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Preparation of carbon nanofibrous mats encapsulating zero-valent Fe nanoparticles as Fe reservoir for removal of organic pollutants†

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Zero-valent iron nanoparticles (ZVI NPs) display promising potential in the removal of organic pollutants and heavy metal ions for environmental remediation. However, it is crucial to prevent the oxidation of ZVI NP and control the release of Fe ions under storage and working conditions. In this study, ZVI NPs are encapsulated in single-axial and co-axial carbon nanofibers by electrospinning polyacrylonitrile (PAN)/Fe³⁺ nanofibrous mats with different structures and then annealing the PAN nanofibrous mats in reduction atmosphere. SEM images show that the diameter of the carbon nanofibers is affected by the structure of the nanofibers and the ZVI NPs content after the annealing treatment. The formation of ZVI NPs is confirmed through XPS spectra and HRTEM characterization. The catalytic degradation of organic pollutants by ZVI NPs encapsulated in the carbon nanofibrous mats is evaluated using methylene blue (MB). The results show that the degradation rate of MB is significantly improved when the ZVI NP content encapsulated in the nanofibers increased. MB is completely degraded by the nanofibrous mats with either the single-axial structure or the co-axial structure, but at a higher degradation rate by the single-axial structure than that by the co-axial structure. These results provide alternatives to utilize the carbon nanofibrous mats encapsulating ZVI NPs as Fe reservoir for the removal of organic pollutants in an emergent or long-term situation for environmental remediation.

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1. Introduction

Organic pollutants especially colourful dyes are widely used in the textile, leather, food, and printing industries. Among these dyes, azo dyes are the most diverse and account for approximately 60–70% of synthetic dyes.¹ Azo dyes, which belong to a class of organic compounds with aryl groups attached to both ends of the azo group, are the most widely used organic dyes in textile printing and dyeing processes.² Most organic dyes do not decompose easily in the natural environment. They are transformed into toxic and carcinogenic substances that can cause serious diseases and even induce cancer, which seriously endangers human life.^{3–5} Considering the harm of azo dyes, various methods, including chemical degradation,^{6–8} biodegradation,^{9,10} chemical oxidation,¹¹ and physical adsorption,^{12,13} are

being used to remove and degrade azo dyes in water. In the traditional Fenton reaction, Fe²⁺ or Fe³⁺ reacts with H₂O₂ to produce 'OH, which destroys azo dyes by utilizing its strong oxidation ability.¹⁴ Fenton reaction equipment is simple and experimental conditions are mild, but its pH value is strictly regulated,¹⁵ and the uncontrollable release of Fe ions leads to uncontrollable catalytic reactions. Considering the limitations of Fenton reaction, zero-valent iron nanoparticles (ZVI NPs) were used to replace Fe salts to initiate a Fenton-like reaction.^{16,17} Compared with the Fenton reaction, the Fenton-like reaction is relatively loose for pH and can catalyse H₂O₂ under neutral conditions.¹⁸ However, ZVI NPs can easily be oxidized, leading to the decrease of catalytic action. In sewage treatment, ZVI NPs are not suitable for storage and transportation, resulting in increased costs and reduced catalytic efficiency. Therefore, controlling the release of Fe ions and preventing the oxidation of ZVI NP are extremely important. Several strategies have been developed to stabilize the structure of ZVI NPs.^{19–22}

Electrospinning is a method used to fabricate nanofibrous mats by spraying and stretching polymer solution or melting under electrostatic force.²³ This method can continuously fabricate nanofibrous mats, given its high porosity and large specific surface area.^{24,25} Specifically, co-axial electrospinning is often used to fabricate core-sheath structured nanofibers,²⁶ in which the sheath layer and the core layer are different from each

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other. Therefore, composite nanofibers with different functions can be fabricated by combining the advantages of sheath layer and core layer spinning solutions.^{27,28}

In this work, polyacrylonitrile (PAN) nanofibrous mats encapsulating Fe(acac)₃ with single-axial and co-axial structure were fabricated by electrospinning. After annealing under reduction condition, ZVI NPs were formed and encapsulated in the carbon nanofibrous mats with single-axial and co-axial structure. The morphology and elemental status were characterized through SEM, HRTEM, and XPS. Furthermore, the catalytic degradation of organic dyes by as-prepared nanofibrous mats in the presence of H₂O₂ was evaluated using azo dye methylene blue (MB) as the model. Results showed that as-prepared carbon nanofibrous mats with either single-axial or co-axial structure encapsulating ZVI NPs displayed high degradation efficiency in MB depending on the content of ZVI NPs encapsulated in the nanofibers. In comparison, carbon nanofibrous mats with ZVI NPs encapsulated in the core of the nanofibers displayed a slower degradation rate than those with single-axial structure, implying that it could be used as Fe reservoir for long-term degradation of organic contaminants.

2. Experimental

2.1. Chemical reagents

Polyacrylonitrile (PAN, (C₃H₃N)_n, $M_w = 15\,000\text{ g mol}^{-1}$, CAS number of 25014-41-9), was purchased from Macklin Biochemical Co., Ltd. (Shanghai, China). Fe(acac)₃ (60% purify) was obtained from Sigma-Aldrich. Hydrogen peroxide (H₂O₂), methylene blue (MB), *N,N*-dimethylformamide (DMF), and ethanol were purchased from Sinopharm Chemical Reagent Co., Ltd. (Shanghai, China).

2.2. Preparation of ZVI NPs-encapsulated carbon nanofibrous mats

Single-axial and co-axial electrospinning was carried out using an electrospinning machine (SS-2535H, Ucalery, China). In the case of single-axial electrospinning, PAN and Fe(acac)₃ were dissolved in DMF at different ratios under magnetically stirring for 12 hours. The solution was then carefully loaded into a 5 mL syringe for electrospinning. In the case of co-axial electrospinning, PAN was dissolved in DMF at a designed concentration under magnetically stirring for 12 hours, which was used as the sheath electrospinning solution; Fe(acac)₃ at a designed concentration was dissolved in DMF under magnetically stirring for 12 hours, which was used as core electrospinning solution. The parameters of the electrospinning processes are listed in Table 1.

The as-prepared PAN mats were pre-oxidized at 250 °C for 2.5 h in an electrothermal blower dryer, then were annealed in a tubular furnace at 550 °C for 3 h under reduction atmosphere (5% of H₂ and 95% of N₂). After annealing treatment, the nanofibrous mats were re-named as S1-a, S2-a, S3-a, C1-a, C2-a, C3-a, and S0-a, corresponding to S1, S2, S3, C1, C2, C3, and S0, respectively.

Table 1 The parameters of the electrospinning process and the ratio of electrospinning solution

No. ^a	Content of PAN	Content of Fe(acac) ₃	Working voltage	Injecting speed
S1	12 wt%	2 wt%	+8.0 to -2.5 kV	0.05 mm min ⁻¹
S2	12 wt%	3 wt%		
S3	12 wt%	4 wt%		
C1	12 wt%	2 wt%		0.0125/0.05 mm min ⁻¹
C2	12 wt%	3 wt%		
C3	12 wt%	4 wt%		
S0	12 wt%	0 wt%		0.05 mm min ⁻¹

^a The sample was re-named as S1-a, S2-a, S3-a, C1-a, C2-a, C3-a, and S0-a, corresponding to S1, S2, S3, C1, C2, C3, and S0, respectively, after annealing treatment.

2.3. Characterizations of as-prepared nanofibrous mats

The phase and surface information of nanofibrous mats before and after annealing treatment was characterized through XRD and FT-IR. The surface topography and structure of the samples were observed using field-emission scanning electron microscope (FESEM, S4800, Hitachi, Japan) and high-resolution transmission electron microscope (HRTEM, Hitachi, Japan). X-ray photoelectron spectroscopy (XPS) characterization was performed on an ESCALAB250Xi (Thermo Fisher Scientific, USA) spectrometer using Al K α radiation as the excitation source. The absorbance of MB solution was detected by enzyme labelling instrument (Thermo Fisher Scientific, USA).

2.4. Catalytic degradation of MB by ZVI NPs-encapsulated carbon nanofibrous mats

In a typical experiment, the carbon nanofibrous mat (10 mg) and H₂O₂ (1 mL) was added into 20 mL of MB (10 mg L⁻¹) aqueous solution. The suspension was stored in a constant temperature shock box at 37 °C. The supernatant was collected and monitored at designated time intervals after centrifugation (8000 rpm, 5 min).

3. Results and discussion

3.1. Morphological characterization

Fig. 1 shows SEM images of as-prepared nanofibrous mats before and after annealing treatment. The nanofibers in mats with either single-axial or co-axial structure were randomly oriented and uniform in diameter. No apparent bead strings or fracture marks were observed on the surface of as-prepared nanofibrous mats. The surfaces of the nanofibers became smoother after annealing treatment. The average diameter of the nanofibers was greatly affected by annealing treatment. The average diameter of the nanofibers before annealing was approximate 890.6 ± 98.4 , 478 ± 37.3 , 390.5 ± 26.5 , 650.8 ± 26.8 , 642.4 ± 49.6 , 578.8 ± 74.8 , 379.6 ± 40.4 nm for S1, S2, S3, C1, C2, C3, and S0, which decreased to 321.3 ± 43.4 , 375.8 ± 48.8 , 316.8 ± 48.2 , 360.1 ± 31.6 , 450.0 ± 49.1 , 492.6 ± 49.4 , and 263.6 ± 17.4 nm for S1-a, S2-a, S3-a, C1-a, C2-a, C3-a, and S0-a,



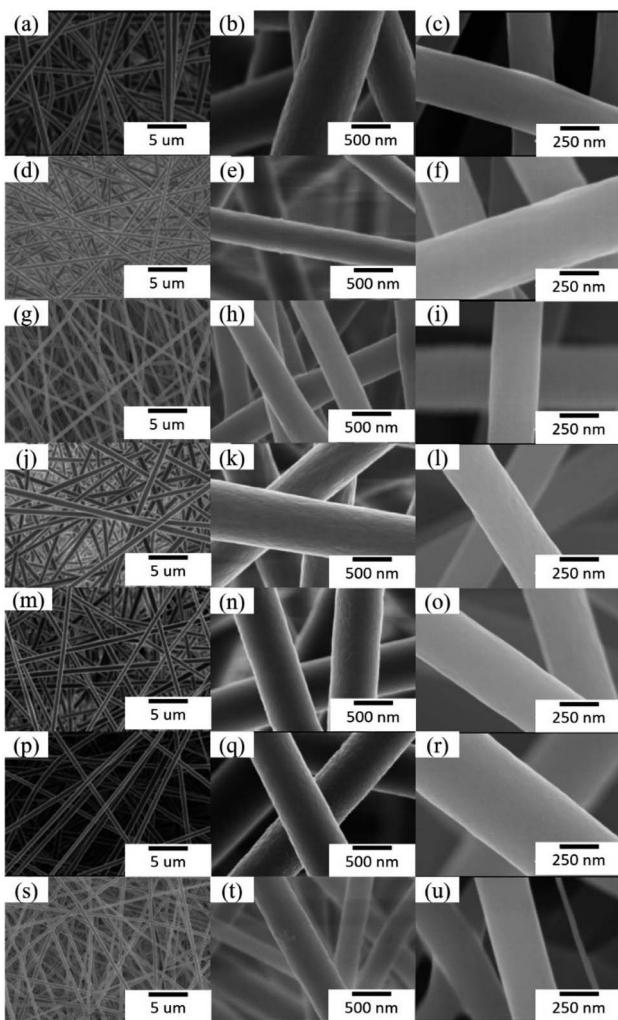


Fig. 1 SEM images of as-prepared nanofibrous mats before (left two lanes) and after annealing treatment (right lane). (a and b) S1, (c) S1-a, (d and e) S2, (f) S2-a, (g and h) S3, (i) S3-a, (j and k) C1, (l) C1-a, (m and n) C2, (o) C2-a, (p and q) C3, (r) C3-a, (s and t) S0, (u) S0-a.

respectively, after annealing treatment. The concentration of Fe salt in the nanofibers also affected the diameter of the carbon nanofibers. Before annealing treatment, the average diameter of the nanofibers with single-axial structure significantly decreased with increasing the concentration of such Fe salt, whereas that of the nanofibers with co-axial structure only slightly decreased. After annealing treatment, however, the average diameter of the nanofibers with single-axial structure further decreased, and such decrement was negatively related to the concentration of Fe salt encapsulated in the nanofibers. In comparison, the average diameter of the nanofibers with co-axial structure significantly increased with increasing such Fe salt concentration after the annealing treatment. The decrease of the average diameter of the nanofibers could be attributed to the further volatilization of residual solvents in the nanofibers and the carbonization of PAN molecules upon annealing treatment. The influence on the average diameter of the nanofibers induced by the concentration of Fe salt also provides an

indirect evidence for the formation of ZVI NPs. In the case of the nanofibers with single-axial structure, ZVI NPs were uniformly formed after reduction annealing and encapsulated in the carbon nanofibers. Therefore, the carbon nanofibers containing ZVI NPs displayed similar diameters, although the amount of ZVI NPs in the carbon nanofiber was different. In the case of the nanofibers with co-axial structure, ZVI NPs were primarily formed in the core layer of the carbon nanofibers reduction after annealing treatment. The higher concentration of Fe salt encapsulated in the core layer of PAN nanofibers may have resulted in the formation of ZVI NPs with a larger size.

3.2. Phase and surface structure

Fig. 2 shows the XRD patterns and FT-IR spectra of the nanofibrous mats after annealing treatment. The broad diffraction peaks at 2θ of approximately 25° were attributed to the (002) plane of amorphous carbon.^{29,30} No evident diffraction peaks assigned to either iron oxides or ZVI NPs were observed in the XRD patterns of as-prepared samples, possibly attributing to the low content and extremely small size of ZVI NPs. The surface information was recorded by FT-IR spectra. Compared with the FT-IR spectrum of S0-a, the peaks near 1584 and 1306 cm^{-1} were shifted to higher wavenumbers of 1595 and 1385 cm^{-1} , respectively, in FT-IR spectra of ZVI NPs-encapsulated nanofibers. These shifts could be attributed to the coordination between iron atoms and the residual carboxyl groups contained in the polymers after annealing. It is noteworthy that there is an obvious peak near 3400 cm^{-1} for the S1-a sample compared with the other samples, which could be derived from the hydroxy group of H_2O .

3.3. Characterization of ZVI NPs

The formation of ZVI NPs was further confirmed by HRTEM characterization. As shown in Fig. 3, numerous ZVI NPs were observed on the surface of the nanofibers. The average diameter of ZVI NPs was in the range of $5\text{--}10\text{ nm}$. The well aligned lattices indicate high crystallization of ZVI NPs. The average interplanar distance between adjacent lattices was approximately $0.203 \pm 0.010\text{ nm}$, which could be indexed to the (110) planes of Fe (JCPDS card no. 06-0696). The amount of ZVI NPs formed on the surface of the nanofibers increased with the enhancement of the Fe salt concentration in the nanofibers. It is interesting that

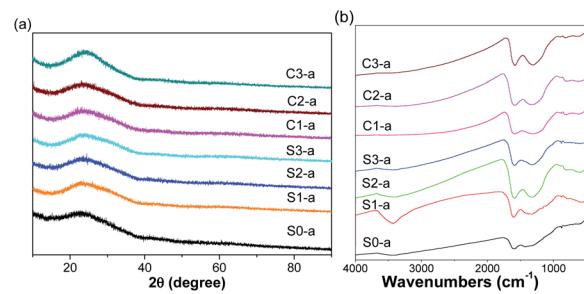


Fig. 2 XRD patterns (a) and FT-IR spectra (b) of the as-prepared nanofibrous mats after annealing.



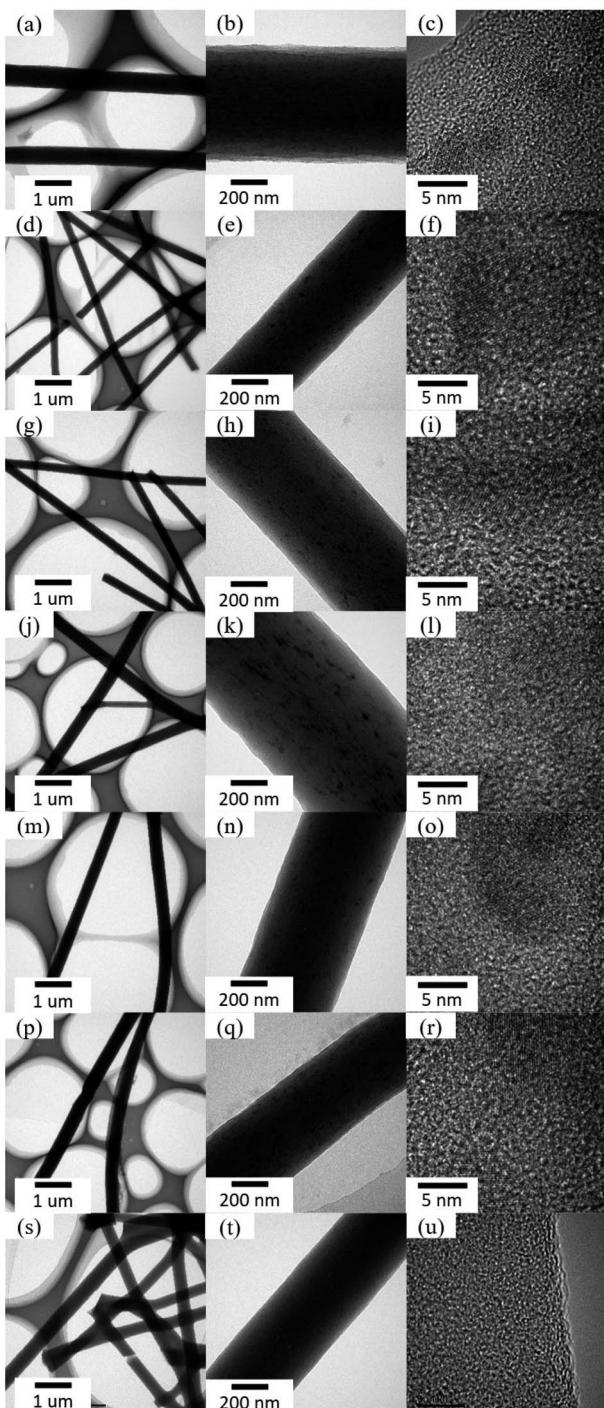


Fig. 3 TEM and HRTEM images of as-prepared nanofibrous mats after annealing treatment. (a–c) S1-a, (d–f) S2-a, (g–i) S3-a, (j–l) C1-a, (m–o) C2-a, (p–r) C3-a, and (s–u) S0-a.

ZVI NPs also appeared on the surface of the nanofibers with coaxial structure. We speculate that a portion of Fe ions seeped from the core layer to the sheath layer during annealing treatment and formed ZVI NPs under reduction condition. The formation of ZVI NPs in both the core layer and the sheath layer resulted in the larger diameter of the nanofibers with co-axial structure than that of the nanofibers with single-axial structure.

To further identify the status of ZVI NPs encapsulated in the nanofibers with different structures, XPS characterization was carried out. Fig. S1† shows the XPS spectra of as-prepared nanofibrous mats after annealing treatment. The peaks that appeared at approximate 284.5, 399, 531, and 710 eV were indexed to C 1s, N 1s, 1s, and Fe 2p, respectively. The existence of N and O could be attributed to the incomplete reduction of PAN under the annealing treatment. No signal indexed to Fe element was detected in XPS spectrum of S0-a. Fig. 4 shows the XPS spectra of Fe 2p in the nanofibers with different structures after annealing treatment. Two major peaks at approximate 709–711.3 and 723–725.6 eV could be indexed to $\text{Fe} 2\text{p}_{3/2}$ and $\text{Fe} 2\text{p}_{1/2}$ spin-orbit coupling, respectively. The peak of $\text{Fe} 2\text{p}_{1/2}$ and $\text{Fe} 2\text{p}_{3/2}$ was at approximately 724.8 and 711 eV for Fe^{3+} and approximately 709.5 and 723.2 eV for Fe^{2+} , respectively.³⁰ Hence, $\text{Fe}^{3+}/\text{Fe}^{2+}$ existed in the nanofibers, possibly attributing to the incomplete reduction of Fe salt encapsulated in the nanofibers. The characteristic peak for Fe^0 at approximately 705 eV was observed in the XPS spectra of all the samples, confirming the formation and encapsulation of ZVI NPs in the nanofibers.

3.4. Catalytic degradation of MB

To evaluate the degradation of MB treated with as-prepared nanofibrous mats in the presence of H_2O_2 , we monitored the absorbance of MB solution at different time intervals. pH value of solution was 7.27, 3.75, and 3.78, corresponding to MB solution, MB solution in the presence of ZVI NPs-encapsulated carbon mats and H_2O_2 at the time of 0 and 6 h, respectively. As

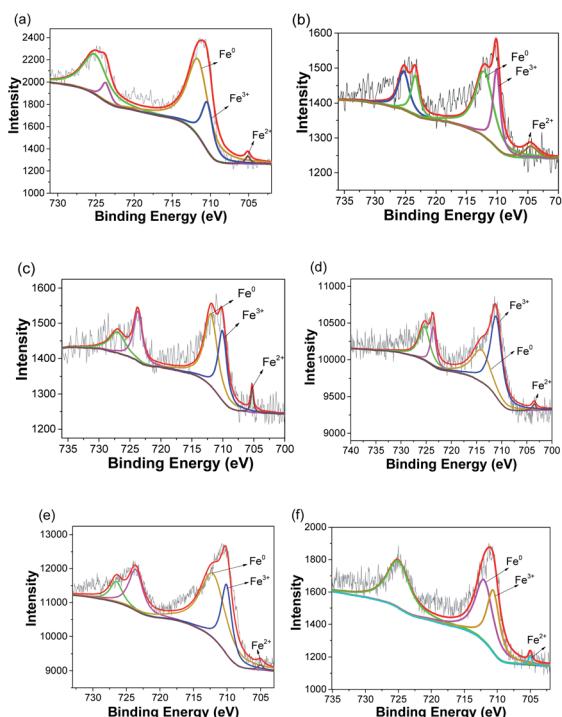


Fig. 4 XPS spectra of Fe 2p in as-prepared nanofibrous mats after annealing treatment. (a) S1-a, (b) S2-a, (c) S3-a, (d) C1-a, (e) C2-a, (f) C3-a.

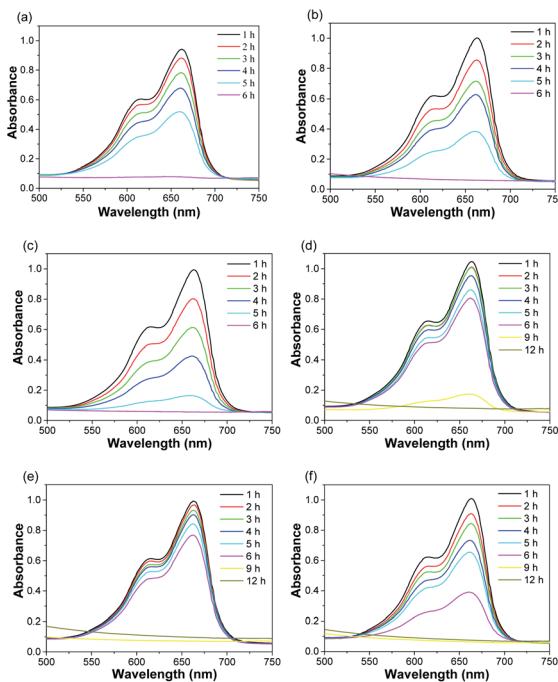


Fig. 5 UV-Vis spectra of MB solution treated with as-prepared nanofibrous mats in the presence of H_2O_2 at different time intervals. (a) S1-a, (b) S2-a, (c) S3-a, (d) C1-a, (e) C2-a, (f) C3-a.

shown in Fig. 5a-c, MB was completely decomposed by the nanofibrous mats with single-axial structure in the presence of H_2O_2 at the end of 6 h. Noticeably, the degradation rate of MB was positively related to the content of ZVI NPs encapsulated in the nanofibers. The higher the content of ZVI NPs encapsulated in the nanofibrous mat, the higher the degradation efficiency of MB. Similar results were observed in the case of treating MB solution with co-axial nanofibrous mats. At the initial 6 h, the higher content of ZVI NPs displayed a higher degradation efficiency of MB, although the entire MB in solution was completely decomposed at the end of 12 h. Thus, the degradation efficiency of MB was lower in the case of ZVI NPs encapsulated in the nanofibers with co-axial structure than that in the nanofibers with single-axial structure. Reasonably, the sheath of co-axial structure prevented the direct and fast release of Fe ions from ZVI NPs, which were encapsulated in the core of nanofibers. Therefore, ZVI NPs encapsulated in the core of the nanofibers with co-axial structure could serve as Fe reservoir for a long-term release of Fe ions.

To explore the degradation mechanism of MB under treatment with as-prepared nanofibrous mats and H_2O_2 , we used a pseudo-first-order kinetic model to evaluate the degradation kinetic of MB, as expressed below:^{31,32}

$$\ln(A_0/A_t) = kt \quad (1)$$

where A_0 represents the initial absorbance value, A_t represents the absorbance value at a certain time, k represents the apparent kinetic constant of the pseudo-first-order reaction model, and t represents the time. Considering t as abscissa and

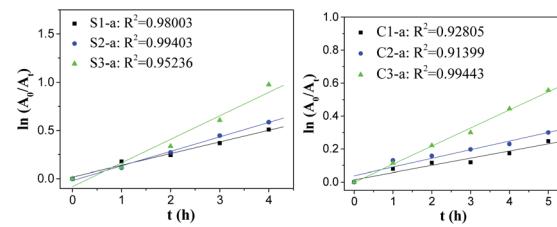
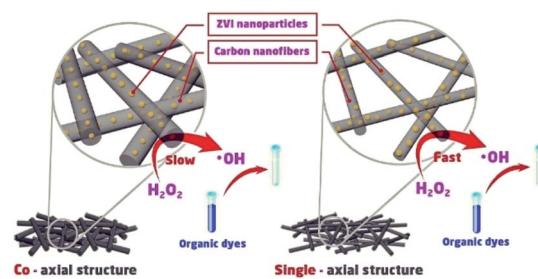


Fig. 6 Kinetic curves of degradation reactions catalysed by as-prepared nanofibrous mats with different structures.



Scheme 1 The speculated degradation mechanism of MB catalysed by as-prepared carbon nanofibrous mats with either the single-axial or co-axial structure.

$\ln(A_0/A_t)$ as ordinate, kinetic curves were generated for the first group, and the slope indicates the apparent kinetic constant. Fig. 6 shows that dye degradation reactions by the nanofibrous mats with single-axial structure have a linear relationship between $\ln(A_0/A_t)$ and t , and the correlation coefficient R^2 is close to 1, implying that dye degradation reactions catalysed by the nanofibrous mats with single-axial structure agreed with the first-order kinetic characteristics. Compared with the slope, the reaction rate constant increased as the Fe content increased, indicating that the ZVI NP content had a significant impact on dye degradation. In the case of degradation reactions catalysed by the nanofibrous mats with co-axial structure, the curves only fitted the first-order kinetic model at the initial 5 h. Scheme 1 shows the speculated catalytic degradation of MB treated by as-prepared nanofibrous mats with either single-axial or co-axial structure. MB in the solution was decomposed by hydroxyl radicals generated by the ZVI NPs-initiated Fenton-like reaction in the presence of H_2O_2 . With the consumption of ZVI NPs dispersed in the shell layer of the nanofibers, iron ions could be continuously released for ZVI NPs encapsulated in the core layer of the nanofibers due to the permeation of the acid H_2O_2 solution, which ensures the complete degradation of MB in the solution.

4. Conclusion

In summary, a series of PAN/Fe³⁺ nanofibrous mats with single-axial and co-axial structures were successfully prepared by electrospinning. ZVI NPs were formed by annealing the PAN nanofibrous mats in a reduction atmosphere and encapsulated in either single-axial or co-axial carbon nanofibers. The results



on the catalytic degradation of organic dye show that the degradation rate significantly improved as the ZVI NP content encapsulated in the nanofibers increased. Moreover, the nanofibrous mats with single-axial structure displayed a higher degradation rate than those with co-axial structure, although the organic dye was completely degraded by the nanofibrous mats with either single axial structure or co-axial structure. ZVI NPs encapsulated in the core layer of the nanofibers can avoid burst release of $\text{Fe}^{2+}/\text{Fe}^{3+}$ during the practical application, thereby providing an alternative as Fe reservoir for long-term degradation of organic contaminants. Therefore, these results provide alternatives for carbon nanofibrous mats encapsulating ZVI NPs for the removal of organic dyes or heavy metal ions in environmental remediation. In a short-term and emergent situation, carbon nanofibrous mats with single-axial structure can catalytically degrade organic dyes quickly, whereas in a long-term and protective situation, carbon nanofibrous mats with co-axial structure can be utilized as the Fe reservoir for controlled degradation of organic dyes.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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