


Cite this: *RSC Adv.*, 2021, 11, 7621

Effects of A-site composition of perovskite ($\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$) oxides on H atom adsorption, migration, and reaction†

Yuta Tanaka,^a Kota Murakami,^a Sae Doi,^a Kazuharu Ito,^a Koki Saegusa,^a Yuta Mizutani,^a Sasuga Hayashi,^a Takuma Higo,^a Hideaki Tsuneki,^a Hiromi Nakai^b and Yasushi Sekine^{a*}

Hydrogen (H) atomic migration over a metal oxide is an important surface process in various catalytic reactions. Control of the interaction between H atoms and the oxide surfaces is therefore important for better catalytic performance. For this investigation, we evaluated the adsorption energies of the H atoms over perovskite-type oxides ($\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$; $0.00 \leq x \leq 0.50$) using DFT (Density Functional Theory) calculations, then clarified the effects of cation-substitution in the A-site of perovskite oxides on H atom adsorption, migration, and reaction. Results indicated local distortion at the oxide surface as a key factor governing H atom adsorption. Subtle Ba^{2+} substitution for Sr^{2+} sites provoked local distortion at the $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ oxide surface, which led to a decrement in the H atom adsorption energy. Furthermore, the effect of $\text{Sr}^{2+}/\text{Ba}^{2+}$ ratio on the H atoms' reactivities was examined experimentally using a catalytic reaction, which was promoted by activated surface H atoms. Results show that the surface H atoms activated by the substitution of Sr^{2+} sites with a small amount of Ba^{2+} ($x = 0.125$) contributed to enhancement of ammonia synthesis rate in an electric field, which showed good agreement with predictions made using DFT calculations.

Received 9th January 2021
Accepted 8th February 2021

DOI: 10.1039/d1ra00180a

rsc.li/rsc-advances

1 Introduction

Migration of hydrogen (H) atoms over a metal oxide surface is a particularly important process in various heterogeneous catalytic reactions including hydrogenation, hydrogenolysis, and ammonia (NH_3) synthesis. In general, H atom diffusion from metal particles to an oxide surface is known as "hydrogen spillover," which is driven by a concentration gradient of H atoms adsorbed onto an oxide surface. The diffusion brings two benefits for active metal supported catalysts: one is the supply of H atoms from an active site, which can dissociate H_2 to an inactive site; the other is the removal of excess strongly adsorbed H atoms over the active metal surface. These effects, which are broadly studied, are applied to various catalytic reactions.^{1–4} In addition, the active migration of H atoms promotes heterogeneous catalytic reactions.^{5–7} In such cases, the migration of H atoms on metal oxide surfaces (proton, H^+) induced by an electric field contributes to cleavage of the rigid bonding in the stable molecules such as $\text{N}\equiv\text{N}$ bonding (for NH_3 synthesis^{8–12}), and C–H bonding (for steam reforming of CH_4 ,^{13–15}

dehydrogenation of methylcyclohexane^{16,17}). H atoms mobility and reactivity play important roles in the catalytic reactions described above. Consequently, controlling the interaction between H atoms and oxide surfaces is necessary to pave the way to a deductive optimization of catalysts.

Perovskite-type oxides (ABO_3) are well known as proton conductors.^{18–24} They are investigated extensively because their properties can be tailored readily^{25–29} using a subtle substitution of A-site cation, which alters the H atom adsorption energy.¹¹ Herein, we investigated the influences of the A-site cation composition on the H atom adsorbed onto $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($0.00 \leq x \leq 0.50$) oxides using density functional theory (DFT) calculations. Moreover, we examined the effects of the H atoms adsorption on catalytic NH_3 synthesis performances in the electric field using $\text{Ru}/\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($0.00 \leq x \leq 0.50$) catalysts. The obtained results indicated that the H atom adsorption was dominated by local distortion at the oxide surface. Furthermore, results show that the trend of the H atom adsorption energies obtained using the theoretical calculations was in close agreement with that of experimentally obtained results on a model reaction of NH_3 synthesis in the electric field.

2 Experimental

2.1. Computational details

We conducted DFT calculations to evaluate the H atom adsorption energies over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25, 0.50$,

^aDepartment of Applied Chemistry, Waseda University, 3-4-1, Okubo, Shinjuku, Tokyo, 169-8555, Japan. E-mail: ysekine@waseda.jp

^bDepartment of Chemistry and Biochemistry, Waseda University, 3-4-1, Okubo, Shinjuku, Tokyo 169-8555, Japan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ra00180a



0.75, 1.00) surfaces. All calculations were conducted using the Vienna *ab initio* simulation package (VASP 5.4.4).^{30–33} A plane-wave basis set with energy cutoff of 400 eV was used; the core–valence interactions were described using the projector augmented wave (PAW) method.³⁴ The generalized gradient approximation (GGA) of the Perdew–Burke–Ernzerhof (PBE) was applied to express the exchange correlation.³⁵ The spin was polarized in all calculations. Electronic occupancies were determined with Gaussian smearing of 0.05 eV. We optimized the slab models with $1 \times 1 \times 1$ *k*-point meshes using the Monkhorst–Pack scheme.³⁶ The convergence threshold of the slab energy was 10^{-5} eV criteria during geometry optimizations. The van der Waals dispersion forces were described using the DFT-D3 method proposed by Grimme.³⁷

We constructed the surfaces of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25, 0.50, 0.75, 1.00$) oxides by the following scheme. First, the bulk models of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$, which had the reported symmetry previously ($x = 0.00$, *Pnma*; $x = 0.25$, *Pnma*; $x = 0.50$, *Imma*; $x = 0.75$, *I4/mcm*; $x = 1.00$, *Pm3m*)³⁸ were prepared. The obtained equilibrium lattice constants were fixed in subsequent calculations of the slab models. Next, each slab model of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ was constructed so that the most stable AO ($\text{Sr}_{1-x}\text{Ba}_x\text{O}$)-terminated surface was exposed. Here, it is noted that the AO-terminated surface is rather more stable than the BO-terminated one,³⁹ and that the exposed surface was the same as the plane from which a maximum diffraction intensity in the X-ray diffraction (XRD) was observed. The slab models consisted of four A–O and B–O layers with a 20 Å vacuum gap separating the adjacent slabs in the *z*-direction. The bottom two layers of the slab were fixed during the optimization of the slab geometry. Finally, a H atom was adsorbed onto the 16 exposed surface oxygen sites. The adsorption energies of the H atom ($E(\text{H adsorption})$) were calculated using eqn (1).

$$E(\text{H adsorption}) = E(\text{surf-H}) - E(\text{surf}) - 0.5E(\text{H}_2) \quad (1)$$

Hereinafter, $E(\text{surf-H})$ and $E(\text{surf})$ respectively stand for the total energies of slab models with and without a H atom. $E(\text{H}_2)$ presents the total energies of gaseous H_2 molecules. Here, the gaseous H_2 molecule was placed in a $10 \times 10 \times 10$ Å cubic box. The adsorption energies were calculated at the gamma point. All calculated values were depicted using VESTA software.⁴⁰

Furthermore, to discuss the effects of the surface oxygen electronic state and the lattice distortion on the H atom adsorption, we performed Bader charge analyses^{41–44} and density of state (DOS) calculations.

2.2. Catalyst preparation

For an experimental confirmation, we prepared catalysts by loading 5 wt% ruthenium (Ru) on $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.000, 0.0625, 0.125, 0.250, 0.375, 0.500$) supports. Each $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ powder was prepared with a complex polymerization method using excessive ethylene glycol, citric acid monohydrate (Kanto Chemical Co. Inc.), and stoichiometric metal precursors $\text{Sr}(\text{NO}_3)_2$, $\text{Ba}(\text{NO}_3)_2$, and $\text{ZrO}(\text{NO}_3)_2 \cdot 2\text{H}_2\text{O}$ (Kanto Chemical Co. Inc.). The obtained polyester resins were calcined twice, at 873 K for 5 h to pyrolyze the rigid polyester network, and at 1373 K for

10 h to decompose the residual carbons completely and to obtain pure oxides in an air atmosphere. Next, Ru was loaded using an impregnation method. As a metal precursor, tris(acetylacetonato)ruthenium(III) ($\text{Ru}(\text{acac})_3$, Tanaka Holdings Co., Ltd.) was used. After $\text{Ru}(\text{acac})_3$ and $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ powders were added to acetone solvents, the slurries were stirred for 2 h. Subsequently, the organic solvents were evaporated at 373 K and dried at 393 K. The obtained powders were reduced at 723 K for 2 h in the stream of $\text{H}_2/\text{Ar} = 1/1$ (total 100 SCCM) to remove ligands of the $\text{Ru}(\text{acac})_3$. For catalytic activity tests, after the prepared samples were pressed into pellets, they were crushed and sieved to obtain grains of 355–500 μm diameter.

2.3. Activity tests

Catalytic activities were evaluated using a fixed-bed, flow-type reactor made of a quartz tube (6.0 mm i.d., 8.0 mm o.d.), as shown in Fig. S1.† Two stainless steel rods (2 mm o.d.) were inserted into the reactor as electrodes. They were attached to the top and bottom of the catalyst bed. To apply the electric field, a direct current (0.0 or 6.0 mA) was imposed using a power supply device. Catalyst bed temperatures and response voltages were monitored respectively by a thermocouple and a digital phosphor oscilloscope (TDS 2001C; Tektronix Inc.). The response voltages were stable during all tests; no plasma formation was observed. Before the activity tests, each 100 mg $\text{Ru}/\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ catalyst was pre-reduced at 723 K for 2 h under a reaction atmosphere. Activity tests were performed in the stream of $\text{N}_2/\text{H}_2 = 1/3$ (0.1 MPa, total 240 SCCM). Under this condition, we confirmed that NH_3 synthesis reaction proceeded in the kinetic region. The produced NH_3 was dissolved in distilled water. Then it was analyzed using an ion chromatograph (IC-2001; Tosoh Co. Inc.).

2.4. Characterization of the catalysts

The crystalline structures of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ supports, fresh and spent 5 wt% $\text{Ru}/\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ catalysts were characterized using powder XRD (SmartLab 3; Rigaku Corp.). These measurements were conducted with filtered Cu-K α radiation operating at 40 mA and 40 kV. Data were recorded from 10° to 90° at a scanning speed of 2° min^{-1} . X-ray fluorescence spectra (XRF, ZSX-Primus II; Rigaku Corp.) were analyzed to verify the Ru loading weight accurately. The particle size distributions of loaded Ru were observed using a field emission transmission electron microscope equipped with an energy dispersive X-ray spectrometer (FE-TEM/EDX, JEM-2100F; JEOL Ltd.). The particle sizes of Ru were found using mean values of about 200 particles treated as spherical objects. Using the obtained Ru particle sizes, the catalytic performances were also evaluated as turn-over frequencies (TOF). The TOF values were calculated following the eqn (2). As previously reported,^{7,11–13} it was elucidated that N_2 reacted with activated protons at the three-phase boundary (TPB) between gas phase, Ru metal, and the oxide support in the electric field. Thus, TOF was calculated by dividing the NH_3 synthesis rates by the number of Ru atoms at TPB; active sites for this reaction.



$$\text{TOF} [\text{h}^{-1}] = \frac{\text{Produced NH}_3 \text{ molecules per unit time} [\text{mol h}^{-1}]}{\text{Number of Ru atoms at TPB} [\text{mol}]} \quad (2)$$

Specific surface areas of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ supports, fresh and spent 5 wt% $\text{Ru}/\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ catalysts were measured using a nitrogen adsorption isotherm at 77 K (Gemini VII; Micromeritics Instrument Corp.). They were ascertained using the Brunauer–Emmett–Teller (BET) method. Before measurements, all samples were pretreated at 473 K for 1 h under He flow to remove adsorbates on the catalyst surfaces. We applied X-ray photoelectron spectroscopy (XPS, Versa Probe II; Ulvac-Phi Inc.) to analyze the amounts of the hydroxy (OH) groups on the $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25, 0.50$) surfaces. These measurements were conducted using Al K α X-ray source. The binding energies were calibrated with the C 1s peak at 284.8 eV.^{24,45–47} Before measurements, the $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ samples were reduced under H_2 atmosphere ($\text{N}_2/\text{H}_2 = 1/3$, total 240 SCCM) at 723 K for 2 h, which was the same procedure as the activity tests. These samples were transferred to the XPS apparatus without an air exposure using a transfer vessel. The obtained C 1s and O 1s peaks were fitted using the Proctor–Sherwood–Shirley method.^{48,49}

3 Results and discussion

3.1. Optimization of computational models

The constructed bulk and slab models of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($0.00 \leq x \leq 1.00$) are presented in Table S1† and are depicted in Fig. S2†. We confirmed that the calculated lattice constants of the optimized bulk models were within 1% error of the measured values obtained by XRD patterns (shown in Fig. S3 and S4†). Regarding the slab models, a discontinuous change between $x \leq 0.50$ and $x \geq 0.75$ was confirmed. Calculations revealed the most stable exposed surface as the (112)-terminated surface for $x = 0.00, 0.25, 0.50$, the (200)-terminated surface for $x = 0.75$, and (101)-terminated surface for $x = 1.00$. The Ba^{2+} -exposed surface tended to be more energetically favourable than the Sr^{2+} -exposed surface for $x > 0.50$. Presumably, this change in the most stable exposed surface was attributable to the difference in the ionic radius between Sr^{2+} and Ba^{2+} . Furthermore, this discontinuous trend was detected in the electronic state of loaded Ru and the basicity of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ supports (see Fig. S5 and S6†). For these reasons, we conducted subsequent investigations using only the samples for $0.00 \leq x \leq 0.50$.

3.2. Effects of perovskite A-site composition on H atom adsorption

The H atom adsorption energies over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ surfaces ($x = 0.00, 0.25, 0.50$) were evaluated using DFT calculations. Fig. 1 exhibits box plots for the calculated H atom adsorption energies on these surfaces. Consequently, the H atom over the surface for $x = 0.25$ exhibited the lowest H atom adsorption energy, signifying a high H^+ -donation ability. Then, we calculated the Bader charges of the surface oxygens and the density of states

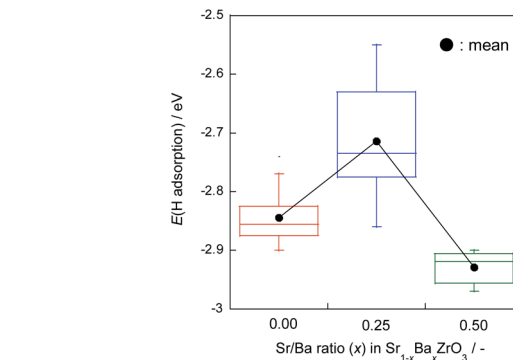


Fig. 1 Box plots for H atom adsorption energies over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25$, and 0.50) oxides. Whisker length is set as 1.5 times of interquartile range. Black plots show mean values.

(DOS) of the slab models to elucidate the effects of $\text{Sr}^{2+}/\text{Ba}^{2+}$ composition ratio on the electronic state of the surface oxygens. The calculated adsorption energies and Bader charges are presented in Table S2† and Fig. 2. In Fig. 2, only the uppermost lattice oxygens are depicted as spheres. Red spheres and blue ones respectively show electron-rich and electron-deficient oxygens. Additionally, the figures in the spheres stand for the H atom adsorption energies. Results show that the $\text{Sr}_{0.75}\text{Ba}_{0.25}\text{ZrO}_3$ possessed evidently electron-rich surface oxygens around Ba^{2+} , and that the H atoms were weakly bound over such electron-rich sites. Reportedly, H atoms tend to join electron-

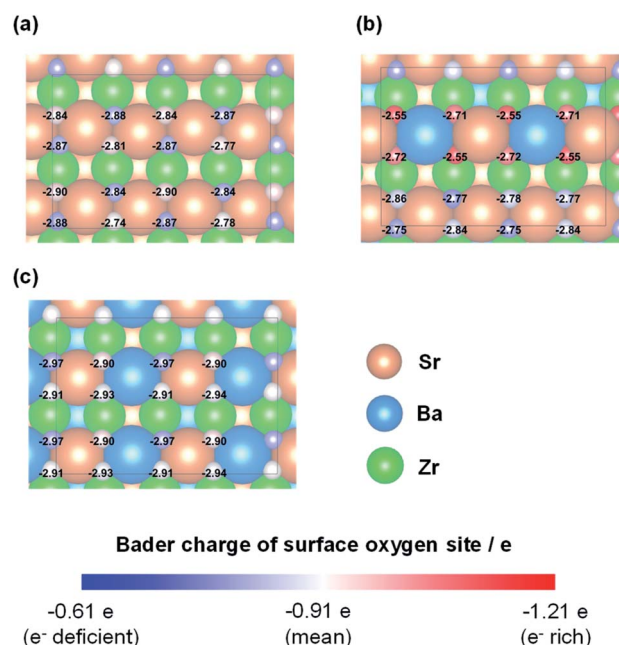


Fig. 2 H atom adsorption energies on each oxygen site at $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ surfaces for (a) $x = 0.00$, (b) $x = 0.25$, and (c) $x = 0.50$. Only the uppermost lattice oxygens are displayed and color-coded following the results of the Bader charge analysis. Red spheres and blue ones show the electron-rich and the electron-deficient oxygen sites, respectively. The figures in the spheres stand for the calculated H atom adsorption energies.

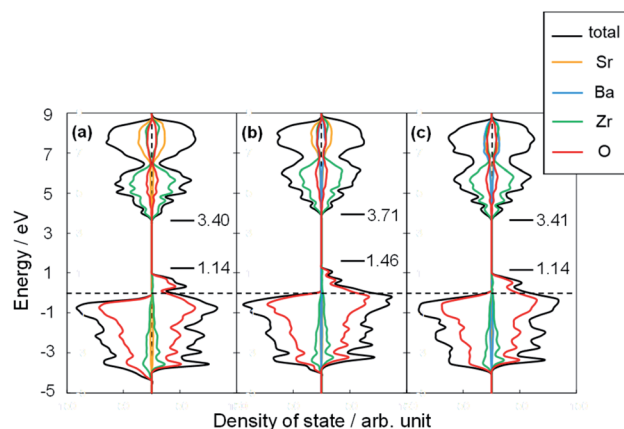


Fig. 3 Total and partial density of state (DOS) shapes of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ slab models for (a) $x = 0.00$, (b) $x = 0.25$, and (c) $x = 0.50$. The total DOS is drawn with a black line, whereas the partial DOS of Sr is drawn with an orange line, Ba with blue, Zr with green, and O with red. In each figure, the two values indicate the energies of the top of the valence band and the bottom of the conduction band.

deficient surface oxygen sites because the electron of the H atom can transfer easily to oxygen sites (known as Lewis acid sites).^{47,50–52} Therefore, we concluded that the formation of the electron-rich surface oxygens by the subtle Ba^{2+} substitution provoked the decrease in the H atom adsorption energy. Furthermore, the obtained DOS shapes (depicted in Fig. 3) show that the position of the valence band (derived from O 2p) and the conduction band (derived from Zr 4d) changed to the higher energy level only for $x = 0.25$, indicating decreased electrophilicity. This change presumably arose from local lattice distortion at the oxide surface induced by the subtle Ba^{2+} addition. Reportedly, lattice distortion in a perovskite-type oxide such as the cation displacements (off-centering), the tilting of BO_6 octahedra, and the elongation and the shortening of some B–O bonds provoke the change in perovskite electronic band structures, especially in d-orbital bands because of its anisotropy (Jahn–Teller effect^{53,54}). Moreover, it has been reported for some perovskite materials that the position of the conduction band level decreases along with the altered hybridization of the electron orbitals when the lattice symmetry is reduced.^{55–57} Therefore, we concluded that the H atom adsorption over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($0.00 \leq x \leq 0.50$) was governed by local lattice distortion, which can be tuned by the A-site $\text{Sr}^{2+}/\text{Ba}^{2+}$ ratio.

3.3. Quantitative analysis for the hydroxy groups (OH)

To evaluate the hydroxy groups (OH) over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ surfaces quantitatively, XPS measurements were conducted. Fig. S7 and S8[†] represent the obtained C 1s and O 1s XPS spectra of $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25, 0.50$) samples. The C 1s peak can be decomposed into three components. The peaks at around 285 eV, 286 eV, and 289 eV were assigned respectively to C–C or C–H, C–O, and O–C=O.^{58,59} The O 1s peak was also decomposed into three components. The peaks at 528 eV, 530 eV, and 532 eV were assigned respectively to lattice oxygens (O_{lat}), O–H or C–O or O–C=O, and physisorbed H_2O .^{46,59–61} Based on these

assignments, the amounts of OH species are calculable following eqn (3) below, as

$$\text{OH ratio} = \frac{((\text{area of "OH"} + \text{"C-O"} + \text{"O-C=O"} \text{ in O 1s spectra}) - \text{RSF}_{\text{O1s}}/\text{RSF}_{\text{C1s}}(\text{area of "C-O"} + \text{"O-C=O"} \text{ in C 1s spectra))}{(\text{area of "O}_{\text{lat}}" \text{ in O 1s spectra})} \quad (3)$$

where RSF_{O1s} and RSF_{C1s} denote the relative sensitivity factor distributed by Ulvac Phi Inc.

The results of quantitative analysis for the OH groups ratio over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.00, 0.25, 0.50$) oxides are shown in Fig. S9.[†] This analysis revealed that the ratio of OH groups over the oxide surface decreased at $x = 0.25$. In other words, the H atoms bonded weakly to the surface oxygen sites for $x = 0.25$. This trend experimentally ensured the results of the H atom adsorption energies predicted by the DFT suggestions.

3.4. Experimental confirmation of the obtained theoretical (DFT) results with activity tests

Based on the DFT suggestions, we presumed that the H atoms over $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ at around $x = 0.25$ exhibited the highest mobility and reactivity. To verify this assumption experimentally, we examined the NH_3 synthesis rate in the electric field as a model reaction because the reaction is controlled by surface H atoms migration. XRD patterns for the synthesized $\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($x = 0.000, 0.063, 0.125, 0.250, 0.375, 0.500$) are shown in Fig. S3.[†] We confirmed that the peak shifted toward the lower angle region with the increase in Ba^{2+} content, meaning the increase in the lattice constants. The calculated lattice parameters changed linearly in accordance with Vegard's law.^{38,62,63} This result demonstrated that Ba^{2+} cation was ideally incorporated into SrZrO_3 at any ratio as in an earlier report, obtained by a synchrotron X-ray diffraction.³⁸ Furthermore, no impure phases were detected under these synthetic conditions.

Previous results verified that activated H atoms over the oxide surface can facilitate N_2 cleavage *via* unstable species (N_2H intermediate in the case of NH_3 synthesis) in the electric field.⁴⁷ In this mechanism, the formation rate of N_2H intermediate determines the overall reaction rate. Therefore, the ability of donating H atoms to N_2 is an important factor for NH_3 synthesis in the electric field. Results of activity tests using $\text{Ru}/\text{Sr}_{1-x}\text{Ba}_x\text{ZrO}_3$ ($0.000 \leq x \leq 0.500$) catalysts are presented in Table 1 and Fig. 4. The NH_3 synthesis rates were evaluated under a low temperature region (about 473–493 K), where N_2 can only dissociate *via* unstable N_2H intermediate. Therefore, the obtained results strongly reflect the H atoms mobility and reactivity. Results demonstrate that the NH_3 synthesis rate achieved the highest value at $x = 0.125$. Here, we confirmed that the catalyst structures were almost unaltered during these catalytic activity tests. In addition, there were no marked differences in the particle size distribution of Ru and the specific surface area among these catalysts, and the TOF values exhibited the same trend as the NH_3 synthesis rates. Moreover, the results detected by an optical measurement and an oxide surface characterization revealed that the influences of $\text{Sr}^{2+}/\text{Ba}^{2+}$ ratio on the Ru electronic state were slight for $0.00 \leq x \leq$



Table 1 Results of activity tests with 5 wt% Ru/Sr_{1-x}Ba_xZrO₃ (0.000 ≤ x ≤ 0.500) catalysts in the electric field (0.1 MPa, 6.0 mA). Catalyst-bed temperatures, response voltages, and NH₃ synthesis rates are the mean values of the multiple measurements. TOFs were calculated using the eqn (2) described in the Section 2.4

Sr/Ba ratio (x) in 5 wt% Ru/Sr _{1-x} Ba _x ZrO ₃	Catalyst-bed temperature (K)	Response voltage (kV)	NH ₃ synthesis rate (μmol g ⁻¹ h ⁻¹)	TOF (h ⁻¹)	Ru particle size (FE-TEM) (nm)	Specific surface area (BET method) (m ² g ⁻¹)
0.000	482.4	0.37	723.3	187	8.79	4.78
0.063	484.0	0.40	1108.9	248	8.17	5.20
0.125	490.0	0.39	1167.3	331	9.20	5.28
0.250	495.2	0.43	1021.5	267	8.83	5.85
0.375	485.1	0.39	799.9	260	9.23	5.62
0.500	479.5	0.39	865.7	219	8.68	6.58

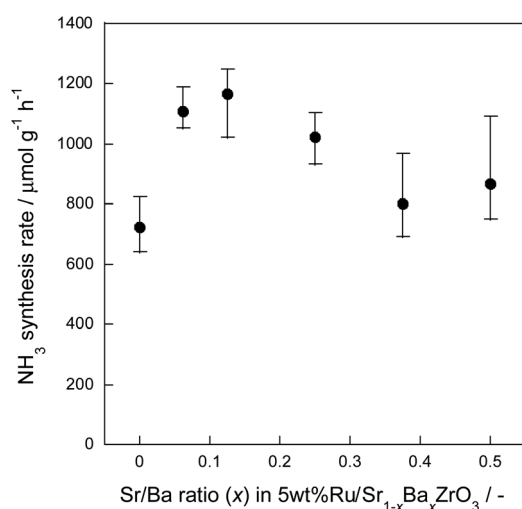


Fig. 4 NH₃ synthesis rates over 5 wt% Ru/Sr_{1-x}Ba_xZrO₃ (0.000 ≤ x ≤ 0.500) in the electric field (about 473 K, 0.1 MPa, 6.0 mA). The error bar indicates the range in the measured values of the multiple measurements.

0.50 (see Fig. S5 and S6[†]). Therefore, the increase in NH₃ synthesis rate at x = 0.125 was attributed not to a change in the Ru electronic state, the Ru size, and the specific surface areas but a change in the H⁺ donating ability derived from the subtle substitution of Sr²⁺ sites by Ba²⁺. In conclusion, it became readily apparent that the evaluations and the predictions of H atoms adsorption by theoretical methods provide clear guidelines for manipulating catalytic reactions involving H atoms over oxide surfaces.

4 Conclusion

We investigated and clarified the effects of the perovskite A-site composition on H atoms adsorption using Sr_{1-x}Ba_xZrO₃ (0.00 ≤ x ≤ 0.50) oxides with DFT calculations. Results revealed that the H atoms adsorbed weakly onto the surface for x = 0.25 because of the presence of electron-rich surface oxygen sites. Local lattice distortion at the oxide surface induced by a subtle Ba²⁺ substitution contributed to this change in the electronic state of surface oxygens.

We verified the roles of H atoms reactivities tuned by Ba²⁺ addition using a model reaction of NH₃ synthesis in the electric field because the reaction is controlled by surface H atoms migration. Consequently, NH₃ synthesis rate reached the maximum value when Sr²⁺ sites were replaced by a small amount of Ba²⁺ (x = 0.125), which agreed well with results obtained using the DFT suggestions. In summary, this investigation demonstrated that theoretical evaluations can predict catalytic performance related to surface H atoms migration over metal oxide surfaces.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research was supported as a JST MIRAI.

Notes and references

- 1 S. Khoobiar, *J. Phys. Chem.*, 1964, **68**, 411–412.
- 2 W. C. Conner and J. L. Falconer, *Chem. Rev.*, 1995, **95**, 759–788.
- 3 R. Prins, *Chem. Rev.*, 2012, **112**, 2714–2738.
- 4 K. Murakami and Y. Sekine, *Phys. Chem. Chem. Phys.*, 2020, **22**, 22852–22863.
- 5 S. Ogo and Y. Sekine, *Chem. Rec.*, 2017, **17**(8), 726–738.
- 6 M. Torimoto, K. Murakami and Y. Sekine, *Bull. Chem. Soc. Jpn.*, 2019, **92**(10), 1785–1792.
- 7 Y. Sekine and R. Manabe, *Faraday Discuss.*, DOI: 10.1039/c9fd00129h, in press.
- 8 R. Manabe, H. Nakatsubo, A. Gondo, K. Murakami, S. Ogo, H. Tsuneki, M. Ikeda, A. Ishikawa, H. Nakai and Y. Sekine, *Chem. Sci.*, 2017, **8**, 5434–5439.
- 9 K. Murakami, R. Manabe, H. Nakatsubo, T. Yabe, S. Ogo and Y. Sekine, *Catal. Today*, 2018, **303**, 271–275.
- 10 K. Murakami, Y. Tanaka, R. Sakai, K. Toko, K. Ito, A. Ishikawa, T. Higo, T. Yabe, S. Ogo, M. Ikeda, H. Tsuneki, H. Nakai and Y. Sekine, *Catal. Today*, 2020, **351**, 119–124.
- 11 K. Murakami, Y. Tanaka, S. Hayashi, R. Sakai, Y. Hisai, Y. Mizutani, A. Ishikawa, T. Higo, S. Ogo, J. G. Seo,



- H. Tsuneki, H. Nakai and Y. Sekine, *J. Chem. Phys.*, 2019, **151**, 064708.
- 12 K. Murakami, Y. Tanaka, R. Sakai, Y. Hisai, S. Hayashi, Y. Mizutani, T. Higo, S. Ogo, J. G. Seo, H. Tsuneki and Y. Sekine, *Chem. Commun.*, 2020, **56**, 3365–3368.
- 13 R. Manabe, S. Okada, R. Inagaki, K. Oshima, S. Ogo and Y. Sekine, *Sci. Rep.*, 2016, **6**, 38007.
- 14 S. Okada, R. Manabe, R. Inagaki, S. Ogo and Y. Sekine, *Catal. Today*, 2018, **307**, 272–276.
- 15 A. Takahashi, R. Inagaki, M. Torimoto, Y. Hisai, T. Matsuda, Q. Ma, J. G. Seo, T. Higo, H. Tsuneki, S. Ogo, T. Norby and Y. Sekine, *RSC Adv.*, 2020, **10**, 14487–14492.
- 16 K. Takise, A. Sato, S. Ogo, J. G. Seo, K. Imagawa, S. Kado and Y. Sekine, *RSC Adv.*, 2019, **9**, 27743–27748.
- 17 M. Kosaka, T. Higo, S. Ogo, J. G. Seo, K. Imagawa, S. Kado and Y. Sekine, *Int. J. Hydrogen Energy*, 2020, **45**(1), 738–743.
- 18 H. Iwahara, H. Uchida, K. Ono and K. Ogaki, *J. Electrochem. Soc.*, 1988, **135**(2), 529–533.
- 19 K. D. Kreuer, *Chem. Mater.*, 1996, **8**(3), 610–641.
- 20 T. Norby, *Solid State Ionics*, 1999, **125**(1), 1–11.
- 21 K. D. Kreuer, *Annu. Rev. Mater. Sci.*, 2003, **33**, 333–359.
- 22 S. Fop, K. S. McCombie, E. J. Wildman, J. M. S. Skakle, J. T. S. Irvine, P. A. Connor, C. Savaniu, C. Ritter and A. C. McLaughlin, *Nat. Mater.*, 2020, **19**(7), 752–757.
- 23 R. Manabe, S. Stub, T. Norby and Y. Sekine, *Solid State Commun.*, 2018, **270**, 45–49.
- 24 Y. Hisai, K. Murakami, Y. Kamite, Q. Ma, E. Völlestad, R. Manabe, T. Matsuda, S. Ogo, T. Norby and Y. Sekine, *Chem. Commun.*, 2020, **56**, 2699–2702.
- 25 J. L. G. Fierro, *Catal. Today*, 1990, **8**, 153–174.
- 26 H. Tanaka and M. Misono, *Curr. Opin. Solid State Mater. Sci.*, 2001, **5**(5), 381–387.
- 27 M. A. Peña and J. L. G. Fierro, *Chem. Rev.*, 2001, **101**(7), 1981–2018.
- 28 N. Labhasetwar, G. Saravanan, S. K. Megarajan, N. Manwar, R. Khobragade, P. Doggali and F. Grasset, *Sci. Technol. Adv. Mater.*, 2015, **16**(3), 036002.
- 29 J. Hwang, R. R. Rao, L. Giordano, Y. Katayama, Y. Yu and Y. Shao-Horn, *Science*, 2017, **358**, 751–756.
- 30 G. Kresse and J. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1993, **47**, 558.
- 31 G. Kresse and J. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **49**, 14251.
- 32 G. Kresse and J. Furthmüller, *Comput. Mater. Sci.*, 1996, **6**(1), 15–50.
- 33 G. Kresse and J. Furthmüller, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169.
- 34 G. Kresse and D. Joubert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758.
- 35 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 36 H. J. Monkhorst and J. D. Pack, *Phys. Rev. B: Solid State*, 1976, **13**, 5188–5192.
- 37 S. Grimme, J. Antony, S. Ehrlich and S. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 38 B. J. Kennedy, C. J. Howard, G. J. Thorogood and J. R. Hester, *J. Solid State Chem.*, 2001, **161**, 106–112.
- 39 R. A. Evarestov, A. V. Bandura and V. E. Alexandrov, *Surf. Sci.*, 2007, **601**, 1844–1856.
- 40 K. Momma and F. Izumi, *J. Appl. Crystallogr.*, 2011, **44**, 1272–1276.
- 41 W. Tang, E. Sanville and G. Henkelman, *J. Phys.: Condens. Matter*, 2009, **21**, 084204.
- 42 E. Sanville, S. D. Kenny, R. Smith and G. Henkelman, *J. Comput. Chem.*, 2007, **28**, 899–908.
- 43 G. Henkelman, A. Arnaldsson and H. Jónsson, *Comput. Mater. Sci.*, 2006, **36**, 354–360.
- 44 M. Yu and D. R. Trinkle, *J. Chem. Phys.*, 2011, **134**, 064111.
- 45 A. Sato, S. Ogo, Y. Takeno, K. Takise, J. G. Seo and Y. Sekine, *ACS Omega*, 2019, **4**, 10438.
- 46 T. Higo, K. Ueno, Y. Omori, H. Tsuchiya, S. Ogo, S. Hirose, H. Mikami and Y. Sekine, *RSC Adv.*, 2019, **9**, 22721.
- 47 K. Murakami, Y. Mizutani, H. Sampei, A. Ishikawa, Y. Tanaka, S. Hayashi, S. Doi, T. Higo, H. Tsuneki, H. Nakai and Y. Sekine, *Phys. Chem. Chem. Phys.*, DOI: 10.1039/d0cp05752e, in press.
- 48 D. A. Shirley, *Phys. Rev. B: Solid State*, 1972, **5**(12), 4709.
- 49 A. Proctor and P. M. A. Sherwood, *Anal. Chem.*, 1982, **54**(1), 13–19.
- 50 Z. Hu, B. Li, X. Y. Sun and H. Metiu, *J. Phys. Chem. C*, 2011, **115**, 3065–3074.
- 51 E. W. McFarland and H. Metiu, *Chem. Rev.*, 2013, **113**, 4391–4427.
- 52 K. Toko, K. Ito, H. Saito, Y. Hosono, K. Murakami, S. Misaki, T. Higo, S. Ogo, H. Tsuneki, S. Maeda, K. Hashimoto, H. Nakai and Y. Sekine, *J. Phys. Chem. C*, 2020, **124**(19), 10462–10469.
- 53 J. A. Alonso, M. J. Martínez-Lope, M. T. Casais and M. T. Fernández-Díaz, *Inorg. Chem.*, 2000, **39**(5), 917–923.
- 54 J. M. Rondinelli and N. A. Spaldin, *Adv. Mater.*, 2011, **23**(30), 3363–3381.
- 55 Q. J. Liu, Z. T. Liu, Y. F. Liu, L. P. Feng, H. Tian and J. G. Ding, *Solid State Commun.*, 2010, **150**, 2032–2035.
- 56 R. Prasanna, A. Gold-Parker, T. Leijtens, B. Conings, A. Babayigit, H. G. Boyen, M. F. Toney and M. D. McGehee, *J. Am. Chem. Soc.*, 2017, **139**(32), 11117–11124.
- 57 C. A. J. Fisher, A. Taguchi, T. Ogawa and A. Kuwabara, *J. Fuel Cell Technol.*, 2020, **19**(3), 95–201.
- 58 M. Mantel and J. P. Wightman, *Surf. Interface Anal.*, 1994, **21**, 595–605.
- 59 E. McCafferty and J. P. Wightman, *Surf. Interface Anal.*, 1998, **26**, 549–564.
- 60 J. T. Newberg, D. E. Starr, S. Yamamoto, S. Kaya, T. Kendelewicz, E. R. Mysak, S. Porsgaard, M. B. Salmeron, G. E. Brown Jr, A. Nilsson and H. Bluhm, *Surf. Sci.*, 2011, **605**, 89–94.
- 61 G. P. López, D. G. Castner and B. D. Ratner, *Surf. Interface Anal.*, 1991, **17**, 267–272.
- 62 L. Vegard and H. Schjelderup, *Phys. Z.*, 1917, **18**, 93.
- 63 A. Yoko, M. Akizuki, N. Umezawa, T. Ohno and Y. Oshima, *RSC Adv.*, 2016, **6**, 67525–67533.

