


Cite this: *RSC Adv.*, 2021, 11, 19836

Thermodynamically stable structure of hydroxy alkane sulfonate surfactant monomers in water achieving high water solubility and a low CMC†

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Although sodium internal olefin sulfonates (IOSs) derived from petrochemicals are known to act as anionic surfactants, these compounds have few practical applications and their interfacial properties have not been examined in detail because they have complex compositions and are challenging to prepare. However, IOSs obtained from vegetable oils have recently been applied to detergents. The present work represents the first study of the IOS sodium 5-hydroxyhexadecane-4-sulfonate (C16-4S-5OH) as a pure substance. The properties of C16-4S-5OH in aqueous solution were assessed by differential scanning calorimetry, equilibrium surface tension measurements, Fourier transform infrared spectroscopy and proton nuclear magnetic resonance spectroscopy. This compound was found to have a low Krafft point and a low CMC, both of which are necessary for a modern sustainable surfactant. Although these two characteristics are normally difficult to achieve simultaneously, this work demonstrated that C16-4S-5OH in aqueous solution has a highly fixed stereostructure that enables both properties to be achieved. The C16 main carbon chain of this compound acts as a C12 and a C4 chain both oriented in the same direction, while the hydrophilic sulfo and hydroxy head groups form a rigid cyclic moiety including two carbons of the C16 alkyl chain based on hydrogen bonding. This highly fixed molecular structure of the C16-4S-5OH in water is maintained in the monomer and micellar states below and above the CMC, respectively.

Received 8th February 2021
Accepted 26th May 2021

DOI: 10.1039/d1ra01046h

rsc.li/rsc-advances

Introduction

Surfactants have numerous uses in industry. In particular, anionic surfactants are able to act as inexpensive cleaning agents. Although the raw materials of most general anionic surfactants are mainly biobased, synthetic ones are also used. Therefore, the development of completely biobased anionic surfactants is urgently required to satisfy global sustainability demands.¹ Sodium internal olefin sulfonates (IOSs), which can be synthesized by the sulfonation of alkenes with internal double bonds produced by the petrochemical industry, has been known as an anionic surfactant since the 1970s (Fig. 1).² Those analogues in which the sulfo group is instead located at a terminal carbon of the alkyl chain are termed sodium α -olefin sulfonates, and there have been many reports concerning the properties of these compounds in aqueous solution.^{3,4} Conversely, there have been only a few studies on the aqueous solution properties of IOS^{3–5} presumably because these are typically complex mixtures (Fig. 1). Specifically, as-synthesized IOSs generally exhibit a wide distribution of sulfo group

positions and also typically comprise a combination of hydroxy alkane sulfonates (**1a**) and alkene sulfonates (**1b**). IOSs have not found applications in industry, likely because it is challenging to prepare colorless IOSs,⁵ although a high quality IOS product was recently obtained from vegetable oils.^{6,7} Because this bio-based IOS comprised linear alkenes derived from C16 and C18 saturated fatty acids (which have not been extensively employed industrially) rather than from C12 and C14 fatty acids that are currently widely used as raw materials for surfactants, it may

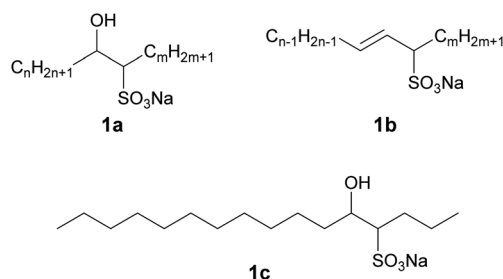


Fig. 1 The molecular structures (**1a** and **1b**) of typical IOSs, where $n + m + 2$ is the total carbon number and a hydrophilic head group is located at some point in the carbon chain. In this study, **1c** (**1a** with $n = 11$, $m = 3$; C16-4S-5OH; sodium 5-hydroxyhexadecane-4-sulfonate) was utilized.

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ra01046h



represent a sustainable anionic surfactant.⁸ This IOS may also allow the effective utilization of surplus saturated long-chain fats.

The practical application of a surfactant requires high solubility in water and a low critical micelle concentration (CMC). The former is necessary to dissolve in water while the latter reduces the amount of surfactant required. Even so, it is very difficult to simultaneously obtain both properties, because the former requires high hydrophilicity while the latter necessitates significant hydrophobicity. In addition, an ideal surfactant will function under various conditions (such as at low temperatures or high water hardness values) and has to work with less amount from the viewpoint of sustainability. Therefore, the development of a design strategy for surfactant molecules that can simultaneously satisfy the aforementioned apparently incompatible requirements is desirable.

In the present study, the aqueous solution properties of IOS were subsequently assessed using the pure sodium 5-hydroxyhexadecane-4-sulfonate (C16-4S-5OH) (**1c**), which is one of the most representative components included in the bio-based IOS showing a low Krafft point (Kp) and low CMC. It is the first-ever investigation of a completely pure single-component IOS. The molecular structure of this IOS is characterized by two hydrophilic groups (that is, sulfo and hydroxy groups) located at the 4- and 5-carbons of the alkyl chain. The relationship between molecular structure and interfacial properties is one of the most interesting and widely researched aspects of such compounds. However, the stereostructures of amphiphiles in solution must be examined not only based on their theoretical molecular structures but also by considering their actual forms induced by hydrophobic effects or hydration of the hydrophilic groups.^{9–12} Therefore, the actual state and stereostructure of C16-4S-5OH in water were studied in detail. Moreover, the relationship between the stereostructure of the surfactant in water and the aqueous solution properties was examined.

Experimental

Materials

In this study, the raw material for the synthesis of the C16-4S-5OH (4-hexadecene) was first made. During this process, 1432.6 g butyltriphenylphosphonium bromide (3.59 mol; Tokyo Chemical Industry Co., Ltd., Tokyo, Japan) was dissolved in 4.50 L tetrahydrofuran (FUJIFILM Wako Pure Chemical Corp., Osaka, Japan) by stirring at 0 °C. Following this, 382.6 g potassium *tert*-butoxide (3.41 mol; FUJIFILM Wako) in 1.50 L tetrahydrofuran was added and the mixture was cooled to –10 °C and stirred for 1.50 h. Subsequently, the reaction mixture was further cooled to –60.0 °C and 507.6 g dodecanal (2.75 mol; Tokyo Chemical Industry) in 0.50 L tetrahydrofuran was added dropwise and the mixture was stirred for 3.00 h at room temperature. Following this, the reaction was quenched by adding a saturated aqueous solution of ammonium chloride (FUJIFILM Wako). The organic phase was washed with brine and dried with magnesium sulfate (FUJIFILM Wako), after which the solvent was removed by evaporation. The resulting

residue was purified by column chromatography utilizing silica gel (60 N; 63–200 µm; Kanto Chemical Co., Inc., Tokyo, Japan) and hexane (FUJIFILM Wako). As in prior studies, gas chromatography (GC) was used to confirm that the 4-hexadecene was a mixture of stereoisomers (*cis/trans* = 97.1/2.9).^{13,14}

C16-4S-5OH was obtained by the sulfonation of 4-hexadecene followed by neutralization with sodium hydroxide.⁷ The reaction product was recrystallized from ethanol (>99.5%, FUJIFILM Wako) three times to give the high-purity C16-4S-5OH monohydrate (purity, >99.9%). The concentrations of impurities such as inorganic salts and residual 4-hexadecene were confirmed to be below the detection limits of the respective analytical techniques during analyses by ion chromatography and gas chromatography. The molecular structure of the product was determined using proton nuclear magnetic resonance (¹H NMR) (Fig. S1, ESI†) and liquid chromatography with tandem mass spectrometry (Fig. S2, ESI†). Additionally, the purity of the material was confirmed by elemental analysis (Table S1, ESI†).

All the aqueous solution samples utilized for the subsequent experiments were prepared with ultrapure water obtained from an RFD250RB system (Advantec Toyo Kaisha, Ltd., Tokyo, Japan). The NMR and Fourier transform infrared (FT-IR) analyses utilized deuterium oxide (FUJIFILM Wako) and sodium isethionate (Combi-Blocks, Inc., San Diego, USA) as received without further purification.

Methods

Determination of Kp using differential scanning calorimetry. The Kp value of the C16-4S-5OH was determined by differential scanning calorimetry (DSC), employing a Discovery DSC (TA Instruments, New Castle, PA, USA). An approximately 70.0 mg sample of a 1.00 wt% aqueous solution of the compound was transferred into a stainless-steel pan (TA Instruments), which was then sealed and heated at a rate of 0.30 °C min^{–1} from –40.0 to 60.0 °C after an initial annealing below –20 °C.

Equilibrium surface tension measurements. The equilibrium surface tension (EST) of the various concentrations of aqueous C16-4S-5OH solution was assessed at 26.0 °C using the Wilhelmy plate method, utilizing a platinum plate with a K-100MK2 tensiometer and a temperature-controlled vessel (Krüss GmbH, Hamburg, Germany). These measurements were performed in the equilibrium state, defined as that in which the fluctuation in the values was less than 0.10 mN m^{–1} for 60 min. All data were acquired after the samples had been allowed to stand for longer than 12 h in a lidded Teflon Petri dish.

¹H NMR spectroscopy. The ¹H NMR spectra were acquired with an ECA-600 instrument (600 MHz; JEOL Ltd., Tokyo, Japan) at 26.0 °C. The chemical shifts were determined relative to a deuterium peak centered at 4.65 ppm, based on previous reports.^{12,15}

Hydrogen/deuterium (H/D) exchange analyses. C16-4S-5OH was dissolved in deuterium oxide at concentrations that were either above or below the CMC and left to stand at 60.0 °C for 3.00 h. Thereafter, the sample was lyophilized. In trials with



sodium isethionate, a 1 wt% deuterium oxide solution was utilized, following the same procedure as was applied with the C16-4S-5OH.

Attenuated total reflectance FT-IR spectroscopy. Attenuated total reflectance (ATR) FT-IR spectra were acquired using a Nicolet 6700 spectrometer (Thermo Fisher Scientific, Inc., Massachusetts, USA) with 2.00 cm^{-1} resolution.

Results and discussion

Aqueous solution properties of C16-4S-5OH

Kp. The Kp of C16-4S-5OH as a 1.00 wt% (approximately $2.9 \times 10^{-2}\text{ mol dm}^{-3}$) aqueous solution was determined to be $25.6\text{ }^{\circ}\text{C}$ by DSC. This Kp is compared to those for similar surfactants in Table 1.^{16,17} Sodium 2-hydroxyalkane-1-sulfonate (C16-1S-2OH), in which the sulfo group is on a terminal carbon, is an analogue of C16-4S-5OH with the two functional groups in different positions. Interestingly, both sodium hexadecyl sulfonate (C16-1S) and C16-1S-2OH, which are also C16 compounds, have much higher Kp values than C16-4S-5OH. These data do not suggest that these surfactants become more water soluble if they possess more head groups. Moreover, even sodium 2-hydroxydodecane-1-sulfonate (C12-1S-2OH), which has a shorter C12 alkyl chain, has a Kp that is $41.9\text{ }^{\circ}\text{C}$ higher than that of C16-4S-5OH. It is well known that longer alkyl chains tend to increase Kp and, comparing C16-1S-2OH with C12-1S-2OH, the linear alkyl chain is shortened by four carbons such that Kp was reduced by $20.5\text{ }^{\circ}\text{C}$. However, comparing C16-1S-2OH with C16-4S-5OH, the sulfo group is shifted from the 1-carbon to the 4-carbon of the alkyl chain and the Kp is decreased by $62.4\text{ }^{\circ}\text{C}$. These results indicate that an internal relocation of the hydrophilic groups is more effective at lowering Kp than reducing the alkyl chain length. The remarkably low Kp of C16-4S-5OH resulting from its internal hydrophilic groups would be expected to allow its use around room temperature, even though most conventional surfactants possessing a C16 alkyl chain precipitate at this temperature. Generally, the Kp values of surfactants with long linear alkyl chains are higher because of the strong hydrophobic effect between the alkyl chains and the aqueous environment. In the case of C16-4S-5OH, the hydrophilic groups divide the linear C16 chain asymmetrically into C12 and C4 chains. This is

thought to affect the alkyl chain conformation so as to reduce interactions in water.

EST of a C16-4S-5OH aqueous solution. Fig. 2 plots EST (γ) against $\log C$ (where C is the total concentration) for a series of aqueous C16-4S-5OH solutions at $26.0\text{ }^{\circ}\text{C}$ (just above the Kp value). The γ values are seen to have decreased linearly as C increased up to the point where they eventually plateaued, corresponding to the CMC. Interestingly, the plot displays a small minimum point close to the CMC. This phenomenon is typically attributed to the presence of a low concentration of hydrophobic impurities, which in the present case might be a byproduct (**1b** in Fig. 1) or positional isomers in which the hydroxy and/or sulfo groups are at different carbons. However, these impurities were not detected by the analyses described above and the elemental analysis data were in good agreement with the theoretical values. Thus, any such impurities are thought to have been present at negligible levels. The linear decrease in γ with increasing concentration in the more dilute regime corresponds to the Gibbs adsorption equation:

$$\Gamma = -\frac{1}{4.606RT} \left(\frac{\partial \gamma}{\partial \log C} \right); \quad (1)$$

$$A_{\min} = \frac{1}{N_A \Gamma}, \quad (2)$$

Here, Γ is the surface excess concentration of the surfactant, R is the gas constant, T is the absolute temperature, A_{\min} is the area per molecule at the air–water interface, and N_A is Avogadro's number. The surface tension of C16-4S-5OH in the vicinity of its CMC (γ_{CMC}) and its A_{\min} values (calculated using eqn (2)) are provided in Table 2, where they are compared with values for conventional anionic surfactants (sodium C12–16 alkyl sulfates). These surfactants have very similar molecular structures to those of sodium alkyl sulfonates but are more water-soluble, and so their reported data are more reliable than those for alkyl sulfonates.^{18–21}

Table 1 Krafft points (Kp) and molecular structures of C16-4S-5OH and other anionic surfactants

	Molecular structure	Kp/ $^{\circ}\text{C}$
C16-4S-5OH		25.6
C16-1S		59.0 (ref. 16)
C16-1S-2OH		88.0 (ref. 17)
C12-1S-2OH		67.5 (ref. 17)

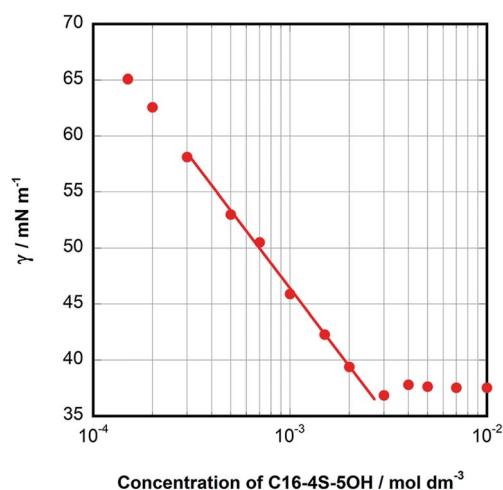


Fig. 2 EST versus the concentration of aqueous C16-4S-5OH solutions at $26.0\text{ }^{\circ}\text{C}$. The straight line represents the Gibbs adsorption isotherm.



Table 2 Interfacial properties of aqueous solutions of C6-4S-5OH and conventional anionic surfactants (sodium alkyl sulfates (AS))^{18–21}

	CMC/mmol dm ⁻³	$\gamma_{\text{CMC}}/\text{mN m}^{-1}$	$\Gamma_{\text{max}}/10^{-6} \text{ mol m}^{-2}$	$A_{\text{min}}/\text{nm}^2$
C16-4S-5OH ^a	3.00	37.8	1.89	0.88
C12 AS ^b	7.94 (ref. 18)	39.9 (ref. 19)	3.16 (ref. 19)	0.53 (ref. 19)
C14 AS ^b	2.05 (ref. 18)	35.2 (ref. 20)	3.70 (ref. 20)	0.45 (ref. 20)
C16 AS ^c	0.50 (ref. 21)			

^a At 26 °C. ^b At 25 °C. ^c At 45 °C. The values for C16-4S-5OH were calculated using eqn (1) and (2) employing the experimental data in Fig. 2. C_n AS: C_nH_{2n+1}-OSO₃Na.

The CMC for C16-4S-5OH was determined to be $3.00 \times 10^{-3} \text{ mol dm}^{-3}$, which was lower than that of C12 AS, almost the same as that of C14 AS and much higher than that of C16 AS. These results clearly suggest that C16-4S-5OH can be considered as comprising asymmetric C12 and C4 chains rather than a single C16 alkyl chain. For double chain surfactants, empirical data show that the subchains are expected to exert almost half the effect of the main chain on the CMC values proportional to the carbon number.²² On this basis, the CMC of C16-4S-5OH should be almost equal to that of C14 AS. Although a C4 alkyl chain is too short to have a hydrophobic effect,^{10,23} the observed decrease in the CMC demonstrates that the C4 chain in this compound was able to function as a hydrophobic chain together with the C12 chain. The γ_{CMC} of C16-4S-5OH was intermediate between those of C12 AS and C14 AS, as was also the case for the CMC values. In contrast, the A_{min} value for C16-4S-5OH was 0.88 nm^2 , which was much larger than those for C12 AS and C14 AS (Table 2). This large A_{min} is not attributed to the presence of the C12 and C4 chains because the A_{min} value of the representative double-tailed anionic surfactant sodium 4-(dodecan-4-yl)benzenesulfonate is smaller (0.55 nm^2).²⁴ It is well known that surfactants reduce surface tension as a result of surface adsorption and coverage.²² Interestingly, this comparison of the A_{min} values suggests the opposite of the anticipated result, since C16-4S-5OH would be expected to form a more densely packed monolayer than C12 AS at the air–water interface.

As noted in the Introduction, a high degree of solubility in water and a low CMC are essential for modern surfactants. Although these two properties are normally not achieved simultaneously, the C16-4S-5OH exhibits both a low K_p and CMC. Common qualitative correlations, including that of A_{min} , between theoretical molecular structure and interfacial properties cannot adequately explain this result. Thus, the stereostructure of this compound in aqueous media was studied, as discussed in the next sub-section.

Analysis of the stereostructure of C16-4S-5OH in water

The actual states of the hydrophobic and hydrophilic portions of the C16-4S-5OH in water were studied to clarify the true stereostructure, which might be responsible for its various characteristics of interest.

¹H NMR. The conformations of the alkyl chains of C16-4S-5OH in water was investigated using ¹H NMR spectroscopy (Fig. S1, ESI†). Fig. 3 presents the ¹H NMR peaks generated by

the 4- and 5-methine protons based on spectra obtained from a $1.00 \times 10^{-3} \text{ mol dm}^{-3}$ C16-4S-5OH solution in D₂O (that is, a concentration lower than the CMC value), together with the corresponding coupling constants. The coupling constants of the 3-methylene and 4-methine protons were found to be significantly nonequivalent (they had different coupling constants; $J_{\text{HH}} = 5.20$ and 6.00 Hz , respectively). Furthermore, the 5-proton and the two 6-protons in the C12 chain were also nonequivalent. These results indicated that the rotations of both the C4 and C12 chains were highly restricted. The coupling constant between the two methine protons (the 4-H and 5-H) was 2.40 Hz , which also suggests that the C4 and C12 chains were tightly fixed and oriented in the same direction. Additionally, if the C–C bond between the 4-C and 5-C was able to rotate freely, the C4 and C12 chains would be expected to orient themselves in opposite directions to reduce steric repulsion, which in turn would increase the coupling constant. It should be noted that the coupling constant associated with the 5-proton and hydroxy proton was difficult to determine because the 5-hydroxy proton peak at approximately 4.6 ppm (Fig. S1, ESI†) completely overlapped the solvent peak. These ¹H NMR results suggest that the configurations of the two asymmetric alkyl chains that were directly attached at the 4- and 5-C positions of the C16-4S-5OH monomer were strongly fixed in this medium.

The effects of the solution concentration on this characteristic stereostructure of the monomer were also studied. In particular, the changes in the ¹H NMR spectra associated with micelle formation were examined by comparing spectra acquired below and above the CMC. In Fig. 4, the difference in the δ values of the peaks for the methyl protons on the terminal carbons (that is, 1- and 16-H) are plotted against the inverse

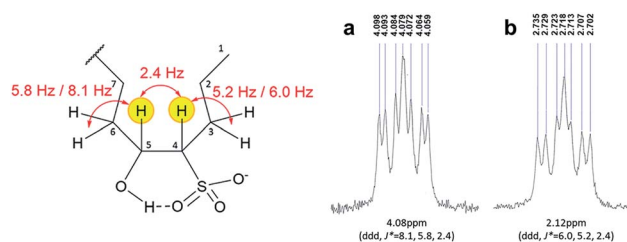


Fig. 3 ¹H NMR spectra showing peaks related to the (a) carbon 4 and (b) carbon 5 protons of C16-4S-5OH. The position numbers are indicated in the molecular structure.



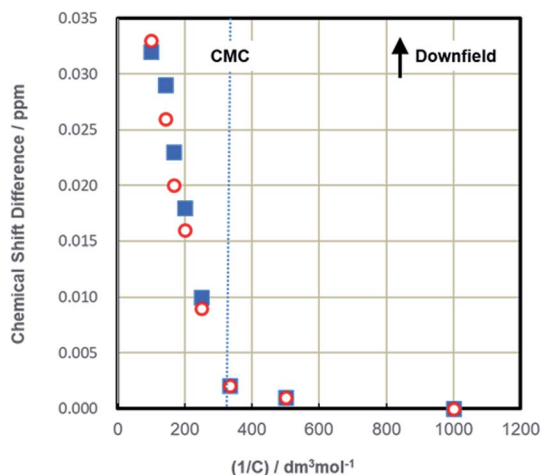


Fig. 4 Difference in the ^1H NMR chemical shifts of the terminal methyl protons of C4 alkyl (○) and C12 alkyl (■) chains of C16-4S-5OH versus the inverse concentration of aqueous C16-4S-5OH solutions at 26 °C.

C16-4S-5OH concentration at values around the CMC. Although both peaks exhibited little change with variations in concentration below the CMC, they suddenly shifted downfield above the CMC, both to the same extent. In the case of DBS, which has two alkyl chains, the variations in this δ difference for the terminal methyl groups below and above the CMC have been studied in detail.^{12,15} Prior work showed that the peak generated by the methyl proton of the longer alkyl chain moved upfield by approximately 0.03 ppm as micelles were formed and aggregated in water. In contrast, the peak related to the shorter alkyl chain moved downfield by about 0.04 ppm. It has been established that the opposite movements of these methyl protons reflect the different positions of the methyl groups in the micelle aggregates. Specifically, the methyl group of the longer alkyl chain is located in the hydrocarbon core while the other is near the hydrophilic groups (that is, the palisade layer). However, the ^1H NMR results in Fig. 4 show that the peaks for both the C12 and C4 chains of C16-4S-5OH moved downfield as micelles formed. This result suggests two possibilities concerning stereostructure in the micelles and the positions of the terminal methyl groups. One explanation is that the highly fixed stereostructures of the monomers below the CMC were maintained when the micelles were formed above the CMC. In this case, the short C4 subchains would function as hydrophobic groups, as was discussed above in relation to CMC values. Another possibility is that the terminal methyl groups continued to experience the deshielding effect of the highly electron-withdrawing sulfo groups, since they were close to these hydrophilic head groups. Downfield shifts of the terminal methyl group peaks were observed in conjunction with the formation of micelles of surfactants having mixed alkyl chain lengths.¹⁵ It has been proposed that this phenomenon occurs when the alkyl chains curl up or bend such that they approach the head groups in the micelles. Although the C16-4S-5OH in the present work was a single compound, the eight carbon difference in the alkyl chain lengths was large enough that these

chains might be expected to bend and loop to fill the spaces in the micelle cores. The δ values for the 4- and 5-protons around the head groups were almost unaffected by increases in the solution concentration. This finding provides support for the hypothesis that the downfield shifts of the peaks for the two terminal methyl protons were caused by their environment in the micelle.

Observation of H/D exchange by FT-IR. The restricted motions of the C16-4S-5OH alkyl chains in water were examined by studying the states of the two hydrophilic head groups. These two functional groups were expected to readily form a six-membered ring containing the 4- and 5-C positions, as a result of intramolecular hydrogen bonding. This structure could potentially affect the overall stereostructure of the C16-4S-5OH and the degree of restriction of the two alkyl chains in water.

H/D exchange experiments were employed to evaluate the presence of hydrogen bonding and its effect on stability in solvents.^{25–27} Initially, the H/D exchange of the hydroxy proton of C16-4S-5OH in D_2O was investigated. Fig. 5(a) presents the FT-IR spectra of solid C16-4S-5OH samples obtained by lyophilizing solutions of C16-4S-5OH in D_2O at various concentrations. The O–H and O–D stretching bands were observed at approximately 3000–3500 and 2500–3000 cm^{-1} , respectively. The strong, sharp peaks in the range of 2800–3000 cm^{-1} were assigned to C–H stretching. The C16-4S-5OH samples obtained from D_2O solutions at 60 °C with concentrations from $1.00 \times 10^{-3} \text{ mol dm}^{-3}$ (below the CMC) to $1.00 \times 10^{-2} \text{ mol dm}^{-3}$ (above the CMC) all generated O–H stretching bands but not O–D stretching bands. The 5-hydroxy proton of the C16-4S-5OH was evidently inactive during the H/D exchange regardless of the solution concentration (and therefore the aggregation state). The formation of a stable hydrogen bond between the hydroxy proton and sulfonyl oxygen could possibly have inhibited the H/D exchange of this hydrogen.

The same H/D exchange experiments were repeated using sodium isethionate, which has a molecular structure that corresponds to the hydrophilic groups of C16-4S-5OH without the hydrophobic alkyl chains (Fig. 5). In these trials, a $1.00 \times 10^{-2} \text{ mol dm}^{-3}$ solution was employed and the FT-IR spectra acquired before and after the D_2O treatment are shown in Fig. 5(b). In these spectra, the O–H stretching band around 3400 cm^{-1} is seen to have decreased in intensity while an O–D stretching band around 2500 cm^{-1} appeared after the D_2O treatment. These results confirmed that an H/D exchange proceeded. The very different behaviors of the C16-4S-5OH and sodium isethionate can likely be ascribed to the presence of long alkyl chains in the former. It was inferred from the ^1H NMR results that the long and short alkyl chains of C16-4S-5OH were highly restricted and also both oriented in the same direction in water. This fixed orientation of the two chains may have promoted the intramolecular cyclization to form a ring comprising the 4-CH, 5-CH, the oxygen and hydrogen of the hydroxy group, and the sulfur and oxygen of the sulfo group (Fig. 5(a)). Conversely, the cyclization of the head groups might have fixed the orientation of the two alkyl chains.

The H/D exchange of C16-4S-5OH in D_2O was inhibited even in the monomer and micelle states. In aqueous solutions below



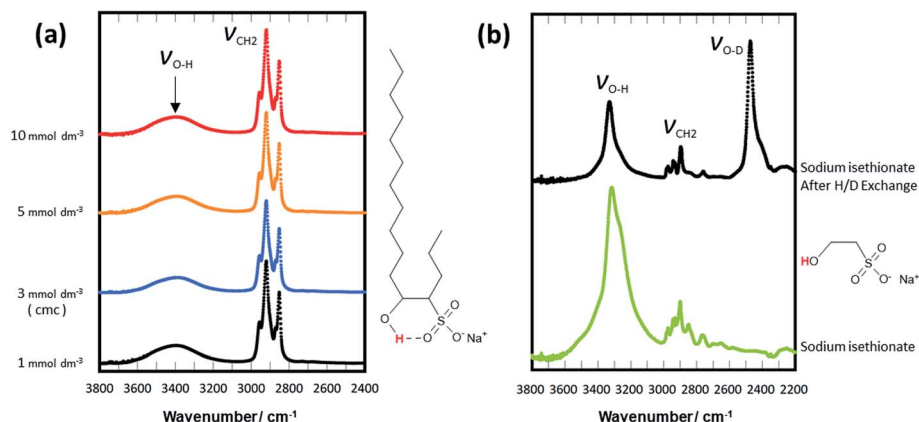


Fig. 5 Evidence for H/D exchange in the FT-IR spectra of C16-4S-5H and sodium isethionate (26 °C). (a) Effect of the concentration of C16-4S-5OH on H/D exchange. (b) Change in the FT-IR spectra of sodium isethionate before and after the H/D exchange treatment.

the CMC, the C16-4S-5OH existed in its monomeric form. Because the distance between monomers was relatively large at these concentrations, intermolecular interactions could not readily proceed. Conversely, there intermolecular interactions occurred easily above the CMC. The minimal H/D exchange at both concentrations strongly suggests that the molecular structure was able to independently stabilize the hydroxy proton by intramolecular cyclization regardless of the aggregation state.

Actual stereostructure of C16-4S-5OH in water

As discussed above, the three aspects of the true molecular structure of C16-4S-5OH in water were clarified. Firstly, the geometries of the C12 and C4 alkyl chains and of the hydrophilic groups were strictly fixed. Secondly, the hydroxy protons were highly stabilized by intramolecular hydrogen bonding. Lastly, the hydroxy protons were stabilized by the two alkyl chains. The actual stereostructure of the two alkyl chains in water was determined to be a *cis* configuration (4*R*, 5*S* or 4*S*, 5*R*) based on the ¹H NMR coupling constants (Fig. 6). Thus, C16-4S-5OH in water appears to consist of a hydrophobic unit containing long and short alkyl chains with highly ordered orientations and a large cyclic hydrophilic head unit containing the sulfo and hydroxy groups.

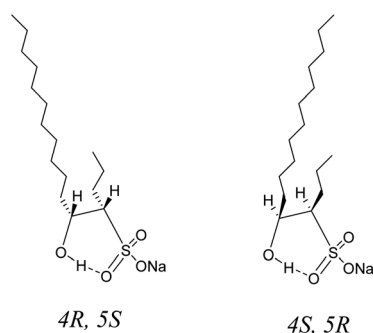


Fig. 6 Possible molecular structures of sodium 5-hydroxyhexadecane-4-sulfonate (C16-4S-5OH) in water.

Two head groups. In the case of a surfactant adsorbed on the air–water interface, the intermolecular distance greatly depends on the size of the hydrophilic groups, including the effects of electrostatic and steric interactions.²² Therefore, it is helpful to discuss the larger A_{\min} value (0.88 nm²) obtained for the C16-4S-5OH (Table 2). Generally, a double alkyl chain surfactant such as DBS possessing only one hydrophilic group will not exhibit restriction of the alkyl chain orientation, and will have an A_{\min} value in the range of 0.45–0.65 nm². These values are almost the same as those of the alkyl sulfates (Table 2) and much smaller than that of C16-4S-5OH.²⁴ Therefore, it appears that the head group of C16-4S-5OH was much larger than a phenyl sulfonate group. This result confirms the successful formation of a large intramolecular cyclic hydrophilic group rather than an intermolecular group, because intermolecular hydrogen bonding would have decreased the intermolecular distance and produced a close packing state.

Dissymmetric alkyl double chains. DBS also possesses an alkyl chain that is divided into a main chain and a subchain by a single hydrophilic sulfo-phenyl group. The two alkyl chains and single hydrophilic group are attached to one chiral methine carbon. The short subchain of DBS is evidently located in a palisade layer near the hydrophilic group away from the main chain based on differences in the chain length, even in the case that the subchain is longer than four carbons.²⁸ This subchain is much different from that of C16-4S-5OH because it is oriented in the same direction as the main chain and also might situate in the micelle core along with the main chain (Fig. 4). Because the saturated carbon atom in DBS has four sigma bonds, the two alkyl chains in DBS each bond to the carbon at an angle of 109.5°, which could hinder the alignment of the two alkyl chains. Conversely, the two alkyl chains of C16-4S-5OH could be oriented close to each other because they are linked to different carbons.

The difference between C16-4S-5OH and a heterogemini surfactant

Gemini surfactants, which comprise two surfactants linked by a spacer unit, have attracted much attention because of their exceptional surface properties, such as low CMCs, good surface

tension reduction characteristics and low Kp values. However, these compounds do exhibit large A_{\min} compared with common surfactants possessing both a single chain and single head group.^{29–31} Those surfactants exhibiting asymmetry in their hydrophilic groups and/or alkyl chains are termed heterogemini surfactants.^{32–36} There are three important factors related to these compounds: the dissymmetry between the two alkyl chain lengths, the spacer chain length, and the type of hydrophilic head group. Although there have been many reports regarding heterogemini surfactants, such prior work has always involved subchains longer than six carbons, possibly because shorter chains may not be hydrophobic.^{10,23} Most previously reported spacer chains were also longer than two carbons because it is challenging to prepare gemini surfactants with shorter spacers. Moreover, such compounds might not be water insoluble because of the strong hydrophobic characteristics of the two alkyl chains in proximity to one another. As an example, even aqueous solutions of gemini surfactants possessing C2 spacers have exhibited very high viscosities at relatively low concentrations.³⁷ There have also been a few reports of heterogemini surfactants, such as C16-4S-5OH, with shorter spacers (C1 or C0).^{38–41} Additionally, surfactants in which both the chain lengths and hydrophilic groups were different have been examined.^{38,39,41}

C16-4S-5OH could be easily regarded as a kind of heterogemini surfactant based on a casual analysis, although its alkyl subchain is too short to be considered as a hydrophobic moiety. Moreover, the spacer is only a C–C bond (that is, C0) while the two head groups are different and form an intramolecular ring structure. In fact, a heterogemini surfactant with the same molecular structure as C16-4S-5OH around the two head groups has been reported.³⁸ However, unlike C16-4S-5OH, this compound possesses two long hydrophobic chains that are believed to be in a *trans* configuration as a result of the steric repulsion between them. These two long chains likely prevent the two head groups from approaching one another to the extent necessary to initiate hydrogen bonding. The A_{\min} value for this surfactant is reported to be above 1.2 nm², which is average for a gemini surfactant but much larger than that of C16-4S-5OH (0.88 nm²). Thus, although the A_{\min} of C16-4S-5OH is greater than that of most common surfactants, it is very compact compared with gemini surfactants. Therefore, the highly fixed molecular structure of C16-4S-5OH could result from both the dissymmetric chains, which do not undergo steric repulsion, and the formation of the intramolecular ring by the two head groups. As such, C16-4S-5OH should not be regarded as a gemini surfactant but rather as an asymmetric double chain surfactant with a large head group.

Origin of the low Kp and CMC of C16-4S-5OH

Low Kp. The two asymmetric alkyl chains of this surfactant naturally affect the packing state between the molecules in water. On the basis of the well-studied phase transition temperatures of phospholipids, increasing the asymmetry of the two alkyl chains would be expected to reduce the chain packing density.⁴² The results of the present study (Table 1)

suggest that the extreme asymmetry of the two alkyl chains could inhibit aggregation, thus reducing Kp. The C16-1S-2OH structural isomer of C16-4S-5OH exhibits a much higher Kp than that of C16-4S-5OH, and the Kp of C12-1S-OH is even higher, although the alkyl chain length is the same as that of the main chain of C16-4S-5OH. The decrease in Kp was therefore caused by the decrease in the packing density of the two asymmetric alkyl chains.

Low CMC. As noted above, C16-4S-5OH can be regarded as a surfactant possessing a C12 alkyl chain as its main hydrophobic group rather than one possessing a C16 hydrophobic chain, because of its low Kp (that is, its higher water solubility). However, surfactants that are highly water-soluble generally have higher CMC values. Therefore, it is of interest to establish the manner in which C16-4S-5OH provides both increased water solubility and a reduced CMC. The results of this study clearly suggest that the C4 alkyl subchain of C16-4S-5OH might act as a hydrophobic moiety, although alkyl chains of less than four carbons would be expected to remain in the palisade layers of micelles in proximity to the hydrophilic head groups. The main driving force for micelle formation by surfactant monomers in water is the entropy gain caused by the release of hydrated water molecules around the hydrophobic chains (meaning the hydrophobic effect). In the case of C16-4S-5OH, both the C12 and C4 alkyl chains could have contributed to the hydrophobic effect. Moreover, the 14 extra carbons could have increased the entropy gains, even though the two carbons to which the two head groups were directly attached were likely situated in the palisade layers of the micelles and did not function as members of the hydrophobic group. The similarity between the CMC values of C16-4S-5H and C14 AS (Table 2) could therefore result from the total number of hydrophobic carbons and not the packing states of the two alkyl chains.

Conclusions

An IOS derived from a C16 oleo resource exhibited both excellent water solubility and a low CMC, which are typically trade-off properties, despite possessing a long hydrophobic chain. Therefore, this compound is expected to have applications as a sustainable anionic surfactant. The C16-4S-5OH IOS synthesized in this work was isolated and purified and its actual stereostructure in water and the effect of its molecular structure on its physicochemical properties were investigated. The highly fixed stereostructure of this molecule was found to be maintained in aqueous solutions both below and above its CMC. The formation of highly oriented and largely hydrophobic asymmetric C12 and C4 chains was confirmed. In addition, the generation of a stable cyclic hydrophilic unit consisting of two carbons of the C16 alkyl chain was shown to occur *via* hydrogen bonding. The actual molecular structure of this surfactant lowered the packing density of the two alkyl chains and reduced its Kp, leading to exceptional water solubility. Conversely, only 16 carbons contributed to the hydrophobic effect during micelle formation, thereby affording a low CMC value. In this manner, the trade-off between water solubility and a low CMC was circumvented. The nature of these IOS characteristics depended



not only on the heterogemini structure of the surfactant but also on the formation of the cyclic hydrophilic unit. The information gained in this work regarding the molecular design of such compounds could contribute to the development of sustainable surfactants.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors did not receive a specific grant from any funding agency in the public, commercial, or not-for-profit sectors. The authors thank the staff of the Kao Corporation who assisted in the development of the biobased IOS.

References

- 1 V. B. Fainerman, D. Mobius and R. Miller, *Surfactants: Chemistry, Interfacial Properties, Applications*, 2001.
- 2 B. V. Vora, G. A. Peterson, S. W. Sohn and M. G. Riley, *Handbook of Detergents Part F: Production*, CRC Press, 2009.
- 3 H. W. Stache, *Anionic Surfactants: Organic Chemistry*, Marcel Dekker, 1996.
- 4 M. E. Tuvel, G. O. Kuehnhan, G. D. Heidebrecht, P. O. Hu and A. D. Zielinski, *J. Am. Oil Chem. Soc.*, 1978, **55**, 70–80.
- 5 J. Stapersma, H. H. Deuling and R. van Ginkel, *J. Am. Oil Chem. Soc.*, 1992, **69**, 39–43.
- 6 S. Takada, WIPO (PCT), WO2011052732A1, 2011.
- 7 H. Hori, Y. Yoshikawa, Y. Mitsuda, T. Fujioka and Y. Nishimoto, WIPO (PCT), WO2015098415A1, 2015.
- 8 Y. Tabuchi, *11th World Surfactant Congress*, Munich, 2019.
- 9 R. Vijay, A. B. Mandal and G. Baskar, *J. Phys. Chem. B*, 2010, **114**, 13691–13702.
- 10 S. Z. Can, D. D. Mago, O. Esenturk and R. A. Walker, *J. Phys. Chem. C*, 2007, **111**, 8739–8748.
- 11 Y. Hou, Y. Han, M. Deng, J. Xiang and Y. Wang, *Langmuir*, 2010, **26**, 28–33.
- 12 S. Das, R. G. Bhurud, N. Nayyar, K. S. Narayan and V. V. Kumar, *J. Phys. Chem.*, 1992, **96**, 7454–7457.
- 13 W. G. Francis, *Chem. Phys. Lipids*, 1981, **29**, 369–374.
- 14 A. Shibahara, K. Yamamoto, T. Nakayama and G. Kajimoto, *J. Jpn. Oil Chem. Soc.*, 1985, **34**, 618–625.
- 15 P. Goon, S. Das, C. J. Clemett, G. J. T. Tiddy and V. V. Kumar, *Langmuir*, 1997, **13**, 5577–5582.
- 16 T. Gu and J. Sjöblom, *Acta Chem. Scand.*, 1991, **45**, 762–765.
- 17 S. H. F. Püschel and B. Strobel, *Tenside*, 1967, **4**, 349–352.
- 18 B. D. Flockhart, *J. Colloid Sci.*, 1961, **16**, 484–492.
- 19 M. Dahanayake, A. W. Cohen and M. J. Rosen, *J. Phys. Chem.*, 1986, **90**, 2413–2418.
- 20 H. Lange and M. J. Schwuger, *Kolloid-Z. Z. Polym.*, 1968, **223**, 145–149.
- 21 I. J. Lin and P. J. Somasundaran, *J. Colloid Interface Sci.*, 1971, **37**, 731–743.
- 22 M. J. Rosen and J. T. Kunjappu, *Surfactants and Interfacial Phenomena*, Wiley, 4th edn, 2012.
- 23 J. A. Mondal, V. Namboodiri, P. Mathi and A. K. Singh, *J. Phys. Chem. Lett.*, 2017, **8**, 1637–1644.
- 24 J. G. Ma, B. J. Boyd and C. J. Drummond, *Langmuir*, 2006, **22**, 8646–8654.
- 25 C. S. Maier and M. L. Deinzer, *Methods Enzymol.*, 2005, **402**, 312–360.
- 26 M. M. G. Krishna, L. Hoang, Y. Lin and S. W. Englander, *Methods*, 2004, **34**, 51–64.
- 27 S. W. Englander, N. W. Downer and H. Teitelbaum, *Annu. Rev. Biochem.*, 1972, **41**, 903–924.
- 28 E. Caponetti, R. Trilo, C. H. O. Patience, J. S. Johnson, L. J. Magid, P. Butler and K. A. Payne, *J. Colloid Interface Sci.*, 1987, **116**, 200–210.
- 29 F. M. Menger and C. Littau, *J. Am. Chem. Soc.*, 1991, **113**, 1451–1452.
- 30 R. Zana, *Adv. Colloid Interface Sci.*, 2002, **97**, 205–253.
- 31 M. J. Rosen, *Chem. Tech.*, 1993, **23**, 30–33.
- 32 R. Oda, I. Huc and S. J. Candau, *Chem. Commun.*, 1997, **21**, 2105–2106.
- 33 G. Bai, J. Wang, Y. Wang, H. Yan and R. K. Thomas, *J. Phys. Chem. B*, 2002, **106**, 6614–6616.
- 34 D. A. Jaeger, B. Li and T. Clark, *Langmuir*, 1996, **12**, 4314–4316.
- 35 E. Alami, K. Holmberg and J. Eastoe, *J. Colloid Interface Sci.*, 2002, **247**, 447–455.
- 36 E. Alami and K. Holmberg, *Adv. Colloid Interface Sci.*, 2003, **100**, 13–46.
- 37 R. Zana, *J. Colloid Interface Sci.*, 2002, **248**, 203–220.
- 38 Y. Takamatsu, N. Iwata, K. Tsubone, K. Torigoe, T. Endo, K. Sakai, H. Sakai and M. Abe, *J. Colloid Interface Sci.*, 2009, **338**, 229–235.
- 39 K. Sakai, Y. Sangawa, Y. Takamatsu and T. Kawai, *J. Oleo Sci.*, 2010, **59**, 541–548.
- 40 K. Sakai, N. Umemoto, W. Matsuda, Y. Takamatsu, M. Matsumoto, H. Sakai and M. Abe, *J. Oleo Sci.*, 2011, **60**, 411–417.
- 41 S. Liu, X. Wang, L. Chen, L. Hou and T. Zhou, *Soft Matter*, 2014, **10**, 9177–9186.
- 42 G. Cevc, *Biochemistry*, 1991, **30**, 7186–7193.