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A depth-suitable and water-stable trap for CO₂ capture†

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In terms of CO₂ capture and storage (CCS), it is highly desired to substitute of high efficiency process for the applied one which is far from the ideal one. Physical processes cannot capture CO₂ effectively, meanwhile CO₂ desorption is energy-intensive in chemical processes. Herein, a depth-suitable and water-stable trap for CO₂ capture was discovered. Carboxylates can react with polybasic acid roots by forming united hydrogen bonds. Carboxylate ionic liquid (IL) aqueous solutions can absorb one equimolar CO₂ chemically under ambient pressure, and its CO₂ desorption is easy, similar to that in physical absorption/desorption processes. When used as aqueous solutions, carboxylate ILs can replace alkanolamines directly in the applied CCS process, and the efficiency is enhanced significantly due to the low regenerating temperature. CO₂ (or polybasic acids) can be used as a polarity switch for ILs and surfactants. A new method for producing carboxylate ILs is also proposed.

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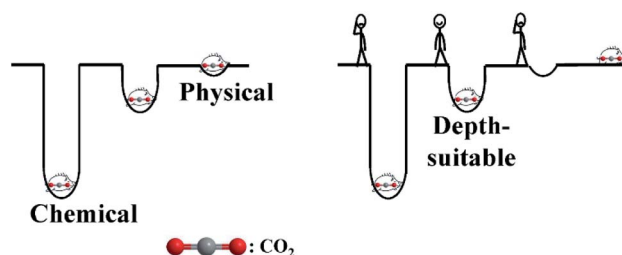
Being the main greenhouse gas, CO₂ capture and storage (CCS) is indispensable for achieving carbon neutrality, and it has been a hot topic for many years.¹ Besides CO₂ generated from the energy demand of human being's daily activities, energy requirement in CCS also produces more CO₂, which comes from fuel consumption, and this will lead to an increase in electricity prices.^{2,3} So, the enhancement of energy efficiency in CCS is highly desired. Alkanolamines are the main absorbents applied for CCS in industry, and their regeneration processes are energy intensive, as shown in Scheme 1 (chemical trap). There are more inherent shortages, such as amine degradation and volatilize, and hence, the CCS process applied in this study is far from the ideal one.

Having many unique features, such as extremely low vapor pressure, ionic liquids (ILs) have aroused the interest of many scientists,⁴⁻⁶ including their ability to absorb CO₂ physically.⁷ Physical absorption is conducted under high pressure and anhydrous conditions. There are several shortcomings as follows: the sorption capacity is limited at low pressure, water is invariably present in almost all ILs,⁸ CO₂ and water are the two main exhaust products in fossil fuel burning. It is difficult to get satisfactory results in physical absorption, as shown in Scheme 1. Functional ILs have also been tested for absorbing CO₂ (ref. 9 and ¹⁰) and SO₂ (ref. 11 and ¹²) chemically, and some authors claimed that the products obtained are carbamic acids or amidates while absorbing CO₂. Recently, Dupont *et al.* considered that carbonates/

bicarbonates are generated because water is inevitable in most cases,⁸ and this is similar to the alkanolamine aqueous solution in the applied CCS process. The regeneration of those ILs is energy intensive, and the stronger the functional ILs used for absorbing CO₂, the more energy is needed in the regeneration process. The high viscosity of ILs is also unfavourable for capturing CO₂ effectively. Finding a depth-suitable trap (as shown in Scheme 1) with low viscosity and water-stability for CO₂ capture is highly desirable.

Various chemical reactions are going on at this moment in nature and our bodies, and many of them have been discovered and explained clearly.¹³ Most of them form new chemical bonds, including covalent bonds, ionic bonds, and metal bonds.¹⁴ In comparison, hydrogen bonding is a weak interaction that has been known for a century¹⁵ and has been redefined recently.¹⁶ Although weak, hydrogen bonds are vital for water keeping its state as we know generally¹⁷ and for life passing on its genetic code.^{18,19}

Herein, a depth-suitable and water-stable trap for CO₂ capture was discovered. Carboxylate roots can react with polybasic acid roots by forming united hydrogen bonds, and this



Scheme 1 Schematic diagram of a depth-suitable trap for CO₂ capture.

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lower energy of the products make them more stable even in water. Carboxylate IL aqueous solutions can replace alkanolamine aqueous solutions directly in the applied CCS process, and the efficiency is enhanced significantly, which comes from the low regenerating temperature. CO₂ (or polybasic acids) can be used as a polarity switch for ILs and surfactants. A new way for producing carboxylate ILs is also proposed.

Finding the trap

Strong acids reacting with weak acid salts to produce weak acids are classic reactions that are widely applied in industry. Most carboxylic acids, such as formic acid and acetic acid, are stronger than carbonic acid, and can react with carbonates to form carboxylates, CO₂, and water. On the contrary, some carboxylates can absorb CO₂ under certain conditions.^{20–26} So, there must be an intermediate in this reaction, which is relatively stable under certain conditions. Bicarbonate is a reasonable explanation, and Yasaka *et al.* considered that bicarbonate is stabilized by a hydrogen bond with a carboxylic acid.²⁵ However, those systems are viscous and not stable in water, and the CO₂ sorption shrinks down sharply with an increase in water concentration.

In this study, 1 mole of sodium 2,2-dimethylbutyrate with 0.5 mol water could absorb 0.5 mol CO₂ under ambient conditions in one bottle. In another bottle, 2,2-dimethylbutyric acid and sodium carbonate (molar ratio of 2 : 1) were added and stirred, and there was no CO₂ bubbling out as the bottle was covered with a cap. Proportions of the 2,2-dimethylbutyrate root, sodium ion, and carbonic acid (CO₂ plus H₂O) in the two bottles are the same, so it is reasonable to consider that the structure of the products from the left and right sides of eqn (1) (in Scheme 2) are the same, and the product has a lower energy level than the reactants of two sides of the reaction under certain conditions. It is an equilibrium reaction, and the equilibrium will shift to the right side by moving CO₂ out or adding water. The stability of the intermediate varies with carboxylate anions. As the caps of the bottles are opened, the CO₂ release

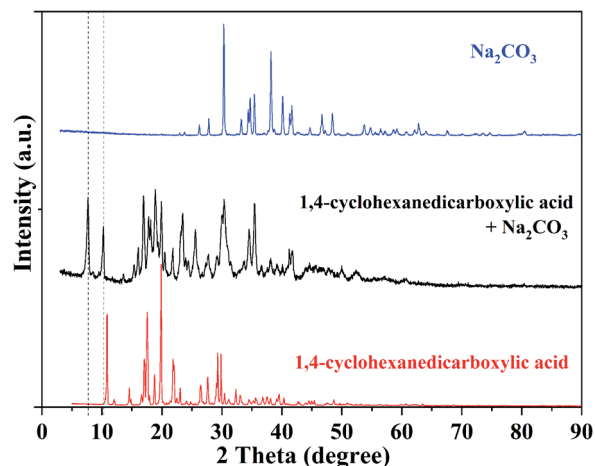
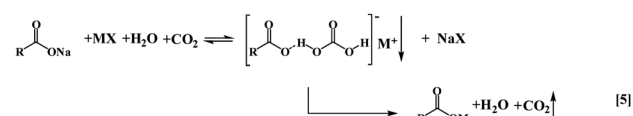
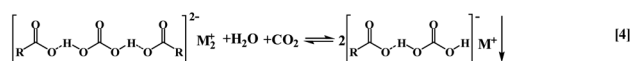
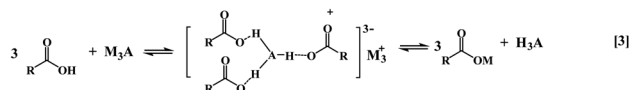
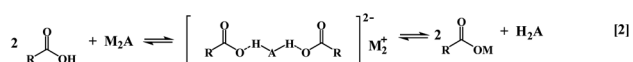
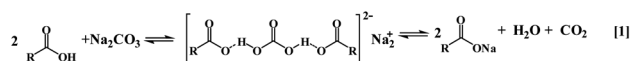


Fig. 1 Powder XRD patterns of 1,4-cyclohexanedicarboxylic acid + Na₂CO₃.

rate from formic acid and sodium carbonate is faster than that from acetic acid and sodium carbonate. X-ray diffraction (XRD) and infrared radiation (IR) analysis results of the product of sodium carbonate reacting with 1,4-cyclohexanedicarboxylic acid are shown in Fig. 1 and S1a,[†] respectively. The results show that the product is different from both the reactants. The new peaks at 7.7° and 10.3° in XRD indicate that an ordered structure is formed between the two reactants.

The stability of the intermediates also varies with different cations, including inorganic and organic cations. Tributylhexylphosphonium 2-ethylhexanoate (1) is a hydrophilic IL and miscible with water. When 1 is selected as the carboxylate in eqn (1), the product after the absorption of CO₂ is stable in water, and the sorption remains constant with water being added. When the aqueous solution (50 wt%) is exposed to the CO₂ atmosphere, a new liquid phase appears during the absorption of CO₂, as shown in Fig. 2, and the sorption is 0.49 mol CO₂ per mol IL under ambient temperature. This suggests that the sorption of CO₂ changes the polarity of the reaction system and induces product separation from water.

With the above understanding, we proposed that there are two hydrogen bonds in the intermediates of eqn (1), and the two united hydrogen bonds tie two carboxylate roots and one carbonate root together, as shown in the spatial conformation



Scheme 2 Reactions between carboxylates and polybasic acid roots. M represents inorganic or organic cation.

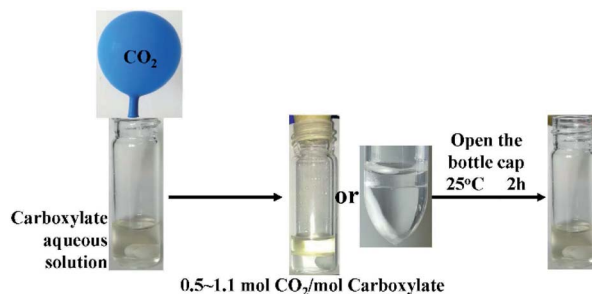


Fig. 2 Phase behaviour of some carboxylate IL aqueous solutions with or without CO₂.



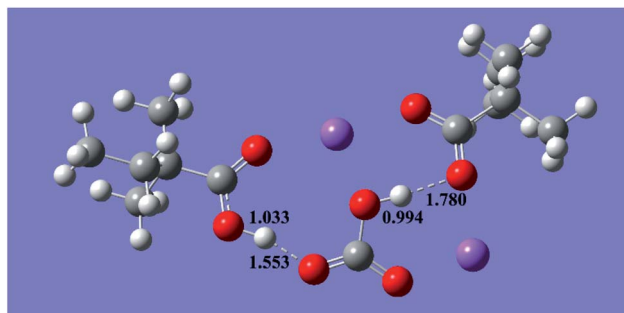


Fig. 3 Spatial conformation of the product from 2,2-dimethylbutyric acid + Na_2CO_3 obtained from DFT calculation.

obtained from DFT calculation (see Fig. 3 and S2a, b†). The united hydrogen bonds lower the energy of the products and enhance their stabilities even in an aqueous solution. The united hydrogen bonds also change the polarity of the product and cause its separation from water.

Reaction of carboxylates with other polybasic acid roots

This reaction is also tested with sulfur dioxide (SO_2), hydrogen sulfide (H_2S), and phosphoric acid, and they can all react with an aqueous solution of **1** and form two liquid phases. The top phase of aqueous solution **1** absorbing SO_2 was tested with ^1H NMR spectroscopy (Fig. S3†), and the results show that tributylhexylphosphonium ions and 2-ethylhexanoate ions in the top phase are equimolar. This implies that the top organic phase separating from the aqueous solution is not 2-ethylhexanoic acid, which may come from 2-ethylhexanoate ions reacting with sulfite acid. Some products of 1,4-cyclohexanedicarboxylic acid and 1,3,5-benzenetricarboxylic acid reacting with sodium sulfite and dipotassium hydrogen phosphate were tested with XRD (Fig. S4a–c†), IR spectroscopy (Fig. S1b and c†), and scanning electron microscopy (SEM, Fig. S5a–d†). XRD and IR analysis of the products indicate that the products are different from the raw reactants, and new ordered structures are generated. The SEM images show the shapes of those products, and there is one kind of substance in each photo, while not mixture of two reactants. This represents the reactants reacting with each other and forming new products. The products from 1,3,5-benzenetricarboxylic acid + K_2HPO_4 (Fig. S5a†) and 1,3,5-benzenetricarboxylic acid + Na_2SO_3 (Fig. S5d†) are sheets, and this implies that the united hydrogen bond products from 1,3,5-benzenetricarboxylic acid are ready to form a sheet. The shapes of products from 1,4-cyclohexanedicarboxylic acid + K_2HPO_4 (Fig. S5c†) and 1,4-cyclohexanedicarboxylic acid + Na_2SO_3 (Fig. S5b†) are different. It is seen from the four images that the shape varies with the type of carboxylate acids and salts. The spatial conformations of sulfite acid and H_2S reacting with **1** from DFT calculation (Fig. S2c and d†) show that there are two hydrogen bonds in both of their products.

The product of tripotassium phosphate reacting with 1,4-cyclohexanedicarboxylic acid was tested with XRD (Fig. S4d†)

and SEM (Fig. S5e†). The product of tripotassium phosphate reacting with 1,3,5-benzenetricarboxylic acid was tested using IR spectroscopy (Fig. S1d†). The results demonstrate that the product is different from that of dipotassium hydrogen phosphate plus carboxylate acid. This means that carboxylate roots can react with phosphate root in two ways. One is two carboxylate ions combining with one hydrogen phosphate ion, and forming two hydrogen bonds (Fig. S2e†). The other is three carboxylate ions combining with one phosphate ion, and forming three hydrogen bonds (Fig. S2f†).

In order to show the reaction clearly, the general equations are given as eqn (2) and (3). The united hydrogen bonds enhance the stability of the products as in DNA,¹⁹ and some of them are stable in water. The united hydrogen bonds tie two or three carboxylate ions together, and this induces different polarities of the products with reactants.

Influence of different anions and cations

The influence of different anions and cations on the system of carboxylates absorbing CO_2 in water was studied and listed in Table 1. When the salts are strongly hydrophilic, such as sodium carboxylates, they are soluble in water, and there is no change in the phase behaviour of their aqueous solutions with the addition of CO_2 . When the carboxylate ILs have suitable hydrophilicity, they are miscible with water, and new phases appear in their aqueous solutions with the addition of CO_2 . Some of them form new liquid phases, and some others form solid phases, as shown in the table. CO_2 sorption in the systems forming two phases was measured and marked in the table. For tributylhexylphosphonium aqueous solutions, there are four anions that can form a new liquid phase with the addition of CO_2 . For **1** and tributylhexylphosphonium 2-propylpentanoate, CO_2 sorption is about 0.5 mole per mol IL, and this implies that these two ILs can form stable, united hydrogen bonds in water. For tributylhexylphosphonium 2,2-dimethylbutyrate and tributylhexylphosphonium *n*-octanoate, the sorptions are 0.11 and 0.18 mol CO_2 per mol IL, respectively. This indicates that the united hydrogen bonds in their products are not stable, resulting from the strong hydrophilicity of anions.

When the anion of ILs are selected as 2-ethylhexylate, several IL aqueous solutions form solids with the addition of CO_2 , and there are great differences in CO_2 sorption. For tributyl-*n*-octylphosphonium-2-ethylhexylate (**2**), the sorption of CO_2 can get to 1.1 mole per mol **2**. During CO_2 absorption by the aqueous solution of **2**, a state of liquid–liquid–solid was found, as shown in Fig. 2. The three-phase state was also found in the CO_2 release of **2** + H_2O + CO_2 system. A liquid–liquid state was found in this system, absorbing CO_2 at 80 °C. All these findings suggest that the united hydrogen bond product shown in eqn (1) is an intermediate, and it reacts with CO_2 and water further and generates the solid as shown in eqn (4). The solid is the crystallization of the product with some more carbonic acid. The solid has lower energy than the liquid united hydrogen bond product shown in eqn (1) and can tie more CO_2 . The



generating carbonates and dicarbonates, and the organic solvents transform from nonpolar to polar. In this work, the united hydrogen bond reaction could change the polarity of carboxylate systems, and they can be used as switchable solvents and surfactants. A series of carboxylate ILs were found that can be separated from water and form a new liquid phase with the addition of CO₂ or other polybasic acids (SO₂ + H₂O, H₃PO₄, etc.). CO₂ can be released at room temperature in a short time, and it will be miscible with water again. Selecting long-chain tetraalkylammonium carboxylates and tetraalkylphosphonium carboxylates as surfactants and switching their polarities by CO₂ may realize the separation of surfactants from water or other solvents. For tributylhexadecylphosphonium ion, it can be separated out with 2-ethylhexanoate root from aqueous solution by dissolving tributylhexadecylphosphonium bromide and sodium 2-ethylhexanoate, even at a concentration of 1 wt%, in several minutes with CO₂ bubbling (Fig. S6[†]). The separation process is revisable. The mixtures are miscible again and become clear in three minutes as they are exposed to air. This process can separate tetraalkylphosphonium (tetraalkylammonium or other ions) carboxylates from water (solid or a new liquid phase) and leave the corresponding ions in water as shown in eqn (5), and so it provides a new way for producing carboxylate ILs.

Conclusions

In summary, we discovered a depth-suitable and water-stable trap for CO₂ capture that can replace the one currently in use, and the efficiency is significantly enhanced. Carboxylates can react with polybasic acid roots by forming united hydrogen bonds, and the united hydrogen bonds enhance the stability of the products even in aqueous solution. Some united hydrogen bond carbonates can interact with CO₂ and water further and form water-stable solids. It can be used to capture acidic gases (CO₂, SO₂, and H₂S) chemically. The CO₂ sorption can exceed one equimolar for carboxylates under ambient pressure. Having suitable heats of absorption, carboxylates can absorb CO₂ chemically, and its CO₂-desorption process is easy, almost like that in physical absorption/desorption processes. CO₂ (or other polybasic acids) can be used as a polarity switch for carboxylate ionic liquids and surfactants. This reaction also provides a new method for producing carboxylate ILs.

Author contributions

Z. Z. proposed the idea, performed most experiments and wrote the manuscript. S. L. and T. W. performed the characterization. J. M. conducted DFT calculations.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 N. M. Dowell, P. S. Fennell, N. Shah and G. C. Maitland, *Nat. Clim. Change*, 2017, **7**, 243–249.
- 2 R. S. Haszeldine, *Science*, 2009, **325**, 1647–1652.
- 3 M. E. Boot-Handford, J. C. Abanades, E. J. Anthony, M. J. Blunt, S. Brandani, N. Mac Dowell, J. R. Fernández, *et al.*, *Energy Environ. Sci.*, 2014, **7**, 130–189.
- 4 J. S. Wilkes and M. J. Zaworotko, *J. Chem. Soc., Chem. Commun.*, 1992, 965–967.
- 5 Y. L. Wang, *Chem. Rev.*, 2020, **120**, 5798–5877.
- 6 Z. F. Zhang, Y. Xie, W. J. Li, S. Q. Hu, J. L. Song, T. Jiang and B. X. Han, *Angew. Chem., Int. Ed.*, 2008, **47**, 1127–1129.
- 7 L. A. Blanchard, D. Hancu, E. J. Beckman and J. F. Brennecke, *Nature*, 1999, **399**, 28–29.
- 8 M. Zanatta, N. M. Simon and J. Dupont, *ChemSusChem*, 2020, **13**, 3101–3109.
- 9 E. D. Bates, R. D. Mayton, I. Ntai and J. H. Davis Jr., *J. Am. Chem. Soc.*, 2002, **124**, 926–927.
- 10 D. Hospital-Benito, J. Lemus, C. Moya, R. Santiago and J. Palomar, *Chem. Eng. J.*, 2020, **390**, 124509.
- 11 W. Z. Wu, B. X. Han, H. X. Gao, Z. M. Liu, T. Jiang and J. Huang, *Angew. Chem., Int. Ed.*, 2004, **43**, 2415–2417.
- 12 C. M. Wang, G. Cui, X. Luo, Y. Xu, H. Li and S. Dai, *J. Am. Chem. Soc.*, 2017, **133**, 11916–11919.
- 13 J. J. Li, *Name Reactions*, Springer-Verlag Berlin Heidelberg, 4th edn, 2009.
- 14 L. Pauling, *The nature of the chemical bond*, Cornell University Press, Ithaca, 3rd edn, 1960.
- 15 W. M. Latimer and W. H. Rodebush, *J. Am. Chem. Soc.*, 1920, **42**, 1419–1433.
- 16 E. Arunan, G. R. Desiraju, R. A. Klein, J. Sadlej, S. Scheiner, *et al.*, *Pure Appl. Chem.*, 2011, **83**, 1619–1636.
- 17 F. N. Keutsch, J. D. Cruzan and R. J. Saykally, *Chem. Rev.*, 2003, **103**, 2533–2577.
- 18 J. D. Watson and F. H. C. Crick, *Nature*, 1953, **171**, 737–738.
- 19 S. E. Harding, G. Channell and M. K. Phillip-Jones, *Biochem. Soc. Trans.*, 2018, **46**, 1171–1182.
- 20 R. Quinn, *Synth. React. Inorg. Met.-Org. Chem.*, 2001, **31**, 359–369.
- 21 R. Quinn, J. B. Appleby and G. P. Pez, *J. Am. Chem. Soc.*, 1995, **117**, 329–335.
- 22 G. Wang, W. Hou, F. Xiao, J. Geng, Y. Wu and Z. Zhang, *J. Chem. Eng. Data*, 2011, **56**, 1125–1133.
- 23 K. Anderson, M. P. Atkins, J. Estager, Y. Kuah, S. Ng, A. A. Oliferenko, N. V. Plechkova, A. V. Puga, K. R. Seddon and D. F. Wassell, *Green Chem.*, 2015, **17**, 4340–4354.
- 24 Y. Yasaka and Y. Kimura, *J. Chem. Eng. Data*, 2016, **61**, 837–845.
- 25 Y. Yasaka, Y. Saito and Y. Kimura, *ChemPhysChem*, 2018, **19**, 1674–1682.
- 26 D. J. Yeadon, J. Jacquemin, N. V. Plechkova, M. Maréchal and K. R. Seddon, *ChemPhysChem*, 2020, **21**, 1369–1374.
- 27 P. G. Jessop, D. J. Heldebrant, L. Xiaowang, C. A. Eckert and C. L. Liotta, *Nature*, 2005, **436**, 1102.
- 28 Y. X. Liu, P. G. Jessop, M. Cunningham, C. A. Eckert and C. L. Liotta, *Science*, 2006, **313**, 958–960.

