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First-principles studies of substituent effects on squaraine dyes

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Dye molecules that absorb light in the visible region are key components in many applications, including organic photovoltaics, biological fluorescent labeling, super-resolution microscopy, and energy transport. One family of dyes, known as squaraines, has received considerable attention recently due to their favorable electronic and photophysical properties. In addition, these dyes have a strong propensity for aggregation, which results in emergent materials properties, such as exciton delocalization. This will be of benefit in charge separation and energy transport along with fundamental studies in quantum information. Given the high structural tunability of squaraine dyes, it is possible that exciton delocalization could be tailored by modifying the substituents attached to the π -conjugated network. To date, limited theoretical studies have explored the role of substituent effects on the electronic and photophysical properties of squaraines in the context of DNA-templated dye aggregates and resultant excitonic behavior. We used *ab initio* theoretical methods to determine the effects of substituents on the electronic and photophysical properties for a series of nine different squaraine dyes. Solvation free energy was also investigated as an insight into changes in hydrophobic behavior from substituents. The role of molecular symmetry on these properties was also explored *via* conformation and substitution. We found that substituent effects are correlated with the empirical Hammett constant, which demonstrates their electron donating or electron withdrawing strength. Electron withdrawing groups were found to impact solvation free energy, transition dipole moment, static dipole difference, and absorbance more than electron donating groups. All substituents showed a redshift in absorption for the squaraine dye. In addition, solvation free energy increases with Hammett constant. This work represents a first step toward establishing design rules for dyes with desired properties for excitonic applications.

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1. Introduction

The aggregation of dyes gives rise to Frenkel exciton delocalization in molecular composite systems.^{1,2} The study of dye aggregates can provide insight into the implementation of aggregate systems, such as organic photovoltaics,^{3,4} near infrared medical imaging⁵ and molecular photoswitch applications⁶ that draw from a well-established theoretical framework exploiting unique exciton transfer properties.⁷ Dye

aggregate behavior is well described by the molecular exciton model, formulated by Davydov and Kasha, in which intermolecular dipole–dipole interactions lead to the mixing of excited-state wavefunctions to access nondegenerate states.^{2,8,9} When dye monomers aggregate, the monomeric excited-state energy levels split into the excited states of the aggregate, in which excitons are distributed in a wave-like fashion, called exciton delocalization.^{10–12} The new excited-state behavior (*i.e.*, exciton delocalization) in the dye aggregate generally manifests as an energy shift.^{13,14} The inclusion of a double-body exciton interaction introduces a second order molecular excitation interaction, *i.e.*, the exciton–exciton interaction energy $K_{m,n}$, which is crucial for modeling exciton–exciton interaction behavior in the dye aggregate.^{14–16}

It is pertinent to also discuss the environment, in which dyes may aggregate. DNA is an attractive choice for dye templating due to its customization at the nanoscale and an ability to promote the exciton delocalization of dyes.^{17–34} DNA has been shown to negligibly change the electronic properties of visible-

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light absorbing dye monomers.³⁵ As such, the electronic properties of the dye monomers can be evaluated as free-dyes to screen their potential utility as DNA-templated dyes. The customizability of DNA templates and the options for binding dyes to different sites are further enhanced *via* DNA origami methods to construct multidimensional scaffolds.^{31,36}

Expanding the number of dyes that can potentially be incorporated will no doubt increase the functional capabilities of DNA templating.^{27,29,35,37–48} Squaraine (SQ) dyes, a family of dyes similar to the widely used cyanine dyes, but with a central squaraine ring, have advantageous properties, such as strong absorption in the visible spectrum⁴⁹ and resistance to photobleaching.⁵⁰ They can potentially be structurally tailored for a wide variety of applications.^{51–53} First synthesized by Treibs and Jacobs,⁵⁴ the central feature of SQ dyes includes an electron-deficient squaric moiety, combined with electron-rich groups in a symmetric manner by means of a methine bridge.⁵² The photophysical features and extensive structural tunability make squaraines well-suited candidates for the investigation of exciton delocalization when assembled.^{18,24,28,55} Although the dipole–dipole coulombic coupling between dyes must be considered to accurately predict aggregate absorption spectra, monomer transition dipole moments provide estimates of the strength of exciton delocalization for various dye configurations *via* the extended dipole approximation, which can be used as a guide for exciton applications.^{15,56}

There is also a robust body of work demonstrating the customizability of squaraine dyes, offering opportunities for the tunability of dye properties through the engineering of functional groups to yield desired properties.^{50,52,57–59} Previous work on photophysical property engineering of squaraine dyes *via* substitution is present in the literature with an emphasis on the changing of the donor groups that flank the central squaraine group or the finetuning of dyes largely for efficient light-to-electrical conversion and imaging.^{52,60,61} In the context of promoting stronger excitonic interaction, customizing squaraine dyes enables the introduction of substituents that may alter their electronic structures to make them more favorable for dipole interactions without detrimentally affecting photophysical properties.^{62–64} There has also been a research interest in manipulating two key excitonic coupling factors, including exciton hopping (or exchange) energy, $J_{m,n}$, and exciton–exciton interaction energy, $K_{m,n}$, by controlling the transition dipole μ and static dipole difference Δd of a dye monomer.^{56,65–67} The maximization of μ within a single absorption band is also benefitted by the minimization of vibronic coupling of dyes.⁶⁸ Maximizing exciton–exciton interaction energy depends upon the maximization of Δd . This should be concomitant with maintaining or increasing μ and is a primary target for the selection of dyes and their substituents in this study. The maximization of these quantities increases $J_{m,n}$ and $K_{m,n}$, leading to a larger excitation energy.⁵⁶

Substituent effects may also alter the propensity for dye aggregation by changing dye solubility.^{35,39} Local environment impacts the orientation of the dye by introducing steric effects when hydrophobic substituents are added. In the case of DNA-templated squaraines, the local environment can consider both

a solvent environment as well as DNA. The balance of substituents' ability to influence electronic, photophysical, and hydrophobic behaviors is key to the promotion of ideal dyes for excitonic device performance. To further investigate the effect of substituents, this study focuses on substituents that can increase the hydrophobicity of a squaraine dye, because this is expected to promote dye packing in order to influence dipole interactions.³⁹ The study of substituent effects on the electronic, photophysical, and solubility properties of monomers can therefore provide information on candidates for aggregation.

In this work, first-principles methods were used to address the potential for indolenine-based squaraine dyes to be tuned for excitonic applications. Specifically, we used density functional theory (DFT)-based methods to calculate the ground- and excited-state properties of nine squaraine dyes, *i.e.*, SQ-H₂, SQ-N(CH₃)₂, SQ-(N(CH₃)₂)₂, SQ-CH₃, SQ-(CH₃)₂, SQ-Cl, SQ-(Cl)₂, SQ-NO₂, and SQ-(NO₂)₂. A range of substituents on a free squaraine dye were studied to evaluate their influence on μ , Δd , absorbance, and hydrophobicity, which were subsequently validated against experimental data for DNA-templated squaraine monomers. By adding functional groups to an unsubstituted squaraine, *i.e.*, SQ-H₂, its μ , Δd , absorbance, and hydrophobicity could be altered. In addition, three different conformers were investigated for each dye to examine the impact of structural changes.

2. Methodology

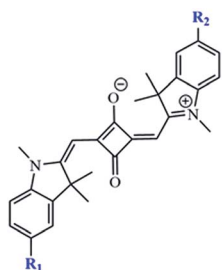
2.1 Computational methods

The Gaussian 16 software package⁶⁹ was used to perform density functional theory (DFT) and time-dependent density functional theory (TD-DFT) calculations. DFT has proven to provide insight in the investigation of dye properties,^{65–67,70} and there has been extensive work addressing best practices in employing this method, such as appropriate exchange–correlation functionals to represent electron–electron interactions in the fluorescing dyes similar to squaraine dyes.^{71–74} The dyes were built and initially relaxed with the molecular editing software Avogadro⁷⁵ using the UFF⁷⁶ method. All calculations were performed using the 6-31+G(d,p) basis set with the M06-2X⁷⁷ exchange–correlation functional, because this showed a good agreement with experimental results in comparison with sets of similar fluorescent dyes.^{72,73} M06-2X is a hybrid meta-generalized gradient approximation exchange–correlation functional. This nonlocality denotes an inclusion of Hartree–Fock (HF) exchange energy, which is advantageous for non-metal systems. The term “meta” indicates that the functional is constrained to be optimized using empirical data. Specifically, for dipole and absorption calculations, M06-2X was found to be the most reliable in comparison with other popular pure and hybrid exchange–correlation functionals.^{72,74,78} Jacquemin *et al.* also conducted an extensive survey of exchange–correlation functionals applicable to different dyes, including squaraine and so-called push–pull dyes.^{71,73,78–80} Molecules were built according to the structures in Fig. 1. Comparisons were made to monomers incorporated to DNA. DNA was found to negligibly affect the absorbance data of a monomer.^{18,35}



Atomic structures were optimized using a tight root mean square residual force of $1 \times (10)^{-5}$ Hartree/Bohr and an ultra-fine integration grid of 99 radial shells and 590 angular points per shell. The ground-state optimization of these molecules was verified by ground-state frequency calculations to ensure that no imaginary frequencies were present, because imaginary frequencies represent unstable geometry. Dyes exhibiting desirable Δd (*i.e.*, large Δd) were selected for further calculations to determine vibrationally-resolved absorption spectra. To do so, the optimized first excited-state geometry was used to calculate the excited-state frequency to ensure that an optimized structure was achieved. This procedure resulted in an adiabatic transition by including the zero-point vibrational energies, which accounted for vibrational energies at the states' respective minima.⁸⁰ The ground- and excited-state frequencies were then used to calculate an absorption spectrum for each molecule with the Franck–Condon (FC) approximation. The FC approximation assumes that nuclear motion is frozen on the timescale of the electronic transition. Our previous studies showed that the calculated absorption spectra of cyanine dyes with the FC approximation agreed well with experiments.³⁵ Here, the changes in the bond lengths of ground and excited states of squaraine dyes were analyzed. We found that the bond lengths among the squaric moiety carbons elongated. The methine chain lengths, connecting the trimethylindolenine groups, and the carbon–oxygen bonds of the squaric moiety shortened. However, all the changes between ground and excited states were on the order of 0.01 Å. This finding indicated that the FC approximation was adequate for squaraine dyes, and agreed well with Bassal *et al.*⁷⁴ The vertical excitation was expected to be indicative of the excitation behavior of the squaraine dyes. For dyes in solution, nonequilibrium solvent–solute conditions were considered. Permanent dipole information was generated by taking the vertical excitation from the optimized ground state. This vertical excitation could result in a difference between ground- and excited-state static dipoles, *i.e.* Δd , written as:⁷²

$$\Delta d = \sqrt{(d_x^{\text{ES}} - d_x^{\text{GS}})^2 + (d_y^{\text{ES}} - d_y^{\text{GS}})^2 + (d_z^{\text{ES}} - d_z^{\text{GS}})^2} \quad (1)$$



	R ₁	R ₂
SQ-H ₂	H	H
SQ-Cl	Cl	H
SQ-Cl ₂	Cl	Cl
SQ-NO ₂	NO ₂	H
SQ-(NO ₂) ₂	NO ₂	NO ₂
SQ-CH ₃	CH ₃	H
SQ-(CH ₃) ₂	CH ₃	CH ₃
SQ-N(CH ₃) ₂	N(CH ₃) ₂	H
SQ-(N(CH ₃) ₂) ₂	N(CH ₃) ₂	N(CH ₃) ₂

Fig. 1 Indolenine-based squaraines. H atoms are located at the R₁ and R₂ positions, forming SQ-H₂. If H is replaced with substituents, it forms either symmetric (R₁ = R₂ = substituent) or asymmetric (R₁ = H and R₂ = substituent). Substituents this study focuses on include N(CH₃)₂, CH₃, Cl, and NO₂.

where d_i^j is the Cartesian component of the permanent dipole moment, i , in either the excited ES or ground GS state, j .

Solvation energy, ΔG_{solv} , has been shown to correspond to hydrophobic behavior and qualitatively imply the stability of dye aggregate.³⁵ To investigate the likelihood of dye aggregation, the solvation free energy ΔG_{solv} was calculated by taking a difference in the ground state energies calculated using SMD (*i.e.*, Solvation Model based on Density) water and vacuum given as:^{35,81}

$$\Delta G_{\text{solv}} = E_{\text{solv}} - E_{\text{v}} \quad (2)$$

where E_{solv} is the ground-state energy calculated in solvent and E_{v} is the vacuum or gas-phase ground-state energy. Previous studies have introduced squaraine conformers present in solution. Their population percentages were calculated using the Boltzmann distribution.⁵⁸ The energies used to compare likely populations were the ground-state energies of each system in vacuum and in a water environment at 25 °C. In addition to water solvent, pyridine, quinoline, and isoquinoline were also investigated to approximate a DNA environment. In order to reduce computational time, nitrogen heterocycle solvents were used as suitable analogs to nitrogenous purine and pyrimidine nucleobases in a DNA scaffold.

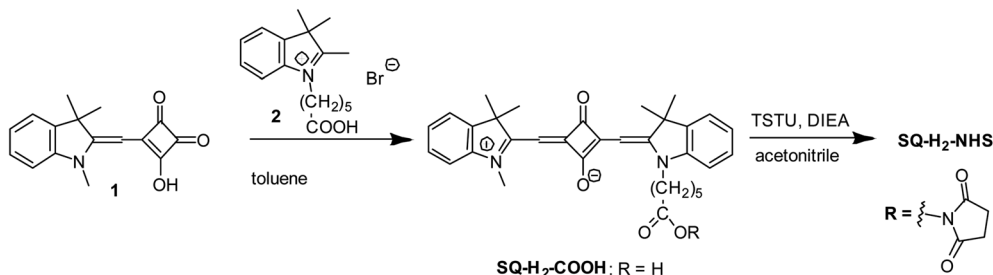
2.2 Experimental methods

N-Hydroxysuccinimide ester of 2-((1-(5-carboxypentyl)-3,3-dimethylindolin-2-ylidene)methyl)-3-oxo-4-((1,3,3-trimethyl-3*H*-indol-1-ium-2-yl)methylene)cyclobut-1-en-1-olate (SQ-H₂-NHS) was synthesised similar to the procedure described in Kolosova *et al.*⁵⁸ (Fig. 2).

3-Hydroxy-4-((1,3,3-trimethylindolin-2-ylidene)methyl)cyclobut-3-ene-1,2-dione (**1**) (150 mg, 0.56 mmol) and 1-(5-carboxypentyl)-2,3,3-trimethyl-3*H*-indolium bromide (**2**) (200 mg, 0.56 mmol) were heated under reflux in toluene (10 mL) for 10 h. The solvent was removed under reduced pressure by a rotary evaporator. The residue was purified by a column chromatography (Silica gel 60, 0–8% methanol–chloroform) to give SQ-H₂-COOH (190 mg, 65%) as a dark blue solid with a golden sheen. ¹H-NMR (200 MHz, DMSO-*d*₆), δ , ppm: 7.52 (2H, d, 7.3 Hz, arom.), 7.44–7.25 (4H, m, arom.), 7.25–7.04 (2H, m, arom.), 5.79 (1H, s, CH), 5.76 (1H, s, CH), 4.06 (2H, t, 7.4 Hz, NCH₂), 3.57 (3H, s, NCH₃), 2.21 (2H, t, 6.7 Hz, CH₂COOH), 1.68 (12H, s, (CH₃)₂), 1.80–1.29 (6H, m). MALDI-TOF MS, m/z calcd. for [C₃₃H₃₆N₂O₄] 524.27, found: 525.32 [M + H]⁺. Anal. calcd. (%) for C₃₃H₃₆N₂O₄: C, 75.55; H, 6.92; N, 5.34. Found C, 75.43; H, 6.89; N, 5.31. UV-Vis: λ_{max} (Abs) 630 nm (methanol); λ_{max} (Em) 639 nm (methanol); λ_{max} (Abs) 622 nm, ϵ 285 000 M⁻¹ cm⁻¹ (phosphate buffer); λ_{max} (Em) 632 nm (phosphate buffer).

SQ-H₂-COOH (30 mg, 57 μ mol), *N,N,N',N'*-tetramethyl-*O*-(*N*-succinimidyl)uronium tetrafluoroborate (TSTU) (26 mg, 86 μ mol), and *N,N*-diisopropylethylamine (DIEA) (16 mL, 92 mmol) were dissolved in acetonitrile (3 mL). The solution was stirred at room temperature for 20 min. The solvent was removed under reduced pressure by a rotary evaporator. The residue was purified by a column chromatography (Silica gel 60, 0–3% methanol–chloroform) to give SQ-H₂-NHS. Yield: 18 mg (51%). ¹H-



Fig. 2 Synthesis of SQ-H₂.

NMR (200 MHz, DMSO-*d*₆), δ , ppm: 7.53 (2H, d, 7.4 Hz, arom.), 7.46–7.24 (4H, m, arom.), 7.24–7.01 (2H, m, arom.), 5.79 (1H, s, CH), 5.76 (1H, s, CH), 4.06 (2H, t, 7.4 Hz, NCH₂), 3.58 (3H, s, NCH₃), 2.81 (4H, s, CH₂ (NHS)), 2.67 (2H, t, 6.0 Hz, CH₂-COONHS), 1.68 (12H, s, (CH₃)₂), 1.80–1.29 (6H, m).

For computation-experiment validation, we assembled a four-arm DNA Holliday junction with the unsubstituted squaraine SQ-H₂ covalently attached to one of the oligonucleotides. Three unlabeled oligonucleotides and one oligonucleotide labeled with SQ-H₂ (SETA BioMedicals, Urbana-Champaign, IL) *via* the nucleosidic sequence modifier C6 dT were obtained from Integrated DNA Technologies (Coralville, IA). Squaraine-labeled and unlabeled DNA oligonucleotides were rehydrated in ultrapure water (Barnstead Nanopure, Thermo Scientific) to prepare a 100 μ M stock solution. Concentrations of DNA samples were determined spectroscopically on NanoDrop One Microvolume UV-Vis (Thermo Scientific) using a calculated extinction coefficient. DNA Holliday junctions were prepared by combining equimolar amounts of

complementary functionalized and non-functionalized oligonucleotides in 1 \times TBE 15 mM MgCl₂ buffer solution, to a final DNA concentration 1.5 μ M. Samples were annealed in a Mastercycler Nexus PCR cyclor (Eppendorf) according to the following protocol: 4 min at 94 $^{\circ}$ C, followed by a cooling rate: 0.1 $^{\circ}$ C per 15 s from 94 $^{\circ}$ C to 64 $^{\circ}$ C, and 10 $^{\circ}$ C per minute from 64 $^{\circ}$ C to room temperature. UV-Vis spectra were recorded in duplicates at room temperature on a dual-beam Cary 5000 UV-Vis-NIR spectrophotometer (Agilent Technologies) in a cuvette with a 10 mm pathlength. Absorbance spectra were monitored over a wavelength range of 230–800 nm. Spectra were normalized at dye absorption maximum in UV-Vis range using OriginPro 2019.

3 Results

3.1 Boltzmann populations of conformers

To further complement changes associated with composition *via* substitution, this study considered three conformers of the

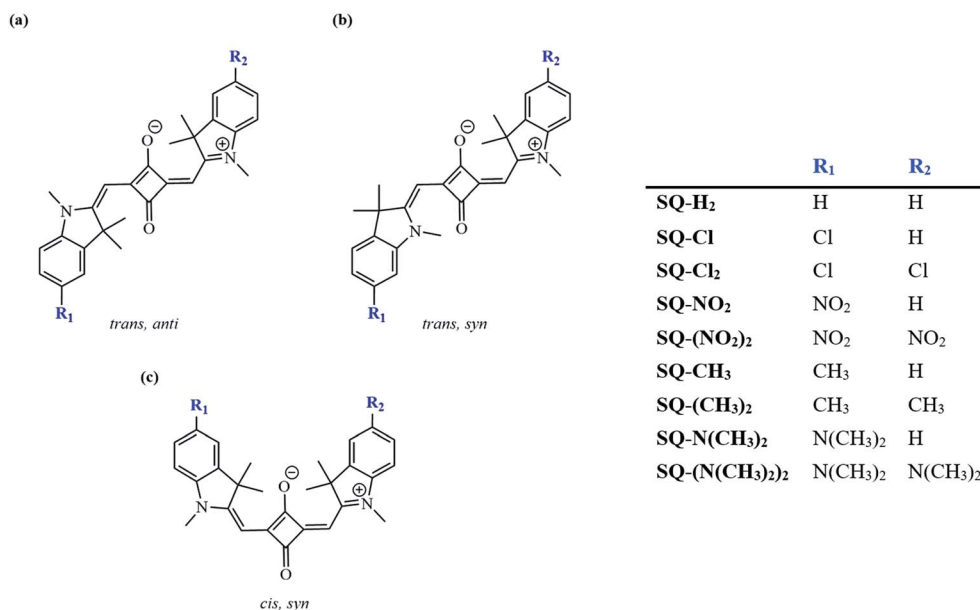


Fig. 3 Three possible indolenine squaraine conformers: (a) *trans,anti*, (b) *trans,syn*, and (c) *cis,syn*. Substituents this study focuses on include N(CH₃)₂, CH₃, Cl, and NO₂.



indolenine-based squaraine, including *trans,syn* with C_s symmetry; *cis,syn* with C_{2v} symmetry; and *trans,anti* with C_{2h} symmetry using a Boltzmann distribution calculation at room temperature.⁵⁸ The conformations of dyes were considered based on previously reported indolenine-based squaraines *via* photoisomerization.^{58,82} The energies of the optimized ground-state squaraine dyes were used to calculate the energy differences and corresponding Boltzmann populations of three conformers of the indolenine-based squaraine dyes, as shown in Fig. 3. The energy differences between the optimized ground-state conformers are shown in Table 1. For all dyes, a more stable conformer is represented by a more negative total energy. Comparing the differences of total energies of different conformers shows that the *trans,anti* conformer is about 5 kJ mol⁻¹ more stable than the *cis,syn* conformer (*trans,anti-cis,syn*) and about 8 kJ mol⁻¹ more stable than the *trans,syn* conformer (*trans,anti-trans,syn*). The energy difference results, when used as states for a Boltzmann distribution, demonstrate that the two major conformers of indolenine-squaraines are *trans,anti* and *cis,syn*. The *trans,anti* conformer is the most energetically favorable for all the indolenine-based squaraine dyes with the most negative total energy. These energy differences have been described as a result of steric strain from the dimethyl group of the indolenine rings in agreement with previous computational and experimental studies.^{58,83,84} The *trans,anti* conformer further benefits from a staggered conjugation, favoring minimized steric effects.^{85,86} However, we found that substituents have a minor effect on the Boltzmann populations with respect to the unsubstituted *trans,anti* squaraine (SQ-H₂) at most by only 3%. The greatest variation of conformer population from the unsubstituted *cis,syn* SQ-H₂ conformer occurs from SQ-Cl by 3%. The Boltzmann populations of different *trans,syn* conformers are lower than 4% and have the greatest variation with respect to the unsubstituted *trans,syn* SQ-H₂ is also from SQ-Cl by 1%. Computational results show that the *cis,syn* and *trans,anti* conformers are popular, so our further studies focus on these conformers.

3.2 Comparison of SQ-H₂ with experiment

To validate our approach for the calculations of ground- and excited-state properties of substituted squaraines, the

vibrationally-resolved absorption spectra of SQ-H₂ were generated using the Franck-Condon (FC) method in the TD-DFT framework for the *trans,anti* and *cis,syn* conformers. The calculated absorption spectra are shown in Fig. 4 along with the experimental absorption profile.

TD-DFT satisfactorily reproduces the lineshape of the experimental absorption spectrum, which exhibits a strong absorption peak at 638 nm and a smaller vibronic shoulder around 590 nm. The TD-DFT calculated peak absorption λ_{\max} is found to be 647 nm for the *trans,anti* conformer and 651 nm for the *cis,syn* conformer. Compared to the experimental data, the absorbance data calculated by TD-DFT exhibits a peak absorption λ_{\max} within 0.027 eV of experiment (as calculated for the *trans,anti* conformer). Furthermore, the calculated *trans,anti* peak absorption λ is closest to experiment, which suggests that the *trans,anti* conformer dominates the dye populations, agreeing well with the Boltzmann population results (Table 1). The deviation from the theoretical calculation is a known artifact in TD-DFT calculations when computing absorption spectra for dyes. This is due to small perturbations in the excited electronic density that the hybrid exchange-correlation functional cannot adequately model in the TD-DFT scheme.⁸⁷

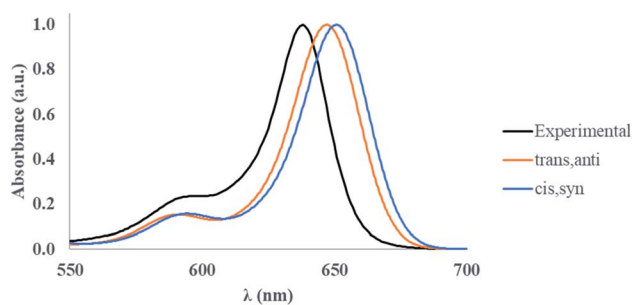


Fig. 4 Experimental and calculated vibrational-resolved absorption spectra for SQ-H₂. The calculated absorption spectra were obtained using the FC approach with the optimized ground- and excited-state structures of a free dye in a water solvent. The experimental spectrum was obtained for SQ-H₂ covalently attached to a DNA HJ; the concentration of SQ-H₂ – DNA HJ construct was 1.5 μM in 1× TBE 15 mM MgCl₂ aqueous buffer solution.

Table 1 Ground-state total energy differences between *trans,anti* and *cis,syn* conformers and between *trans,anti* and *trans,syn* conformers, as well as associated Boltzmann populations calculated from the energy differences at 25 °C in vacuum

Dye	Energy difference (kJ mol ⁻¹)		Boltzmann populations (%)		
	<i>trans,anti-cis,syn</i>	<i>trans,anti-trans,syn</i>	<i>trans,anti</i>	<i>trans,syn</i>	<i>cis,syn</i>
SQ-(N(CH ₃) ₂) ₂	-4.81	-8.39	84.93	2.88	12.2
SQ-N(CH ₃) ₂	-4.78	-8.18	84.58	3.12	12.3
SQ-(CH ₃) ₂	-4.77	-8.33	82.14	1.31	16.54
SQ-CH ₃	-4.76	-8.32	84.65	2.95	12.39
SQ-H ₂	-4.81	-8.31	82.69	2.57	14.74
SQ-Cl	-4.70	-8.18	80.49	1.49	18.02
SQ-(Cl) ₂	-4.72	-8.46	83.02	2.54	14.44
SQ-NO ₂	-4.69	-8.85	84.84	2.38	12.78
SQ-(NO ₂) ₂	-4.95	-8.73	85.83	2.53	11.63



3.3 Solvation free energy calculations

To determine the effects of electron donating and withdrawing substituents on the solvation energies of squaraine dyes, DFT ground-state optimization calculations were performed in vacuum and solvent. From the solvated and vacuum energies, the solvation free energy, ΔG_{solv} , was determined using eqn (2) for each dye, as shown in Fig. 5. As with other studies,^{88,89} the values of ΔG_{solv} were calculated to estimate the solubility of the dyes in the given solvent. As shown in Fig. 1, the substituted dye structures consist of symmetrically substituted dyes ($R_1 = R_2$) and asymmetrically substituted dyes ($R_1 = \text{H} \neq R_2$).

Unsubstituted squaraine has a ΔG_{solv} of -0.81 eV in water for both the *trans,anti* and *cis,syn* conformers. For all dyes, the conformer does not affect ΔG_{solv} , and, overall, the values of ΔG_{solv} are only slightly affected by substitution. Upon substitution, most dyes exhibit more negative ΔG_{solv} values in water, except for **SQ-Cl**, which has the least negative ΔG_{solv} value, indicating being the most hydrophobic. In contrast, **SQ-N(CH₃)₂** has the most negative ΔG_{solv} for the water solvated dyes and so is taken to be the most hydrophilic. The ΔG_{solv} for the dyes in pyridine, quinoline, and isoquinoline follow similar trends as those in water. The values of ΔG_{solv} in pyridine, quinoline, and isoquinoline are also unaffected by conformers. ΔG_{solv} for the dyes in water are less negative, indicating that it is more energetically favorable for the dyes to form solutes in pyridine, quinoline, and isoquinoline. These three solvents are taken to mimic the molecular structure of DNA bases.

3.4 Excited-state calculations

The optimized water-solvated ground-state structures of unsubstituted and substituted squaraines were used for single-step TD-DFT calculations in the first excited singlet state to determine the effects that electron donating and electron withdrawing substituents have on electronic excited-state properties (*i.e.*, μ and Δd). We introduced the empirically derived Hammett constant (σ_p), which could quantify the strength of a substituent as electron withdrawing (positive) or electron donating (negative).^{90,91} A list of these constants is

provided in Table 2. **N(CH₃)₂** and **NO₂** are strong donating and withdrawing substituents, respectively. **CH₃** and **Cl** are weak donating and withdrawing substituents, respectively. The Hammett constant has been shown to relate to the characteristics of an electronic structure.^{67,92} To establish a relationship between the strengths of the electron donating and electron withdrawing substituents (as quantified with σ_p) and calculated properties, μ , Δd , and $\Delta\lambda_{\text{max}}$ are plotted against the σ_p of attached substituents, as discussed in the following sections.

3.4.1 Transition dipole moments. Compared to unsubstituted squaraine, Fig. 6 shows that substituents with the larger magnitudes of σ_p yield dyes with larger μ . The calculated values of μ for unsubstituted squaraine (**SQ-H₂**) are 14.7 D for the *trans,anti* conformer and 14.4 D for the *cis,syn* conformer and were determined to be along the long axis of the dye. The dyes exhibiting the largest values of μ are *trans,anti* **SQ-N(CH₃)₂** (15.9 D) and **SQ-(NO₂)₂** (16.3 D), which are the dyes with the strongest electron donating and electron withdrawing substituents, respectively. In general, symmetrically substituted dyes have the larger values of μ than asymmetrically substituted ones. Furthermore, the *trans,anti* conformers exhibit the larger values of μ than *cis,syn* conformers. In comparison with the unsubstituted squaraine, the largest change is **SQ-(NO₂)₂** with 1.6 D, and all substituents contribute an increase in μ .

3.4.2 Static dipole differences. Similar to μ , the Δd of squaraine increases with the magnitude of σ_p , as shown in Fig. 7. Substitution type and conformation also influence Δd .

Table 2 Substituents used in this study and their corresponding Hammett constants,^{90,91} σ_p . The empirically derived Hammett constant (σ_p) is used to quantify the strength in which a substituent is electron withdrawing (positive) or electron donating (negative)

Substituent	Hammett constant (σ_p)	Classification
N(CH₃)₂	-0.83	Donating
CH₃	-0.17	
H	0	
Cl	0.23	Withdrawing
NO₂	0.78	

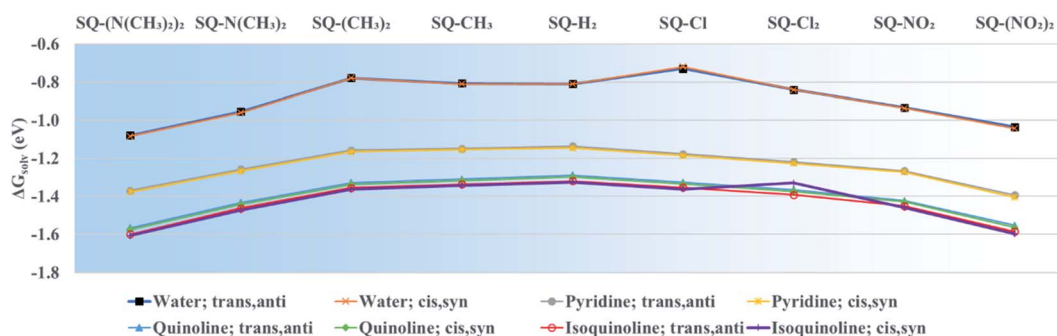


Fig. 5 Solvation free energy (ΔG_{solv}) of unsubstituted and symmetrically substituted squaraine dyes in water, pyridine, quinoline, and isoquinoline calculated using eqn (2). Geometry optimizations were done using the M06-2X functional. The lines added to the data are to highlight trends of the data and are not meant to infer quantitative behavior. The x-axis is in order of increasing donating and withdrawing strength as its position away from **SQ-H₂**.



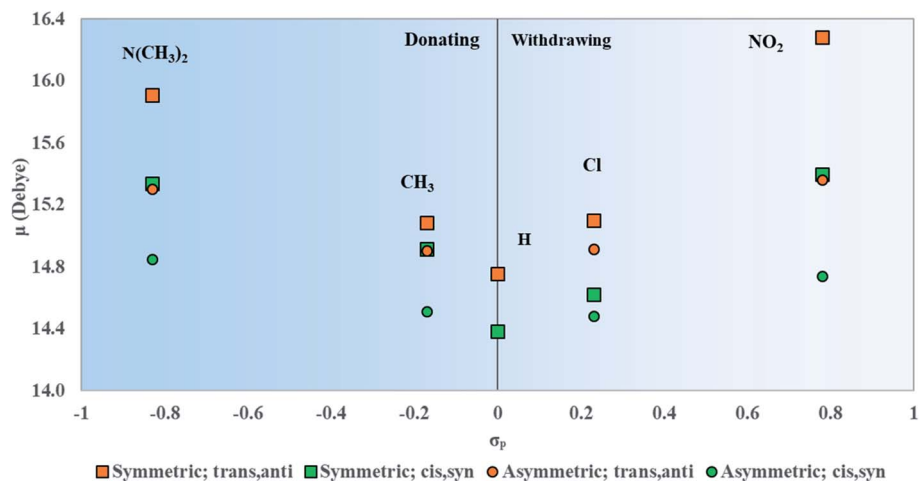


Fig. 6 Transition dipole moment magnitudes (μ) for symmetrically and asymmetrically substituted squaraines plotted against the Hammett constant of the substituent(s) attached to the dye.

Symmetrically substituted **SQ-N(CH₃)₂** and **SQ-NO₂** dyes have the higher Δd than their asymmetrical ones. Symmetrically substituted **SQ-(CH₃)** and **SQ-Cl** dyes have the lower Δd than their asymmetrical ones. Furthermore, symmetrically substituted dyes in the *trans,anti* conformations all have Δd values of nearly 0 D while others have non-zero Δd . In general, *trans,anti* conformers, characterized by C_{2h} type symmetry, have the lower values of Δd compared to *cis,syn* conformers, characterized by C_{2v} symmetry. The largest values of Δd belong to *cis,syn* **SQ-Cl** (0.8 D) in the asymmetrically substituted dyes and *cis,syn* **SQ-NO₂** (3.0 D) in the symmetrically substituted dyes. This indicates that electron withdrawing substituents have the larger impact on Δd than electron donating substituents.

3.4.3 Absorption spectra of asymmetric squaraines. To further test the effects of substituents on the excited-state properties of dyes, vibrationally-resolved absorption spectra were calculated. The asymmetrically substituted dyes in the *trans,anti* conformations with the larger Δd values were

chosen for further studies with the FC method. The normalized results are shown in Fig. 8.

Upon asymmetric substitution, the lineshapes of the spectra remain relatively unaffected, with a main absorption peak between 1.97 and 1.84 eV (630 and 675 nm) and a smaller vibronic shoulder around 2.1 eV (600 nm). However, the λ_{\max} values for asymmetrically substituted dyes are redshifted compared to the unsubstituted squaraine dye. **SQ-Cl** has the smallest redshift of 0.04 eV. The largest redshifts of 0.08 eV and 0.09 eV belong to **SQ-N(CH₃)₂** and **SQ-NO₂**, indicating that the stronger electron donating or electron withdrawing substituents have a larger effect on λ_{\max} .

Like μ and Δd , the values for $\Delta\lambda_{\max}$ were plotted against the values of σ_p for the substituents attached to the dyes. $\Delta\lambda_{\max}$ is defined as the redshift of the λ_{\max} of the dyes in Fig. 8 from unsubstituted squaraine (**SQ-H₂**). As shown in Fig. 9, the larger σ_p values of the substituents promote the larger $\Delta\lambda_{\max}$.

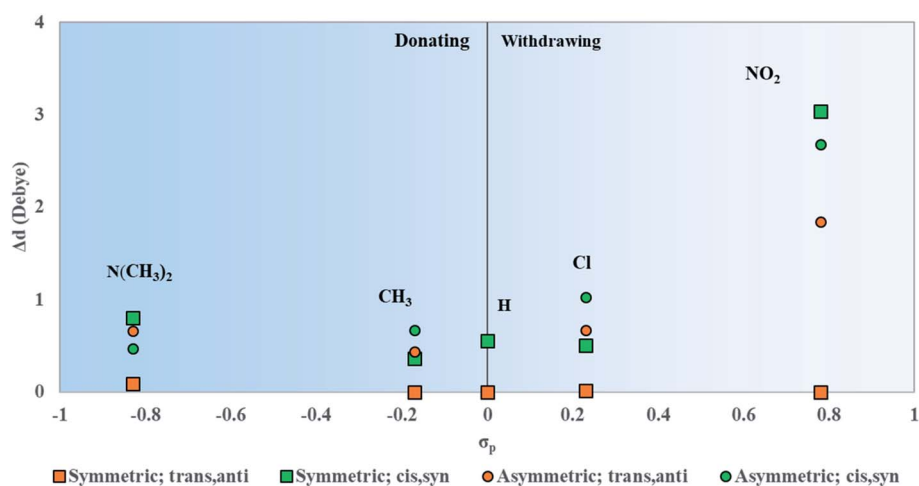


Fig. 7 Static dipole difference magnitudes (Δd) for symmetrically and asymmetrically substituted squaraines plotted against the Hammett constant of the substituent(s) attached to the dye.



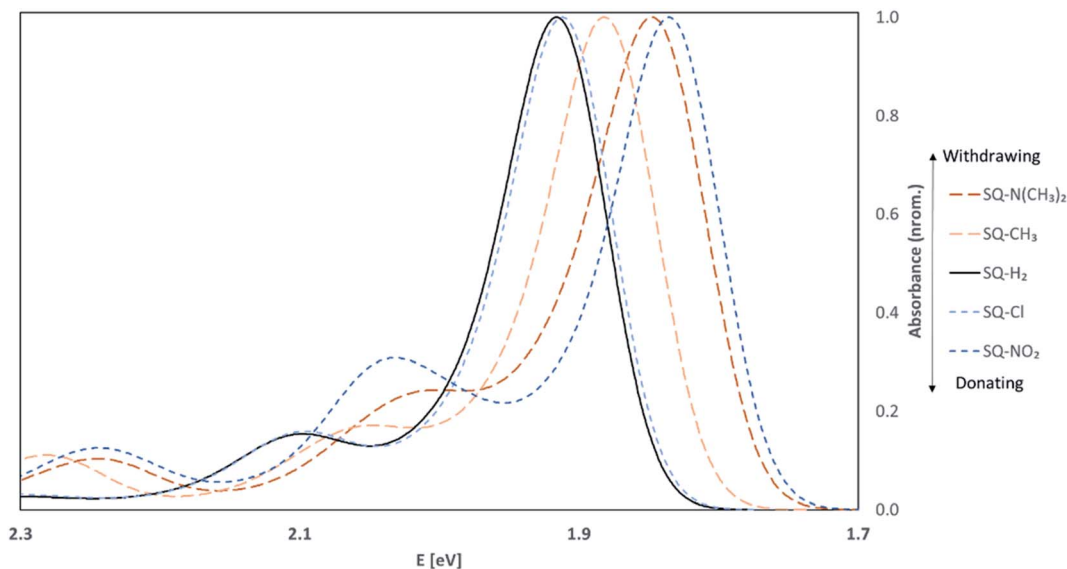


Fig. 8 Vibrationally-resolved absorption spectra for *trans,anti* asymmetrically substituted squaraine dyes calculated using the FC method in implicit water. The solid black line is the unsubstituted squaraine (SQ-H₂) absorbance spectrum.

4. Discussion

In regard to the changes to Δd and μ upon substitution, all substituents lead to changes in electronic structures as compared with the reference squaraine SQ-H₂. In the frame of treating the squaraine as a donor-acceptor-donor dye, the substituents modify the donating behavior of the trimethylindolenine groups attached to the accepting squaric moiety. The electron withdrawing substituents in this study appear to have a stronger effect on the change of Δd and μ . Electronically, this may be the result of decreasing the donating ability of the indolenine groups towards the squaraine center. In contrast to the withdrawing substituents, the donating substituents are

expected to increase the donating strength of the indolenine groups. The contribution of the electron donating substituents to an increase in Δd and μ is relatively smaller than that of the electron withdrawing substituents (Fig. 6 and 7).

ΔG_{solv} is sensitive to solvent and becomes more negative with the nitrogen heterocyclic solvents pyridine, quinoline, and isoquinoline, as shown in Fig. 5. This trend increases from pyridine to quinoline and isoquinoline and can be attributed to π - π interactions between the solvent and the solute.⁹³ The magnitude of ΔG_{solv} also increases with σ_p . An increase in the magnitude of ΔG_{solv} indicates a greater stability in a solvent environment. This would suggest that the dyes have

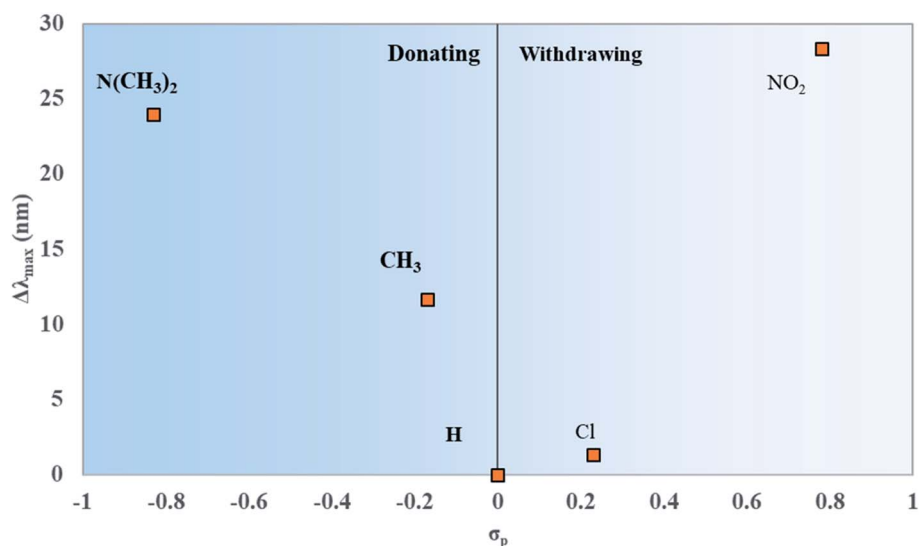


Fig. 9 Maximum absorption wavelength differences from unsubstituted squaraine ($\Delta\lambda_{\text{max}}$) for asymmetrically substituted squaraines plotted against the Hammett constant of the substituent(s) attached to the dye.



a propensity for a DNA environment and are more likely to aggregate in a DNA template rather than dissolve in aqueous solution.

Recent squaraine dye applications have revealed the effect of conformers on device performance. These effects are due to the alteration of energy transfer pathways.^{53,82,94–97} Conformers have also been shown to change transition dipole orientation, further potentially affecting exciton delocalization.⁹⁸ Certain applications may benefit from different conformers, such as dye-sensitized solar cells in part because of accommodation of different anchoring options on substrates.⁵³ Boltzmann distribution calculations have been used to predict the population of conformers existing in solution.⁵⁸ Squaraine dyes with unaltered central squaric moieties have been found most likely to exist in a *trans,anti* (C_{2h}) symmetry state.

Upon both symmetric and asymmetric substitution, μ is relatively unaffected by the type of substitution. Because of this, the exciton hopping interaction occurring between dyes in an aggregate should remain unchanged and may be slightly augmented with substituents. For symmetric substitution, most dyes in the *trans,anti* conformations exhibit a Δd of ~ 0 D. As shown in Fig. 3, dyes in the *trans,anti* conformations have C_{2h} symmetry, which, along with the donor–acceptor–donor electronic pattern of squaraine, results in small changes in Δd . In contrast, dyes in the *cis,syn* conformations are characterized by C_{2v} symmetry and exhibit a non-zero Δd . By only including a single substituent on the dye (*i.e.* asymmetric substitution), the structural symmetry is distorted. Furthermore, substituents with non-zero σ_p would increase Δd . Specifically for the substituents studied in this work, the electron withdrawing substituent NO_2 increases Δd the most. Our computational results suggest that σ_p can guide the selection of dye candidates with desired electronic and photophysical properties.

Based on the results of this study, the substitution of the squaraine indolenine rings can enhance the dye's excitonic properties. Upon substitution, μ is slightly increased for most dyes, indicating that the excitonic hopping interaction between dyes in an aggregate should be enhanced rather than diminished. Similarly, substituents can promote an increase in Δd , which could improve the exciton–exciton interaction energy of the squaraine aggregate.

5. Conclusion

Squaraine dyes with varied substituents were investigated to compare their solvation free energy ΔG_{solv} , static dipole difference Δd , transition dipole moment μ , and absorption wavelength λ_{max} using DFT and TD-DFT. Changes in these values upon substitution were compared to the empirically derived Hammett constant σ_p and experimental absorption profiles for the unsubstituted squaraine dye. It was found that the magnitude of σ_p correlated with ΔG_{solv} , μ , Δd , and λ_{max} . ΔG_{solv} becomes more negative with a larger σ_p value in water and in solvents similar to a DNA environment. μ increases with σ_p for symmetric substitution patterns. Δd increases with asymmetric substitution and σ_p . λ_{max} also increases with σ_p . These findings on the electronic, photophysical, and hydrophobic properties of

squaraine dyes can guide the selection of substituted dyes. The ability to control dye properties, when coupled with DNA scaffolding, may make it possible to tailor the performance of dye aggregate materials for excitonic systems and applications.

Author contributions

German Barcenas: investigation (lead), writing – original draft (lead), methodology (equal), formal analysis (lead), data curation (lead). Austin Biaggne: investigation (supporting), writing – original draft (supporting), methodology (equal), formal analysis (supporting), data curation (supporting). Olga A. Mass: conceptualization (equal), validation (lead), writing – review and editing (equal), supervision (supporting). Christopher K. Wilson: validation (supporting). Olena M. Obukhova: resources (supporting). Olga S. Kolosova: resources (supporting). Anatoliy L. Tatarets: resources (supporting). Ewald Terpetschnig: resources (supporting), writing – review and editing (equal). Ryan D. Pensack: conceptualization (equal), writing – review and editing (equal). Jeunghoon Lee: conceptualization (equal), writing – review and editing (equal). William B. Knowlton: conceptualization (equal), project administration (supporting), writing – review and editing (equal). Bernard Yurke: conceptualization (equal), writing – review and editing (equal), supervision (supporting). Lan Li: conceptualization (equal), project administration (lead), writing – review and editing (equal), supervision (lead).

Conflicts of interest

There are no conflicts to declare.

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