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Novel electroblowing synthesis of tin dioxide and composite tin dioxide/silicon dioxide submicron fibers for cobalt(II) uptake†

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Nanoscale SnO₂ has many important properties ranging from sorption of metal ions to gas sensing. Using a novel electroblowing method followed by calcination, we synthesized SnO₂ and composite SnO₂/SiO₂ submicron fibers with a Sn : Si molar ratio of 3 : 1. Different calcination temperatures and heating rates produced fibers with varying structures and morphologies. In all the fibers SnO₂ was detected by XRD indicating the SnO₂/SiO₂ fibers to be composite instead of complete mixtures. We studied the Co²⁺ separation ability of the fibers, since ⁶⁰Co is a problematic contaminant in nuclear power plant wastewaters. Both SnO₂ and SnO₂/SiO₂ fibers had an excellent Co²⁺ uptake with their highest uptake/*K_d* values being 99.82%/281 000 mL g⁻¹ and 99.79%/234 000 mL g⁻¹, respectively. Compared to the bare SnO₂ fibers, the SiO₂ component improved the elasticity and mechanical strength of the composite fibers which is advantageous in dynamic column operation.

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1 Introduction

Inorganic materials are interesting as industrial adsorbents due to their good resistance to decomposition at high temperatures and under ionizing radiation and with a wide operating pH range. Furthermore, they are typically much more selective than organic resins for metal ions. The high selectivities stem from their regular, porous and rigid structures that exhibit ion sieve functionality.¹ There has been research on the sorption characteristics of metal oxides, such as SiO₂, TiO₂, ZrO₂ and SnO₂ since the mid-20th century.² Among them, tin dioxide nano- and microparticles have shown promising sorption properties for both organic³ and inorganic^{4–13} compounds. Nanoscale SnO₂ has attracted attention as a multifunctional material and it has also been used in transistors,¹⁴ protective coatings¹⁵ and gas sensors.^{16,17} Similarly, both bare SnO₂ and composite SnO₂/SiO₂ nanofibers have exhibited promising gas sensing,^{18–25} electrochemical,²⁶ optical²⁷ and molecular filtration properties.²⁸ The large specific surface area and porous structure of fibers provide plenty of contact area for target gases^{24,25} and adsorbing species²⁸ and improve the selectivity for adsorbing molecules with different sizes.²⁸ The incorporation of an amorphous SiO₂ component into the SnO₂ fibers has been shown to improve both the gas sensing properties compared to SnO₂

nanoparticles²⁵ and the structural stability²⁴ and mechanical strength²⁸ of the fibers.

For removal of heavy metal ions from aqueous solutions, there are plenty of techniques such as chemical precipitation, ion exchange, adsorption, membrane filtration, electrochemical methods and phytoremediation.^{29,30} Adsorption has certain advantages over the other methods including high removal efficiency even at low concentrations, cost-effectiveness, simple design and less production of toxic sludge.^{29,31} Adsorbents comprise both organic and inorganic materials which may be synthetic or of natural origin, such as agricultural by-products. For example, titanium(IV) iodovanadate particles have been utilized for Pb²⁺ and Hg²⁺ separation³² and coco-peat biomass for Pb²⁺, Cd²⁺, Cu²⁺ and Ni²⁺ removal.³³ Inorganic submicron and nanofibers could be excellent adsorbents for a variety of industrial applications. Inorganic fibers have a large specific surface area due to their high surface-to-volume ratio and often form porous structures which can lead to a good adsorption capacity. Zirconium dioxide submicron fibers have performed well in flow-through column operation and shown better mechanical stability, less pressure build-up in the column and faster adsorption kinetics than the corresponding granules.³⁴ Submicron and nanofibers can also be an eco-friendly alternative for the purification of nuclear power plant (NPP) wastewaters, since they can markedly reduce the volume of solid radioactive waste that requires a permanent repository.

The most common method to prepare polymeric or inorganic submicron and nanofibers is electrospinning. In this technique, a high voltage is applied to a precursor solution

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containing a polymer, solvent(s) and for example a metal salt, depending on the desired fiber composition. During the electrospinning process, repulsive electrostatic forces cause the polymer chains to stretch and form fibers that are deposited on a grounded collector. *Via* post-electrospinning calcination, the polymer may be removed and the final fibrous product, typically a metal oxide, is formed. Different precursor solutions and operational parameters allow the control of the properties and morphology of the fibers.³⁵ Electrospinning is therefore a simple and cost-efficient method and upscalable for industry.^{36–39} Solution blowing (airbrushing) method in contrast uses pressurized gas to draw fibers from the precursor solution.⁴⁰ Solution blowing offers even 15 times faster production rate than electrospinning⁴¹ but it tends to result in bundles of aligned fibers whereas electrospinning yields an entanglement of individual, mainly un-aligned fibers.^{42,43} In electroblowing, the fibers are formed by both electrostatic force and air flow and hence the technique combines electrospinning and solution blowing and their advantages. Moreover, electroblowing enables 2.5 times faster solution feeding rate and thus fiber production rate than solution blowing.^{41,44} Compared to electrospinning, the additional air flow in electroblowing permits the use of more viscous precursor solutions⁴⁵ and produces fibers with smaller diameters and fewer beads.⁴⁶ It should be feasible to scale up electroblowing for industry, similarly to electrospinning.

To tackle climate change, global energy production is increasingly based on clean low-carbon power including nuclear power. However, the production of nuclear energy generates radioactive fission and activation products that pose a risk to human health and environment without proper treatment of nuclear waste effluents. Among the activation products, ⁶⁰Co caused by steel corrosion is one of the most hazardous ones due to its rather long half-life of 5.3 years and high gamma decay energies of 1.17 and 1.33 MeV.⁴⁷ Removal of ⁶⁰Co is therefore essential to diminish workers' radiation exposure, to reduce radioactive emissions to aquatic systems and to create safe waste forms for the final disposal. However, the separation of this radionuclide is quite difficult owing to its low concentration in the liquid waste of NPPs (402 Bq L⁻¹ ⁶⁰Co in floor drain water, Loviisa NPP, Finland).⁴⁸ The process for its uptake has to be very selective, since generally the liquid waste contains large amounts of inactive metal ions (*e.g.* 37 mg L⁻¹ Na⁺ in floor drain water, Loviisa NPP, Finland).⁴⁸

In this research, we prepared SnO₂ and composite SnO₂/SiO₂ submicron fibers by the electroblowing technique and calcination. We also studied the ability of the fibers to remove ⁵⁷Co²⁺ from an aqueous solution. To investigate the effect of SiO₂ on both the structure and Co²⁺ uptake performance of the SnO₂ fibers, we added 25 mol% of silicon in the synthesis solution of the composite fibers. The structure and morphology of all the fibers were characterized and their Co²⁺ removal capability was studied. To the best of our knowledge, this is the first report on the electroblowing synthesis of SnO₂ and SnO₂/SiO₂ submicron fibers and on their Co²⁺ uptake. This is also the first report on the uptake of metal ions by fibrous SnO₂ and SnO₂/SiO₂.

2 Experimental

2.1 Materials

The precursor solutions for the electroblowing experiments were prepared from SnCl₄·5H₂O (≥98%, Sigma-Aldrich), tetraethoxysilane (TEOS, Si(C₂H₅O)₄, 98%, Sigma-Aldrich), polyvinylpyrrolidone (PVP, (C₆H₉NO)_{*n*}, *M_w* = 1 300 000, Alfa Aesar), *N,N*-dimethylformamide (DMF, C₃H₇NO, ≥99.9%, Sigma-Aldrich) ethanol (C₂H₅OH, 96 vol%, GPR RECTAPUR) and deionized water. In the uptake experiments ⁵⁷CoCl₂, Co(NO₃)₂·6H₂O (98%, Sigma-Aldrich), NaNO₃ (≥99%, VWR Chemicals), NaCl (99.75%, Fisher Chemical), CaCl₂ (fused, granular, Fisher Chemical), HCl (1 M, Oy FF-Chemicals Ab), NaOH (1 M, Reag. Ph. Eur., VWR Chemicals) and deionized water were used. A stock solution of 5 kBq mL⁻¹ ⁵⁷Co²⁺ was prepared by dissolving ⁵⁷CoCl₂ in 1 mM HCl.

2.2 Synthesis of the SnO₂ and SnO₂/SiO₂ fibers

SnO₂ and SnO₂/SiO₂ fibers were synthesized by electroblowing a solution comprising the tin or tin and silicon precursors, PVP and solvents followed by calcination. For the synthesis of SnO₂ fibers, a certain mass of SnCl₄·5H₂O was dissolved in deionized water and then mixed with DMF and 18 wt% PVP/EtOH solution. The mass fractions of the constituents in the solution were 10 wt% for SnCl₄·5H₂O, 9 wt% for PVP, 11 wt% for H₂O, 39 wt% for EtOH and 31 wt% for DMF. For the preparation of SnO₂/SiO₂ fibers, SnCl₄·5H₂O was dissolved in deionized water and then mixed with DMF, 14 wt% PVP/EtOH solution and TEOS. The mass fractions of the constituents in the solution were 10 wt% for SnCl₄·5H₂O, 2 wt% for TEOS, 7 wt% for PVP, 11 wt% for H₂O, 40 wt% for EtOH and 30 wt% for DMF. The solutions were stirred at room temperature until they became clear and homogeneous.

For the electroblowing, a self-assembled apparatus was used.⁴⁹ In a typical experiment, 12 mL of precursor solution was withdrawn into a syringe and a 27 G (inner diam. 0.21 mm) needle was attached to the syringe. The syringe was placed on a syringe infusion pump (KD Scientific Legato® 101) and the solution feed rate was set to 15 mL h⁻¹. This feed rate is 4 to 500 times as high as reported for the electrospinning of SnO₂ and SnO₂/SiO₂ fibers.^{21–24,26,50,51} The needle was pushed through a 3 mm metal adapter of a box enclosing a cylindrical side collector and a planar back collector at 80 cm distance, both being made of a metal wire mesh. The potential difference between the needle and collectors was set to 15 kV with a high voltage power source, and compressed air was delivered through the adapter at a rate of 30 NL min⁻¹. The solution jet erupted from the needle tip was deposited as fibers on the grounded collectors and additional drying air was delivered to the box at a rate of 40 NL min⁻¹ to enhance solvent evaporation and to control the relative humidity within the box (≤20%). The as-electroblown fibrous mats were detached from the collectors and calcined in an air furnace at 400, 450 and 500 °C for 4 hours with a heating rate of 1 °C min⁻¹ in order to remove the polymer and to form the desired ceramic material. The as-electroblown SnO₂/SiO₂/PVP fibers were also calcined at 400 °C for 4 hours with heating rates of 5 and 10 °C min⁻¹. The yields of SnO₂ and



SnO₂/SiO₂ fibers were at best 0.54 and 0.53 g per hour of electroblowing, respectively.

2.3 Characterization of the SnO₂ and SnO₂/SiO₂ fibers

Morphology of the fibers was analysed by imaging with secondary electrons (SE) and transmitted electrons (TE) with a Hitachi S-4800 field emission SEM. Prior to the imaging with secondary electrons, the samples were placed on carbon tape and sputter coated with a 4 nm layer of Au/Pd alloy to improve conductivity. Elemental analysis of the fibers including elemental mapping was conducted with an Oxford INCA 350 energy dispersive X-ray spectroscopy (EDX) system connected with the Hitachi S-4800. A Quanta 3D 200i focused ion beam SEM (FIB-SEM) equipped with an Oxford INCA 350 EDX system and an Omniprobe nanomanipulator was used for extracting single SnO₂/SiO₂ fibers onto a copper grid for cross-sectional elemental mapping. The average diameters of the fibers and their grains were determined with a Fiji ImageJ software. The crystallinity of the fibers was analysed with a PANalytical X'Pert PRO MPD X-ray diffractometer using Cu K α radiation and focusing optics. The fiber samples were powdered prior to the analysis. The mean crystallite sizes and their weight ratios were determined from the XRD data by the Rietveld refinement using a MAUD software.⁵² Thermogravimetric analysis (TGA) of the as-electroblown fibers was conducted with a NETZSCH STA 449 F3 Jupiter® system using a heating rate of 10 °C min⁻¹ in a temperature range of 25 to 1000 °C in a flow of air (50 mol%) and N₂ (50 mol%, the purge gas). The specific surface area and porosity of the fibers were measured by N₂ physisorption at 77 K with a Micromeritics ASAP 2020 Gas sorption analyser.

2.4 Co²⁺ uptake studies

2.4.1 Effect of calcination temperature and heating rate. Co²⁺ uptake by the SnO₂ and SnO₂/SiO₂ fibers calcined at

different temperatures and with different heating rates was studied at a pH of 6. 20 mg of ground fibers was weighed into 20 mL scintillation vials and 10 mL of 0.01 M NaNO₃ solution containing 30 Bq mL⁻¹ ⁵⁷Co²⁺ was added into the vials. The pH of the solution was adjusted to 6 with a small volume of NaOH. The samples were equilibrated in a constant rotary mixer (50 rpm) for 24 hours after which the equilibrium pH was measured. The samples were phase separated by centrifugation at 4000 rpm (2100g) and syringe filtration (Acrodisc LC PVDF, 0.2 μ m). The ⁵⁷Co²⁺ uptake efficiency of each sample was determined by pipetting 5 mL of the filtrate into a scintillation vial and measuring the remaining activity with a PerkinElmer Wallac Wizard 3" 1480 automatic gamma counter. The ⁵⁷Co²⁺ uptake results are presented by means of distribution coefficient K_d , that describes the distribution of the adsorbate between the adsorbent and solution:

$$K_d = \frac{(c_0 - c_{eq})V}{c_{eq}m} \quad (1)$$

where c_0 (Bq L⁻¹) is the initial concentration, c_{eq} (Bq L⁻¹) is the equilibrium concentration, V (mL) is the volume of the solution and m (g) is the mass of dry adsorbent. Background activity was subtracted before the calculations. Uncertainty of K_d was calculated using the error propagation law.

2.4.2 Effect of pH. Co²⁺ uptake by SnO₂ fibers calcined at 500 °C was investigated in the pH range of 4 to 12. The batch samples were prepared as described above and the pH of the solution was adjusted with a small volume of either HCl or NaOH. The equilibrium pH was measured after the 24 hours of constant rotary mixing. The ⁵⁷Co²⁺ uptake by the fibers was calculated by means of the distribution coefficient K_d as described above.

2.4.3 Effect of coexisting ions. Selectivity of SnO₂ fibers towards Co²⁺ in the presence of competing ions Na⁺ and Ca²⁺ was examined. The batch samples were prepared as reported

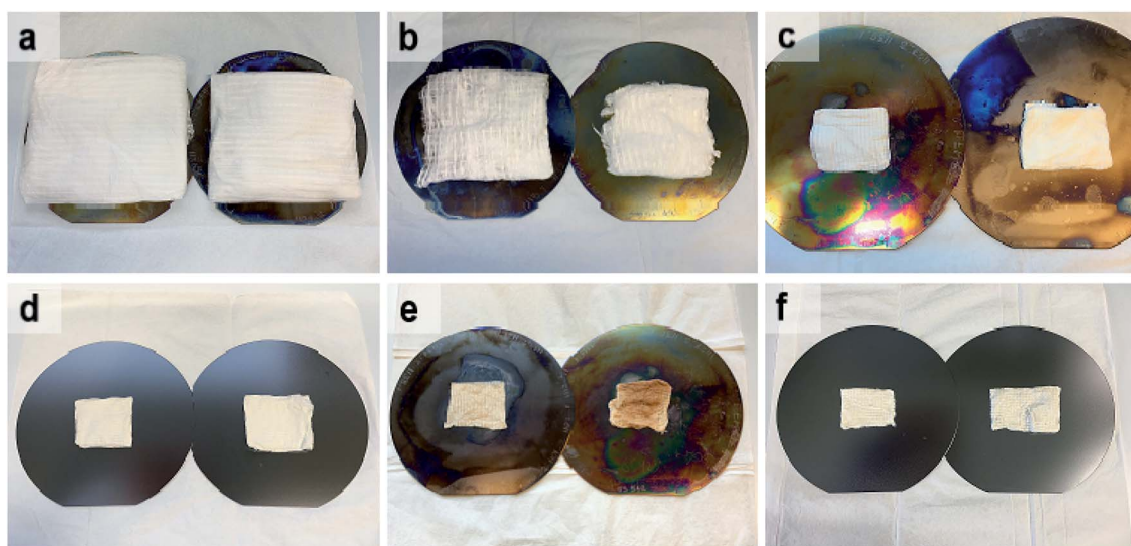


Fig. 1 As-electroblown SnO₂/PVP (a) and SnO₂/SiO₂/PVP (b) fibers, SnO₂ fibers calcined at 400 °C (c) and 500 °C (d) as well as SnO₂/SiO₂ composite fibers calcined at 400 °C (e) and 500 °C (f) on 150 mm silicon wafers. In the calcination, a heating rate of 1 °C min⁻¹ was used. The calcined fiber mats are shrunken due to the removal of the polymer.



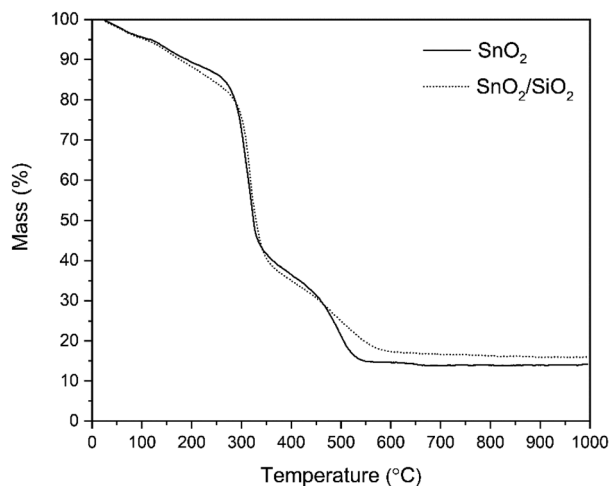


Fig. 2 TG curves of as-electroblown SnO_2/PVP and $\text{SnO}_2/\text{SiO}_2/\text{PVP}$ fibers in air (50 mol%) and N_2 (50 mol%, the purge gas).

above and the initial pH was adjusted to 6. Four concentrations of the competing ion were used: 0.001, 0.01, 0.1 and 1 M for NaCl , and 0.001, 0.01, 0.1 and 0.5 M for CaCl_2 . The $^{57}\text{Co}^{2+}$ removal by the fibers was calculated as the distribution coefficient K_d .

2.4.4 EDX analysis. Elemental analysis of the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers after adsorption of Co^{2+} including elemental mapping was conducted. For the analysis, 20 mg of fibers was

weighed into 20 mL scintillation vials and 10 mL of 0.01 M NaNO_3 solution containing 1 mM non-radioactive Co^{2+} ($\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$) was added into the vials. The pH of the solution was adjusted to 6. The samples were equilibrated for 24 h, phase separated and dried in an oven at 70 °C overnight.

3 Results and discussion

3.1 Electron microscopy and TG analysis of the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers

Photographs of as-electroblown SnO_2/PVP and $\text{SnO}_2/\text{SiO}_2/\text{PVP}$ fibers as well as SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers calcined at 400 and 500 °C are shown in Fig. 1. The calcined fiber mats have shrunk because the polymer has been removed. Except for the $\text{SnO}_2/\text{SiO}_2$ fibers calcined at 400 °C, the colour of the calcined fibers is white, which implies no major amounts of carbon residues *i.e.* efficient combustion of the polymer. In regard to the $\text{SnO}_2/\text{SiO}_2$ fibers with a brownish hue, both the lower calcination temperature of 400 °C and the presence of SiO_2 probably cause incomplete combustion of the PVP. Based on TG analysis,⁵³ combustion of the bare PVP fibers is not complete until at 700 °C and hence there may be carbon residues in the calcined fibers although not necessarily to a visible extent. The amorphous SiO_2 phase in the composite fibers may be more prone to retain the amorphous polymer compared to the crystalline SnO_2 fibers and SiO_2 and PVP are known to form hydrogen bonds.^{54,55}

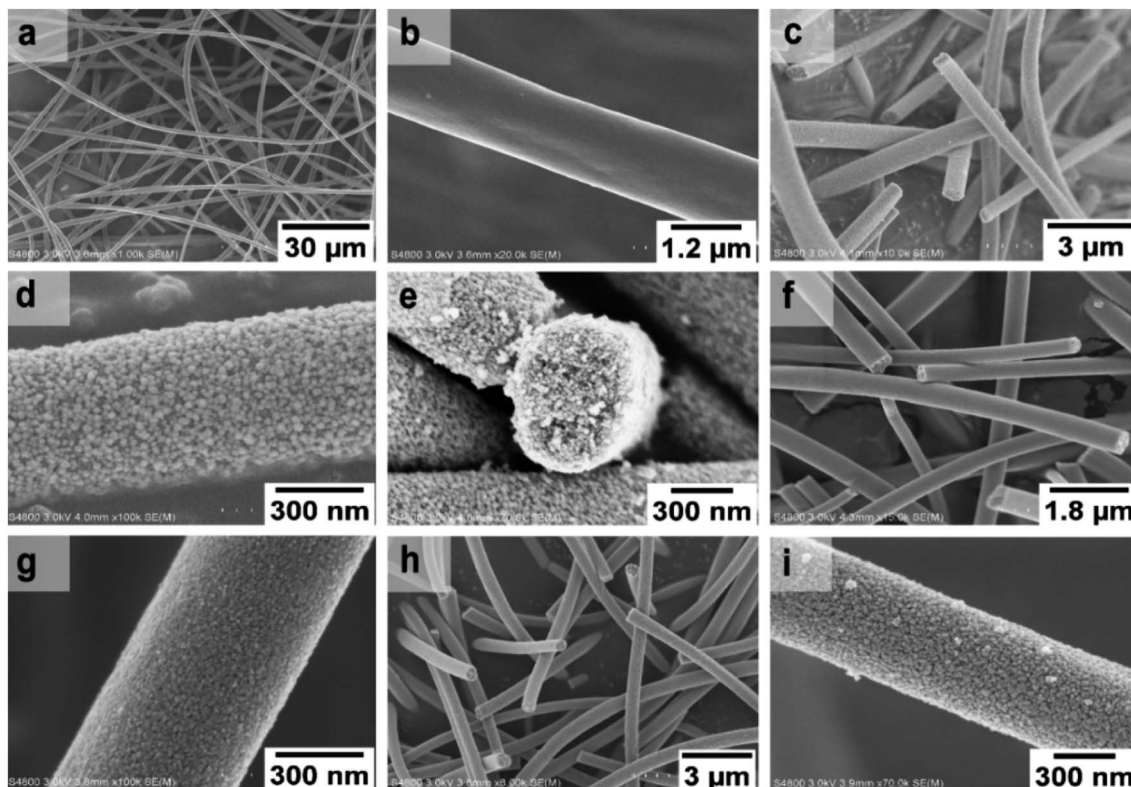


Fig. 3 FESEM images at low (a) and high (b) magnification of uncalcined SnO_2/PVP fibers and SnO_2 fibers calcined at 400 °C (c to e), 450 °C (f and g) and 500 °C (h and i).



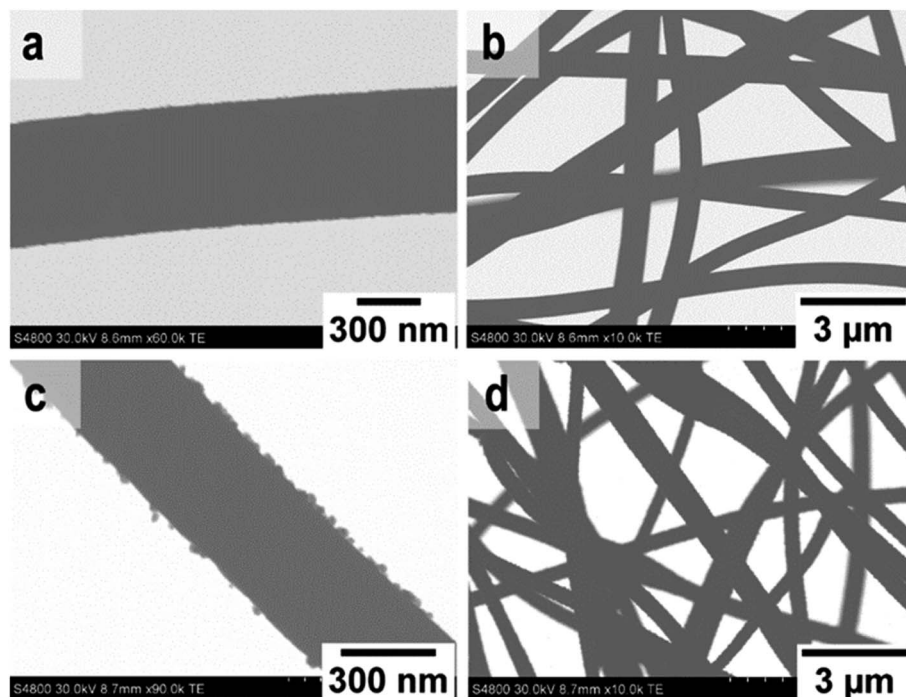


Fig. 4 TE images at low and high magnification of SnO₂ fibers (a and b) and SnO₂/SiO₂ composite fibers (c and d) calcined at 500 °C.

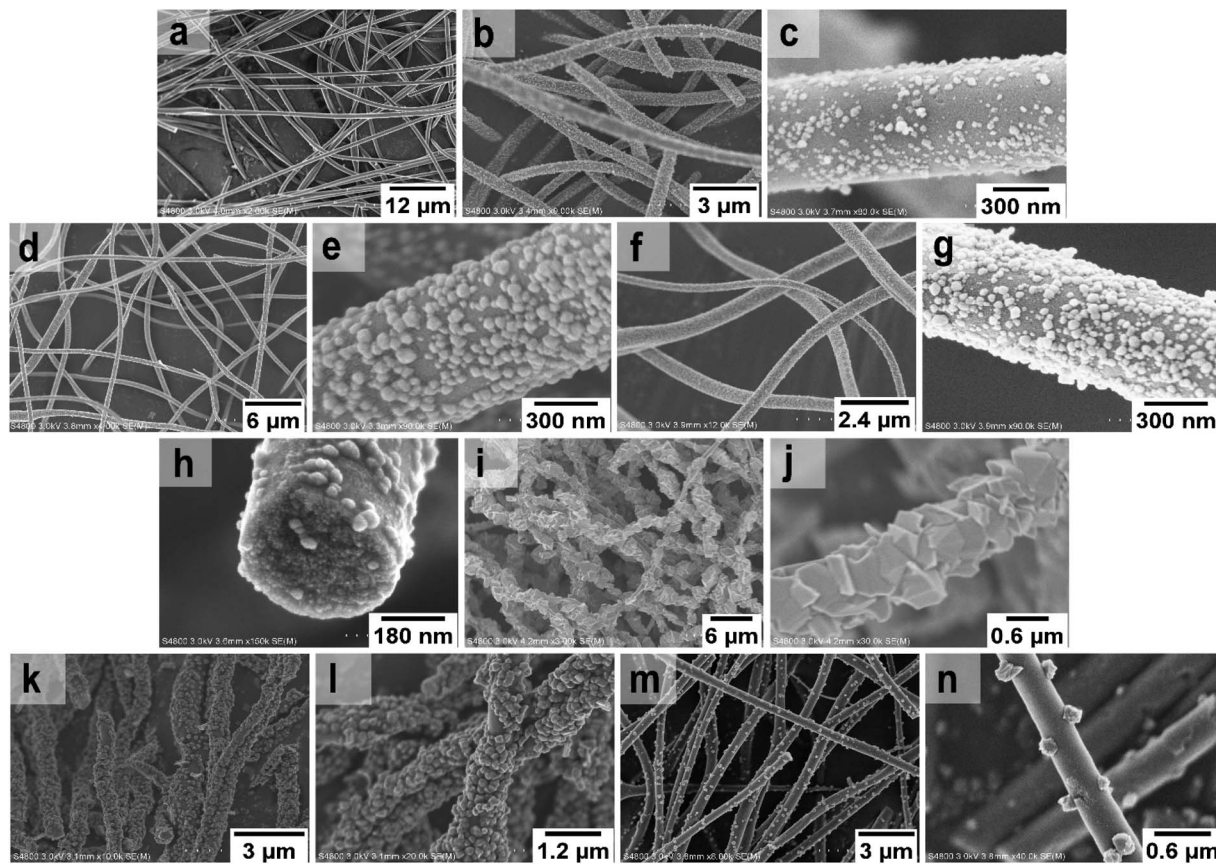


Fig. 5 FESEM images of uncalcined SnO₂/SiO₂/PVP fibers (a), and images at low and high magnification of SnO₂/SiO₂ composite fibers calcined at 400 °C (b and c), 450 °C (d and e) and 500 °C (f–j) with a heating rate of 1 °C min⁻¹ as well as SnO₂/SiO₂ composite fibers calcined at 400 °C with heating rates of 5 °C min⁻¹ (k and l) and 10 °C min⁻¹ (m and n).



According to the TG analysis of the as-electroblown SnO₂/PVP and SnO₂/SiO₂/PVP fibers (Fig. 2), the combustion of the PVP in the SnO₂/SiO₂ composite fibers is not complete until at 600 °C whereas it is complete at 500 °C in the case of the bare SnO₂ fibers. Furthermore, it can be seen that the mass of pure SnO₂ is *ca.* 15% and the mass of pure SnO₂/SiO₂ is *ca.* 20% of the mass of the as-electroblown material. The measured weights of the calcined SnO₂ and SnO₂/SiO₂ fibers are in line with this, also for the lowest calcination temperature of 400 °C. This implies a more effective combustion of the polymer in the calcining furnace which is probably due to a higher oxygen content as the furnace atmosphere comprises air instead of air/N₂ mixture and a long 4 hour duration of the calcination. Owing to the presence of SiO₂, the calcined SnO₂/SiO₂ composite fiber mats were elastic and could be bent without causing fractures to them. The SnO₂ fiber mats were more brittle and fractured when bent. The differences in elasticity between the SnO₂ and SnO₂/SiO₂ fibers are demonstrated in the video “Bending experiments with the SnO₂ and SnO₂/SiO₂ fibers”.

FESEM images of the uncalcined SnO₂/PVP fibers and SnO₂ fibers calcined at 400, 450 and 500 °C are presented in Fig. 3. The average diameter of the uncalcined fibers containing the polymer was 1.5 μm whereas the average diameter of the calcined fibers was 560 nm. The morphology of the calcined SnO₂ fibers was similar regardless of the calcination temperature. The fibers seemed to have a uniform structure that consisted of roundish grains approximately 20 nm in diameter, although there was some variation between individual grains. A similar granular structure has also been reported for both dense^{21,23,26,50} and hollow^{22,56} electrospun SnO₂ nanofibers. The

dense character of the fibers in the current study is proved by the FESEM image of a fiber cross-section (Fig. 3e) and was verified by imaging with transmitted electrons (Fig. 4a and b).

FESEM images of the uncalcined SnO₂/SiO₂/PVP fibers and SnO₂/SiO₂ fibers calcined at 400, 450 and 500 °C with a heating rate of 1 °C min⁻¹ as well as at 400 °C with heating rates of 5 and 10 °C min⁻¹ are presented in Fig. 5. The average diameters of the uncalcined and calcined fibers were 1.2 μm and 580 nm, respectively. Irrespective of the calcination temperature, the

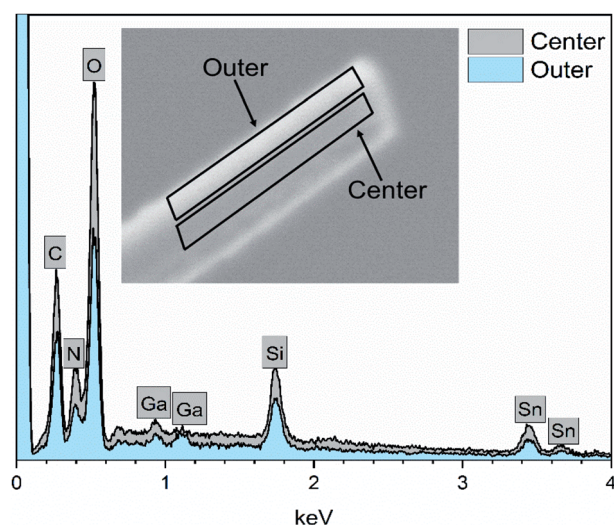


Fig. 7 EDX spectra of a single SnO₂/SiO₂ composite fiber calcined at 500 °C measured at the longitudinal cross-section surface. Ga peaks are due to the ion beam used for fiber milling.

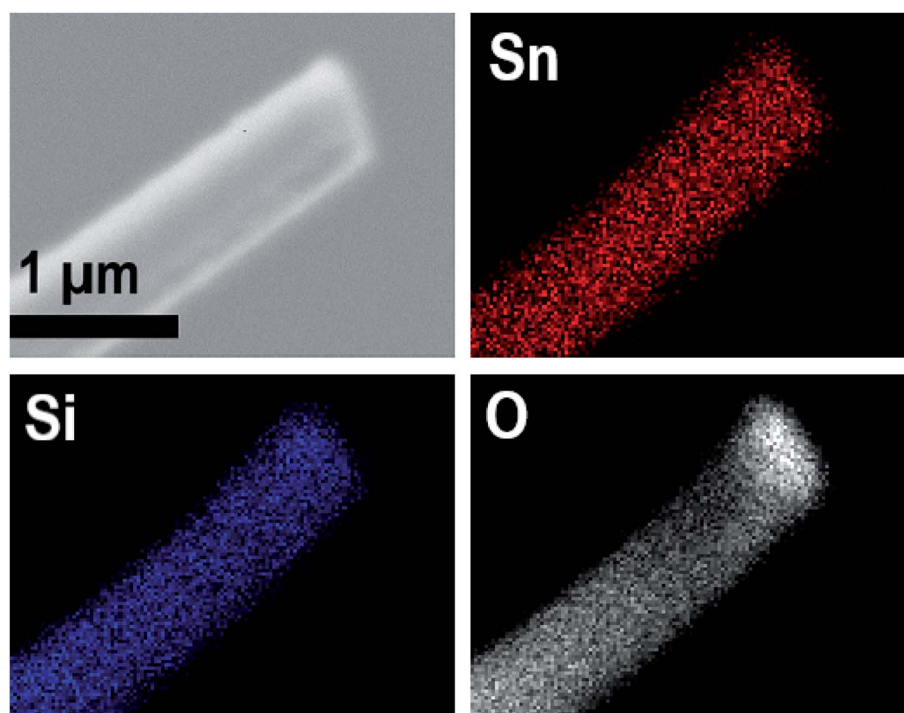


Fig. 6 EDX elemental maps of a single SnO₂/SiO₂ composite fiber calcined at 500 °C.



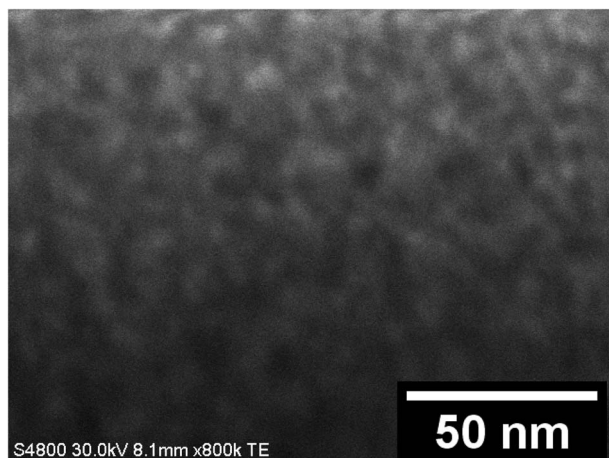


Fig. 8 A TEM image of the inner section of a single $\text{SnO}_2/\text{SiO}_2$ composite fiber calcined at $500\text{ }^\circ\text{C}$.

appearance of the fibers calcined with a heating rate of $1\text{ }^\circ\text{C min}^{-1}$ was the same (Fig. 5b–h). The fibers seemed to have a smooth core that was rather sparsely covered with roundish grains approximately 30 nm in diameter. A rather similar structure comprising a smooth core and granular surface has been reported previously for $\text{SnO}_2/\text{SiO}_2$ nanofibers, although with a higher proportion of Si from 50 to 83 mol% compared to ours of 25 mol%.²⁴ As for the fibers calcined with the faster heating rates, they also had a smooth core sparsely covered with grains that were quite angular in shape and rather large (Fig. 5k–n). The average diameter of the grains was *ca.* 200 nm in the fibers calcined with heating rates of 5 and $10\text{ }^\circ\text{C min}^{-1}$,

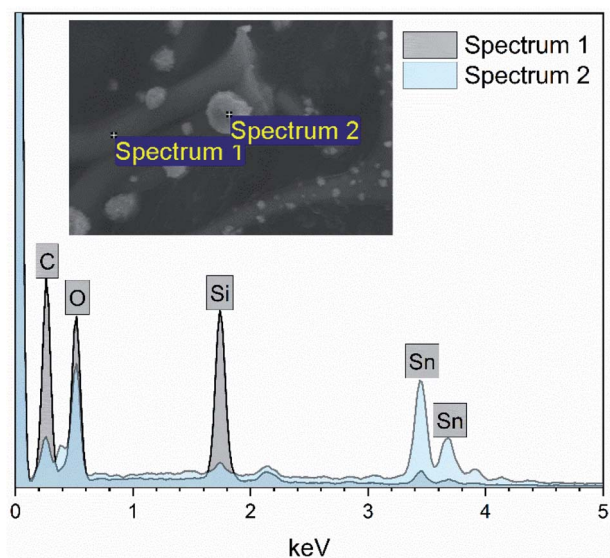


Fig. 10 EDX spectra measured at the smooth part (spectrum 1) and surface grain (spectrum 2) of a single $\text{SnO}_2/\text{SiO}_2$ composite fiber calcined at $400\text{ }^\circ\text{C}$ with a heating rate of $10\text{ }^\circ\text{C min}^{-1}$.

although there was some variation between individual fibers. Among the fibers calcined at $500\text{ }^\circ\text{C}$ with the slow heating rate of $1\text{ }^\circ\text{C min}^{-1}$, there were also some fibers with a special morphology with the fiber core covered with large, irregular grains of 500 nm to $1\text{ }\mu\text{m}$ in diameter (Fig. 5i and j). FESEM (Fig. 5h) and TEM imaging (Fig. 4c and d) confirmed that the structure of the $\text{SnO}_2/\text{SiO}_2$ composite fibers was dense like that of the bare SnO_2 fibers.

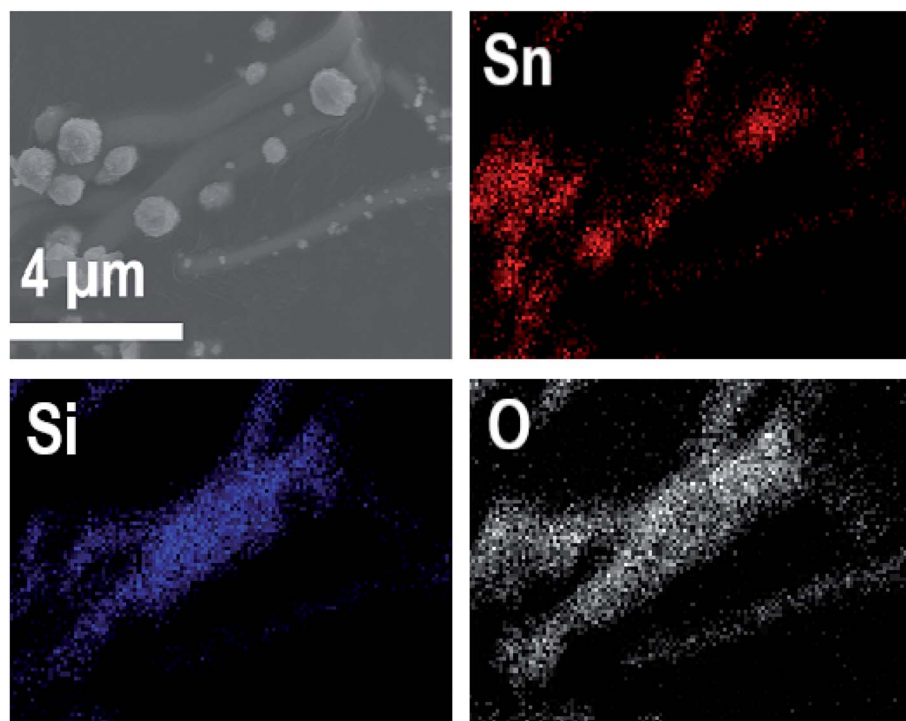


Fig. 9 EDX elemental maps of $\text{SnO}_2/\text{SiO}_2$ composite fibers calcined at $400\text{ }^\circ\text{C}$ with a heating rate of $10\text{ }^\circ\text{C min}^{-1}$.



3.2 EDX analysis of the SnO₂ and SnO₂/SiO₂ fibers

EDX spectra of the SnO₂ fibers are shown in Fig. S1 and S3.† As for the SnO₂/SiO₂ composite fibers, the quantitative Sn : Si EDX results were 77 at% : 23 at% and 80 at% : 20 at% for the fibers calcined at 400 and 500 °C, respectively (Fig. S2 and S4†). Considering possible variation in sample homogeneity, the results are as expected and indicate a successful synthesis. Owing to the difficult quantification of light elements with EDX and because detected carbon may also originate in the environment, the amount of residual carbon in the fibers could not be reliably determined.

EDX elemental maps of a single SnO₂/SiO₂ fiber calcined at 500 °C are shown in Fig. 6. EDX spectra recorded from both inner and outer section of the smooth part of the fiber are presented in Fig. 7. A TE image of a thin inner section of the fiber is shown in Fig. 8. For the EDX measurement, a portion of the fiber was FIB milled away to expose a flat longitudinal cross-section surface from which the spectra were recorded. For the TE imaging, equal longitudinal portions from both sides of the fiber were FIB milled away leaving a thin slab in the middle of which the imaging was done. The elemental maps prove the presence of Sn, Si and O in the fiber but on the basis of them it is difficult to tell any difference between the distribution or the concentration of Sn and Si in the fiber. The EDX spectra of the center and outer smooth parts of the fiber are quite identical and they both show the presence of Sn. This implies that the core of the SnO₂/SiO₂ composite fibers in this study is different from the single-phase amorphous SiO₂ core reported for the SnO₂/SiO₂ composite fibers with a higher proportion of Si.²⁴ The TE image from inside the fiber supports this as it suggests that the inner part of the fiber is not homogeneous but consists of distinct nanoscale domains, possibly small SnO₂ grains embedded in a SiO₂ matrix. Previously it has been proved that more than *ca.* 1 mol% of SnO₂ cannot be dissolved in SiO₂ but SnO₂ forms crystalline nanoclusters dispersed in amorphous SiO₂.^{57,58}

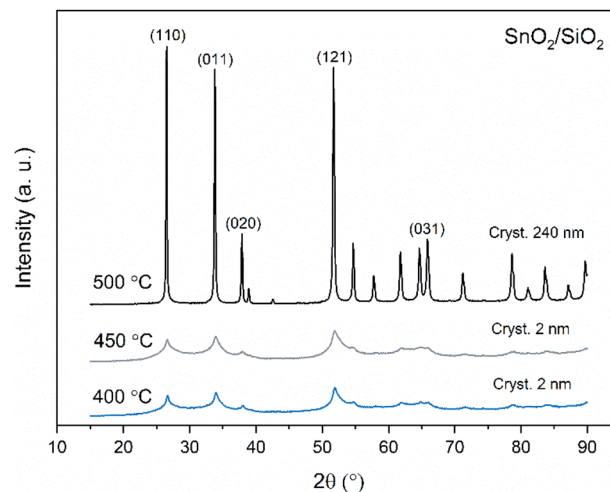


Fig. 12 X-ray diffraction patterns of the SnO₂/SiO₂ fibers calcined at 400, 450 and 500 °C with a heating rate of 1 °C min⁻¹.

EDX elemental maps of the SnO₂/SiO₂ composite fibers calcined at 400 °C with a heating rate of 10 °C min⁻¹ are presented in Fig. 9. As can be seen, Sn is more concentrated at the sites where the large grains are located while Si is more concentrated at the smooth part of the fibers. This implies that the large grains on the surface of the fibers consist primarily of SnO₂. More evidence is provided by EDX spectra (Fig. 10) recorded at the smooth part (spectrum 1) and at a surface grain (spectrum 2) of the fibers. In the spectrum recorded at the grain there are strong signals for Sn and a very weak signal for Si while in the spectrum recorded at the smooth part the case is the opposite, *i.e.* the signal for Si is very intense and the signals for Sn are quite low.

3.3 Crystal structure of the SnO₂ and SnO₂/SiO₂ fibers

X-ray diffraction patterns of both bare SnO₂ and composite SnO₂/SiO₂ fibers are presented in Fig. 11–13. All the fibers

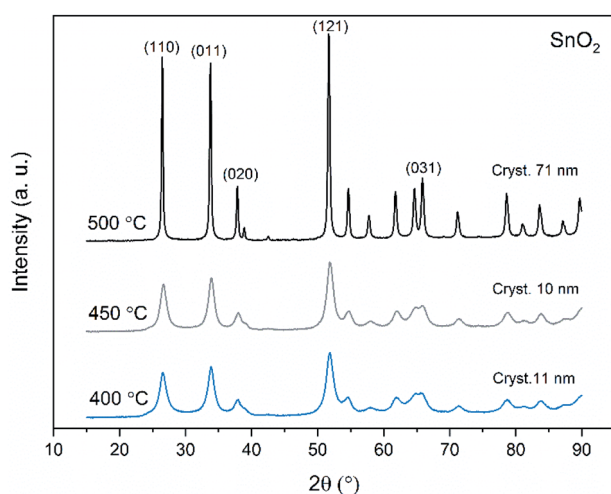


Fig. 11 X-ray diffraction patterns of the SnO₂ fibers calcined at 400, 450 and 500 °C. Cryst. refers to average crystallite size in this and Fig. 12 and 13.

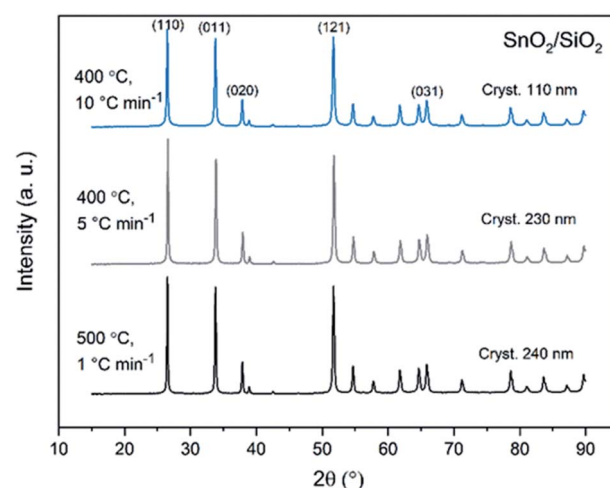


Fig. 13 X-ray diffraction patterns of the SnO₂/SiO₂ fibers calcined at 500 °C with a heating rate of 1 °C min⁻¹ and at 400 °C with heating rates of 5 and 10 °C min⁻¹.



irrespective of the calcination temperature or heating rate have the tetragonal rutile structure. The presence of SiO₂ doesn't seem to affect the crystalline phase nor the lattice parameters compared to bare SnO₂ fibers but instead it affects the crystallite size. Interestingly, in both SnO₂ and SnO₂/SiO₂ fibers there were crystallites of two different sizes for which the weight ratios were determined (Table 1). Refining the XRD data with a single crystallite size led to unsatisfactory results while clear improvement was achieved by using two different crystallite sizes with varying weight ratios. This was pronounced with the composite fibers. However, the bimodal size distribution may not be sufficient and a more complex distribution is probable, but from the perspective of this study it is adequate to show that there are at least two different crystallite sizes present. The fraction of one size was always much higher than that of the other size, ranging from 63 to 93 wt%. In all the fibers the larger crystallites were predominant over the smaller ones, except for the composite fibers calcined at 400 and 450 °C with a heating rate of 1 °C min⁻¹ (Fig. 5b–e). The sizes of the proportionally larger crystallites varied between 10 and 243 nm; the fibers calcined at 400 and 450 °C with slow heating possessed the smallest of them and the fibers calcined at 500 °C or at 400 °C with fast heating the largest of them. The sizes of the proportionally smaller crystallites varied from 2 to 15 nm and also here the size followed calcination temperature and heating rate: the fibers calcined at 400 and 450 °C with slow heating possessed the smallest of them and the fibers calcined at 500 °C or at 400 °C with fast heating the largest of them. Rietveld refined X-ray diffractograms of SnO₂ and SnO₂/SiO₂ fibers calcined at 500 °C are presented in Fig. S13 and S14.†

The crystallites in the SnO₂ fibers calcined at 400 and 450 °C are in average 11 and 10 nm in size, respectively, whereas in the corresponding SnO₂/SiO₂ fibers they are only 2 nm in size. The silica in the composite fibers probably acts as a crystal growth inhibitor, as is also the case with SnO₂ nanoparticles embedded in a SiO₂ network.⁵⁹ It should be noted though that the refinement results of the broad bumps in this data (Fig. 12) leading to below 10 nm sized crystallites represent rather the lack of long range order than an actual precise size, since severe overlapping of the peaks showing poor crystallinity undermines the accuracy of the results. In both the bare and composite fibers, the average crystallite size increases substantially as the calcination temperature is increased and in the composite fibers the increase is quite steep (Table 1). It is known that the crystallite size of ceramic submicron fibers depends on the calcination temperature.⁴⁴ In the bare SnO₂ fibers, the crystallite growth is probably accompanied with a grain growth at the higher temperatures²¹ although no prominent difference in grain size between the fibers calcined at 400, 450 and 500 °C was observed on the basis of the FESEM analysis (Fig. 3). As for the composite fibers calcined at 500 °C with a heating rate of 1 °C min⁻¹ and at 400 °C with heating rates of 5 and 10 °C min⁻¹, both the high calcination temperature and fast heating rates could explain the large crystallite sizes of 243, 230 and 111 nm, respectively (Table 1). It is likely that the large, irregular grains on the fibers (Fig. 5i–n) consist mainly of SnO₂ that is known to form various shapes in nanoscale.^{50,60} High calcination temperatures and fast

Table 1 Average crystallite sizes of the SnO₂ and SnO₂/SiO₂ fibers and weight ratios of small (low crystallinity) and large (high crystallinity) crystallites. Calc. param. refer to calcination parameters, *i.e.* calcination temperature and heating rate. The calcination duration was 4 h

Fibers	Calc. param.	Crystallinity	Av. cryst. size (nm)	Wt%
SnO ₂	400 °C, 1 °C min ⁻¹	High/low	11/3	63/37
SnO ₂	450 °C, 1 °C min ⁻¹	High/low	10/2	93/7
SnO ₂	500 °C, 1 °C min ⁻¹	High/low	71/12	87/13
SnO ₂ /SiO ₂	400 °C, 1 °C min ⁻¹	Low/high	2/14	83/17
SnO ₂ /SiO ₂	450 °C, 1 °C min ⁻¹	Low/high	2/11	80/20
SnO ₂ /SiO ₂	500 °C, 1 °C min ⁻¹	High/low	243/11	79/21
SnO ₂ /SiO ₂	400 °C, 5 °C min ⁻¹	High/low	230/15	74/26
SnO ₂ /SiO ₂	400 °C, 10 °C min ⁻¹	High/low	111/9	74/26

heating rates may influence both the size and shape of crystals,⁶¹ which is reflected in the morphology of grains comprising the crystals.

The presence of crystallites of two different sizes in the SnO₂/SiO₂ composite fibers is reasonable in the light of the electron microscopy and EDX analysis results. Based on EDX analysis, both Sn and Si are present in the core of the composite fibers but the surface grains consist primarily of SnO₂ (Fig. 7, 9 and 10). It is seen in the TEM image (Fig. 8) that the core of the fibers comprises different nanoscale domains probably of SnO₂ and SiO₂. As discussed, SiO₂ in the core of the fibers presumably prevents SnO₂ crystallite growth. This might cause the presence of small SnO₂ crystallites and grains in the core of the fibers. The SnO₂ crystallites can grow more freely on the surface of the fibers resulting in larger crystallite and grain sizes. For the SnO₂/SiO₂ fibers calcined at 400 and 450 °C with a heating rate of 1 °C min⁻¹ the relatively high amount of small 2 nm crystallites (83 and 80 wt%, respectively, Table 1) and the moderate size of the larger crystallites (14 and 11 nm, respectively, Table 1) are probably due to the low calcination temperature and slow heating rate. As regards the bare SnO₂ fibers, the presence of the smaller crystallites (3, 2 and 12 nm for the fibers calcined at 400, 450 and 500 °C, respectively, Table 1) might be at least partly due to some PVP residues hindering crystal growth. This may especially be the case for the fibers calcined at the lowest temperature of 400 °C for which the fraction of the small crystallites, 37 wt%, is the highest of the SnO₂ fibers. The average crystallite sizes of 11, 10 and 12 nm (SnO₂, Table 1) as well as 14, 11 and 11 nm (SnO₂/SiO₂, Table 1) for the fibers calcined at 400, 450 and 500 °C with slow heating, respectively, coincide with some SnO₂ grain sizes seen in the FESEM images (Fig. 3d, e, g, i and 5c, e, g). Previously it has been found that the crystal size of granular polycrystalline SnO₂ nanofibers is approximately the same as the size of the SnO₂ grains meaning that each SnO₂ grain is possibly a single crystal.²⁶ Perhaps this is the case for some SnO₂ grains in the fibers of this study as well.

3.4 Specific surface area, pore volume and pore size analysis of the SnO₂ fibers

Specific surface area of the SnO₂ fibers calcined at 500 °C was analysed by the Brunauer–Emmett–Teller (BET) method and



total pore volume, pore size distribution and average pore diameter by the Barrett–Joyner–Halenda (BJH) method using nitrogen gas adsorption and desorption. The specific surface area, total pore volume and average pore size of the fibers were $41 \text{ m}^2 \text{ g}^{-1}$, $0.15 \text{ cm}^3 \text{ g}^{-1}$ and 10 nm , respectively. The surface area of the fibers is somewhat higher than that found in literature for both dense⁵¹ and hollow^{56,62} electrospun SnO_2 fibers calcined at 500 or 600 °C, 6.9 to $36 \text{ m}^2 \text{ g}^{-1}$. The same is true of the total pore volume, as a smaller pore volume of $0.079 \text{ cm}^3 \text{ g}^{-1}$ has been reported by Mudra *et al.*⁶² while a larger average pore size of 16 nm has been reported by Xia *et al.*⁵⁶ The surface area of the SnO_2 fibers is quite high considering the large 71 nm crystallites, as large crystallites commonly result in a small surface area. For ZrO_2 fibers of almost the same diameter (570 nm) crystallites of 9 and 63 nm produced surface areas of 14 and $1.7 \text{ m}^2 \text{ g}^{-1}$, respectively.⁴⁴ The large surface area of the SnO_2 fibers of this study can probably be explained by the granular and porous structure. The average pore size of 10 nm seems to match quite well with the occasional interstices between the grains of the fibers (Fig. 3d, e, g and i). N_2 adsorption and desorption isotherms and pore size distribution of the SnO_2 fibers are presented in Fig. S15 and S16,[†] respectively.

3.5 Co^{2+} uptake by the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers

The adsorption ability of electroblown submicron fibers is greatly affected by their crystal structure and size, morphology and the amount of possible polymer residues. These characteristics, in turn, depend on the calcination temperature.⁴⁴ Therefore, we investigated the effects of calcination temperature and heating rate on the Co^{2+} uptake by the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers. As revealed by Table 2, all the fibers have a good Co^{2+} uptake performance with an average uptake of 99.71% and 99.53% for SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers, respectively. However, there were some differences between the fibers. As seen from the K_d values, the bare SnO_2 fibers have somewhat better Co^{2+} uptake than the $\text{SnO}_2/\text{SiO}_2$ composite fibers. The SiO_2 in the composite fibers seems to impair their Co^{2+} uptake, although chemisorption of Co^{2+} on SiO_2 is known.⁶³ In this study, however, the crystalline SnO_2 appears to be the major adsorbent. The high surface area of the SnO_2 fibers is likely to enhance the adsorption even more as it should provide plenty of

adsorption sites. EDX elemental maps and spectra of the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers calcined at 400 and 500 °C with a heating rate of 1 °C min^{-1} after adsorption of Co^{2+} are shown in Fig. S5–S12.[†]

In regard to the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers calcined with a heating rate of 1 °C min^{-1} , the calcination temperature had a different effect on their Co^{2+} uptake performance. Calcination temperatures of 500 and 400 °C produced the best adsorption performance for SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers, respectively (Table 2 and Fig. 14). Lower calcination temperature tends to produce smaller crystallites and thus increase the surface area of the material, which, in turn, improves its adsorption properties.⁴⁴ Therefore it is surprising that the SnO_2 fibers calcined at 500 °C possessing 71 nm crystallites perform better than the SnO_2 fibers calcined at 400 and 450 °C possessing 11 and 10 nm crystallites, respectively. One possible reason is that the lower calcination temperatures leave some PVP residues in the material blocking some of the adsorption sites. However, despite their different crystallite size, the K_d value of the SnO_2 fibers calcined at 450 °C is almost as high as that of the fibers calcined at 500 °C. The crystallite size doesn't seem to be a determining factor in the Co^{2+} uptake performance of the SnO_2 fibers.

In the $\text{SnO}_2/\text{SiO}_2$ fibers calcined with a heating rate of 1 °C min^{-1} the Co^{2+} uptake correlated better with the crystallite size, since the fibers calcined at 400 °C and with the smallest 2 nm crystallites performed the best (Table 2 and Fig. 14). It should be noted that although the fibers calcined at 450 °C also possessed 2 nm crystallites, their sorption and K_d values were lower. One reason for the weaker Co^{2+} uptake performance of the $\text{SnO}_2/\text{SiO}_2$ fibers calcined at 450 and 500 °C with a heating rate of 1 °C min^{-1} could be that due to the higher calcination temperature, they have fewer surface hydroxyl groups that are assumed to play a significant role in the uptake process. As for the fibers calcined at 500 °C, the peculiar morphology of some

Table 2 $^{57}\text{Co}^{2+}$ uptake by the SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers in 0.01 M NaNO_3 at pH 6.0. Calc. param. refer to calcination parameters, i.e. calcination temperature and heating rate. The calcination duration was 4 h

Fibers	Calc. param.	Uptake (%)	K_d (mL g^{-1})
SnO_2	400 °C, 1 °C min^{-1}	99.51 ± 0.02	$173\,000 \pm 3000$
SnO_2	450 °C, 1 °C min^{-1}	99.81 ± 0.02	$272\,000 \pm 4000$
SnO_2	500 °C, 1 °C min^{-1}	99.82 ± 0.01	$281\,000 \pm 5000$
$\text{SnO}_2/\text{SiO}_2$	400 °C, 1 °C min^{-1}	99.69 ± 0.02	$163\,000 \pm 3000$
$\text{SnO}_2/\text{SiO}_2$	450 °C, 1 °C min^{-1}	99.18 ± 0.02	$58\,100 \pm 1000$
$\text{SnO}_2/\text{SiO}_2$	500 °C, 1 °C min^{-1}	99.28 ± 0.02	$69\,700 \pm 1300$
$\text{SnO}_2/\text{SiO}_2$	400 °C, 5 °C min^{-1}	99.70 ± 0.02	$168\,000 \pm 3000$
$\text{SnO}_2/\text{SiO}_2$	400 °C, 10 °C min^{-1}	99.79 ± 0.02	$234\,000 \pm 4000$

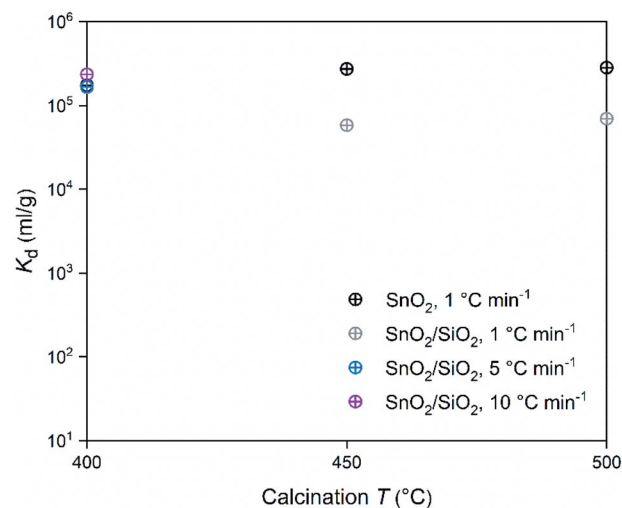


Fig. 14 Distribution coefficient of 30 Bq mL^{-1} $^{57}\text{Co}^{2+}$ on SnO_2 and $\text{SnO}_2/\text{SiO}_2$ fibers as a function of calcination temperature and heating rate in 0.01 M NaNO_3 at pH 6.0 (2 g L^{-1} SnO_2 or $\text{SnO}_2/\text{SiO}_2$).



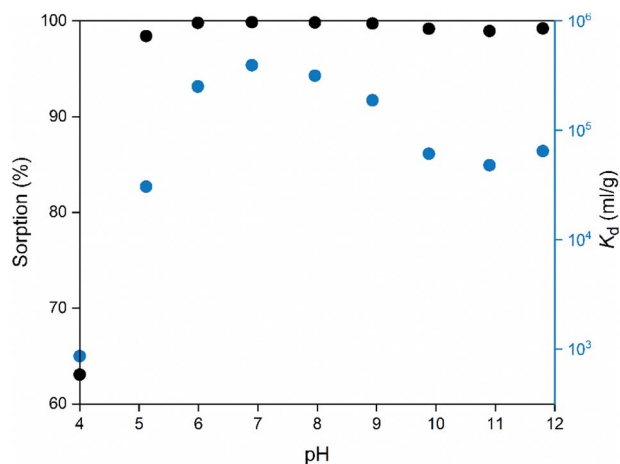


Fig. 15 Sorption and distribution coefficient of $30 \text{ Bq mL}^{-1} {}^{57}\text{Co}^{2+}$ on SnO_2 fibers calcined at 500°C as a function of pH in 0.01 M NaNO_3 ($2 \text{ g L}^{-1} \text{ SnO}_2$).

of the fibers (Fig. 5i and j) and large crystallite size of 243 nm that decreases the surface area, may also impair their Co^{2+} uptake performance. The $\text{SnO}_2/\text{SiO}_2$ fibers calcined at 400°C with the fast heating rates of 5 and $10^\circ\text{C min}^{-1}$ exhibited excellent Co^{2+} uptake despite their large crystallite sizes. This might be explained by the roundish shape and quite sparse distribution of their SnO_2 grains (Fig. 5k–n) compared to the angular and rather tightly packed SnO_2 grains of the fibers calcined at 500°C (Fig. 5i and j). The roundish and sparsely distributed surface SnO_2 grains might be better accessible to adsorbing ions in an aqueous solution. Moreover, the lower target temperature of 400°C may cause more hydroxyl groups remaining on the surface of the fibers enhancing the uptake of Co^{2+} .

The pH of the solution may have a major impact on the uptake properties of a material because it affects both the

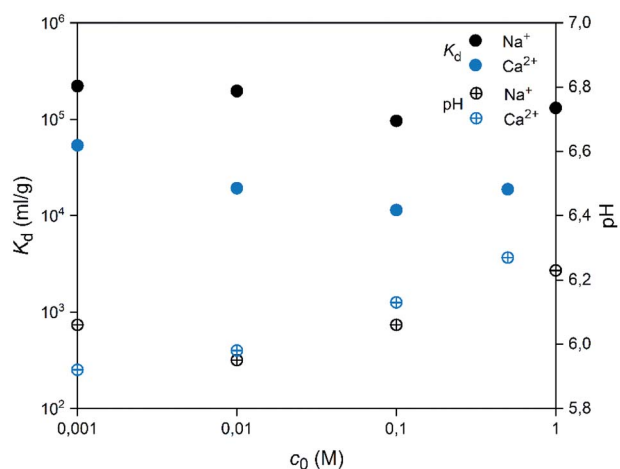


Fig. 16 Effects of concentrations of competing ions Na^+ or Ca^{2+} on the uptake of $30 \text{ Bq mL}^{-1} {}^{57}\text{Co}^{2+}$ by SnO_2 fibers calcined at 500°C . The initial pH was 6 and the pH at the end of the experiment is shown ($2 \text{ g L}^{-1} \text{ SnO}_2$).

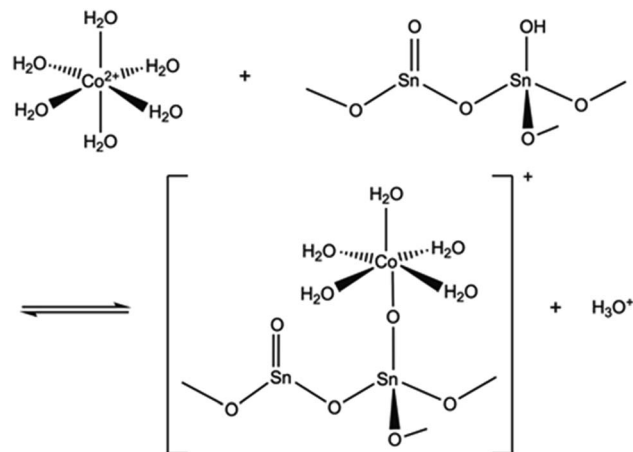


Fig. 17 An assumed ion exchange reaction between Co^{2+} and a proton of surface hydroxyl group of SnO_2 .⁵

speciation of the adsorbate and the surface charge of the adsorbent. As for Co^{2+} , it exists as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the pH range of 2 to 9.⁶⁴ It can be inferred from the K_d values (Fig. 15), that the SnO_2 fibers adsorb Co^{2+} the best in the neutral to mildly basic pH region, from pH 6 to 9, and they reach the highest K_d value at the pH of 7. This is in accordance with previous research⁷ and promising for the use of the fibers in purification of NPP waste waters, since the pH of the primary coolant water in NPPs is about 7. The point of zero charge (pH_{pzc}) of SnO_2 lies at a pH of 4,⁷ and above this pH the surface charge of pure SnO_2 is negative. The Co^{2+} uptake by the rutile structured SnO_2 is evidently based on electrostatic forces⁶⁵ which explains the good uptake in the neutral to mildly basic pH region where the charges of the Co^{2+} species and the surface of the SnO_2 fibers are the opposite. The adsorption of Co^{2+} on hydrous SnO_2 most probably occurs *via* substitution of H^+ ions of the surface water molecules or hydroxyl groups by Co^{2+} ions.^{4–6,66} An assumed ion exchange reaction between hydrous SnO_2 and Co^{2+} is illustrated in Fig. 17.⁵

A good selectivity of the adsorbent is crucial when a trace amount of a specific ion is separated from a solution containing much higher concentrations of other ions. Thus, we examined how the Co^{2+} adsorption on SnO_2 fibers is influenced by Na^+ and Ca^{2+} ions that are among the most common cations in natural and nuclear waste waters with concentrations of 0.47 and 0.01 M in sea water, respectively.⁶⁷ As Fig. 16 reveals, the Co^{2+} uptake remains high in the presence of Na^+ ions irrespective of their concentration. By contrast, there is a marked weakening of the uptake in the CaCl_2 solution at concentrations of 0.01 M or higher. With both of the ions, the interfering effect increases with concentration up to 0.1 M , after which there is no further decreasing effect on the uptake of Co^{2+} . The good selectivity of the fibers for Co^{2+} over Na^+ is promising for their use in decontamination of radioactive waste water, since Na^+ is the most abundant coexisting ion in nuclear waste effluents.⁶⁷



4 Conclusions

We have synthesized SnO₂ and composite SnO₂/SiO₂ submicron fibers with a Sn : Si molar ratio of 3 : 1 and studied the ability of the fibers to remove Co²⁺ from an aqueous solution. For the synthesis, a novel and efficient electroblowing method was used. The as-electroblown fibers were calcined in air at 400, 450 and 500 °C with varying heating rates in order to produce the desired ceramic material and to investigate the effect of calcination temperature and heating rate on the structure and Co²⁺ uptake by the fibers. The bare SnO₂ fibers had a granular structure in the tetragonal rutile phase with an average diameter of 560 nm. The SnO₂/SiO₂ composite fibers had a smooth core possibly comprising small SnO₂ grains in a SiO₂ matrix with large SnO₂ grains dispersed on the core and the average diameter of the fibers was 580 nm. The morphology of the surface SnO₂ grains of the composite fibers was dependent on the calcination temperature and heating rate. Irrespective of the calcination temperature, a heating rate of 1 °C min⁻¹ yielded surface SnO₂ grains that were roundish and *ca.* 30 nm in diameter. In the composite fibers calcined at 500 °C with a heating rate of 1 °C min⁻¹ or at 400 °C with faster heating rates of 5 and 10 °C min⁻¹ there were also some surface SnO₂ grains that were irregularly shaped and 200 nm to 1 μm in diameter.

Both bare SnO₂ and composite SnO₂/SiO₂ fibers had a high Co²⁺ uptake with SnO₂ fibers exhibiting slightly more efficient Co²⁺ separation. Calcination temperature and heating rate affected the Co²⁺ uptake by the fibers. Among the SnO₂ fibers, the fibers calcined at 500 °C performed the best. Among the composite fibers, the fibers calcined at 400 °C performed the best, and within them the fibers calcined with a heating rate of 10 °C min⁻¹ were superior to the fibers calcined with heating rates of 1 and 5 °C min⁻¹. The SnO₂/SiO₂ composite fibers were more elastic and durable and easier to handle than the somewhat brittle SnO₂ fibers which makes them ideal for use in flow-through separation columns. Overall, on the basis of our results, the mechanical strength of SnO₂ submicron fibers can be enhanced by adding a moderate amount of SiO₂ without compromising their adsorption performance too much. This approach might also be extended to other ceramic submicron and nanofibers that require improved mechanical properties in various applications.

Author contributions

J. P.: conceptualization, methodology, investigation, data curation, formal analysis, validation, visualization, writing – original draft. S. W.: investigation. S. L.: investigation, formal analysis. M. H.: formal analysis. M. V.: investigation, formal analysis. M. K.: investigation, formal analysis. T. H.: investigation, formal analysis. M. R.: funding acquisition, supervision, writing – review & editing. R. K.: funding acquisition, supervision, writing – review & editing.

Conflicts of interest

There are no conflicts of interest to declare.

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References

- 1 P. Yang, *The Chemistry of Nanostructured Materials*, World Scientific, Singapore, 2003.
- 2 A. Clearfield, *Inorganic Ion Exchange Materials*, CRC Press, Boca Raton, Florida, 1982.
- 3 E. Abdelkader, L. Nadjia and V. Rose-Noelle, *Int. J. Ind. Chem.*, 2016, 7, 53.
- 4 J. D. Donaldson and M. J. Fuller, *J. Inorg. Nucl. Chem.*, 1968, 30, 1083.
- 5 N. Jaffrezic-Renault, *J. Inorg. Nucl. Chem.*, 1978, 40, 539.
- 6 N. Z. Misak, E. I. Shabana, E. M. Mikhail and H. F. Ghoneimy, *React. Polym.*, 1992, 16, 261.
- 7 R. Răutiu and D. A. White, *Solvent Extr. Ion Exch.*, 1996, 14, 721.
- 8 D. A. White and R. Răutiu, *Chem. Eng. J.*, 1997, 66, 85.
- 9 A. Lagashetty and A. Venkataraman, *Bull. Mater. Sci.*, 2004, 27, 491.
- 10 R. Koivula and R. Harjula, *Presented in part at Waste management 2007 Conference*, Tucson, Arizona, USA, 25 Feb. to 1 Mar., 2007.
- 11 A. Nilchi, T. Shariati Dehaghan and S. Rasouli Garmarodi, *Desalination*, 2013, 321, 67.
- 12 K. Yogesh Kumar, T. N. Vinuth Raj, S. Archana, S. B. Benaka Prasad, S. Olivera and H. B. Muralidhara, *Journal of Water Process Engineering*, 2016, 13, 44.
- 13 S. Haq, W. Rehman, M. Waseem, M. Rehman and K. H. Shah, *J. Inorg. Organomet. Polym. Mater.*, 2020, 30, 1197.
- 14 R. E. Presley, C. L. Munsee, C.-H. Park, D. Hong, J. F. Wager and D. A. Keszler, *J. Phys. D: Appl. Phys.*, 2004, 37, 2810.
- 15 A. P. Rizzato, C. V. Santilli, S. H. Pulcinelli, Y. Messaddeq and P. Hammer, *J. Sol-Gel Sci. Technol.*, 2004, 32, 155.
- 16 Y. Wang, X. Jiang and Y. Xia, *J. Am. Chem. Soc.*, 2003, 125, 16176.
- 17 N. Pinna, G. Neri, M. Antonietti and M. Niederberger, *Angew. Chem., Int. Ed.*, 2004, 43, 4345.
- 18 Y. Wang, I. Ramos and J. J. Santiago-Aviles, *IEEE Sens. J.*, 2007, 7, 1347.
- 19 Y. Zhang, X. He, J. Li, Z. Miao and F. Huang, *Sens. Actuators, B*, 2008, 132, 67.
- 20 Y. Zhang, J. Li, G. An and X. He, *Sens. Actuators, B*, 2010, 144, 43.
- 21 J. Y. Park, K. Asokan, S.-W. Choi and S. S. Kim, *Sens. Actuators, B*, 2011, 152, 254.
- 22 R. Ab Kadir, Z. Li, A. Z. Sadek, R. Abdul Rani, A. S. Zoofakar, M. R. Field, J. Z. Ou, A. F. Chrimes and K. Kalantar-Zadeh, *J. Phys. Chem. C*, 2014, 118, 3129.
- 23 J. P. Santos, M. J. Fernández, J. L. Fontecha, D. Matatagui, I. Sayago, M. C. Horrillo and I. Gracia, *Sensors*, 2014, 14, 24231.



- 24 Y. Liu, P. Yang, J. Li, K. Matras-Postolek, Y. Yue and B. Huang, *RSC Adv.*, 2016, **6**, 13371.
- 25 S. Gunji, M. Jukei, Y. Shimotsuma, K. Miura, K. Suematsu, K. Watanabe and K. Shimano, *J. Mater. Chem. C*, 2017, **5**, 6369.
- 26 Z. Yang, G. Du, C. Feng, S. Li, Z. Chen, P. Zhang, Z. Guo, X. Yu, G. Chen, S. Huang and H. Liu, *Electrochim. Acta*, 2010, **55**, 5485.
- 27 G. Granger, C. Restoin, P. Roy, R. Jamier, S. Rougier, J.-R. Duclere, A. Lecomte, R. Dauliat and J.-M. Blondy, *Micro-Structured and Specialty Optical Fibres IV*, 2015, vol. 9507, p. 95070J.
- 28 H. Shan, X. Wang, F. Shi, J. Yan, J. Yu and B. Ding, *ACS Appl. Mater. Interfaces*, 2017, **9**, 18966.
- 29 L. A. Malik, A. Bashir, A. Qureshi and A. H. Pandith, *Environ. Chem. Lett.*, 2019, **17**, 1495.
- 30 S. Muthusarayanan, N. Sivarajasekar, J. S. Vivek, T. Paramasivan, M. Naushad, J. Prakashmaran, V. Gayathri and O. K. Al-Duaij, *Environ. Chem. Lett.*, 2018, **16**, 1339.
- 31 T. S. Anirudhan, F. Shainy and J. R. Deepa, *Chem. Ecol.*, 2019, **35**, 235.
- 32 M. Naushad, Z. A. AlOthman, Md. R. Awual, M. M. Alam and G. E. Eldesoky, *Ionics*, 2015, **21**, 2237.
- 33 K. Vijayaraghavan, S. Rangabhashiyam, T. Ashokkumar and J. Arockiaraj, *Sep. Sci. Technol.*, 2016, **51**, 2725.
- 34 S. Lönnrot, V. Suorsa, J. Paajanen, T. Hatanpää, M. Ritala and R. Koivula, *RSC Adv.*, 2019, **9**, 22355.
- 35 S. Ramakrishna, K. Fujihara, W.-E. Teo, T.-C. Lim and Z. Ma, *An Introduction to Electrospinning and Nanofibers*, World Scientific, Singapore, 2005.
- 36 S. Thenmozhi, N. Dharmaraj, K. Kadirvelu and H. Y. Kim, *Mater. Sci. Eng., B*, 2017, **217**, 36.
- 37 Elmarco, <https://elmarco.com/>, accessed January 2021.
- 38 Fanavaran Nano-Meghyas (Fnm co. Ltd.), <http://en.fnm.ir/>, accessed January 2021.
- 39 Inovenso Ltd., <https://www.inovenso.com/>, accessed January 2021.
- 40 J. L. Daristotle, A. M. Behrens, A. D. Sandler and P. Kofinas, *ACS Appl. Mater. Interfaces*, 2016, **8**, 34951.
- 41 Z. Huang, A. Kolbasov, Y. Yuan, M. Cheng, Y. Xu, R. Rojaee, R. Deivanayagam, T. Foroozan, Y. Liu, K. Amine, J. Lu, A. L. Yarin and R. Shahbazian-Yassar, *ACS Appl. Mater. Interfaces*, 2020, **12**, 16200.
- 42 W. Tutak, S. Sarkar, S. Lin-Gibson, T. M. Farooque, G. Jyotsnendu, D. Wang, J. Kohn, D. Bolikal and C. G. Simon Jr., *Biomaterials*, 2013, **34**, 2389.
- 43 E. N. Bolbasov, K. S. Stankevich, E. A. Sudarev, V. M. Bouzunik, V. L. Kudryavtseva, L. V. Antonova, V. G. Matveeva, Y. G. Anissimov and S. I. Tverdokhlebov, *Mater. Chem. Phys.*, 2016, **182**, 338.
- 44 J. Paajanen, S. Lönnrot, M. Heikkilä, K. Meinander, M. Kemell, T. Hatanpää, K. Ainassaari, M. Ritala and R. Koivula, *Nanoscale Adv.*, 2019, **1**, 4373.
- 45 I. C. Um, D. Fang, B. S. Hsiao, A. Okamoto and B. Chu, *Biomacromolecules*, 2004, **5**, 1428.
- 46 H.-Y. Hsiao, C.-M. Huang, Y.-Y. Liu, Y.-C. Kuo and H. Chen, *J. Appl. Polym. Sci.*, 2012, **124**, 4904.
- 47 WWW Table of Radioactive Isotopes, <http://nucleardata.nuclear.lu.se/toi/nuclide.asp?iZA=270060>, accessed January 2021.
- 48 R. Harjula, J. Lehto, A. Paajanen and L. Brodtkin, presented in part at Waste Management'99 Conference, Tucson, Arizona, USA 28 Feb.–4 Mar, 1999.
- 49 J. Holopainen and M. Ritala, *J. Eur. Ceram. Soc.*, 2016, **36**, 3219.
- 50 S. Santibenchakul, S. Chaiyasith and W. Pecharapa, *Integr. Ferroelectr.*, 2016, **175**, 130.
- 51 A. S. Reynolds, T. H. Pierre, R. McCall, J. Wu and W. E. Gato, *J. Environ. Sci. Health, Part A: Toxic/Hazard. Subst. Environ. Eng.*, 2018, **53**, 986.
- 52 L. Lutterotti, D. Chateigner, S. Ferrari and J. Ricote, *Thin Solid Films*, 2004, **450**, 34.
- 53 E. Santala, R. Koivula, R. Harjula and M. Ritala, *Environ. Technol.*, 2019, **40**, 3561.
- 54 C. N. Hsiao and K. S. Huang, *J. Appl. Polym. Sci.*, 2005, **96**, 1936.
- 55 L. M. Al-Harbi, S. A. Kosa, M. K. Baloch, Q. A. Bhatti and E. E. H. El-Mossalamy, *Int. J. Polym. Sci.*, 2016, **2016**, 2417292.
- 56 X. Xia, X. J. Dong, Q. F. Wei, Y. B. Cai and K. Y. Lu, *EXPRESS Polym. Lett.*, 2012, **6**, 169.
- 57 N. Chiodini, F. Meinardi, F. Morazzoni, J. Padovani, A. Paleari, R. Scotti and G. Spinolo, *J. Mater. Chem.*, 2001, **11**, 926.
- 58 B. I. Denker, B. I. Galagan, L. D. Iskhakova, S. E. Sverchkov and E. M. Dianov, *Appl. Phys. B: Lasers Opt.*, 2015, **120**, 13.
- 59 C. Leonhardt, S. Brumm, A. Seifert, G. Cox, A. Lange, T. Ruffer, D. Schaarschmidt, H. Lang, N. Jöhrmann, M. Hietschold, F. Simon and M. Mehring, *ChemPlusChem*, 2013, **78**, 1400.
- 60 H. Ohgi, T. Maeda, E. Hosono, S. Fujihara and H. Imai, *Cryst. Growth Des.*, 2005, **5**, 1079.
- 61 I. Denry, J. A. Holloway and P. K. Gupta, *J. Biomed. Mater. Res., Part B*, 2012, **100**, 1198.
- 62 E. Mudra, I. Shepa, O. Milkovic, Z. Dankova, A. Kovalcikova, A. Annušová, E. Majkova and J. Dusza, *Appl. Surf. Sci.*, 2019, **480**, 876.
- 63 R. Trujillano, F. Villain, C. Louis and J.-F. Lambert, *J. Phys. Chem. C*, 2007, **111**, 7152.
- 64 R. N. Collins and A. S. Kinsela, *Chemosphere*, 2010, **79**, 763.
- 65 R. Koivula, R. Harjula and J. Lehto, *Microporous Mesoporous Mater.*, 2002, **55**, 231.
- 66 F. Granados, V. Bertin, S. Bulbulian and M. Solache-Ríos, *Appl. Radiat. Isot.*, 2006, **64**, 291.
- 67 J. Lehto, R. Koivula, H. Leinonen, E. Tusa and R. Harjula, *Sep. Purif. Rev.*, 2019, **48**, 122.

