorption

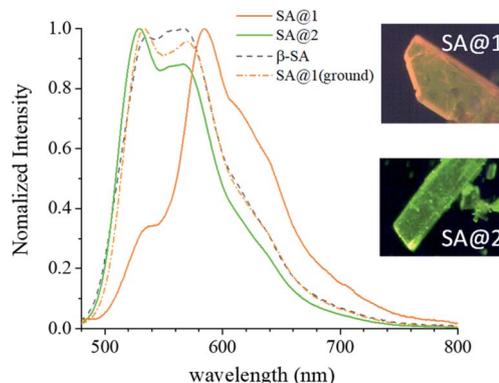


Fig. 3 Emission spectrum of SA@1, SA@2, as compared with that of β -SA. Spectra of SA@1 after grinding with a mortar is overlaid.



of a 500–600 nm region of the emission by the photo-product. The first reason is rather unlikely in view of a more planar structure (angle between two phenyl rings is 2.8°) of SA@2 than that (23°) of SA@1. Interestingly, the luminescence spectra of SA@1 appeared similar to those of SA@2 after the SA@1 crystals were pulverized in an agate mortar; however, its photochromism was maintained. This phenomenon implies that the pulverization induced a partial collapse of the pillared-layer structure, but leaves some room for discussion about the coexistence of yellowish-green luminescence and photochromism that seems contradictory to the second reason above. These observations will be a clue to a complete understanding of the current system, and worth further inspecting in the future.

The duration of photo-coloration for SA@1 was notably longer than that for α_1 -SA. This was investigated *via* an examination of the bleaching process by measuring the absorbance decay curves at 480 nm, starting from the photo-stationary state at 15°C (Fig. 4). As previously noted,^{18,19} the decay curves do not fit well to a single exponential function; hence, a time constant cannot be extracted. Instead, the values of T_{50} and T_{95} (durations at 50% and 95% decay, respectively) were compared, which revealed that the lifetime of SA@1 was 30 times longer than that of α_1 -SA (Table 1). Although the lifetime is known to vary with conditions such as the temperature or intensity of ambient light, these results clearly showcase the overwhelming kinetic stability of the photoproduct in SA@1. These experiments were repeated under the same conditions with *N*-3,5-di-*tert*-butylsalicylidene-3-nitroaniline (SA-NO₂), which is known to possess an extremely long lifetime (half-life of ~ 40 days and 1200 min as obtained by Kawato *et al.*²¹ and Ohashi *et al.*,²² respectively; these values were measured in the dark); the results suggest that SA@1 possesses a longer lifetime ($\sim 2\times$) than that of SA-NO₂ (Fig. 4).

The crystal structure of SA@1 was solved under the $P\bar{1}$ space group, which means that the SA molecule was assumed to have a disorder ratio of 1 : 1 for opposite orientations, with respect to the pseudo-inversion point at the center of the C=N bond. In addition, the phenyl rings of biphenyl-4,4'-disulfonate were assumed to have disorders for rotations of $17\text{--}40^\circ$ with respect

Table 1 Comparison of the lifetimes (s) of SA@1, α_1 -SA, and SA-NO₂

Compound	T_{50}	T_{95}
SA@1	329	910
α_1 -SA	11	31
SA-NO ₂	223	361

to the axis along S...S (Fig. 5). These observations suggest that the cavity inside **1** can be flexibly shaped in accordance with the structure of the guest molecule(s). For reference, the crystal structure of the *o*-dichlorobenzene clathrate of **1** (CCDC refcode XAQUWH),²³ which has a unit cell and packing structure similar to those of SA@1, was solved without assuming any disorders, and revealed an almost parallel arrangement of the biphenyl moieties in a planar conformation. It is worth mentioning that this arrangement overlays decently with one of the disordered structures (shown in blue and magenta in Fig. 5) in SA@1 (Fig. S5†).

The torsion angle (θ) between the two phenyl rings in the SA molecule has been noted to indicate the possibility of photochromism. Previous studies have suggested that SA becomes thermochromic if the molecule is nearly planar ($\theta < \sim 30^\circ$), whereas it appears photochromic if the molecule is twisted ($\theta > \sim 30^\circ$).²⁻⁴ However, SA@1 ($\theta = 23^\circ$) narrowly fails at being categorized as a photochromic crystal in accordance with this criterion, whereas SA@2 ($\theta = 2.8^\circ$) is appropriately predicted as non-photochromic. Another requisite for photochromism involves a loosely packed (open) structure that allows *cis*-to-*trans* photoisomerization.¹⁻³ However, there is a lack of definitive discussions on the quantitative relationship between the degree of the open structure and photochromic activity. In addition to the difficulty in quantifying the open structure, the conventional dualistic classification into photochromic and thermochromic (non-photochromic) crystals is also an issue; it is possible for a photochromic crystal with an extremely loose packing to have a very high bleaching speed, and get classified as non-photochromic. Therefore, a sufficient long life of the photoproduct can be another requisite for photochromism. For

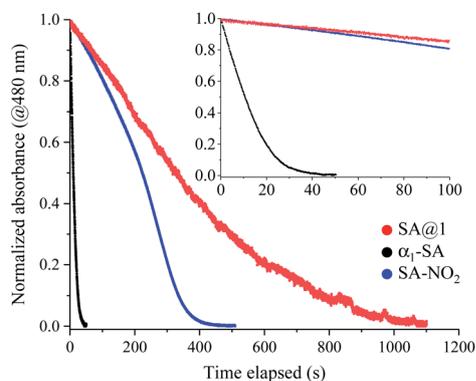


Fig. 4 Comparison of relaxation times in the photo bleaching of SA@1, SA-NO₂, and α_1 -SA.

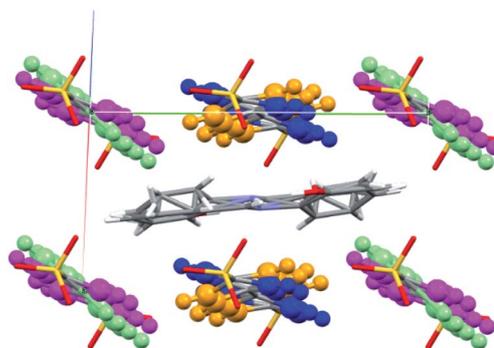


Fig. 5 Crystal structure of the biphenyl moieties surrounding the SA molecule in SA@1 (view along 101 direction). The structures are shown using an overlay of two different conformations, in blue and orange, and in magenta and pale green.



SA-NO₂ and polymorphs of its carboxylate analogue, a positive correlation between the lifetime and stability of the *trans*-keto form has been inferred.²²

Combining the above analyses, an explanation for the long-life photochromism of SA@1 can be proposed. The putative apo-host 1 possesses a cavity volume that is sufficient to accommodate the SA molecule, and the cavity shape is flexibly adjusted by the rotation of the phenyl rings of the pillars. When photo-isomerization occurs, the phenyl rings can rotate to follow the structural change. The rings rotate almost freely; however, they do not rotate independently because of the inter-pillar distance of 6.2 Å, which is insufficient for the two phenyl rings to avoid van der Waals contacts. Therefore, the biphenyl moieties serve as meshing gears to prevent the photoproduct from relaxing to the thermodynamically stable enol form; this mechanism can be compared to that of a ratchet gear. By contrast, SA@2, in which the cavity (480 Å³) is slightly smaller than that (546 Å³) of SA@1,²⁴ neither has enough volume to accommodate the twisting of the SA molecule, nor the flexibility to follow the structural change upon isomerization.

In summary, drastic changes in the photochemical properties of SAs were observed *via* the formation of clathrates of GS compounds. Since the discovery of GS compounds by Etter and Ward,^{11–13} their pillared-layer structures have been extensively applied for the isolation and stabilization of guest molecules.²⁵ The inclusion of functional dyes such as azobenzene,¹⁴ coumarin,¹⁵ and ROY²⁶ into GS hosts has been accomplished. It is worth noting that the torsion angle of ROY can be adjusted by changing the type of bifunctional organosulfonic acid employed. The present study demonstrates the possibility of GS clathrates to facilitate new systems with which the various properties of photochromic compounds can be examined in a controlled molecular environment.

Author contributions

The preparation and characterization of the compounds were conducted by K. Nakayama and T. Manako. The microscope absorption spectra and related data were collected by R. Koibuchi. The crystal structures were analyzed by I. Yoshikawa. The whole research project was supervised by H. Houjou. The original draft was written by K. Nakayama, and it was reviewed and edited by H. Houjou.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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